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Electrochemical synthesis of n-type ZnS layers on p-Cu₂O/n-ZnO heterojunctions with different deposition temperatures

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Metal oxide p–n heterojunctions consisting of p-Cu₂O/n-ZnO/n-ZnS nanostructures were deposited on an ITO substrate by three-step electrodeposition. The effect of ZnS layer deposition temperature on the properties of the heterojunction was investigated by different techniques. The Mott–Schottky analysis confirmed the n-type conductivity for ZnO and ZnS and p-type conductivity for the Cu₂O layer, respectively. Also, it showed a decrease of ZnS donor concentration with increasing deposition temperature. The X-ray diffraction (XRD) analysis confirms a pure phase of hexagonal ZnO, cubic ZnS and cubic Cu₂O structures, respectively. The heterojunction with ZnS deposited at 60 °C shows high crystallinity. The morphological measurements by scanning electron microscopy (SEM) indicate that the deposition temperature has a significant influence on the morphology of ZnO and the atomic force microscopy (AFM) images revealed the improvement of Cu₂O morphology by increasing the ZnS deposition temperature. The UV-Vis response shows strong absorption in the visible region and the profile of optical absorption spectra changes with the ZnS deposition temperature. The current–voltage (*I*–*V*) characteristics of the Au/p-Cu₂O/n-ZnO/n-ZnS/ITO heterojunction display well-defined rectifying behavior for the heterojunction with ZnS deposited at 60 °C.

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1. Introduction

In the last 20 years, solar cells based on inorganic materials have been extensively investigated due to their excellent photovoltaic properties.^{1,2} Cuprous oxide (Cu₂O) is a promising candidate material for low-cost solar cell production owing to its low cost, non-toxicity and long-term stability.^{3,4} Cu₂O is a p-type semiconductor with a direct band gap ($E_g = 1.9\text{--}2.2$ eV), and a high absorption coefficient (10^5 to 10^6 cm⁻¹) for the photon energies higher than 2.7 eV. In addition, it has a theoretical maximum efficiency of 20% under AM1 solar illumination.⁵ Actually, Cu₂O has been employed in many heterojunctions such as CdO/Cu₂O, In₂O₃/Cu₂O and ZnO/Cu₂O with ZnO as the most successful n-type counterpart.^{6–8} To date, the highest efficiency of Cu₂O based solar cells is 8.1%, obtained for a polycrystalline Cu₂O absorber and a Zn_{1-x}Ge_xO buffer layer.⁹

Also, zinc oxide (ZnO) and zinc sulfide (ZnS) are interesting II–IV semiconductors that attract intense attention due to their wide applications in nanodevices and nano-systems, such as solar cells,¹⁰ sensors¹¹ and light-emitting diodes.¹² The ZnO is n-type semiconductor with wideband gap (3.37 eV), high transparency in the visible range and a large exciton energy (60 meV) at room temperature.¹³ As well, the ZnS shows a high optical transparency from the ultraviolet (UV) to the infrared (IR) region¹⁴ and a large direct band gap of 3.68 eV.¹⁵ Compared to the ZnO and ZnS individual layers, ZnO/ZnS heterostructure presents a better electronic, structural and optical properties.^{16–18} Due to that, the ZnO/ZnS heterostructures have already proved a high potential for application in multi-functional devices such as UV detectors and biological applications.¹⁹ Moreover, the theoretical calculation of ZnO/ZnS heterostructures optical properties predicted that such structure could be used as window layer in photovoltaic devices.²⁰ Therefore, such heterostructures are suitable as window layer in Cu₂O based solar cells. To our knowledge, the fabrication of Cu₂O/ZnO/ZnS by three-step electrochemical deposition has not yet been reported.

Recently, several efforts have been made to provide convenient and efficient approaches for depositing Cu₂O based heterojunction. However, these techniques are still too expensive and complicated. As another approach,

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electrochemical deposition was recognized as competitive method for manufacturing semiconductor thin films due to the following advantages: (i) easy control of the thickness and morphology of the film (ii) simplicity and low equipment cost (iii) the possibility of large area thin films elaboration.

In our group, high quality of ZnO and Cu₂O nanostructures were prepared under the optimal condition.^{21,22} In this study, the Cu₂O/ZnO/ZnS heterojunction was prepared by three step electrochemical deposition on indium-doped tin oxide (ITO) substrate. The effect of the deposition temperature of ZnS layer on the structural, morphological, optical and electrical properties of the heterojunction was studied in details.

2. Experimental procedure

The electrochemical deposition of Cu₂O/ZnO/ZnS heterojunction was carried out in a typical three-electrode cell using a computer-controlled potentiostat/galvanostat (Voltalab 40) as a potential source. All potentials were reported with reference to the SCE scale. Note that the working electrode was indium tin oxide coated glass (ITO) with sheet resistance of 30 Ω cm⁻², a saturated calomel electrode (SCE, +0.241 V vs. SHE) and a platinum electrode were used as reference electrode and counter electrode, respectively. Prior to the electrodeposition the substrates were ultrasonically cleaned with acetone, ethanol and distilled water for 10 min, respectively. The ZnS was directly electro-deposited onto ITO substrate from aqueous electrolyte bath containing 10⁻⁴ M ZnSO₄ and 10⁻³ M Na₂S₂O₃, the solution pH was adjusted to 2.4 using sulphuric acid. The deposition temperature of ZnS layers was varied from 40 to 70 °C. Then, the ZnO layer was electrodeposited onto ZnS/ITO substrate by pulse potential deposition with V_{on} of -1.3 V and V_{off} of -0.1 V were applied during 1 s for 35 cycles.

The ZnO deposition bath was containing 0.08 M Zn(NO₃)₂ and 1 M KNO₃ and temperature of 70 °C and pH 6. Then, ZnO/ZnS/ITO was used as substrate for the electro-deposition of Cu₂O layer at fixed temperature of 65 °C and applied potential of -0.5 V. The Cu₂O deposition bath was containing a 0.05 M of CuSO₄ stabilized with 0.05 M citric acid as chelating agent. The pH value was adjusted to 11 using NaOH solution. After the deposition, the samples were washed with distilled water, dried in air and kept for analysis, as show in Fig. 1.

The thickness (d) of the electrodeposited layers can be estimated from the charge quantity (Q) following Faraday's law:³⁶

$$d = \frac{QM}{nFA\rho} \quad (1)$$

where M is the molecular weight of the deposited semiconductor, n is the number of electrons transferred per atom discharged, A is the working electrode area, ρ is specific density and F is the Faraday constant.

In our study, the thicknesses of ZnO, ZnS and Cu₂O layers were fixed at 200, 200 and 300 nm, respectively.

The electrochemical study was carried out using computer-controlled potentiostatic/galvanostatic (Voltalab 40). The Mott-Schottky analysis was realised in 1.0 M KNO₃ and employed a frequency of 0.2 kHz for ZnO and ZnS layers and in 0.5 M Na₂SO₄ solution under applied frequency of 2 kHz for Cu₂O layer. Nyquist and Bode electrochemical impedance spectroscopy (EIS) measurements of Cu₂O layer and Cu₂O/ZnO/ZnS heterojunctions was performed using an alternative current voltage of 10 mV, in a 0.5 M Na₂SO₄ solution, at frequency range of 10⁵ to 10⁻² Hz. The photocurrent response of Cu₂O layer and Cu₂O/ZnO/ZnS heterojunctions was recorded at -0.5 V vs. SCE in 0.5 M Na₂SO₄, under chopped irradiation of 500 W xenon (Xe) lamps. The morphological characterization of the films was performed by field emission scanning electron microscopy (FESEM) (JEOL JSM 7000F). The roughness (root-mean-square height deviation) was obtained directly from the software of the AFM (PicoScan 5.3 from molecular imaging) and the crystallographic structure were identified using X-ray diffraction (XRD) Rigaku Smartlab® X-ray diffractometer with a Cu K α_1 radiation (45 kV, 200 mA, $\lambda = 0.154056$ nm). The optical properties were investigated using UV-Vis-NIR spectrophotometer (Shimadzu UV-1800) in the UV-visible region. The current-voltage (I - V) measurements were recorded using a picoammeter (Keithley 617) at room temperature in the dark. The ohmic behavior of the electrodes was confirmed by I - V measurement between the gold pads deposited on Cu₂O surface and ITO substrate.

3. Results and discussion

3.1 Mott-Schottky analysis

Mott-Schottky analysis (M-S) is well suitable to have information about the semiconductor characteristics, such as the flat band potential and the acceptor/donor concentration without damaging its surface. It is based on examining the semiconductor/solution interface and measuring the capacitance of the semiconductor depletion layer as function of the applied voltage following Mott-Schottky relation.^{23,24}

$$\frac{1}{C^2} = \frac{2}{\varepsilon\varepsilon_0eN_D} \left(E - E_{FB} - \frac{kT}{e} \right) \quad \text{for n-type semiconductor} \quad (2)$$

$$\frac{1}{C^2} = \frac{2}{-\varepsilon\varepsilon_0eN_A} \left(E - E_{FB} - \frac{kT}{e} \right) \quad \text{for p-type semiconductor} \quad (3)$$

where, e is the electron charge, ε_0 is the vacuum permittivity, ε is the dielectric constant of 8.5 for ZnO²⁵ and 7.6 for Cu₂O,²⁶ k is the Boltzmann constant, T is the absolute temperature, N_D and N_A are donor and acceptor concentration of the n-type and p-type semiconductor, respectively and E_{fb} is the flat band potential.

