



Cite this: *J. Mater. Chem. C*, 2018, 6, 12652

Synthesis of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) solid solutions from solventless thermolysis of metal xanthate precursors†

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Tris(*O*-ethylxanthate)bismuth(III) (**1**) and tris(*O*-ethylxanthate)antimony(III) (**2**) were synthesized and used as precursors for the preparation of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ alloys ($x = 0, 0.2, 0.4, 0.6, 0.8$ and 1) using solventless thermolysis. The p-XRD peaks at all ratios corresponded well to an orthorhombic crystal structure between that of Bi_2S_3 and Sb_2S_3 . The lattice parameters a and b were found to increase linearly as the Sb content increases in the $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ samples, whereas c decreases gradually with increasing Sb. For all samples the elemental composition and uniform stoichiometry was determined by EDX analysis and ICP-OES. Incorporation of Sb into Bi_2S_3 led to an increase in the band gap from 1.66 eV at $x = 0$ to 2.19 eV at $x = 1$. The morphologies of the alloys were investigated using SEM, revealing different texturing as the composition changes from Bi_2S_3 to Sb_2S_3 .

Received 16th May 2018,
Accepted 18th September 2018

DOI: 10.1039/c8tc02374c

rsc.li/materials-c

Introduction

The family of V–VI semiconductor materials (M_2X_3 , $\text{M} = \text{As}, \text{Sb}, \text{Bi}$; $\text{X} = \text{S}, \text{Se}, \text{or Te}$) have received widespread attention in previous decades due to their semiconducting properties, high refractive indices, good electrical conductivity and potential use in optoelectronic^{1–5} and thermoelectric applications.^{6–10}

Antimony sulfide (Sb_2S_3) is an earth abundant useful V–VI binary chalcogenide. It has an orthorhombic crystal structure with a melting point of 650 °C. Sb_2S_3 has a direct band gap, which ranges between 1.5 to 2.5 eV, depending on the shape, size, and crystallinity of the nanostructure. Moreover, it has exceptional semiconducting and photoconductivity properties that allows potential applications in solar energy conversion, optoelectronics, and thermoelectric devices.^{2,11,12} Bismuth sulfide (Bi_2S_3) is another V–VI semiconductor with orthorhombic crystal structure. It has attracted considerable attention largely owing to its direct band gap, which lies between 1.2 to 1.7 eV, and high absorption coefficient (in the order of 10^4 to 10^5 cm^{-1}).

Bi_2S_3 nanoparticles and thin films have found various applications in gas sensors, thermoelectric, photovoltaic, electronic and optoelectronic devices.^{13–17} Indeed, the crystal structures of both Sb_2S_3 and Bi_2S_3 are similar, with orthorhombic unit cell dimensions differing by a maximum of 3.5%,¹⁸ which suggests the potential for complete solid solubility between the two phases (Fig. 1).^{18,19} The Bi_2S_3 – Sb_2S_3 mineral series are found in hydrothermal ore deposits. Discrete compositional variations occur along the path of flow of the mineralising fluids, with the group extremes being typified by a prevalence of mole ratios of $\text{Bi}/(\text{Sb}+\text{Bi})$ and a notable lack of intermediate compositions.^{19,20} These contrasts in the molar ratios of $\text{Bi}/(\text{Sb}+\text{Bi})$ have been ascribed to the operation of so-called semi-metal boiling, during which a sudden alteration in the $\text{Sb}_2\text{S}_3/\text{Bi}_2\text{S}_3$ (Sb/Bi) ratio arises at a specific point along the flow path due to preferential volatilisation and loss of antimony from the ore fluid.^{20–22}

Numerous approaches have been employed for the preparation of Sb_2S_3 and Bi_2S_3 nanoparticles and thin films including solvothermal,^{23,24} hydrothermal,^{25,26} single-source routes,^{27–30} colloidal methods,^{16,31,32} melt thermolysis,³³ dip-dry techniques,³⁴ chemical bath deposition,^{35,36} aerosol-assisted chemical vapour deposition (AA-CVD),³⁷ spray pyrolysis,^{38,39} and successive ionic layer adsorption reactions (SILAR) have also been used.^{40,41}

Solventless thermolysis has advantages over other routes, as it is a simple method in which solid state decomposition of a precursor is achieved by thermal treatment under inert conditions. This approach has proven to be an effective way of producing metal chalcogenide nanomaterials with a wide range

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† Electronic supplementary information (ESI) available: p-XRD patterns of Bi_2S_3 and Sb_2S_3 materials and absorption spectra of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$. See DOI: 10.1039/c8tc02374c

