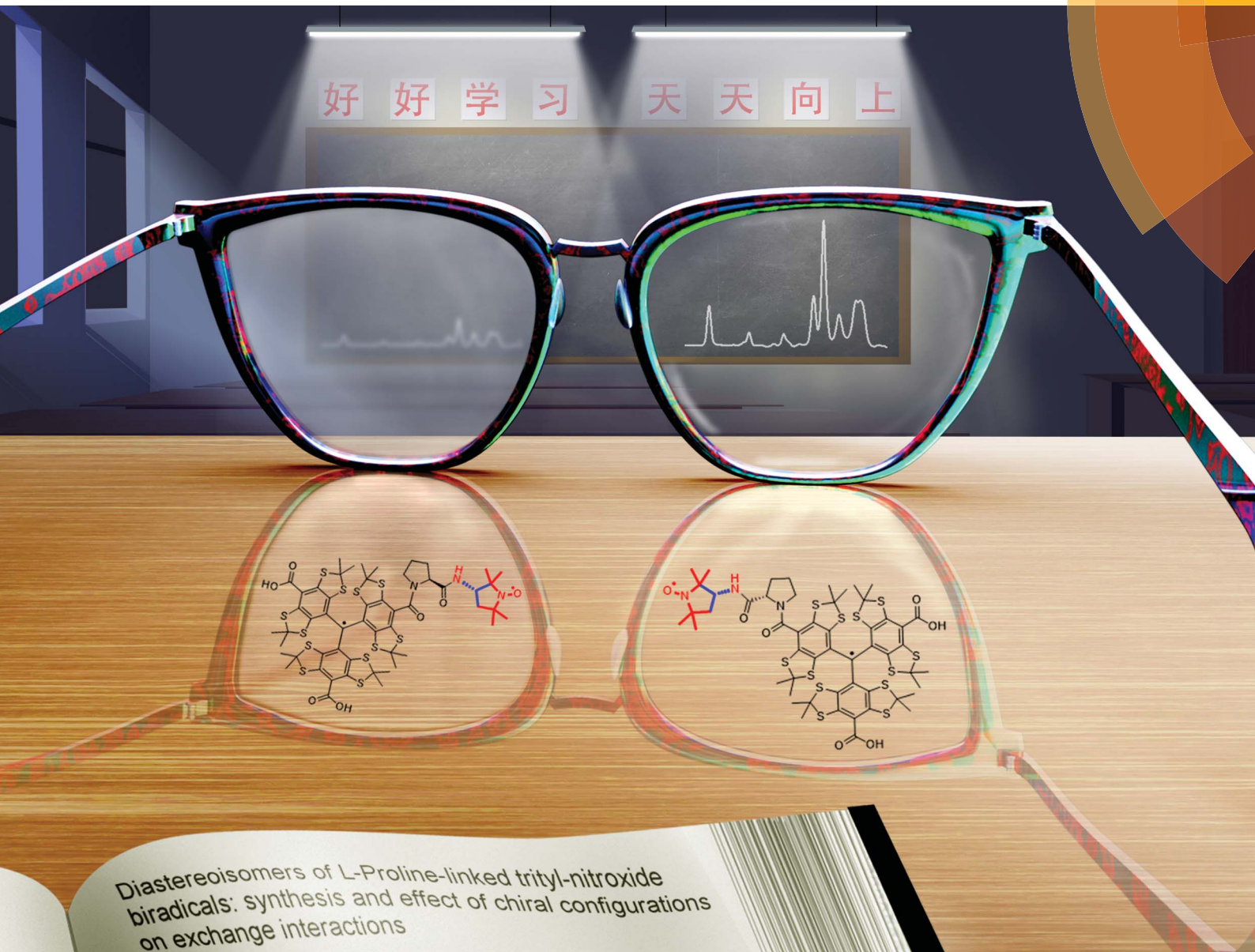


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Diastereoisomers of L-proline-linked trityl-nitroxide biradicals: synthesis and effect of chiral configurations on exchange interactions†

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The exchange (J) interaction of organic biradicals is a crucial factor controlling their physicochemical properties and potential applications and can be modulated by changing the nature of the linker. In the present work, we for the first time demonstrate the effect of chiral configurations of radical parts on the J values of trityl-nitroxide (TN) biradicals. Four diastereoisomers (TNT₁, TNT₂, TNL₁ and TNL₂) of TN biradicals were synthesized and purified by the conjugation of a racemic (R/S) nitroxide with the racemic (M/P) trityl radical *via* L-proline. The absolute configurations of these diastereoisomers were assigned by comparing experimental and calculated electronic circular dichroism (ECD) spectra as (M, S, S) for TNT₁, (P, S, S) for TNT₂, (M, S, R) for TNL₁ and (P, S, R) for TNL₂. Electron paramagnetic resonance (EPR) results showed that the configuration of the nitroxide part instead of the trityl part is dominant in controlling the exchange interactions and the order of the J values at room temperature is TNT₁ (252 G) > TNT₂ (127 G) >> TNL₂ (33 G) > TNL₁ (14 G). Moreover, the J values of TNL₁/TNL₂ with the S configuration in the nitroxide part vary with temperature and the polarity of solvents due to their flexible linker, whereas the J values of TNT₁/TNT₂ are almost insensitive to these two factors due to the rigidity of their linkers. The distinct exchange interactions between TNT_{1,2} and TNL_{1,2} in the frozen state led to strongly different high-field dynamic nuclear polarization (DNP) enhancements with $\epsilon = 7$ for TNT_{1,2} and 40 for TNL_{1,2} under 800 MHz DNP conditions.

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Introduction

Exchange-coupled biradicals have recently attracted tremendous interest owing to their unique physicochemical properties and potential applications in high-frequency dynamic nuclear polarization (DNP),^{1,2} molecule-based magnetism,^{3–5} and molecular charge transfer^{6–8} as well as molecular sensing.^{9,10} A key point in the design of new biradicals is to control the magnitude and sign of the spin–spin exchange interaction that determines their physicochemical properties and potential applications. The intramolecular exchange interaction of biradicals strongly depends on the number of chemical bonds in the linkage between the two spins, the linker conformation, and

the σ/π contributions to the bonding as well as the environment (*e.g.*, temperature and solvent). This exchange interaction can be through-bond and/or through-space, and its value varies by many orders of magnitude due to different linkers between the two radical moieties.^{11,12} In general, for the biradicals with π -conjugated backbones, the exchange interaction is mediated more effectively *via* through-bond mechanisms. The large exchange interactions have been achieved by enhancing the degree of π -orbital overlap between the spacer and two radical moieties through a conformational constraint to enforce their coplanarity.^{13–16} Using donor–bridge–acceptor biradical systems, the dependence of magnetic exchange interaction on the degree of π -orbital overlap and torsional rotations between different parts has been well demonstrated.^{17–19} Comparatively, the exchange interactions in the biradicals with nonconjugated spacers are much weaker and can be through-bond and/or through-space, depending on the flexibility of the spacers. Recently, weakly coupled nitroxide biradicals have attracted intense attention as high-field DNP polarizing agents which boost the sensitivity of solid-state nuclear magnetic resonance spectroscopy.^{20–27} Nonconjugated linkers were applied in these biradicals and the rigidity and conformation of the linkers were well tailored in order to optimize the through-bond exchange and dipolar interactions and simultaneously realize the

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matching conditions of the EPR frequency. In addition, the supramolecular interaction with host molecules has been an effective approach to modulate the exchange interaction in nitroxide biradicals.^{28–31} Despite these extensive studies, to the best of our knowledge, there is no related study on the chiral effect of the linker and radical moieties on the exchange interaction of biradicals.

