Merohedral icosahedral $M_{48}$ ($M = \text{Co}^{II}$, \text{Ni}^{II}$) cage clusters supported by thiacalix[4]arene$^\ddagger$

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Cage clusters are a discrete chemically and topologically diverse family of molecule-based functional materials. Presented here are two isosctructural $M_{48}$ ($M = \text{Co}^{II}$ for \text{LSHU01}, \text{Ni}^{II}$ for \text{LSHU02}) cage clusters with a merohedral icosahedral cage structure featuring 12 M$_4$-TC4A (H$_4$TC4A, $p$-tert-butylthiacalix[4]arene) second building units as vertices and 18 asymmetric 5-(1-tetrazol-1-yl)isophthalate ligands as faces. They are the highest-nuclearity cage compounds of Co$^{II}$ and Ni$^{II}$. The activated Co$_{48}$ cage exhibited high selectivity in the sorption of C$_2$H$_6$ over C$_4$H$_8$ under ambient conditions. Frequency response experiments indicated that the extrinsic voids and matrix interface of the activated crystalline samples are primarily responsible for the observed gas adsorption performance.

Introduction

Polyhedral coordination clusters (PCCs) are an intriguing class of compounds because of their structural aesthetics, interesting physical and chemical properties, and potential application as molecule-based functional materials.$^1$–$^3$ Numerous such species varying in both structure and composition have been reported. Although their nuclearities and overall structures generally cannot be predicted a priori, in many of such species smaller and structurally well-defined complex units are clearly recognizable.$^5$–$^{10}$ For example, the Mn$_{16}$ torus can be formally built by alternating linear [Mn$_4$O$_4$] and cubic [Mn$_8$O$_8$(OMe)$_2$] units;$^11$ the record-setting Fe$_{64}$ cage consists of Fe$_8$ units linked by H$_3$tea (triethanolamine) ligands,$^12$ while the spheroidal Cu$_{36}$ cluster can be formally constructed from metal–organic units.$^13$–$^{14}$ The structural modularity of such PCCs suggests the possibility of assembling even larger polynuclear species by using preformed cluster building units in combination with various bridging ligands, in a stepwise fashion or formally from a one-pot reaction.$^{15}$–$^{17}$ Efforts along this line have indeed produced much success. For example, the largest Co$^{II}$ cluster known (Co$_{36}$) can be viewed as an assembly of a cubane core of Co$_{12}$ clamped by two identical Co$_{12}$ wings via a 2,3-dicarboxypyrazine ligand.$^{18}$

Typical calixarenes and their derivatives are used as multi-dentate ligands for assembly of cage clusters, e.g. Co$_{10}$/Ni$_{10}$ and Ni$_{34}$/Co$_{34}$ cage clusters, that have been reported by Atwood.$^{19}$–$^{23}$ In comparison, $p$-tert-butylthiacalix[4]arene (H$_4$TC4A, Scheme S1a, ESI†) and its derivatives$^{24}$–$^{26}$ with four potentially bridging coordinating groups ([S, SO, and SO$_2$] and four OH groups form tetranuclear M$_4$-thiacalixarene compounds that have been used as secondary building units (SBUs) with appropriate bridging ligands for the construction of more higher nuclearity PCCs including the highest nuclearity Co$_{132}$ and Ni$_{440}$ cages.$^{27}$–$^{30}$ Recently, these PCCs were also found to have new applications in molecule-based materials, and it was demonstrated that the obtained properties highly rely on the cluster shapes and components.$^{31}$–$^{36}$ One interesting observation is that when bridging units of an apparent structural symmetry were used, clusters of regular polyhedral structures including tetrahedral,$^{37}$ cubic,$^{39}$ and octahedral$^{32}$–$^{34}$,38–45 platonic solids were obtained. In comparison, the use of a less symmetric bridging ligand afforded an even larger cage cluster such as the largest known Ni$_{40}$ cage containing 10 Ni$_4$TC4A units.$^{36}$ Although the profound ligand effect is clear, how exactly the ligand dictates the assembly of the clusters eventually formed remains unclear. More such species with various compositions and shapes will help understand how these giant cage clusters may have been assembled. With such an understanding, we report here two isosctructural M$_{48}$ cage clusters of the general formula [M$_{48}$-\{TC4A\}$_{12}$\{L$_{10}$Cl$_{12}$\(H$_2$O)$_6$\}$(+\text{solvents})$ \{\text{LSHU01}, M = \text{Co}; \text{LSHU02}, M = \text{Ni}\} consisting of 12 M$_4$-TC4A units and 18 deprotonated 5-(1H-tetrazol-1-yl)isophthalic acid (H$_2$L, Scheme S1b, ESI†) ligands. They represent the largest known PCCs of Co$^{II}$ and Ni$^{II}$. 