Fig. 2 presents Mott-Schottky plots ZnS layers deposited at different temperatures, ZnO layer and Cu₂O layer,



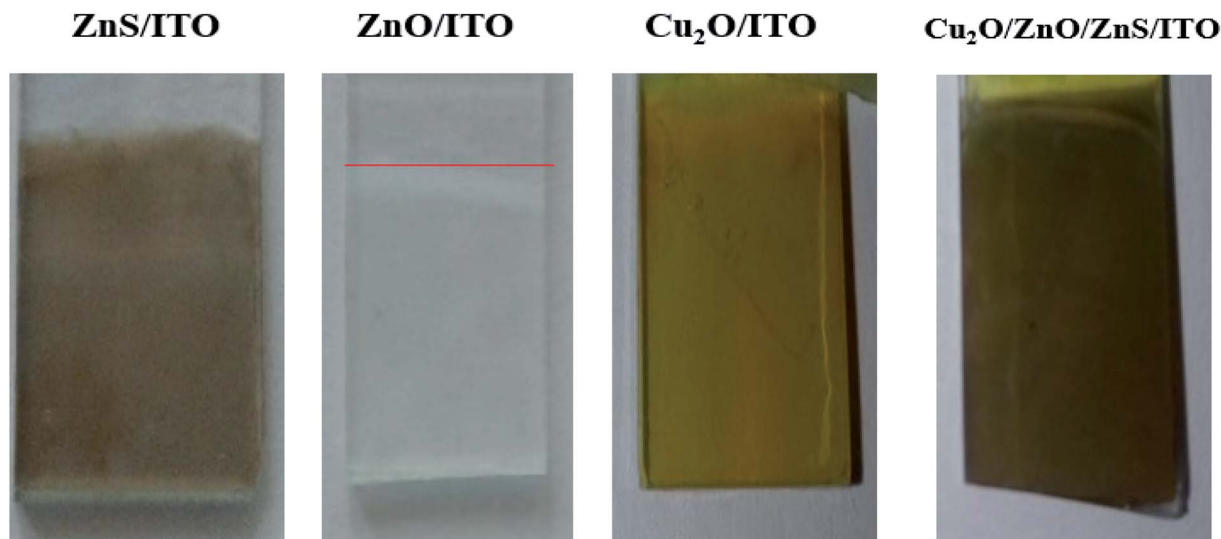


Fig. 1 Photographs of ZnS, ZnO, Cu₂O individual electrodeposited layers and Cu₂O/ZnO/ZnS heterojunction deposited on ITO substrate, respectively.

respectively. It is clear that the Mott-Schottky equation is valid in a wide range of potential, which indicates well-defined electronic surface states. The M-S plots in Fig. 2a

and b reveal straight lines with positive slopes, which confirm the n-type characteristic of ZnS and ZnO layers, respectively. While, the negative slope of M-S plot in Fig. 2c

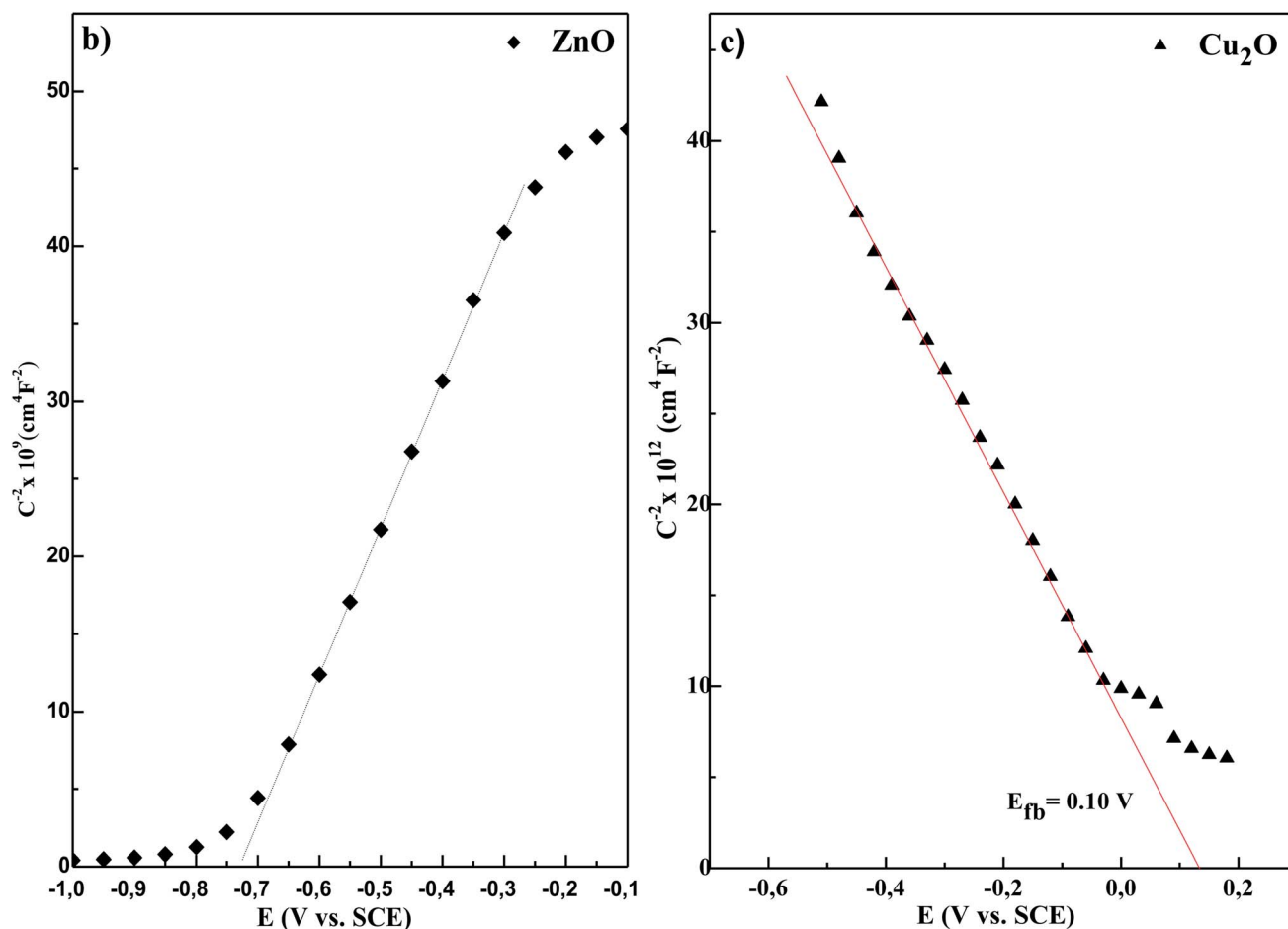


Fig. 2 Mott-Schottky plots recorded in 1.0 M KNO₃ electrolyte for: (a) ZnS layers deposited at different temperatures, (b) ZnO layer and (c) recorded in 0.5 M Na₂SO₄ for Cu₂O layer. The corresponding flat band potential values are indicated.



confirms the p-type conductivity of Cu₂O layer. In addition, the presence of multiple linear regions in M–S plots of ZnO and ZnS layers indicates the existence of multiple donor levels.²⁷

The flat band potentials were determined from the extrapolation of the linear portion of M–S plots to $\frac{1}{C^2} = 0$, while the donor and acceptor concentration values were calculated from the slope of M–S plots ($\frac{2}{\epsilon\epsilon_0 e N_D}$, the obtained values are set out in Table 1). As can be seen from the data in Table 1, the acceptor concentration was $5.01 \times 10^{17} \text{ cm}^{-3}$ for Cu₂O layer, which is within the typical value (10^{15} to 10^{18} cm^{-3}) reported in the literature.²⁸ Turning to the ZnO layer, high donor concentration of $1.8 \times 10^{20} \text{ cm}^{-3}$ was obtained, which is in accordance with other studies.²⁹ On the other hand, the ZnS layers show high donor concentration. When the deposition temperature of ZnS increases from 40 to 70 °C, the donor concentration decreases from 2.51×10^{21} to $7.28 \times 10^{19} \text{ cm}^{-3}$. Thus, the ZnS layer deposited at higher temperature have the lowest conductivity. This may be due to the higher density of grain boundaries associated with the smaller crystallites size of this sample observed in XRD patterns.³⁰ Accordingly, the flat band potential of ZnS layers increases from -0.94 to -0.51 V vs. SCE with increasing the deposition temperature from 40 to 70 °C, this may be due to the decrease in donor concentration in n-type ZnS layers.³¹

3.2 Electrochemical impedance spectroscopy analysis (EIS)

EIS analysis is one of the most common procedures for evaluating the electrical properties of materials due to its extraordinary sensitivity. Fig. 3 presents the Nyquist and Bode diagrams of Cu₂O and Cu₂O/ZnO/ZnS heterojunctions, with different deposition temperatures of ZnS layer, recorded in 0.5 M Na₂SO₄ solution and frequency range from 10^5 to 10^{-2} Hz . As displayed in Fig. 3a and b, the Nyquist EIS spectra of all samples show semicircle at high frequencies followed by diagonal line at low frequencies (Warburg impedance). It is well known that the semicircle is equal to the charge transfer resistance (R_{ct}), while the diagonal line is related to the diffusion process. In our case, the Warburg line was dominating, in such system, the mass transport is rate limiting due to the high charge-transfer rate and the low

rate of charge recombination which will definitely improve the performance of solar cell device.⁵

The reduction in the diameter of semicircle is related to the decrease of the interfacial charge transfer resistance (R_{ct}) across the electrode/electrolyte interface.³² By fitting the Nyquist plots we obtained the charge transfer resistance (R_{ct}) values for Cu₂O layer and Cu₂O/ZnO/ZnS heterojunctions, with different deposition temperatures of ZnS layer, as summarized in Table 1. It is clear that R_{ct} values of Cu₂O/ZnO/ZnS heterojunctions are lower than that of Cu₂O layer indicating that the resistance of the charge movement was significantly lower in Cu₂O/ZnO/ZnS heterojunction interface. Thus, the Cu₂O/ZnO/ZnS heterojunction improves significantly the separation efficiency of photogenerated electron–hole pairs.³³ In exception of Cu₂O/ZnO/ZnS heterojunctions with ZnS deposited at 40 °C and 70 °C that show R_{ct} higher than Cu₂O layer.

