Mercury (Hg) is one of the major toxic heavy metals, harmful to the environment and human health. Thus, it is significantly important to find an easy and quick method to detect Hg$^{2+}$. In this study, we designed and synthesized a simple fluorescent probe with excellent properties, such as high sensitivity and selectivity, rapid response, and outstanding water solubility. When Hg$^{2+}$ (5 μM) was added to the probe solution, it exhibited a very large fluorescent enhancement (about 350-fold stronger than the free probe) with the help of hydrogen peroxide (H$_2$O$_2$). Probe HCDC could quantitatively detect Hg$^{2+}$ in the range of 0–10 μM using the fluorescence spectroscopy method and the detection limit was measured to be about 0.3 nM (based on a 3σ/slope). Analytical application was also studied, and the probe HCDC exhibited excellent response to Hg$^{2+}$ with the addition of H$_2$O$_2$ in real water samples. So, our proposed probe HCDC provided a practical and promising method for determining Hg$^{2+}$ in the environment.

1. Introduction

Heavy metal ion pollution has become a global environmental problem. As a common metal element, mercury (Hg) is extremely toxic and shows a great threat to the environment and human health due to its bioaccumulation, duration, and easy absorption by aquatic organisms. Once the mercury enters the organism, it will lead to the dysfunctions of the central nervous systems, the brain, and the kidneys. As we all know, the maximum contaminant level of mercury in drinking water is 2 ppb (0.01 mg/L) which was set by the Environmental Protection Agency (EPA). It can be seen that mercury could pose a serious threat to humans at very low concentrations. Thus, developing a convenient, rapid, highly selective and sensitive method for the determination of Hg$^{2+}$ is vitally important.

Nowadays, several classical methods have been applied to determine Hg$^{2+}$ including electrochemical analysis, atomic absorption/emission spectroscopy, electrochemical analysis, and inductively coupled plasma mass spectrometry. Although these methods are selective and accurate to detect ions, most of them require long time, expensive instruments, and tedious sample pre-treatment. Thus, the techniques of fluorescent probes were widely studied due to the advantages of highly selective, operational simplicity, and relatively low-cost. Many fluorescent probes for detecting Hg$^{2+}$ have been reported; however, most of them were still limited by the disadvantages of low sensitivity, bad water solubility, and long response time. So, a simple water-soluble fluorescent probe for determining Hg$^{2+}$ with high sensitivity and quick response became our primary task.

Very recently, we have reported several carbonothioate-based fluorescent probe for detecting Hg$^{2+}$. These probes respectively chose the 2-(2'-hydroxyphenyl)benzothiazole (HBT), dichlororesorufin, and seminaphthorhodaflool as the fluorophores, and a carbonothioate moiety was used as the recognition receptor. All these showed quick response and high selectivity for detecting Hg$^{2+}$. In connection with our continuing research, we designed a new probe HCDC (Scheme 1) for detecting Hg$^{2+}$ with superior properties in this paper, and the property of probe HCDC was compared with other reported studies in Table 1. A dimethylthiocarbamic ester group and 7-hydroxycoumarin were respectively chosen as the recognition receptor and the fluorophore. We assumed that the thioester group changed into a fluorophore, and a carbonothioate moiety was used as the recognition receptor. Thus, the probe showed quick response and high selectivity for detecting Hg$^{2+}$. When Hg$^{2+}$ was added to the probe solution, a fluorescent enhancement (about 350-fold stronger than the free probe) was observed (Figure 1). The probe HCDC could quantitatively detect Hg$^{2+}$ in the range of 0–10 μM using the fluorescence spectroscopy method and the detection limit was calculated to be about 0.3 nM (based on a 3σ/slope). Analytical application was also studied, and the probe HCDC exhibited excellent response to Hg$^{2+}$ with the addition of H$_2$O$_2$ in real water samples. So, our proposed probe HCDC provided a practical and promising method for determining Hg$^{2+}$ in the environment.
2. Experimental

