



Interaction of oxygen with halide perovskites†

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We systematically investigate the consequences of exposing halide perovskites to an oxygen-containing atmosphere, in particular methylammonium lead iodide, the archetypal compound used as photo-absorber in perovskite solar cells. For this purpose, we study oxygen solubility and global reaction with oxygen both in the dark and under light, and we refer to the kinetics in terms of surface reaction and bulk diffusion. As thermodynamics reveals, the material is unstable against oxygen, primarily because of the large driving force of water formation. While under light the material quickly degrades, in the dark the surface reaction kinetics – not the bulk transport – is very sluggish and keeps it metastable. For the same reason, oxygen incorporation into the lattice is negligible in the dark. On illumination, an accelerated oxygen in-diffusion occurs (as shown by ¹⁸O incorporation experiments) that severely modifies the electronic and ionic conductivities in the way that is expected for an acceptor dopant. Lastly, we investigate the impact of cation as well as anion mixing on the degradation and stress the necessity of using encapsulation during solar cell operation.

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Mixed organic–inorganic halide perovskites (MOIHPs) have been extensively studied in recent years due to their prospective application as light harvesters in solar cells.^{1,2} This new class of devices, named perovskite solar cells, offers both simple processing and outstandingly high photo-conversion efficiencies. Unfortunately, the development and application of such devices is impeded by the poor stability of the halides. In particular, degradation processes in MOIHPs can be easily induced by temperature,^{3–5} water or high humidity^{6,7} and even high-intensity light.^{8,9} A major source of instability is exposure to oxygen.^{10–13} Notwithstanding the fact that oxygen is omnipresent during material synthesis as well as device operation, the interaction of O₂ with MOIHPs has received only minor attention. Some of these reports show that oxygen can strongly and reversibly improve photoluminescence in MOIHPs.^{14–16} Other investigations deal with the formation of reactive superoxide species,^{10,17,18} attributed to charge transfer between O₂ molecules and the photo-excited MOIHP.^{11,19,20} Notably, several commonly used organic hole- or electron-transport materials are known to degrade when exposed simultaneously to oxygen and light,^{21–25} making it crucial to distinguish the influence of O₂ on the halide perovskite layer from the effect on the hole- or electron-transport materials used in perovskite solar cell devices. Surprisingly, there are no studies on oxygen solubility in MOIHPs, even though O-defects will act as dopants and

severely influence not only their electronic²⁶ but also ionic transport properties. For these reasons, we investigate reactivity and solubility of oxygen in the pure archetypal halide perovskite, *i.e.* methylammonium lead iodide (MAPbI₃), comparing dark and illuminated conditions. We study both thermodynamics and kinetics of oxygen interaction, ranging from solubility to phase conversion (degradation). We observe that oxygen incorporation in MAPbI₃ is modest in the dark, but strongly enhanced under illumination. By combining ¹⁸O chemical diffusion techniques with electrical characterizations and with UV-Vis and XRD degradation studies, as well as thermodynamic data, we obtain a consistent picture of the effects of O₂ not only on transport properties, but also phase stability and degradation behavior of MAPbI₃. Based on this, as a final point we also analyzed the stability of mixed halide (I, Br) and mixed cation (MA, FA) perovskite formulations, since such solid solutions represent the state of the art in terms of device applications.

Oxygen dissolution

Let us start with discussing the oxygen dissolution in MAPbI₃ and its effect on transport properties (next section) under conditions where overall the perovskite phase remains intact. Owing to the configurational entropy involved, a certain solubility of oxygen in the MAPbI₃ lattice must always occur. Considering the situation on the atomistic scale, atomic oxygen species can be introduced substitutionally, occupying an iodine site or an iodine vacancy and hereby forming a substitutional oxygen defect with a negative effective charge (denoted as O_i[−]). Alternatively, since the ionic radius of O^{2−} is only 140 pm (I[−] is

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210 pm), oxygen can be introduced interstitially, thus forming a doubly negatively charged defect (denoted as O_i''). For the purpose of a direct experimental observation of oxygen dissolution in MAPbI_3 , we performed chemical diffusion experiments with a ^{18}O tracer (Fig. 1). As shown in Fig. 1b (inset), in the dark we find incorporation of O_2 in thin films to be poor and sluggish. This remains true even when elevated temperatures and long waiting times are involved, as shown for pellets (Fig. 1d). This is in striking contrast to the situation under illumination, where oxygen incorporation is substantially accelerated and enhanced (Fig. 1b). XRD studies show that the MAPbI_3 phase was almost entirely preserved after these treatments (Fig. S3†). The detailed analysis of the diffusion profiles (Fig. 1c) reveals that the light-enhanced kinetics is due to an accelerated surface rate constant (k^δ) while the impurity diffusion coefficient (D^δ) remains rather unchanged. While the effect on k^δ is not unexpected for perovskites, the insensitivity of D^δ deserves some explanation, as our previous work showed a largely increased ionic carrier concentration under light.²⁷ An independence of the impurity diffusion constant follows directly for the motion of interstitial oxygen defects (O_i''), but it may also apply for the diffusion of substitutional defects (O_i'), provided such species move strongly coupled with the induced iodine vacancy (such coupling is, in fact, expected for room temperature, as we detail in the ESI Section 5†). As far as the increased rate constant of the surface step under light is concerned, this may be attributed to an accelerated O_2 reduction