Comparing the R_{ct} of Cu₂O/ZnO/ZnS heterojunctions, the ZnS deposited at 60 °C reveals the lowest resistance ($0.41 \times 10^3 \Omega \text{ cm}^2$), which is expected to have a good electrical property compared to others samples.

Fig. 3c presents Bode phase plot of Cu₂O/ZnO/ZnS heterojunctions with different deposition temperatures of ZnS layer. In the Bode plot, the peak frequency of the mid-frequency region (1 to 100 Hz) is inversely related to the lifetime of the photogenerated electrons.^{34,35} Fig. 3c reveals that increasing the deposition temperature of ZnS results in an obvious drop of the peak frequency. The heterojunction with ZnS deposited at 60 °C shows the lowest peak frequency indicating high lifetime of photogenerated electrons. This confirms that the ZnS layer deposited at 60 °C act as an effective surface modification of the window layer to help increase the electron lifetime within the Cu₂O/ZnO/ZnS heterojunction.

The separation of photogenerated electron hole pairs in pure Cu₂O and Cu₂O/ZnO/ZnS heterojunctions will be further investigated in the following section.

3.3 Photoelectrochemical measurements

The photo-activity of pure Cu₂O and Cu₂O/ZnO/ZnS heterojunction was investigated using photoelectrochemical measurement in order to determine the influence of ZnS deposition temperature on the photo-generation and separation of electron–hole pairs in Cu₂O/ZnO/ZnS heterojunction. The transient photocurrent responses performed in 0.5 M Na₂SO₄ solution and an applied potential of -0.5 V under chopped illumination are shown in Fig. 4.

As can be seen from Fig. 4a, the cathodic photocurrent confirms that Cu₂O layer behaves as p-type semiconductor, which is in good agreement with Mott–Schottky plot (Fig. 2c). Under irradiation, the pure Cu₂O layer grown on ITO generates photocurrent of 245.6 μA . The deposition of Cu₂O on ZnO/ZnS heterostructures at different deposition temperatures caused an increase of photocurrent density. In addition, it can be seen from Fig. 4b that the photocurrent density of the Cu₂O deposited onto ZnO/ZnS/ITO layer

Table 1 Carrier density, flat band potential and charge transfer resistance of all layers in Cu₂O/ZnO/ZnS/ITO heterojunction with different electrodeposition temperatures of ZnS layer

Samples	$N_{D,A} (\text{cm}^{-3})$	$V_{fb} (\text{V per SCE})$	$R_{ct} (\text{k}\Omega \text{ cm}^2)$
Cu ₂ O/ZnO/ZnS (40 °C)	2.51×10^{21}	-0.94	1.088
Cu ₂ O/ZnO/ZnS (50 °C)	4.94×10^{20}	-0.72	0.45
Cu ₂ O/ZnO/ZnS (60 °C)	5.06×10^{19}	-0.59	0.41
Cu ₂ O/ZnO/ZnS (70 °C)	7.28×10^{19}	-0.51	0.57
ZnO	1.80×10^{20}	-0.74	—
Cu ₂ O	5.01×10^{17}	$+0.10$	—



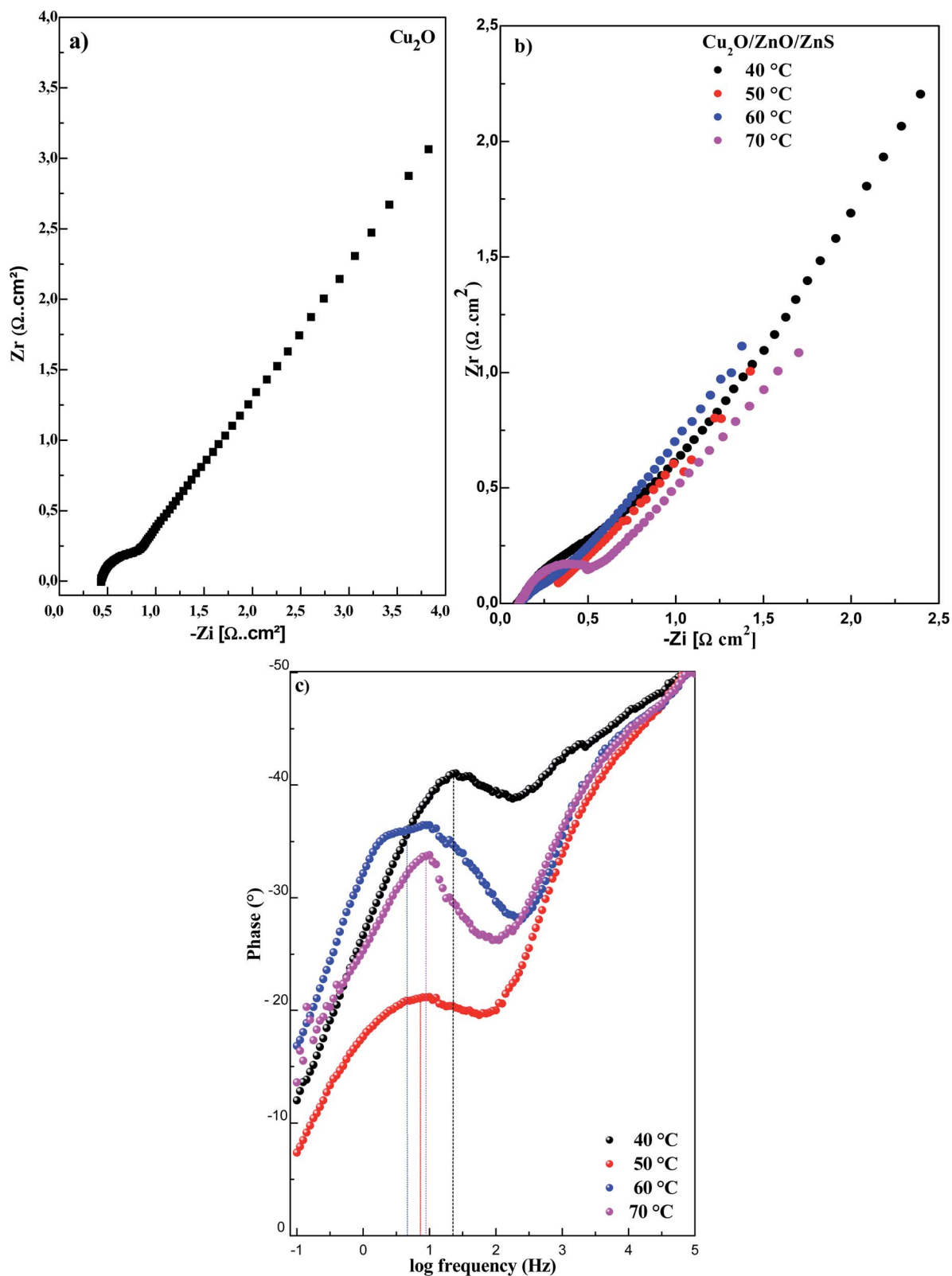


Fig. 3 Nyquist EIS diagrams recorded in 0.5 M Na_2SO_4 solution with a frequency range from 10^5 Hz to 10^{-2} Hz for: (a) Cu_2O layer and (b) $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}$ with different deposition temperatures of ZnS layer. (c) Bode EIS plots $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}$ with different deposition temperatures of ZnS layer.

with different deposited temperatures of ZnS layer high lights a photocurrent value between 78.9 and 480.9 μA . The Cu_2O layer deposited on ZnO/ZnS (60 °C)/ITO shows

a photocurrent density of 480.9 μA . It is clear that this heterojunction has a better performance compared to pure Cu_2O layer and the other samples. Hence, it has a better



