



Cite this: *Chem. Sci.*, 2018, **9**, 6703

All publication charges for this article have been paid for by the Royal Society of Chemistry

Received 7th April 2018
Accepted 8th July 2018

DOI: 10.1039/c8sc01587b
rsc.li/chemical-science

Dewetting transitions coupled to K-channel activation in cytochrome c oxidase[†]

Shreyas Supekar and Ville R. I. Kaila  *

Cytochrome c oxidase (CcO) drives aerobic respiratory chains in all organisms by transducing the free energy from oxygen reduction into an electrochemical proton gradient across a biological membrane. CcO employs the so-called D- and K-channels for proton uptake, but the molecular mechanism for activation of the K-channel has remained elusive for decades. We show here by combining large-scale atomistic molecular simulations with graph-theoretical water network analysis, and hybrid quantum/classical (QM/MM) free energy calculations, that the K-channel is activated by formation of a reactive oxidized intermediate in the binuclear heme a_3/Cu_B active site. This state induces electrostatic, hydration, and conformational changes that lower the barrier for proton transfer along the K-channel by dewetting pathways that connect the D-channel with the active site. Our combined results reconcile previous experimental findings and indicate that water dynamics plays a decisive role in the proton pumping machinery in CcO.

Introduction

Cytochrome *c* oxidase (CcO) functions as the terminal electron acceptor in respiratory chains of mitochondria and aerobic bacteria.^{1,2} CcO catalyses the reduction of oxygen to water and it employs the released free energy to pump four protons across the mitochondrial inner membrane or the bacterial cytoplasmic membrane.^{1,2} This establishes a proton motive force that is utilized to thermodynamically drive ATP synthesis and active transport.³

CcO receives electrons from the soluble cytochrome *c*, and transfers them to oxygen, which is reduced to water in the binuclear heme a_3/Cu_B active site (binuclear centre, BNC; Fig. 1A). Reduction of the BNC is coupled with uptake of protons *via* the D- and K-channels, named after the conserved Asp-91 and Lys-319 residues (bovine numbering is used if not otherwise stated, Fig. 1A). All four pumped protons and two chemical protons used for the oxygen reduction chemistry in the oxidative half of the catalytic cycle (Fig. 1, $P_M \rightarrow O_H$) originate from the D-channel. In contrast, the K-channel is activated in the reductive half of the catalytic cycle ($O_H \rightarrow R$), and it is employed for the uptake of two chemical protons (Fig. 1). Despite extensive experimental^{4–8} and computational studies,^{9–14} the molecular mechanism for this channel-switching remains unclear.^{2,15,16} Interestingly, certain bacterial oxidases entirely lack the D-channel and utilize only their K-

channel analogue to transfer both chemical and pumped protons.^{17–19}

The catalytic cycle of CcO is initiated from the reduced R state ($a_3^{\text{II}}/\text{Cu}^{\text{I}}/\text{Tyr-OH}$) that binds and splits dioxygen to yield the P_M state ($a_3^{\text{IV}}=\text{O}^{2-}/\text{Cu}^{\text{II}}-\text{OH}^{\text{-}}/\text{Tyr-O}^{\text{-}}$) (Fig. 1B). Subsequent reduction of the BNC results in the P_R state ($a_3^{\text{IV}}=\text{O}^{2-}/\text{Cu}^{\text{II}}-\text{OH}^{\text{-}}/\text{Tyr-O}^{\text{-}}$),² which is coupled with transfer of a (pumped) proton from the conserved Glu-242 to a transient proton-loading site (PLS), located in the vicinity of heme a_3 .^{20–28} Protonation of the PLS leads to transfer of a second (chemical) proton from Glu-242 to the BNC, which ejects the PLS proton to the positively charged side (P-side) of the membrane and yields the F state ($a_3^{\text{IV}}=\text{O}^{2-}/\text{Cu}^{\text{II}}-\text{H}_2\text{O}/\text{Tyr-O}^{\text{-}}$).² Further reduction along with internal proton reorganization within the BNC yields the F_R state ($a_3^{\text{III}}-\text{OH}^{\text{-}}/\text{Cu}^{\text{II}}-\text{OH}^{\text{-}}/\text{Tyr-O}^{\text{-}}$),²⁹ which upon protonation decays to the oxidized O_H state, following a similar sequence as in the $P_M \rightarrow F$ transition.

The $O_H \rightarrow E_H$ transition is also believed to follow a similar sequence as the $P_M \rightarrow F$ and $F \rightarrow O_H$ transitions, but instead, the K-channel is employed for uptake of a chemical proton in the one-electron reduced O_H state (here the $O_{H,R}$, $a_3^{\text{III}}-\text{OH}^{\text{-}}/\text{Cu}^{\text{I}}-\text{H}_2\text{O}/\text{Tyr-O}^{\text{-}}$).^{15,20,30} This yields the E_H state,^{8,31} which by reduction and further protonation from the K-channel yields the R state, completing the catalytic cycle. The overall pumping cycle thus results in consumption of four electrons from the P-side, four protons from the N-side, and one oxygen molecule, producing two water molecules and pumping of four protons across the membrane from the N-side to the P-side.²

In addition to the pumping O_H state, CcO can also relax into a resting oxidized O state, but in contrast to the former, the reduction of the O state does not couple with proton pumping

Department Chemie, Technische Universität München, Lichtenbergstraße 4, D-85748 Garching, Germany. E-mail: ville.kaila@ch.tum.de

† Electronic supplementary information (ESI) available: Simulation details, analysis of hydration dynamics, pK_a calculations, DFT models, and QM/MM simulations. See DOI: 10.1039/c8sc01587b

across the membrane.³² The exact molecular structure of this low-potential O form is unclear and it is spectroscopically indistinguishable from the activated O_H state.³³ Based on computational studies,^{9,11,34} it has been suggested that the high and low potential forms of the O state could be connected to the electronic equilibrium between Cu_B (Cu^{II/1}) and Tyr-244 (Tyr-O[·]/Tyr-O[−]) that is modulated by the ligand state of Cu_B (H₂O/OH[−]). We show here that this unique Cu^{II}/Tyr-O[−] state is important for regulating the hydration state of the channels.

