

Cite this: *Chem. Sci.*, 2018, 9, 4703Received 28th January 2018
Accepted 22nd April 2018

DOI: 10.1039/c8sc00450a

rsc.li/chemical-science

Catalytic synthesis of functionalized (polar and non-polar) polyolefin block copolymers†

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Herein, we report a methodology for the synthesis of polyolefin containing block-copolymers using a catalytic postpolymerization modification strategy. The most common polyolefin grades are converted into macroinitiators using a cross-metathesis reaction. These functionalized polyolefins are then used to initiate living: coordinative ring opening polymerization of lactide, anionic ring opening polymerization of epoxide, and radical polymerization of styrene to yield the corresponding block copolymers. The high activity of the catalysts employed in the different steps offers improved practicality for scalable synthesis.

Introduction

Over the past few decades, block copolymers have emerged as a class of soft materials with a wide range of technological applications.^{1–3} Due to the high tunability of their chemical structure (morphology, architecture, and domain size), block copolymers have been utilized as surfactants, thermoplastic elastomers, nano-templates, membranes, *etc.*⁴ Controlled polymerizations such as ionic, controlled radical (ATRP, NMP, RAFT), and ring opening polymerizations have been the standard means for producing block copolymers.¹ While these methods have proven successful, they lack the ability to homopolymerize the world's two most produced and inexpensive monomers: ethylene and propylene.^{5,6}

Polyolefins today are produced industrially *via* catalytic insertion (co)polymerization of ethylene, propylene, and linear α -olefins on the scale of 70×10^6 metric tons per year.^{5,7} The tunability of the polymer's crystallinity offers a mixture of properties such as toughness, elasticity, solvent resistance, *etc.* which are difficult to reproduce economically by other monomers.^{6,8} Therefore, the incorporation of polyolefins into block copolymers is of great value, as it would further expand the usefulness of the largest family of polymers in our society. However, due to the high oxophilicity of the insertion metal catalysts used in industry, commercial polyolefin block copolymers have been limited to non-polar monomers, which also limits applications.^{6,7,9} Hence, the synthesis of functionalized block copolymers containing polyolefins remains a modern challenge for synthetic chemists.

Over the years, alternative approaches have been developed to synthesize functionalized polyolefin block copolymers.

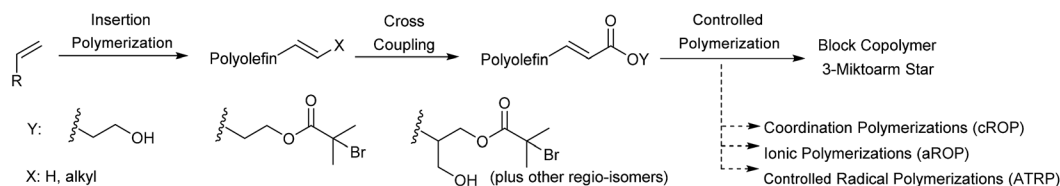
Nearly all approaches rely on the formation of (semi-) telechelic polymers which can be used to introduce a polar functionality.^{10,11} The most popular of these approaches employs the anionic polymerization of butadiene quenched by an epoxide, followed by the hydrogenation, of the polymer to yield a hydroxyl terminated linear low density polyethylene (LLDPE) block.¹² This approach is efficient, but is not compatible with the synthesis of polypropylene blocks and uses stoichiometric amounts of a pyrophoric initiator. Ring opening metathesis (ROMP) of cyclic alkenes has also been implemented to yield telechelic high density polyethylene (HDPE).^{13,14} Despite being successful, employing monomers such as butadiene and cyclic alkenes remains less favorable than the direct polymerization of ethylene and their inability to yield block copolymers containing stereoregular polypropylene (PP) is limiting. Living insertion polymerizations have been reported to produce telechelic polyolefins from ethylene and propylene, however, these systems only produce one polymer chain per metal which drastically limits their commercial potential.¹⁵ Chain transfer insertion polymerizations have been reported to yield multiple chains of polymer per catalytic site. However, this approach still requires high loadings of metal(loid) chain transfer agents and also provides low stereocontrol.^{16–19}