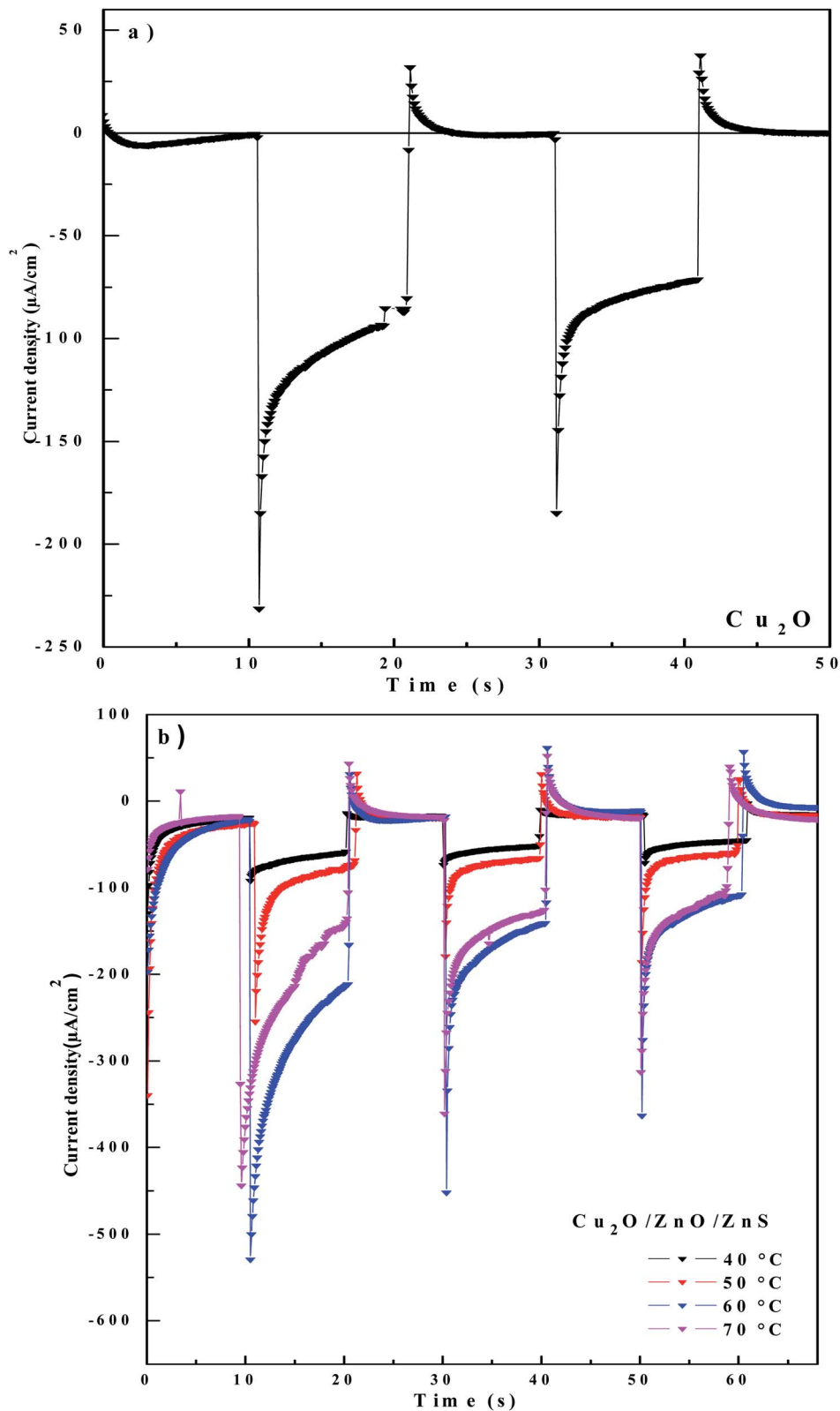


Fig. 4 Transient photocurrent response under bias potential of -0.5 V vs. SCE of: (a) Cu_2O layer and (b) $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunction with different deposition temperatures of ZnS layer.

efficiency of carrier's photo-generation and separation, which is in accord with Nyquist and Bode impedance results reported earlier (Fig. 3b and c).

3.4 Morphological characterization

The surface morphology and topography were examined using scanning electron microscopy (SEM) and atomic force



microscopy (AFM) images. Fig. 5a–d present top view SEM images of ZnO layer grown on ZnS/ITO with different deposition temperatures of ZnS layer. It is clear that increasing the deposition temperature of ZnS affects greatly the morphology of ZnO layer. As shown in Fig. 5a, spherical grains of ZnO were spread over the surface of ZnS layer deposited at 40 °C. The higher magnification (insert image) reveals a continuous layer of thin nanosheets of ZnO covering the grains. Increasing the deposition temperature of ZnS to 50 °C (Fig. 5b) increased the size of ZnO grains and we noted the formation of few nanorods of ZnO on ZnS deposited at 60 °C (Fig. 5c). The ZnO forms a dense film of large spherical grain and nanorods covered with thin layer of nanosheets, it is important to note that the size and the number of nanorods increased. Increasing the deposition temperature of ZnS to 70 °C (Fig. 5d) increased the size and amount of ZnO nano-rods and the granular morphology was almost disappeared. It is clear that the amount of ZnO grown on ZnS deposited at 60 °C is higher than other

samples and it covered entirely the ZnS surface. This may be due to the high crystallinity of this sample confirmed in XRD measurements. Fig. 6 displays 2D and 3D AFM images of Cu₂O/ITO and Cu₂O/ZnO/ZnS/ITO with different deposition temperatures of ZnS layer. As shown in Fig. 6a, the Cu₂O layer grown on ITO substrate reveals homogenous surface of pyramidal grains with low mean square roughness of 22.5 nm compared to Cu₂O layers deposited on ZnO/ZnS/ITO (Fig. 6b–e). On the other hand, the AFM images of Cu₂O deposited on ZnO/ZnS/ITO (Fig. 6b–e) show a great effect of ZnS deposition temperature on the morphology of Cu₂O. As ZnS deposition temperature increased the surface morphology changed from granular to pyramid shapes. The average surface roughness was 69.2, 65.4, 109.2 and 63.6 nm for of Cu₂O layer grown on ZnS deposited at 40, 50, 60 and 70 °C, respectively. The Cu₂O grown on ZnS deposited at 60 °C shows the highest roughness (109.2 nm). This value is consistent with the observation. The ZnO deposited on ZnS (60 °C) shows a large grain and good coverage of ZnS surface

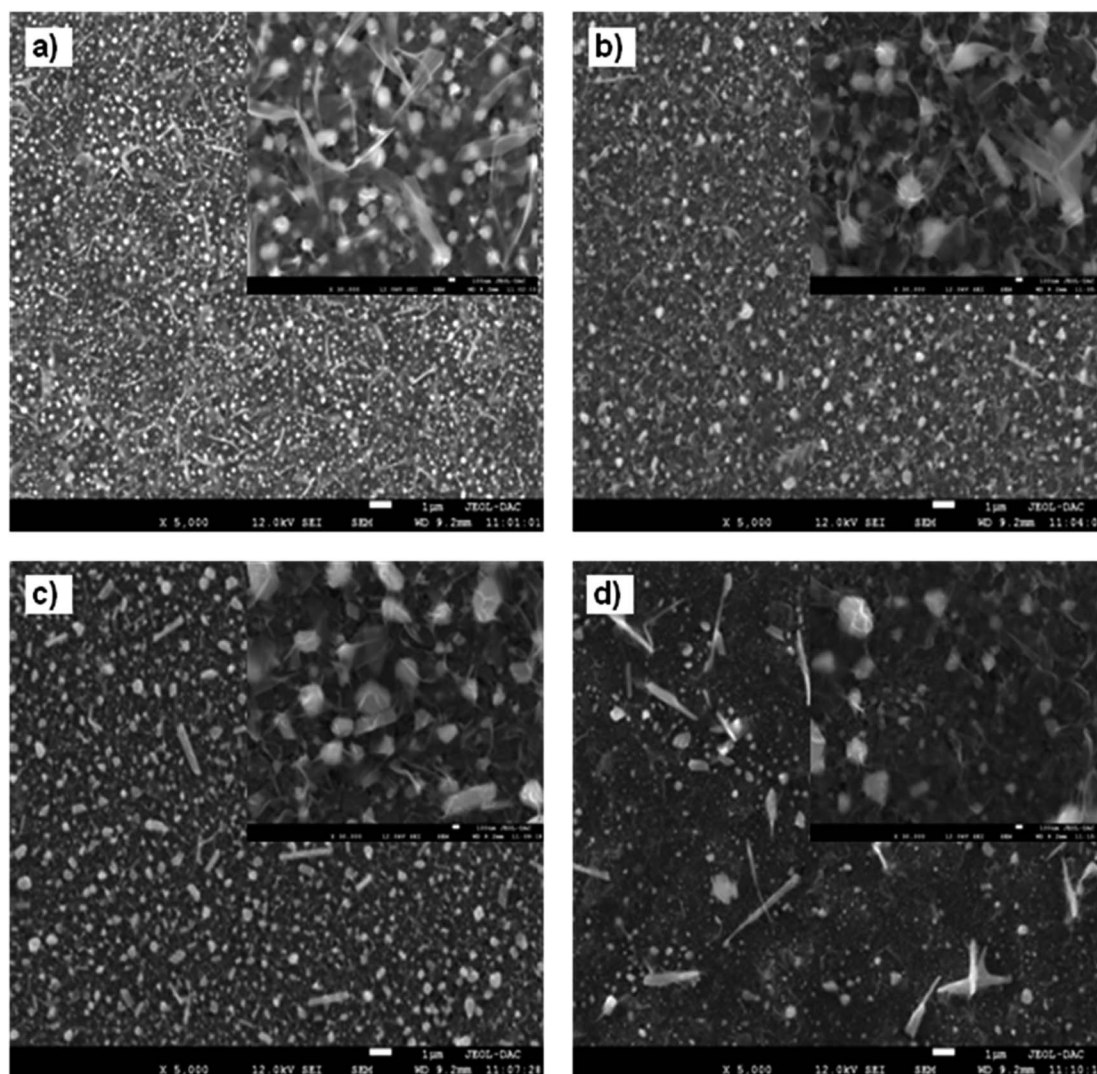


Fig. 5 FE-SEM images of ZnO films deposited on ZnS/ITO with different deposition temperatures of ZnS layer: (a) 40 °C, (b) 50 °C, (c) 60 °C and (d) 70 °C. Insert represents the higher magnification of the respective samples.