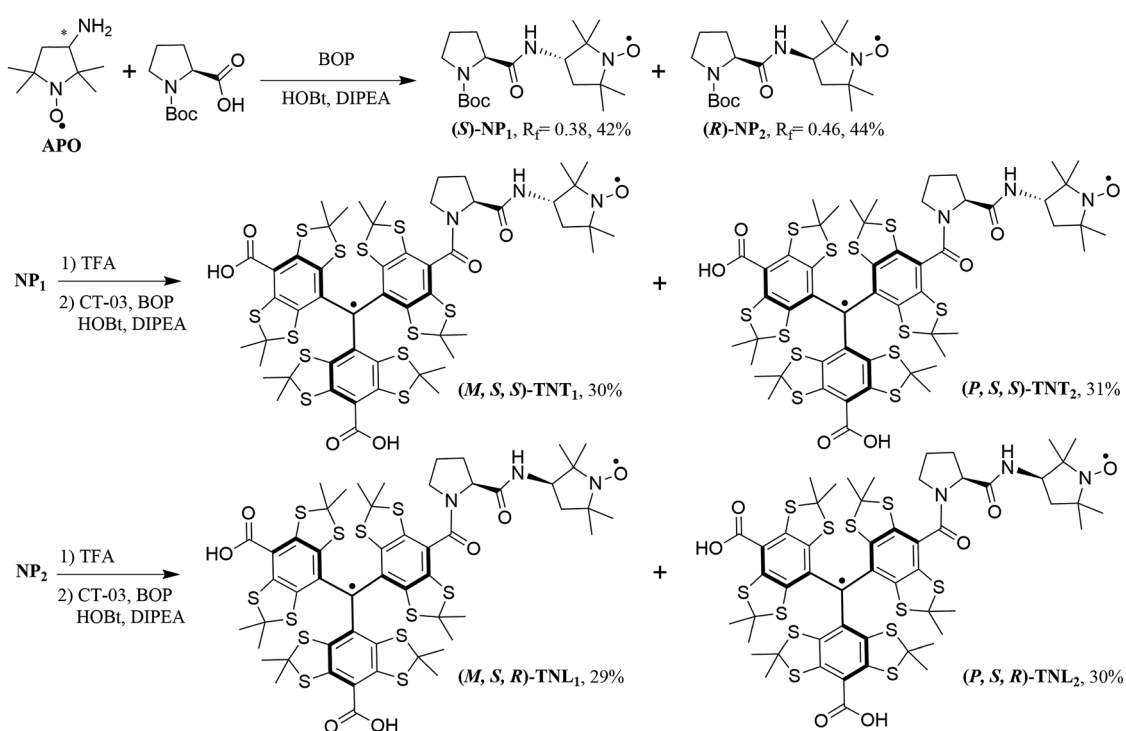
In the past few years, we have developed trityl-nitroxide (TN) biradicals which combine two different radical properties into unique molecules.^{32–36} Similar to other biradicals, the exchange interaction of TN biradicals is a key factor for their applications. We have fine-tuned the exchange interactions of TN biradicals by structural modification of the linkers³⁷ as well as by supramolecular interaction with cyclodextrins.³⁸ The largest DNP enhancement at a high magnetic field (18.8 T) has been achieved by using these TN biradicals as polarizing agents,^{35,36} in part due to the lack of nuclear depolarization.³⁹ In the present work, we utilize the chiral effect of the linker and two radical parts to modulate the spin–spin exchange interaction of TN biradicals. The trityl radical has a propeller configuration which affords the right-handed (*P*) and left-handed (*M*) helices. Due to the large steric bulk of three aryl groups, the interconversion between the two enantiomeric helices is slow and they are separable at room temperature.^{40,41} Thus, the conjugation of the trityl radical CT-03 with the racemic mixture of the nitroxide APO through *L*-proline leads to four enantiomerically pure diastereoisomers (TNT₁, TNT₂, TNL₁ and TNL₂, Scheme 1) which were separated by reversed-phase (TNT₁ and TNT₂) or chiral (TNL₁ and TNL₂) HPLC. The configurations of these diastereoisomers were determined by the comparison of experimental

and calculated electronic circular dichroism (ECD) spectra. The effect of the configurations of the trityl and nitroxide radicals on the exchange interactions was investigated by EPR at different temperatures and solvents. Moreover, the DNP properties of TNT_{1,2} and TNL_{1,2} were also studied at a high field (18.8 T).

Results and discussion

Synthesis of the four enantiomerically pure diastereoisomers

Scheme 1 shows the synthetic procedure of the four biradical diastereoisomers. The racemic mixture of the nitroxide APO was firstly coupled with *N*-Boc-*L*-proline to afford the nitroxide diastereoisomers NP₁ (yield, 42%) and NP₂ (yield, 44%) which were easily separated using column chromatography on silica gel. Then NP₁ and NP₂ underwent deprotection by TFA, followed by conjugation with the trityl radical CT-03 to give the biradicals TNT_{1,2} and TNL_{1,2}, respectively. TNT_{1,2} displayed two close peaks (11.34 and 11.58 min) with an area ratio of 1 : 1 on reversed-phase high performance liquid chromatography (RP-HPLC, Fig. S1†) due to the *P/M* propeller configurations of the trityl part.^{40,41} Therefore, it is possible to separate the two enantiomerically pure diastereoisomers of TNT_{1,2} by semi-preparative RP-HPLC and ~2 mg of both diastereoisomers (TNT₁ and TNT₂) were obtained. In contrast, only one peak was observed in the RP-HPLC chromatogram of TNL_{1,2} under the same conditions (Fig. S1†). Attempts to separate the two diastereoisomers of TNL_{1,2} by optimizing the RP-HPLC conditions failed. In addition, attempts to obtain the pure diastereoisomers of TNL_{1,2} through their esterification using chiral agents such as *S*(–)-1,1'-binaphthyl-2,2'-diol were unsuccessful. As



Scheme 1 Synthesis of the four enantiomerically pure diastereoisomers of TN biradicals. *M/P* represents the helices of the trityl radical, while *R/S* indicates the configurations of *L*-proline and the nitroxide APO.



such, the separation of TNL_1 and TNL_2 was achieved by analytical HPLC with a chiral stationary phase (Fig. S3†) and this separation procedure was repeated ~ 50 times to afford ~ 0.4 mg of TNL_1 and ~ 0.4 mg of TNL_2 . The amounts obtained for the four diastereoisomers are sufficient for subsequent ECD and EPR experiments.

Configuration of the four biradical diastereoisomers

In order to determine the absolute configurations of the four biradical diastereoisomers, the experimental electronic circular dichroism (ECD) spectra of NP_1/NP_2 and the four diastereoisomers were recorded in methanol. Meanwhile, the theoretical ECD spectra of NP_1 , TNT_2 and TNL_2 were also calculated using quantum chemical methods. All calculated ECD spectra with the TD-DFT method show an acceptable fitting with the corresponding experimental spectra in terms of the spectral pattern and sign of the bands with respect to the wavelength. As illustrated in Fig. 1A, NP_1 exhibits a relatively strong ECD signal with a negative Cotton effect at 217 nm and a broad signal with a positive Cotton effect at 420 nm. Good agreement between the experimental and calculated ECD spectral patterns of NP_1 confirms the *S* absolute configuration of the nitroxide part. Accordingly, the absolute configuration of the nitroxide part of NP_2 is assigned as *R*.

Based on the absolute configuration of the nitroxide part, the absolute configurations of the four biradical diastereoisomers were further determined. The ECD spectrum of TNT_1 is characteristic of the trityl radical with two strong positive Cotton effects at 384 nm and 467 nm and two relatively weak negative Cotton effects at 324 and 352 nm.⁴² Interestingly, the ECD spectra of TNT_1 and TNT_2 are mirror images most likely because the ECD signals of the nitroxide part in these biradical diastereoisomers are very weak and negligible as compared to the signals of the trityl part. Using TD-DFT calculations, the absolute configurations of TNT_1 and TNT_2 were assigned as (*M*, *S*, *S*) and (*P*, *S*, *S*), respectively. Likewise, the configurations of TNL_1 and TNL_2 were assigned as (*M*, *S*, *R*) and (*P*, *S*, *R*), respectively (Fig. S10†).

