A combined experimental and computational study on the reaction of fluoroarenes with Mg–Mg, Mg–Zn, Mg–Al and Al–Zn bonds†

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Through a combined experimental and computational (DFT) approach, the reaction mechanism of the addition of fluoroarenes to Mg–Mg bonds has been determined as a concerted S$_2$Ar-like pathway in which one Mg centre acts as a nucleophile and the other an electrophile. The experimentally determined Gibbs activation energy for the addition of CsF$_6$ to a Mg–Mg bond of a molecular complex, $\Delta G^\ddagger_{328 \text{K(experiment)}} = 21.3$ kcal mol$^{-1}$ is modelled by DFT with the ωB97X functional, $\Delta G^\ddagger_{328 \text{K(DFT)}} = 25.7$ kcal mol$^{-1}$. The transition state for C–F activation involves a polarisation of the Mg–Mg bond and significant negative charge localisation on the fluoroarene moiety. This transition state is augmented by stabilising closed-shell Mg···F$_{\text{ortho}}$ interactions that, in combination with the known trends in C–F and C–M bond strengths in fluoroarenes, provide an explanation for the experimentally determined preference for C–F bond activation to occur at sites flanked by ortho-fluorine atoms. The effect of modification of both the ligand coordination sphere and the nature and polarity of the M–M bond (M = Mg, Zn, Al) on C–F activation has been investigated. A series of highly novel β-diketiminate stabilised complexes containing Zn–Mg, Zn–Zn–Zn, Zn–Al and Mg–Al bonds has been prepared, including the first crystallographic characterisation of a Mg–Al bond. Reactions of these new M–M containing complexes with perfluoroarenes were conducted and modelled by DFT. C–F bond activation is dictated by the steric accessibility, and not the polarity, of the M–M bond. The more open coordination complexes lead to enhanced Mg···F$_{\text{ortho}}$ interactions which in turn lower the energy of the transition states for C–F bond activation.

Introduction

The activation of strong carbon–fluorine bonds remains a key academic and industrial challenge that is motivated by the prevalence of fluorine-containing molecules in the pharmaceutical, agricultural and materials industries. Recent advances have focused on the functionalisation of fluoroarenes, with aim of making synthetically useful building blocks. In particular, C–F activation and subsequent borylation has been reported to occur catalytically using a range of late transition metal catalysts. The activation of aromatic sp$^2$ C–F bonds with transition metal complexes is known to proceed by several mechanisms, most common of which are oxidative addition (OA) and nucleophilic aromatic substitution (S$_n$Ar). Oxidative addition is typically observed with 14- or 16-electron complexes of the form ML$_2$ (M = Ni, Pd, Pt) or ML$_3$X (M = Rh, Ir), where higher oxidation states are easily accessible.

The concept of ligand-assisted C–F bond activation was introduced to rationalise experimentally observed fluorophosphines, formed during the reaction of perfluoroarenes with group 9 and 10 transition metal phosphine complexes. Originally reported by Milstein in the early 1990s, the reaction of CsF$_6$ with [([Et$_3$P)$_2$IrMe] forms [([Et$_3$P)$_2$FeT$_2$P]IrCsF$_6$] along with ethene and methane. Density functional calculations were used to substantiate a pathway that occurs by 1,2-addition of the C–F to the Ir–P bond: nucleophilic attack of the electron-rich metal on the fluorocarbon occurs with simultaneous activation, and trapping, of the fluoride by the phosphine. The concept has been expanded to boryl-assisted C–F activation using a [([Et$_3$P)$_2$RhBpin] complex and the mechanism used to explain the experimentally observed activation of pentafluoropyridine at the 2-position (Fig. 1 – ligand assisted OA). Lewis acid-assisted C–F bond activation, has been explored in a detailed experimental and theoretical study by Nakamura and co-workers. This mechanism is conceptually related to the ligand-assisted pathway but differs in that the Lewis acidic and nucleophilic metal centers are not directly bound to one.

† Electronic supplementary information (ESI) available: Experimental procedures, details of the DFT studies, single crystal X-ray data and multinuclear NMR spectra (DOI: 10.1039/c7sc05059c). X-ray crystallographic data for 2a, 2a-hexane, 2b, 3, 3-ThF, 3a, 4, 6 and 7 ( CIF), CCDC 1575549–1575558, 1577209. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7sc05059e

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Figs. 1 and 2. Selected transition states for C–F bond activation of fluoroarenes.