The K-channel starts at residue Glu-62 (subunit II) at the N-side of the membrane and leads *via* Lys-319 and Thr-316 to the cross-linked Tyr-244 of Cu_B, providing a proton conduit from the negatively charged side (N-side) to the BNC (Fig. 1A).^{15,35}

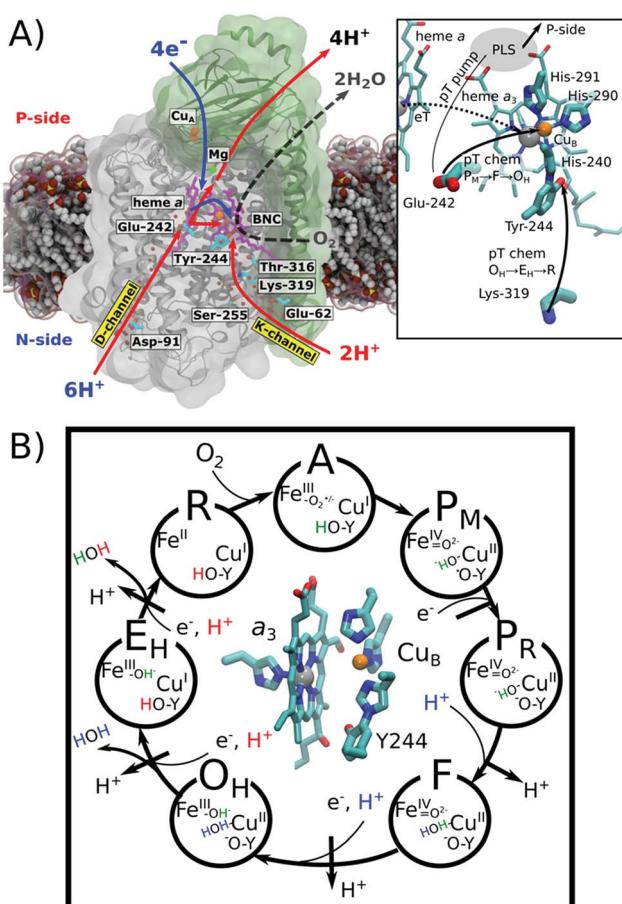


Fig. 1 Structure and function of cytochrome c oxidase (CcO). O₂ reduction to H₂O drives the electron transfers (blue arrow) to the binuclear centre (BNC: heme a₃/Cu_B) and proton transfers (red arrow) to the BNC for oxygen chemistry and for pumping of protons to the positively charged (P) side across the membrane. Protons are taken up from the negatively charged (N) side of the membrane through conserved D- and K-channels, terminating at Glu-242 and Tyr-244, respectively. Water molecules occupying the channels are shown as red spheres. Hemes are shown in purple and the copper centres are depicted in orange. Inset: close-up of the structure and proton transfer (pT)/electron transfer (eT) pathways between Glu-242, Lys-319, the proton loading site (PLS),^{20–27} heme a, and heme a₃/Cu_B. (B) Catalytic cycle of CcO showing the BNC and the catalytic intermediates. Protons (H⁺) marked in blue and red refer to chemical protons taken up from the D- and K-channels, respectively. Protons pumped across the membrane are indicated in black.

Mutation of Lys-319 drastically slows down the reduction of the fully oxidized state, whereas oxidation of the fully reduced state remains unaffected. This suggests that the K-channel provides chemical protons in the reductive half of the catalytic cycle.^{4,6} Interestingly, electrometric studies of that K319M mutant suggest that reduction of the BNC is coupled with an electrogenic charge movement within the K-channel³¹ that could arise from flipping Lys-319 towards the BNC, with the positive charge density facilitating the reduction of the BNC.

Molecular dynamics (MD) simulations of the K-channel have highlighted the role of hydrogen-bonded water molecules that are important for proton transfer reactions.^{10,14} Recent computational voltage studies³⁶ have reinforced the Lys-319 “up”-flip assignment from electrometric studies³¹ that probe voltage changes associated with charge movement along the membrane axis. Moreover, continuum electrostatics calculations based on the X-ray structures of the oxidized and reduced forms of CcO suggested that Lys-319 is in its protonated form, whereas pK_a calculations based on MD structures suggested that Lys-319 is deprotonated throughout the catalytic cycle, except in the one-electron reduced states of the BNC (e.g. P_R), where Lys-319 was predicted to be protonated.^{13,14} Quantum chemical calculations also suggest that Lys-319 might modulate the pK_a of Tyr-244.^{9,11,12,34} However, despite these important insights, the molecular mechanism behind the K-channel activation process still remains elusive.

In order to probe the K-channel activation process, we study here the P → F and O_H → E_H transitions (Fig. 1B) using multi-scale quantum and classical molecular simulations. To this end, we performed classical MD simulations of CcO embedded in a biological membrane environment on the microsecond timescales of each catalytic state, with the enzyme redox-states modelled based on quantum chemical calculations. Based on the MD simulations, we further characterized water networks connecting the D- and K-channels with the BNC using a graph-theoretical approach, and computed free energy profiles for the proton transfer process using DFT-based quantum mechanics/classical mechanics (QM/MM) free energy simulations.

Computational models

Classical MD simulations

The X-ray structure of subunits I-II bovine CcO [PDB ID: 1V54 (ref. 37)] was embedded in a 1-palmitoyl-2-oleoyl-sn-glycero-3-phosphocholine (POPC) membrane and was solvated with TIP3P water to yield a lipid-membrane-water-ion environment. The BNC was modelled in the P_M ($a_3^{IV}=\text{O}^{2-}/\text{Cu}^{II}-\text{OH}^-/\text{Tyr-O}^\cdot$), P_R ($a_3^{IV}=\text{O}^{2-}/\text{Cu}^{II}-\text{OH}^-/\text{Tyr-O}^\cdot$), O_H ($a_3^{III}-\text{OH}^-/\text{Cu}^{II}-\text{OH}_2/\text{Tyr-O}^\cdot$) and O_{H,R} ($a_3^{III}-\text{OH}^-/\text{Cu}^I-\text{OH}_2/\text{Tyr-O}^\cdot$) states using the CHARMM36 force field^{52a} and in-house DFT based parameters of the co-factors.^{52b} 500 ns MD simulations of each state were performed with NAMD 2.9–2.12³⁸ at $T = 310$ K using a 2 fs integration timestep, and by treating long-range electrostatics using the particle mesh Ewald approach. To probe the protonic connectivities from Glu-242 and Lys-319 to the BNC, we calculated the shortest water-mediated pathway using Dijkstra's algorithm,³⁹ and evaluated the longest (“rate-limiting”)



distance, ζ , along the shortest path (see the ESI and ESI-Fig. 13†). The employed protonation states are given in ESI-Table 2† and also described in the extended Methods section of the ESI.† Notably, Lys-319 was modelled in both protonated (Fig. 2) and deprotonated (neutral, ESI-Fig. 10†) states, and Glu-242 and Asp-364 in their protonated states (see the ESI†). Continuum electrostatic titration profiles were obtained by