Room-temperature EPR studies

Fig. 2 shows the EPR spectra of the four diastereoisomers in aqueous solutions at room temperature. Both TNT_1 and TNT_2 in

phosphate buffer (PB, 20 mM, pH 7.4) exhibit well-resolved triplets with line separations of ~ 8.2 G, about half the ^{14}N hyperfine splitting (α_{N} , 16.0 G) of APO. This spectral feature is characteristic of strong intramolecular exchange interaction between the two spins.³⁷ Of note is that this EPR triplet pattern was only observed for the directly linked TN biradicals in previous studies.^{32–34,37} Therefore, the strong exchange interactions for TNT_1 and TNT_2 are unexpected given that there are multiple σ bonds (proline linker) between the two radical moieties. In contrast to $\text{TNT}_{1,2}$, $\text{TNL}_{1,2}$ had much more complicated and less symmetric EPR spectra with an intense central line surrounded by multiple weak peaks (TNL_1) or one broad peak (TNL_2) at both sides. The broad EPR peak of TNL_2 at the left side indicates that it has a wide conformational distribution possibly due to the relatively high flexibility of the linker.³⁷ Since the anisotropic hyperfine and dipolar interactions of TN biradicals are averaged out in the liquid state, the exchange interactions (*J*) can be estimated from their EPR spectra in aqueous solution. Using the well-established simulation program,⁴³ the *J* values of TNL_1 and TNL_2 at room temperature are estimated to be 14 G and 32 G (Table 1), respectively. As expected, TNT_1 (252 G) and TNT_2 (128 G) have much higher *J* values than $\text{TNT}_{1,2}$. Interestingly, the β and γ relaxation parameters in the line width formula of the nitrogen hyperfine pattern obtained for $\text{TNL}_{1,2}$ are larger than those of $\text{TNT}_{1,2}$, indicating that the molecular rotation of $\text{TNL}_{1,2}$ is slower than that of $\text{TNT}_{1,2}$. The slow molecular rotation further demonstrates that $\text{TNL}_{1,2}$ has loose geometries where the effective rotational radius is rather large and the two radical moieties are far away from each other. This spatial separation of the two radical moieties in $\text{TNL}_{1,2}$ may further explain why they have small *J* values as compared to $\text{TNT}_{1,2}$. Taken together, the much stronger exchange interaction of $\text{TNT}_{1,2}$ than $\text{TNL}_{1,2}$ suggests that the *R/S* configurations of the nitroxide part play more important roles in modulating their exchange interactions than the *M/P* helices of the trityl part. This is unsurprising since the molecular structure of the trityl radical has high symmetry. Thus, it can also be deduced that the relative geometric conformation between the trityl part and the proline linker is similar for $\text{TNT}_1/\text{TNT}_2$ or $\text{TNL}_1/\text{TNL}_2$. To verify this hypothesis, the four diastereoisomers of TN biradicals were reduced with ascorbate to the corresponding trityl

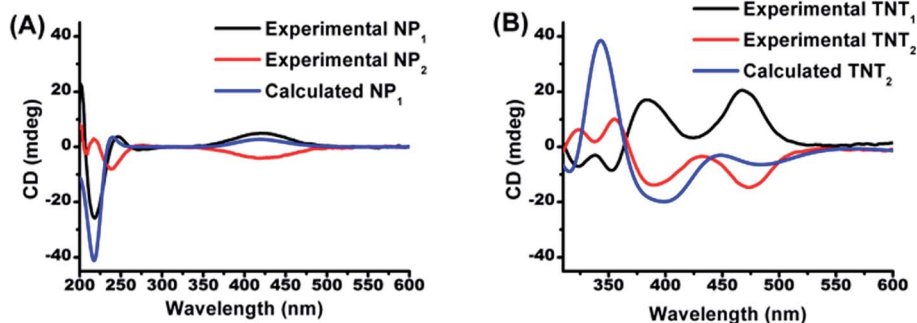


Fig. 1 Comparison of the calculated ECD spectrum of (A) NP_1 with the experimental ones of NP_1 and NP_2 and (B) of TNT_2 with the experimental ones of TNT_1 and TNT_2 .



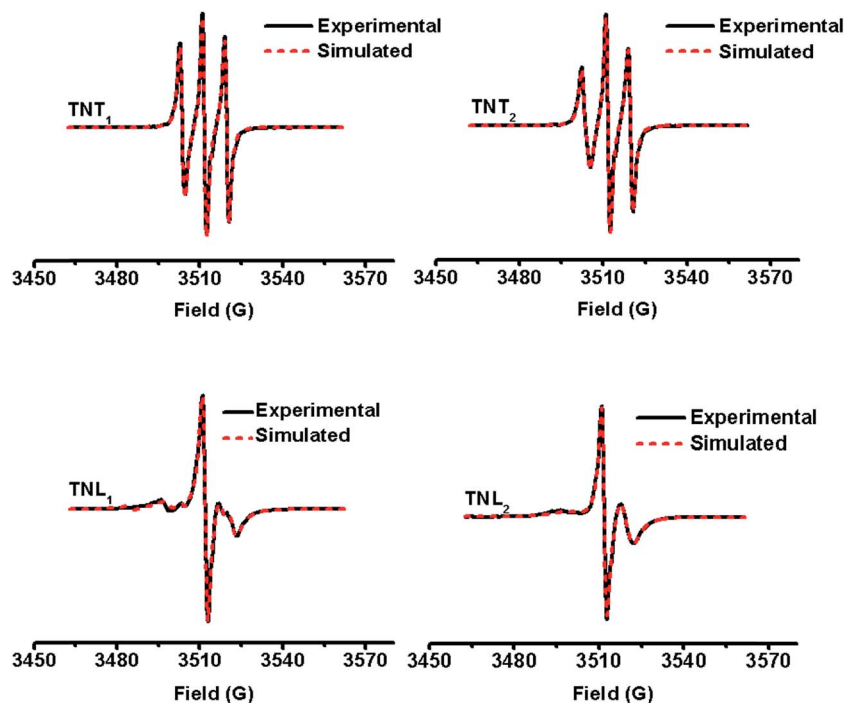


Fig. 2 Experimental (black solid line) and simulated (red dotted line) EPR spectra of TNT₁, TNT₂, TNL₁ and TNL₂ in phosphate buffer (20 mM, pH 7.4) at room temperature.

Table 1 Exchange (J , G) and dipolar (D , G) interactions of TN biradicals at room temperature (298 K) and low temperature (\sim 220 K)

Biradical	J (RT)	J (LT)	D
TNT ₁	252	196	9
TNT ₂	127	172	11
TNL ₁	14	5.4	8
TNL ₂	33	6.3	8

monoradicals and their EPR spectra were recorded. As shown in Fig. S5,[†] the four trityl monoradicals had almost identical and partially overlapped triplet signals under anaerobic conditions due to the hyperfine splitting (α_N , 0.18 G) of the nitrogen from the proline linker. The identical α_N value confirms that the relative geometric conformation between the trityl part and the proline linker is the same for the four monoradicals and the M/P helices of the trityl part have a minor effect on their exchange interactions.

Effect of temperature on the exchange interactions

The through-space exchange interaction of the biradicals with flexible linkers can be enhanced at high temperature due to the rapid rotation of the chemical bonds in the linkers which increases the collisional frequencies of the two radical parts. As mentioned above, TNL₂ has a flexible linker as evidenced by its broad EPR low-field peak. To further investigate the effect of the linker flexibility of the four diastereoisomers on their exchange interactions, their EPR spectra were recorded at high temperatures in the range of 305 K to 355 K. As shown in Fig. 3A, the EPR

spectra of TNL₂ exhibited outstanding changes with temperature with two relatively stronger peaks at both sides than the central peak at high temperature. Similar results were also observed for TNL₁. However, the EPR spectra of the two diastereoisomers TNT_{1,2} at 355 K did not exhibit a significant change except for slightly narrow linewidths (Fig. S8[†]). Spectral simulation revealed that the J values of TNL₁ and TNL₂ increased with temperature from 17 G (305 K) to 93 G (355 K) and 53 G (305 K) to 149 G (355 K), respectively. The positive response of J values with temperature for TNL_{1,2} can be explained by their flexible linkers which allow for fast rotation of the chemical bonds in the linker, increase the collisional frequencies of the two radical moieties and further enhance the through-space exchange interactions.⁴⁴ In contrast, the EPR spectra of TNT_{1,2} had no noticeable changes at different temperatures and their J values kept constant throughout the temperature range studied, implying that these two biradical diastereoisomers have rigid linkers (Fig. S8[†]). Therefore, the configuration of the nitroxide part plays a critical role in modulating the flexibility of the linkers and further the dependence of the exchange interactions on temperature.

