



Cite this: *Chem. Sci.*, 2018, 9, 1604

Regioselective Simmons–Smith-type cyclopropanations of polyalkenes enabled by transition metal catalysis†

Jacob Werth and Christopher Uyeda *

A [i -PrPDI]CoBr₂ complex (PDI = pyridine-diimine) catalyzes Simmons–Smith-type reductive cyclopropanation reactions using CH₂Br₂ in combination with Zn. In contrast to its non-catalytic variant, the cobalt-catalyzed cyclopropanation is capable of discriminating between alkenes of similar electronic properties based on their substitution patterns: monosubstituted > 1,1-disubstituted > (*Z*)-1,2-disubstituted > (*E*)-1,2-disubstituted > trisubstituted. This property enables synthetically useful yields to be achieved for the monocyclopropanation of polyalkene substrates, including terpene derivatives and conjugated 1,3-dienes. Mechanistic studies implicate a carbenoid species containing both Co and Zn as the catalytically relevant methylene transfer agent.

Received 10th November 2017
Accepted 23rd December 2017

DOI: 10.1039/c7sc04861k

rsc.li/chemical-science

Introduction

Cyclopropanes are common structural elements in synthetic and natural biologically active compounds.¹ The Simmons–Smith cyclopropanation reaction was first reported over half a century ago but remains today one of the most useful methods for converting an alkene into a cyclopropane.² As compared to diazomethane, which is shock sensitive and must be prepared from complex precursors, CH₂I₂ is both stable and readily available, making it an attractive methylene source. Additionally, the stereospecificity of the Simmons–Smith reaction allows diastereomeric relationships in cyclopropanes to be established with a high degree of predictability. Several advances have addressed many of the limitations of the initial Simmons–Smith protocol. For example, Et₂Zn can be used in the place of Zn to more reliably and quantitatively generate the active carbenoid reagent.³ Acidic additives, such as CF₃CO₂H⁴ and substituted phenols,⁵ have been found to accelerate the cyclopropanation of challenging substrates. Finally, Zn carbenoids bearing dialkylphosphate anions⁶ or bipyridine ligands⁷ are sufficiently stable to be stored in solution at low temperatures (Fig. 1).

Despite the many notable contributions in Zn carbenoid chemistry, a persistent limitation of Simmons–Smith-type cyclopropanations is their poor selectivity when attempting to discriminate between multiple alkenes of similar electronic properties. For example, the terpene natural product limonene

possesses a 1,1-disubstituted and a trisubstituted alkene. Friedrich reported that, under a variety of Zn carbenoid conditions, the two alkenes are cyclopropanated with similar rates, resulting in mixtures of monocyclopropanated (up to a 5 : 1 ratio of regioisomers) and dicyclopropanated products.⁸ This issue is exacerbated by the challenge associated with separating the two monocyclopropane regioisomers, which only differ in the position of a non-polar CH₂ group. In general, synthetically useful regioselectivities in Simmons–Smith reactions are only observed for substrates containing directing groups.⁹

In principle, catalysis may provide an avenue to address selectivity challenges in Simmons–Smith-type cyclopropanations; however, unlike diazoalkane transfer reactions,

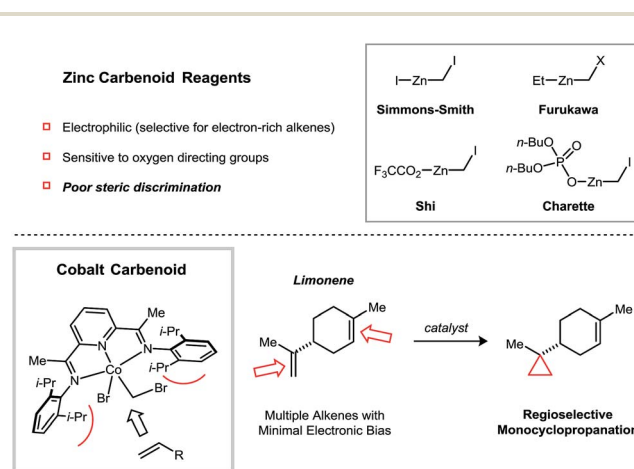


Fig. 1 Factors governing alkene selectivity in Zn carbenoid-mediated cyclopropanation reactions. Cobalt-catalyzed reductive cyclopropanation exhibiting high regioselectivities in polyalkene substrates based on alkene substitution patterns.

