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Experimental and theoretical evidence of a supercritical-like transition in an organic semiconductor presenting colossal uniaxial negative thermal expansion†

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Thermal expansion coefficients of most materials are usually small, typically up to 50 parts per million per kelvin, and positive, *i.e.* materials expand when heated. Some materials show an atypical shrinking behavior in one or more crystallographic directions when heated. Here we show that a high mobility thiophene-based organic semiconductor, BHH-BTBT, has an exceptionally large negative expansion between 95 and 295 K ($-216 < \alpha_2 = \alpha_b < -333 \text{ MK}^{-1}$), being compensated by an even larger positive expansion in the perpendicular direction ($287 < \alpha_1 < 634 \text{ MK}^{-1}$). It is shown that these anomalous expansivities are completely absent in C8-BTBT, a much studied organic semiconductor with a closely related molecular formula and 3D crystallographic structure. Complete theoretical characterization of BHH-BTBT using *ab initio* molecular dynamics shows that below ~ 200 K two different α and β domains exist of which one is dominant but which dynamically exchange around and above 210 K. A supercritical-like transition from an α dominated phase to a β dominated phase is observed using DSC measurements, UV-VIS spectroscopy, and X-ray diffraction. The origin of the extreme negative and positive thermal expansion is related to steric hindrance between adjacent tilted thiophene units and strongly enhanced by attractive S...S and S...C interactions within the highly anharmonic mixed-domain phase. This material could trigger the tailoring of optoelectronic devices highly sensitive to strain and temperature.

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Introduction

Organic and inorganic compounds displaying negative thermal expansion (NTE) in one or more crystallographic directions have attracted great interest in recent years for fundamental reasons,^{1,2} as well as for their applications in functional materials.³ NTE is characterized by a negative sign of the thermal expansion coefficient $\alpha_l = (\partial l / \partial T)_P / l$, with l being one of the crystallographic axes

(a , b , or c) or one of the principal (orthogonal) axes of the strain tensor (1–3), where usually the temperature range is indicated over which the NTE occurs. Potential applications of NTE materials include their use as fillers in controlled thermal expansion composites,^{2,4} as thermomechanical actuator components^{5,6} or as precursor compounds for the design of zero thermal expansion materials by chemical modifications.³

It has become clear that porous materials, such as metal-organic frameworks (MOFs), are more likely candidates to display NTE than are dense materials, whether they are organic, inorganic or hybrid organic/inorganic. MOFs that display NTE include the cubic empty pore MOF-5 compound⁷ with $\alpha_a = -16 \text{ MK}^{-1}$, Ag [(ethylenediamine)]NO₃ with $\alpha_b = \alpha_2 = -89 \text{ MK}^{-1}$,⁸ MCF-82 ($\alpha_3 = -218 \text{ MK}^{-1}$),⁹ and [Zn₂(fu-L1)₂]_n ($\alpha_1 = -380 \text{ MK}^{-1}$).¹⁰

Dense materials displaying NTE have mostly smaller (less negative) linear thermal expansion coefficients, but occasionally these exceed those of porous materials. Examples include the inorganic phases ZrW₂O₈ ($\alpha_a = \alpha_1 = -9 \text{ MK}^{-1}$)¹¹ and Ca₂RuO_{3.74} ($\alpha_b = \alpha_2 = -115 \text{ MK}^{-1}$),¹² the metal-organic “wine-rack” phases Ag₃[Co(CN)₆]_n ($\alpha_c \approx -125 \text{ MK}^{-1}$),¹³ α -Cu-tricyanomethanide ($\alpha_b = \alpha_2 = -407 \text{ MK}^{-1}$),¹⁴ and thermosalient (phenylazophenyl)palladium hexafluoroacetylacetonate ($\alpha_c = -80 \text{ MK}^{-1}$),¹⁵ and the organic compounds nitromethane

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18-crown-6 ($\alpha_1 = -129 \text{ MK}^{-1}$),¹⁶ octa-3,5-diyne-2,7-diol ($\alpha_c = \alpha_3 = -204 \text{ MK}^{-1}$)¹⁷ and thermosalient *N'*-2-propylidene-4-hydroxybenzohydrazide (IMACET) ($\alpha_c = \alpha_3 = -290 \text{ MK}^{-1}$).¹⁸ The latter compounds are exceptional since NTE is assumed to be less common in purely organic compounds than in dense organometallic or inorganic compounds.^{17,19}

It is clear that the mechanisms underlying the phenomenon of NTE are likely to be different in dense and porous materials. We present here a case study of a dense organic benzothienobenzothiophene structure (bis(hydroxyhexyl)[1]benzothieno[3,2-*b*][1]benzothiophene (**BHH-BTBT**)) that displays in the absence of an obvious structural discontinuity an exceptional negative uniaxial thermal expansion along the crystallographic *b*-axis of the same order of magnitude as that in IMACET, *i.e.* $\alpha_b = \alpha_2 = -333 \text{ MK}^{-1}$ at 295 K, and an average (linear) value of -274 MK^{-1} between 95 and 295 K. The negative thermal expansion in IMACET appears to be the onset of a first-order phase transition, whereas in **BHH-BTBT** a more complex supercritical-like transition without obvious structural discontinuity takes place as will be shown below.

The uniaxial NTE in **BHH-BTBT** is more than compensated by a very high positive thermal expansion (PTE) along the *a*-axis of 580 MK^{-1} at 295 K (linear value between 95 K and 295 K: 441 MK^{-1}), so that the volume expansion displays a net positive effect. Thermal expansion of most compounds does not exceed 50 MK^{-1} . The negative expansion in **BHH-BTBT** can be labeled as 'colossal' according to the classification proposed and used by Goodwin *et al.*,¹³ van Heerden *et al.*,²⁰ and Henke *et al.*¹⁰ where 'giant' means $|\alpha| \geq 25 \text{ MK}^{-1}$, 'colossal' $|\alpha| \geq 100 \text{ MK}^{-1}$ and 'massive' $|\alpha| \geq 400 \text{ MK}^{-1}$. We also show that this exceptional thermal expansion is completely absent in a closely related compound, the well-known, much studied,^{21–23} and commercially available bis(octyl)[1]benzothieno[3,2-*b*][1]benzothiophene (**C8-BTBT**, Scheme 1), despite a very similar crystal packing.