Solvent effect on the exchange interactions

Solvents have an important influence on the exchange interaction of biradicals by changing the population of the conformations or collisional frequencies of their two radical parts. Thus, the solvent effect on the exchange interactions of the four biradical diastereoisomers was also investigated in seven solvents (PB, MeOH, DMF, acetonitrile, acetone, THF and dioxane) at room temperature. EPR spectra were recorded



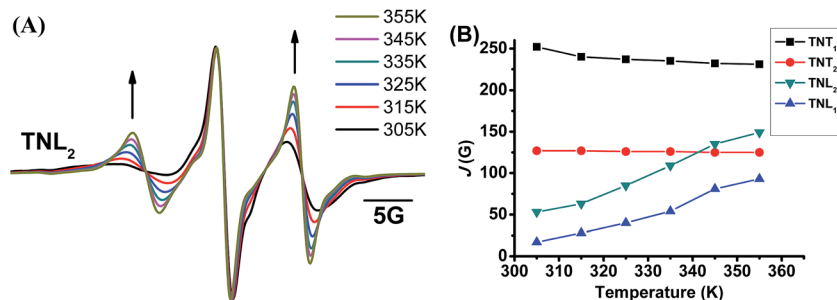


Fig. 3 (A) Normalized EPR spectra of TNL_2 in phosphate buffer (PB, 20 mM, pH 7.4) at different temperatures. (B) J values of the four biradical diastereoisomers as a function of temperature (305–355 K).

(Fig. S9†) under anaerobic conditions where the potential effect of oxygen concentration on the EPR linewidth can be eliminated. These four diastereoisomers, in general, have relatively low magnitudes of exchange interactions in polar solvents (Table 2). For example, the J values of the two diastereoisomers of $\text{TNL}_{1,2}$ exhibited a negative correlation with the solvent polarities. As for $\text{TNT}_{1,2}$, strong exchange interactions were consistently observed with no marked difference in all organic solvents. Considering the large J values of the four diastereoisomers in viscous dioxane (1.54), enhanced exchange interactions in organic solvents most likely arises from the conformational conversion instead of the increased intramolecular collisional frequencies between the two radical parts. Since $\text{TNT}_{1,2}$ have very rigid linkers, their conformations are strongly solidified and little dependence of their exchange interactions on solvents is expected.

Solid-state EPR studies and spectral simulation

It has been demonstrated that the magnitudes of both dipolar and exchange interactions are very critical for the applications of biradicals in DNP which is carried out at low temperature. For this reason, the estimation of these two parameters from the frozen-solution spectra was attempted. Fig. 4 shows the EPR spectra of the four diastereoisomeric biradicals in a 60 : 40 (v/v) glycerol/ H_2O glass-forming solution at low temperature (~ 220 K). At the first glimpse, the two diastereoisomers of $\text{TNT}_{1,2}$ have strong exchange interactions as reflected from small overall separations between the outermost lines (*i.e.*, 52 G for TNT_1 and 51 G for TNT_2) in their EPR spectra which are much smaller than $2A_{zz}$ (~ 70 G) of the nitroxide radical. Comparatively, the

exchange interactions of the two diastereoisomers of $\text{TNL}_{1,2}$ are identical and quite weak since there is a large overall separation of 73 G between the outermost lines in their EPR spectra. In order to reliably determine D and J values from the frozen solution spectra, especially when the J coupling is smaller than the dipolar and hyperfine interactions, we modified our previous EPR simulation program (ROKI/EPR)³⁷ by taking into account the anisotropic dipolar interaction besides g and hyperfine tensors. Note that this new program (called ROKI/DNP) enables the precise measurement of D and J values even determining the sign of J . Furthermore, it can also determine two polar angles which define the orientation of the linker in the principal direction of the nitroxide moiety. If the spectra are recorded at a high resonance frequency, the three Euler angles can be estimated giving the relative orientation of the principal axes of g and hyperfine tensors between the nitroxide and the trityl moieties (see more details in the Experimental section). As shown in Table 1, the order of the J values for the four biradical diastereoisomers in the frozen state is TNT_1 (196 G) \sim TNT_2 (172 G) \gg TNL_2 (6.3 G) \sim TNL_1 (5.4 G). This trend is similar to the results obtained at room temperature. As mentioned above, the flexible linkers in $\text{TNL}_{1,2}$ allow for the spatial proximity of the two radical parts and their exchange interactions are mostly based on a through-space mechanism above room temperature. In the frozen state, their molecular conformations were immobilized and their through-space exchange interactions were inhibited. Thus, the very weak exchange interactions observed for $\text{TNL}_{1,2}$ in the frozen state are most likely due to the through-bond mechanism. The close J values of TNL_1 and TNL_2 indicate that they have similar preferential conformation(s) at low temperature. The similar conformations of $\text{TNL}_{1,2}$ in the frozen state were further confirmed by their identical dipolar interactions ($D = 8$ G) which indicate the same distances between the two spins in $\text{TNL}_{1,2}$. Comparatively, the exchange interactions of $\text{TNT}_{1,2}$ did not significantly vary with temperature and had similar J values at ~ 220 K and room temperature. These temperature-insensitive exchange interactions are due to the rigid linkers in $\text{TNT}_{1,2}$ which prevent the rotation of the chemical bonds in the linker and solidify their molecular conformations.

Dynamic nuclear polarization studies

Since TN biradicals are very promising DNP polarizing agents especially at a high field,^{35,36} the DNP performances of $\text{TNT}_{1,2}$

Table 2 Exchange interaction (J , G) of $\text{TNT}_{1,2}$ and $\text{TNL}_{1,2}$ in various solvents at room temperature^a

Biradical	PB	DMF	MeOH	ACN	DMK	THF	Diox
TNT_1	252	301	342	322	310	313	373
TNT_2	127	303	308	356	378	380	383
TNL_1	14	58	58	209	227	281	351
TNL_2	33	110	101	233	276	282	293

^a PB, phosphate buffer. DMF, *N,N*-dimethylformamide. MeOH, methanol. ACN, acetonitrile. DMK, acetone. THF, tetrahydrofuran. Diox, 1,4-dioxane.



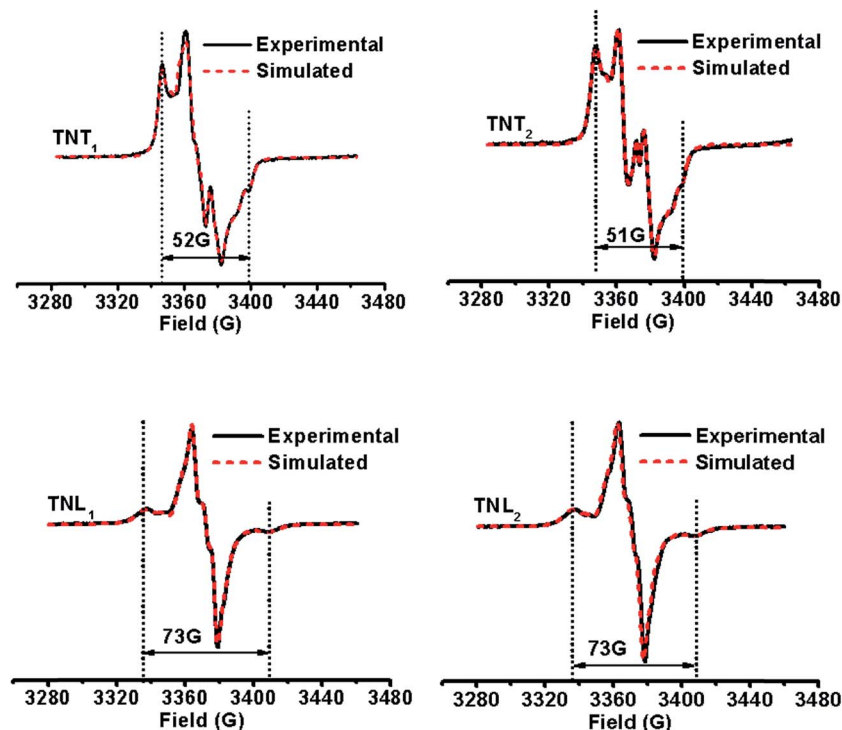


Fig. 4 Experimental (black solid line) and simulated (red dotted line) EPR spectra of TNT₁, TNT₂, TNL₁ and TNL₂ in glycerol/water (v/v, 60/40) at low temperatures (~220 K).

