Large-bite diboranes for the $\mu$(1,2) complexation of hydrazine and cyanide†

Chang-Hong Chen and François P. Gabbai*  

Received 24th April 2018  
Accepted 27th May 2018  
DOI: 10.1039/c8sc01877d  
rsc.li/chemical-science

Introduction

The chemistry of main group-based polydentate Lewis acids has drawn considerable attention over the past decades, leading to applications in anion sensing, anion transport, small molecule activation and catalysis. Diboranes featuring the rigid 1,8-naphthalenediyI or ortho-phenylene backbones are the most studied examples of such systems. Owing to the short spacing of the two boron centres, these diboranes are well-suited for the chelation of monoatomic anions such as hydrazine, fluoride and chloride or polyatomic anions amenable to $\mu$(1,1) ligation such as azide. The strength of the diborane host-anionic guest interaction in these systems has led to the development of application in selective anion sensing as well as catalysis. Although no structural evidence has been obtained for the ditopic complexation of neutral molecules by diboranes of type A and B, it has been demonstrated recently that 9,10-dihydroanthracenes of type C are able to engage 1,2-diazines in complex formation and catalyse their Diels Alder chemistry. The ability of C to complex 1,2-diazines is reminiscent of the adduct formed by 2,2'- diborabiphenyl and pyridazine.

Bearing in mind that the selectivity of these diboranes for their respective guests is dictated by the size match that exists between the host and the guest, we have recently become interested in diboranes with an increased separation between the Lewis acidic centres. It occurred to us that such systems may display a different selectivity and may become well adapted to the $\mu$(1,2) chelation of diatomic molecules. We also speculated that the rigidity of the targeted diboranes may inform the molecular recognition properties of these bidentate Lewis acidic hosts.

† Electronic supplementary information (ESI) available: Experimental details, NMR spectra of novel compounds, crystallographic data, absorption and emission spectra, DFT optimization results. CCDC 1838442, 1827003-1827005. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8sc01877d
In this article, we report a series of results concerning the synthesis and properties of bidentate diboranes in which the two boron atoms are separated by more than 4.5 Å. We also show that this increased separation allows for the selective $\mu(1,2)$ complexation of hydrazine and cyanide, two diatomic molecules that are known for their high toxicity.

**Results and discussion**

**Synthesis, characterization and properties of the diboranes**

We have recently described the synthesis of 1,8-bis(dimesitylboryl)biphenylene (1), a dibaborane in which the two boron atoms are separated by 4.566(5) Å. To assess the influence of the backbone over the properties of such large-bite diboranes, we have now decided to prepare its 1,8-triptycenediyl analogue (2). By analogy with the approach we employed to access 1, diborane 2 was obtained via the dilithiation of 1,8-dibromotriptycene, followed by metathesis with Mes$_2$BF (Scheme 1).

Compound 2 is a colourless air-stable solid which has been isolated in moderate yield. The $^1$H NMR signals of the triptycene backbone are consistent with a symmetrical structure. The detection of six sharp signals arising from the methyl groups and the four meta- H signals of the mesityl groups suggest that the molecule retains mirror symmetry in solution. The $^{11}$B NMR spectrum shows a broad peak at 73.0 ppm which falls within the typical range expected for triarylboranes. The cyclic voltammogram (CV) of compound 2 shows two quasi-reversible reduction waves at $-2.62$ and $-3.00$ V vs. Fe/Fe$^+$, suggesting that the molecule can be reduced by two electrons (Fig. 1c). Compared to the biphenylene derivative 1 for which the two reduction waves are observed at $E_{1/2} = -2.23$ and $-2.74$ V (Fig. 1c), the first reduction of 2 is shifted toward cathodic potentials by 390 mV. This observation indicates that the biphenylene backbone is substantially more electron withdrawing. This conclusion is in agreement with two previous studies dealing with related biphenylene boron derivatives. Also, the shorter separation observed between the first and second reduction wave in 2 ($\Delta E_{1/2} = 0.51$ V for 1 vs. $\Delta E_{1/2} = 0.38$ V for 2) signals a decreased electronic communication between the boron atoms, an effect that we assign to the absence of extended conjugation in the triptycene backbone. The importance of electronic communication in 1 is further supported by the fact that the $\\Delta E_{1/2}$ measured for 1 (0.51 V) is comparable to the largest value measured for naphthalene-based diboranes ($\Delta E_{1/2}$ range = 0.3–0.52 V), a class of compounds in which the two boron atoms are separated by only 3.002–3.385 Å.