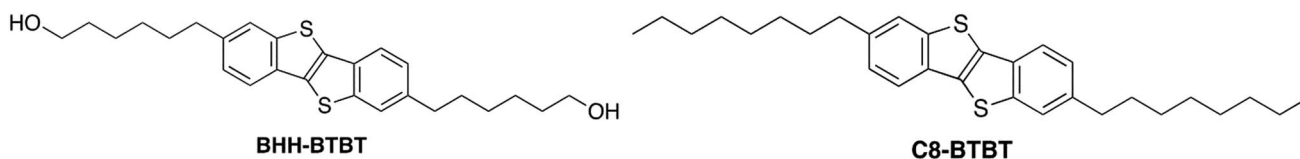
This type of π -conjugated thiophene based material is of particular interest in the field of organic electronics, used as a semiconductor in organic field-effect transistors (OFETs).^{24–26} In this context, some of us recently reported thermally evaporated **BHH-BTBT** based OFETs exhibiting a high mobility of up to $0.17 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ measured under an inert atmosphere but also in ambient air.²⁷ The crystal packing features, morphology and charge transport in organic semiconductors are intimately related,²⁸ so in order to obtain a deeper understanding of the exceptional NTE in **BHH-BTBT** we have carried out dynamic *ab initio* density functional theory (DFT) molecular calculations. Thermal expansion was shown to be the main factor responsible for thermal bandwidth narrowing in organic semiconductors

such as pentacene (with a slight negative thermal expansion) and rubrene.²⁹ The modeling of thermal expansion coefficients by means of *ab initio* calculations is challenging; in most cases, they are estimated by computing the quasiharmonic vibrational effects at different volumes to estimate the anharmonic contribution to the energy expansion.^{30–33} In the present case, because the material is dominated by weak forces, this approach is not expected to work, in particular for high temperatures. Therefore, *ab initio* dynamic calculations at constant temperature and pressure have been done for the first time to our knowledge for such a large and weakly bonded compound. They give insight into the structural and electronic changes as a function of temperature.

Structure description and thermal expansion

We reported previously the structure of **BHH-BTBT** at 295 K.²⁷ It crystallizes in space group *I2/a* with two short *a* and *b* axes of 7.8 and 5.9 Å, respectively, and one very long *c*-axis of 49.5 Å (95 K, Fig. S1†). The crystal packing is very characteristic of thiophene based OFET structures, either with **BTBT** cores or fused thiophene-based structures with, for example, dithieno-thiophene cores, with linear or branched side chains.^{34–38} The more or less stretched molecules stack according to a herringbone organization (Fig. S25†) as layers parallel to the *ab*-plane in such a way as to have close S...S and S...C contacts, responsible for the conducting behavior of the materials. In **BHH-BTBT** four such contacts below 4 Å are present (Fig. 1).

The shortest S...S contact at 3.430 Å is between adjacent parallel molecules with an angle between the two planes through the **BTBT** cores of strictly 0°, since they are shifted by one lattice translation along the *b*-axis. These molecules are said to be in an edge-to-edge (EE) conformation although they are in fact largely shifted along the *b*-axis. The other relatively short S...S contact (3.911 Å) is between adjacent molecules which are in a face-to-edge (FE) conformation (FE) with respect to the **BTBT** cores at an angle of 53° at 95 K. This herringbone angle σ_h is very close to that in similar **BTBT** systems where this angle varies between 50 and 54° (Table S3†). The FE conformation also gives rise to two very short atomic distances (3.309 and 3.384 Å) from the S atom of one molecule to the C3 and C4 carbon atoms of two different aromatic bonds (Fig. 1b). The S...C distances to the two other carbon atoms are slightly larger than the sum of the van der Waals radii of carbon. The distinguishing structural feature of this thiophene based semiconductor compared to other reported structures is the end hydroxyl functional group which stabilizes the crystal packing by a hydrogen-bonded network between **BHH-BTBT** layers (Fig. S3†).



Scheme 1 Chemical schemes for **BHH-BTBT** and **C8-BTBT**.



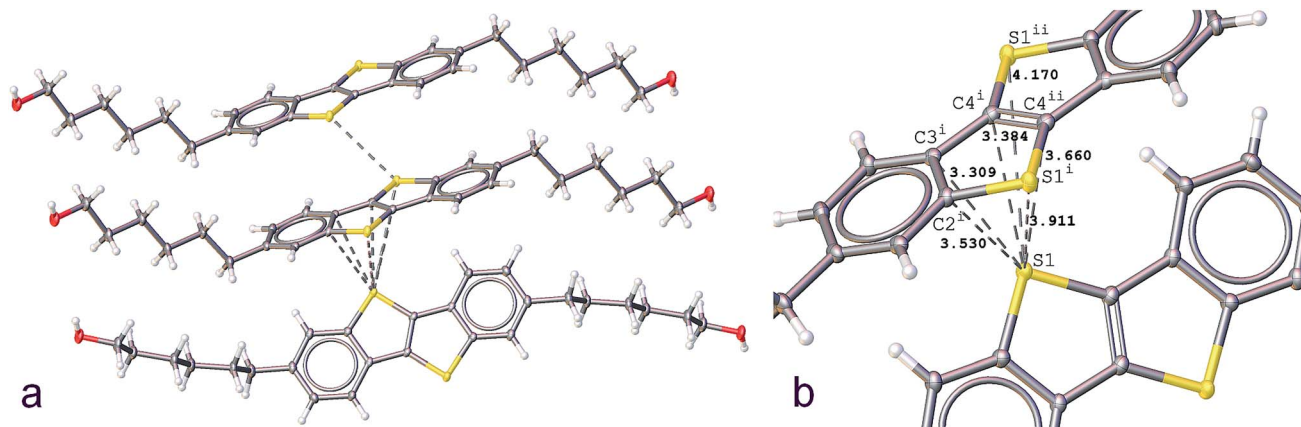


Fig. 1 (a) View of S...S and S...C contacts in the structure at 95 K with a parallel edge-to-edge conformation between the upper two molecules and a tilted face-to-edge conformation between the two bottom molecules. (b) Close contacts between an edge sulfur and the atoms of the thiophene ring of an adjacent tilted BTBT core. The shortest S1...S1ⁱⁱⁱ contact (3.430 Å) is visible in part (a) of this figure as the dashed line between the two upper molecules. Symmetry codes: (i) $1/2 + x, 1 - y, z$; (ii) $2 - x, 1/2 + y, 1.5 - z$; and (iii) $1.5 - x, 1.5 - y, 1.5 - z$.