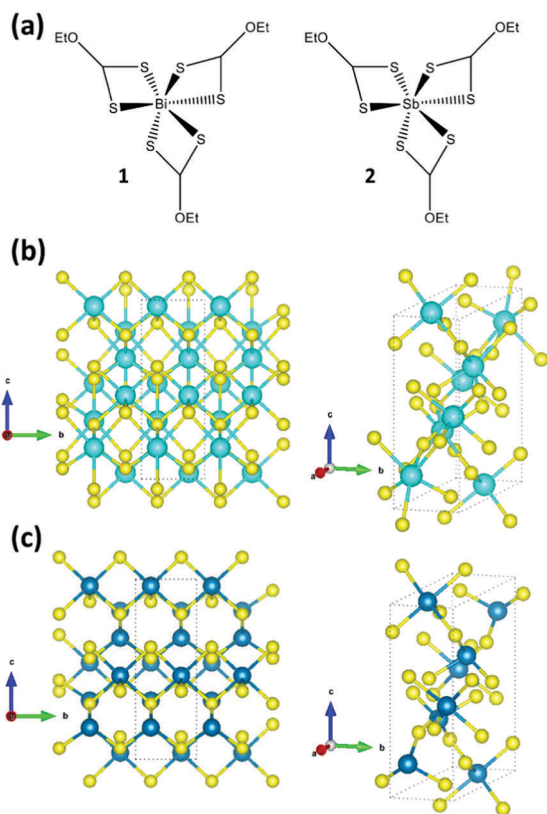


Fig. 1 (a) Chemical structure of the precursors used in this study: $\text{Bi}[\text{S}_2\text{COEt}]_3$ (1), and $\text{Sb}[\text{S}_2\text{COEt}]_3$ (2). (b) Structure of Bi_2S_3 with the orthorhombic unit cell marked ($a = 11.316 \text{ \AA}$, $b = 3.971 \text{ \AA}$, $c = 11.178 \text{ \AA}$ with $\alpha = \beta = \gamma = 90^\circ$) and (c) structure of Sb_2S_3 with the orthorhombic unit cell marked ($a = 11.319 \text{ \AA}$, $b = 3.833 \text{ \AA}$, $c = 11.233 \text{ \AA}$ with $\alpha = \beta = \gamma = 90^\circ$), after Kyono and Kimata.²² Blue atoms represent Bi/Sb and the yellow atoms are S.

of morphologies such as nanorods,⁴² nanowires,⁴³ nanospheres,⁴⁴ and nanodisks.⁴⁵ Compared to the other solution based chemical approaches melt thermolysis offers a simple and cost-effective way to scale up production. Furthermore, it offers economic and environmental benefits, eliminates the need for harsh reactants, and typically yields are usually high.⁴⁶

The use of single-source precursors offers significant advantages over other methods for the preparation of a number of semiconducting materials. This approach is simple, can produce high-quality materials with better control over composition, the potential for low-growth temperatures, and is highly suitable for mass production.⁴⁷ Indeed, great success has been achieved using the thermal decomposition of antimony/bismuth complexes of dithiocarbamates,^{27,37,48–53} dithiophosphates,^{27,30,54} and thiourea,^{55,56} as single-source precursors. The use of metal xanthate precursors for the preparation of Sb_2S_3 ^{27,57,58} and Bi_2S_3 ^{59–61} is promising owing to the low decomposition temperature of metal xanthate complexes (100–200 °C) and clean abstraction of volatile organic material from the crystal matrix by Chugaev elimination during thermolysis.^{59,62–64} For example, it has been shown that it is preferable to use the xanthate precursors in order to produce oxide-free metal sulfide products

compared to the dithiocarbamate complexes.²⁷ Despite considerable efforts devoted to the study of Sb_2S_3 and Bi_2S_3 , only a little information on the preparation of the mixed bismuth antimony sulfide system is available in the literature.^{65,66}

In this work, we investigate the use of $\text{Bi}[\text{S}_2\text{COEt}]_3$ (1) and $\text{Sb}[\text{S}_2\text{COEt}]_3$ (2) as single source precursors for the solventless synthesis of Bi_2S_3 , Sb_2S_3 and $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) via a facile melt method. The thermogravimetric analysis (TGA) reveals that both precursors exhibit complete decomposition in a similar temperature range. The materials produced are investigated in detail using powder X-ray diffraction (p-XRD), scanning electron microscopy (SEM), energy dispersive X-ray spectroscopy (EDX), inductively coupled plasma optical emission spectrometer (ICP-OES) analysis, Raman spectroscopy and UV-vis spectroscopy.

Experimental

Chemicals

Potassium ethyl xanthogenate (96%, Sigma-Aldrich), carbon disulphide (low benzene $\geq 99.9\%$, Sigma-Aldrich), potassium hydroxide ($> 85\%$, Fisher Scientific), methanol ($\geq 99.8\%$, Sigma-Aldrich), ethanol ($\geq 99.8\%$, Sigma-Aldrich), chloroform (anhydrous, $\geq 99\%$, Sigma-Aldrich), antimony(III) chloride ($> 99\%$, Fisher Scientific), and bismuth(III) chloride (98%, Sigma-Aldrich) were used with no further purification.

Instrumentation

Elemental analysis of the precursors was carried out by the chemistry microanalysis laboratory at the University of Manchester. TGA was conducted from 25 °C to 500 °C under nitrogen using Mettler Toledo TGA/DCS system. Fourier transform infrared (FTIR) spectra were obtained using a Specac single reflectance ATR and melting points were obtained using a Barloworld SMP10 apparatus. Nuclear magnetic resonance (NMR) spectra were obtained using a 400 MHz Bruker instrument. Powder-X-ray diffraction (p-XRD) of all the samples was carried out using a Bruker Xpert diffractometer. The samples were scanned between 20° and 80° using $\text{CuK}\alpha$ radiation with a step size of 0.02°. A Philips XL30 FEG scanning electron microscope was used to obtain the scanning electron microscopy images for all the samples. EDX spectra were obtained using a DX4 spectrometer attached to the same microscope. All the samples were carbon coated using an Edwards E306A coating instrument before SEM and EDX studies. Transmission electron microscope (TEM) imaging and diffraction was performed using Tecnai F30 TEM operated at 300 kV.