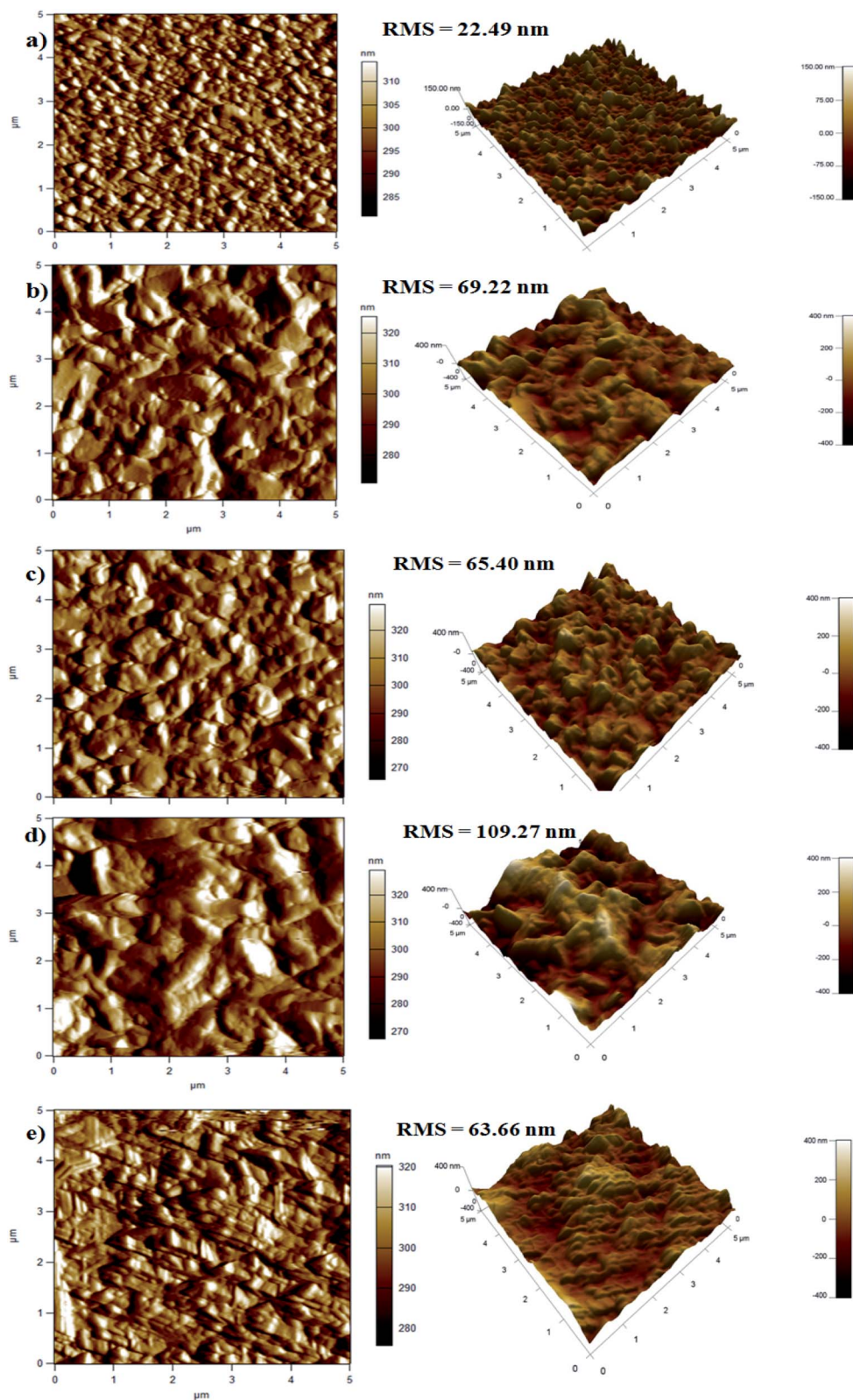


Fig. 6 Tapping mode 2D and 3D AFM images ($10 \times 10 \mu\text{m}^2$) of: (a) $\text{Cu}_2\text{O}/\text{ITO}$, (b–e) $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions with different electrodeposition temperatures of ZnS layer: (b) 40°C , (c) 50°C , (d) 60°C and (e) 70°C .

(insert in Fig. 5c) which facilitates the deposition of Cu_2O layer and consequently, the formation of rough surface of Cu_2O .

3.5 Structural characterization

To prove the crystalline structure, X-ray diffraction (XRD) was carried out for $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions with



different deposition temperature of ZnS layer. As shown from XRD pattern in Fig. 7, no impurity peaks were detected besides the diffraction peaks of the indium tin oxide (ITO) substrate (marked by stars in the pattern). The diffraction peaks observed at $2\theta \sim 36.45, 42.22^\circ$ and 61.35° are indexed respectively to the crystalline orientation (111), (200) and (221) of Cu_2O cubic structure (JCPDS card no 00-034-1354). While, the diffraction peak observed at $2\theta \sim 36.45^\circ$ can be indexed to the crystalline orientation (101) of ZnO hexagonal wurtzite structure according to (JCPDS card no 00-036-1451). Moreover, the supplement diffraction peak at $2\theta \sim 33.01^\circ$ corresponds to the (200) orientation of cubic ZnS according to the (JCPDS card no 00-005-0566). These results confirm the formation of pure $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions by three-step electrodeposition process. It is important to note that the preferential orientations are (111), (101) and (200) for Cu_2O , ZnO and ZnS, respectively. The presence of different diffraction peaks from basic planes of Cu_2O layer indicates that the dense films were composed of Cu_2O crystallites with random orientation. The heterojunction with ZnS layer deposited at 60°C shows the highest of (200) ZnS peak which indicates the high crystallinity of this sample.

The crystallites size of the as prepared films have been estimated from the main (111) Bragg peak of cubic Cu_2O using:³⁶

$$D = \frac{0.94\lambda}{\beta \cos \theta} \quad (4)$$

where λ is the wavelength of $\text{Cu } K\alpha_1$ radiation ($\lambda = 1.54056 \text{ \AA}$), β is the full width at half maximum and θ is the diffraction angle. The values of average vertical coherence length (crystallites size) of Cu_2O layers are 41.9, 44, 52.7 and 58.4 nm corresponding the deposition temperature 40, 50, 60 and 70°C of ZnS layer, respectively. Therefore, the different crystallites size contains different dislocation defects, which have a great effect on the electronic property of the heterojunction.³⁷

3.6 Optical characterization

The optical properties of our heterojunction are very important for many of their technological applications. The absorbance spectra of $\text{ZnO}/\text{ZnS}/\text{ITO}$ heterostructure and $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions with different deposited temperatures of ZnS layer were recorded for wavelength range from 300 to 1000 nm as shown in Fig. 8. From the inset of Fig. 8, the $\text{ZnO}/\text{ZnS}/\text{ITO}$ layer shows strong absorption in the UV region extends from 300 to 450 nm. The absorption was significantly improved after the electrodeposition of Cu_2O on $\text{ZnO}/\text{ZnS}/\text{ITO}$ layer. As can be seen from Fig. 8, the absorption edge of the $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions shifted into the visible region (red shift). This amelioration will ensure good performance to the photovoltaic device. Comparing the absorption of the heterojunctions it is clear that $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ with ZnS layer deposited at 60°C shows the highest absorbance, which is in accordance with the strong photocurrent generated by this sample under AM 1.5 illumination, as observed in the photoelectrochemical analysis (Fig. 4b).

The optical band gap energies can be determined plotting $(\alpha h\nu)^2$ as a function of photon energy $h\nu$ using the Tauc relation:³⁸

$$(\alpha h\nu) = A(h\nu - E_g)^n \quad (5)$$

where α is the absorption coefficient and $h\nu$ is the photon energy. As ZnO is semiconductor with a direct band gap, we use $n = 1/2$ as a characteristic of the direct band gap absorption without phonon assistance. Therefore, the optical band gap is obtained from Tauc's plot by extrapolating the straight lines downwards to the axis of photon energy.³⁹

The Tauc plots of $\text{ZnO}/\text{ZnS}/\text{ITO}$ and $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions are presented in Fig. 9. The band gap obtained for $\text{ZnO}/\text{ZnS}/\text{ITO}$ was 3.67 eV (inset of Fig. 9). Lower band gap energies were determined for $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions due to the narrow band gap of Cu_2O . As can be seen from Fig. 9, the estimated band gap was between 2.01 and 2.36 eV. These values are intermediate, lower than that of ZnS (approximately 3.7 eV) and larger than that of Cu_2O (approximately 2.17 eV). The variation of $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ energy band gap with the deposition temperature of ZnS layer may be explained by the crystallinity deterioration with increasing ZnS deposition temperature.

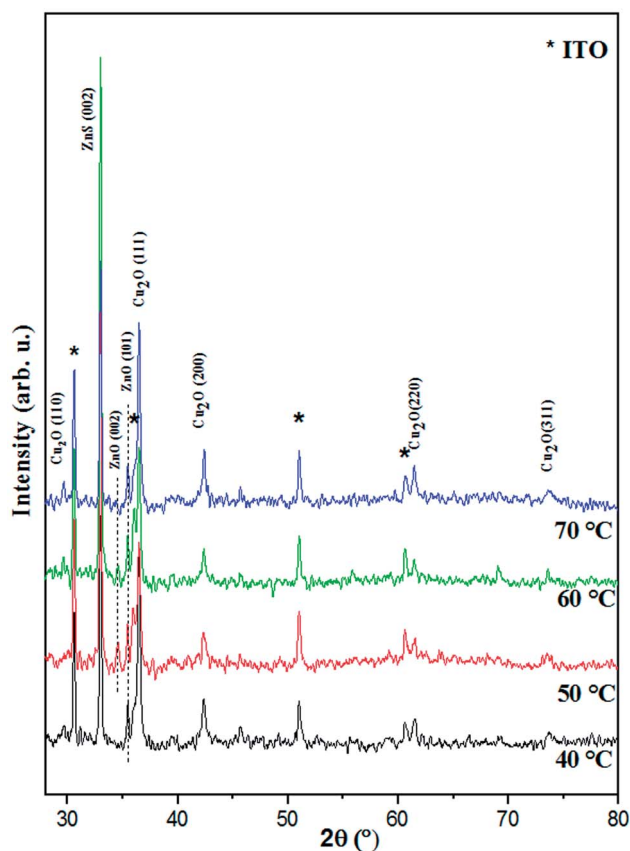


Fig. 7 XRD patterns of $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunction with different electrodeposition temperatures of ZnS layer.