The normalized thermal expansion values for the three crystallographic axes and the volume are shown in Fig. 2. Whereas the *c*-axis shows a very classical thermal expansion of less than 2% over the studied temperature range between 95 K and 295 K, the behavior of the *a* and *b* axes is very atypical with an exceptional NTE of the *b*-axis (-333 MK^{-1} , *i.e.* nearly 6%) at the highest temperatures and a more than compensating very strong PTE of the *a*-axis (580 MK^{-1} , *i.e.* nearly 9%). The value of the cell volume is mostly determined by the value of the normally increasing very large *c*-axis and the very large increase of the *a*-axis and augments by nearly 5%. The expansion is far from linear throughout the investigated temperature interval, but no discontinuities can be detected. Since the structure is monoclinic and the β -angle varies as well with temperature the expansivities along the *a* and *c* axes are not along identical directions at different temperatures. The difference is not large

because the β -angle is close to 90° , varying between 95.8° at the lowest temperature to 92.8° at the highest temperature. For a better comparison with other systems the expansivities along orthogonal axes are reported as well (Fig. S9[†]): $287 < \alpha_1 < 634 \text{ MK}^{-1}$; $-216 < \alpha_2 = \alpha_b < -333 \text{ MK}^{-1}$; $23 < \alpha_3 < 60 \text{ MK}^{-1}$. It is noted that there is no space group change throughout the investigated temperature interval.

The impact of the large positive and negative expansions of notably the *a* and *b* parameters can be seen by plotting several geometrical parameters calculated from the atomic models resulting from the structural refinements as a function of temperature (Fig. 3). The principal effect is a very large tilting of a BTBT core with respect to another in FE conformation. The tilting is accompanied by nearly constant S...C3ⁱ and S...C4ⁱ contact distances and a much larger increase of the S...C2ⁱⁱ and S...C4ⁱⁱ distances which were initially already longer than the sum of the carbon van der Waals radii. A closer inspection of the distance and angle variations seems to indicate a change of the slope in between approximately 200 and 220 K, suggesting a second-order phase transition. Other geometrical parameters behave normally as a function of temperature, in particular those related to intramolecular distances, angles, and torsional angles showing that the conformation of the molecules hardly changes as a function of temperature (Fig. S13 and S15[†]), which is also expressed in the nearly constant eigenvalues of the molecule's tensor of inertia (Fig. S12[†]).

The crystal packing of BHH-BTBT is by far not unique: a search in the Cambridge Structural Database³⁹ for structures with BTBT cores yielded 14 unique hits, including BHH-BTBT, and a search for structures with analogous bisthienothiophene (BTB) cores resulted in 16 hits (excluding co-crystals). Of these, depending on the side functional groups, 9 structures with BTBT cores have very similar packing structures to that of BHH-BTBT with alternating EE and FE orientations, and 7 structures with BTB cores (Table S3[†]). The crystal structures of all structures, except BHH-BTBT, have only been measured at one temperature, mostly 300 K, since this is the usual operating

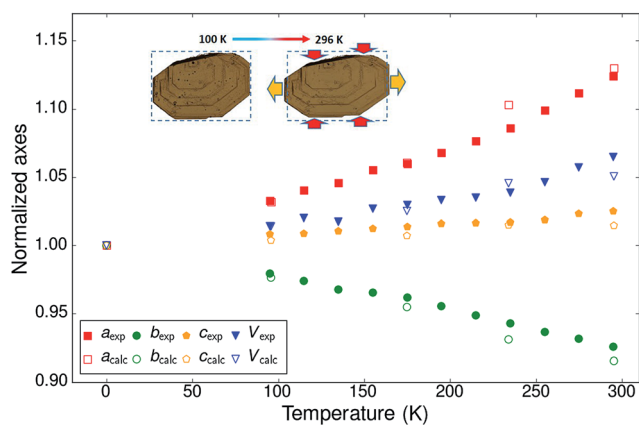


Fig. 2 Normalized experimental and calculated cell parameters and volume as a function of temperature for BHH-BTBT. The inset shows optical microscopy pictures of a crystal of BHH-BTBT at 100 K and 296 K. The crystallographic *a*-axis is to the right of the paper and the *c*-axis perpendicular to the paper. The dotted rectangles have exactly the same dimensions (see also Fig. S38[†]).



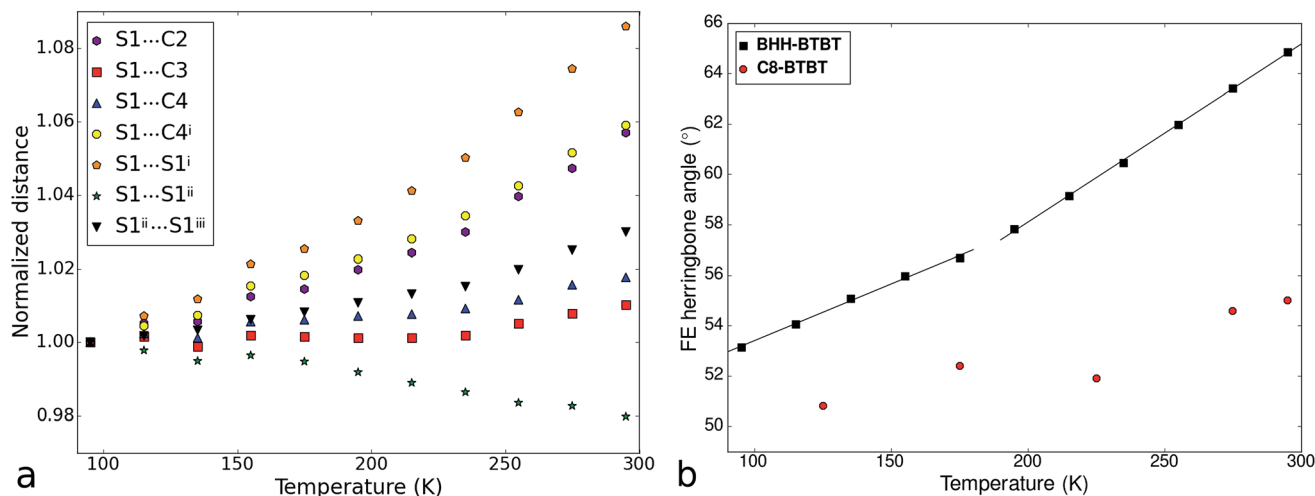


Fig. 3 (a) Temperature dependence of atomic distances between the sulfur atom of a BTBT core and the atoms of the thiophene ring of an adjacent BTBT core in the structure of BHH-BTBT. For the atom labelling, see Fig. 1b. (b) Temperature dependence of the herringbone angle σ_h between the least-squares planes of BTBT cores in the edge and face positions. Two straight lines were fitted to a low and a high temperature interval, respectively. σ_h for C8-BTBT is given as well.