Synthesis of bismuth(III) ethylxanthate, $\text{Bi}[\text{S}_2\text{COEt}]_3$ (1)

Potassium ethylxanthate (4.0 g, 24.9 mmol) was dissolved in deionised water (40 mL), followed by the drop-wise addition of bismuth(III) chloride (2.62 g, 8.30 mmol) solution in deionised water (25 mL) under constant stirring at room temperature. A dark yellow precipitate was formed immediately, and the stirring was continued for thirty minutes. The product was



filtered under vacuum and recrystallized from chloroform. Yield (3.8 g, 80%), m.p. 110 °C, elemental analysis: found (%): C, 19.14; H, 2.65; S, 33.46; Bi, 36.47; calc. (%) C, 18.86; H, 2.64; S, 33.53; Bi, 36.49. IR (cm⁻¹): 1206 ν (C–O); 1018 ν (C–S). ¹H NMR (CDCl₃): δ 4.69 (q, J = 7.1 Hz, 2H, CH₂), δ 1.49 (t, J = 7.1 Hz, 3H, CH₃). ¹³C NMR: δ 206.99 (S₂C), δ 71.3 (CH₂), δ 13.9 (CH₃).

Synthesis of antimony(III) ethylxanthate, Sb[S₂COEt]₃ (2)

Antimony(III) chloride (1.9 g, 8.3 mmol) was dissolved in 25 mL of methanol and added drop-wise to a solution of potassium ethylxanthate (4 g, 24.9 mmol) in methanol (100 mL) with constant stirring. The reaction mixture was stirred at room temperature for approximately thirty minutes to ensure the completion of the reaction. The precipitates formed were filtered and dried under vacuum, resulting in a bright yellow solid. The product was re-crystallised by dissolving the precipitates in chloroform and removing the insoluble impurities by filtration. A clear yellow solution was obtained which on evaporation yielded the pure product. Yield: (3.2 g) 79%, m.p. 90.3 °C, elemental analysis: found (%): C, 22.61; H, 3.07; S, 39.71; Sb, 24.15; calc. (%) C, 22.25; H, 3.11; S, 39.55; Sb, 25.08. IR (cm⁻¹): 1220 ν (C–O); 1020 ν (C–S). ¹H-NMR (CDCl₃): δ 4.61 (q, J = 7.1 Hz, 2H, CH₂), δ 1.44 (t, J = 7.1, 3H, CH₃). ¹³C NMR: δ 207.02 (S₂C), δ 72.07 (CH₂), δ 13.9 (CH₃).

Synthesis of bismuth–antimony sulfide system

Bismuth–antimony sulfide Bi_{2–2x}Sb_{2x}S₃ with different mole fractions of Sb (x = 0, 0.2, 0.4, 0.6, 0.8 and 1) were prepared using the melt method. Precursors (1) or (2) were used for the synthesis of Bi₂S₃ and Sb₂S₃ materials. Approximately 0.2 g of the solid precursor was placed into a ceramic boat in a tube furnace and annealed at 200, 250, and 300 °C under nitrogen for one hour. For the mixed samples, Bi[S₂COEt]₃ (1) and Sb[S₂COEt]₃ (2) were mixed in different molar ratios (ESI† Table S1) and dissolved in small amount of chloroform to ensure uniform mixing of the precursors. The solid mixture obtained after evaporation of chloroform was then placed into a ceramic boat inside a reaction tube in a furnace and annealed at 250 °C under nitrogen flow for one hour. The obtained black residue was then characterised by p-XRD, EDX, ICP-OES and SEM.

Results and discussion

TGA of the Bi[S₂COEt]₃ (1) and Sb[S₂COEt]₃ (2) precursors was conducted from 25 °C to 500 °C under nitrogen. The Bi[S₂COEt]₃ (1) exhibits one-step decomposition with rapid mass loss of 54% between 120 and 160 °C. The total residual mass (46%) is in good agreement with the calculated value for Bi₂S₃ (45%). Similarly, the TGA for (2) shows major-step decomposition in the temperature range of 100–160 °C. The minor weight loss of approximately 2% can be ascribed to the removal of sulfur (ESI† Fig. S1). The final observed residual mass (35%) matches the calculated value for Sb₂S₃ (35%).

The thermal decomposition of (1) or (2) performed at temperatures 200, 250, and 300 °C, under inert conditions,

for one hour gave a black residue which was analysed by p-XRD. For pure Bi₂S₃, the p-XRD of the powder (ESI† Fig. S2a) prepared from (1) matched very well with the orthorhombic phase of Bi₂S₃ (bismuthinite, ICDD: 01-089-8965),⁶⁷ with no indication of impurities or secondary phases. The powder prepared at 300 °C showed relatively narrower and sharper peaks than those prepared at lower temperatures. Such sharp narrow p-XRD peaks give an indication of the high crystallinity of the product. For pure Sb₂S₃, the p-XRD patterns obtained from decomposition of (2) at 200, 250, and 300 °C are shown in ESI† Fig. S2b. All diffraction peaks fully matched with the orthorhombic phase of Sb₂S₃ (stibnite, ICDD: 01-075-1310).⁶⁸ Antimony ethylxanthate was also decomposed at 250 °C for two hours to observe the effect of time but it was found that this additional time did not have any significant effect on the crystallinity or phase as indicated by p-XRD. The p-XRD was quite similar to the pattern obtained at the same temperature for one hour with no extra peaks from impurities or secondary phases.