Computational studies show that the lowest unoccupied molecular orbital (LUMO) of 1 (Fig. 1a) is dominated by the two boron p orbitals which are in conjugation with the biphenylene $\pi$ system. The makeup of this LUMO underscores the effective electronic communication that exists between the two boron centres. In the case of 2, which lacks mirror symmetry because of steric repulsions between the mesityl substituents, the LUMO bears a larger contribution from the p$_a$ orbital of one boron atom. The LUMO + 1, which lies only 0.14 eV above the LUMO, shows a large contribution from the opposite boron atom (see ESI†). More importantly, $\pi$-conjugation between the two boron atoms in 2 is interrupted by the sp$^3$ triptycene bridgehead carbon atoms. These features corroborate the conclusion that the two boron atoms of 1 are in closer electronic communication than in 2, as confirmed from the CV measurements.

Single crystals of compound 2 could be obtained by layering a solution of 2 in CH$_2$Cl$_2$ with MeOH. The solid-state structure of 2 shows that the boron atoms adopt a trigonal planar geometry as indicated by the sum of the C$_{aryl}$–B–C$_{aryl}$ angles ($\sum C-B1-C = 359.7^\circ$, $\sum C-B2-C = 359.4^\circ$) (Fig. 2). The B1–B2 separation of 5.559(4) Å in 2 is notably larger than that in 1.

Reactions of the diboranes with neutral diatomic molecules

With 1 and 2 at our disposal, we became eager to investigate the behaviour of these two compounds in the presence of simple diatomic molecules including hydrogen peroxide, hydroxylamine and hydrazine. While complexes containing doubly coordinated peroxide\(^{16a,b}\) and hydrazine\(^{16b}\) molecules have generated much interest in the context of energy related research, we also note that coordinated hydrazine has been identified as a possible intermediate in nitrogen fixation reactions.\(^{16,17}\) Because of its acute toxicity, there are also numerous ongoing efforts aimed at developing selective receptors for molecular recognition and sensing applications.\(^{20}\)

Inspired by the recent discovery that perfluorinated boranes can doubly complex the peroxide and superoxide anions,\(^{21}\) we first tested the reaction of 1 and 2 with 30% H\(_2\)O\(_2\)(aq). Unfortunately, we observed decomposition, illustrating the vulnerability of these non-perfluorinated diboranes under oxidative conditions. We also tested the reaction of 1 and 2 with 50% NH\(_3\)OH\(_{aq}\), which also resulted in decomposition of the hosts. By contrast, reaction of diborane 1 with hydrazine monohydrate (N\(_2\)H\(_4\)\_H\(_2\)O) in THF proceeded smoothly to afford the corresponding hydrazine complex 1-\(\mu_2\)-N\(_2\)H\(_4\) as an air- and moisture-stable off-white solid (Scheme 2). This reaction results in a distinct loss of the yellow colour indicating coordination of the hydrazine to each boron atom. The same effect is responsible for the observed quenching of the green fluorescence (Fig. 3). In stark contrast, diborane 2 does not complex hydrazine under the same condition, even with a large excess of N\(_2\)H\(_4\)\_H\(_2\)O. The lack of reactivity of 2 toward hydrazine may be correlated to its lower Lewis acidity as suggested by the CV measurements. It is also possible that the

---

**Fig. 2** Solid-state structure of 2. Thermal ellipsoids are drawn at the 50% probability level. The hydrogen atoms are omitted for clarity.