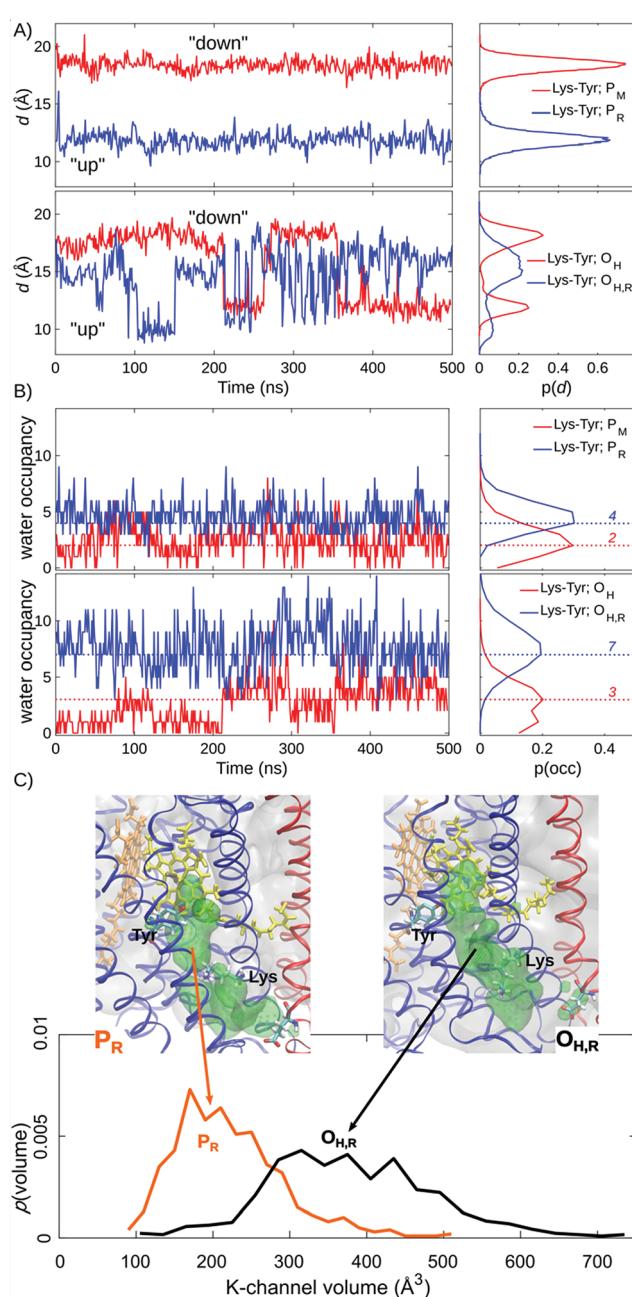


Fig. 2 Lys-319 sidechain dynamics, water occupancy in the K-channel above Lys-319 and K-channel volume (in \AA^3) from 500 ns of MD simulations. (A) Distances between Lys-319 and Tyr-244 from simulations of P_M , P_R , O_H , and $O_{H,R}$ states. (B) Water occupancies in the K-channel between Lys-319 and Tyr-244 from simulations of P_M , P_R , O_H , and $O_{H,R}$ states. Modal water occupancies are indicated as dotted lines. (C) Water accessible volume in the K-channel for the P_R and $O_{H,R}$ states.

solving the linearized Poisson–Boltzmann equation with MEAD.⁴⁰ To this end, the protein was described using partial atomic charges, embedded in an inhomogeneous dielectric continuum with dielectric constants of $\epsilon = 4$ for the protein, a dielectric membrane slab with $\epsilon = 2$, and $\epsilon = 80$ for the water. The probe radius was 1.4 Å and the ionic strength was 100 mM potassium chloride. In the PB calculations, all internal protein-bound water molecules were replaced by a polarizable medium with $\epsilon = 80$. Monte-Carlo pK_a evaluations based on the MD trajectories were performed with Karlsberg.⁴¹ The solvent accessible volume in the K-channel was calculated using Hollow.⁴²

DFT calculations

Quantum chemical clusters were optimized at the D3-BP86 level^{43–45} using the multipole accelerated resolution of identity approximation (RI-MARIJ).⁴⁶ The protein environment was modelled using a polarizable dielectric medium with $\epsilon = 4$ using the conductor-like screening model (COSMO).⁴⁷ The models included heme α_3 (without the propionic groups), Cu_B, and Tyr-244, His-290, His-291, His-240, His-376, Thr-316 and Lys-319, as well as 10–15 water molecules obtained from the MD simulations (see below). The amino acids were cut at the C β -positions, which were fixed during the structure optimization to account for protein strain. Single point energy calculations of the optimized structures were performed using the D3-B3LYP functional.^{43,48,49} In all models, def2-TZVP (Fe, Cu) and def2-SVP basis sets (C, H, O, N) were employed for geometry optimizations, while def2-TZVP basis sets were employed for single point energy calculations.⁵⁰ The DFT calculations were performed with TURBOMOLE v. 6.6–7.1.⁵¹

QM/MM free energies

In QM/MM calculations, link atoms were introduced between the C α and C β carbons of the amino acids in the QM region comprising *ca.* 120–130 atoms, while the remaining system was treated classically with the classical CHARMM36 force field.^{52a,b} For the QM/MM reaction pathway optimization, the QM region was described with the D3-B3LYP functional^{43,48,49} and def2-TZVP (Fe, Cu)/def2-SVP(C, H, O, N) basis sets.⁵⁰ QM/MM umbrella sampling (QM/MM US) simulations were performed with the D3-BP86 functional^{42–44} using the MARIJ approximation.⁴⁶ For QM/MM US calculations, each restraint window along the reaction coordinate was simulated for 1 ps at $T = 310$ K. No cutoffs were employed in the electrostatic embedding scheme. Due to the high computational cost of DFT-based QM/MM free energy simulations, the dynamics sampling is shorter than in previous semi-empirical-QM/MM or reactive force field/MM calculations of CcO.^{22–25} Although the starting structures for the QM/MM simulations were obtained from the 500 ns MD simulations, we nevertheless expect large statistical uncertainties in the PMF due to the limited QM/MM-sampling. Our DFT-based QM/MM dynamics is not based on pre-parametrized potentials which is why we expect an energetically well-balanced description of chemical intermediates involved in the pT reactions to the BNC, which is challenging to accurately describe at



the semi-empirical level. To sample the proton transfer path, the structures were subjected to harmonic restraints based on linear combination of all water O–H bond-distances between Glu-242 and Cu_B, and Lys-319 and Tyr-244 with a force constant of $k = 500$ kcal mol⁻¹ Å⁻² for QM/MM optimizations and $k = 100$ kcal mol⁻¹ Å⁻² for QM/MM MD US calculations. The CHARMM/TURBOMOLE⁵³ interface was employed for the QM/MM calculations (see the ESI† for further simulation details).

