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Redox-active ligand controlled selectivity of vanadium oxidation on Au(100)[†]

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Metal–organic coordination networks at surfaces, formed by on-surface redox assembly, are of interest for designing specific and selective chemical function at surfaces for heterogeneous catalysts and other applications. The chemical reactivity of single-site transition metals in on-surface coordination networks, which is essential to these applications, has not previously been fully characterized. Here, we demonstrate with a surface-supported, single-site V system that not only are these sites active toward dioxygen activation, but the products of that reaction show much higher selectivity than traditional vanadium nanoparticles, leading to only one V-oxo product. We have studied the chemical reactivity of one-dimensional metal–organic vanadium – 3,6-di(2-pyridyl)-1,2,4,5-tetrazine (DPTZ) chains with O₂. The electron-rich chains self-assemble through an on-surface redox process on the Au(100) surface and are characterized by X-ray photoelectron spectroscopy, scanning tunneling microscopy, high-resolution electron energy loss spectroscopy, and density functional theory. Reaction of V-DPTZ chains with O₂ causes an increase in V oxidation state from V^{II} to V^{IV}, resulting in a single strongly bonded (DPTZ²⁻)V^{IV}O product and spillover of O to the Au surface. DFT calculations confirm these products and also suggest new candidate intermediate states, providing mechanistic insight into this on-surface reaction. In contrast, the oxidation of ligand-free V is less complete and results in multiple oxygen-bound products. This demonstrates the high chemical selectivity of single-site metal centers in metal–ligand complexes at surfaces compared to metal nanoislands.

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Introduction

Single-site catalysts have long been sought after as the next step in heterogeneous catalyst development.¹ A variety of strategies have been used to create single-site reaction centers, including reaction sites formed in metal–organic frameworks^{2–5} and single-site alloy systems on metals and oxides.^{6–10} There has also been extensive work, including reactivity studies, on tethered metal complexes on metal nanoparticles¹¹ and oxides.^{12–18} In general, single-sites in these systems show higher selectivity than typical heterogeneous catalysts¹⁹ and the ability to avoid coking in a variety of chemical reactions.²⁰ A newer method for forming single-sites is on-surface metal–ligand coordination to

produce metal–organic networks on a solid surface. These have generated growing interest due to the uniformity of the single-site metal centers that they produce.^{21–25} The structure of on-surface metal–ligand networks is determined by the coordination of metal atoms to organic ligands, which form through self-assembly. This strategy for single-site formation has many advantages, such as the previously mentioned uniformity of the metal sites, which should allow for high selectivity, and their open axial coordination sites, which makes these metal centers good potential reaction sites, in addition to being well suited for gas sensor applications.²⁶ Contact with the underlying surface is also an advantage, as the surface can play a beneficial role in reaction mechanisms, which allows for additional, more complicated, chemistry to occur. One specific example of an on-surface metal–ligand coordination network, which has been studied in detail both experimentally^{27,28} and theoretically,²⁹ is the formation of 1D chains by on-surface redox-active self-assembly of platinum atoms and dipyridinyltetrazine (DPTZ). The Pt atoms in this structure have been oxidized to Pt^{II} and the DPTZ ligands have been reduced.

As often as the above advantages of on-surface metal–organic coordination networks have been discussed, they have rarely

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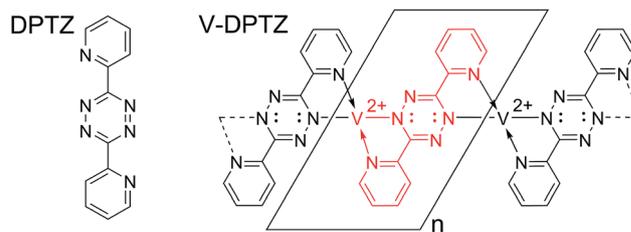


been demonstrated. In most previous studies investigating the interaction of gases with on-surface coordination networks, the gas molecules simply adsorb, with no further chemistry occurring until they later desorb intact. For example, nitric oxide (NO) was found to bind to Fe and Co metal centers (bound in porphyrins) where it was seen that the strength of the metal bonds to the Ag(111) surface were noticeably weakened after complexation with NO.³⁰ The NO coordination was reversed between 500–600 K. A previous study also shows the physisorption of CO₂ on 1D metal–ligand chains formed from Au adatoms and 1,4-phenylene diisocyanide ligands on Au(111) and Au(100),³¹ where a temperature of 90 K was needed for the CO₂ to adsorb, indicating a weak gas interaction.