**Fig. 3** UV-vis absorption (left) and emission (right) spectra of 1 (solid line) and 1-\(\mu_2\)-N\(_2\)H\(_4\) (dash line) (4.52 \(\times\) 10\(^{-5}\) M, \(\lambda_{ex} = 370\) nm) in THF. The pictures of the solutions were taken at the same concentration. The images showing the fluorescence emission were illuminated with a hand-held UV lamp.

**Scheme 2** Synthesis of 1-\(\mu_2\)-N\(_2\)H\(_4\) and the reaction of 1-\(\mu_2\)-N\(_2\)H\(_4\) with benzaldehyde. (a) 2.5 equiv. N\(_2\)H\(_4\)\_H\(_2\)O, THF, rt. (b) 1 equiv. benzaldehyde, CDCl\(_3\), 60 °C.
increased spacing between the two boron atoms, the lower rigidity of the backbone and the presence of a bridgehead methine group interfere with the ability of 2 to bind hydrazine.

The $^1$H NMR spectrum recorded for the resulting hydrazine adduct 1-$\mu_2$N$_2$H$_4$ in CD$_2$Cl$_2$ features six methyl-proton and four meta-proton resonances, indicating that the complex adopts C$_2$ symmetry in solution. The most conspicuous evidence for the formation of a B-NH$_2$-NH$_2$-B bridge is the detection of two broad multiplets at 5.15 and 6.29 ppm, indicating that the B-NH$_2$-NH$_2$-B unit gives rise to an AA'BB' spin system. These spectroscopic features also indicate that the conformation of the hydrazine molecule in the diboron pocket is rigidly locked in solution. The solid-state structure of 1-$\mu_2$N$_2$H$_4$ offers a consistent picture (Fig. 5). This neutral complex crystallizes in the monoclinic C2/c space group with a half molecule in the asymmetric unit. The hydrazine nitrogen atoms are coordinated to the boron atoms via a B-N bond length of 1.688(2) Å which is longer than that in H$_2$B-NH$_2$-NH$_2$-BH$_3$ (1.609 Å) but comparable to that recently reported by Szymczak for complex D-$\mu_2$N$_2$H$_4$ (1.697(2) Å and 1.698(2) Å) (Scheme 3).

The N–N bond length of 1.469(2) Å is similar to the N–N bond length determined previously in the literature. This observation clearly indicates that diborane 1 is well suited for the chelation of hydrazine molecule. We also note that D can bind two hydrazine molecules, one at each boron atom. This is not the case of diborane 1 which only forms a 1 : 1 complex, even in the presence of an excess of hydrazine. This difference in behaviour illustrates the selectivity imparted by the rigid preorganization of the two boron atoms in 1.

To further establish the high affinity that 1 displays for hydrazine, we have studied the complexation reaction under dilute conditions. We observed that diborane 1 is capable of capturing hydrazine from dilute aqueous solutions under biphasic conditions. Quantitative formation of 1-$\mu_2$N$_2$H$_4$ was observed by $^1$H NMR spectrometry when a solution of 1 in CH$_2$Cl$_2$ (1 mL, [1] = 7.7 mM) was stirred for 24 hours with an aqueous solution (2.5 mL) containing 1 weight% of N$_2$H$_4$. Complex 1-$\mu_2$N$_2$H$_4$ is remarkably stable. It can be stored in air and shows no evidence of decomposition for months. It is also resistant to reactions with aldehydes. For example, 1-$\mu_2$N$_2$H$_4$ failed to react with benzaldehyde in CDCl$_3$ at room temperature over the course of 20 h. However, upon heating to 60 °C, a clean, yet slow reaction is observed leading to the quantitative formation of the free diborane 1 and benzaldehyde hydrazone after 20 h as confirmed by $^1$H NMR spectroscopy (Scheme 2).