caused by the photo-generated electrons (ESI Section 3, eqn (S10)†), as in the case of SrTiO_3 .²⁸ This is in also in agreement with previous studies showing how perovskite solar cells stability can be extended by improving the extraction of photo-generated carriers.^{11,17,20} These considerations also imply that in the dark oxygen incorporation is limited by the slow surface reaction. We note that the tracer profiles we observe here are typical of a bulk process, and also that the property changes discussed in detail below (e.g. doping effect) clearly show that the bulk of the material is affected by oxygen treatment. Nevertheless, even though we have no indications of comparatively fast grain boundary contribution (see ESI Section 13†) in our samples, we recognize that other preparation procedures could result in more open grain boundaries that would even enhance the incorporation rate.^{17,29,30}

Dependence of transport properties on oxygen partial pressure

Upon oxygen incorporation, be it substitutional or interstitial, a negative dopant is generated (O_i' or O_i''), which gives rise to an electronic acceptor effect (increase of hole, decrease of excess electron concentration) as well as production of positively charged iodine vacancies (V_i'), that we have identified as decisive ionic charge carriers in this material.³¹ For the effect of such doping to be significant, the incorporated oxygen content needs

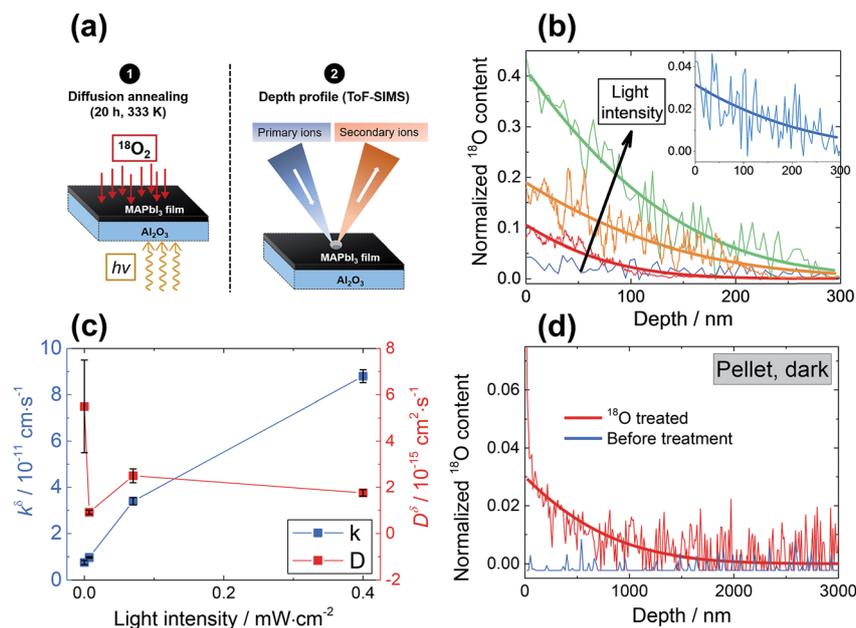


Fig. 1 (a) Schematic of the ^{18}O incorporation experiment performed on MAPbI_3 films deposited on Al_2O_3 substrates. First (1), the samples are annealed under oxygen; when necessary, light is simultaneously applied. After the treatment (2), the oxygen content is analyzed as a function of depth from the surface using Time of Flight Secondary Ion Mass Spectroscopy (ToF-SIMS). (b) ToF-SIMS depth profiles of four MAPbI_3 thin films exposed to pure $^{18}\text{O}_2$ at 333 K for 20 hours. One of the films was kept in dark (— blue), while the others illuminated with 1 (— green), 0.3 (— orange) and 0.015 (— red) mW cm^{-2} light. Insets show a magnification of the profile of the film kept under dark, where the profile is still visible and in agreement with the one observed for the pellet of panel (d). (c) Chemical diffusion coefficients (D^δ) and surface rate constants (k^δ) extracted from the curves in panel (b) (at 333 K). (d) ToF-SIMS depth profiles of a MAPbI_3 pellet, before and after exposure to $^{18}\text{O}_2$ at 378 K for 46 hours under dark conditions (D_{pellet}^δ (378 K) = $3 \times 10^{-14} \text{ cm}^2 \text{ s}^{-1}$, k_{pellet}^δ (378 K) = $1.2 \times 10^{-11} \text{ cm s}^{-1}$). The amount of ^{18}O found before treatment is too little to be properly detected, and appears lower than natural abundance, thus the normalized content lies below 0.