Fig. 1. Selected transition states for C–F bond activation of fluoroarenes.

Fig. 2. C–F bond activation of fluoroarenes using Mg–Mg reagents.

pathway, with no radical intermediates detected through trapping experiments. A plausible transition state for this reaction involves the 1,2-addition of the C–F bond across the apolar Mg–Mg bond.

Herein, we report a joint experimental and computational analysis of the carbon–fluorine bond activation of fluorinated arenes using low-valent complexes which contain M–M bonds (M = Mg, Zn, Al). We provide strong evidence for the proposed concerted pathway and a transition state reminiscent of those represented in Fig. 1 that involves dual nucleophilic and electrophilic sites that polarise and activate the C–F bond. We introduce additional experiments and prepare new M–M complexes in order to interrogate the key transition state for C–F activation. Specifically, the effects of steric accessibility and M–M bond polarity on reaction rates are reported.

Results

Concerted addition of C₆F₆ to 1

The reaction mechanisms reported herein were studied by DFT using the Gaussian 09 suite [revision D.01] and optimisations employed the ωB97X functional. Mg and Zn centers were described with Stuttgart SDDAll RECPs and associated basis sets whereas 6-31G* basis sets was used for all other atoms. Free energies are corrected for both benzene solvent using the Polarisable Continuum Model (PCM) approach and dispersion effects using ωB97X-D.

The potential energy surface for the reaction of C₆F₆ with 1 is presented in Fig. 3a. C–F bond activation is initiated upon weak coordination of C₆F₆ to the ligand sphere of 1 to create the encounter complex, Int-A. Substrate association to 1 occurs by a network of non-covalent interactions (e.g. C–H···F, C–H···π, π···π) which although individually weak contribute to a modest binding enthalpy of ΔH° = −9.2 kcal mol⁻¹ (Fig. S27†).

The increased ordering of the system is, however entropically unfavourable and thus fluoroarene binding is slightly endergonic (ΔG° = 1.1 kcal mol⁻¹). The ligands in Int-A are oriented perpendicular to one another and rotation around the Mg–Mg bond is required to access the more reactive conformer Int-A'. Int-A' contains a reactive pocket and a sterically accessible Mg–Mg bond. C–F bond activation occurs via the concerted transition state TS-1 (ΔG°₂₉₈K = 25.7 kcal mol⁻¹). Decay from TS-1 leads to the C–F activation product Int-B in which a newly formed Mg–C bond and bridging Mg–F–Mg bond contribute to a 6-membered ring structure. Overall C–F bond activation is extremely exergonic (ΔG°₂₉₈K = −112.2 kcal mol⁻¹).

An Eyring analysis on the reaction of 1 with C₆F₆ under pseudo first-order conditions, conducted over the temperature range 258–288 K, yielded activation parameters ΔH° = 10.8 kcal mol⁻¹ and ΔS° = −35 cal K⁻¹ mol⁻¹ (Fig. 3b and c). The experimental negative activation entropy is consistent with a highly ordered transition state, while the activation enthalpy is suggestive of a transition state with only minimal bond making and breaking having occurred. The Gibbs activation energy ΔG°₂₉₈K = 21.3 kcal mol⁻¹ matches reasonably well with that determined from DFT, ΔG°₂₉₈K = 25.7 kcal mol⁻¹.
The transition state for C–F bond activation

The calculated TS-1 is suggestive of an asynchronous concerted process: considerable Mg–C bond formation has occurred before any significant C–F bond breaking (Fig. 4). Formally, the transition state involves the concerted oxidative addition of the C–F bond to the Mg–Mg bond. Related pathways have been calculated for the addition of C–F bonds of fluoroarenes to single-site Al(I) and Si(II) centres.40–42

The bond lengths and angles around the ipso-carbon (Fig. S28a†), along with the charge distribution about the aromatic ring, however are consistent with an SNAr-like mechanism for C–F activation with one Mg center acting as a nucleophile, attacking the aromatic ring, while the other acts as an electrophile, binding to and polarising the breaking C–F bond.