**Reaction of the diboranes with the cyanide anion**

Next in our survey of the properties of these diboranes, we decided to focus on the case of anions and especially the cyanide anion which is known for its acute toxicity and which has often been considered as a target in molecular recognition assays, including using boron-based hosts.

Both diboranes quickly react with KCN in a CH$_2$Cl$_2$/MeOH (1/1 vol) mixture containing dibenzo-18-crown-6. As expected, this reaction is accompanied by a quenching of both the absorption and emission band (Fig. 4). In the case of 1, these changes result in a distinct loss of both the yellow colour and the green fluorescence of the starting diborane (Fig. 4a). Workup of these reactions afforded the corresponding cyanide complexes [1-$\mu_2$CN] and [2-$\mu_2$CN] as air-stable [K(dibenzo-18-crown-6)] salts (Scheme 4). The $^{11}$B NMR signals detected for these anionic complexes (−15.8 ppm for [1-$\mu_2$CN] and −11.7 ppm for [2-$\mu_2$CN]) are consistent with the existence of four-coordinate boron atoms. These signals are somewhat broad, possibly indicating the presence of two closely spaced and overlapping resonances corresponding to the N$_{CN}$-bound and C$_{CN}$-bound boron atoms, respectively. The presence of the cyanide anion in these complexes is confirmed by IR spectroscopy which shows that, in both cases, the cyanide stretching frequency ($v_{CN}$ = 2229 cm$^{-1}$ for 1 and 2184 cm$^{-1}$ for 2) is higher than that of KCN ($v_{CN}$ = 2158 cm$^{-1}$). The higher energy of these vibrations is typical of cyanoborates. This effect can be rationalized by invoking a stabilization of the π molecular orbitals of the cyanide anion. The absence of π-backbonding can also be invoked as a cause for this effect. The greater $v_{CN}$ observed for [1-$\mu_2$CN] can be corroborated to the CV results and again suggest that the biphenylene backbone is more electron withdrawing than the triptycene backbone. The methyl region of the
H NMR spectrum recorded for both [1-μ2-CN]− and [2-μ2-CN]− as [K(dibenzo-18-crown-6)] as [K(dibenzo-18-crown-6)]+ salts, no interaction with HCO3 or N terminus of the cyanide anion.

Interestingly, diborane 1 and 2 appear to be highly selective for cyanide since no interaction with HCO3, HSO4, H2PO4, CH3COO, Cl, Br, I, and N3− is observed in CHCl3/MeOH (1/1 vol) by UV-vis spectrometry. Boranes are known to also react with LiHBEt3 in dry CDCl3.