SEM-EDX was performed on the powder obtained at 250 °C to determine the composition of the product (ESI† Fig. S3). For Bi₂S₃, it is clearly seen that no additional peaks of other elements except Bi and S are observed in the spectrum. The atomic ratio of Bi to S is found to be 43.6:56.4 (Bi_{2.18}S_{2.82}), indicating that the Bi₂S₃ powder is slightly rich in bismuth. Similarly, for Sb₂S₃, the atomic ratios of Sb to S was found to be 40.6:59.4 corresponding to Sb_{2.03}S_{2.97}, which is close to the expected value of Sb₂S₃ (40:60).

Bismuth–antimony–sulfide (Bi–Sb–S) system

The orthorhombic crystal structures of both Bi₂S₃ and Sb₂S₃ exist with comparable lattice constants.¹⁸ Hence the formation of the Bi–Sb–S system is feasible, over entire range of compositions, as both antimony and bismuth can be used interchangeably in the crystal lattice. Moreover, the decomposition temperature range of bismuth(III) ethylxanthate and antimony(III) ethylxanthate are similar, so the mixture of these precursors can be used to produce solid solutions of Bi–Sb–S between the two end phases. A series with varying stoichiometry was prepared by mixing (1) and (2) in specific molar ratios *i.e.* Sb/Sb+Bi = 0.2, 0.4, 0.6 and 0.8. The mixture was then annealed at 250 °C under nitrogen for one hour. A black residue was obtained in all the pyrolysis experiments.

p-XRD patterns of Bi_{2–2x}Sb_{2x}S₃ ($0 \leq x \leq 1$) with different mole fractions of Sb are presented in Fig. 2a. The p-XRD peak positions fall in between those of orthorhombic Bi₂S₃ (ICDD: 01-089-8965) and orthorhombic Sb₂S₃ (ICDD: 01-075-1310).^{67,68} The diffraction peaks at 23.6°, 24.9°, 31.7°, 32.8°, 39.8°, 46.5° and 52.5° gradually shift to higher values of 24.5°, 25.8°, 32.3°, 33.4°, 40.4°, 46.8° and 52.9°, respectively, as the composition changes from Bi₂S₃ (x = 0) to Sb₂S₃ (x = 1), as shown in Fig. 2b.

Table 1 shows the unit cell lattice parameters a , b and c and unit cell volume V for Bi_{2–2x}Sb_{2x}S₃ ($0 \leq x \leq 1$), which were calculated from the X-ray-diffraction data using the lattice relation for orthorhombic structure; $1/d^2 = h^2/a^2 + k^2/b^2 + l^2/c^2$ and $V = abc$, where d is lattice plane spacing and (hkl) are the Miller indices of the plane.⁶⁹ The lattice parameters of Bi₂S₃



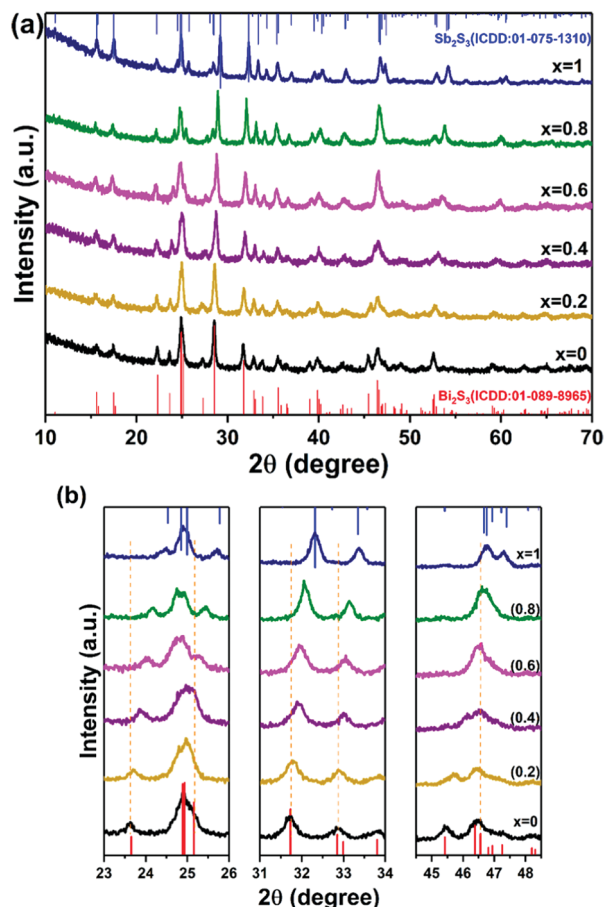


Fig. 2 (a) XRD patterns of Bi_2S_3 , Sb_2S_3 and $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ with variations in Sb/Sb+Bi molar ratios at 250 °C with standard patterns of Bi_2S_3 (bottom) and Sb_2S_3 (top). (b) Shows the shift of diffraction peaks to larger values as the composition x changes from 0 to 1.