Department of Chemistry, Purdue University, West Lafayette, IN 47907, USA. E-mail: cuyeda@purdue.edu

† Electronic supplementary information (ESI) available: Experimental procedures and characterization data. CCDC 1584851. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7sc04861k



which are catalyzed by a broad range of transition metal complexes,^{9b,10} there has been comparatively little progress toward the development of robust catalytic strategies for reductive cyclopropanations.¹¹ Lewis acids in substoichiometric loadings have been observed to accelerate the Simmons–Smith reaction, but in many cases, this rate effect is restricted to allylic alcohol substrates.^{12,13}

Recently, our group described an alternative approach to catalyzing reductive cyclopropanation reactions using a transition metal complex that is capable of activating the dihaloalkane reagent by C–X oxidative addition. A dinickel catalyst was shown to promote methylene¹⁴ and vinylidene¹⁵ transfer using CH₂Cl₂ and 1,1-dichloroalkenes in combination with Zn as a stoichiometric reductant. Here, we describe a mononuclear [PDI]Co (PDI = pyridine-diimine) catalyst¹⁶ that imparts a high degree of steric selectivity in the cyclopropanation of polyalkene substrates. Mechanistic studies suggest that the key intermediate responsible for methylene transfer is a heterobimetallic conjugate of Co and Zn.

Results and discussion

4-Vinyl-1-cyclohexene contains a terminal and an internal alkene of minimal electronic differentiation and thus provided a suitable model substrate to initiate our studies (Table 1).^{8,17} Under standard CH₂I₂/Et₂Zn conditions (entry 1), there is a modest preference for cyclopropanation of the more electron-rich disubstituted alkene (rr = 1 : 6.7) with increasing amounts of competing dicyclopropanation being observed at higher conversions (entries 2 and 3). Other modifications to the conditions, including the use of a Brønsted acid⁴ (entry 4) or a Lewis acid additive^{12b,18} (entry 5), did not yield any improvements in selectivity. Likewise, an Al carbenoid generated using CH₂I₂ and AlEt₃ afforded a similar preference for cyclopropanation of the endocyclic alkene (entry 6).¹⁹

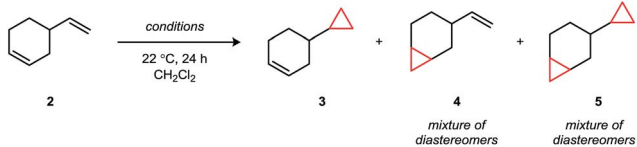
In a survey of transition metal catalysts, the [ⁱ-PrPDI]CoBr₂ complex **1** was identified as a highly regioselective catalyst for

the cyclopropanation of 4-vinyl-1-cyclohexene, targeting the less hindered terminal alkene (Table 2). CH₂Br₂ and Zn alone do not afford any background levels of cyclopropanation (entry 1); however, the addition of 6 mol% [ⁱ-PrPDI]CoBr₂ (**1**) provided monocyclopropane **3** (81% yield) with a >50 : 1 rr and <1% of the dicyclopropane product (entry 5). The steric profile of the catalyst appears to be critically important for yield. For example, the mesityl- (entry 6) and phenyl-substituted variants (entry 7) of the ligand provided only 58% and 4% yield respectively under the same reaction conditions. Related N-donor ligands similarly afforded low levels of conversion (entries 8–12) as did the use of other first-row transition metals, including Fe (entry 14) and Ni (entry 15), in the place of Co.

In order to define the selectivity properties of catalyst **1**, we next conducted competition experiments using alkenes bearing different patterns of substitution (Fig. 2). Reactions were carried out using an equimolar amount of each alkene and run to full conversion of the limiting CH₂Br₂ reagent (1.0 equiv.). Mono-substituted alkenes are the most reactive class of substrates using **1** but are not adequately differentiated from 1,1-disubstituted alkenes (3 : 1). By contrast, terminal alkenes are significantly more reactive than internal alkenes, providing synthetically useful selectivities (≥31 : 1). Furthermore, a model *Z*-alkene was cyclopropanated in preference to its *E*-alkene congener in a 33 : 1 ratio. Using catalyst **1**, trisubstituted alkenes are poorly reactive, and no conversion is observed for tetrasubstituted alkenes.

The synthetic applications of the catalytic regioselective cyclopropanation were examined using the terpene natural products and derivatives shown in Fig. 3. In all cases, the selectivity properties follow the trends established in the competition experiments. Substrates containing ether or free alcohol functionalities (*e.g.*, **7**, **10**, and **11**) exhibit a strong directing group effect under classical Simmons–Smith conditions; however, catalyst **1** overrides this preference and targets the less hindered alkene. Additionally, the presence of electron-deficient α,β-unsaturated carbonyl systems (*e.g.*, **9**, **13**, and **14**) do not perturb the expected steric selectivity.