temperature of the functional device. The distinguishing feature between BHH-BTBT and other structures with similar packings is the value of the tilt angle between the BTBT cores in face-edge orientation, which is 65° in BHH-BTBT at room temperature and $50\text{--}55^\circ$ for other BTBT structures, and $44\text{--}46^\circ$ for BTT core structures regardless whether the structure has been measured at low temperatures or at ambient temperatures. We determined the thermal expansion of C8-BTBT (Scheme 1) whose room temperature structure was determined previously.³⁷ Although the crystal packing is very similar to that in BHH-BTBT, the thermal expansion is completely different with a normal PTE along the in-plane (*i.e.* the plane perpendicular to the stacking direction of the C8-BTBT *b*-axis) and a close to zero thermal expansion for the other in-plane *a*-axis and along the stacking direction (Fig. S24[†]). The evolution of the herringbone angle is between 52.6° at 125 K and 55.0° at 295 K (Fig. 3b), which is quite normal and much smaller than that of BHH-BTBT. The presence of the hydrogen-bonded network between the stacked BHH-BTBT layers thus appears to be a determining factor for the outstanding NTE and PTE properties of BHH-BTBT.

Theoretical investigations

The *ab initio* DFT molecular calculations have been carried out in the *P1* space group. A symmetry analysis of the optimized DFT structures shows that the actual *I2/a* space group symmetry is very well retained (Table S6[†]). The cell parameters of the computed BHH-BTBT structure at 0 K (called α -0K) optimized from the 95 K experimental structure are in excellent agreement with the extrapolated cell parameters using the experimental negative and positive linear thermal expansion in the 95–200 K temperature range (Fig. 2). The evolution of the average cell parameters with temperature was determined from the constant particle, pressure and temperature Langevin dynamics (NPT⁴⁰). Below 200 K, the results are in excellent agreement with

the experimental observations. The time correlation of the *a* and *b* lattice parameters in Fig. 4 shows that the trajectory possesses only one unique attractor around the average values and also that there is a linear correlation between *a* and *b* keeping the *a* × *b* product approximately constant. For temperatures beyond 240 K a qualitative agreement for the positive and negative expansion remains but successive runs at the same temperature lead to a dispersion of about 10% on *a* and *b*, whereas *a* × *b*, *c* and *V* show a dispersion of only 1%, 0.5% and 0.6%, respectively. The system thus enters a new phase space domain beyond the pure harmonic vibrational limit associated with large correlated fluctuations of *a* and *b*. The *a/b* time correlation diagram for different runs (Fig. S28[†]) at 295 K shows 2 attractors for the trajectories characterized by the (*a*, *b*)-coordinates (9.1, 5.1), called α -295K, and (7.9, 5.9), called β -295K, with the intermediate region being traveled by the different runs. This suggests that above 200 K the system does not oscillate around only one equilibrium structure but accesses the phase space around two different local minima corresponding to two different phases, α and β . These two structures can be extracted by computing the average position of the atoms corresponding to the runs around the two different attractors (Fig. 4). One domain, α , corresponds to a short *a* parameter and has a structure very similar to the α -0K structure while the other domain, β , having a long *a* parameter and a different connectivity for, in particular, the intermolecular S...S contacts, is associated with the metastable β -0K phase.

The total structure averaged from all different runs is in very good agreement with the experimental structure not only for the unit cell parameters but also for the average local bonding. This suggests that the experimental structure at temperatures above 200 K is a dynamic mix of both α and β domains. A close analysis of the time-averaged α and β structures in the high-temperature regime shows rather important structural