The structures of the cyanide complexes as [K(dibenzo-18-crown-6)]+ salts have been investigated by single crystal X-ray diffraction which confirms the μ(1,2) chelation of the cyanide anions (Fig. 6). Both structures were solved on the basis of a model in which the cyanide anion is disordered over two overlapping head-to-tail positions, which as indicated by the refinements, contribute almost equally to the observed structures. In both structures, the boron atoms are distinctly pyramidal as indicated by the sum of the Caryl-B-Caryl angles which fall in the 341.7(4)-343.2(4)° range for [1-μ2-CN]− and 342.0(5)-344.6(5)° for [2-μ2-CN]−. The accuracy of the crystallographic measurements does not allow for a comparison of the B-C=NC bond and B-N=CN bond distances which fall in the 1.54(4)-1.62(3) Å for [1-μ2-CN]− and 1.63(5)-1.67(8) Å for [2-μ2-CN]−. The small difference in the B1-B2 separations in [1-μ2-CN]− (4.187 (4) Å)
and 1 (4.566(5) Å) reflects the rigidity of the molecule. Compared to 1, the large variation of the B1–B2 separation on going from 2 (B1–B2 5.559(4) Å) to [2-µ2-CN]– (B1–B2 4.321(5) Å) speaks to the flexibility of the triptycene backbone and its pincer ability toward the cyanide anion. The behaviour of 1 and 2 and their ability to coordinate to both ends of the cyanide anion is reminiscent of the behaviour of dicopper(II) complexes and their ability to coordinate to both ends of the cyanide anion. The behaviour of 1 and 2 is a stronger Lewis acid than 2, and we became eager to experimentally confirm that 1 would indeed outcompete 2 with regard to cyanide binding. In accordance with this prediction, we observed, using 1H NMR spectroscopy, the quantitative transfer of the cyanide anion from 2 to 1 when [Bu4N][2-µ2-CN] was mixed with an equimolar quantity of 1 in CDCl3/CD3OD (1/1 vol) at 60 °C over the course of 12 h (Scheme 6). This observation confirms the higher cyanide ion affinity of 1 in protic solution, which is consistent with the electrochemical and IR measurements described above. Furthermore, a competition experiment was also performed with the monofunctional model borane Mes2BPh (3). In these tests, the cyanide anion was transferred from [3-CN]– to 1 quantitatively in CDCl3/CD3OD (1/1 vol) at 60 °C (Scheme 6). However, no transfer was observed between [3-CN]– and 2. The absence of a reaction points to the weaker cyanide affinity of 2 when compared to 1. The reversibility of the cyanide binding was also investigated by allowing [Bu4N][1-µ2-CN] and [Bu4N][2-µ2-CN] to react with one equivalent of C6F5 in CDCl3/CD3OD (1/1 vol). Under these conditions, [2-µ2-CN]– was readily converted into 2 while [1-µ2-CN]– remained untouched, again supporting the superior Lewis acidity of 1.

Scheme 5 Structure of the cyanide complexes of ditopic hosts E–G.

Scheme 6 Competition reactions of [Bu4N][2-µ2-CN] and 1 (top) and [Bu4N][3-CN] and 1 (bottom). The [Bu4N]+ cations are not shown for clarity.

Conclusions

In summary, we describe two “large-bite” diboranes that are ideally suited for the selective µ(1,2) complexation of hydrazine and cyanide, two diatomic molecules that are known for their high toxicity. The high selectivity displayed for these two specific molecules highlights the defining role played by the backbone. The most potent binder is the biphenylene-based diborane which complexes both hydrazine and cyanide while the triptycene derivative only binds cyanide. A similar picture emerges from competition experiments which show that 1 is a better molecular recognition unit for cyanide than 2. Our results and analyses indicate that the enhanced properties of 1 arise from the electron withdrawing nature of the biphenylene backbone. We also propose that the rigidity of this diborane is a favourable factor that helps sequester hydrazine and the cyanide anion in the diboron pocket. This binding appears to be irreversible in the case of the cyanide anion. In the case of the hydrazine complex, we observe that slow release can be triggered upon elevation of the temperature and in the presence of a reaction partner such as benzaldehyde. This last feature shows that such bidentate hosts could be used for the slow release of reactive compounds.