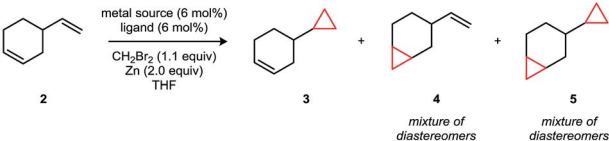
Table 1 Regioselectivity studies using Zn and Al carbenoid reagents^a



| Entry | Reaction conditions | Yield (3 + 4) | rr (3 : 4) | Yield 5 |
|-------|---|---------------|------------|---------|
| 1 | CH ₂ I ₂ (1.0 equiv.), Et ₂ Zn (0.5 equiv.) | 28% | 1 : 6.7 | 3% |
| 2 | CH ₂ I ₂ (1.0 equiv.), Et ₂ Zn (1.0 equiv.) | 33% | 1 : 4.6 | 5% |
| 3 | CH ₂ I ₂ (2.0 equiv.), Et ₂ Zn (2.0 equiv.) | 53% | 1 : 6.5 | 16% |
| 4 | CH ₂ I ₂ (1.0 equiv.), Et ₂ Zn (1.0 equiv.), 3,5-difluorobenzoic acid (2.0 equiv.) | 28% | 1 : 3.5 | 19% |
| 5 | CH ₂ I ₂ (2.0 equiv.), Et ₂ Zn (2.0 equiv.), TiCl ₄ (0.2 equiv.) | 13% | 1 : 4.6 | 1% |
| 6 | CH ₂ I ₂ (1.2 equiv.), AlEt ₃ (1.2 equiv.) | 38% | 1 : 3.1 | 9% |

^a Reaction conditions: 4-vinylcyclohexene (0.14 mmol), CH₂Cl₂ (1.0 mL), 24 h, 22 °C. Yields and ratios of regioisomers were determined by GC analysis against an internal standard.



Table 2 Catalyst structure–activity relationship studies^a


^{i-Pr}PDI: Ar = 2,6-(*i*-Pr)₂C₆H₃
^{Me}PDI: Ar = 2,4,6-(Me)₃C₆H₂
^{Ph}PDI: Ar = C₆H₅
^{i-Pr}DAD: Ar = 2,6-(*i*-Pr)₂C₆H₃
^{i-Pr}IP: Ar = 2,6-(*i*-Pr)₂C₆H₃
 terpy
^{t-Bu}PyBOX: R = *t*-Bu
^{Ph}PyBOX: R = C₆H₅

| Entry | Metal source | Ligand | Yield (3 + 4) | rr (3 : 4) | Yield 5 |
|-------|------------------------|----------------------------|---------------|------------|---------|
| 1 | — | — | <1% | — | <1% |
| 2 | CoBr ₂ | — | <1% | — | <1% |
| 3 | Co(DME)Br ₂ | — | <1% | — | <1% |
| 4 | — | ^{i-Pr} PDI | <1% | — | <1% |
| 5 | CoBr ₂ | ^{i-Pr} PDI | 81% | >50 : 1 | <1% |
| 6 | CoBr ₂ | ^{Me} PDI | 58% | >50 : 1 | <1% |
| 7 | CoBr ₂ | ^{Ph} PDI | 4% | — | <1% |
| 8 | CoBr ₂ | ^{i-Pr} DAD | <1% | — | <1% |
| 9 | CoBr ₂ | ^{i-Pr} IP | 2% | — | <1% |
| 10 | CoBr ₂ | Terpy | 4% | — | <1% |
| 11 | CoBr ₂ | ^{t-Bu} PyBOX | <1% | — | <1% |
| 12 | CoBr ₂ | ^{Ph} PyBOX | <1% | — | <1% |
| 13 | CoBr ₂ | PPh ₃ (12 mol%) | <1% | — | 0% |
| 14 | FeBr ₂ | ^{i-Pr} PDI | 3% | — | <1% |
| 15 | NiBr ₂ | ^{i-Pr} PDI | <1% | — | <1% |

^a Reaction conditions: 4-vinylcyclohexene (0.14 mmol), THF (1.0 mL), 24 h, 22 °C. Yields and ratios of regioisomers were determined by GC analysis against an internal standard.