2.1 Materials and general methods

All the chemicals were gained from commercial suppliers and were used without further purification. The Sartorius Arium 611DI system can produce the ultra-pure water which was used during the entire process of the experiment. Column chromatography using Silica gel 200–300 mesh from the Qingdao Haiyang Chemical Co was used to purify the reaction product. The $^1$H NMR and $^{13}$C NMR were all recorded on the Bruker AV-400 spectrometer with chemical shifts reported as ppm (in DMSO- $d_{6}$, TMS as internal standard). Fluorescence spectra were gained from a Horiba FluoroMax-4 spectrophotometer. All the fluorescence spectra were uncorrected. Probe HCDC was prepared according to the previous paper.42

2.2 General procedure for analysis

Stock solutions of probe HCDC (1 mM) were prepared in ethanol. And the parent stock solutions (1 mM) of Hg$^{2+}$, other

<table>
<thead>
<tr>
<th>Probe</th>
<th>LOD (nM)</th>
<th>Solution (v/v)</th>
<th>Time (min)</th>
<th>References</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>1.8</td>
<td>Ethanol : PBS = 2 : 1</td>
<td>30</td>
<td>34</td>
</tr>
<tr>
<td></td>
<td>10</td>
<td>$C_2H_5OH : water = 1 : 1$</td>
<td>40</td>
<td>35</td>
</tr>
<tr>
<td></td>
<td>34</td>
<td>$CH_3CN : water = 1 : 1$</td>
<td>—</td>
<td>36</td>
</tr>
<tr>
<td></td>
<td>18.8</td>
<td>$C_2H_5OH : water = 2 : 8$</td>
<td>20</td>
<td>37</td>
</tr>
<tr>
<td></td>
<td>100</td>
<td>DMSO : $H_2O = 9 : 1$</td>
<td>30</td>
<td>38</td>
</tr>
<tr>
<td></td>
<td>0.3</td>
<td>Aqueous solution</td>
<td>10</td>
<td>This work</td>
</tr>
</tbody>
</table>
metal ions and common anions species (10 mM) were prepared in ultrapure water. The test solution contained 50 µL of the probe stock solution and moderate amount of each stock solution, and eventually diluted to 10 mL in a test tube with 100% ultrapure water containing HEPES (5 mM, pH 7.4).

Before Hg^{2+} and other ions were added to the test solution, H_{2}O_{2} (3 mM) was added. All the tests were conducted at room temperature (25 °C).

2.3 Determination of the detection limit

According to the previous paper,^{43,44} the method of fluorescence titration was used to calculate the detection limit. The fluorescence emission spectrum of probe HCDC was measured by fifteen times in the absence of Hg^{2+} in order to achieve the standard deviation of blank measurement. To gain the slope, the fluorescence intensity at 455 nm was plotted as a concentration of Hg^{2+}. Therefore, the formula for calculating the detection limit is described as following:

\[
\text{Detection limit} = \frac{3\sigma}{k}
\]

where \(\sigma\) is the standard deviation of the five spectrums among the fifteen spectrums which peaks were nearest, \(k\) is the slope between the Hg^{2+} concentration versus the fluorescence intensity.

2.4 Analytical application

Three water samples were chosen to test the analytical application of the probe HCDC. One water sample was collected from Jia Zi Lake, University of Jinan, China, and the other two samples were respectively collected from the Xiu Jiang River and Yellow River, at Jinan, China. All the samples needed to filter through filter paper before the measurement.

3. Results and discussion

3.1 Characteristic spectra

In this paper, the fluorescence and absorption spectra of probe HCDC was shown as Fig. 1. In the absence of Hg^{2+} and H_{2}O_{2}, the probe solution exhibits a very weak fluorescence emission peak at 455 nm. When only Hg^{2+} was added to the probe solution, the reaction was very slow. And, the reaction was not carried out when only the H_{2}O_{2} was added. Once Hg^{2+} and

![Scheme 2](image)

Scheme 2 Reaction Mechanism of probe HCDC for Hg^{2+} with the help of H_{2}O_{2}.
H₂O₂ were all added to the probe HCDC solution, the maximum fluorescence emission peak exhibits a very large enhancement at 455 nm (Fig. 1a). The fluorescence intensity peak of the probe solution containing Hg²⁺ and H₂O₂ was about 350-fold stronger than that of the blank probe solution. Meanwhile, in the absorption spectra, with the addition of Hg²⁺ and H₂O₂, the probe solution exhibited a new absorption peak at 364 nm (Fig. 1b). The result showed that Hg²⁺ could generate the cleavage of an eater group in the presence of H₂O₂ (Scheme 2).