Results

Redox-state dependent lysine dynamics

Our MD simulations suggest that water molecules in the K-channel provide a hydrogen-bonded connectivity between Lys-319 and Tyr-244 via Thr-316, consistent with results from previous MD studies.^{10,12–14} In the P_M state, the Lys-319 side-chain flips “down”, towards the N-side of the membrane, and remains in this conformation for the complete 500 ns simulation (Fig. 2A). In contrast, in the one-electron reduced P_R state, Lys-319 flips “up” towards Tyr-244 (Fig. 2A). We observe that reduction of the BNC links to an average increase of two water molecules between Lys-319 and Tyr-244, suggesting that the charge state of the BNC might regulate the hydration of the proton channel (Fig. 2B). Similar electrostatic field effects have also been suggested to be involved in other redox-driven proton pumps, *e.g.*, in respiratory complex I,^{54–56} and light-driven ion-pumps.⁵⁷ The observed “up”-state of Lys-319 is also consistent with results from previous electrochemical experiments on the K319M mutant,³¹ suggesting that the “up” conformation can stabilize the reduced BNC, prior to its protonation, and also the oxidative phase where the K-channel is not employed for proton uptake.

In the O_H and O_{H,R} intermediates, where the K-channel is employed for proton uptake,^{4,6,15,16,32} we find that the lysine side-chain flickers between the “up” and “down” conformations, suggesting that the flipping barrier is low (Fig. 2B, ESI-Fig. 1†). In these simulations, we also observe a significant increase in the water occupancy of the K-channel from four water molecules in the P_R state to seven water molecules in the O_{H,R} state (Fig. 2B, ESI-Fig. 2†), whereas the water occupancies in the P_M and O_H states resemble each other (Fig. 2). We also observe conformational changes in the tertiary structure of the surrounding helices VIII, IX, and X (subunit I), and helix II (subunit II), which make the K-channel wider to accommodate the increased hydration state (Fig. 2C, ESI-Fig. 3 and 4†) in the O_{H/R} intermediates, but not in the P_{M/R} intermediates. These findings are consistent with conformational changes reported in previous X-ray structures of CcO.⁵⁸ Our pK_a calculations at the Poisson-Boltzmann (PB) continuum electrostatics level with Monte Carlo (MC) sampling of possible protonation states suggest that Lys-319 remains predominantly protonated in the P and O states (ESI-Fig. 5,† but *cf.* also ref. 14). Moreover, based on experimental data from the O_H → E_H transition, it has been suggested that the Lys-319 “up” conformation might be a prerequisite for reduction of the BNC, which itself is coupled to PLS protonation.⁵⁹ This further underpins the functional importance of Lys-319 in the proton pumping machinery of CcO.

Proton-pathway connectivities regulate the channel activity

The increased K-channel hydration, subsequent conformational changes, and flipping of Lys-319 suggest that the BNC might become more accessible from the K-channel upon formation of the O_{H,R} state. In order to probe the water-mediated protonic networks in the D- and K-channels, we employed a graph-theoretical approach for evaluating protonic connectives along the water-mediated pathways, which are important in catalysing Grotthuss-type pT reactions in proteins.⁶⁰ To this end, the longest (“rate-limiting”) distance along the shortest hydrogen-bonded pathway, ζ , connecting the proton donor (Lys-319 or Glu-242) and proton acceptor (Cu_B, heme a_3 , and Tyr-244) was evaluated over the MD trajectories (see the ESI†).

The ζ -reaction coordinate (see the ESI†) employed here should be considered as a qualitative measure of the hydration state and connectivity between the donor and acceptor paths. Previous studies suggest that proton pathways are not directly linked to the proton transfer barriers.^{22–25} However, if the proton donor and acceptor groups are distant as in CcO, and not connected by a water chain, as indicated by a high ζ value >4–5 Å, the proton transfer barriers are also significantly higher.² These observations are also consistent with recent reactive force field calculations by Liang *et al.*²⁵

We observe a persistent hydrogen-bonded connectivity along the K-channel between Lys-319 and Tyr-244 in the P_R, O_H, and O_{H,R} states with a high population of configurations with $\zeta < 2.5$ Å, but no connectivity between Lys-319 and Tyr-244 in the simulations of the P_M state, possibly due to the “down” conformation of Lys-319 (Fig. 3, ESI-Fig. 6†). In stark contrast, the connectivity from the terminal D-channel residue, Glu-242 to the BNC is persistent in the P states, whereas no hydrogen-bonded connectivity is observed between these sites in the O_H or O_{H,R} states (Fig. 3, ESI-Fig. 7†). The electric field originating from the P_R state strongly increases the hydrogen-bonded connectivity between Glu-242 and Cu_B, consistent with previous studies,^{2,61} but we also sample hydrogen-bonded connections between Glu-242 and Cu_B in the P_M state, suggesting that the orientation of the water chain alone is unlikely to serve as a gate for premature proton transfer reactions to the BNC. We also observe that the water occupancy of the non-polar cavity connecting Glu-242 with the BNC sharply drops in the O_{H,R} state, which could explain the lack of connectivity, whereas the water occupancy increases in the P_R state, favouring the formation of water wires (Fig. 3C, ESI-Fig. 7 and 8†). The free energy for non-polar cavity hydration near Glu-242 was recently probed by Son *et al.*⁶² for the *Rhodobacter sphaeroides* (Rs) CcO. In the P_R state, our simulations predict a somewhat higher cavity hydration in comparison to the Rs-CcO simulations that might arise from modelling CcO from a different organism. Moreover, here we have not considered the PLS protonation^{21–26} or Glu-242 deprotonation^{64–67} reactions in the classical simulations, which may also affect the cavity hydration state.⁶⁷ Although our study focuses on the connections between the D-/K-channels and the BNC, we note that hydrogen-bonded pathways to the PLS also remains open when 3–4 water molecules are present in the non-