An electron density difference map illustrates the depletion of electron density between the breaking C–F and Mg–Mg bonds along with the simultaneous accumulation of electron density between Mg and C at the forming Mg–C bond (Fig. 4a). Electron density accumulation also occurs at both ortho- and para-carbons of the fluoroarene. The electron density redistribution in TS-1 is mirrored by the NBO analysis showing the difference in the NPA charges between TS-1 and Int-A′ (Fig. 4b). The loss of electron density at the Mg centers (Δ = +0.52 e) is redistributed to the fluoroarene ring (Δ = −0.53 e) and approximately half of the redistributed electron density is localised at ipso-carbon (Δ = −0.27 e).

The effect of functional group substitution on the reaction rates of concerted SNAr pathways is well understood: stabilisation of the negative charge on aromatic ring through inductive and mesomeric effects leads to an increase in reaction rate. Rate constants were measured for the reaction of 1 with a series of fluoroarenes of the form C6F5R. The group para- to the breaking C–F bond was varied and rate constants were measured under pseudo first-order conditions (Table S1†). Broadly the relative rate increases with stabilisation of negative charge on the ipso-carbon, leading to the general reactivity trend: R = CF3 > F > C6F5 > CH3.

Importance of secondary Mg···Fortho interactions in the TS for C–F activation

In addition to the expected Mg···Fipso interaction derived from the breaking C–F bond, there is a significant Mg···Fortho interaction in TS-1 that is retained in Int-B. The presence of a strong Mg···Fortho interaction is evidenced by a short Mg···Fortho distance (2.34 Å). The QTAIM molecular graph of TS-1 (Fig. 4c) identifies the presence of a bond critical point (BCP) between

Fig. 3 (a) The calculated reaction pathway for C–F activation with 1. inset shows Mg···Mg separation. (b) Kinetics of the reaction of 1 with 10 equiv. of C6F6 at various temperatures and Eyring analysis with determined activation parameters for C–F activation.

Fig. 4 (a) Energy density difference map (EDDM) for TS-1. (b) the change (Δ) in NPA charges from Int-A′ to TS-1 and (c) QTAIM data on TS-1: ρ(r) values in bold, V^2ρ values in parentheses. Full QTAIM for TS-1 in Fig. S28b.†
Mg and F$_{ortho}$. The attributes of the BCP (small $\rho_{(s)}$, small positive $\nabla^2 \rho$) are indicative of a weak, closed shell interaction. The DFT prediction of this secondary Mg--F interaction is particularly noteworthy, as experimentally the addition of C--F bonds of fluorooarenes to the Mg--Mg bond of 1 always proceeds to give regioisomers in which at least one *ortho*-fluorine atom is adjacent to the C--F bond that breaks.

The effect of the position of fluorine substitution and the importance of the secondary Mg--F$_{ortho}$ interaction on the C--F bond cleavage transition state was interrogated by calculating the pathway for the addition of C$_6$F$_5$H to 1. Fig. 5 compares the change in Gibbs activation energy for the regioisomeric transition states (relative to TS-1). DFT calculations predict that substitution at positions X$^2$ and X$^3$ should lead to lowest energy transition states. Experimentally, regioisomeric products of C--F activation of both these positions of C$_6$F$_5$H are observed.$^{27}$ Substitution (of fluorine for hydrogen) at X$^1$ is highly dis-favored and removes the *ortho*-fluorine substituent that anchors the substrate in the C--F bond activation TS ($\Delta \Delta G = +5.1$ kcal mol$^{-1}$). For comparison, X$^1$ is not involved in appreciable Mg--F$_{ortho}$ bonding in the transition state and substitution here only raises the TS by +1.8 kcal mol$^{-1}$. Substitution at X$^2$ also raises the energy of the C--F activation TS ($\Delta \Delta G = +2.8$ kcal mol$^{-1}$). This position is expected to exert a secondary influence on the Mg--F$_{ortho}$ binding energy. The fluorine atom in position X$^2$ is a weak $\pi$-donor and will increase the electron density on X$^3$. Replacing X$^3$ with a hydrogen atom will remove this effect diminishing the electrostatic interaction between Mg and F$_{ortho}$.