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$^\dag$Electronic supplementary information (ESI) available: Scheme, magnetic susceptibility measurements, MALDI-TOF mass spectra, PXRD, TGA, FT-IR spectra, UV-vis-NIR spectra, additional figures and tables. CCDC 1526121 and 1526122. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8sc03193b
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Results and discussion

Structure of M₄₈ cage clusters

Crystallographic studies (Table S1, ESI†) revealed that LSHU01 and LSHU02 are isomers and there are 2 M₄–TC4A units, 3 L₂ ligands, 2Cl⁻ ions, and an aqua ligand in the asymmetric unit. The tetranuclear unit of M₄–TC4A, with its TC4A ligand adopting a cone conformation, is capped at the bottom by μ₂-Cl⁻ showing a C₄ᵥ symmetry (Fig. 1a). This unit structure is the same as previously reported.²⁴ The expected metric values of the M–O, M–N, and M–S bond lengths (Table S2, ESI†), the charge-balanced composition as determined crystallographically, and the bond valence sum calculations (BVS) indicate that the cobalt and nickel ions are divalent. Both magnetic susceptibility measurements and XPS investigations are consistent with this conclusion (Fig. S1, ESI†).

Further structural analysis revealed subtle differences between the two M₄–TC4A units in the asymmetric unit with the one (SBU-1, Fig. 1c) containing five L₂ ligands while the other (SBU-2, Fig. 1d) has four L₂ ligands. The L₂ ligands in SBU-1 exhibit two different coordination modes: three adjacent L₂ ligands each use one of its two carboxylate groups to bridge the four metal ions, while the remaining two L₂ ligands each use one of the N atoms in its tetrazolyl group to coordinate in a monodentate fashion as reported in the literature (Fig. S2, ESI†). In comparison, SBU-2 can be viewed as being derived from SBU-1 by replacing one of the two L₂ ligands that use the tetrazolyl N atom for coordination with an aqua ligand. SBU-1 and SBU-2 units, 6 of each kind, are organized into an icosahedron with the tetranuclear units occupying its 12 vertices (Fig. 1e). A more careful analysis of the structure revealed that the SBU-2 units are in an octahedral arrangement, while the SBU-1 units are in a chair-like arrangement, much like the more stable conformation of cyclohexane. Unlike in the arrangement of the SBU-1 units, the three adjacent SBU-2 units are not directly linked by an L₂ ligand. Rather, they are hinged by three L₂ ligands to form a metallamacrocycle with three aqua ligands, one on each SBU-2 unit, disposed within the ring structure (Fig. S3, ESI†). As such, the overall structure of the cage cluster can be best described as a merohedral convex icosahedron with 18 triangular faces rather than a regular icosahedron with 20 triangular faces (Fig. 1e). We note that this structure type is rarely observed in metal–organic systems despite its common existence in pure inorganic compounds including B₁₂, Keggin-type polyoxometalates, and metal-centered endedohedral clusters.²⁷–³⁰ With the assembly of UO₂⁺ and calix[5]arene-carboxylate being viewed as a dodecahedron,³¹ the present M₄₈ cage clusters complete the five structural types of platonic solids in metal-calixarene chemistry (Fig. S4, ESI†).²⁹,³³–³⁶ MALDI-TOF mass spectra confirmed the presence of the Co₄₈ cluster (Fig. S5, ESI†).

The cage cluster has an outside dimension of 37.3 × 35.8 × 33.8 Å³ ([C₄butyl-C₄butyl]) with an inner cavity ([Cl···Cl⁺]) of 22.3 × 20.8 × 20.8 Å³; the latter is ca. 1.5 times larger than that of the Ni₄₀ cage,³⁰ although the former is similar. The solvent-accessible voids of LSHU01 and LSHU02 were estimated to be 54 653.5 Å³/64 029.6 Å³ per cell 96 152.0 Å³ and 106 649.0 Å³, respectively, corresponding to 56.8%/60.0% of the total crystal volume. It is plausible that the different estimated solvent-accessible voids between the two isocrystallographic cage clusters are due to the presence of different solvent molecules as well as the different collecting temperatures of the crystallographic data. That the hollow cages have two opposite triangle windows between two metal-lamacrocyclic rings partly occupied by aqua ligands is interesting. The edge of the triangle is ca. 7.6 Å with an in-circle diameter of ca. 5.4 Å (O_water···O_water, Fig. S3, ESI†). Furthermore, irregular pockets between three cages are also observed with dimensions ranging from 4.0 to 11.7 Å (C₄butyl···C₄butyl, Fig. S6, ESI†). By packing the cages along the c axis, one can find irregular channels with permanent pores (pore A, within cages) and packing pores (pore B, between adjacent cages) in ABBA mode (Fig. 2, S7 and S8, ESI†). The porous structure is expected to be useful for the uptake of gas molecules upon activation by desolvation.