and TNL_{1,2} were also evaluated on an 800 MHz/527 GHz DNP/MAS solid-state NMR system.⁴⁵ The enhancements of ¹³C NMR signals of ¹³C-¹⁵N proline were measured using TNT_{1,2} and TNL_{1,2} as polarizing agents. As shown in Fig. 5, TNL_{1,2} exhibits an over 40-fold signal enhancement, whereas TNT_{1,2} has only a 7-fold enhancement. The distinct enhancements imply that the nitroxide chirality and thus the amplitude of the exchange interactions significantly influence the DNP properties of TNT_{1,2} and TNL_{1,2}. Our present study experimentally

verified the adverse effect of the too strong exchange interactions of TNT_{1,2} on its DNP enhancement although a similar result was theoretically suggested in a previous study.⁴⁶ In addition, the exchange interaction may affect the electron spin relaxation of biradicals⁴⁷ that has been shown to be a crucial factor controlling the DNP properties of nitroxide biradicals.^{48,49} Thus, the effect of the exchange interaction is multiple. Detailed theoretical simulations may quantitatively describe the effect of the exchange interactions on the DNP properties of TN biradicals.^{39,50} On the other hand, the 40-fold signal enhancement of TNL_{1,2} is somewhat lower than that previously seen for TEMTriPol-1 ($\epsilon = 65$ for ¹³C-urea) which differs from TNL_{1,2} only by the linker and has an exchange coupling of 17 G determined by the new simulation program, which was reported to be 26 G in a previous paper.³⁶ Since TNL_{1,2} ($J \sim 6$ G) has weaker exchange interaction than TEMTriPol-1 ($J = 17$ G), an optimal J value for TN biradicals may exist that is large enough for highly efficient polarization transfer but is so small as not to interfere with the frequency matching required for DNP.⁵¹

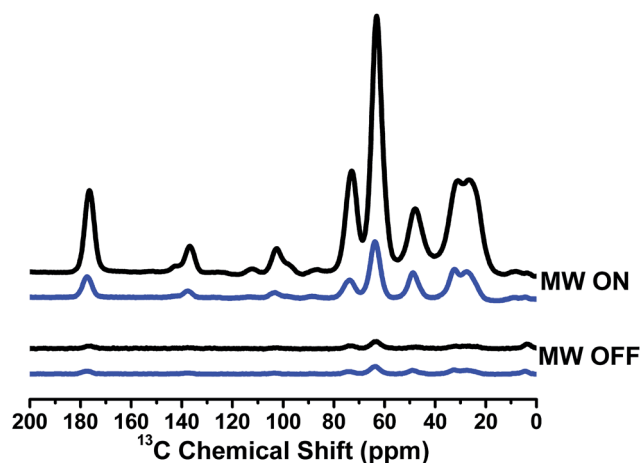


Fig. 5 ¹³C CP-MAS spectra of TNT_{1,2} (blue) and TNL_{1,2} (black) measured on the sample containing 0.25 M U-¹³C-¹⁵N proline with 10 mM TNT_{1,2} or TNL_{1,2} in [D₈]glycerol/D₂O/H₂O (60/30/10, volume ratio) at 18.8 T (800 MHz). The measurements were carried out at 100 K with a MAS rate of 8 kHz.

Conclusions

We synthesized four enantiomerically pure diastereoisomers (TNT₁, TNT₂, TNL₁ and TNL₂) of TN biradicals with different chiral configurations in the nitroxide and trityl parts. The chiral configuration of the nitroxide part plays important roles in modulating the exchange interactions of these diastereoisomers and their dependence on temperature and solvents due to the different flexibilities of the linkers. The diastereoisomeric mixture TNL_{1,2} with a small and similar J value (~6 G) has



a much higher DNP enhancement ($\epsilon = 40$) than $\text{TNT}_{1,2}$ ($f > 190$ G, $\epsilon = 7$), accounting for the adverse effect of too strong exchange interactions on the DNP enhancement.³⁶ This study demonstrates that the chiral configuration of the radical part is an effective factor modulating the exchange interaction of TN biradicals and this strategy can be used to develop new biradicals with improved DNP properties.

Experimental section

General information

All reactions were carried out under an argon atmosphere. Dichloromethane (CH_2Cl_2) was redistilled with CaH_2 and dimethylformamide (DMF) were passed through a column of molecular sieves. Boc-L-proline, 1-hydroxybenzotriazole (HOBt), (benzotriazol-1-yloxy)tris(dimethylamino)phosphoniumhexafluoro-phosphate (BOP), *N,N*-diisopropylethyl-amine (DIPEA), 2,2,5,5-tetramethyl-3-amino-pyrrolidine-1-oxyl free radical (APO) and ascorbic acid were purchased and used without purification. CT-03 was prepared according to a previously reported method.⁵² Thin layer chromatography was performed on 0.25 mm silica gel plates. Flash column chromatography was employed using silica gel with a 200–300 mesh. Thin layer chromatography plates were visualized by exposure to UV light. High-resolution mass spectrometry was carried out in methanol employing electrospray ionization (ESI) methods. High resolution mass spectrometry (HRMS) analyses were performed on a LCMS-IT-TOF Shimadzu liquid chromatograph mass spectrometer. EPR measurements were carried out on a Bruker EMX-plus X-band spectrometer. ECD spectra were recorded on a Jasco J715 spectropolarimeter (Jasco Corporation, Tokyo, Japan). Analytical HPLC was carried out on an Agilent 1100 equipped with a G1315B DAD detector and G1311A pump. Semipreparative HPLC was carried out on an SSI 1500 equipped with a UV/vis detector and versa pump.

Synthesis

NP₁ and NP₂. To a solution of Boc-L-proline (65 mg, 0.3 mmol), HOBt (111 mg, 0.82 mmol) and DIPEA (0.232 mL, 1.37 mmol) in CH_2Cl_2 (3 mL), BOP (145 mg, 0.33 mmol) was added. The resulting solution was stirred at ambient temperature for 0.5 h. Then, a solution of the nitroxide APO (43 mg, 0.27 mmol) in CH_2Cl_2 (1 mL) was added and the reaction mixture was stirred at 25 °C for another 3 h. CH_2Cl_2 (15 mL) was added and the organic layer was washed successively with 6% citric acid (10 mL), saturated solution of NaHCO_3 (10 mL) and brine (10 mL). The organic layer was dried over anhydrous sodium sulfate, filtered and concentrated *in vacuo*. The crude residue was purified by flash column chromatography on silica gel using EtOAc/petroleum ether (1 : 2) as an eluent to give two diastereoisomers NP₁ (41 mg, 42% yield) as a yellow oil and NP₂ (43 mg, 44% yield) as a yellow solid. NP₁, MS (ESI, $[\text{M} + \text{H}]^+$, m/z): 355.2 (measured), 355.2 (calculated); ($[\text{M} + \text{Na}]^+$, m/z): 377.2 (measured), 377.2 (calculated). EPR: $\alpha_{\text{N}} = 15.7$ G in MeOH/ H_2O (v/v, 1 : 2) and 14.5 G in CH_2Cl_2 . NP₂, MS (ESI, $[\text{M} + \text{H}]^+$, m/z): 355.2 (measured), 355.2 (calculated); (ESI, $[\text{M} + \text{Na}]^+$, m/z): 377.3

(measured), 377.2 (calculated). EPR: $\alpha_{\text{N}} = 15.7$ G in MeOH/ H_2O (v/v, 1/2) and 14.5 G in CH_2Cl_2 .

TNT₁ and TNT₂. A solution of NP₁ (38 mg, 0.10 mmol) in CH_2Cl_2 (1 mL) was treated with trifluoroacetic acid (TFA, 1 mL) and the resulting solution was stirred at 25 °C for 2 h. After removing the solvents under *vacuo*, the residue was redissolved in DMF (2 mL). To a solution of CT-03 (53 mg, 0.053 mmol), HOBt (22 mg, 0.16 mmol) and DIPEA (90 μL , 0.53 mmol) in DMF (3 mL), BOP (24 mg, 0.05 mmol) was added. The resulting solution was stirred at 24 °C for 0.5 h and then mixed with the solution of the deprotected NP₁ in DMF (2 mL). After stirring at 25 °C for another 18 h, the reaction mixture was poured into EtOAc (10 mL) and 1 M HCl (10 mL). The organic layer was separated and washed with brine (2×10 mL), dried over anhydrous sodium sulfate, filtered and concentrated *in vacuo*. The resulting residue was dissolved in phosphate buffer (0.2 M, pH 7.4) and purified by column chromatography on a reversed-phase C18 using water followed by 0–40% MeOH in H_2O as eluents to give the diastereoisomeric mixture TNT_{1,2} (43 mg, 61%). HPLC analysis of the mixture on the reversed-phase C18 column showed that there were two close peaks with the retention times of 11.34 (TNT₁) and 11.58 min (TNT₂) (Fig. S1†). Subsequently, TNT₁ and TNT₂ were separated by semi-preparative HPLC on a Grace Apollo column C18 (5 μm , 150 mm \times 22 mm). Column temperature: 25 °C, UV detection at 254 nm, flow rate: 5 mL min^{-1} , and eluent: acetonitrile/ NH_4OAc (20 mM, pH 6.8), 30%/70% (0 min)–50%/50% (20 min). An amount of ~ 2 mg was obtained for each diastereoisomer. TNT₁, HRMS (ESI, $[\text{M} - 2\text{H}]^{2-}$, m/z): 617.5496 (measured), 617.0540 (calculated). TNT₂, HRMS (ESI, $[\text{M} - 2\text{H}]^{2-}$, m/z): 617.5486 (measured), 617.0540 (calculated).