to exceed the intrinsic defect concentrations but may be still very small when compared with the number of regular lattice constituents. The quantitative consequences of oxygen doping on electronic and ionic carrier concentrations can be deduced from mass action considerations. The results of these considerations are reflected by the defect diagram shown in Fig. 2c and, more in detail, in the ESI (Section 3†). To analyze the oxygen effect quantitatively, we monitor the changes in electronic and ionic conductivities of a MAPbI₃ film as a function of oxygen partial pressure, as reported in Fig. 2 (kinetics considerations are given in ESI Section 8†). As the comparison of the defect diagram of Fig. 2c with Fig. 2b (and Fig. S5†) demonstrates, the behavior under illumination well agrees with the predicted doping effect: At low oxygen pressures ($P(\text{O}_2) < 10^{-2}$ bar), the electronic conductivity increases with $P(\text{O}_2)$, while the ionic conduction appears invariant (this behavior is also observed in the dark, see below). At higher pressures ($P(\text{O}_2) > 10^{-2}$ bar), however, the situation changes drastically, in that here the ionic conductivity increases as well, while the electronic conductivity changes more steeply. We note that, in

principle, also superoxide ions could cause such effect, but only if incorporated interstitially, which is thermodynamically not expected. If superoxide species would, as proposed, occupy iodine vacancies,¹⁷ no heterovalent doping effect would be achieved, in contrast with the experimental observations (see ESI Section 14† for more details). Having established that under illumination oxygen incorporation takes place and leads to the expected acceptor doping, we now turn to the conductivity effects caused by oxygen exposure in the dark, which appears puzzling at a first glance but fits nicely into the thermodynamic and kinetic picture given below. As the ¹⁸O incorporation experiments show, without illumination only little oxygen incorporation takes place (Fig. 1), hence one would not expect a significant effect on the transport properties of MAPbI₃. We observe, however, that the samples (Fig. 2a) exhibit an increase in electronic conduction upon increasing $P(\text{O}_2)$, while the ionic conductivity stays rather constant. This behavior is not the one expected from O-doping, but is qualitatively comparable to what is observed upon exposure of MAPbI₃ to an increasing iodine partial pressure.³¹ In fact the variations completely disappear