In this work, we develop a general methodology to produce HDPE, PP, LLDPE and HBPE (hyperbranched polyethylene) containing block copolymers (Scheme 1). Our approach consists of performing three consecutive reactions: first a traditional insertion polymerization to yield alkene terminated polyolefins, second a cross metathesis reaction to functionalize the polyolefin end-group, third a controlled living polymerization initiated by the end functionalized polyolefins. Our approach takes advantage of the selectivity, productivity and chemical versatility of the different catalytic reactions employed. In step 1, the catalytic olefin polymerizations yield exclusively olefin terminated polymers through β -X elimination or chain transfer to monomer.²⁰ In step 2, the selectivity of the

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† Electronic supplementary information (ESI) available: Experimental procedures, NMR spectra, GPC chromatogram, DSC traces. See DOI: 10.1039/c8sc00450a





Scheme 1 Generalized route for block copolymer synthesis.

metathesis reaction between olefins and acrylates quantitatively converts the polyolefins into macroinitiators. In step 3, the living coordinative Ring Opening Polymerization (cROP) of lactide,¹⁴ anionic Ring Opening Polymerization (aROP) of epoxide,²¹ and Atom Transfer Radical Polymerization (ATRP) of styrene²² are employed to quantitatively synthesize the block copolymers from the polyolefin macroinitiators (Scheme 1). Finally, we also demonstrate the advantage of using exclusively catalytic reactions to synthesize olefin containing block copolymers by comparing the overall productivity of our process to previously reported strategies.

Prior to our work, other groups have employed post-polymerization modification techniques to convert vinylic terminated polyolefins into macroinitiators. Hydrosilation,²³ thiol-ene,²⁴ hydroalumination^{25,26} and esterification²⁷ reactions have been applied with moderate to high conversions. It is worth noting that these reactions were only performed on low molecular weight vinylic terminated polyolefins ($M_n < 5 \text{ kg mol}^{-1}$). The non-quantitative conversion of most of these reactions and the exclusive reactivity toward the vinyl terminated polymers (which is catalyst dependent and often not the most common end-group) drastically restricts the impact of these previous methods. Additionally, *in situ* cross coupling compatibilization has been attempted by Duchateau *et al.* resulting in a process with less than 50% efficiency for producing the desired block copolymer.²⁸ The approach reported here, aims to address these limitations by quantitatively converting mono and di-substituted alkene terminated polyolefins, of any molecular weight into block copolymers.

Results and discussion

A series of single site homogenous insertion catalysts were used to produce a polyolefin library containing HDPE, PP, LLDPE

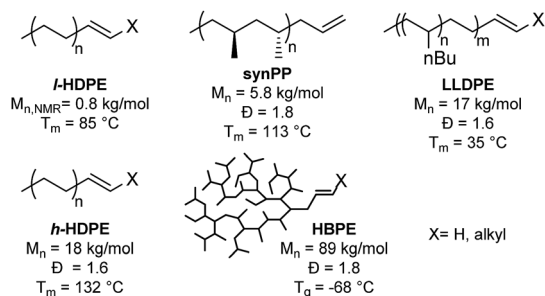


Chart 1 Polyolefin library.

and HBPE (Chart 1).^{29–32} Two grades of linear polyethylene were synthesized. Low molecular weight linear polyethylene (*l*-HDPE) with an $M_n = 800 \text{ g mol}^{-1}$ was produced from a phosphinosulfonate nickel catalyst,³⁰ and serves as a model system since there is a high concentration of end-groups which eases the characterization of the product by NMR spectroscopy. While a higher molecular weight semi-crystalline HDPE (*h*-HDPE), $M_n = 18\,000 \text{ g mol}^{-1}$, with an industrially relevant melting temperature ($T_m = 132 \text{ °C}$) was produced from a phosphinosulfonate palladium complex.²⁹ *h*-HDPE was implemented to show that the methodology is quantitative for various molecular weights and for a variety of polymerization catalysts. A bis(phenoxyimine)titanium dichloride catalyst³¹ was implemented to synthesize the stereoselective synPP and LLDPE. Finally, a Pd-diimine catalyst³² was used to produce an amorphous polyethylene that we refer to as HBPE, with a comparatively higher molecular weight of $M_n = 89\,000 \text{ g mol}^{-1}$. The molecular structure of this polymer makes it soluble at room temperature in most organic solvents which is advantageous for the characterization of the material as the low concentration of end-group makes it otherwise challenging to quantify conversion.