Vinylcyclopropanes are a valuable class of synthetic intermediates that engage in catalytic strain-induced ring-opening reactions.²⁰ The monocyclopropanation of a diene represents an attractive approach to their synthesis but would require a catalyst that is capable of imparting a high degree of regioselectivity and avoiding secondary additions to form dicyclopropane products.²¹ These challenges are addressed for a variety of diene classes using catalyst **1** (Fig. 4). Over the substrates that we have examined, the selectivities for cyclopropanation of the terminal over the internal double bond of the diene system are uniformly high. Additionally, the catalyst is tolerant of vinyl bromide (**15**) and vinyl boronate (**23**) functional groups, which are commonly used in cross-coupling reactions.

Like the non-catalytic Simmons–Smith reaction,^{2c} the cyclopropanation using **1** is stereospecific within the limit of detection, implying a mechanism in which the two C–C σ -bonds are either formed in a concerted fashion or by a stepwise process that does not allow for single bond rotation. For example,

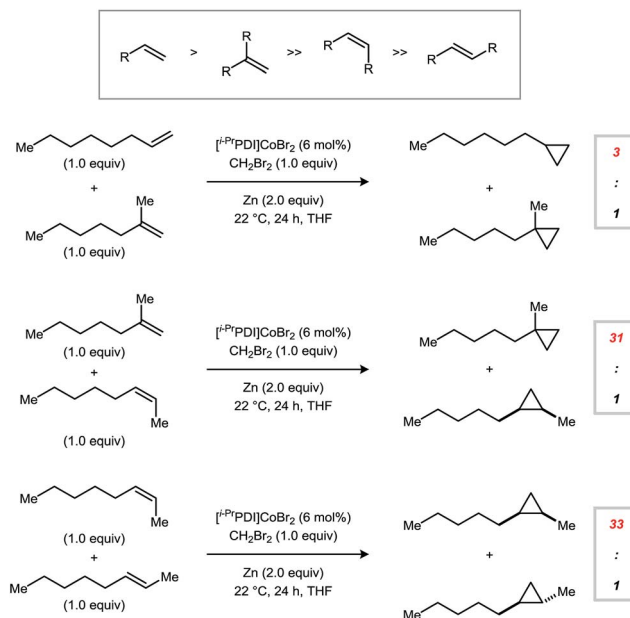


Fig. 2 Intermolecular competition experiments probing selectivity based on alkene substitution patterns. Reactions were conducted using a 1 : 1 : 1 molar ratio of the two alkene starting materials and CH₂Br₂.

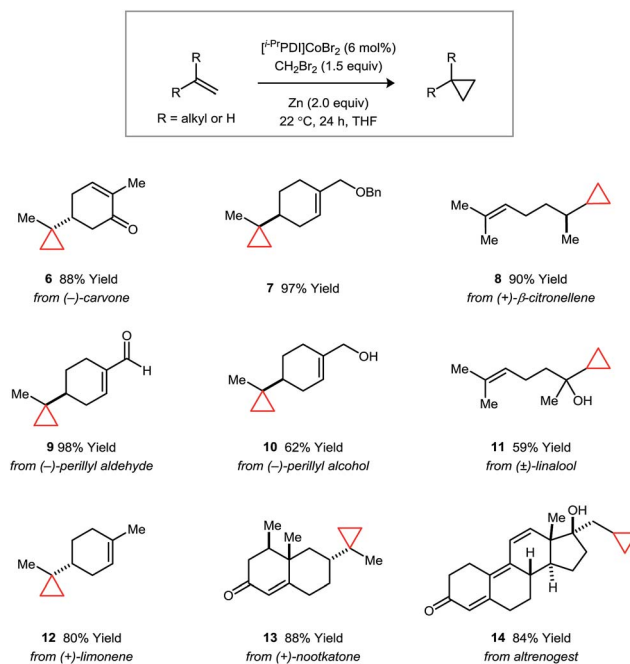


Fig. 3 Catalytic regioselective monocyclopropanations of terpene natural products and derivatives. Isolated yields following purification are averaged over two runs.

cyclopropanation of the *Z*-alkene **24** affords the *cis*-disubstituted cyclopropane **25** in 95% yield as a single diastereomer (Fig. 5a). Furthermore, the vinylcyclopropane substrates **26** and **28**, commonly used as tests for cyclopropylcarbinyl radical intermediates, react without ring-opening to afford products **27** and **29** (Fig. 5b).