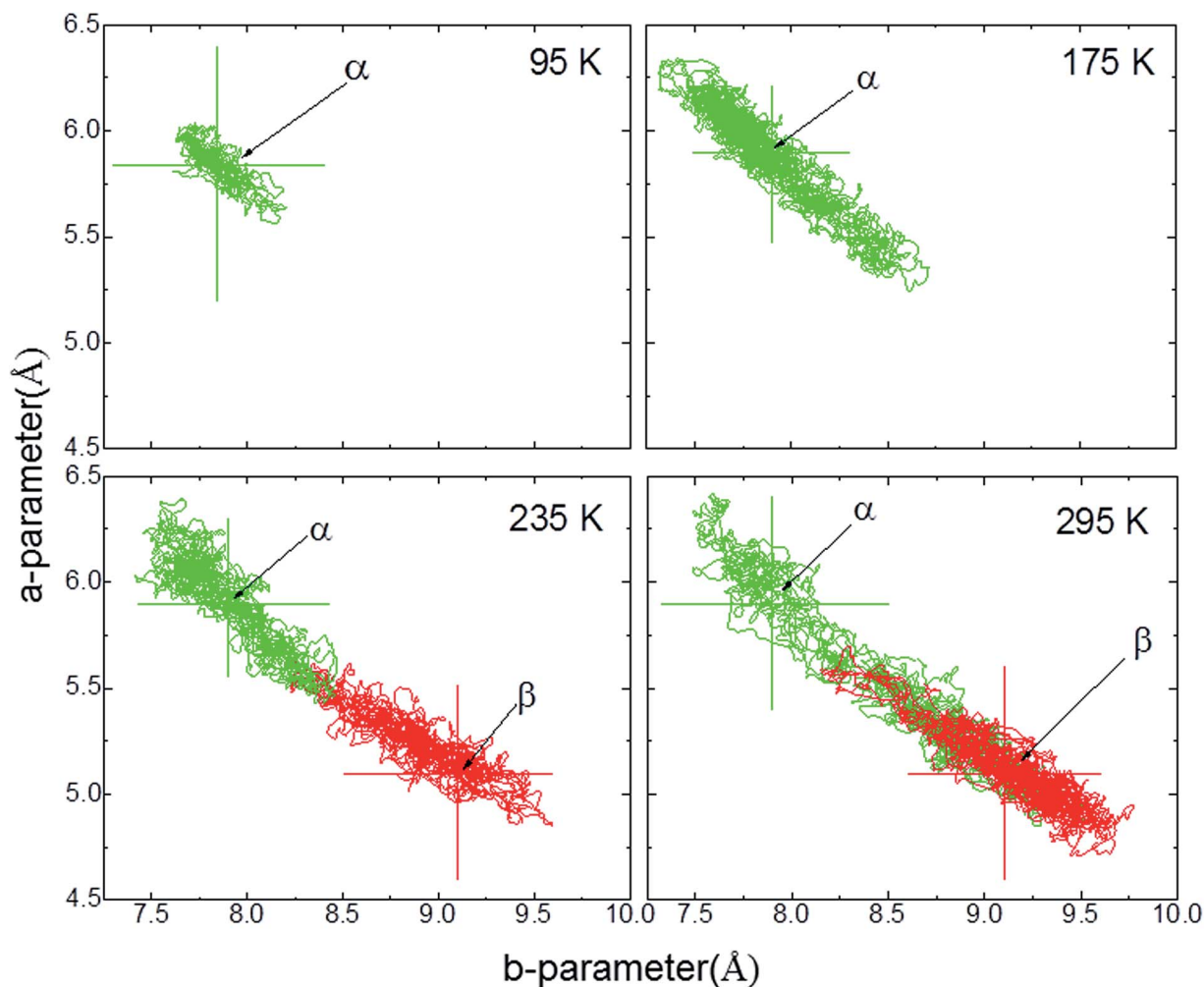


Fig. 4 *a*-parameter versus *b*-parameter for trajectories from 95 to 295 K. The large crosses give the equilibrium points for the α -0K (green) and β -0K (red) structures.

differences between them (Table 1 and Fig. S32†) for the axes and thus also for all derived geometrical parameters. The experimental values are always close to the mean of the values of

the α and β structures. It can be seen that the β structure is solely responsible for the deviating large herringbone angle σ_h in the experimental structure compared to related structures.

Table 1 Comparison of the averaged DFT structures at 235 K and 295 K in the two branches, α and β , and the experimental structure. Since the DFT optimizations have been done in the *P1* space group, α and γ angles deviate slightly from 90° . It also implies that there are two independent herringbone angles σ_h for each structure. The interatomic distances refer to the distances in the *I2/a* symmetrized structure. Symmetry codes are as in the caption of Fig. 1

	α -235K	β -235K	235K-exp	α -295K	β -295K	295K-exp
<i>a</i> (Å)	7.9191	8.8788	8.16147	8.1376	8.8560	8.4460
<i>b</i> (Å)	5.8957	5.2477	5.6917	5.7948	5.2953	5.5894
<i>c</i> (Å)	49.6875	50.3981	49.8698	49.6180	50.1623	50.2534
α (°)	90.0553	89.0994	90	89.9519	89.7405	90
β (°)	95.9899	91.3073	93.819	95.5113	93.2687	92.830
γ (°)	89.9097	89.5738	90	90.0081	89.9775	90
<i>V</i> (Å ³)	2307.2	2347.2	2311.5	2329.0	2348.5	2369.5
σ_h (°)	53.73/52.86	74.17/73.56	60.46	57.22/57.05	72.59/73.35	64.86
S1 ⁱⁱ ...S1 ⁱⁱⁱ (Å)	3.415	3.561	3.482	3.447	3.555	3.533
S1...S1 ⁱⁱ (Å)	4.195	3.784	4.113	4.146	3.892	4.086
S1...S ⁱ (Å)	3.992	4.429	4.107	4.099	4.451	4.247
S...C3 ⁱ (Å)	3.393	3.392	3.315	3.397	3.370	3.343
S...C4 ⁱ (Å)	3.420	3.345	3.415	3.429	3.652	3.444



The different connectivities between the **BTBT** cores in α and β also give rise to very different electronic structures. At 0 K (for the stable α structure), the percolation pathway of electrons and holes can be visualized in real space using the Fukui functions (Fig. S28†). The conductivity can directly be linked with very diffuse $3p_z$ sulphur orbitals interacting with each other and with the $C4^I/C4^{II}$ of the π -system of aromatic carbons for the valence band and with the π -system of the $C4^I$ aromatic carbons for the conduction band giving rise to electron-hole percolation pathways mostly in the ab -plane.

Fig. S29† shows the time-averaged temperature-dependent vertical Kohn–Sham band gap for three irreducible k -points extracted from NPT dynamics. An average decrease of the band gap of -0.8 meV K^{-1} for all k -points between 0 and 195 K is observed. Above 300 K, as the kinetic behavior of the system can be decomposed into the two domains α and β , existing locally for a time scale of larger than 10 ps, the optical band gap is dominated by the phase with the lower band gap, in this case the β domain. For this specific domain the band gap minimum is not anymore for G1 but for G2/G3 and the band gap decrease becomes -1.1 meV K^{-1} . This is coherent with the experimental UV-VIS spectroscopy results (Fig. 5) showing a decrease of -0.39 meV K^{-1} below 200 K and -0.53 meV K^{-1} above 200 K. The band gap in **C8-BTBT** is 0.02 meV K^{-1} , nearly temperature independent.

The theoretical over-estimation of the gap variation is most probably linked to the underestimation of the Kohn–Sham gaps compared to the experimental ones. The change in behavior for the k -points can be qualitatively understood by looking at the full band structure from the time-averaged structures at different temperatures shown in Fig. S31.† All α -domains present a similar band structure with a band structure minimum around the Γ -point.

For this structure, the four highest occupied bands and the four lowest unoccupied bands present a decrease of the band dispersion caused by the dynamic lowering of the local symmetry due to the splitting of the band degeneracy, which causes the decrease of the band gap tending to an asymptotic lower limit around 200 K. Above 200 K the β -domain starts to

contribute to the system dynamics, having a different band gap because of the different connectivity of the intermolecular $S\cdots S$ bonds (in particular $S\cdots S^I$ and $S\cdots S^{II}$). Now the total band gap decrease is determined by that of the β -domain because the temperature increase favors the decrease of the $S\cdots S^{II}$ contact distance and the electronic delocalization inducing an increase of the band dispersion.

Discussion

In order to explain the anomalous behavior of the thermal expansion in **BHH-BTBT** we could look at analogous compounds, but temperature-dependent structural data on organic semiconductors are very scarce. Only one comparative structural study was published on the acenes naphthalene, anthracene, tetracene and pentacene,^{41,42} which all show the typical herringbone packing also observed in **BHH-BTBT**. Pentacene presents negative uniaxial thermal expansion an order of magnitude smaller than in **BHH-BTBT** and has a much smaller temperature dependence of the herringbone angle. The main reason for the negative thermal expansion in pentacene was shown to be the librational motion along the long molecular axis in conjunction with an increase of σ_h , and frustrated by steric hindrance between the herringbone-packed molecules. This is expected to play a role as well in **BHH-BTBT** since librational motion in **BHH-BTBT** is twice as large (Fig. S12†) as in pentacene, but other factors enhance NTE and PTE significantly.