As recently illustrated by Mecking *et al.*, implementing a ruthenium catalyzed cross coupling metathesis represents an efficient approach to functionalizing amorphous polyolefins.³³ However, in order to implement cross coupling metathesis for the functionalization of the full library of polyolefins, it was necessary to identify reaction conditions compatible with the elevated temperature (120 °C) necessary to solubilize semi-crystalline polyolefins. We found that by adding 1 mol% of Hoveyda–Grubbs II catalyst³⁴ (HG-catalyst) simultaneously with a cross coupling partner, 2-hydroxyethyl acrylate (HEA) or 2-(2-bromoisobutyryloxy) ethyl acrylate (BIEA), over 30 minutes was highly effective for converting *l*-HDPE end groups. Slow feeding of the olefin metathesis catalyst was necessary to thwart its thermal deactivation. Fig. 1 presents the ¹H NMR for *l*-HDPE, illustrating quantitative conversion (see ESI† Section 9 for sensitivity experiments) of both terminal (5.9 ppm and 5.0 ppm) and internal olefins (5.5 ppm).³³

Next, we applied similar reaction conditions to the full library of polyolefins with only minor adjustments to HEA feed procedure depending on the ratio of internal to terminal double bonds present. Quantitative conversions of the polyolefin double bonds were achieved for all cross coupling reactions for all polyolefins as determined by ¹H NMR (Table 1). Stacked spectrums for the different polymers can be found in the ESI† (S1–S8). The absence of internal olefinic double bonds in the ¹H



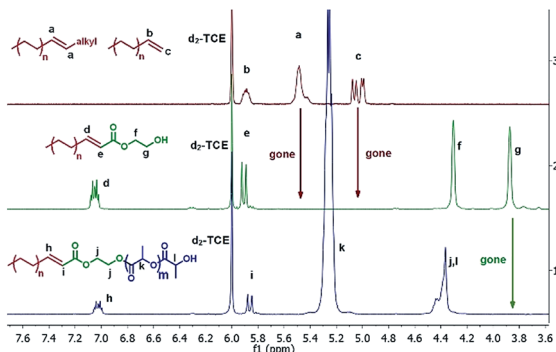


Fig. 1 Spectrum of *l*-HDPE progressing through the methodology. Top: alkene functionalized *l*-HDPE. Middle: *l*-HDPE after cross metathesis, Table 1 entry 1-1. Bottom: *l*-HDPE-*b*-PLA, Table 2 entry 2-1.

Table 1 Cross metathesis reactions

Entry ^a	Polyolefin	Acrylate	Conv. ^b (%)	M_n^c (kg mol ⁻¹)	\mathcal{D}^c
1-1	<i>l</i> -HDPE	HEA	Quant	0.9 ^d	— ^e
1-2	<i>h</i> -HDPE		Quant	19	1.5
1-3	synPP		Quant	5.5	1.8
1-4	LLDPE		Quant	17	1.5
1-5	HBPE		Quant	95	1.7
1-6	<i>l</i> -HDPE	BIEA	Quant	1.1 ^d	— ^e

^a See ESI for conditions. ^b Determined from ¹H NMR of precipitated product. ^c Determined from GPC or HT-GPC. ^d Determined by NMR. ^e Molecular weight too low for GPC analysis.