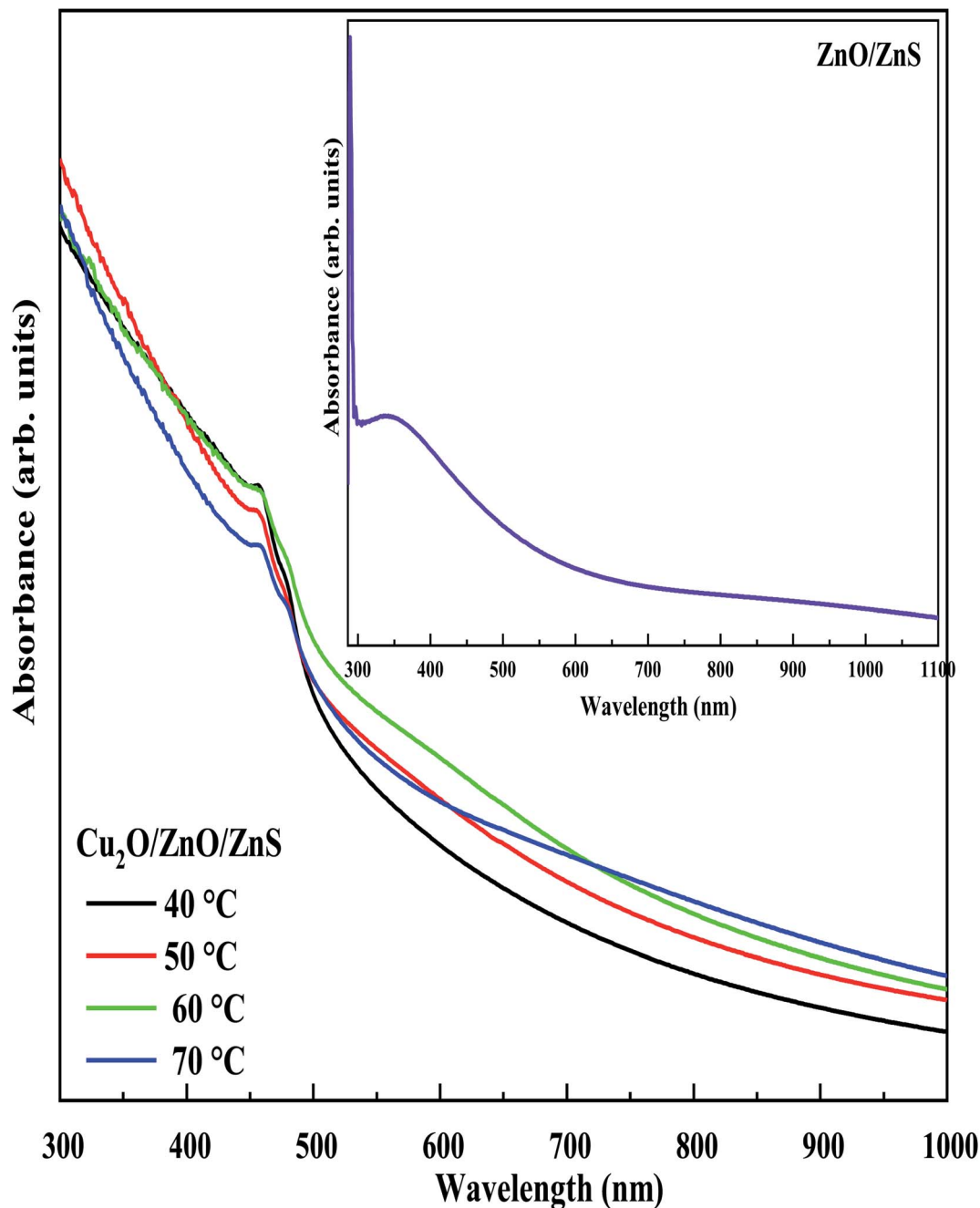


Fig. 8 UV-Vis absorption spectra of Cu₂O/ZnO/ZnS/ITO heterojunction with different electrodeposition temperatures of ZnS layer. Inset: the UV-Vis absorption spectra of ZnO/ZnS/ITO sample.

3.7 Current–voltage measurements

The *I*–*V* curves of Au/Cu₂O/ZnO/ZnS/ITO heterojunctions, with different deposition temperatures of ZnS layer, measured at room temperature in the dark are shown in Fig. 10. This method based on the quantification of the current crossing a junction under the effect of an external electric field. As can be seen from Fig. 10a, Cu₂O/ZnO/ZnS/ITO heterojunction with ZnS deposited at 40, 50 and 70 °C shows an excellent ohmic response. However, good rectifying behavior with a low leakage current of 5×10^{-7} A at

a reverse bias of -0.5 V was obtained for the Cu₂O/ZnO/ZnS/ITO with ZnS (60 °C), as presented in Fig. 10b, which confirms the formation of p–n junction for this sample. This may be due to the high crystallinity confirmed by XRD.

A schematic diagram of the energy-band structure of the p-Cu₂O/n-ZnO/n-ZnS heterojunction was constructed simulated using the model proposed by Anderson⁴⁰ is shown in Fig. 11. When the p- and the n-type semiconductors are in contact, they form a p–n junction with a space-charge region at the interfaces due to the diffusion of electrons and holes



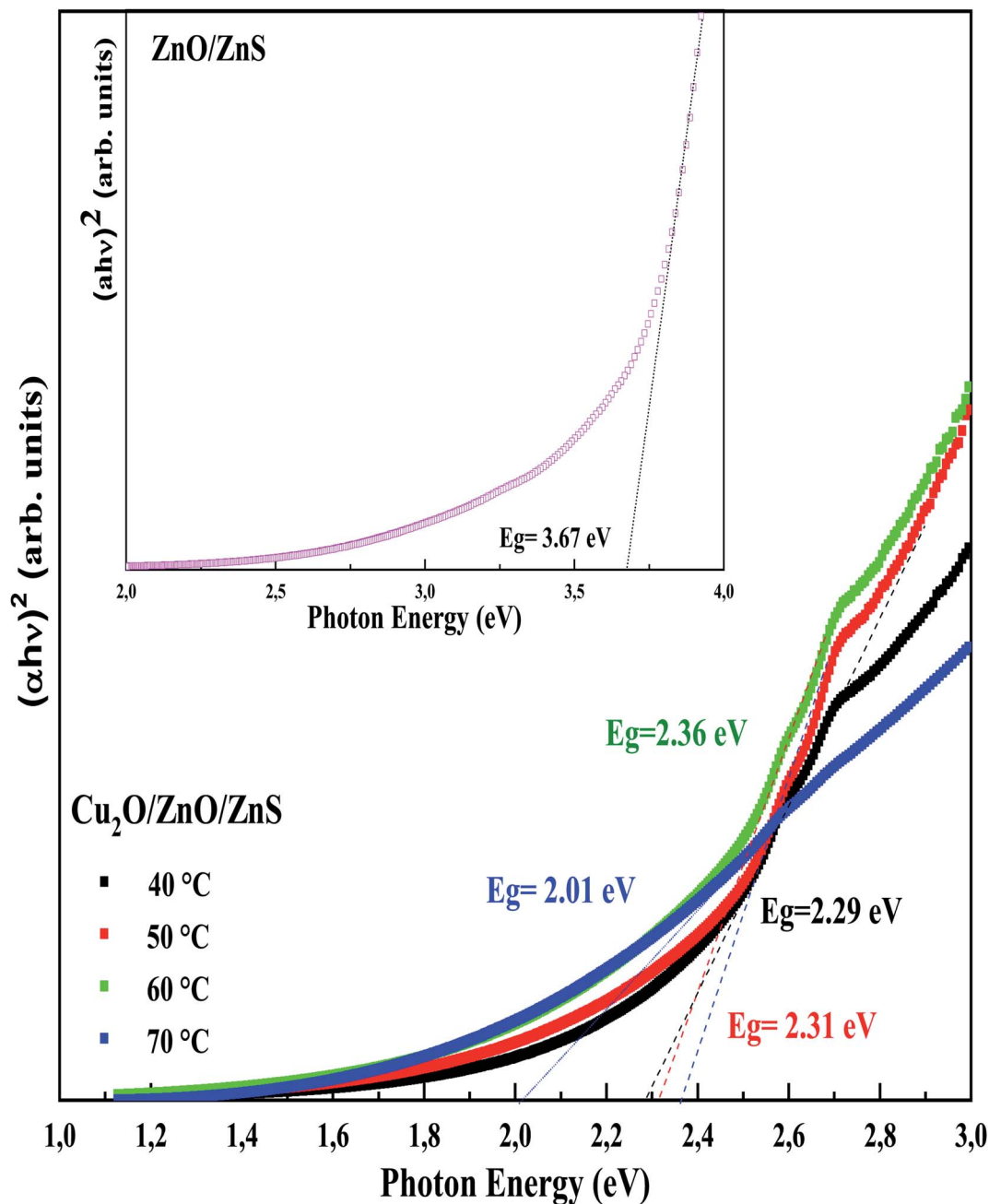


Fig. 9 Tauc's plot of $\text{Cu}_2\text{O}/\text{ZnO}/\text{ZnS}/\text{ITO}$ heterojunctions with different electrodeposition temperatures of ZnS layer. Inset: Tauc's plot of $\text{ZnO}/\text{ZnS}/\text{ITO}$ sample.