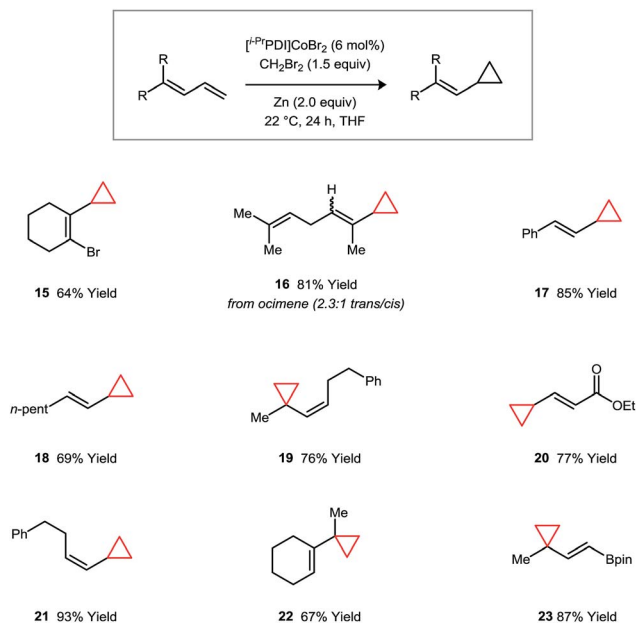


Fig. 4 Catalytic regioselective monocyclopropanations of 1,3-dienes. Isolated yields following purification are averaged over two runs.

Under standard catalytic conditions, the reaction mixtures using **1** adopt a deep violet color, which persists until complete consumption of the alkene. The UV-vis spectrum of the catalytic mixture at partial conversion is consistent with a Co(I) resting state (Fig. 6a). The authentic $[i\text{-Pr}]\text{PDI}]\text{CoBr}$ complex (**30**) can be prepared by stirring the $[i\text{-Pr}]\text{PDI}]\text{CoBr}_2$ complex **1** over excess Zn metal.²² Cyclic voltammetry data (Fig. 6b) indicates an $E_{1/2}$ for the Co(II)/Co(I) redox couple of $-1.00\text{ V vs. Fc/Fc}^+$. The large peak-to-peak separation (0.96 V in $0.3\text{ M }[n\text{-Bu}_4\text{N}][\text{PF}_6]/\text{THF}$) is characteristic of a slow bromide dissociation step following electron transfer. The second Co(I)/Co(0) reduction event is significantly more cathodic at -1.93 V and is inaccessible using Zn.

In order to decouple the cyclopropanation steps of the mechanism from catalyst turnover, we conducted stoichiometric reactions with the isolated $[i\text{-Pr}]\text{PDI}]\text{CoBr}$ complex in the

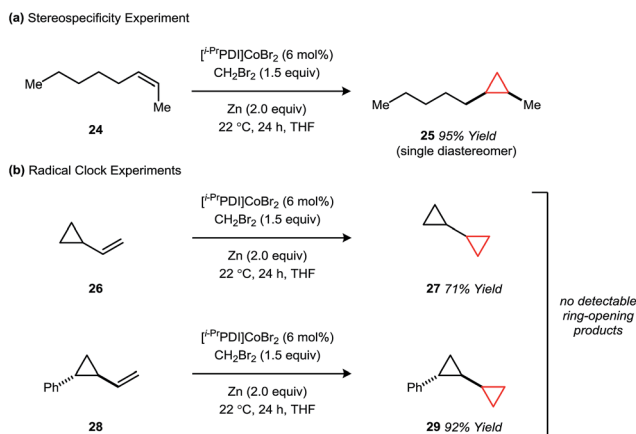


Fig. 5 Mechanistic studies probing the concertedness of cyclopropane formation.