3.2 Effect of H₂O₂

Different concentrations of H₂O₂ were respectively added to the solution of probe HCDC (5 μM) and Hg²⁺ (5 μM) in the aqueous solution which contained HEPES (5 mM, pH 7.4). It was shown in inset of Fig. 2, the fluorescence intensity was increased first and then decreased with the increasing concentrations of H₂O₂ (the range of 1 to 40 mM). The enhancement of fluorescence reached the maximum when the concentration of H₂O₂ was 3 mM. The results implied that the low concentration of H₂O₂ (1 to 3 mM) could promote the cleavage of ester group and amplify the response of probe HCDC to Hg²⁺. But when the concentration is over 3 mM (4 to 40 mM), H₂O₂ will inhibit the response of probe HCDC to Hg²⁺. Thus, we chose 3 mM H₂O₂ as the amplification reagent in the detection of Hg²⁺.

3.3 Time dependence of detecting Hg²⁺

Response time is a significant element to test the effect of the probe. So, the response time of probe HCDC was investigated under the analytical conditions which were mentioned above. The result was shown in Fig. 3. When only Hg²⁺ was added to H₂O₂ were all added to the probe HCDC (5 μM) in the presence of Hg²⁺ (5 μM) and different concentrations of H₂O₂ (final concentration: 0, 1, 2, 3, 4, 5, 10, 15, 20, 25, 30, 35, 40 mM) under HEPES (5 mM, pH 7.4) aqueous solution. Excitation wavelength = 330 nm. Error bar = RSD (n = 3).

![Fig. 2](image-url)

**Fig. 2** Fluorescence spectra of probe HCDC (5 μM) in the presence of Hg²⁺ (5 μM) and different concentrations of H₂O₂ (final concentration: 0, 1, 2, 3, 4, 5, 10, 15, 20, 25, 30, 35, 40 mM) under HEPES (5 mM, pH 7.4) aqueous solution. Excitation wavelength = 330 nm. Error bar = RSD (n = 3).

![Fig. 3](image-url)

**Fig. 3** Reaction time for probe HCDC (5 μM) with the addition of only Hg²⁺ (5 μM) (a), and Hg²⁺ (5 μM) and H₂O₂ (3 mM) (b).
the probe solution, the reaction was very slow. While, the fluorescence intensity at 455 nm was increased as soon as Hg²⁺ and H₂O₂ were added, and increased gradually until the reaction time exceeded 10 min. Thus, all the measurements of the spectrum were carried out 10 minutes after the addition of Hg²⁺. The results showed that the probe HCDC could provide a rapid analytical method for detecting Hg²⁺.

3.4 Quantification of Hg²⁺
As shown in Fig. 4, the continuous enhancement of fluorescence intensity at 455 nm can be seen with the continuous addition of Hg²⁺ (the range of 0 to 10 μM). And, there was a good linearity (y = 225439x + 3899.2, R² = 0.998) between the Hg²⁺ concentrations of 0–2 μM and the fluorescence intensity with a detection limit of 0.3 nM (based on 3σ/slope). The detection limit of the probe HCDC was far lower than the standard of Hg²⁺ (10 nM) in drinking water which set up by the Environmental Protection Agency (EPA). The above results showed that the probe HCDC could detect Hg²⁺ quantitatively by the fluorescence spectrometry method with excellent sensitivity.

3.5 Selectivity to Hg²⁺
High selectivity is also an important factor for an excellent probe. The selectivity of HCDC toward Hg²⁺ was evaluated under the above-mentioned conditions. The effect of the metal ions (Mg²⁺, Zn²⁺, Ni²⁺, Ca²⁺, Na⁺, K⁺, Pb²⁺, Cd²⁺, Cr⁶⁺, Fe²⁺, Fe³⁺, Sn⁴⁺) and the common anions species (NO₃⁻, NO₂⁻, HCO₃⁻, H₂PO₄⁻, F⁻, CO₃²⁻, Cl⁻, SO₄²⁻) on fluorescence spectra of probe HCDC were investigated. Firstly, Hg²⁺, other metal ions and common anions species were respectively added to the solution, and only Hg²⁺ resulted in a considerable fluorescence enhancement (Fig. 5a and c). The results exhibited that HCDC possesses high selectivity toward Hg²⁺.