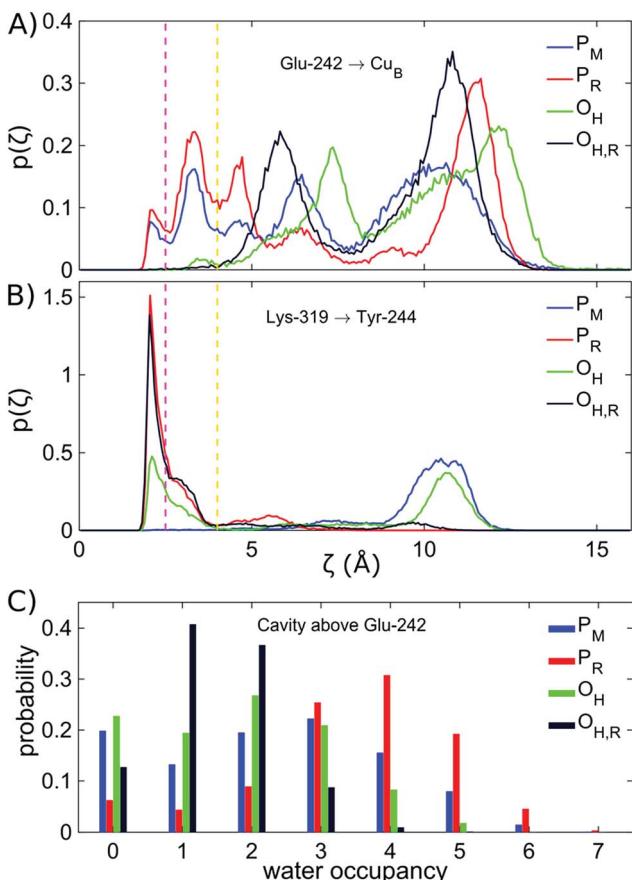


Fig. 3 Longest non-bonded connectivity along the shortest path connecting the proton donor and proton acceptor, ζ , calculated using Dijkstra's algorithm in graph theory. (A) Water-mediated connectivity from Glu-242 to Cu_B in P_M , P_R , O_H , and $\text{O}_\text{H,R}$ states. Thresholds for hydrogen-bonding connectivity are shown in magenta (2.5 Å) and yellow (4.0 Å) dashed lines. (B) Water-mediated connectivity from Lys-319 to Tyr-244 in P_M , P_R , O_H , and $\text{O}_\text{H,R}$ states. (C) Water occupancy in the non-polar cavity above Glu-242 in P_M , P_R , O_H , and $\text{O}_\text{H,R}$ states.

polar cavity (cf. ref. 21). The MD timescales employed here are longer than in previous work, but our cavity hydration profiles are nevertheless similar to previous simulations of the *Bos taurus* enzyme.^{2,21,61}

The dehydration of the non-polar cavity in the O_H states might arise from the weak aquo-ligand of Cu_B that is unable to fully stabilize the water pathway from Glu-242, an effect which is expected to be even more pronounced upon dissociation of the aquo-ligand. As discussed above, it should be noted that the exact structure of the O_H state is still unclear, but previous studies have suggested that heme α_3 is coordinated by a hydroxo ligand, and Cu_B has a trigonal planar geometry, weakly interacting with a water molecule.^{11,34,63} Moreover, FTIR experiments suggest that Tyr-244 is deprotonated in the O_H state.³⁰ Our calculations also suggest that the dominant pathway forms between Glu-242 and Cu_B , but not between Glu-242 and heme α_3 (ESI-Fig. 7†), and also upon formation of the F_R state, whose ζ profiles resemble those computed for the P_R state (ESI-Fig. 9†). These findings thus indicate that the dewetting of the non-polar cavity between the D-channel terminus (Glu-242) and the BNC

in the reductive phase could trigger the activation of the K-channel.

To address the dynamics of Lys-319 after pT to the BNC, we performed additional 500 ns MD simulations with Lys-319 in its neutral form, starting from the MD simulations of the P_R state with Lys-319 flipped towards the BNC. We find that in the neutral state, Lys-319 flips “down” towards the N-side of the membrane in the one-electron reduced states, P_R and $\text{O}_\text{H,R}$, whereas there is an increased population in the “up”-flipped population in the P_M and O_H states (ESI-Fig. 10A†). Interestingly, we also observe a strong reduction of the K-channel hydration in all four states with the neutral lysine, suggesting that the charged state of Lys-319 might modulate the hydration state of the K-channel (ESI-Fig. 10B†), similar to that in complex I^{54–56} and bacterial ion-pumps.⁵⁷ The sidechain dynamics of the neutral Lys-319 and the consequent loss of a hydrogen-bonded connectivity could help in preventing the chemical proton from leaking backwards towards the N-side. This could increase the overall efficiency of the pump, similar to our previously proposed glutamate-gate that might prevent leaks along the D-channel⁶⁴ (cf. also ref. 65–67). The observed hydration/dehydration effects in the K-channel might be further linked to the two experimentally known forms of the oxidized state: the “resting”, non-pumping O state, and the physiologically relevant, pumping O_H state. As discussed above, these states are spectroscopically indistinguishable,³³ but it is known that the resting O state can be activated by reduction and re-oxidation to yield the activated O_H state. This leads to synthesis of new water molecules in the BNC, which could in turn lead to hydration of the D- and K-channels, and the transient proton-loading site to sustain the proton pumping function.²¹ Moreover, wetting/dewetting transitions in the non-polar cavity above Glu-242 have been shown to play a key role in the proton pumping machinery of CcO in the $\text{P} \rightarrow \text{F}$ transition.^{62,67}