Experimental

Synthesis of 2

BuLi (2.65 M, 1 mL, 2.65 mmol) was slowly added to a solution of 1,8-dibromotriptycene (500 mg, 1.2 mmol) in dry THF (8 mL) under N2 at −78 °C. After 30 min of stirring at low temperature, a solution of Mes2BF (715 mg, 2.6 mmol) in THF (5 mL) was slowly transferred into the reaction flask using a cannula. The resulting solution was stirred for an additional 12 h at room
temperature. The solution was then treated with saturated NH4Cl(aq) (1 mL) and the solvent was removed under vacuum. The resulting white solid was then dissolved in CH2Cl2 (5 mL). The resulting solution was filtered and brought to dryness under vacuum. The off-white solid was then washed with MeOH (15 mL) twice and dried under vacuum to afford diborane 2 as a white powder in 60% yield. 1H NMR (399.50 MHz, 25 °C, CDCl3): δ 7.49 (d, 2H, 3H-J = 6.6 Hz, triptycene-CH), 7.23 (d, 1H, 3H-J = 7.0 Hz, triptycene-CH), 6.97 (pseudo t, 2H, 3H-J = 7.4 Hz, triptycene-CH), 6.90–6.86 (m, 3H, triptycene-CH), 6.82 (s, 2H, mes-CH), 6.77 (s, 2H, mes-CH), 6.73 (s, 2H, mes-CH), 6.61 (pseudo t, 1H, 3H-J = 7.4 Hz, triptycene-CH), 6.35 (s, 2H, mes-CH), 5.40 (s, 1H, triptycene-CH), 5.30 (d, 1H, 3H-J = 7.4 Hz, triptycene-CH), 5.16 (s, 1H, triptycene-CH), 2.34 (s, 6H, –CH3), 2.30 (s, 6H, –CH3), 2.17 (s, 6H, –CH3), 2.13 (s, 6H, –CH3), 1.79 (s, 6H, –CH3), 0.60 (s, 6H, –CH3) ppm. 13C NMR (125.61 MHz, 25 °C, CDCl3): δ 150.86, 145.50, 145.23, 144.42, 144.27, 143.12, 141.52, 141.17, 140.90, 140.77, 139.89, 138.58, 138.57, 137.13, 129.92, 129.90, 128.47, 128.14, 127.90, 126.58, 126.17, 125.05, 124.16, 123.50, 121.60, 55.45 (triptycene-CH), 54.26 (triptycene-CH), 25.20 (24.96 (–CH3), 22.37 (–CH3), 22.03 (–CH3), 21.39 (–CH3), 21.25 (–CH3) ppm. 11B (128.16 MHz, 25 °C, CDCl3): δ 73.00 (br) ppm. Elemental analysis calculated (%) for C56H56B2: C, 21.39; H, 2.30 (s, 6H, –CH3), 2.17 (s, 6H, –CH3), 2.13 (s, 6H, –CH3), 1.79 (s, 6H, –CH3), 0.60 (s, 6H, –CH3) ppm. 13C NMR (100.46 MHz, 25 °C, CDCl3): not observed.

Synthesis of 1H-μ2-NH4

A solution of N2H4·H2O (20 mg, 0.4 mmol) in THF (2 mL) was combined with a solution of diborane 1 (100 mg, 0.154 mmol) in THF (5 mL) at room temperature. The resulting solution was stirred for an additional hour and brought to dryness under vacuum. The resulting solid was washed twice with MeOH (5 mL) to afford 1H-μ2-NH4 as a white-off powder in 90% yield. 1H NMR (499.48 MHz, 25 °C, CD2Cl2): δ 6.85 (s, 2H, mes-CH), 6.80 (s, 2H, mes-CH), 6.68 (s, 2H, mes-CH), 6.57 (s, 2H, mes-CH), 6.55–6.50 (m, 6H, biphenylene-CH), 6.31–6.28 (m, 2H, –NH2), 5.17–5.13 (m, 2H, –NH2), 2.25 (s, 6H, –CH3), 2.24 (s, 6H, –CH3), 2.15 (s, 6H, –CH3), 1.80 (s, 6H, –CH3), 1.74 (s, 6H, –CH3), 1.67 (s, 6H, –CH3) ppm. 13C NMR (125.61 MHz, 25 °C, CD2Cl2): δ 151.26, 149.80, 143.94, 141.61, 139.50, 137.63, 137.42, 136.01, 135.58, 132.50, 131.24, 130.65, 130.60, 128.00, 115.17, 26.77 (–CH3), 25.72 (–CH3), 23.61 (–CH3), 23.21 (–CH3), 21.07 (–CH3), 20.87 (–CH3) ppm. 11B (128.16 MHz, 25 °C, CD2Cl2): not observed. Elemental analysis calculated (%) for C58H62B2N2·0.93 × CHCl3: C, 74.24; H, 6.99; found C, 74.22; H, 7.03. This EA result indicates the crystal sample obtain from CHCl3/MeOH contains interstitial solvent molecules. This view is consistent with the crystallographic measurements which show the presence of disordered interstitial solvents equivalent to an electron count of 59 as indicated by application of the SQUEEZE protocol.13 This electron count corresponds almost exactly to the number of electrons of a chloroform molecule.58 These EA results suggest partial loss of the interstitial chloroform molecule.