and Sb_2S_3 were in agreement with literature values (ICDD: 01-089-8965 for Bi_2S_3 and ICDD: 01-075-1310 for Sb_2S_3).^{67,68} Lattice parameters are plotted against the variations in Sb/Sb+Bi molar ratios (Fig. 4). It is clearly seen that a and b increase linearly as the mole fraction of Sb increases, whereas c and V decrease gradually with an increase in the Sb content. A decrease of almost 3.5% in the volume was observed as moving from Bi_2S_3 to Sb_2S_3 . This is in accordance with the smaller atomic radius of antimony compared to bismuth and with literature X-ray data for $\text{Bi}_{2-x}\text{Sb}_x\text{S}_3$ films prepared by the dip-dry method.⁷⁰

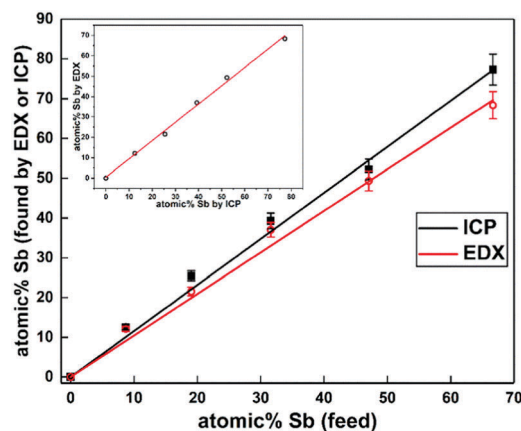


Fig. 3 Comparison of change in Sb content found in elemental analysis of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) with increasing antimony molar ratio by EDX and ICP. The inset shows the plots of ICP versus EDX.

The substitution of Sb^{3+} for Bi^{3+} ions in the Bi–Sb–S solid solution series causes the contraction of the inter-rod distance, which is the main factor for increasing the a and b unit cell parameters from the Bi to the Sb forms.⁷¹ For the c unit cell parameter, a continuous contraction is observed in going along the solid solution which may be related to the contraction of rods along this direction as a consequence of the decrease of the M–S bond lengths from Bi_2S_3 to Sb_2S_3 .⁷¹

The elemental compositions of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) samples were examined through EDX and ICP-OES. EDX spectra (ESI,† Fig. S3) of the Bi–Sb–S alloys show the characteristic peaks of bismuth $\text{M}\alpha$, antimony $\text{L}\alpha$, and sulfur $\text{K}\alpha$. Table 1 shows the atomic percentages of bismuth, antimony, and sulfur for all compositions calculated by EDX. The analysis was performed for several different areas which all yielded similar results, suggesting a uniform elemental distribution. The results of both EDX and ICP-OES are consistent and confirm that the atomic ratios of all these samples have the expected Sb content in the precursor mixture (Fig. 3). All samples were found to be slightly sulfur deficient (Table 1), perhaps due to the high partial pressure of the sulfur, which causes the evaporation of sulfur at high temperature.⁷²

Raman spectroscopy was used to investigate the $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) samples prepared from mixtures of bismuth and antimony xanthates at 250 °C (Fig. 5). The pure Bi_2S_3 sample shows a weak peak at 184 cm^{-1} and other two dominant peaks

Table 1 Lattice parameters (a , b , c), unit cell volume (V), band gap (E_g) and EDX compositional analysis of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) with variations in Sb/Sb+Bi molar ratios

x	Composition	Lattice parameters				EDX (at%)			E_g (eV)
		a (Å)	b (Å)	c (Å)	V (Å ³)	Bi	Sb	S	
0	Bi_2S_3	11.173 ± 0.00078	11.322 ± 0.00079	3.989 ± 0.00028	504.61 ± 0.062	43 ± 2.1	0	56 ± 2.8	1.66
0.2	$\text{Bi}_{1.6}\text{Sb}_{0.4}\text{S}_3$	11.185 ± 0.00078	11.324 ± 0.00079	3.981 ± 0.00028	504.23 ± 0.061	37 ± 1.8	11 ± 0.5	52 ± 2.6	1.72
0.4	$\text{Bi}_{1.2}\text{Sb}_{0.8}\text{S}_3$	11.209 ± 0.00078	11.325 ± 0.00079	3.936 ± 0.00028	499.64 ± 0.060	28 ± 1.3	18 ± 0.8	55 ± 2.7	1.75
0.6	$\text{Bi}_{0.8}\text{Sb}_{1.2}\text{S}_3$	11.233 ± 0.00077	11.328 ± 0.00078	3.878 ± 0.00027	493.46 ± 0.059	17 ± 0.8	27 ± 1.3	56 ± 2.8	1.88
0.8	$\text{Bi}_{0.4}\text{Sb}_{1.6}\text{S}_3$	11.245 ± 0.00077	11.332 ± 0.00078	3.855 ± 0.00027	491.24 ± 0.059	8 ± 0.4	33 ± 1.6	59 ± 2.9	2.08
1	Sb_2S_3	11.248 ± 0.00076	11.334 ± 0.00078	3.829 ± 0.00026	488.14 ± 0.058	0	41 ± 2.0	59 ± 2.9	2.19