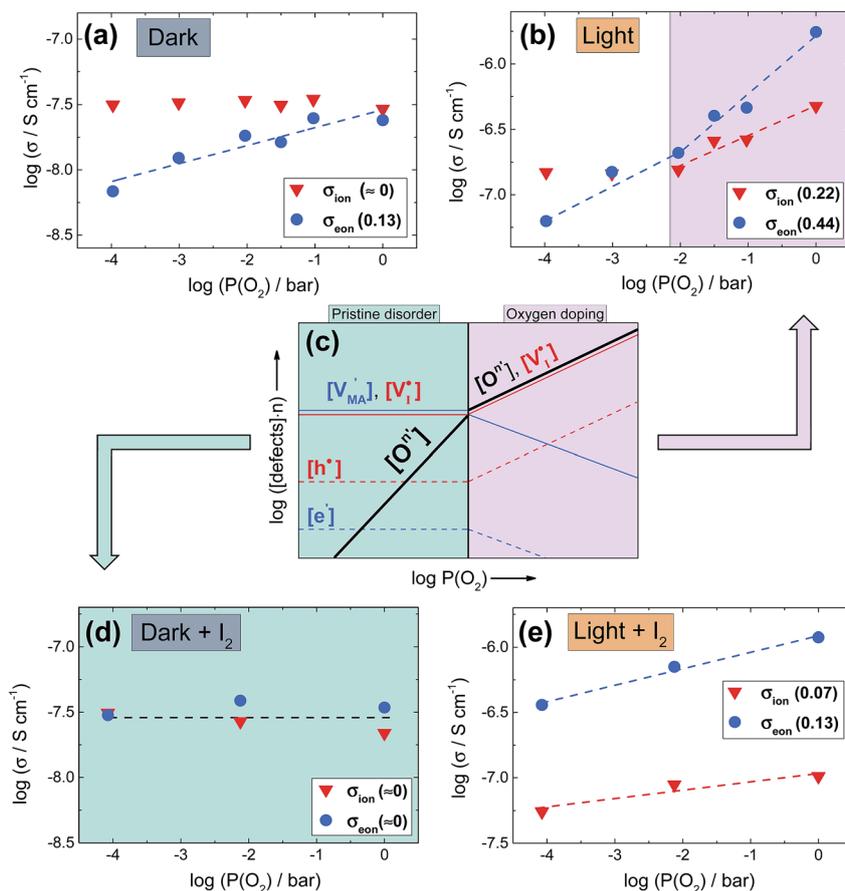


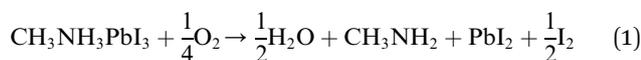
Fig. 2 (a and b) Ionic and electronic contribution to the conductivity of MAPbI₃ thin films as a function of oxygen partial pressure, extracted from d.c. galvanostatic polarization experiments performed at 333 K, in the dark and under 0.5 mW cm^{-2} illumination. Slopes are given in the legends. The behavior under higher light intensity is reported in Fig. S5.† (c) Kröger–Vink diagram representing defect concentrations as a function of oxygen partial pressure. The incorporation of O₂ is noted as a general atomic O-defect of effective negative n -charge (O^{\prime}). The relative slopes obtained when this defect overcomes the intrinsic equilibrium and achieves doping (violet region) are dependent on the charge of the defect, and are given in ESI Fig. S2.† (d and e) Conductivity measurements as a function of $P(\text{O}_2)$ (analogous as panel (a) and (b)), this time carried out under constant $P(\text{I}_2)$ (1.4×10^{-6} bar), in the dark and under 0.5 mW cm^{-2} illumination. Note that, in the dark, conductivity is not affected by oxygen pressure variations, while under illumination only a weak dependence remains.



when the outer iodine partial pressure is fixed (compare Fig. 2a and d). These results show that the O₂ effect in the dark originates from the surface, where in simple terms (ESI Section 3†) oxygen converts I⁻ to iodine, thus changing the iodine activity. Unlike for oxygen, the communication of the lattice with iodine is rather quick.³¹ Below we will show that such a surface conversion reaction consuming oxygen and producing iodine is indeed expected. Also under illumination, the simultaneous presence of iodine and oxygen severely weakens the O-influence on the transport properties (compare Fig. 2b and e), in agreement with a reduced iodine removal due to the fixed external I₂ partial pressure (details in ESI Section 3†). A further interesting point is the observed reversibility (Fig. S6†), in spite of oxygen incorporation leading to iodine excorporation (ESI Section 3†); it follows that iodine is not excorporated into the gas phase, but rather it appears to be accommodated at higher dimensional defects, or even in the interstitial lattice.²⁷

Degradation thermodynamics

We now discuss the phase stability of halide perovskites with respect to oxygen. Thermodynamically, here we refer to oxygen partial pressures that lead to an oxygen content exceeding the limit of oxygen solubility. Using recent experimental values^{4,32,33} in conjunction with tabulated standard data (ESI Section 2†) we can give standard reaction Gibbs energies for a variety of possible decomposition reactions (ESI Table S3†). Solubility effects, though highly relevant for transport issues, are negligible as far as their impact on phase energetics is concerned. As pointed out in ref. 34, reactions that involve C–N bond breaking (and conversion of MAPbI₃ into CH₃I and NH₃) are kinetically hindered and not relevant for room temperature, while the expected decomposition reactions gives PbI₂, HI and CH₃NH₂ (Table S3, eqn (S2)†). The related Δ_rG⁰ value is distinctly positive (>120 kJ mol⁻¹), so that in spite of the fact that the educts can be removed under a dynamic gas flow, this reaction can only occur to a very minor extent. A closer look into the thermodynamics shows that HI is marginally unstable with respect to decomposition, slightly facilitating the MAPbI₃ decomposition reaction into I₂ and H₂ (Table S3, eqn (S3)†). Much more energetically favourable, however, is the decomposition if O₂ is involved (Δ_rG⁰ = 5.3 kJ mol⁻¹), owing to the enormous energy gained when instead of HI (or H₂), water is formed:



We note that the same reaction was proposed for superoxide-induced degradation.^{10,17} Indeed, the kinetic mechanism behind phase decomposition (but also behind oxygen incorporation) is expected to involve reactive intermediates such as superoxide species (as it is also the case for perovskites at high temperatures),³⁵ while leaving the thermodynamic picture unchanged. It is also conceivable to further oxidize the products of eqn (1), in particular PbI₂ to PbO, resulting in an additional modest free energy gain (Table S3, eqn (S5)†). The fact that only PbI₂ is found as degradation product (as reported later) shows