NMR spectrum at 5.5 ppm confirms that no polyolefin homopolymer remained (see ESI† Section 8 for homopolymer ¹H NMR). The original polyolefins may participate in homo-coupling, but the homo-coupling product will be easily re-activated and will eventually be paired with HEA. Moreover, acrylates are ideal coupling partners because electron deficient olefins are known to be slow at homodimerization.³⁵ This is confirmed by our own experiments when we subjected only HEA to our reaction conditions and did not observe any homo-coupling product formation.

HEA-functionalized polyolefins were used as macroinitiators for cROP of lactide and aROP of *tert*-butyl glycidyl ether (*t*BuGE), while BIEA functionalized HDPE was used to initiate the ATRP of styrene (Table 2). The implementation of controlled polymerizations ensures that all polyolefin chains are converted into block copolymers and that no homopolymer is formed. We performed the immortal cROP of lactide on *l*-HDPE, *h*-HDPE, synPP, LLPDE, and HBPE (Table 2 entry 2-1 to 2-5). Sub-stoichiometric loadings of Sn(Oct)₂ were used to catalyze the cROP. Evidence of complete conversion of the macroinitiator was demonstrated by the total disappearance of the methylene protons adjacent to the hydroxyl group (3.8 ppm) in ¹H NMR, Fig. 1.

To further showcase the versatility of our approach, we performed ATRP of styrene resulting in quantitative conversion of BIEA functionalized *l*-HDPE to yield *l*-HDPE-PS block copolymer (entry 2-6). Additionally, we carried out the immortal aROP of an

Table 2 Living chain growth polymerizations

Entry ^a	Block copolymer	Conv. ^b (%)	f_a^c	M_n^d (kg mol ⁻¹)	\mathcal{D}^d
2-1	<i>l</i> -HDPE-PLA	Quant.	0.54	4.2	1.4
2-2	<i>h</i> -HDPE-PLA	Quant.	0.47	69	1.3
2-3	synPP-PLA	Quant.	0.58	15	1.6
2-4	LLDPE-PLA	Quant.	0.72	29	1.2
2-5	HBPE-PLA	Quant.	0.84	170	1.8
2-6	<i>l</i> -HDPE-PS	Quant.	0.46	5.0	1.1
2-7	<i>l</i> -HDPE-P(<i>t</i> BuGE)	— ^e	0.29	7.5	1.2

^a See ESI for conditions. ^b Determined from ¹H NMR of precipitated product. ^c Mole fraction of polyolefin. ^d Determined from GPC or HT-GPC. ^e Unable to determine due to overlapping signals in the ¹H NMR.

epoxide monomer by using sub-stoichiometric amounts of KN(TMS)₂ (entry 2-7).²¹ However, to perform this anionic polymerization, it was necessary to reduce the acrylate linker of *l*-HDPE-HEA into a primary alcohol to avoid side reactions.³³

More advanced architectures can be achieved beyond simple diblock copolymers with this methodology. An example of this is a 3-miktoarm star. Star polymers have gained much attention in literature recently for their unique ability to yield highly asymmetric lamellae domains, which are not accessible by linear diblock copolymers.³⁶ By cross coupling a polyolefin to a dual initiating coupling partner, one that contains an alcohol and 2-bromoisobutyrate functionality, both cROP of lactide and ATRP of styrene can be done resulting in a *l*-HDPE-PLA-PS 3-miktoarm star (Fig. 2). Efficient synthesis of the 3 arm star is demonstrated by the increase in molecular weight (GPC chromatograms in ESI Fig. 57†) after each step.