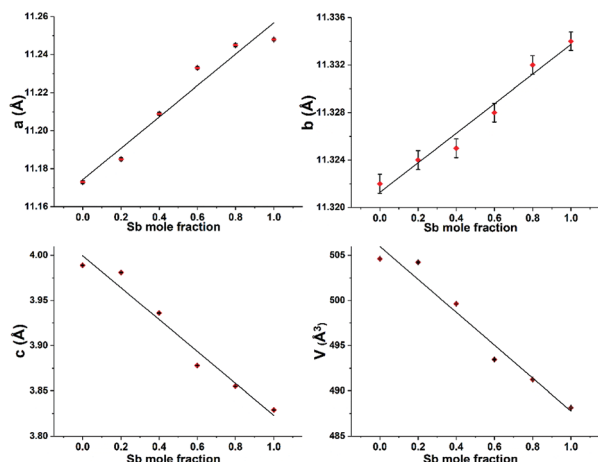


Fig. 4 Lattice parameters a , b , c and unit cell volume V of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) with the different mole fractions of antimony.

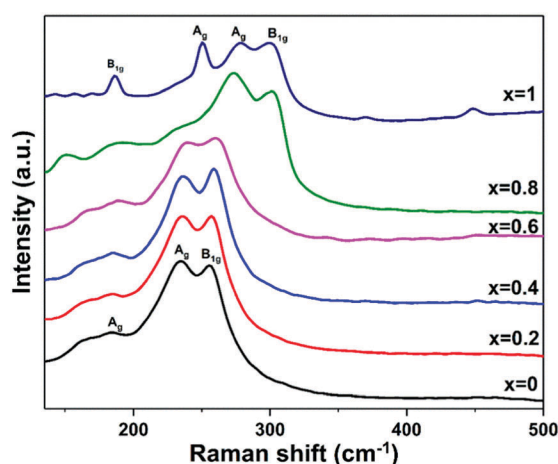


Fig. 5 Raman spectra of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) samples prepared by melt thermolysis at 250°C using bismuth and antimony xanthate precursors with different mole fractions of Sb.

located at 235 and 255.4 cm^{-1} , which is in agreement with the Raman analysis reported for Bi_2S_3 .^{53,59,73} The low intensity peak at 184 cm^{-1} is assigned to the A_g symmetric bending modes, and the peaks at 235 and 255 cm^{-1} are assigned to the A_g and B_{1g} anti-symmetric stretching modes, respectively.^{59,74} Similarly, the Raman spectrum of the pure Sb_2S_3 indicates peaks at 186 , 250 , 280 , 301 and 448 cm^{-1} corresponding to the Raman spectrum of orthorhombic Sb_2S_3 .^{75–79} It is proposed to assign the peak at 186 cm^{-1} to the B_{1g} anti-symmetric S–Sb–S bending modes, the peak at 250 cm^{-1} to the A_g symmetric S–Sb–S bending modes, and the peaks at 280 and 301 cm^{-1} to the A_g and B_{1g} anti-symmetric Sb–S stretching modes, respectively.^{74,77} For the $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($x = 0.2, 0.4, 0.6$ and 0.8), the A_g and B_{1g} modes, which involve the vibration of metal and sulfur atoms are shifted continuously toward higher wavenumbers with increasing antimony content. The large shift of corresponding peaks to higher wavenumbers in Sb_2S_3 is caused by both the lower mass of Sb and the shorter Sb–S distances in comparison with Bi–S.⁷⁴

The morphology of all the samples was observed using SEM. Images of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ obtained at 250°C using $0, 0.2, 0.4, 0.6, 0.8$ and 1 mole fractions of Sb are presented in Fig. 6. The SEM image of pure Bi_2S_3 ($x = 0$) revealed nanorod structures (Fig. 6a), which changed to slightly thicker nanorods as the Sb content increased to 0.4 mole fraction (Fig. 6b and c). The samples with higher mole fractions of Sb showed completely different morphologies; large particles were observed on addition of 0.6 and 0.8 mole fractions of Sb, which themselves consist of nano-sized rods (Fig. 6d and e). Finally, the SEM image of pure Sb_2S_3 ($x = 1$) revealed uneven particles (Fig. 6f). EDX spectrum mapping of the mixed samples prepared at all the compositions x are shown in Fig. 7. It is clearly seen that at this length scale all the elements (Bi, Sb and S) were evenly distributed throughout the entire sample.

TEM electron diffraction (ED) was used to characterise the local $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ crystal structure for two samples ($x = 0.2, 0.8$).