that such oxidation is, as expected, kinetically difficult. Moreover, other reaction channels can be opened by the presence of water^{6,7} formed *in situ*, and also by illumination.^{8,9,27} Focusing on eqn (1), the aforementioned only slightly positive value of Gibbs energy for the decomposition reaction involving O₂ is easily turned into a negative balance if the produced gases are far from 1 bar of partial pressure (note that a reduction of the partial pressure by only one order of magnitude is already equivalent to a variation of -5.7 kJ mol⁻¹ per gas molecule). Hence a slightly positive standard reaction Gibbs energy – given the dynamic gas atmosphere – demands a nearly complete exclusion of oxygen for the reaction in eqn (1) not to be of relevance; without such exclusion, O₂ is thermodynamically expected to degrade MAPbI₃ to PbI₂ and release I₂. Given this clear issue, it is worth considering alternative MOIHP compositions that can avoid a similar degradation. In this respect, we note that other halides, such as MAPbBr₃ and MAPbCl₃ are thermodynamically expected to be much more stable with respect to O₂ (ESI Table S4†). Additionally, even though no thermodynamic data for formamidinium – (FA) containing perovskites have been reported, several works observed an increased stability of pure FAPbI₃ and or FA, MA mixtures with respect to pure MAPbI₃.^{36,37} We will experimentally address these points below.

Degradation kinetics

Now we are armed to discuss the kinetics of degradation of MAPbI₃ under O₂ exposure. Degradation kinetics is experimentally followed by monitoring *in situ* the UV-Vis absorption of a MAPbI₃ film kept under constant gas flow, with conditions purposefully chosen to induce a degradation of the films in a reasonable amount of time. As shown in Fig. 3, no degradation is observable even after more than 150 hours when exposing films to O₂ in dark conditions (at 333 K). When exposing the samples simultaneously to O₂ and light, however, full degradation to PbI₂ is achieved in less than 40 hours. As a reference, we exposed films to Argon under identical illumination conditions (Fig. S8a†), and no degradation was observed. These results are not unexpected, since the rather stable behavior in the dark – even under conditions where the whole phase should be converted into the products of eqn (1) – can be consistently attributed to the kinetic hindrances discussed above. Nonetheless, in the dark one may still expect this reaction to be of relevance for the surface layer, for which a transfer step into the bulk is not necessary. This straightforwardly explains the oxygen-induced “iodide-to-iodine conversion” observed when analyzing the transport properties (Fig. 2a). The aforementioned kinetic hindrances are, however, removed by illumination, allowing the sample to undergo a quick degradation. Interestingly, when following the degradation kinetics under illumination as a function of time (Fig. 3c), one clearly observes an initial relatively constant degradation rate, followed by a sudden accelerated drop, suggesting the opening of a parallel degradation path, such as one involving the water that is generated *in situ* by oxygen exposure. As expected from eqn (1), O₂-induced degradation can be readily slowed down by



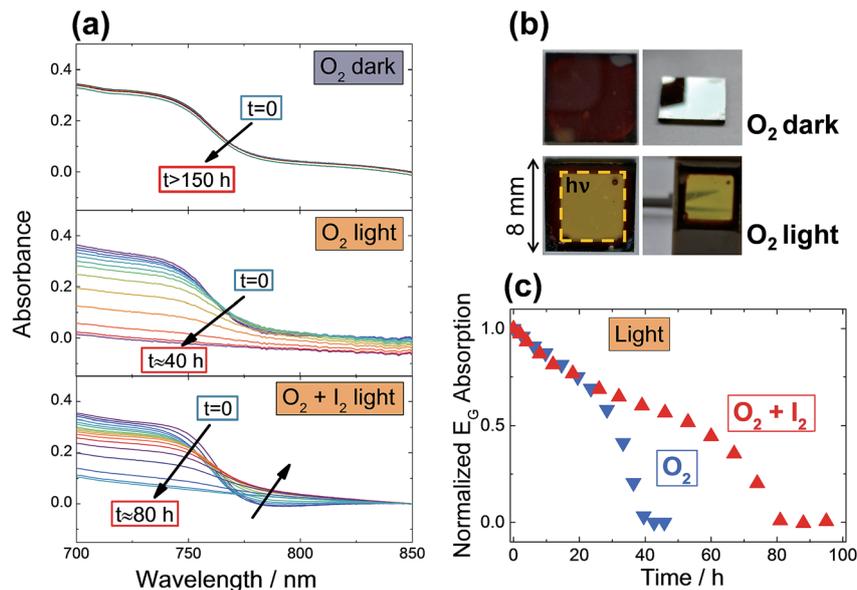


Fig. 3 (a) Stability of MAPbI₃ films under O₂ exposure, in dark and under illumination (4.6 mW cm⁻²), monitored by UV-Vis spectroscopy. The samples are kept at 333 K under constant gas flow. After more than 150 hours, no changes are visible for samples kept in the dark between measurements, while 40 hours are sufficient to completely degrade samples exposed to light. As a comparison, a film kept under the same conditions (O₂ and light) but also under constant iodine partial pressure ($P(\text{I}_2) = 1.4 \times 10^{-5}$ bar), shows doubled stability, with full degradation achieved after 80 hours. Note the formation of a tail in the absorption, indicating amorphization of MAPbI₃. (b) Photos of the sample after the O₂ exposure under dark and light. Note the clear difference between illuminated and non-illuminated area. The sample kept in dark is identical to a pristine sample, with good transparency and reflectivity. (c) shows the degradation rate (here taken as the bandgap absorption of MAPbI₃, see details in ESI Section 1†) as a function of time under light and O₂, with and without simultaneous iodine exposure. These data are extracted from the spectra of panel (a).