The importance of the *ortho*-fluorine substituents was further exemplified by calculating the transition state for C--F bond activation of 1,3-C$_6$F$_5$H$_2$ in which both positions *ortho*- to the breaking C--F bond are occupied by hydrogen atoms (Fig. S30†). The absence of any Mg--F$_{ortho}$ anchor results in a reorganisation of the TS to a conformer in which the breaking C--F bond is parallel with the Mg--Mg bond ($\Delta \Delta G = +9.5$ kcal mol$^{-1}$ relative to TS-1). The geometry of this reorganised TS is consistent with that reported for the addition of CO$_2$ to 1 reported by Maron and co-workers.$^{41}$

**Effect of steric on the TS for C--F activation**

A number of analogues of 1 are known. These include a series of $\beta$-diketiminate complexes which differ in the steric profile of the flanking aromatic groups on the ligand. Consideration of the data above leads to the obvious question, what are the effects of steric in C--F bond activation by addition to a Mg--Mg bond?

The potential energy surface for C--F bond activation of C$_6$F$_6$ by 2 ([MesBDIMg--MgBDIMes]$_2$) was calculated and compared to that of 1 ([Dipp$_4$BDIMg--MgBDIDipp])$_2$. The reduced steric bulk of the mesityl substituents results in greater access to the Mg--Mg bond. Consequently, the substrate coordinates to Mg in 2 through a short Mg--F interaction (2.76 Å) to give Int-C. Overall, the formation of Int-C is facile ($\Delta G = 0.3$ kcal mol$^{-1}$). Rotation about the Mg--Mg bond in Int-C gives the more reactive conformer Int-C’ which contains a stronger Mg--F interaction (2.66 Å) in the form of a 1:1-bound fluorooarene adduct.$^{44,45}$ The transition state for C--F cleavage TS-2 is located 23.0 kcal mol$^{-1}$ above the reactants and leads to Int-D (Fig. 6a). Significantly TS-2 contains two Mg--F$_{ortho}$ interactions as evidenced by the short M--F$_{ortho}$ distances (Fig. S31†) and quantified by QTAIM analysis (Fig. 6b). These electrostatic interactions undoubtedly contribute to transition state stabilisation and explain the lower energy of TS-2 relative to TS-1 (Fig. 6a).

Experimental data were collected to verify the hypothesis that 2 is more reactive to fluorooarenes than 1 (Fig. 6c). The reaction between 2 (0.02 M) and 10 equiv. of C$_6$F$_6$ in C$_6$D$_6$ was facile, with immediate observation of $^{19}$F NMR resonances at $\delta = -118.5$, $-157.1$ and $-158.3$ ppm corresponding to the Mg--C$_6$F$_6$ moiety. A further resonance at $\delta = -183.3$ ppm is assigned to a Mg--F by-product.$^{46}$ A competition reaction between 1, 2, and C$_6$F$_6$ led exclusively to products derived from C--F bond activation with 2. Further experiments replacing C$_6$F$_6$ with 2-(2,3,4,5,6-pentafluorophenyl)-pyridine (2-Py-C$_6$F$_5$) again led to exclusive reaction of 2 with no evidence for the participation of 1 despite its known reaction with the same substrate. As with the analogous reaction using 1, 2 reacts with 2-Py-C$_6$F$_6$ at the 2-position. Both the C--F activation product 2a and the fluoride-containing by-product [{[MesBDIMg]$_2$(μ-F)$_3$}] (2b) have been characterised by single crystal X-ray diffraction, the latter complex forming a fluoride bridged trimer in the solid state.$^{47}$

**Effect of M--M bond polarisation on the TS for C--F activation**

The polarisation of the Mg--Mg bond in TS-1 and TS-2 raises a further question: do reactions of polar M--M bonds proceed faster than apolar M--M bonds? To address this question a series of new heterobimetallic complexes with polar M--M’ bonds (M = Mg, M’ = Zn, Al; M = Zn, M’ = Al) were prepared. Complexes 3, 4, 6, 7 are novel M--M containing species, and their synthesis was inspired by recently described methods, namely metathesis,$^{48,51}$ and insertion of an Al(IV) reagent$^{51-59}$ into metal–carbon bonds.$^{50,63}$ Despite the exploitation of known methods, this represents a highly novel series of main group complexes containing metal–metal bonds, including the first example of a crystallographically characterised Mg–Al bond and rare examples of Mg–Zn, Zn–Al and Zn–Zn–Zn bonds for which there is limited precedent.