In some cases reactivity has been demonstrated, such as O₂ cleavage by manganese porphyrins on the Ag(111) surface.³² The metal centers were oxidized from Mn^{II} to Mn^{III}, with one oxygen atom bound to each Mn center but the mechanism for the migration of one oxygen atom to the neighboring unoxidized Mn center was not discussed. Oxygen was also shown to dissociate, and then recombine, on di-iron centers formed by coordination with terephthalic acid on the Cu(001) surface.³³ Two O₂ molecules react with the di-iron centers, leaving two bound oxygen atoms (one on each Fe) and a recombined O₂ molecule, which leaves. Upon oxidation, the Fe–Fe bond distance in this study decreases from 4.45 Å to 3.82 Å. While these studies show the potential reactivity towards bond cleavage that single-site metals in on-surface coordination networks can have, they do not address the single-site selectivity advantage that should also be present.

The results presented in this paper are in one way distinguished from previous surface syntheses because of the choice of DPTZ as the coordinating ligand. Because it is redox-active, after metal coordination, the reducing power of the metal is retained nearby, stored in the π* orbitals of DPTZ, leaving them available for subsequent use towards the substrate, in this case O₂. DPTZ also contains binding pockets that are well suited for metal coordination. The diverging nature of DPTZ nitrogen lone pairs, with these binding pockets located on either side of the molecule, allows for the formation of extended coordination chains.

Vanadium has been chosen for study here since it exists in a large number of adjacent oxidation states, with facile inter-conversion among these, and because of its wide use in metal–ligand complexes for homogeneous catalysis^{34,35} and its oxides in heterogeneous catalysis.^{36–39} Although vanadium oxide formation has been studied on a variety of surfaces,^{40–46} only rarely has the oxidation of individual metal particles been examined,⁴⁷ which is central to the results we present here. O₂ reacting with V metal centers in metal–ligand networks on surfaces has not been previously studied, but is important to consider, given the importance of vanadium and vanadium oxides in both oxidation reactions^{48,49} and oxygen reduction reactions.⁵⁰ Vanadium phthalocyanine complexes have been synthesized on the Ag(111) surface,⁵¹ and vanadium oxide phthalocyanine complexes on surfaces have also been studied,^{52,53} but these complexes were deposited pre-made, not created *via* an on-surface reaction.



Scheme 1 Di-pyridinyltetrazine (DPTZ) and the M–L Polymer that forms after their redox assembly on the Au(100) surface.

The V complex used in this paper is formed by an on-surface redox process between elemental vanadium and DPTZ (Scheme 1) on a Au(100) surface to form V^{II} sites; this on-surface redox process was the subject of a recent publication.⁵⁴ Here, we address the reactivity of these V single-site complexes with dioxygen. The combination of scanning tunneling microscopy (STM), X-ray photoelectron spectroscopy (XPS), and high-resolution electron energy loss spectroscopy (HREELS) allows clear characterization of the reaction product. The use of density functional theory (DFT) allows for detailed mechanistic and structural understanding of V-DPTZ and its interaction with the Au(100) surface through the reaction with O₂.

This vanadium oxidation reaction is of interest because it produces a variety of products on metallic V nanoislands.^{55,56} However, as we shall show, nearly all of the single site V atoms undergo a two electron oxidation to the same product. To our knowledge, this is the first time that such an oxophilic metal has been used in single-site metal–ligand assembly reactivity studies. Providing a direct comparison between metal/surface and metal–organic/surface behavior in the context of reactivity with a gaseous species has not previously been shown, and allows for a unique insight into the role that ligands play in the increased selectivity of coordinated metal centers compared to metal nanoislands.