exposing the film simultaneously to iodine partial pressure (Fig. 3c). Unsurprisingly, exposure to similar iodine partial pressures in an otherwise inert atmosphere (Ar) does not lead to degradation (Fig. S8b†). For very high I₂ partial pressures, (≥ 1 bar), degradation of MOIHPs was observed,³⁸ which would imply another degradation path under such harsh conditions. The effect of simultaneous iodine exposure on the oxygen interaction with MAPbI₃ is visible also in photoluminescence (PL) spectra, where the enhancement of the PL response observed under pure O₂ (ref. 14–16) is weakened by I₂ (ESI Section 14, Fig S13 and S14†).

An analogous degradation study has been performed using *in situ* XRD measurements (Fig. 4a and b). The results obtained are in full agreement with the UV-Vis study, and here once again we observe an enhanced stability under simultaneous iodine and oxygen exposure. Interestingly, while under O₂ exposure the film degrades yielding crystalline PbI₂, under simultaneous exposure to O₂ and I₂ no PbI₂ peaks are visible after the treatment (Fig. 4a, S9c†). The color change from black (MAPbI₃) to yellow (PbI₂) occurs as expected, but the absence of diffraction peaks speaks for the formation of an amorphous phase. Notably, UV-Vis spectra under O₂ and I₂ (Fig. 3a) show the formation of an absorption tail at longer wavelengths, in line with a loss of crystallinity of the perovskite phase. A similar behavior – though much less pronounced – is present when exposing for long times MAPbI₃ to I₂ in inert atmosphere (under illumination, see Fig. S8b†). We recognize that this finding needs additional experimental studies, hence we will not discuss it further here.

With the intent of obtaining a deeper understanding of the O₂ degradation processes under conditions that are relevant for MOIHPs application, we investigated the influence of the spectral content and of the illumination intensity. As previously shown (Fig. 3), degradation quickly takes place even at low light intensity (4.6 mW cm⁻²) when coupled with a slight heating (333 K). A stronger light intensity (43 mW cm⁻²) can induce a quicker degradation even at room temperature (Fig. 4a and b). In both cases, no UV-light is present, clearly indicating that this is not necessary to induce degradation of MOIHPs under O₂; additionally, the illumination intensity used is significantly below 1 Sun (100 mW cm⁻²). We thus probe the intensity dependence of the degradation process (Fig. 4c); as expected, higher intensities accelerate the degradation process. Interestingly, as shown in Fig. 4d, we also observe a clear wavelength dependence for the degradation. Under identical light intensities, violet (420 nm) light is much more effective than red (660 nm) in inducing degradation. As local heating of the MOIHP surface is not a probable reason for this phenomenon (the samples are at 333 K), we postulate that it is the result of a favored electron transfer to the adsorbed O₂ molecules due to high-energy electrons. We can thus conclude that, under real operative conditions (1 Sun, including high-energy UV-light) the degradation processes will be much more severe than what we observe here, and this makes the O₂-induced degradation of MOIHP most relevant for real photovoltaic devices. In real devices, additional degradation pathways may occur, but note that the increased complexity of a solar cell cannot suppress the