Complex 3, containing a new Mg–Zn bond, was synthesised by the stoichiometric reaction of the corresponding zinc iodide with 2 (Scheme 1). The reaction occurs with concomitant formation of one equiv. of the $\beta$-diketiminate supported magnesium(II) iodide and 3 could be isolated cleanly by exploitation of the difference in solubility between the two products (54%
isolated yield). Compound 3 was found to be unstable in solution, degrading slowly over time (days) to form the mixed bis-$\beta$-diketiminate complex $\text{[Mg(DippBDI)(MesBDI)]}$ along with a black precipitate (presumed to be $\text{Zn}^{2+}$). Compound 4 can, however, be stored over a prolonged period of time at 238 K, as a solid or in hydrocarbon solution. This complex may be trapped by inser-
tion of a carbodiimide ($\text{RN}_2$) into the Mg-$\text{Zn}$ bond and details of these experiments are provided in the ESI.$^{*}$ An analogue of 3 bearing solely mesityl substituted ligands was also prepared but could not be isolated cleanly.

X-ray quality crystals of 3 were grown from n-hexane at 238 K. The Mg-$\text{Zn}$ bond length in 3 is 2.5993(8) Å, which is comparable with the only two previously reported examples of complexes containing a magnesium–zinc bond (Fig. 7, Table 1).$^{24}$ This value lies between the bond lengths of analogous Mg-$\text{Mg}$ (2.8078–2.8700(9) Å) and $\text{Zn}$-$\text{Zn}$ (2.3586(7)–2.3813(8) Å) $\beta$-dike-
timinate complexes.$^{25,21,42–44}$ Comparison of the series of structures reveals that the complexes adopt a number of conformers in the solid state which differ by the dihedral angle between the two $\beta$-diketiminate planes. A low energy difference between these conformers is anticipated, as would be consistent with the small calculated energy difference between Int-A and Int-A' (Fig. 3).

Addition of the coordinating solvents, THF and pyridine, to 3 led to the formation of the coordination complexes 3-THF and 3-Py, respectively. Integration of the $^1$H NMR spectra indicates that in both cases only one equiv. of solvent is bound (Scheme 1). This was confirmed by the X-ray structure of 3-THF, which shows one molecule of tetrahydrofuran coordinated to magnesium, whilst the zinc remains in a three-coordinate geometry (Fig. 7).

The Mg-$\text{Zn}$ bond length of 3-THF (2.6816(8) Å) is signifi-
cantly longer than that of 3, a phenomenon also observed in the di-solvated solid-state structures of 1 (where solv. = THF, pyri-
dine, DMAP).$^{24}$ 3-THF is a loose approximation of Int-C$^*$ (Fig. 6a) and suggests that substrate coordination to a single metal center is a feasible proposition in the pathway for C–F bond activation.

Following isolation of 3, further small amounts of bright yellow crystals were observed to form in the reaction filtrate. Single crystal X-ray diffraction revealed formation of the novel tri-$\text{Zn}$ species, 4. Compound 4 is only the second reported example of a linear tri-$\text{Zn}$ complex.$^{34}$ 4 is $C_3$ symmetric and the $\text{Zn}$-$\text{Zn}$-$\text{Zn}$ bond is linear (180°). The $\text{Zn}$-$\text{Zn}$ bond length of 2.3908(3) Å is slightly longer than in the previously reported example bearing mono-dentate amide ligands. Compound 4 is iso-structural with a previously reported $\text{Zn}$--$\text{Hg}$--$\text{Zn}$ complex.$^{22,66}$

Related Mg-$\text{Al}$ and $\text{Zn}$--$\text{Al}$ complexes were also prepared. The stoichiometric reaction of 5 with magnesium and zinc alkyl complexes in $\text{C}_6\text{D}_6$ resulted in the formation of new asymmetric products, 6 and 7, respectively (Scheme 1). Deep red crystals of 6 and yellow crystals of 7 were grown from concentrated n-hexane solutions at 238 K (Fig. 7). Single crystal X-ray diffraction studies confirmed the presence of a newly formed aluminium–metal bond, which measured 2.7687(5) Å for Al–$\text{Mg}$ and 2.488(1) Å for $\text{Al}$–$\text{Zn}$.$^{66}$ The aluminium centers in 6 and 7 have a distorted tetrahedral geometry, with $\gamma'$ values of 0.87 and 0.93, respectively. Both compounds 6 and 7 were found to be remarkably