Stability and porosity

Both PXRD and thermogravimetric analysis indicated that LSHU01 and LSHU02 lost lattice solvent molecules between room
samples were placed at ambient temperature for several hours before pictures and morphology corresponding to different states. The maintenance of structural integrity (Fig. 3 and S11) suggests complete removal of organic solvents and LSHU01 Weight loss in the range of 150°C found: 3.1%.

of the coordinated water molecules and Cl atoms (calc: 3.3%, found: 3.1%).

In order to assess the porosity of the M48 cages, argon adsorption measurements were performed at 87 K using temperature and 80 °C under vacuum (Fig. S9–S11, ESI†). Negligible weight loss in the temperature range of 80–150 °C for LSHU01 suggests complete removal of organic solvents and maintenance of structural integrity (Fig. 3 and S11–S13, ESI†). Weight loss in the range of 150–250 °C can be attributed to loss of the coordinated water molecules and Cl atoms (calc: 3.3%, found: 3.1%).

In order to assess the porosity of the M48 cages, argon adsorption measurements were performed at 87 K using LSHU01′ and LSHU01′′ (LSHU01 activated at 80 °C and 150 °C, respectively), for example LSHU01 and LSHU02 are isostuctural and the cage clusters showed similar outside dimensions and inner cavities (Fig. S14, ESI†). The gas sorption showed a non-reversible type-I isotherm with large hysteresis upon desorption of LSHU01′. Such a behavior is probably due to the presence of the voids of the cage cluster and the extrinsic voids between cages as evidenced by the broad pore size distributions (PSDs) (0.50–1.10 nm with two peaks at ca. 0.51 nm and 0.65 nm, respectively, Fig. S14, inset†). A typical reversible type-I isotherm was obtained using LSHU01′′; its PSD is similar to that of LSHU01′ but the decrease of surface areas at 0.65 nm suggested different microstructures of these two activated samples. The observation of the PSD at ca. 0.51 nm that can be ascribed to the cage windows for both LSHU01′ and LSHU01′′ suggested the undamaged framework of cage clusters after higher temperature evacuation. For LSHU01′, the Langmuir and Brunauer–Emmett–Teller (BET) surface areas are estimated to be 803.09 m² g⁻¹ and 582.02 m² g⁻¹, while for LSHU01′ the corresponding values are 607.95 m² g⁻¹ and 438.55 m² g⁻¹, respectively. The decreases in micropore areas (from 411.94 to 360.88 m² g⁻¹, 12.39%), external surface areas (from 170.08 to 77.67 m² g⁻¹, 54.33%), and micropore volume (from 0.1776 to 0.1600 cm³ g⁻¹, 9.91%) for LSHU01′ with respect to LSHU01′′ are attributed to the decrease of free space between cages caused by the more compact packing of cages evacuated at higher temperature, consistent with Horváth–Kawazoe (HK) and Barrett–Joyner–Halenda (BJH) absorption PSDs (Fig. S15 and S16, ESI†).

Gas absorption and separation
The CH₄, C₂H₄, C₂H₆, C₃H₈, and CO₂ uptake experiments were conducted at 295 K in the pressure range of 0–100 kPa using LSHU01′ and LSHU01′′ (Fig. 4 right and Fig. S17–S23, ESI†). For CH₄, the absorption isotherms are almost identical and can be attributed to loss of the coordinated water molecules and Cl atoms (calc: 3.3%, found: 3.1%).