The full dynamic characterization by DFT has shown that **BHH-BTBT** presents two polymorphs α and β characterized by different (a,b) -parameters and σ_h angles. The phases are closely related, the main difference being in the connectivity between adjacent **BTBT** moieties leading to quite different electronic structures (Fig. S31†). The most stable phase at 0 K is α but β is only 85 meV per molecule higher in energy, including zero-point vibrational energies. With increasing temperature, the **BTBT** moieties librate more and more around their long axis, asymmetrically towards the β structure as it allows some electron delocalization through the $S\cdots S^{II}$ contact because of the associated length reduction. This leads to a flattened anharmonic potential in the direction of the β structure (Fig. 6).

This strongly anharmonic frustrated rotation couples with the a and b parameters leading to a colossal positive and negative thermal expansion as schematized in Fig. 7. This is also seen in the computed (a,b) trajectory at 95 K that tends to be displaced toward larger a values (Fig. 4). For higher temperature, the fluctuation in the a -parameter becomes more and more important until it becomes as large as the separation between α and β at around 200 K. Therefore α and β cannot be considered as independent phases but as an average phase with average properties between the α and β domains.

In contrast to a “classical” first order phase transition (Fig. 6) where structural fluctuations are smaller than the separation between the two different structures, the structural fluctuations in the present case are of the order of the structural separation between α/β domains. This is reminiscent of a supercritical transition from a liquid to a gas with no discontinuity of the

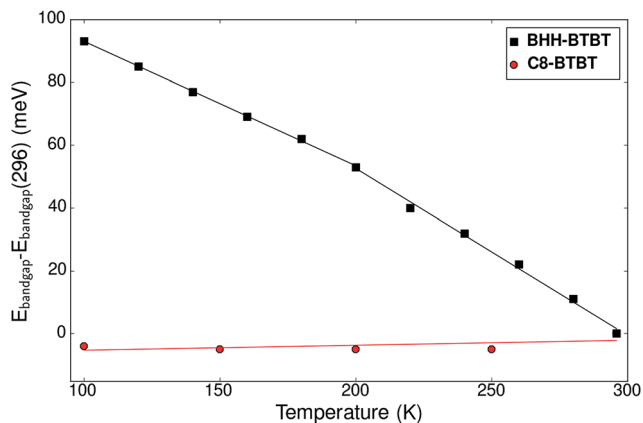


Fig. 5 Experimental evolution of the optical band gap of **BHH-BTBT** and **C8-BTBT** as a function of temperature as determined from UV-vis spectroscopy.



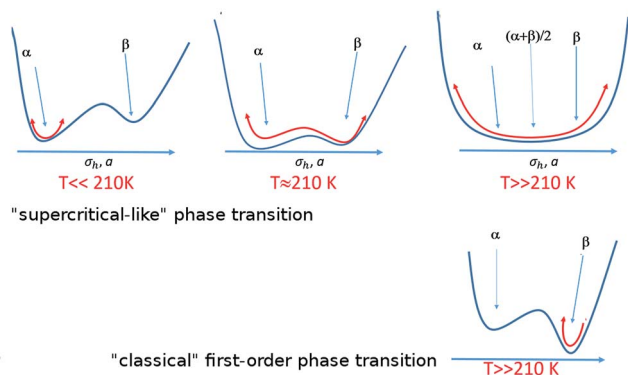


Fig. 6 Schematic view of the evolution of the free energy landscape of the α and β domains in BHH-BTBT in the case of a supercritical-like phase transition (top) and of a conventional first-order phase transition (bottom).

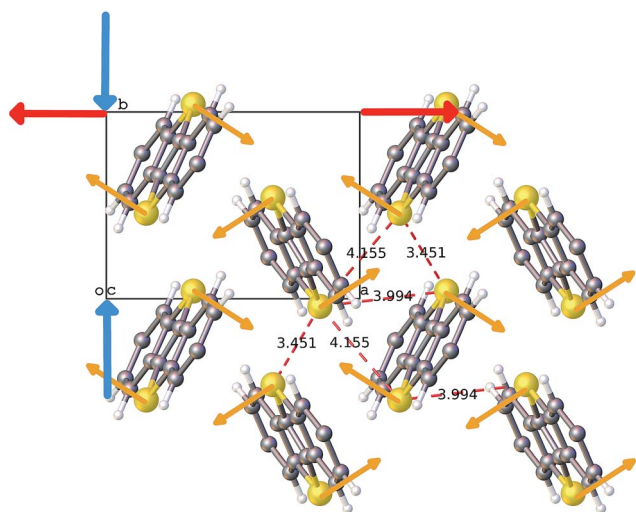


Fig. 7 Scheme for the thermal expansion of BHH-BTBT. Yellow arrows represent the rotation of the moieties with the increase of temperature; the blue and red ones give the direction of the negative and positive unit cell expansion, respectively. Only the aromatic moieties are shown (for the 155 K experimental structure). The numbers refer to close S...S distances in Å.

density. DSC measurements (Fig. S36[†]) show a discontinuity at 210 K with a very weak enthalpy $\Delta H = 170 \text{ J mol}^{-1}$ which is at least an order of magnitude smaller than the energy required to break a classical hydrogen bond (6–30 kJ mol^{-1}). Although the small hysteresis in the DSC plot could indicate a classical one to one first-order phase transition, this is at first sight not very likely in view of the extremely weak energy involved in the transition and the absence of clear discontinuities in the evolution of the cell parameters. From the dynamic DFT calculations we rather deduce that in analogy with supercritical fluids this could correspond to the equivalent of a Widom line⁴³ which is observed above the critical point separating a pseudo-liquid from pseudo gas dominated domains associated with a maximum of the isobaric heat capacity and a pseudo-boiling temperature.⁴⁴ This observation agrees with the classification by

Christy⁴⁵ of an isosymmetric transition, *i.e.* a transition in which the high-temperature phase has the same space group symmetry as the low-temperature phase, as a crossover or diffuse transition without discontinuity in the free-energy derivative leading to supercritical behaviour. Isosymmetric structural transitions are necessarily first-order according to Landau theory and are thus discontinuous. There is however not a minimum size on the discontinuity under the Landau conditions, and so a weakly discontinuous isosymmetric transition may appear continuous.⁴⁶

Below 210 K, with increasing temperature the fluctuations become extremely large and the dominant α phase contains more and more domains that dynamically convert into β -domains even if they are strictly less stable. As seen in the *ab initio* dynamics α/β -domains are locally stable on the typical time-scale of the dynamics (*i.e.* 20 ps) even if the β -domains remain less probable than the α -domains. This creates a mix of α/β sub-domains stabilized by entropy that interconverts with time: on the macroscopic time and space scale, the compound appears to be formed of a unique phase that would correspond to the time and space averaged mix of α/β with possibly a few domains that are pure α structures.