Energetics of proton transfer from the K-channel

In order to probe the energetics for the pT process, we constructed quantum chemical models with *ca.* 160–200 atoms, comprising the BNC and its nearby surroundings based on the MD structures. The models were optimized at the density functional theory (DFT) level. The calculations suggest that transferring the proton from Lys-319 to Tyr-244 in the $\text{O}_\text{H,R}$ state is stabilized by *ca.* 3 kcal mol⁻¹ as compared to the P_R state (ESI-Fig. 11†), which might be linked to the unique electronic equilibrium between $\text{Cu}[\text{i}]/\text{Tyr-O}^\bullet$ and $\text{Cu}[\text{ii}]/\text{Tyr-O}^-$ in the $\text{O}_\text{H} \rightarrow \text{O}_\text{H,R}$ transition.^{11,29,34} The DFT calculations also suggest that the intrinsic proton affinity of the Cu_B hydroxo ligand in the P_R state is up to *ca.* 6 pK-units higher in comparison to the proton affinities of Tyr-244 and *ca.* 14 pK-units higher compared to the oxo ligand of heme α_3 . In the $\text{O}_\text{H,R}$ state, Tyr-244 has a *ca.* 4 pK-units higher proton affinity than the oxygenous ligand of heme α_3 (ESI-Table 1†), suggesting that Tyr-244 is the likely proton acceptor. Nonetheless, the absence of the protonic connectivity between both Glu-242 and Tyr-244 or heme α_3 in the $\text{O}_\text{H,R}$ state further supports the view that the proton is taken up from the K-channel instead (Fig. 3, ESI-Fig. 7†). Previous DFT calculations²⁹



suggest that the BNC could take up an extra proton in the absence of continuous electron flow. This significantly lowers the redox potential, which in turn does not support the proton pumping activity. Our models of the O_H state are consistent with the ligand state of these calculations, and further suggest that the electric field arising from a ligand substitution could drastically affect the hydration dynamics and proton transfer energetics.

Free energy profiles for proton transfer along the D- and K-channels

To probe free energies for the pT reaction from Lys-319 to Tyr-244 and Glu-242 to Cu_B in the presence of dynamic protein surroundings, we performed QM(DFT)/MM umbrella sampling (US) simulations. To this end, we computed the free energy surface for the pT process along water chains connecting the proton donor and acceptor that were formed during the classical MD simulations ($\zeta < 2.5 \text{ \AA}$). We obtain free energy barriers of *ca.* 13 kcal mol⁻¹ and 11 kcal mol⁻¹ for the P_R and $O_{H,R}$ states, respectively. The profiles indicate that the pT is unlikely from the K-channel in the P_R state as the product state is not stabilized, whereas pT is feasible from the D-channel in the P_R state, and from the K-channel in the $O_{H,R}$ state (Fig. 4). This suggests that the K-channel is both thermodynamically and kinetically preferred for pT in the $O_{H,R}$ state, while the D-channel is kinetically and thermodynamically likely to supply chemical protons in the P_R state. The free energy barriers of 13 and 11 kcal mol⁻¹ for pT in the P_R and $O_{H,R}$ states are consistent with the pT rates inferred from experimental electrometric amplitudes with rate-constants of 0.8 ms (*ca.* 14 kcal mol⁻¹),²⁰ and by considering that reaction barriers are often somewhat underestimated at the DFT/GGA level. However, to further explain why the K-channel is not employed in the P_R state, we also studied the energetics for the pT between Lys-319 and Tyr-244 and compared it to the pT profiles between Glu-242 and Cu_B based on MD structures with good hydrogen-bonded connectivities ($\zeta < 2.5 \text{ \AA}$). A similar trend is also reflected in our QM/MM reaction pathway optimizations suggesting that the pT energy barrier along the K-channel (Lys-319 \rightarrow Tyr-244) is *ca.* 5–8 kcal mol⁻¹ higher than that for a pT from the D-channel (Glu-242 \rightarrow Cu_B) in the P_R state (ESI-Fig. 13†). Comparison of the free energy and energy profiles suggests that there could be an entropic contribution of up to 8 kcal mol⁻¹ in the process. Although similar entropic estimates have also been found previously,^{68,69} it is nevertheless likely that these estimates arise, at least in part, from non-equilibrium protein relaxation effects, which are not accurately captured in picosecond timescale QM/MM calculations. Our combined QM/MM results thus suggest that the D-channel supplies the chemical protons in the $P \rightarrow F$ transition due to low kinetic barriers and stable product states, whereas lack of connectivity between the D-channel and the BNC in the $O \rightarrow E$ transition might trigger the activation of the K-channel for delivery of the chemical protons.

Our work predicts that the wetting/dewetting transition is mediated by the redox and ligand states of the BNC, which in turn regulate the proton transfer kinetics along the D- and K-channels. Our putative model could be experimentally

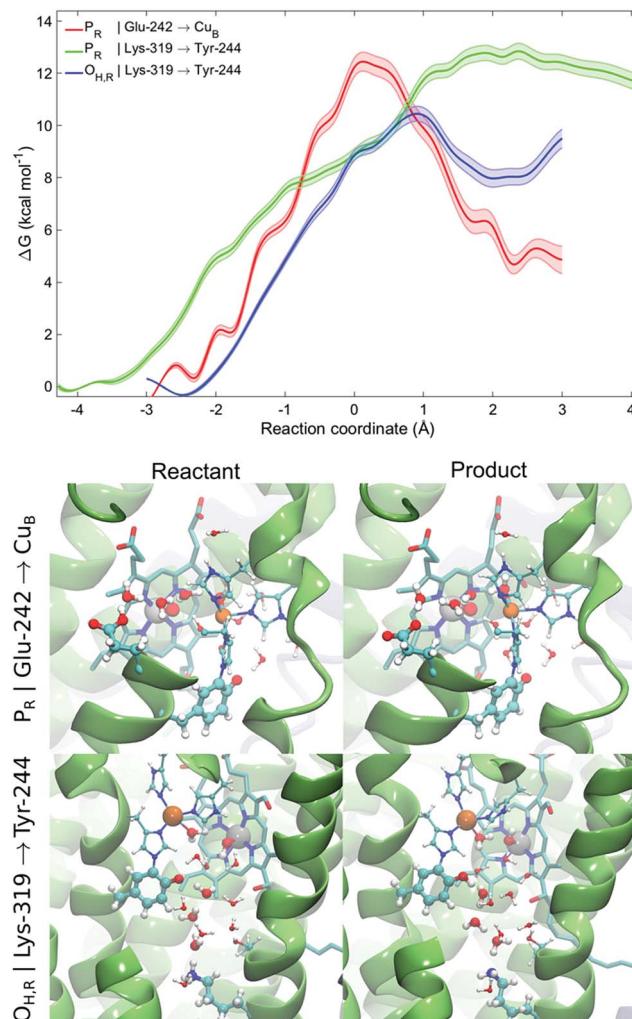


Fig. 4 QM/MM-MD free energy profiles for proton transfer from Glu-242 to Cu_B in the P_R state and from Lys-319 to Tyr-244 in the P_R and $O_{H,R}$ states. Lighter traces indicate the statistical error for the free energy profile. Snapshots of the reactants and products for the proton transfer profiles from Glu-242 to Cu_B in the P_R state and from Lys-319 to Tyr-244 in the $O_{H,R}$ state shown in the bottom panel (see ESI-Fig. 12 for more snapshots†).

validated by kinetically accelerating proton uptake along the K-channel that is expected to rescue the catalytic activity of CcO with a blocked D-channel. This could be achieved by stabilizing an open conformation of the K-channel by removing bulky residues at the hydrophobic gate area close to Lys-319, and/or by levelling the pK_a difference between the Lys-319 and Tyr-244, *e.g.*, by replacing the former by a tyrosine residue. In addition to the crystallographic studies that support the subtle helix motions predicted to couple with the K-channel opening, this effect could also be probed, for example, in fluorescence resonance energy transfer (FRET)-experiments by labelling the helices surrounding the K-channel.