The interference of the metal ions and common anions species on the detection of Hg²⁺ was also studied. Even though the mental ions and common anions species were respectively existed in the solution, the probe HCDC also had a good response to Hg²⁺ (Fig. 5b and d). These studies clearly indicated that the probe HCDC could be used for the selective detection of Hg²⁺ without interference from anions.

3.6 Analytical application
Based on the above results, the probe HCDC owned the excellent properties containing a very low detection limit, quicker response, excellent water solubility, selectivity and anti-interference. So, we attempted to investigate the practical application of the probe for the selective detection of Hg²⁺ in three real water samples. The results were shown in Table 2. At first, there were no Hg²⁺ in three water samples, and then 2 μM, 5 μM, or 10 μM Hg²⁺ was respectively added to the solution of real water (5 mM, HEPES, pH 7.4) containing 5 μM probe HCDC and 3 mM H₂O₂. Every sample was repeatedly configured and tested for 3 times. Therefore, we obtained a good method for detecting Hg²⁺ in the environment.
4. Conclusion

In a word, we have designed and synthesized a highly sensitive and selective probe for detecting Hg$^{2+}$ by the fluorescence spectrum. The enhancement of fluorescence intensity of the probe solution containing Hg$^{2+}$ and H$_2$O$_2$ was very significant, which could reach 350-fold stronger than that of the blank probe solution. There was a good linear relationship in the range of 0–2 mM Hg$^{2+}$, and the low detection limit was 0.3 nM (on the basis of 3σ/slope). The probe showed excellent solubility in water. More importantly, the response of probe HCDC to Hg$^{2+}$ in real water samples was also satisfactory. All these remarkable advantages of this new probe suggest that this probe will have excellent application prospect in the detection of mercury.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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References


Table 2  Analysis results of Hg$^{2+}$ in three real water samples $^a$

<table>
<thead>
<tr>
<th>Real water sample</th>
<th>Found Hg$^{2+}$ (μM)</th>
<th>Found Hg$^{2+}$ (μM)</th>
<th>Recovery (%)</th>
<th>RSD (n = 3) (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Sample A</td>
<td>2</td>
<td>1.82 ± 0.06</td>
<td>91.06</td>
<td>2.89</td>
</tr>
<tr>
<td></td>
<td>5</td>
<td>4.52 ± 0.11</td>
<td>90.47</td>
<td>2.10</td>
</tr>
<tr>
<td></td>
<td>10</td>
<td>9.61 ± 0.40</td>
<td>96.06</td>
<td>4.02</td>
</tr>
<tr>
<td>Sample B</td>
<td>2</td>
<td>1.85 ± 0.11</td>
<td>92.52</td>
<td>5.36</td>
</tr>
<tr>
<td></td>
<td>5</td>
<td>4.97 ± 0.29</td>
<td>99.33</td>
<td>5.83</td>
</tr>
<tr>
<td></td>
<td>10</td>
<td>10.11 ± 0.25</td>
<td>101.11</td>
<td>2.51</td>
</tr>
<tr>
<td>Sample C</td>
<td>2</td>
<td>2.24 ± 0.13</td>
<td>111.93</td>
<td>6.56</td>
</tr>
<tr>
<td></td>
<td>5</td>
<td>5.15 ± 0.21</td>
<td>102.96</td>
<td>4.19</td>
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<tr>
<td></td>
<td>10</td>
<td>9.73 ± 0.37</td>
<td>97.28</td>
<td>3.72</td>
</tr>
</tbody>
</table>

$^a$ Sample A, lake water from JiaZi Lake, University of Jinan; sample B, from the Xiu Jiang River, at Jinan, China; sample C, from the Yellow River, at Jinan, China.
18 L. Yang, Y. Su, Y. Geng, H. Xiong, J. Han, Q. Fang and X. Song, Inorg. Chem., 2018, 57, 5036–5042.