Around 210 K, the free energy of the α and β structures has becomes equal, the relative α/β populations become slowly constant and only a small fraction of the α structures converts into β structures. From the DSC enthalpy value and the computed energy difference between α and β structures it can be estimated that only 2% of the α structures dynamically converts to β structures. As the conversion from α to β structures implies a nucleation process, this could explain the hysteresis in the DSC curve. At higher temperatures, the fluctuations in the parameters become even much larger than the separation between the sub-domains that become more and more difficult to separate: α/β -domains melt into a unique dynamic phase. The global phase appears to be homogeneous with properties that are the average of the averaged lattice and bond length properties that compare very well with the experimental data. The increase of the thermal expansion coefficient along a continues to increase even more (Fig. S8[†]) since the system explores a larger phase space in the α/β -domain “melt”, inducing a more important anharmonic contribution to the thermal expansion.

The transition between the low and high temperature systems can be characterized by computing $\langle E_{\text{tot}} \rangle - \langle 2E_{\text{kin}} \rangle$ with $\langle E_{\text{tot}} \rangle$ the average total energy and $\langle E_{\text{kin}} \rangle$ the kinetic energy. The virial theorem requires that $\langle E_{\text{tot}} \rangle - \langle 2E_{\text{kin}} \rangle = 0$ for perfect harmonic oscillators. Fig. S30[†] shows that for temperatures below 200 K the computed dynamics have a moderate anharmonicity behavior while above 200 K there is a global increase of the anharmonic behavior of the α -domains associated with an important deviation from harmonicity for the β -domains.

Conclusions

In summary we have shown that the organic semiconductor BHH-BTBT, belonging to a wider class of structurally similar compounds based on fused thiophene units, displays extremely



large negative and positive thermal expansion over the investigated temperature range of 200 K and that it displays anomalies around 210 K in derived geometrical parameters as well as in UV-VIS spectroscopy data. The unique feature of this study combining state-of-the-art experiments and theoretical calculations is that we were able to show that there exists a low-temperature state consisting of one domain and a high-temperature phase consisting of two dynamic concomitant domains permanently exchanging. The transition between the two states is not a classical discontinuous phase-to-phase transition, but presents a quasi-continuity of all geometric parameters very similar to the ones observed in supercritical transitions. Nevertheless it is responsible for the anomalies observed in geometrical parameters derived from X-ray diffraction experiments, UV-VIS spectroscopy and DSC measurements. It is expected that this solid-state supercritical transition is not an isolated case limited to an organic semiconductor, but that isosymmetrical transitions in other compounds will be recognized as such as well, as was done, for example, by Swainson *et al.* for the transition in ammonium hexafluorophosphate.⁴⁶

The dynamic co-existence of the two domains is in itself not the only factor that causes the extreme expansions but it enhances it considerably by amplifying the anharmonicity; steric hindrance between adjacent thiophene units is the primary structural reason for the NTE in this compound. Combining the excellent electrical properties with the atypical thermo-mechanical properties could lead to multipurpose temperature dependent electronic devices. It is not uncommon that a NTE material exhibits coexisting negative properties, such as a negative Poisson ratio or negative compressibility,⁴⁷ but this has not yet been investigated for **BHH-BTBT**.

The question why **BHH-BTBT** has exceptionally atypical thermal expansion, and the related compound **C8-BTBT** not, is difficult to answer. The presence of the hydrogen-bonded network between the alternating layers could be a decisive factor in the sense that it appears to switch the orientation of two adjacent **BTBT** cores in FE stacking by 180° (Fig. S27†) compared to all other **BHH-BTBT** structures in the CSD, but in the absence of any structural data of other **BHH-BTBT** compounds with hydrogen bonded networks, this is not more than a hypothesis. From DFT calculations it results that the relative orientation of two adjacent FE **C8-BTBT** cores is about 6 meV per core more stable than that in **BHH-BTBT**. With such a small energy difference other structural factors or kinetic effects during the crystallization could be more decisive for the orientation that emerges.

Conflicts of interest

The authors declare no conflict of interest.