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For binary mixtures composed of equimolar C₂-hydrocarbons or CO₂ and CH₄, the selectivity is below 30 for both LSHU01₀ and LSHU01₀₀ (Fig. 4 right), which is comparable to many systems previously reported. However, the initial selectivity for the C₃H₈/CH₄ mixture was 140.1 and 70.5 and reached 912.6 and 189.8 at 295 K and 1 bar for LSHU01₀ and LSHU01₀₀, respectively (Fig. S22 and S23, ESI†). It is highlighted that the selectivity for C₃H₈/CH₄ under ambient conditions of LSHU01₀ is about 30 fold that of the closed Ni₄₀ cage and the best performance among metal–organic polymers (Table S3, ESI†). The extremely high selectivity values for C₃H₈/CH₄ strongly suggest that it is feasible to separate the pair in a vacuum swing adsorption process using the activated Co₄₈ sample as an adsorbent.

Frequency response
To understand why LSHU01₀ showed higher selectivity than LSHU01₀ in the adsorption of C₃H₈ over CH₄ at 295 K, frequency response (FR) spectra of C₃H₈ were obtained in both in-phase (IP) and out-of-phase (OP) FR signals of C₃H₈ can be detected using both LSHU01₀ and LSHU01₀₀ (Fig. 5), but not for CH₄ (Fig. S24, ESI†). The OP curves of C₃H₈ spectra, well fitted by the Yasuda sorption theoretical model, correspond to three adsorption processes (P₁–P₃) for LSHU01₀ and LSHU01₀₀, which can be ascribed to the adsorption in the molecular cages, between cages, and on the interface of the matrix from low frequency to high frequency, respectively. The reduction of the response intensity of IP and OP for LSHU01₀₀ versus LSHU01₀ indicates the lower absorption of C₃H₈ molecules, which is consistent with the decrease in BET surface areas and the C₃H₈-absorption experiments. Furthermore, the fitted OP response peaks P₁–P₃ are shifted to lower frequency (from 0.03 to 0.02 Hz for P₁, from 0.22 to 0.16 Hz for P₂, and from 0.74 to 0.97 Hz for P₃) for LSHU01₀ with respect to LSHU01₀, pointing to the lower absorption rate of C₃H₈ molecules. The observations detailed above together indicate that the free space between the cages and matrix interface plays an important role in the uptake and separation of C₃H₈ over CH₄ under ambient conditions. It is also concluded that the absorption properties are sensitively dependent on the pore size, the moderate surface areas in accordance with gas kinetic diameters, and/or the polarizability of the gas molecules.

Conclusions
In summary, we have obtained and structurally characterized two record-setting high-nuclearity M₄₈ (M = Co, Ni) cage clusters. It has been shown that higher-nuclearity Co and Ni cages can be built by using M₄₄-TC₄₄ A SBUs and asymmetric ligands. The merohedral icosahedron-type structures with large pores and windows allow us to investigate gas sorption properties of the crystalline materials. The activated Co₄₈ crystals were proved to be a competitive candidate for separating C₃H₈ from CH₄. The activation temperature for Co₄₈ crystals has a profound effect on the extrinsic voids between the cage clusters which are believed to be primarily responsible for the observed absorption properties and selectivity. Efforts aimed at the synthesis of even higher-nuclearity cage clusters are ongoing.

Experimental section
Materials and measurements
p-tert-Butylthiacalix[4]arene (H₄TC₄₄ A)⁶⁷ and 5-{1H-tetrazol-1-yl} isophthalic acid (H₄L)⁶⁸ were synthesized according to...
a literature method, respectively. Other reagents were purchased from commercial sources and used as received. TGA was performed on a Perkin Elmer Pyris1 TGA thermogravimetric analyzer. FT-IR spectra using KBr pellets were taken on a Perkin Elmer Spectrum GX spectrometer. UV-vis spectra were recorded on an Agilent Cary 5000 spectrometer. MALDI-TOF mass data were collected on a Bruker Smartbeam MALDI-TOF mass spectrometer. CHNS elemental analysis was performed on an EA1 CE-440 instrument. Magnetic susceptibility measurements were performed on a Quantum Design MPMS XL-5 SQUID system in the temperature range of 2–300 K, and diamagnetic corrections for the sample and sample holder were applied. X-ray photoelectron spectroscopic (XPS) measurements were carried out on an ESCALAB 250Xi using a monochromatic Al Kα X-ray source (1486.6 eV).