In addition to chemical versatility, our approach offers another significant benefit, productivity. The scalability of any material synthesis is a key consideration when thinking about applications.³⁷ In which, one of the main limitations preventing the broader application of block copolymers comes from the overall low productivity of the living polymerization methods. They require using stoichiometric amounts of initiator per polymer chain and/or the use of expensive monomers/initiator. Our system addresses this limitation by implementing catalytic reactions and using industrially relevant monomers. We compared the productivity of our methodology to other

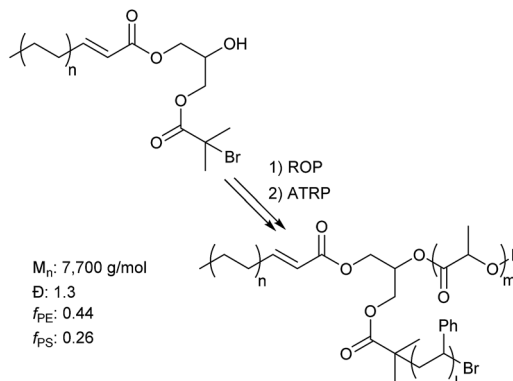
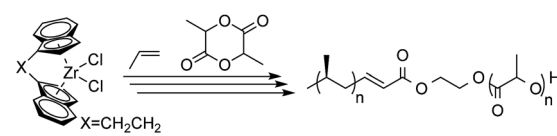


Fig. 2 Synthesis of 3-miktoarm star.



Table 3 Productivity calculations^a


Polymer	Method	g poly/g cat	g poly/g metal
<i>h</i> -HDPE-PLA ^b	Anionic/Pd/AlEt ₃ (ref. 12)	4	70
	ROMP/Pt/Sn(Oct) ₂ (ref. 14)	34	1900
	CTA/O ₂ /Sn(Oct) ₂ , ^c (ref. 16)	50	420
	This work	3200	20 000
isoPP-PLA ^d	This work	29 000	125 000
isoPP	4th Gen. Ziegler Natta, ^e (ref. 39)	—	600 000

^a See ESI for calculations, Section 10. ^b $M_n = 64 \text{ kg mol}^{-1}$, $f_{\text{HDPE}} = 0.5$ (mole frac.). ^c Contains 20% homopolymer. ^d $M_n = 71 \text{ kg mol}^{-1}$, $f_{\text{PP}} = 0.5$, contains 30% homopolymer. ^e $\text{MgCl}_2/\text{ester}/\text{TiCl}_4/\text{AlEt}_3/\text{PhSi}(\text{OEt})_3$.

published methods for the synthesis of *h*-HDPE-PLA with a mole fraction of 0.50 and a $M_n = 64 \text{ kg mol}^{-1}$ (Table 3). Without significant optimization of the reaction conditions, we were able to achieve an overall productivity of 20 000 g of polymer per gram of metal, which is over one order of magnitude more productive than previously reported systems. This productivity could be further improved by applying the methodology to isotactic polypropylene made from a metallocene catalyst.³⁸ Indeed, the high productivity of the catalyst enabled us to achieve an overall productivity of 125 000 g of polymer per gram of metal for the synthesis of isoPP-PLA with a mole fraction of 0.50 and a $M_n = 71 \text{ kg mol}^{-1}$. To put these numbers into perspective, industrially produced polypropylene is predominantly made with Ziegler–Natta catalysts which produces 600 000 grams of polymer per gram of metal.³⁹ However this improvement in productivity comes at the expense of conversion. Indeed the vinylidene terminated polymer chains were unreactive toward the metathesis reaction (30 mol%) resulting in the presence of homopolymer of iPP in the block copolymer.

Conclusions

In summary, we have developed a facile method to synthesize polyolefin containing block copolymers by implementing three catalytic reactions subsequently. A series of olefin terminated polyolefins were synthesized by catalytic insertion polymerizations. These polymers were then functionalized by cross metathesis reactions to yield polyolefin macroinitiators, which could further be converted into various polyolefin containing block copolymers by three different controlled polymerizations. The absence of homopolymer impurities showcases the high conversion of each step of the transformation. The range of non-polar and polar block-copolymers that are accessible by coordinative polymerizations,^{40,41} anionic polymerizations,⁴²

and controlled radical polymerization⁴³ make this system a powerful platform to generate a wide range of polyolefin containing materials. The use of highly active catalysts in each step of the process results in an overall process with unprecedented productivity lending itself to potential industrial applications.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors thank Umicore for the generous gift of Grubbs catalyst and the Department of Chemical and Biomolecular Engineering at the University of Illinois at Urbana–Champaign for funding this research. Major funding for the 500 MHz Bruker CryoProbe was provided by the Roy J. Carver Charitable Trust (Muscatine, Iowa; Grant #15-4521) to the School of Chemical Sciences NMR Lab.