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References

- W. Miller, C. W. Smith and D. S. Mackenzie, *J. Mater. Sci.*, 2009, **44**, 5441–5451.
- C. Lund, *Materials*, 2012, **5**, 1125–1154.
- J. Chen, L. Hu, J. Deng and X. Xing, *Chem. Soc. Rev.*, 2015, **44**, 3522–3567.
- K. D. Tran, T. J. Groshens and J. G. Nelson, *Mater. Sci. Eng., A*, 2001, **303**, 234–240.
- O. Sigmund and S. Torquato, *J. Mech. Phys. Solids*, 1997, **45**, 1037–1067.
- H. Janocha, *Springer Science & Business Media*, 2007.
- W. Zhou, H. Wu, T. Yildirim, J. R. Simpson and A. R. Hight Walker, *Phys. Rev. B*, 2008, **78**, f054114.
- A. Katrusiak and W. Cai, *Nat. Commun.*, 2014, **5**, f4337.
- H.-L. Zhou, Y.-B. Zhang, J.-P. Zhang and X.-M. Chen, *Nat. Commun.*, 2015, **6**, f6917.
- S. Henke, A. Schneemann and R. A. Fischer, *Adv. Funct. Mater.*, 2013, **23**, 5990–5996.
- T. A. Mary, J. S. O. Evans, T. Vogt and A. W. Sleight, *Science*, 1996, **272**, 90–92.
- K. Takenaka, Y. Okamoto, T. Shinoda, N. Katayama and Y. Sakai, *Nat. Commun.*, 2017, **8**, f14102.
- A. L. Goodwin, M. Calleja, M. J. Conterio, M. T. Dove, J. S. O. Evans, D. A. Keen, L. Peters and M. G. Tucker, *Science*, 2008, **319**, 794–797.
- S. J. Hunt, M. J. Cliffe, J. A. Hill, A. B. Cairns, N. P. Funnell and A. L. Goodwin, *CrystEngComm*, 2015, **17**, 361–369.
- M. K. Panda, T. Runčevsk, S. C. Sahoo, A. A. Belik, N. K. Nath, R. E. Dinnebier and P. Naumov, *Nat. Commun.*, 2014, **5**, f4811.
- E. R. Engel, V. J. Smith, C. X. Bezuidenhout and L. J. Barbour, *Chem. Commun.*, 2014, **50**, 4238–4241.
- D. Das, T. Jacobs and L. J. Barbour, *Nat. Mater.*, 2010, **9**, 36–39.
- M. K. Panda, R. Centore, M. Causà, A. Tuzi, F. Borbone and P. Naumov, *Sci. Rep.*, 2016, **6**, f29610.
- G. H. Greyling, “*Negative thermal expansion of organic materials*”, Stellenbosch University, MSc thesis, 2011.
- D. P. van Heerden, C. Esterhuysen and L. J. Barbour, *Dalton Trans.*, 2016, **45**, 4141–4149.
- Y. Yuan, G. Giri, A. L. Ayzner, A. P. Zoombelt, S. C. B. Mannsfeld, J. Chen, D. Nordlund, M. F. Toney, J. Huang and Z. Bao, *Nat. Commun.*, 2014, **5**, f3005.
- R. Janneck, N. Pilet, S. P. Bommanaboyena, B. Watts, P. Heremans, J. Genoe and C. Rolin, *Adv. Mater.*, 2017, **29**, f1703864.
- L. Lyu, D. Niu, H. Xie, Y. Zhao, N. Cao, H. Zhang, Y. Zhang, P. Liu and Y. Gao, *Phys. Chem. Chem. Phys.*, 2017, **19**, 1669–1676.
- A. N. Sokolov, S. Atahan-Evrenk, R. Mondal, H. B. Akkerman, R. S. Sánchez-Carrera, S. Granados-Focil, J. Schrier,



- S. C. B. Mannsfeld, A. P. Zoombelt, Z. Bao and A. Aspuru-Guzik, *Nat. Commun.*, 2011, **2**, f437.
- 25 K. Nakayama, Y. Hirose, J. Soeda, M. Yoshizumi, T. Uemura, M. Uno, W. Li, M. J. Kang, M. Yamagishi, Y. Okada, E. Miyazaki, Y. Nakazawa, A. Nakao, K. Takimiya and J. Takeya, *Adv. Biomater.*, 2011, **23**, 1626–1629.
- 26 Y. Kiyota, T. Kadoya, K. Yamamoto, K. Iijima, T. Higashino, T. Kawamoto, K. Takimiya and T. Mori, *J. Am. Chem. Soc.*, 2016, **138**, 3920–3925.
- 27 G. H. Roche, Y.-T. Tsai, S. Clevers, D. Thuau, F. Castet, Y. H. Geerts, J. J. E. Moreau, G. Wantz and O. J. Dautel, *J. Mater. Chem. C*, 2016, **4**, 6742–6749.
- 28 C. Wang, H. Dong, L. Jiang and W. Hu, *Chem. Soc. Rev.*, 2018, **47**, 422–500.
- 29 Y. Li, V. Coropceanu and J.-L. Brédas, *J. Phys. Chem. Lett.*, 2012, **3**, 3325–3329.
- 30 N. L. Allan, G. D. Barrera, T. H. K. Barron and M. B. Taylor, *Int. J. Thermophys.*, 2001, **22**, 535–546.
- 31 Y. Wang, Z. K. Liu and L. Q. Chen, *Acta Mater.*, 2004, **52**, 2665–2671.
- 32 H. Zhang, S. L. Shang, Y. Wang, A. Saengdeejing, L. Q. Chen and Z. K. Liu, *Acta Mater.*, 2010, **58**, 4012–4018.
- 33 Y. N. Heit and G. J. O. Beran, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2016, **72**, 514–529.
- 34 M.-C. Chen, S. Vegiraju, C.-M. Huang, P.-Y. Huang, K. Prabakaran, S. L. Yau, W.-C. Chen, W.-T. Peng, I. Chao, C. Kim and Y.-T. Tao, *J. Mater. Chem. C*, 2014, **2**, 8892–8902.
- 35 Y. Liu, C.-A. Di, C. Du, Y. Lu, K. Liu, W. Qiu and G. Yu, *Chem.–Eur. J.*, 2010, **16**, 2231–2239.
- 36 Y. Sun, Y. Ma, Y. Liu, Y. Lin, Z. Wang, Y. Wang, C. Di, K. Xiao, X. Chen, W. Qiu, B. Zhang, G. Yu, W. Hu and D. Zhu, *Adv. Funct. Mater.*, 2006, **16**, 426–432.
- 37 T. Izawa, E. Miyazaki and K. Takimiya, *Adv. Mater.*, 2008, **20**, 3388–3392.
- 38 H. Mieno, T. Yasuda, Y. S. Yang and C. Adach, *Chem. Lett.*, 2014, **43**, 293–295.
- 39 C. R. Groom, I. J. Bruno, M. P. Lightfoot and S. C. Ward, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2016, **72**, 171–179.
- 40 D. Quigley and M. I. J. Probert, *J. Chem. Phys.*, 2004, **120**, 11432–11441.
- 41 S. Haas, B. Batlogg, C. Besnard, M. Schiltz, C. Kloc and T. Siegrist, *Phys. Rev. B*, 2007, **76**, f205203.
- 42 T. Siegrist, C. Besnard, S. Haas, M. Schiltz, P. Pattison, D. Chernyshov, B. Batlogg and C. Kloc, *Adv. Mater.*, 2007, **19**, 2079–2082.
- 43 G. G. Simeoni, T. Bryk, F. A. Gorelli, M. Krisch, G. Ruocco, M. Santoro and T. Scopigno, *Nat. Phys.*, 2010, **6**, 503–507.
- 44 D. K. Banuti, *J. Supercrit. Fluids*, 2015, **98**, 12–16.
- 45 A. G. Christy, *Acta Crystallogr., Sect. B: Struct. Sci.*, 1995, **51**, 753–757.
- 46 I. P. Swainson, R. P. Hammond, J. K. Cockcroft and R. D. Weir, *Phys. Rev. B*, 2002, **66**, f174109.
- 47 C. Huang and L. Chen, *Adv. Mater.*, 2016, **28**, 8079–8096.