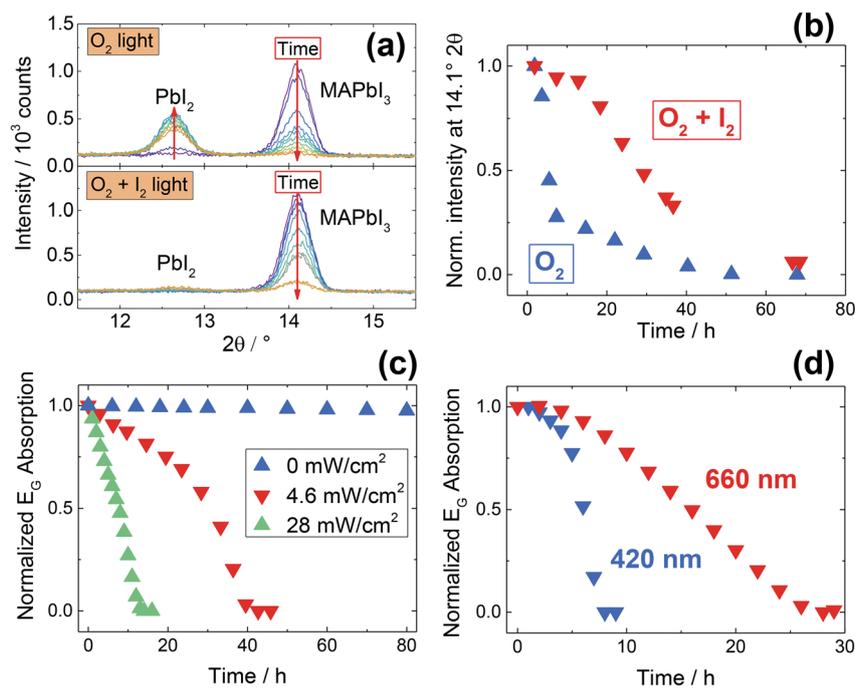


Fig. 4 (a) Stability of MAPbI₃ films under O₂ exposure under illumination (43 mW cm⁻²), with and without simultaneous iodine exposure ($P(I_2) = 4.5 \times 10^{-6}$ bar), measured by *in situ* XRD at room temperature. Exposure to O₂ in the dark gives no significant degradation (Fig. S9a†), while under illumination a severe degradation takes place, with full disappearance of all MAPbI₃ reflections after 50 hours (see also Fig. S9b†). After degradation under oxygen and iodine, no PbI₂ peaks are visible, indicating the formation of an amorphous phase. (b) Normalized intensity of the strongest MAPbI₃ reflection (at 14.1° 2θ), as a function of degradation time, extracted from the diffractograms of panel (a). The film kept under simultaneous iodine exposure shows enhanced stability. We recognize that, due to preferred orientations of the MAPbI₃ crystals, the analysis of the intensity of a single peak can only be semi-quantitative. (c) Degradation kinetics of pure MAPbI₃ films under O₂ measured with UV-Vis spectroscopy under different light intensities. (d) Wavelength dependence of the degradation kinetics, monitored using UV-Vis spectroscopy, under 4 mW cm⁻² illumination supplied by different LED lamps.

inherent instability of the photo-absorber if oxygen access is possible.

Since for MAPbI₃ the degradation processes are clearly very severe, as a last analysis we investigate the stability of a mixed-halide (I, Br) and of a mixed-cation (MA, FA) MOIHP under O₂. Starting with the mixed-halide perovskite, we chose a phase-stable formulation, *i.e.* MAPb(Br_{0.2}I_{0.8})₃, in order to avoid halide de-mixing under illumination.³⁹ Interestingly, the sample does not show any enhanced stability with respect to the pure iodide (Fig. 5a). XRD analysis of the film after degradation, reported in Fig. 5b, shows that the expected removal of I₂ (due to O₂ incorporation in the perovskite) takes place in the mixed-halides as well, resulting in the formation of PbI₂ and MAPbBr₃, while the pristine sample displayed only the crystallographic phase of tetragonal MAPbI₃. This clearly indicates that mixed-halide perovskites still suffer from severe O₂-induced degradation, notwithstanding a free energy gain obtained from mixing entropy (which is expected however to be rather small). As previously predicted by thermodynamics considerations, the newly formed MAPbBr₃ phase easily survives the O₂ exposure, since no PbBr₂ reflections are present after the treatment. As far as the mixed-cation perovskites are concerned, we choose a formulation for which the perovskite phase is structurally stable at room temperature, namely (MA_{0.8}FA_{0.2})PbI₃ (note that pure FAPbI₃ crystallizes in a non-perovskite phase).⁴⁰ Notably, in this case the degradation rate

is significantly lowered – even though not completely suppressed – by mixing FA with MA (Fig. 5c and d, also Fig. S10†). As it is not the minor mixing free energy that can be responsible for the enhanced stability (*cf.* results on mixed halide formulations), we can attribute this finding to the following reasons: (i) thermodynamically, the stability is increased as deprotonation of the FA cation is less favored (formamidinium is a stronger base than methylamine),⁴¹ and so will be the final oxidation driving force. (ii) Kinetically, we found a diminished ion conductivity by FA admixing³⁶ that we directly observe in electrical measurements (Fig. S11†). This leads to a lowering of the rate of the transport step of the degradation. For all these reasons, we expect that higher FA contents, such as the one commonly used in state-of-the-art perovskite solar cells (FA > 80%), could potentially yield more stable, but possibly not phase-pure compounds.⁴² The analysis given in the ESI (Fig. S15†), shows that higher FA contents indeed improve the oxygen stability. Future work is needed to decide if this is due to the chemistry or to the possible two-phase nature of the material. Lastly, we note that in the case of (MA_{0.8}FA_{0.2})PbI₃, no peak belonging to PbI₂ are observed after degradation, indicating formation of an amorphous phase (ESI Section 11†) or pure FAPbI₃. Moreover, UV-Vis spectra (Fig. 5c) show the formation of an absorption tail at longer wavelengths, indicating a loss of crystallinity of the perovskite phase.