Notes and references

- 1 F. S. Bates, M. A. Hillmyer, T. P. Lodge, C. M. Bates, K. T. Delaney and G. H. Fredrickson, *Science*, 2012, **336**, 434–440.
- 2 F. S. Bates and G. H. Fredrickson, *Phys. Today*, 1999, **52**, 32–38.
- 3 A.-V. Ruzette and L. Leibler, *Nat. Mater.*, 2005, **4**, 19–31.
- 4 N. Hadjichristidis, S. Pispas and G. Floudas, *Block Copolymers: Synthetic Strategies, Physical Properties, and Applications*, John Wiley & Sons, Inc., Hoboken, USA, 2002.
- 5 M. Stürzel, S. Mihañ and R. Mühlaupt, *Chem. Rev.*, 2016, **116**, 1398–1433.
- 6 M. C. Baier, M. A. Zuideveld and S. Mecking, *Angew. Chem., Int. Ed. Engl.*, 2014, **53**, 9722–9744.
- 7 J. M. Eagan, J. Xu, R. Di Girolamo, C. M. Thurber, C. W. Macosko, A. M. LaPointe, F. S. Bates and G. W. Coates, *Science*, 2017, **355**, 814–816.
- 8 C. De Rosa and F. Auriemma, *Polym. Chem.*, 2011, **2**, 2155.
- 9 A. Hotta, E. Cochran, J. Ruokolainen, V. Khanna, G. H. Fredrickson, E. J. Kramer, Y.-W. Shin, F. Shimizu, A. E. Cherian, P. D. Hustad, J. M. Rose and G. W. Coates, *Proc. Natl. Acad. Sci.*, 2006, **103**, 15327–15332.
- 10 Y. Zhao, L. Wang, A. Xiao and H. Yu, *Prog. Polym. Sci.*, 2010, **35**, 1195–1216.
- 11 N. M. G. Franssen, J. N. H. Reek and B. de Bruin, *Chem. Soc. Rev.*, 2013, **42**, 5809–5832.
- 12 Y. Wang and M. A. Hillmyer, *J. Polym. Sci., Part A-1: Polym. Chem.*, 2001, **39**, 2755–2766.
- 13 L. M. Pitet and M. A. Hillmyer, *Macromolecules*, 2009, **42**, 3674–3680.
- 14 L. M. Pitet, M. A. Amendt and M. A. Hillmyer, *J. Am. Chem. Soc.*, 2010, **132**, 8230–8231.
- 15 K. Zhang, Z. Ye and R. Subramanian, *Macromolecules*, 2008, **41**, 640–649.