driven by the electric field.⁴¹ The electrons are transferred to the conduction band of the n-type semiconductors and the holes to the valence band of the p-type semiconductors. The conduction band offset (ΔE_c) and the valence band offset (ΔE_v) of the heterojunction are represented by the following equations:

$$\Delta E_c = \chi_1 - \quad (\chi_2 \quad 6)$$

$$\Delta E_v = E_{g_2} - E_{g_1} + \Delta E_c \quad (7)$$

where the electron affinities (χ) are assumed to be 3.9, 4.2 and 3.2 eV for ZnS, ZnO and Cu_2O semiconductors, respectively.^{42–45} The measured band gap energies (E_g) were 3.84, 3.38 and 2.38 eV for ZnS, ZnO and Cu_2O layers, respectively. The Fermi level energies (E_f) were calculated from the flat band potential (E_{fb}) using the following equations:²⁷

$$E_{fb} \text{ (in V vs. NHE)} = -E_f \text{ (in eV vs. vacuum)} + V_H \text{ (in V)} - 4.5(8)$$

$$V_H \text{ (in V)} = 0.059(\text{pH}_{\text{PZZP}} - \text{pH}) \quad (9)$$



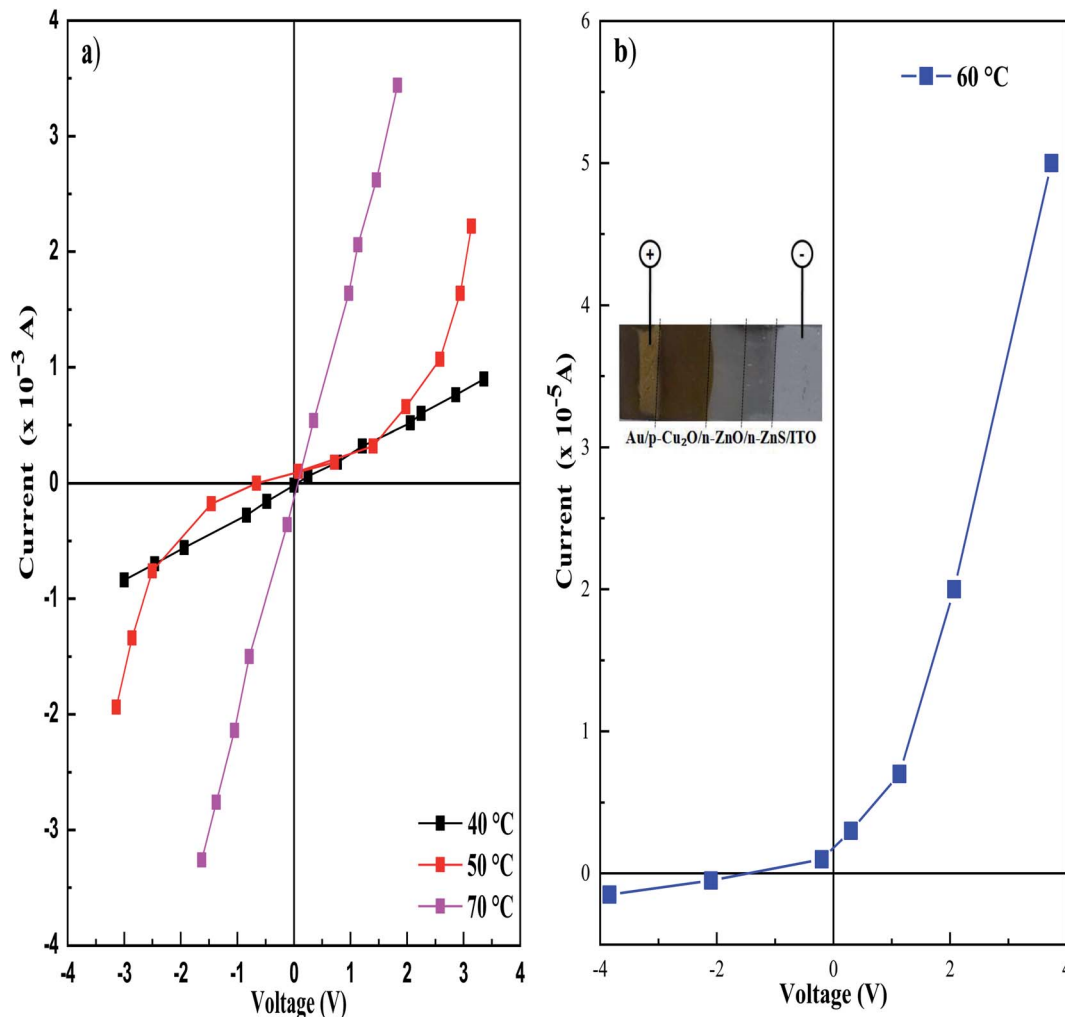


Fig. 10 Current–voltage characteristics of Au/p-Cu₂O/n-ZnO/n-ZnS/ITO heterojunction in the dark with different electrodeposition temperature of ZnS layer. Inset of (b) photograph of the p-Cu₂O/n-ZnO/n-ZnS/ITO heterojunction with Au top contact.

where E_{fb} is -0.46 , -0.52 and 0.36 V per NHE for ZnS, ZnO and Cu₂O, respectively. V_H is the potential drop across the Helmholtz layer. The pH_{PZZP} is the point of zero zeta potential of 1.7, 8.8 and 8.5 for ZnS, ZnO and Cu₂O, respectively.⁴⁵

The estimated Fermi level energies for n-ZnS, n-ZnO and p-Cu₂O were -4.29 , -3.82 eV and -4.72 eV vs. the vacuum, respectively.

Based on the obtained χ , E_g , E_f we proposed an approximate energy-band diagram of the three materials p-type Cu₂O, n-type ZnO and n-ZnS before and after the junction formation (Fig. 11). Putting the three materials in contact creates a p–n junction at their interface. The electrons in the n-ZnS layer transferred to the p-Cu₂O layer through the n-ZnO layer. In one hand, ZnS has a higher conduction band edge with respect to ZnO, which is advantageous in the injection of excited electrons from ZnS.⁴⁶ In the other hand, the holes of Cu₂O layer transferred in the opposite direction until a constant Fermi-level formed at equilibrium. Hence, creates a built-in electrical potential that can direct the

electrons and holes to travel in the opposite direction. The calculated conduction band offset (ΔE_C) of the heterojunction Cu₂O/ZnO was 1.0 eV, and the valence band offset (ΔE_V) was 2.0 eV. The small conduction band offset should result in easy way for electron injection from the n-type ZnO to the p-type Cu₂O. In addition, the large valence band offset should result in a thin tunneling barrier across the p–n junction.⁴⁷

As illustrated in Fig. 11, the built-in potential V_{bi1} (-0.9 V) is equal to the difference between the Fermi levels of Cu₂O and ZnO. Consequently, the p-Cu₂O/n-ZnO/n-ZnS heterojunction, with ZnS deposited at 60 °C, possesses larger built-in potential compared to the other samples. Therefore, p-Cu₂O/n-ZnO/n-ZnS (60 °C) got the highest charge separation efficiency, which results in good rectifying behavior in the dark. Moreover, this explains the high photoresponse of this sample reported in the photoelectrochemical analysis (Fig. 4b).



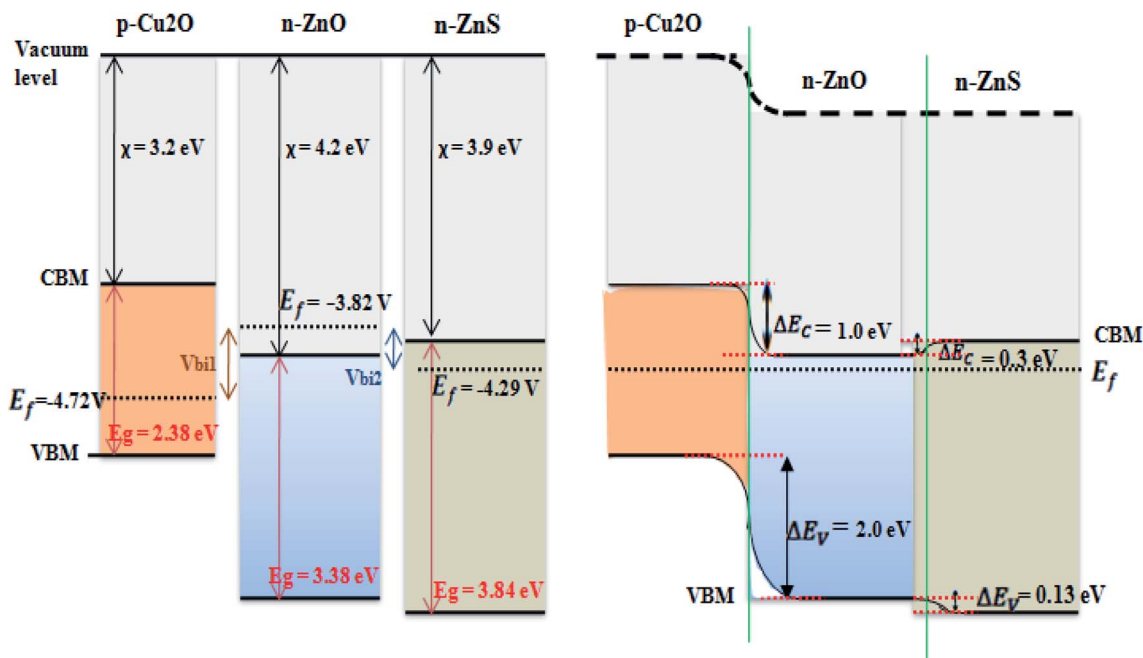


Fig. 11 Band energy diagram of the isolated states of the n-ZnO, n-ZnS, p-Cu₂O and p-Cu₂O/n-ZnO/n-ZnS heterojunction under equilibrium condition.