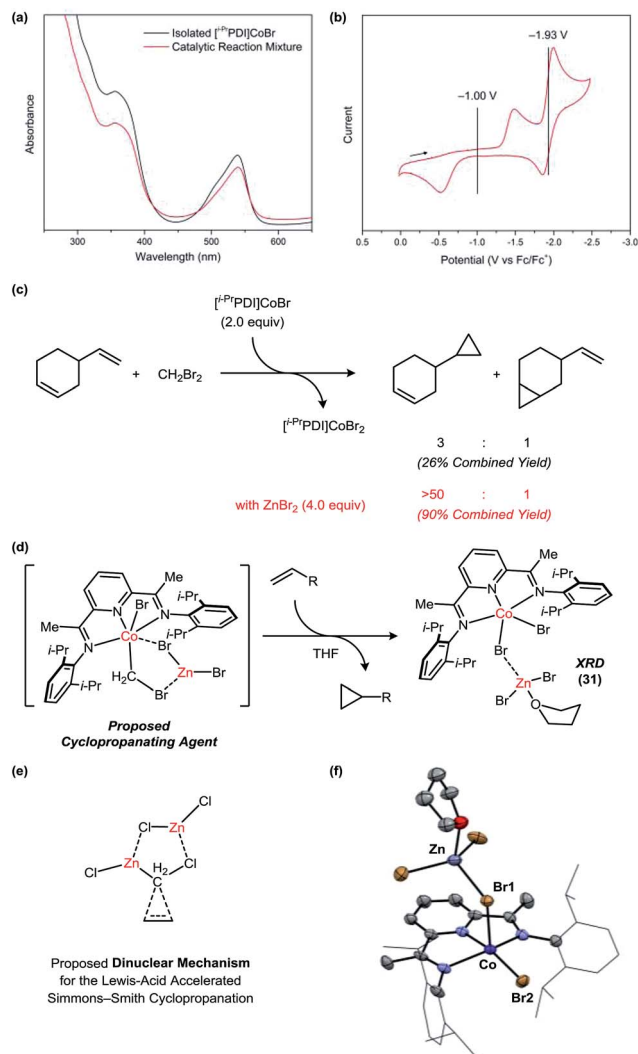


Fig. 6 Mechanistic studies probing the nature of the active carbenoid intermediate. (a) Identification of the catalyst resting state. (b) Cyclic voltammetry data for the $[i\text{-Pr}]\text{PDI}]\text{CoBr}_2$ complex. (c) Stoichiometric cyclopropanation reactions using the $[i\text{-Pr}]\text{PDI}]\text{CoBr}_2$ complex in the absence and presence of ZnBr_2 . (d) A proposed Co/Zn carbenoid species. (e) Proposed dinuclear mechanism for the Lewis acid-accelerated Simmons–Smith reaction. (f) Solid-state structure for the $[i\text{-Pr}]\text{PDI}]\text{CoBr}_2\text{Zn}(\text{THF}/\text{Et}_2\text{O})\text{Br}_2$ complex (**31**). The solvent molecule bound to Zn is disordered between Et_2O and THF. Only the THF-bound structure is shown for clarity.

absence of Zn (Fig. 6c). The reaction of **30** with 4-vinylcyclohexene and CH_2Br_2 generates the $[i\text{-Pr}]\text{PDI}]\text{CoBr}_2$ complex **1** within 24 h at room temperature but forms cyclopropanated products in a relatively low combined yield of 26%, which is not commensurate with the efficiency of the catalytic process. Furthermore, the regioselectivity is only 3 : 1, whereas the catalytic cyclopropanation achieves a $>50 : 1$ selectivity for this substrate. When the same stoichiometric reaction is conducted in the presence of ZnBr_2 , the yield and selectivity of the catalytic process is fully restored.

The Co-containing product (**31**) of the stoichiometric reaction in the presence of ZnBr_2 is green, which is notably distinct



from the tan color of the $[\text{t}^{\text{-Pr}}\text{PDI}]\text{CoBr}_2$ complex **1**. This green species is NMR silent but may be crystallized from saturated solutions in Et_2O to afford **31** (Fig. 6f). The solid-state structure reveals the expected $[\text{t}^{\text{-Pr}}\text{PDI}]\text{CoBr}_2$ fragment in a distorted square pyramidal geometry ($\tau_5 = 0.36$) with a $\text{Zn}(\text{THF}/\text{Et}_2\text{O})\text{Br}_2$ Lewis acid coordinated to one of the Br ligands. This interaction induces an asymmetry in the structure, causing the Co–Br1 distance (2.557(1) Å) to be elongated relative to the Co–Br2 distance (2.358(2) Å).