Conclusions

We have studied here the molecular mechanism for the K-channel activation in CcO by combined MD simulations,



graph-theoretical analysis of hydrogen-bonded networks, and QM/MM free energy simulations. Our combined results suggest that formation of a reactive Cu¹⁺/Tyr-O⁻ state upon reduction of the O_H intermediate increases the proton affinity of Tyr-244, and leads to an increase in the hydration levels of the K-channel that provides an effective hydrogen-bonded connectivity to the BNC. These hydration changes are supported by tertiary structural changes around the K-channel. The increased proton affinities and the elevated hydration state subsequently lower the pT barriers along the K-channel. Importantly, our simulations also suggest that the D-channel cannot provide a proton to the BNC in the reductive phase due to a loss of water-mediated connectivity to the BNC from the D-channel. Consistent with electrometric studies that probe voltage changes associated with charge movement along the membrane axis,³¹ we also find that Lys-319 flips towards the BNC upon its reduction in the P_M → P_R transition that could stabilize electron transfer from heme *a*. The combined results thus suggest that Glu-242, the terminal residue of the D-channel is wired to the oxygenous ligand of Cu_B, whereas the Lys-319 of the K-channel is wired to Tyr-244 and further to heme *a*₃. Upon formation of the aquo-ligand on Cu_B in the O_H state, the connectivity from the D-channel weakens, and thus, the K-channel is kinetically selected for transfer of the protons to the BNC. In light of our present results, it is thus important to elucidate in future work how bacterial heme-copper oxidases employ their K-channel analogues for uptake of both chemical and pumped protons, and how their pumping stoichiometries compare to the corresponding canonical A-type oxidases.^{4,19}

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We thank Prof. Mårten Wikström for insightful discussions. This work was supported by the German Research Foundation (DFG). We acknowledge SuperMuc (grant: pr84gu) at the Leibniz Rechenzentrum for computing time.

References

- 1 M. K. F. Wikström, *Nature*, 1977, **266**, 271–273.
- 2 V. R. I. Kaila, M. I. Verkhovsky and M. Wikström, *Chem. Rev.*, 2010, **110**, 7062–7081.
- 3 P. Mitchell, *Nature*, 1961, **191**(4784), 144–148.
- 4 P. Ädelroth, R. B. Gennis and P. Brzezinski, *Biochemistry*, 1998, **37**, 2470–2476.
- 5 C. Pecoraro, R. B. Gennis, T. V. Vygodina and A. A. Konstantinov, *Biochemistry*, 2001, **40**, 9695–9708.
- 6 D. Riegler, *et al.*, *Biochim. Biophys. Acta, Bioenerg.*, 2005, **1706**, 126–133.
- 7 T. V. Vygodina, C. Pecoraro, D. Mitchell, R. B. Gennis and A. A. Konstantinov, *Biochemistry*, 1998, **37**, 3053–3061.
- 8 M. Ruitenberg, *et al.*, *Proc. Natl. Acad. Sci. U. S. A.*, 2000, **97**, 4632–4636.
- 9 M. R. A. Blomberg, *Biochemistry*, 2016, **55**, 489–500.
- 10 R. I. Cukier, *Biochim. Biophys. Acta, Bioenerg.*, 2005, **1706**, 134–146.
- 11 V. Sharma, K. D. Karlin and M. Wikström, *Proc. Natl. Acad. Sci. U. S. A.*, 2013, **110**, 16844–16849.
- 12 V. Sharma and M. Wikström, *Biochim. Biophys. Acta, Bioenerg.*, 2016, **1857**, 1111–1115.
- 13 A. Tuukkanen, M. I. Verkhovsky, L. Laakkonen and M. Wikström, *Biochim. Biophys. Acta, Bioenerg.*, 2006, **1757**, 1117–1121.
- 14 A. L. Woelke, G. Galstyan and E.-W. Knapp, *Biochim. Biophys. Acta, Bioenerg.*, 2014, **1837**, 1998–2003.
- 15 A. A. Konstantinov, S. Siletsky, D. Mitchell, A. Kaulen and R. B. Gennis, *Proc. Natl. Acad. Sci. U. S. A.*, 1997, **94**, 9085–9090.
- 16 P. R. Rich and A. Marechal, *J. R. Soc., Interface*, 2013, **10**, 20130183.
- 17 H. Y. Chang, J. Hemp, Y. Chen, J. A. Fee and R. B. Gennis, *Proc. Natl. Acad. Sci. U. S. A.*, 2009, **106**, 16169–16173.
- 18 H. Han, *et al.*, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **108**, 14109–14114.
- 19 V. Rauhamäki, M. Baumann, R. Soliymani, A. Puustinen and M. Wikström, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 16135–16140.
- 20 I. Belevich, D. A. Bloch, N. Belevich, M. Wikström and M. I. Verkhovsky, *Proc. Natl. Acad. Sci. U. S. A.*, 2007, **104**, 2685–2690.
- 21 S. Supekar, A. P. Gamiz-Hernandez and V. R. I. Kaila, *Angew. Chem., Int. Ed.*, 2016, **55**, 11940–11944.
- 22 P. Goyal, S. Yang and Q. Cui, *Chem. Sci.*, 2015, **6**, 826–841.
- 23 A. V. Pisliakov, P. K. Sharma, Z. T. Chu, M. Haranczyk and A. Warshel, *Proc. Natl. Acad. Sci. U. S. A.*, 2008, **105**, 7726–7731.
- 24 R. Liang, J. M. J. Swanson, M. Wikström and G. A. Voth, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, **114**, 5924–5929.
- 25 R. Liang, J. M. J. Swanson, Y. Peng, M. Wikström and G. A. Voth, *Proc. Natl. Acad. Sci. U. S. A.*, 2016, **113**, 7420–7425.
- 26 J. Lu and M. R. Gunner, *Proc. Natl. Acad. Sci. U. S. A.*, 2014, **111**, 12414–12419.
- 27 M. Svensson-Ek, J. Abramson, G. Larsson, S. Törnroth, P. Brzezinski and S. Iwata, *J. Mol. Biol.*, 2002, **321**, 329–339.
- 28 M. Iwaki and P. R. Rich, *J. Am. Chem. Soc.*, 2007, **129**, 2923–2929.
- 29 M. R. A. Blomberg and P. E. M. Siegbahn, *Biochim. Biophys. Acta, Bioenerg.*, 2015, **1847**, 364–376.
- 30 E. A. Gorbikova, M. Wikström and M. I. Verkhovsky, *J. Biol. Chem.*, 2008, **283**, 34907–34912.
- 31 H. Lepp, E. Svahn, K. Faxén and P. Brzezinski, *Biochemistry*, 2008, **47**, 4929–4935.
- 32 D. Bloch, I. Belevich, A. Jasaitis, C. Ribacka, A. Puustinen, M. I. Verkhovsky and M. Wikström, *Proc. Natl. Acad. Sci. U. S. A.*, 2004, **101**, 529–533.
- 33 D. Jancura, V. Berka, M. Antalik, J. Bagelova, R. B. Gennis, G. Palmer and M. Fabian, *J. Biol. Chem.*, 2006, **281**, 30319–30325.