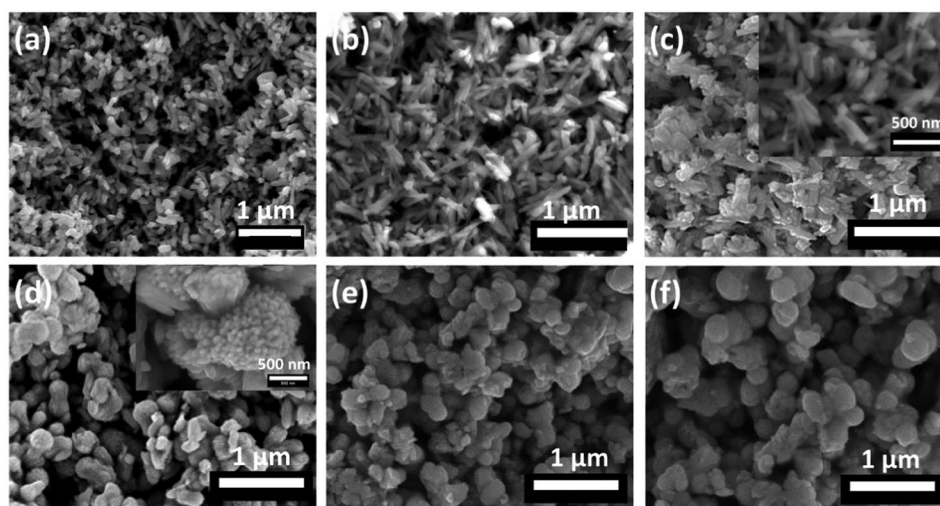


Fig. 6 SEM representative secondary electron SEM images (10 kV) of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) samples prepared by melt thermolysis at 250°C using bismuth and antimony xanthate precursors with different mole fractions of antimony: (a) $x = 0$ (Bi_2S_3), (b) $x = 0.2$, (c) $x = 0.4$, (d) $x = 0.6$, (e) $x = 0.8$ and (f) $x = 1$ (Sb_2S_3). Insets: SEM images of samples at higher magnification.



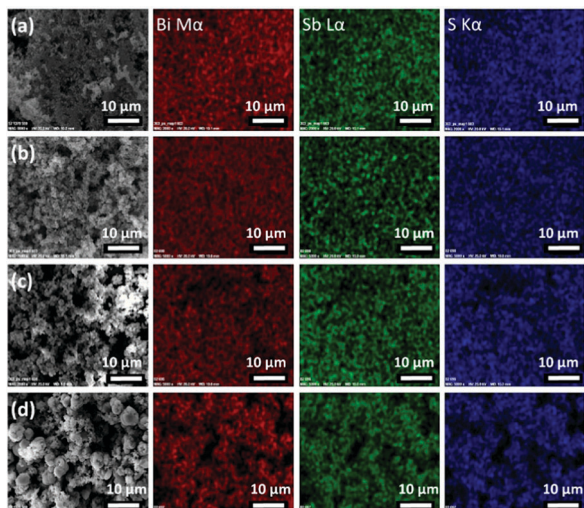


Fig. 7 EDX elemental mapping (20 kV) of Bi M α , Sb L α and S K α for Bi_{2-2x}Sb_{2x}S₃ samples. (a) $x = 0.2$, (b) $x = 0.4$, (c) $x = 0.6$ and (d) $x = 0.8$ mole fractions of antimony.

The Bi_{2-2x}Sb_{2x}S₃ was ground to a fine powder and placed on a holey carbon grid. As with SEM, significant differences in morphology were observed for each sample: for $x = 0.2$ crystalline nanorods were observed, while for $x = 0.8$ the material was in the form of bulky, clustered particles (Fig. S6, ESI†). Fig. 8a shows electron diffraction patterns for each sample with 'spotty' rings characteristic of many randomly orientated crystals being present within the sample region chosen by the selected area aperture. The presence of many small crystallites is confirmed by TEM imaging (Fig. 8b and Fig. S6, ESI†). The radii of the coloured arcs overlaid on the ED patterns in Fig. 8a correspond to the (hkl) lattice spacings determined from p-XRD data. Table S3 (ESI†) compares the lattice spacings measured by radial averaging of the ED data with the expected interplanar spacings predicted from the bulk p-XRD analysis (Table 1). These show a good match within the measurement accuracy of the ED data.

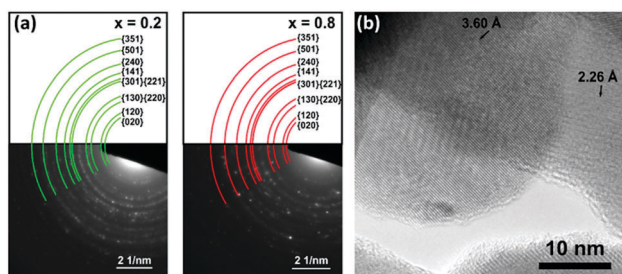


Fig. 8 (a) Electron diffraction patterns of Bi_{2-2x}Sb_{2x}S₃ for $x = 0.2$ (left) and $x = 0.8$ (right). Coloured rings represent d -spacings for $x = 0.2$ (left) and $x = 0.8$ (right) identified by p-XRD. The ED data also contains additional lattice spacings for the crystal structure which are indexed in Table S3 (ESI†). The presence of additional lattice spacings in the ED data is expected as the ED condition for constructive interference is less severe than for p-XRD (due to the small electron wavelength and small size of the crystallites). (b) TEM micrograph showing several individual crystallites for Bi_{2-2x}Sb_{2x}S₃, $x = 0.8$. Lattice spacings of 2.2 Å and 3.6 Å were assigned to the (141) and (130) planes, respectively.