4. Conclusion

In this work, the incorporated n-ZnS layer in p-Cu₂O/n-ZnO heterojunction was successfully prepared by electrodeposition on ITO substrate. Where, the deposition temperature of ZnS layer was varied between 40 and 70 °C. It was found that the Cu₂O properties were mainly affected by the ZnS deposition temperature. Thus, high crystallinity and large grains were confirmed for the p-Cu₂O/n-ZnO/n-ZnS heterojunction with ZnS deposited at 60 °C by XRD analysis and AFM images. In addition, high electron-hole separation efficiency and high photoresponse were confirmed for this sample according to the photoelectrochemical measurements. The UV-Vis spectra show strong absorption in the visible region, which ensures, for this photovoltaic device, a good consumption of the sunlight. The Au/p-Cu₂O/n-ZnO/n-ZnS (60 °C)/ITO heterojunction revealed a well-defined electrical rectification characteristic according to the current-voltage (*I*-*V*) characteristics. This study demonstrates that electrodeposition, which is easily adapted to other chemical systems, can be a promising technique for large-scale fabrication of low-cost heterojunctions solar cells.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 M. Grätzel, J. Photochem and C. Photobiol, *Photochem. Rev.*, 2003, **4**, 145–153.
- 2 D. C. Olson, J. Piris, R. T. Collins, S. E. Shaheen and D. S. Ginley, *Thin Solid Films*, 2006, **496**, 26–29.
- 3 M. E. Abu-Zeid, A. E. Rakhshani, A. A. Aljassar and Y. A. Youssef, *Phys. Status Solidi A*, 1986, **93**, 613.
- 4 W. Siripala, L. D. R. D. Perera, K. T. L. De Silva, J. K. D. S. Jayanetti and I. M. Dharmadasa, *Sol. Energy Mater. Sol. Cells*, 1996, **44**, 251–260.
- 5 A. V. Rokade, S. R. Rondiya, A. A. Jadhavar, S. M. Pandharkar, S. D. Karpe, K. D. Diwate and S. R. Jadhkar, *AIP Conf. Proc.*, 2016, **1724**, 1–7.
- 6 Y. Hameş and S. E. San, *Sol. Energy*, 2004, **77**, 291–294.
- 7 C. J. Dong, W. X. Yu, M. Xu, J. J. Cao, C. Chen, W. W. Yu and Y. D. Wang, *J. Appl. Phys.*, 2011, **110**, 073712.
- 8 Y. Levskaia, R. L. Z. Hoyer, A. Sadhanala, K. P. Musselman and J. L. MacManus-Driscoll, *Sol. Energy Mater. Sol. Cells*, 2015, **135**, 43–48.
- 9 T. Minami, Y. Nishi and T. Miyata, *Appl. Phys. Express*, 2016, **9**, 052301.
- 10 F. M. A. Hernández, M. C. López, V. Donderis, M. González, B. Mar and J. R. Ramos-Barrad, *Thin Solid Films*, 2008, **516**, 1622–1625.
- 11 S. M. Al-Hilli, R. T. Al-Mofarji, P. Klason, M. Willander, N. Gutman and A. J. Sa'ar, *J. Appl. Phys.*, 2008, **103**, 014302.



- 12 M. Sookhakian, Y. M. Amin, W. J. Basirun, M. T. Tajabadi and N. Kamarulzaman, *J. Lumin.*, 2014, **145**, 244–252.
- 13 H. A. Cheol, Y. K. Young, C. K. Dong, K. M. Sanjay and K. C. Hyung, *J. Appl. Phys.*, 2009, **105**, 013502.
- 14 F. Xiaosheng, Z. Tianyou, K. G. Ujjal, L. Liang, W. Limin, B. Yoshio and G. Dmitri, *Prog. Mater. Sci.*, 2011, **56**, 175–287.
- 15 Ü. Neslihan and A. M. Çetin, *J. Mater. Sci.: Mater. Electron.*, 2010, **34**, 287–292.
- 16 J. Jan, X. Fan, L. Zhang, Y. Bando, U. K. Gautam, B. Dierre, T. Sekiguchi and D. Goldberg, *Nano Lett.*, 2008, **8**, 2794–2799.
- 17 W. Zhiqiang, L. Xuedong, G. Jiangfeng, H. Hongbo, G. Shulin and Y. Shaoguang, *Cryst. Growth Des.*, 2008, **8**, 3911–3913.
- 18 M. W. Murphy, X. T. Zhou, J. Y. P. Ko, J. G. Zhou, F. Heigl and T. K. Sham, *J. Chem. Phys.*, 2009, **130**, 084707.
- 19 S. Simmi and C. Santa, *Electron. Mater. Lett.*, 2013, **9**, 267–271.
- 20 J. Schrier, D. O. Demchenko and L. W. Wang, *Nano Lett.*, 2007, **7**, 2377–2382.
- 21 L. Mentar, O. Baka, M. R. Khelladi, A. Azizi, S. Velumani, G. Schmerber and A. Dinia, *J. Mater. Sci.: Mater. Electron.*, 2015, **26**, 1217–1224.
- 22 S. Laidoudi, A. Y. Bioud, A. Azizi, G. Schmerber, J. Bartringer, S. Barre and A. Dinia, *Semicond. Sci. Technol.*, 2013, **28**, 115005–115011.
- 23 Y. F. Cheng, C. Yang and J. L. Luo, *Thin Solid Films*, 2002, **416**, 169–173.
- 24 A. W. Bott, *Electrochemistry of semiconductors*, *Curr. Sep.*, 1998, **17**, 87–91.
- 25 H. Noh, M. Scharrer, M. A. Anderson, R. P. H. Chang and H. Cao, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2008, **77**, 115136.
- 26 Z. Zhang and P. Wang, *J. Mater. Chem.*, 2012, **22**, 2456–2464.
- 27 C. M. McShane and K.-S. Choi, *J. Phys. Chem.*, 2012, **14**, 6112.
- 28 F. Caballero-Briones, J. M. Artes, I. Diez-Perez, P. Gorostiza and F. Sanz, *J. Phys. Chem. C*, 2009, **113**, 1028.
- 29 C. Yilmaz and U. Unal, *Appl. Surf. Sci.*, 2016, **368**, 456–463.
- 30 R. Ghosh, D. Basak and S. Fujihara, *J. Appl. Phys.*, 2004, **96**, 2689–2692.
- 31 Y. Yang, J. Han, X. Ning, W. Cao, W. Xu and L. Guo, *ACS Appl. Mater. Interfaces*, 2014, **6**, 22534–22543.
- 32 Z. Kang, X. Yan, Y. Wang, Z. Bai, Y. Liu, Z. Zhang, P. Lin, X. Zhang, H. Yuan, X. Zhang and Y. Zhang, *Sci. Rep.*, 2015, **5**, 78–82.
- 33 Z. Bai and Y. Zhang, *J. Alloys Compd.*, 2017, **698**, 133–140.
- 34 R. Kern, R. Sastrawan, J. Ferber, R. Stangl and J. Luther, *Electrochim. Acta*, 2002, **47**, 4213–4225.
- 35 M. Guo, K. Xie, J. Lin, Z. Yong, C. T. Yip, L. Zhou, Y. Wang and H. Huang, *Energy Environ. Sci.*, 2012, **5**, 9881–9888.
- 36 B. D. Cullity, *Elements of X-ray Diffraction*, Addison Wesley, Reading, MA, 2nd edn, 1978.
- 37 D. K. Zhang, Y. C. Liu, Y. L. Liu and H. Yang, *Phys. B*, 2004, **351**, 178.
- 38 H. Lahmar, A. Azizi, G. Schmerber and A. Dinia, *RSC Adv.*, 2016, **6**, 68663–68674.
- 39 J. Tauc, in *Optical properties of solids*, ed. F. Abeles, North-Holland Pub Co., Amsterdam, vol. 22, 1970.
- 40 A. G. Milnes and D. L. Feucht, *Heterojunctions and Metal-Semiconductor Junctions*, Academic Press, New-York, 1972.
- 41 L. Jiang, G. Zhou, J. Mi and Z. Wu, *Catal. Commun.*, 2012, **24**, 48–51.
- 42 R. K. Swank, *Phys. Rev.*, 1967, **153**, 844.
- 43 W. Siripala, A. Ivanovskaya, T. F. Jaramillo, S. H. Baeck and E. W. McFarland, *Sol. Energy Mater. Sol. Cells*, 2003, **77**, 229–237.
- 44 H. Kobayashi, H. Mori, T. Ishida and Y. Nakato, *J. Appl. Phys.*, 1995, **77**, 1301–1307.
- 45 Y. Xu and M. A. A. Schoonen, *Am. Mineral.*, 2000, **85**, 543–5565.
- 46 L. P. Wu, Y. L. Zhang, L. Z. Long, C. P. Cen and X. J. Li, *RSC Adv.*, 2014, **4**, 20716–20721.
- 47 T. J. Hsueh, C. L. Hsu, S. J. Chang, P. W. Guo, J. H. Hsieh and I. C. Chen, *Scr. Mater.*, 2007, **57**, 53–56.