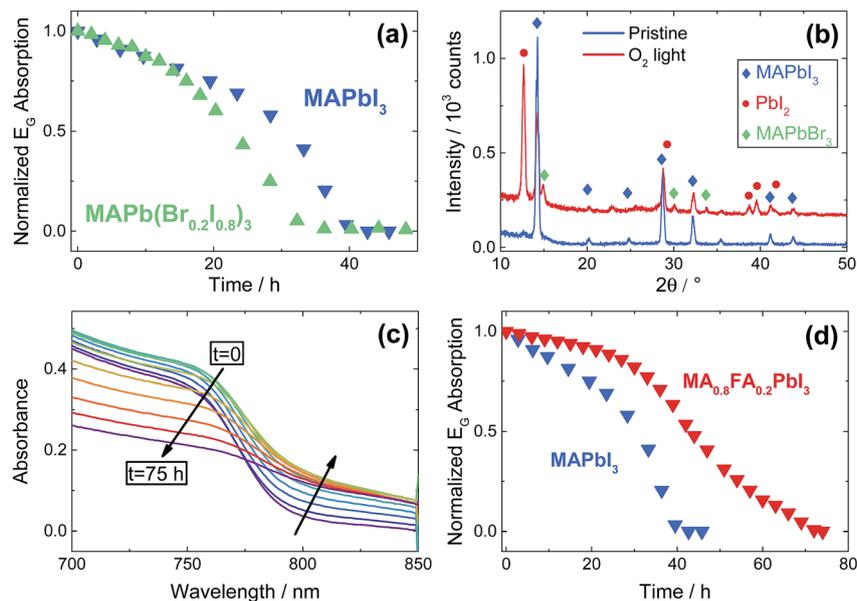


Fig. 5 (a) Comparison between the stability of a pure MAPbI₃ film with a sample containing 20% Br, *i.e.* MAPb(Br_{0.2}I_{0.8})₃. Both films were exposed to O₂ under illumination (4.6 mW cm⁻²) at 333 K and measured by UV-Vis. (b) XRD pattern of the bromine-containing film before and after O₂-induced degradation. Note the appearance of peaks pertaining to the pure MAPbBr₃ phase, that were absent in the pristine film. The presence of weak MAPbI₃ peaks after degradation is because part of the film is shaded from the light by the sample holder, thus avoiding degradation (see Fig. 3b). (c) UV-Vis spectra of (MA_{0.8}FA_{0.2})PbI₃ exposed to O₂ under illumination (4.6 mW cm⁻²) at 333 K. Note the increase of the absorbance at high wavelengths. (d) Comparison between the degradation of pure MAPbI₃ and of (MA_{0.8}FA_{0.2})PbI₃ under O₂ exposure, monitored with UV-Vis under illumination (4.6 mW cm⁻²) at 333 K.

As a conclusion, we end up with the following picture: MAPbI₃ shows a certain oxygen solubility, which would be the stable state only for extremely low oxygen partial pressures. Interestingly, this dissolution is kinetically feasible only under light because otherwise the surface reaction steps are too sluggish (while the bulk diffusion is still acceptable). Oxygen, once introduced, acts as an acceptor dopant owing to O²⁻ replacing I⁻ or being accommodated interstitially. Under realistic conditions, MAPbI₃ is thermodynamically expected to degrade forming water and iodine. Consequently, the degradation can be significantly slowed down by simultaneous mild iodine treatment. In agreement with the above, this reaction proceeds quickly under illumination, while in the dark the surface reaction is sluggish and keeps the material metastable. The kinetic mechanism for both the oxygen incorporation and phase decomposition is expected to involve reactive intermediate such as superoxide species, as well studied for perovskites at high temperatures³⁵ and also in agreement with ref. 10 and 17. Forming mixed anionic compositions with other halides does not substantially improve stability since here the MAPbI₃ portion is selectively and quickly degraded. More encouraging is the cation mixing, where we observed a significant reduction of the degradation rate. Nevertheless, the major tendency still remains and we conclude that under illumination, oxygen exposure is to be avoided to prevent degradation of MAPbI₃ or its mixtures. Naturally, this process must also have a major impact on the performance stability of the related photovoltaic devices, notwithstanding additional effects on the other solar cell components. Hence, full encapsulation or a completely oxygen-free environment is required for a stable performance of perovskite solar cells.

Methods

Experimental methods are given in the ESI Section 1.†

Author contributions

A. S. synthesized the materials and carried out the stability characterizations. T. A. performed the ToF-SIMS measurements. A. S. and G. Y. K. performed the electrical experiments. U. S. and R. M. supervised and discussed the chemical diffusion experiments. J. M. and M. G. supervised and discussed the stability characterization and the electrical measurements. A. S., R. M., G. Y. K. and J. M. wrote the manuscript.

Conflicts of interest

There are no conflicts of interest to declare.

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