**Synthesis of LSHU01 and LSHU02**

H₂L was added to a suspension of MCl₂·6H₂O (M = Co, Ni) (0.95 g, 0.4 mmol) and H₄TC4A (0.09 g, 0.125 mmol) in a m:n (v/v) CHCl₃-CH₃OH mixed solvent (total 10 mL, m = 8, n = 2 for Co; m = 5, n = 5 for Ni) with stirring for 10 min and then transferred into a 20 mL Teflon-lined autoclave which was kept at 130 °C for 3 days and then slowly cooled to room temperature at about 4 °C/h. Red/green crystals were isolated by filtration and then washed with m:n CHCl₃-CH₃OH and dried in a vacuum at room temperature. Yield (0.094 g and 0.043 g): ca. 55% and 26% with respect to H₄TC4A for LSHU01 and LSHU02, respectively. Elemental analysis: calculated (%) for [M₄₈(C₄₀H₄₄S₄O₄)₁₂(C₉H₄N₄O₄)₁₈Cl₁₂(H₂O)₆]: M (Co, Ni) 80.71073, H 3.82, N 6.25, S 9.54; M (Ni, Co) 80.71073, H 3.82, N 6.25, S 9.54; M (Ni) 80.71073, H 3.82, N 6.25, S 9.54; found: M (Co, Ni) 80.71073, H 3.82, N 6.25, S 9.54; M (Ni, Co) 80.71073, H 3.82, N 6.25, S 9.54; M (Ni) 80.71073, H 3.82, N 6.25, S 9.54 (after being dried in a vacuum at 80 °C).

**Single crystal X-ray diffraction**

The intensity data were recorded on a Bruker D8 QUEST system with Mo-Kα radiation (λ = 0.71073 Å). The crystal structures were solved by means of direct methods and refined by employing full-matrix least squares on F² (SHELXTL-2014).39 Even the low temperature data set obtained at about 100 K for the compound LSHU01 reveals severely disordered solvents within the lattice interstices that are difficult to interpret thus complicating efforts to give precise estimates of the molecular formula. The diffraction data were treated by the “SQUEEZE” method as implemented in PLATON (see ESI† for details).39 All non-hydrogen atoms were refined anisotropically, and hydrogen atoms of the organic ligands were generated theoretically on the specific atoms and refined isotropically with fixed thermal factors. Since the crystals do not diffract very well at high angles due to the structural disorder, the determined 2θ is 22.0141° and 20.1027° for LSHU01 and LSHU02, respectively. The θ factors in the final structural refinement are also relatively large but typical of such systems.35,36,40 Refinement parameters and crystallographic data, selected bond distances and BVS calculations for LSHU01 and LSHU02 are shown in Tables S1 and S2 in the ESI† respectively. CCDC 1526121 and 1526122 contain the supplementary crystallographic data of LSHU01 and LSHU02 for this paper, respectively.

**Gas adsorption experiments**

Ultra-high-purity grade gas (>99.99) was used throughout the adsorption experiments. The measured sorption isotherms have been recorded at least two times to confirm the reproducibility within experimental errors. About 150 mg of methanol solvent-exchanged Co₄₈ samples were activated at 80 °C and 150 °C for 10 hours according to the TPD experiment, respectively. Low-pressure gas sorption experiments are carried out on Micromeritics ASAP 2020M automatic volumetric instruments for Ar and on an Intelligent Gravimetric Analyser (IGA-003 Hiden Analytical Ltd., Warrington, UK.) for CO₂, CH₄, C₂H₄, C₂H₆ and C₃H₈, respectively. Buoyancy corrections for the samples were applied for gravimetric measurements. Desorption was achieved by placing the samples under a dynamic vacuum at 80 °C or 150 °C for three hours, respectively. Ar isotherms were measured using a liquid argon bath (87 K). Other gas isotherms, e.g. for CO₂, CH₄, C₂H₄, C₂H₆ and C₃H₈ were measured at 273 K and 295 K, respectively. The specific surface areas are determined using the Brunauer–Emmett–Teller and Langmuir equations from the Ar sorption data. The pore size distribution was obtained from the DFT, HK, and BJH models in the Micromeritics ASAP 2020 software package based on the Ar sorption at 87 K.

**Frequency response measurements**

Frequency response measurements were carried out on a high-accuracy differential Baratron pressure transducer (MKS 698A11TRC). The frequencies were controlled using an on-line computer, which was also used for the acquisition of the pressure data from the Baratron. An accurate amount of sample (ca. 30 mg) was scattered in a plug of glass wool and degassed under a high vacuum (<10⁻³ Pa) at 80 °C or 150 °C for 6 h. Prior to the measurement, the sorbate vapor from the supply side of the vapor reservoir was admitted to the sorption chamber and equilibrated over the sample at a certain pressure.81

**Conflicts of interest**

There are no conflicts to declare.

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Notes and references