Results and discussion

Vanadium nanoisland oxidation on Au(100)

The Au(100) surface, used for all of the experiments presented in this work, undergoes a spontaneous reconstruction of the top-most atomic layer to more closely resemble the packing of the lower-energy Au(111) hexagonal packed surface.^{57–59} One effect of this well-characterized reconstruction is a periodic variation in the heights of the Au atoms or “rows”, which lead to anisotropy in the growth of metal islands on this surface.⁵⁸ Vanadium grows on Au(100) at submonolayer coverages as flat, rectangular nanoislands that are oriented in the direction of the row features created by the surface reconstruction (Fig. 1d), similar to the growth of other metals on the same surface.⁵⁸ The average height of these islands is 2.3 ± 0.2 Å. XPS shows that the V in these islands is in a zero oxidation state (Fig. 1a). The V islands were exposed to a small amount of O₂ gas by back-filling the chamber to 1 × 10^{−6} Torr O₂ for 50 seconds, which is a 50 Langmuir (L) exposure (50 L means that, on average, each surface atom experiences 50 collisions from O₂ gas molecules).



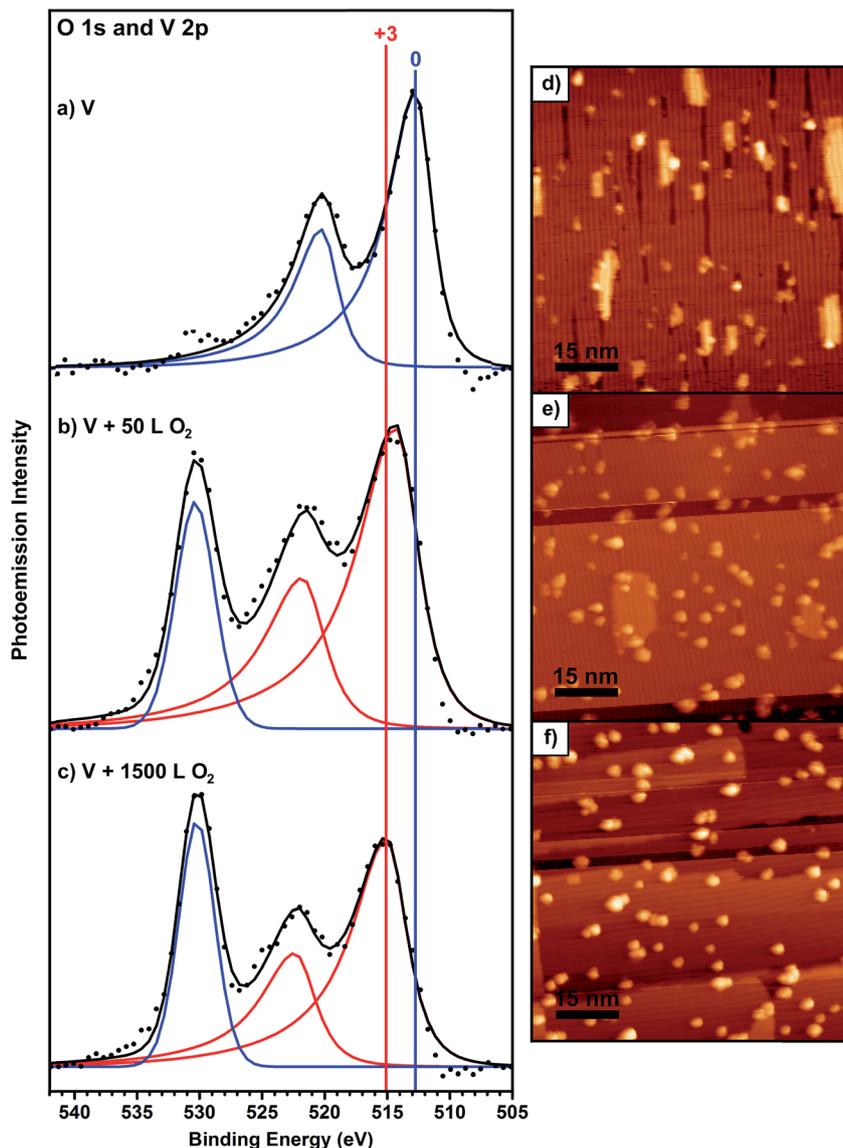


Fig. 1 XPS spectra and STM images showing the reaction of O_2 with V metal deposited on the Au(100) surface while holding the surface at 190°C . (a, b, c) XPS data shown as black points with fit components in red and blue colors and the sum of the fit components as a black line. (a) XPS of metallic V with (d) corresponding STM image of V islands. (b) After 50 L O_2 exposure, an O 1s component is clearly observed along with a shift in the V peak position with (e) corresponding STM image. (c) After 1500 L of O_2 , the surface has been saturated and (f) corresponding STM images showing the resulting VO_x nanoparticles. After reaction with O_2 , the V 2p peaks are noticeably wider, indicating the presence of multiple V oxidation states. The vertical blue and red lines drawn through the XPS data show the positions of the V oxidation states of 0 and +3, respectively. Note that all of these XPS and STM measurements were made in the same UHV system on the same sample.