34 V. R. I. Kaila, M. P. Johansson, D. Sundholm, L. Laakkonen and M. Wikström, *Biochim. Biophys. Acta, Bioenerg.*, 2009, **1787**, 221–233.

35 S. Iwata, C. Ostermeier, B. Ludwig and H. Michel, *Nature*, 1995, **376**, 660–669.

36 I. Kim and A. Warshel, *Proc. Natl. Acad. Sci. U. S. A.*, 2016, **113**, 7810–7815.

37 T. Tsukihara, *et al.*, *Proc. Natl. Acad. Sci. U. S. A.*, 2003, **100**(26), 15304–15309.

38 J. C. Phillips, *et al.*, *J. Comput. Chem.*, 2005, **26**, 1781–1802.

39 E. W. Dijkstra, *Numer. Math.*, 1959, **1**, 269–271.

40 D. Bashford and K. Gerwert, *J. Mol. Biol.*, 1992, **224**, 473–486.

41 (a) G. Kieseritzky and E. W. Knapp, *Proteins*, 2008, **71**, 1335–1348; (b) B. Rabenstein and E. W. Knapp, *Biophys. J.*, 2001, **80**, 1141–1150.

42 B. K. Ho and F. Gruswitz, *BMC Struct. Biol.*, 2008, **8**, 49.

43 S. Grimme, *J. Comput. Chem.*, 2006, **27**, 1787–1799.

44 A. D. Becke, *Phys. Rev. A: At., Mol., Opt. Phys.*, 1988, **38**, 3098–3100.

45 J. P. Perdew, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1986, **33**, 8822–8824.

46 M. Sierka, A. Hogekamp and R. Ahlrichs, *J. Chem. Phys.*, 2003, **118**, 9136–9148.

47 A. Klamt and G. Schüürmann, *J. Chem. Soc., Perkin Trans. 2*, 1993, **2**, 799–805.

48 C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1988, **37**, 785.

49 A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648–5652.

50 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.

51 R. Ahlrichs, M. Bär, M. Häser, H. Horn and C. Kölmel, *Chem. Phys. Lett.*, 1989, **162**, 165–169.

52 (a) A. D. MacKerell, *et al.*, *J. Phys. Chem. B*, 1998, **102**, 3586–3616; (b) M. P. Johansson, V. R. I. Kaila and L. Laakkonen, *J. Comput. Chem.*, 2008, **29**, 753–767.

53 S. Riahi and C. N. Rowley, *J. Comput. Chem.*, 2014, **35**, 2076–2086.

54 A. Di Luca, A. P. Gamiz-Hernandez and V. R. I. Kaila, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, **114**, E6314–E6321.

55 V. R. I. Kaila, M. Wikström and G. Hummer, *Proc. Natl. Acad. Sci. U. S. A.*, 2014, **111**, 6988–6993.

56 V. R. I. Kaila, *J. R. Soc., Interface*, 2018, **15**, 20170916.

57 C. M. Suomivuori, A. P. Gamiz-Hernandez, D. Sundholm and V. R. I. Kaila, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, **114**, 7043–7048.

58 L. Qin, *et al.*, *Biochemistry*, 2009, **48**, 5121–5130.

59 M. Wikström, V. Sharma, V. R. I. Kaila, J. P. Hosler and G. Hummer, *Chem. Rev.*, 2015, **115**, 2196–2221.

60 F. Garczarek and K. Gerwert, *Nature*, 2005, **439**, 109–112.

61 V. Sharma, G. Enkavi, I. Vattulainen, T. Rög and M. Wikström, *Proc. Natl. Acad. Sci. U. S. A.*, 2015, **112**, 2040–2045.

62 C. Y. Son, A. Yethiraj and Q. Cui, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, **114**, E8830–E8836.

63 M. R. A. Blomberg and P. E. M. Siegbahn, *Biochim. Biophys. Acta, Bioenerg.*, 2015, **1847**, 1173–1180.

64 V. R. I. Kaila, M. I. Verkhovsky, G. Hummer and M. Wikström, *Proc. Natl. Acad. Sci. U. S. A.*, 2008, **105**, 6255–6259.

65 A. L. Woelke, G. Galstyan, A. Galstyan, T. Meyer, J. Heberle and E. W. Knapp, *J. Phys. Chem. B*, 2013, **117**, 12432–12441.

66 B. M. Samudio, V. Couch and A. A. Stuchebrukhov, *J. Phys. Chem. B*, 2016, **120**, 2095–2105.

67 P. Goyal, J. Lu, S. Yang, M. R. Gunner and Q. Cui, *Proc. Natl. Acad. Sci. U. S. A.*, 2013, **110**, 18886–18891.

68 A.-L. Johansson, S. Chakrabarty, C. L. Berthold, M. Högbom, A. Warshel and P. Brzezinski, *Biochim. Biophys. Acta, Bioenerg.*, 2011, **1807**, 1083–1094.

69 S. Narayan, D. L. Wyatt, D. S. Crumrine and S. Cukierman, *Biophys. J.*, 2007, **93**, 1571–1579.