TNL₁ and TNL₂. A similar procedure for the synthesis and purification of TNT₁ and TNT₂ was applied. The diastereoisomeric mixture TNL_{1,2} (40 mg) was obtained from NP₂ (40 mg, 0.11 mmol) and CT-03 (55 mg, 0.055 mmol) in a yield of 59%. Attempts to separate TNL₁ and TNL₂ from the mixture by reversed-phase C18 HPLC failed. Fortunately, TNL₁ and TNL₂ were appropriately separated on a chiral HPLC CHIRALPAK® IG column (5 μm , 1.6 mm \times 250 mm) with the retention times of 7.89 and 10.65 min for TNL₁ and TNL₂, respectively (Fig. S2†). Column temperature: 25 °C, UV detection at 254 nm, flow rate: 1 mL min^{-1} , and eluent: ethanol/*n*-hexane 25%/75% (0.1% AcOH). Repeating this separation procedure (~ 50 times) afforded small amounts (~ 0.4 mg) of these two diastereoisomers. TNL₁, HRMS (ESI, $[\text{M} - 2\text{H}]^{2-}$, m/z): 617.5497 (measured), 617.0540 (calculated). TNL₂, HRMS (ESI, $[\text{M} - 2\text{H}]^{2-}$, m/z): 617.5482 (measured), 617.0540 (calculated).

Reduction of the four biradical diastereoisomers by ascorbic acid

Each diastereoisomer (50 μM) was mixed with ascorbic acid (2 mM) in PBS (20 mM, pH 7.4). After the reaction was completed, the resulting solution was transferred into a capillary tube and EPR spectra were recorded using the following EPR parameters: microwave power, 0.5 mW; modulation amplitude, 0.03 G.



EPR spectroscopy

EPR measurements were carried out on a Bruker EMX-plus X-band spectrometer at room temperature (298 K) or low temperature (~220 K). General instrumental settings were as follows: modulation frequency, 100 kHz; microwave power, 10 mW; and modulation amplitude, 1 G for room temperature and 2 G for low temperature. Measurements were performed in 50 μ L capillary tubes. Spectral simulation was performed by using the program developed by Professor Rockenbauer.⁴³ In this study, the exchange, dipolar and hyperfine couplings are given in Gauss units which can be converted into cm^{-1} by multiplying with $g \times 4.6686 \times 10^{-5}$, where g is the respective Zeeman factor. EPR measurements under anaerobic conditions were carried

Then in the Zeeman and hyperfine terms, only the z component of the electron and nuclear spin remains:

$$H_Z = H_0 g_1(\theta_1, \phi_1) \mu_B S_{z1} + H_0 g_2(\theta_2, \phi_2) \mu_B S_{z2}$$

$$H_{\text{HF}} = A_1(\theta_1, \phi_1) S_{z1} I_{z1} + A_2(\theta_2, \phi_2) S_{z2} I_{z2}$$

Here the g -factor is

$$g(\theta, \phi) = \sqrt{g_{xx}^2 \sin^2 \theta \cos^2 \phi + g_{yy}^2 \sin^2 \theta \sin^2 \phi + g_{zz}^2 \cos^2 \theta}$$

The angular dependence of the hyperfine constant is

$$A(\theta, \phi) = \sqrt{g_{xx}^2 A_{xx}^2 \sin^2 \theta \cos^2 \phi + g_{yy}^2 A_{yy}^2 \sin^2 \theta \sin^2 \phi + g_{zz}^2 A_{zz}^2 \cos^2 \theta} / g_1(\theta, \phi)$$

out using a gas-permeable Teflon tube (i.d. = 0.8 mm). Briefly, the sample solution was transferred to the tube which was then sealed at both ends. The sealed sample was placed inside a quartz EPR tube with open ends. Argon gas was allowed to bleed into the EPR tube, and then an EPR spectrum was recorded.

EPR simulation

The EPR spectra are computed by using a perturbation solution of the biradical spin Hamiltonian:

$$H_{\text{SH}} = \vec{H} \hat{g}_1 \vec{S}_1 \mu_B + \vec{H} \hat{g}_2 \vec{S}_2 \mu_B + J \vec{S}_1 \vec{S}_2 + D''(2S_{z1} S_{z2} - S_{x1} S_{x2} - S_{y1} S_{y2}) + \vec{S}_1 \hat{A}_1 \vec{I}_1 + \vec{S}_2 \hat{A}_2 \vec{I}_2$$

The computations are carried out in three different frames:

(1) The (x, y, z) reference frame is chosen as the principal axis of the \hat{g}_1 Zeeman and \hat{A}_1 hyperfine tensor of the nitroxide moiety. The orientation of the magnetic field is given by the θ and ϕ polar angles and defined in the above frame.

$$H_x = H_0 \sin \theta \cos \phi, H_y = H_0 \sin \theta \sin \phi, H_z = H_0 \cos \theta$$

(2) The slightly anisotropic \hat{g}_2 tensor and a small \hat{A}_2 hyperfine tensor of the trityl are defined in the (x', y', z') frame. The relative orientation of the two frames is characterized by the three Euler angles α, β and γ :

$$\hat{g}_2 = \hat{R}^{-1}(\alpha, \beta, \gamma) \hat{g}'_2 \hat{R}(\alpha, \beta, \gamma)$$

The same transformation defines the relationship of the hyperfine tensors \hat{A}_1 and \hat{A}_2 .

(3) The dipolar interaction D'' is defined in a third frame, where the z'' axis shows the direction of the linker and its orientation is given by two polar angles ξ and ζ . A perturbation procedure is applied in the case when the Zeeman term is large compared to the exchange, dipolar and hyperfine interactions.

The angle dependence of the dipolar interaction is given as

$$D = D''/2(1 - 3 \cos^2 \delta)$$

The δ angle between the magnetic field and the z'' axis can be expressed by the polar angles ξ and ζ :

$$\cos \delta = \sin \theta \sin \xi \cos \phi \cos \zeta + \sin \theta \sin \xi \sin \phi \sin \zeta + \cos \theta \cos \xi$$

The biradicals have four energy levels that can be assigned to the $S = 1$ triplet and $S = 0$ singlet state. Both the exchange and dipolar terms can mix the $M_S = 0$ states of the $S = 1$ triplet and $S = 0$ singlet:

$$H_{\text{Exch}} + H_{\text{Dip}} = S_{z1} S_{z2} (2D + J) + 1/2(S_{+1} S_{-2} + S_{-1} S_{+2})(J - D)$$

The diagonal elements in the $|S, M_S\rangle$ basis are

$$\langle 1, 1 | H_{\text{Exch}} + H_{\text{Dip}} | 1, 1 \rangle = \langle 1, -1 | H_{\text{Exch}} + H_{\text{Dip}} | 1, -1 \rangle = J/4 + D/2$$

$$\langle 1, 0 | H_{\text{Exch}} + H_{\text{Dip}} | 1, 0 \rangle = J/4 - D$$

$$\langle 0, 0 | H_{\text{Exch}} + H_{\text{Dip}} | 0, 0 \rangle = -3J/4$$

while the off-diagonal elements are zero.