Collectively, these studies suggest that both Co and Zn are present in the reactive carbenoid intermediate, and that ZnBr_2 may interact with the $[\text{t}^{\text{-Pr}}\text{PDI}]\text{Co}$ complex through Lewis acid–base interactions. There is a notable similarity between the observed Co/Zn effect and previous studies of Lewis acid acceleration in the Simmons–Smith cyclopropanation. For example, Zn carbenoid reactions are known to be accelerated by the presence of ZnX_2 ,^{12c} which is generated as a byproduct of the reaction. DFT calculations conducted by Nakamura have suggested that the origin of this rate acceleration may be due to the accessibility of a five-membered ring transition state, which requires the presence of an additional Zn equivalent to function as a halide shuttle.²³

Conclusions

In summary, transition metal catalysis provides a pathway to accessing unique selectivity in reductive carbenoid transfer reactions. A $[\text{t}^{\text{-Pr}}\text{PDI}]\text{CoBr}_2$ complex functions as a robust catalyst for Simmons–Smith type cyclopropanation using a $\text{CH}_2\text{Br}_2/\text{Zn}$ reagent mixture. This system exhibits the highest regioselectivities that have been observed in reductive cyclopropanations based solely on the steric properties of the alkene substrate. Accordingly, a range of terpenes and conjugated dienes may be converted to a single monocyclopropanated product. Ongoing studies are directed at exploring the applications of transition metal catalysts to other classes of carbenoid transfer reactions.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by Purdue University and the National Institutes of Health (R35GM124791). We thank Dr Matthias Zeller for assistance with XRD experiments. XRD data were collected using an instrument funded by the NSF (CHE-1625543). C. U. is an Alfred P. Sloan Foundation Research Fellow.

Notes and references

- (a) J. Salaun and M. Baird, *Curr. Med. Chem.*, 1995, **2**, 511–542; (b) W. A. Donaldson, *Tetrahedron*, 2001, **57**, 8589–8627; (c) L. A. Wessjohann, W. Brandt and T. Thiemann, *Chem. Rev.*, 2003, **103**, 1625–1648; (d) N. A. Meanwell, *J. Med. Chem.*, 2011, **54**, 2529–2591; (e) D. Y. K. Chen, R. H. Pouwer and J.-A. Richard, *Chem. Soc. Rev.*, 2012, **41**, 4631–4642.
- (a) H. E. Simmons and R. D. Smith, *J. Am. Chem. Soc.*, 1958, **80**, 5323–5324; (b) H. E. Simmons and R. D. Smith, *J. Am. Chem. Soc.*, 1959, **81**, 4256–4264; (c) A. B. Charette and A. Beauchemin, *Org. React.*, 2001, **58**, 1–415.
- (a) J. Furukawa, N. Kawabata and J. Nishimura, *Tetrahedron Lett.*, 1966, **7**, 3353–3354; (b) J. Furukawa, N. Kawabata and J. Nishimura, *Tetrahedron*, 1968, **24**, 53–58.
- (a) Z. Yang, J. C. Lorenz and Y. Shi, *Tetrahedron Lett.*, 1998, **39**, 8621–8624; (b) J. C. Lorenz, J. Long, Z. Yang, S. Xue, Y. Xie and Y. Shi, *J. Org. Chem.*, 2004, **69**, 327–334.
- A. B. Charette, S. Francoeur, J. Martel and N. Wilb, *Angew. Chem., Int. Ed.*, 2000, **39**, 4539–4542.
- A. Voituriez, L. E. Zimmer and A. B. Charette, *J. Org. Chem.*, 2010, **75**, 1244–1250.
- A. B. Charette, J.-F. Marcoux, C. Molinaro, A. Beauchemin, C. Brochu and É. Isabel, *J. Am. Chem. Soc.*, 2000, **122**, 4508–4509.
- There is some conflicting data in the literature regarding the regioselectivity of Zn carbenoid addition to limonene; however, the most recent and comprehensive studies have found that mixtures of regioisomers are formed and that the product ratios depend on the reaction conditions: (a) S. D. Koch, R. M. Kliss, D. V. Lopicke and R. J. Wineman, *J. Org. Chem.*, 1961, **26**, 3122–3125; (b) H. E. Simmons, E. P. Blanchard and R. D. Smith, *J. Am. Chem. Soc.*, 1964, **86**, 1347–1356; (c) E. C. Friedrich and F. Niyati-Shirkhodae, *J. Org. Chem.*, 1991, **56**, 2202–2205.
- (a) A. H. Hoveyda, D. A. Evans and G. C. Fu, *Chem. Rev.*, 1993, **93**, 1307–1370; (b) H. Lebel, J.-F. Marcoux, C. Molinaro and A. B. Charette, *Chem. Rev.*, 2003, **103**, 977–1050.
- (a) H. Nozaki, S. Moriuti, M. Yamabe and R. Noyori, *Tetrahedron Lett.*, 1966, **7**, 59–63; (b) M. P. Doyle, *Chem. Rev.*, 1986, **86**, 919–939; (c) M. P. Doyle, M. A. McKervey and T. Ye, *Modern catalytic methods for organic synthesis with diazo compounds*, Wiley, 1998; (d) H. Pellissier, *Tetrahedron*, 2008, **64**, 7041–7095.
- Photoredox catalysts for reductive cyclopropanation reactions: (a) A. M. del Hoyo, A. G. Herraiz and M. G. Suero, *Angew. Chem., Int. Ed.*, 2017, **56**, 1610–1613; (b) A. M. del Hoyo and M. García Suero, *Eur. J. Org. Chem.*, 2017, **2017**, 2122–2125.
- (a) H. Takahashi, M. Yoshioka, M. Ohno and S. Kobayashi, *Tetrahedron Lett.*, 1992, **33**, 2575–2578; (b) A. B. Charette and C. Brochu, *J. Am. Chem. Soc.*, 1995, **117**, 11367–11368; (c) S. E. Denmark and S. P. O'Connor, *J. Org. Chem.*, 1997, **62**, 584–594; (d) A. B. Charette, C. Molinaro and C. Brochu, *J. Am. Chem. Soc.*, 2001, **123**, 12168–12175; (e) H. Shitama and T. Katsuki, *Angew. Chem., Int. Ed.*, 2008, **47**, 2450–2453; (f) J. Balsells and P. J. Walsh, *J. Org. Chem.*, 2000, **65**, 5005–5008.
- J. Long, H. Du, K. Li and Y. Shi, *Tetrahedron Lett.*, 2005, **46**, 2737–2740.
- Y.-Y. Zhou and C. Uyeda, *Angew. Chem., Int. Ed.*, 2016, **55**, 3171–3175.