Synthesis of [K(dibenzo-18-crown-6)]2-μ2-CN

A solution of KCN (10 mg, 0.15 mmol) in MeOH (2 mL) was slowly added at room temperature into a CH2Cl2 solution (10 mL) containing 1 (100 mg, 0.15 mmol) and dibenzo-18-crown-6 (56 mg, 0.15 mmol). The resulting colourless solution was stirred for an additional 1 h at room temperature. The solvent was removed under vacuum and the resulting solid was washed with Et2O (10 mL) twice to afford [K(dibenzo-18-crown-6)]2-μ2-CN as a white powder in 90% yield. 1H NMR (499.53 MHz, 25 °C, CDCl3): δ 7.00–6.97 (m, 4H, –CH2–), 6.84–6.79 (m, 4H, –CH2–), 6.66 (t, 2H, 3H-J = 8.78 Hz, biphenylene-CH), 6.50 (br, 8H, mes-CH), 6.20–6.16 (m, 2H, biphenylene-CH), 6.10 (d, 3H-J = 6.34 Hz, biphenylene-CH), 4.03–4.01 (m, 8H, –CH2–), 3.76–3.74 (m, 8H, –CH2–), 2.13 (br, 12H, –CH2–), 1.84 (br, 24H, –CH3) ppm. 13C NMR (125.62 MHz, 25 °C, CDCl3): δ 158.33, 157.36, 150.05, 146.14, 137.01, 136.28, 132.48, 132.37, 128.55, 124.36, 122.25, 112.02, 111.77, 80.29, 68.97 (–CH2–), 66.59 (–CH2–), 24.44 (–CH3), 20.82 (–CH3) ppm. 11B (128.16 MHz, 25 °C, CDCl3): δ –15.8 ppm. IR νCN = 2229 cm⁻¹. Elemental analysis calculated (%) for C60H52B2KO6·0.55 × CH2Cl2·C, 75.48; H, 6.84; found C, 75.49; H, 6.78. These EA results indicate partial loss of the interstitial solvent molecules.
(–CH3), 25.48 (–CH3), 24.79 (–CH3), 24.48 (–CH3), 21.24 (–CH3), 20.86 (–CH3) ppm. 11B (128.16 MHz, 25 °C, CDCl3): δ = 11.7 ppm. IR υ(C═O) 2184 cm−1. Elemental analysis calculated (%) for C77H80B2KNO6: C, 76.70; H, 6.71. These EA results indicate partial loss of the interstitial solvent molecules.

Computational details

Density functional theory (DFT) structural optimizations of 1 and 2 were carried with the Gaussian 09 program. In all cases, the structures were optimized using the B3LYP functional and the following mixed basis set: C/H, 6-31g; B, 6-31g+(d). For all optimized structures, frequency calculations were carried out to confirm the absence of imaginary frequencies. The molecular orbitals were visualized and plotted using the Jmip2 program. The TD-DFT and NTO calculations were carried out using the MPW1PW91 functional and the above-mentioned basis sets.

Conflicts of interest

The authors declare no conflict of interest.

Acknowledgements

Acknowledgement is made to the donors of the American Chemical Society Petroleum Research Fund for partial support of this research (Grant 56871-N03). We also acknowledge support from the Welch Foundation (Grant A-1423), Texas A&M University (Arthur E. Martell Chair of Chemistry), and the Laboratory for Molecular Simulation at Texas A&M University (software and computational resources) is gratefully acknowledged.

References