The Zeeman and hyperfine interactions have only diagonal elements in the $|S, M_S, M_{I1}, M_{I2}\rangle$ basis if $M_S = 1$ or -1 :

$$\langle 1, 1 | H_Z + H_{\text{HF}} | 1, 1 \rangle = -\langle 1, -1 | H_Z + H_{\text{HF}} | 1, -1 \rangle = 1/2[(g_1 + g_2) \mu_B H_0 + A_1 M_{I1} + A_2 M_{I2}]$$

For the $M_S = 0$ state, these interactions give zero diagonal elements, but in this case the off-diagonal elements can be non-zero:



$$\langle 1, 0 | H_Z + H_{\text{HF}} | 0, 0 \rangle = 1/2[(g_1 - g_2)\mu_B H_0 + A_1 M_{11} - A_2 M_{12}]$$

Since the $|1, 1\rangle$ and $|1, -1\rangle$ triplet states are well separated from the central states, the eigenvalues can be obtained directly from the diagonal elements. We introduce the indices $i = 1$ and 4 for the highest $|1, 1\rangle$ and lowest $|1, -1\rangle$ states, respectively. Furthermore $i = 2$ and 3 denote the two central states with $M_S = 0$.

$$E_{1, M_{11}, M_{12}} = -E_{4, M_{11}, M_{12}} = J/4 + D/2 + 1/2(g_1 + g_2)\mu_B H_0 + 1/2(A_1 M_{11} + A_2 M_{12})$$

The mixing of $|1, 0\rangle$ and $|0, 0\rangle$ states can be obtained by the solution of the secular equation:

$$|2\rangle = p_1|1, 0\rangle + p_2|0, 0\rangle$$

$$|3\rangle = p_2|1, 0\rangle - p_1|0, 0\rangle$$

The mixing coefficients can be expressed by the factor R :

$$R^2 = (J - D)^2 + ((g_1 - g_2)\mu_B H_0 + A_1 M_{11} - A_2 M_{12})^2$$

$$p_1^2 = 1/2|R + J - D|/R, p_2^2 = 1/2|R - J + D|/R$$

It is worth noting that the transition probability and energy depend on the relative sign of D and J , which offers a possibility to decide from the frozen solution EPR spectra whether the Heisenberg exchange is ferro- or anti-ferromagnetic. Then, the energy expressions for the central states are

$$E_{i, M_{11}, M_{12}} = -J/4 - D/2 \pm R/2, i = 2, 3.$$

Due to the triplet-singlet mixing, four transitions can take place in the whole M_{11}, M_{12} subspace. The two resonances have the transition probability

$$W_{12} = W_{24} = |\langle 1 | S_+ | 2 \rangle|^2 = 2p_1^2$$

The respective frequency is

$$\hbar\omega_{12 \text{ or } 24} = 1/2(g_1 + g_2)\mu_B H_0 + 1/2(A_1 M_{11} + A_2 M_{12}) \pm (J/2 + D - R/2).$$

The above resonances can be called “allowed”, since their intensities are larger than 1.0 and the line separations are smaller than J . The other pair of resonances has the transition probability

$$W_{13} = W_{34} = |\langle 1 | S_+ | 3 \rangle|^2 = 2p_2^2$$

This transition can be considered as “forbidden” due to the smaller transition probability: $2p_2^2 < 1.0$ and the frequency spread can be extended in a broad range comparable with the value of J :

$$\hbar\omega_{13 \text{ or } 34} = 1/2(g_1 + g_2)\mu_B H_0 + 1/2(A_1 M_{11} + A_2 M_{12}) \pm (J/2 + D + R/2)$$

For a fixed ω_{EPR} microwave frequency the above relationships allow the determination of the corresponding four resonance fields $H_{i, M_{11}, M_{12}}(\theta, \phi)$, where the i index denotes the four transitions.

The ROKI/DNP program can use either Lorentzian or Gaussian, or mixed line shapes $f(H)$ and the powder pattern can be obtained by integrating over the angles θ and ϕ :

$$\text{EPR}(H) = \sum_{i=1}^4 \sum_{M_{11}=-I}^I \sum_{M_{12}=-I}^I \int_{\theta=0}^{\pi} \int_{\phi=0}^{2\pi} f(H - H_{i, M_{11}, M_{12}}(\theta, \phi)) \sin \theta d\theta d\phi$$

All parameters can be optimized to achieve the best fit of the experimental spectra; the method of optimization is the same as that described in the ROKI/EPR program.⁴³ Due to the large number of optimized parameters ambiguity can occur. This problem can be reduced by recording the frozen solution spectra at different resonance frequencies. In X band spectra the determination of the three Euler angles is rather problematic due to the small g anisotropy of the trityl moiety. The polar angles, however, are more reliable since the nitroxide part has a large Zeeman and hyperfine anisotropy.

DNP/ssNMR spectroscopy of the diastereoisomeric mixtures TNT_{1,2} and TNL_{1,2}

DNP experiments were performed on frozen solutions of 10 mM biradical in d8-glycerol : D₂O : H₂O 60 : 30 : 10 v/v/v with 0.25 M U-¹³C-¹⁵N proline. DNP experiments at 800 MHz were performed on a Bruker BioSpin 527 GHz solid-state NMR DNP spectrometer. This spectrometer is equipped with a Bruker 800 WB/RS Plus magnet with a sweep coil, an Avance III NMR console, and a low-temperature 3.2 mm triple-resonance DNP MAS NMR probe.⁴⁵ A gyrotron microwave source emits microwaves at a frequency of 527.043 GHz. In the DNP experiments the nuclear polarization is measured through the spectrum of ¹³C-¹⁵N proline, which is observed *via* ¹³C-¹H cross-polarization (CP). A CP spin-locking field of 50 kHz is applied on ¹³C, while the spin-locking field on ¹H is ramped from 36 to 45 kHz. The contact time was set to 2 ms. During the acquisition SPINAL-64 decoupling⁵³ was used at 83 kHz. Each CP acquisition was preceded by a presaturation sequence consisting of 300 90° ¹H pulses 20 μs apart at an RF power of 100 kHz, followed by a period of polarization build-up, which was set to 1.26T₁ for optimal sensitivity. Each spectrum was acquired with a 4-step phase cycle and acquired three times to confirm stability and reproducibility. The MAS frequency was set to 8 kHz and the sample temperature was kept at 103 K, unless noted otherwise.