Scheme 1 Preparation of polar $M$–$M$ heterobimetals.
stable in solution, with no degradation observed when heated for 48 h at 353 K.\textsuperscript{67}

The effect of polarity on C–F activation with main group compounds containing M–M’ bonds was investigated. Reaction of the heterobimetallic species 3, 6 and 7, with fluoroarenes were conducted in hydrocarbon solution. No evidence of C–F bond activation was observed upon reaction of Mg–Al (6) or Zn–Al (7) containing compounds with a range of fluoroarenes (e.g. C₆F₆, C₆F₅H and 2-Py-C₆F₅) between 298 and 353 K. While 3 does react slowly with C₆F₆ at room temperature this reaction is competitive with decomposition of the Mg–Zn reagent.\textsuperscript{68,69} A clean reaction was observed between 3 and 2-Py-C₆F₅ to form the Zn complex 3a and the corresponding magnesium fluoride (Scheme 2, Fig. 7).

The reaction pathway for the addition of C₆F₆ to 3 was calculated and shown to be analogous to that described for 1. C–F bond cleavage occurs via an asymmetric transition state TS-3 (Fig. S32\textsuperscript{†}) which leads to Zn–C and Mg–F bond formation and is 26.4 kcal mol\textsuperscript{−1} higher in energy than the separated reactants. Consideration of the difference in electronegativity between Mg (\(\chi_p = 1.31\)) and Zn (\(\chi_p = 1.65\)) explains the geometry of the transition state and the destination of the fluorine atom in the reaction. All attempts to locate a transition state that resulted in Mg–C and Zn–F bond formation led to TS-3 and it is clear that polarisation of the M–M’ bond is a key factor in determining the outcome of the reaction (Fig. 8a).

In combination, these experiments suggest that Mg–Zn, Mg–Al and Zn–Al reagents prepared herein are all less reactive toward fluoroarenes than 1 or 2. This finding was borne out by competition experiments using both C₆F₆ and 2-Py-C₆F₅. The rate of reaction of 3 with 2-Py-C₆F₅ is marginally slower than

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**Table 1** Selected bond lengths from the X-ray data of 3–7\textsuperscript{a}

<table>
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<tr>
<th></th>
<th>3</th>
<th>2a</th>
<th>3a</th>
<th>4</th>
<th>6</th>
<th>7</th>
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<td>M–M</td>
<td>2.5993(8)</td>
<td>2.6816(8)</td>
<td>—</td>
<td>2.3908(3)</td>
<td>2.7687(8)</td>
<td>2.4877(10)</td>
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<tr>
<td>fsr</td>
<td>0.99</td>
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<td>0.98</td>
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<tr>
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<td>2.0782(19)</td>
<td>2.037(3)</td>
<td>1.979(2)</td>
<td>1.9761(18)</td>
<td>2.0862(10)</td>
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<td>2.0540(19)</td>
<td>2.0999(19)</td>
<td>2.035(3)</td>
<td>1.973(2)</td>
<td>2.016(3)</td>
<td>1.9739(10)</td>
</tr>
</tbody>
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\textsuperscript{a} fsr is the formal shortness ratio defined as the bond length/sum of the single bond radii (Pauling).
with 2, \( t_{1/2} = 1 \text{ h} \) versus <0.5 h, but is significantly faster than degradation of 3. Competition experiments between 3 and 2 confirm the relative rates and led to reaction of the Mg-Mg reagents in all cases (Fig. 8b). While competition experiments between 3 and 1 also resulted in selective formation of the organomagnesium species, in this instance trace amounts of 3a are also formed. The relative reactivity is predicted by DFT calculations and TS-3 is higher in energy than both TS-2 (\( \Delta G = 3.4 \text{ kcal mol}^{-1} \)) and TS-1 (\( \Delta G = 0.7 \text{ kcal mol}^{-1} \)). Given the small energy difference involved, the computational methods do not result in a quantitative prediction of the reaction outcome in terms of an accurate product distribution but rather a qualitative prediction of the most reactive species. In line with these predictions, C-F bond activation by 6 and 7 are calculated to occur by concerted transition states TS-4 and TS-5 (Fig. S34) with \( \Delta G_{298}^\ddagger \) of 40.0 and 65.1 kcal mol\(^{-1}\) respectively. While in both cases the calculations predict the formation of C-Al bonds upon C-F activation with 6 or 7, the activations barriers are prohibitively high and these reactions are not observed experimentally.