- 15 S. Pal, Y.-Y. Zhou and C. Uyeda, *J. Am. Chem. Soc.*, 2017, **139**, 11686–11689.
- 16 Seminal studies and selected reviews of catalytic reactions using (PDI)Co complexes: (a) B. L. Small, M. Brookhart and A. M. A. Bennett, *J. Am. Chem. Soc.*, 1998, **120**, 4049–4050; (b) G. J. P. Britovsek, V. C. Gibson, S. J. McTavish, G. A. Solan, A. J. P. White, D. J. Williams, G. J. P. Britovsek, B. S. Kimberley and P. J. Maddox, *Chem. Commun.*, 1998, 849–850; (c) V. C. Gibson, C. Redshaw and G. A. Solan, *Chem. Rev.*, 2007, **107**, 1745–1776; (d) C. Bianchini, G. Giambastiani, L. Luconi and A. Meli, *Coord. Chem. Rev.*, 2010, **254**, 431–455; (e) Z. Flisak and W.-H. Sun, *ACS Catal.*, 2015, **5**, 4713–4724; (f) P. J. Chirik, *Angew. Chem., Int. Ed.*, 2017, **56**, 5170–5181.
- 17 U. Burger and R. Huisgen, *Tetrahedron Lett.*, 1970, **11**, 3057–3060.
- 18 E. C. Friedrich, S. E. Lunetta and E. J. Lewis, *J. Org. Chem.*, 1989, **54**, 2388–2390.
- 19 K. Maruoka, Y. Fukutani and H. Yamamoto, *J. Org. Chem.*, 1985, **50**, 4412–4414.
- 20 (a) T. Hudlicky, T. M. Kutchan and S. M. Naqvi, *Org. React.*, 1985, **33**, 247–335; (b) J. E. Baldwin, *Chem. Rev.*, 2003, **103**, 1197–1212; (c) T. Hudlicky and J. W. Reed, *Angew. Chem., Int. Ed.*, 2010, **49**, 4864–4876; (d) Y. Gao, X.-F. Fu and Z.-X. Yu, in *CC Bond Activation*, Springer, 2014, pp. 195–231; (e) L. Soullart and N. Cramer, *Chem. Rev.*, 2015, **115**, 9410–9464.
- 21 A. J. Anciaux, A. Demonceau, A. F. Noels, R. Warin, A. J. Hubert and P. Teyssié, *Tetrahedron*, 1983, **39**, 2169–2173.
- 22 M. J. Humphries, K. P. Tellmann, V. C. Gibson, A. J. P. White and D. J. Williams, *Organometallics*, 2005, **24**, 2039–2050.
- 23 (a) E. Nakamura, A. Hirai and M. Nakamura, *J. Am. Chem. Soc.*, 1998, **120**, 5844–5845; (b) M. Nakamura, A. Hirai and E. Nakamura, *J. Am. Chem. Soc.*, 2003, **125**, 2341–2350.