For optical bandgap measurements, thin-film samples of Bi_{2-2x}Sb_{2x}S₃ were prepared by spin coating different molar concentrations of precursors (1):(2) on glass substrate to yield mole fractions $x = 0, 0.2, 0.4, 0.6, 0.8$ and 1, and then annealed in quartz tube under nitrogen at 250 °C for an hour. Optical absorption spectra of these samples have been recorded in wavelength range of 400–1000 nm (Fig. S7, ESI†). The band gap of the films can be calculated using the formula $(\alpha h\nu)^n = A(h\nu - E_g)$, where E_g is the optical band gap, $h\nu$ is the photon energy, α is the absorption coefficient, A is a constant characteristic of the material, and $n = 2, 1/2$ for allowed direct and allowed indirect transitions, respectively. Bi₂S₃ and Sb₂S₃ are direct band gap materials.^{80,81} Plots of the $(\alpha h\nu)^n$ versus $h\nu$ for $n = 2$ and $1/2$ show a linear behaviour for $n = 2$, which confirms the presence of a direct transition in the Bi_{2-2x}Sb_{2x}S₃ samples for all compositions (Fig. S8, ESI†). Fig. S8, ESI† presents a Tauc plot to determine the band gaps for all Bi_{2-2x}Sb_{2x}S₃ samples. For the binary Bi₂S₃ and Sb₂S₃ compositions, the direct band gaps were found to be 1.66 and 2.19 eV, respectively; very close to the reported values for these materials.^{3,35,80–83} The values of the band gap for the mixed samples Bi_{2-2x}Sb_{2x}S₃ were estimated to be 1.72, 1.75, 1.88 and 2.08 eV for $x = 0.2, 0.4, 0.6$ and 0.8, respectively, values which lie between those of pure Bi₂S₃ and Sb₂S₃ (Table 1). Fig. 9 represents the dependence of band gap of the Bi_{2-2x}Sb_{2x}S₃ samples on composition, demonstrating that the band gap of Bi_{2-2x}Sb_{2x}S₃ films increases from 1.66 to 2.19 eV with increasing Sb concentration. This gradual change in the band gap confirms the formation of solid solution. A deviation from a linear trend is probably due to the slight stoichiometric variations in the synthesized solid solution.⁸⁴ Similar observations for (Bi_{1-x}Sb_x)₂S₃ thin films grown by arrested precipitation and spray pyrolysis techniques have been reported.^{66,85} Patil *et al.* found that the films have a band gap of 1.8 eV for Bi₂S₃ and 2.3 eV for Sb₂S₃, whereas the mixed films have band gaps in the range of 1.85–2.27 eV with increasing Sb content in the films.⁶⁶ It has also been found that incorporation of only 3% bismuth into Sb₂S₃ narrowed the band gap from 1.72 to 1.59 eV.⁸⁶

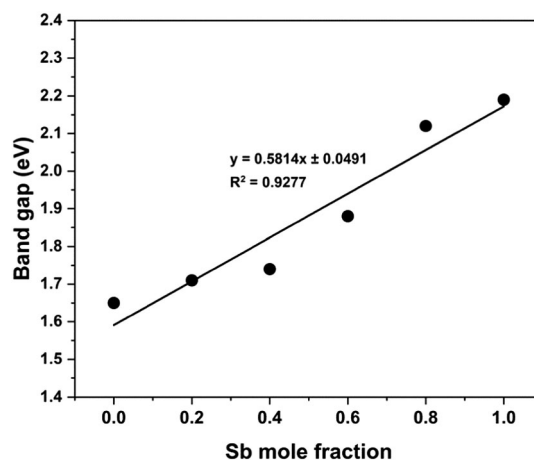


Fig. 9 Variation of the band gap for Bi_{2-2x}Sb_{2x}S₃ ($0 \leq x \leq 1$) samples as a function of mole fraction (x) of antimony.



Conclusions

$\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ ($0 \leq x \leq 1$) solid solution, with controllable compositions were successfully synthesized *via* solventless thermolysis using molecular precursors. Structural analysis *via* p-XRD and TEM-ED indicates the successful incorporation of antimony as a solid solution into the parent bismuth sulfide. SEM images showed changes in the morphology as the amount of antimony was increased in the samples. The elemental compositions of all the samples were examined *via* ICP-OES and EDX spectroscopic mapping, with the latter technique revealing uniform spatial distribution of elements in all samples. The band gaps of $\text{Bi}_{2-2x}\text{Sb}_{2x}\text{S}_3$ films were found to vary from 1.66 eV for Bi_2S_3 to 2.19 eV for Sb_2S_3 with increasing antimony mole fraction (x). These ternary bismuth–antimony chalcogenides may find applications as photoelectrocatalysts for solar water splitting,⁸⁷ as their band gap lies within the suitable range for such applications. The approach provides a simple, scalable and efficient synthesis route to produce ternary sulfide materials with electronic properties that are tuneable across the full compositional range, which, given the rich range of metal xanthate precursors available, gives an exciting opportunity for the low temperature synthesis of a range of semiconductors.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

T. A. thanks the Imam Abdulrahman Bin Faisal University for the funding support and for allowing a period of study in the U.K., and also, the Saudi Arabian Cultural Bureau in London for their assistance and support. S. J. H. and D. J. K. thank the Engineering and Physical Sciences (EPSRC) U.K (grants EP/P009050/1 and the NowNANO Graphene CDT), the Defense Threat Reduction Agency (HDTRA1-12-1-0013) and the European Research Council (ERC) under the European Union's Horizon 2020 research and innovation program (ERC-2016-STG-EvoluTEM-715502) for funding.

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