- 16 S. K. T. Pillai, W. P. Kretschmer, M. Trebbin, S. Förster and R. Kempe, *Chem.–Eur. J.*, 2012, **18**, 13974–13978.
- 17 R. Briquel, J. Mazzolini, T. Le Bris, O. Boyron, F. Boisson, F. Delolme, F. D'Agosto, C. Boisson and R. Spitz, *Angew. Chem., Int. Ed.*, 2008, **47**, 9311–9313.
- 18 I. German, W. Kelhifi, S. Norsic, C. Boisson and F. D'Agosto, *Angew. Chem., Int. Ed.*, 2013, **52**, 3438–3441.
- 19 A. L. Mogstad and R. M. Waymouth, *Macromolecules*, 1992, **25**, 2282–2284.
- 20 G. Cecchin, G. Morini and F. Piemontesi, in *Kirk-Othmer Encyclopedia of Chemical Technology*, John Wiley & Sons, Inc., Hoboken, NJ, USA, 2003, pp. 502–554.
- 21 M. Erberich, H. Keul and M. Möller, *Macromolecules*, 2007, **40**, 3070–3079.
- 22 Q. Ma and K. L. Wooley, *J. Polym. Sci., Part A-1: Polym. Chem.*, 2000, **38**, 4805–4820.
- 23 K. Matyjaszewski, J. Saget, J. Pyun, M. Schlögl and B. Rieger, *J. Macromol. Sci. Part A*, 2002, **39**, 901–913.
- 24 S. Norsic, C. Thomas, F. D'Agosto and C. Boisson, *Angew. Chem., Int. Ed.*, 2015, **54**, 4631–4635.
- 25 H. Kaneko, S. Kojoh, N. Kawahara, S. Matsuo, T. Matsugi and N. Kashiwa, *J. Polym. Sci., Part A-1: Polym. Chem.*, 2005, **43**, 5103–5118.
- 26 H. Kaneyoshi and K. Matyjaszewski, *J. Appl. Polym. Sci.*, 2007, **105**, 3–13.
- 27 Y. Inoue and K. Matyjaszewski, *J. Polym. Sci., Part A-1: Polym. Chem.*, 2004, **42**, 496–504.
- 28 C. Descour, T. Macko, I. Schreur-Piet, M. P. F. Pepels and R. Duchateau, *RSC Adv.*, 2015, **5**, 9658–9666.
- 29 D. Zhang, D. Guironnet, I. Göttker-Schnetmann and S. Mecking, *Organometallics*, 2009, **28**, 4072–4078.
- 30 D. Guironnet, T. Rünzi, I. Göttker-Schnetmann and S. Mecking, *Chem. Commun.*, 2008, 4965–4967.
- 31 A. E. Cherian, E. B. Lobkovsky and G. W. Coates, *Macromolecules*, 2005, **38**, 6259–6268.
- 32 L. K. Johnson, C. M. Killian and M. Brookhart, *J. Am. Chem. Soc.*, 1995, **117**, 6414–6415.
- 33 T. Wiedemann, G. Voit, A. Tchernook, P. Roesle, I. Göttker-Schnetmann and S. Mecking, *J. Am. Chem. Soc.*, 2014, **136**, 2078–2085.
- 34 S. B. Garber, J. S. Kingsbury, B. L. Gray and A. H. Hoveyda, *J. Am. Chem. Soc.*, 2000, **122**, 8168–8179.
- 35 A. K. Chatterjee, T.-L. Choi, D. P. Sanders and R. H. Grubbs, *J. Am. Chem. Soc.*, 2003, **125**, 11360–11370.
- 36 W. Shi, A. L. Hamilton, K. T. Delaney, G. H. Fredrickson, E. J. Kramer, C. Ntaras, A. Avgeropoulos and N. a. Lynd, *J. Am. Chem. Soc.*, 2015, **137**, 6160–6163.
- 37 T. P. Lodge, *Macromol. Chem. Phys.*, 2003, **204**, 265–273.
- 38 L. Resconi, F. Piemontesi, I. Camurati, O. Sudmeijer, I. E. Nifant'ev, P. V. Ivchenko and L. G. Kuz'mina, *J. Am. Chem. Soc.*, 1998, **120**, 2308–2321.
- 39 R. Mühlaupt, in *Ziegler Catalysts*, Springer Berlin Heidelberg, Berlin, Heidelberg, 1995, pp. 35–55.
- 40 S. Paul, Y. Zhu, C. Romain, R. Brooks, P. K. Saini and C. K. Williams, *Chem. Commun.*, 2015, **51**, 6459–6479.
- 41 A. Duda, ROP of Cyclic Esters. Mechanisms of Ionic and Coordination Processes, *Polymer Science: A Comprehensive Reference*, Elsevier, 2012, vol. 4.
- 42 S. Penczek and A. Duda, *Makromol. Chem., Macromol. Symp.*, 1993, **67**, 15–42.
- 43 W. A. Braunecker and K. Matyjaszewski, *Prog. Polym. Sci.*, 2007, **32**, 93–146.