**Conclusions**

**Sterics versus bond polarisation**

Comparison of TS-1-3 and consideration of the competition experiments definitively shows that sterics are the dominant factor at play in determining the reactivity of M-M bonds toward C-F activation of fluoroarenes. Despite polarisation of the M-M bond occurring in the transition state to C-F bond cleavage, introduction of polarity into the M-M' bond by variation of the metals, while determining the selectivity of addition, does not lead to faster rates of reaction. Moreover, blocking access to the M-M' bond through increasing the coordination number of one of the metals from 3- to 4-coordinate (in 6 and 7) completely shuts down reactivity. These results can be rationalised based on the ease of the formation of the C-F bond activation transition state. The polar heterobimetals (such as 3) have contracted M-M' bonds relative to apolar ones, due to the ionic contribution to bonding, and this bond length contraction closes up the steric pocket that is required for formation of the transition state for C-F bond activation (Table S7, Fig. S34†).

**A thermodynamic ortho-fluorine influence and kinetic ortho-fluorine effect**

In our initial communication of this work, we showed that C-F bond activation with 1 occurs at positions that are flanked by at least one ortho-fluorine substituent.27 Substrate control of regioselectivity is rigorously enforced in these systems and an ortho-fluorine substituent appears essential for the reactions to proceed.

Reactivity trends of fluorinated arenes with transition metal complexes are increasingly well understood in both qualitative and quantitative terms.1,5,7-9,71 Experimentally, the activation of C-F bonds becomes kinetically more challenging as the number of fluorine atoms in \( C_6F_{n-1}H_n \) decreases.27 The influence of fluorine substitution on C-M bond strengths has been the subject of numerous studies. The C-M bond strength increases with an increasing number of ortho-fluorine substituents and this phenomenon is in part accountable for the “ortho-fluorine effect” in reactions that break C-H bonds of fluoroarenes.

Whittlesey and Macgregor have calculated the influence of substitution patterns on the properties of fluoroarenes and found that substituting hydrogen for fluorine decreases C-F bond strengths in the order ortho -> meta -> para.72 Hence, there is a clear thermodynamic rationale for C-F bond activation with M-M reagents to occur at positions flanked by one or more fluorine substituents. These positions are likely to have the weakest C-F bonds and lead to the formation of strongest C-M bonds.

Consideration of TS-1-3 reveals that a kinetic factor also needs to be considered. M...Fortho fluorine interactions are not only present in the key transition states for C-F bond activation, they determine the geometry of these transition states and significantly lower their energy. The role of electrostatic M...F bonding in C-F bond activation with organocerium complexes by a ‘harpoon mechanism’ has been highlighted by Maron and coworkers.74 Furthermore, Whittlesey and Macgregor have also commented on the effect of ortho-fluorine substitution on stabilising transition states during concerted C-F bond activation with transition metal complexes.72 To clarify the dual role of ortho-fluorine substitution in the reaction pathways reported
herein we suggest that C–F bond activation by M–M reagents could be best described as being subject to:

(i) an ortho-fluorine influence that results in destabilisation of the starting materials and stabilisation of the reaction products due to weakened C–F bond strengths and increased M–C bond strengths.

(ii) an ortho-fluorine effect that results in a rate-acceleration through stabilisation of the transition state for C–F bond activation by secondary M···F interactions.

Conflicts of interest
There are no conflicts to declare.

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References
30 Full reference in ESI.
39 A series of functionals were tested and benchmarked against experimental data (activation parameters, molecular coordinates of stationary points). These studies are presented in the ESI.† While the B3PW91 functional most accurately modelled the experimental Gibbs activation energy associated with TS-1, the ωB97X functional was not only within a reasonable margin of error for this data point but also captured all the qualitative experimental trends reported herein.
The formation of high nuclearity species such as trimers is isolation and separation of the reaction products was hindered by the complex Schlenk equilibria which persist even in the presence of a strongly coordinating co-ligand such as pyridine or 4-dimethylaminopyridine (Fig. S7†).

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