The enhancement, ϵ , was obtained from comparison of the ^{13}C signal amplitude with and without microwaves. T_1 (microwaves off) and T_B (microwaves on) were measured *via* a saturation recovery experiment. T_2 was determined from the decay of the echo intensity during a rotor-synchronized Hahn echo sequence. To obtain the CE DNP enhancement field profiles the magnetic field is swept and at each field position the nuclear polarization is measured *via* the cross-polarization experiment as described above.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- Q. Z. Ni, E. Daviso, T. V. Can, E. Markhasin, S. K. Jawla, T. M. Swager, R. J. Temkin, J. Herzfeld and R. G. Griffin, *Acc. Chem. Res.*, 2013, **46**, 1933–1941.
- A. J. Rossini, A. Zagdoun, M. Lelli, A. Lesage, C. Coperet and L. Emsley, *Acc. Chem. Res.*, 2013, **46**, 1942–1951.
- H. Iwamura and N. Koga, *Acc. Chem. Res.*, 1993, **26**, 346–351.
- A. Rajca, *Chem. Rev.*, 1994, **94**, 871–893.
- M. Abe, *Chem. Rev.*, 2013, **113**, 7011–7088.
- E. A. Weiss, M. J. Ahrens, L. E. Sinks, A. V. Gusev, M. A. Ratner and M. R. Wasielewski, *J. Am. Chem. Soc.*, 2004, **126**, 5577–5584.
- A. B. Ricks, K. E. Brown, M. Wenninger, S. D. Karlen, Y. A. Berlin, D. T. Co and M. R. Wasielewski, *J. Am. Chem. Soc.*, 2012, **134**, 4581–4588.
- M. A. Harris, A. K. Mishra, R. M. Young, K. E. Brown, M. R. Wasielewski and F. D. Lewis, *J. Am. Chem. Soc.*, 2016, **138**, 5491–5494.
- G. I. Roshchupkina, A. A. Bobko, A. Bratasz, V. A. Reznikov, P. Kuppussamy and V. V. Khramtsov, *Free Radicals Biol. Med.*, 2008, **45**, 312–320.
- E. A. Legenzov, S. J. Sims, N. D. Dirda, G. M. Rosen and J. P. Kao, *Biochemistry*, 2015, **54**, 6973–6982.
- D. A. Shultz, R. M. Fico, H. Lee, J. W. Kampf, K. Kirschbaum, A. A. Pinkerton and P. D. Boyle, *J. Am. Chem. Soc.*, 2003, **125**, 15426–15432.
- D. A. Shultz, A. K. Boal and G. T. Farmer, *J. Am. Chem. Soc.*, 1997, **119**, 3846–3847.
- A. Rajca, A. Olankitwanit and S. Rajca, *J. Am. Chem. Soc.*, 2011, **133**, 4750–4753.
- A. Olankitwanit, M. Pink, S. Rajca and A. Rajca, *J. Am. Chem. Soc.*, 2014, **136**, 14277–14288.
- N. M. Gallagher, A. Olankitwanit and A. Rajca, *J. Org. Chem.*, 2015, **80**, 1291–1298.
- N. M. Gallagher, J. J. Bauer, M. Pink, S. Rajca and A. Rajca, *J. Am. Chem. Soc.*, 2016, **138**, 9377–9380.
- M. L. Kirk, D. A. Shultz, E. C. Depperman, D. Habel-Rodriguez and R. D. Schmidt, *J. Am. Chem. Soc.*, 2012, **134**, 7812–7819.
- M. L. Kirk, D. A. Shultz, D. E. Stasiw, D. Habel-Rodriguez, B. Stein and P. D. Boyle, *J. Am. Chem. Soc.*, 2013, **135**, 14713–14725.
- D. E. Stasiw, J. Zhang, G. Wang, R. Dangi, B. W. Stein, D. A. Shultz, M. L. Kirk, L. Wojtas and R. D. Sommer, *J. Am. Chem. Soc.*, 2015, **137**, 9222–9225.
- C. Song, K. N. Hu, C. G. Joo, T. M. Swager and R. G. Griffin, *J. Am. Chem. Soc.*, 2006, **128**, 11385–11390.
- C. Ysacco, E. Rizzato, M. A. Virolleaud, H. Karoui, A. Rockenbauer, M. F. Le, D. Siri, O. Ouari, R. G. Griffin and P. Tordo, *Phys. Chem. Chem. Phys.*, 2010, **12**, 5841–5845.
- M. K. Kiesewetter, B. Corzilius, A. A. Smith, R. G. Griffin and T. M. Swager, *J. Am. Chem. Soc.*, 2012, **134**, 4537–4540.
- E. L. Dane, B. Corzilius, E. Rizzato, P. Stocker, T. Maly, A. A. Smith, R. G. Griffin, O. Ouari, P. Tordo and T. M. Swager, *J. Org. Chem.*, 2012, **77**, 1789–1797.
- C. Sauvee, M. Rosay, G. Casano, F. Aussenac, R. T. Weber, O. Ouari and P. Tordo, *Angew. Chem., Int. Ed.*, 2013, **52**, 10858–10861.
- A. P. Jagtap, M. A. Geiger, D. Stoppler, M. Orwick-Rydmark, H. Oschkinat and S. T. Sigurdsson, *Chem. Commun.*, 2016, **52**, 7020–7023.
- D. J. Kubicki, G. Casano, M. Schwarzwald, S. Abel, C. Sauvee, K. Ganesan, M. Yulikov, A. J. Rossini, G. Jeschke, C. Coperet, A. Lesage, P. Tordo, O. Ouari and L. Emsley, *Chem. Sci.*, 2015, **7**, 550–558.
- C. Sauvee, G. Casano, S. Abel, A. Rockenbauer, D. Akhmetzyanov, H. Karoui, D. Siri, F. Aussenac, W. Maas and R. T. Weber, *Chem.–Eur. J.*, 2016, **22**, 5598–5606.
- J. Y. Chen, N. Jayaraj, S. Jockusch, M. F. Ottaviani, V. Ramamurthy and N. J. Turro, *J. Am. Chem. Soc.*, 2008, **130**, 7206–7207.
- S. Yi, B. Captain, M. F. Ottaviani and A. E. Kaifer, *Langmuir*, 2011, **27**, 5624–5632.
- C. Casati, P. Franchi, R. Pievo, E. Mezzina and M. Lucarini, *J. Am. Chem. Soc.*, 2012, **134**, 19108–19117.
- F. Romano, R. Manoni, P. Franchi, E. Mezzina and M. Lucarini, *Chem.–Eur. J.*, 2015, **21**, 2775–2779.
- Y. P. Liu, F. A. Villamena, A. Rockenbauer and J. L. Zweier, *Chem. Commun.*, 2010, **46**, 628–630.
- Y. P. Liu, F. A. Villamena, Y. G. Song, J. Sun, A. Rockenbauer and J. L. Zweier, *J. Org. Chem.*, 2010, **75**, 7796–7802.
- Y. P. Liu, Y. G. Song, A. Rockenbauer, J. Sun, C. Hemann, F. A. Villamena and J. L. Zweier, *J. Org. Chem.*, 2011, **76**, 3853–3860.
- K. N. Hu, V. S. Bajaj, M. Rosay and R. G. Griffin, *J. Chem. Phys.*, 2007, **126**, 44512.



- 36 G. Mathies, M. A. Caporin, V. K. Michaelis, Y. P. Liu, K. N. Hu, D. Mance, J. L. Zweier, M. Rosay, M. Baldus and R. G. Griffin, *Angew. Chem., Int. Ed.*, 2015, **54**, 11770–11774.
- 37 Y. P. Liu, F. A. Villamena, A. Rockenbauer, Y. G. Song and J. L. Zweier, *J. Am. Chem. Soc.*, 2013, **135**, 2350–2356.
- 38 X. L. Tan, Y. G. Song, H. Q. Liu, Q. W. Zhong, A. Rockenbauer, F. A. Villamena, J. L. Zweier and Y. P. Liu, *Org. Biomol. Chem.*, 2016, **14**, 1694–1701.
- 39 F. Mentink-Vigier, G. Mathies, Y. P. Liu, A. L. Barra, M. A. Caporini, D. Lee, S. Hediger, R. G. Griffin and G. D. Paepe, *Chem. Sci.*, 2017, **8**, 8150–8163.
- 40 B. Driesschaert, R. Robiette, F. Lucaccioni, B. Gallez and J. Marchand-Brynaert, *Chem. Commun.*, 2011, **47**, 4793–4795.
- 41 V. M. Tormyshev, A. M. Genaev, G. E. Salnikov, O. Y. Rogozhnikova, T. I. Troitskaya, D. V. Trukhin, V. I. Mamatyuk, D. S. Fadeev and H. J. Halpern, *Eur. J. Org. Chem.*, 2012, **3**, 623–629.
- 42 J. Suk, V. R. Naidu, X. Liu, M. S. Lah and K. S. Jeong, *J. Am. Chem. Soc.*, 2011, **133**, 13938–13941.
- 43 A. Rockenbauer and L. Korecz, *Appl. Magn. Reson.*, 1996, **10**, 29–43.
- 44 S. Sankarapandi, M. Sukumar, P. Balaram and P. T. Manoharan, *Biochem. Biophys. Res. Commun.*, 1995, **213**, 439–446.
- 45 E. J. Koers, E. A. W. van der Crujisen, M. Rosay, M. Weingarh, A. Prokofyev, C. Sauvee, O. Ouari, J. van der Zwan, O. Pongs, P. Tordo, W. E. Maas and M. Baldus, *J. Biomol. NMR*, 2014, **60**, 157–168.
- 46 K. N. Hu, G. T. Debelouchina, A. A. Smith and R. G. Griffin, *J. Chem. Phys.*, 2011, **134**, 125105.
- 47 H. Sato, V. Kathirvelu, G. Spagnol, S. Rajca, A. Rajca, S. S. Eaton and G. R. Eaton, *J. Phys. Chem. B*, 2008, **112**, 2818–2828.
- 48 A. Zaghdoun, G. Casano, O. Ouari, M. Schwarzwald, A. J. Rossini, F. Aussenac, M. Yulikov, G. Jeschke, C. Coperet, A. Lesage, P. Tordo and L. Emsley, *J. Am. Chem. Soc.*, 2013, **135**, 12790–12797.
- 49 D. J. Kubicki, G. Casano, M. Schwarzwald, S. Abel, C. Sauvee, K. Ganesan, M. Yulikov, A. J. Rossini, G. Jeschke, C. Coperet, A. Lesage, P. Tordo, O. Ouari and L. Emsley, *Chem. Sci.*, 2016, **7**, 550–558.
- 50 D. Mance, P. Gast, M. Huber, M. Baldus and K. L. Ivanov, *J. Chem. Phys.*, 2015, **142**, 234201.
- 51 K. N. Hu, C. Song, H. H. Yu, T. M. Swager and R. G. Griffin, *J. Chem. Phys.*, 2008, **128**, 052302.
- 52 Y. P. Liu, F. A. Villamena, J. Sun, Y. Xu, I. Dhimitruka and J. L. Zweier, *J. Org. Chem.*, 2008, **73**, 1490–1497.
- 53 B. M. Fung, A. K. Khitrin and K. Ermolaev, *J. Magn. Reson.*, 2000, **142**, 97–101.