After a 50 L O_2 exposure while holding the surface at 190°C , the originally flat rectangular V islands become rounded particles with an average height of $6.1 \pm 1.5 \text{ \AA}$, with only a small amount of flat islands remaining (Fig. 1e). Some rounded particles are observed on the few remaining flat V islands, usually at the edges of the islands (Fig. 1e), which may indicate a higher reactivity of the island edges compared to the centers. XPS data shows the appearance of an oxygen peak (Fig. 1b) and a number ratio of oxygen to V of 0.8 : 1. There is also a noticeable change in the V 2p peak width and position in XPS compared to before oxygen exposure (Fig. 1b). The full width at half maximum (FWHM) of both the V $2p_{3/2}$ and $2p_{1/2}$ peaks increase

significantly, from 3.1 eV to 4.3 eV. The binding energy (BE) of the V $2p_{3/2}$ peak increases by 1.6 eV to 514.5 eV, indicating oxidation of the V. Based on the increased FWHM of the V 2p XPS peaks, we can conclude that not all of the V is oxidized. A fit model deconvoluting this data into V^0 and V^{III} components is shown in Fig. S6.†

High-resolution electron energy loss spectroscopy (HREELS) is a surface-sensitive vibrational spectroscopy technique that allows for direct chemical characterization of reactions and bonding on a surface and is particularly effective for organic compounds.^{60,61} Details regarding the experimental procedure for HREELS can be found in the ESI.† When the V/Au(100)



surface is exposed to 50 L of O₂, a new HREELS peak is observed at 1029 cm⁻¹ corresponding to a vanadyl terminal VO stretching mode, showing O/O bond scission, as well as a broad feature at 684 cm⁻¹ due to the V–O–V bridging features that result from O binding to multiple V atoms in various conformations (Fig. 2).^{55,56,62,63} These indicate that there is not one uniform V oxide structure, but multiple binding sites for O on the V nanoparticles. As the O₂ exposure is increased to 500 L, a new peak appears between the 684 and 1029 cm⁻¹ peaks. We note that no significant spectral features are observed for V/Au before O₂ exposure and that the clean Au(100) surface is completely inert to these O₂ exposures as determined by HREELS.

Longer O₂ exposure experiments were conducted to test whether VO_x could be driven to a more uniform oxidation state. Further O₂ exposures at 190 °C reveal that the surface saturates at a total exposure of 1500 L (Fig. 1c); exposures beyond that did not result in any changes in the XP spectra. At this point, the V 2p_{3/2} peak shifted by an additional 0.8 eV to reach a position of 515.3 eV, indicating oxidation of more V, with an O : V number ratio of 1.2. The FWHM decreased slightly, to 3.9 eV, but remains much larger than V alone on the surface (3.1 eV), showing that although the V oxidation states are better defined than at lower exposures, V still exists in multiple oxidation states. STM reveals that the surface is populated by bright, round particles with an average height of 8.9 ± 1.6 Å (Fig. 1f). The increased height of these particles is due to both their increased oxygen content and sintering as a result of the 190 °C annealing done during the exposure steps; this reshaping is expected under these thermal conditions.⁶⁴ To further confirm the effect of oxygen on the size of the particles, the average particle volume was calculated by tabulating many particles

from several images. Although the average particle volume decreases after oxygen exposure, as the reaction with O₂ breaks up the larger V islands into smaller particles (Fig. 1), the total volume per 80 nm × 80 nm STM image increases noticeably. Before O₂ exposure, the V islands occupy an average volume of 89 nm³ per image. After 50 L of total exposure, this volume increased to 182 nm³ and it increased further to 237 nm³ after 1500 L of total exposure. This significant increase in the total mass of the adsorbed islands is due to the binding of O during oxidation. This value did not increase after further O₂ exposure, providing additional evidence that the surface was already saturated.

To provide a comparison to these results, an additional experiment was performed where V was deposited onto the Au(100) surface in an oxygen background of 1.0 × 10⁻⁶ Torr. This procedure resulted in the formation of V₂O₃, as evidenced by the oxidation of all of the V to V^{III} (V 2p_{3/2} peak position of 515.2 eV), with an O : V ratio of 1.7 (Fig. S3†). After annealing the surface at 350 °C, V₂O₃ remains on the surface, with a V 2p_{3/2} peak position of 514.9 eV and an O : V ratio of 1.6 (Fig. S4†). This is consistent with prior reports of vanadium deposition in an oxygen background to create surface oxides.^{40,47,65} These different experiments show that V atoms in 2D islands do not react in a uniform fashion to form one product.

V-DPTZ chain structure

The formation of V-DPTZ chains on the Au(100) surface was recently reported by our collaboration⁵⁴ and provides an excellent platform here for the study of chemical activity of metal single-sites. We will briefly review key points of the structural and chemical characterization of those chains before examining their reactivity with O₂ gas in the next section, which was not previously studied. DPTZ molecules and V metal were sequentially vapor deposited onto the Au(100) surface in UHV and annealed to form one-dimensional metal–organic chains by an on-surface redox reaction. As described in a previous publication,⁵⁴ V metal is oxidized to V^{II} when it reacts with DPTZ on the Au(100) surface, as evidenced in XPS by a 1.1 eV shift in the V 2p_{3/2} binding energy (BE) from 512.7 eV for metallic V to 513.8 eV for V^{II}. A 1D chain structure is formed, illustrated in Scheme 1 and imaged in Fig. 3d. The chains predominately orient themselves at ±43 ± 4° relative to the Au(100) surface reconstruction rows, although some chains grow along the direction of the reconstruction rows. The chains are also straight and uniform, with few kink defects. The redox assembly into this structure does not depend on the deposition order of the metal and the ligand. This structure of formula V-DPTZ is stable against thermal desorption up to at least 200 °C.⁵⁴ DPTZ was chosen over another possible ligand, bipyrimidinyl tetrazine (BMTZ), because DPTZ is less oxidizing. BMTZ would further deplete reducing power from the attached metal, lowering its potential for high reactivity with added gas.

DFT calculations were carried out for isolated V-DPTZ chain (gas phase) and on the reconstructed Au(100) surface for structural analysis (see Section S2 in ESI† for computational details). For gas phase V-DPTZ chain, our calculations show that

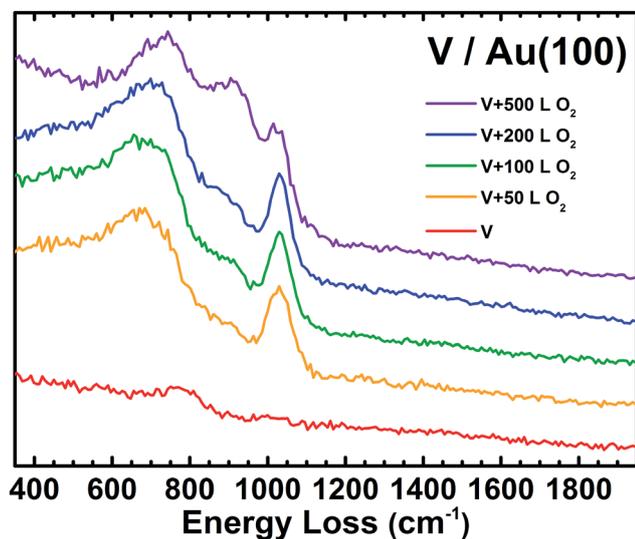


Fig. 2 HREELS of V deposited onto Au(100) with increasing oxygen exposures. The V sample was initially annealed at 190 °C to remove weakly adsorbed contamination. Each O₂ exposure was performed while the sample was held at 190 °C and HREELS was acquired after cooling to RT. The vanadyl peak (V=O) is observed at 1029 cm⁻¹. At lower energies there is a very broad feature correlating to V–O–V bending modes. The broad shape results from the several possible vanadium oxide structures that could form.



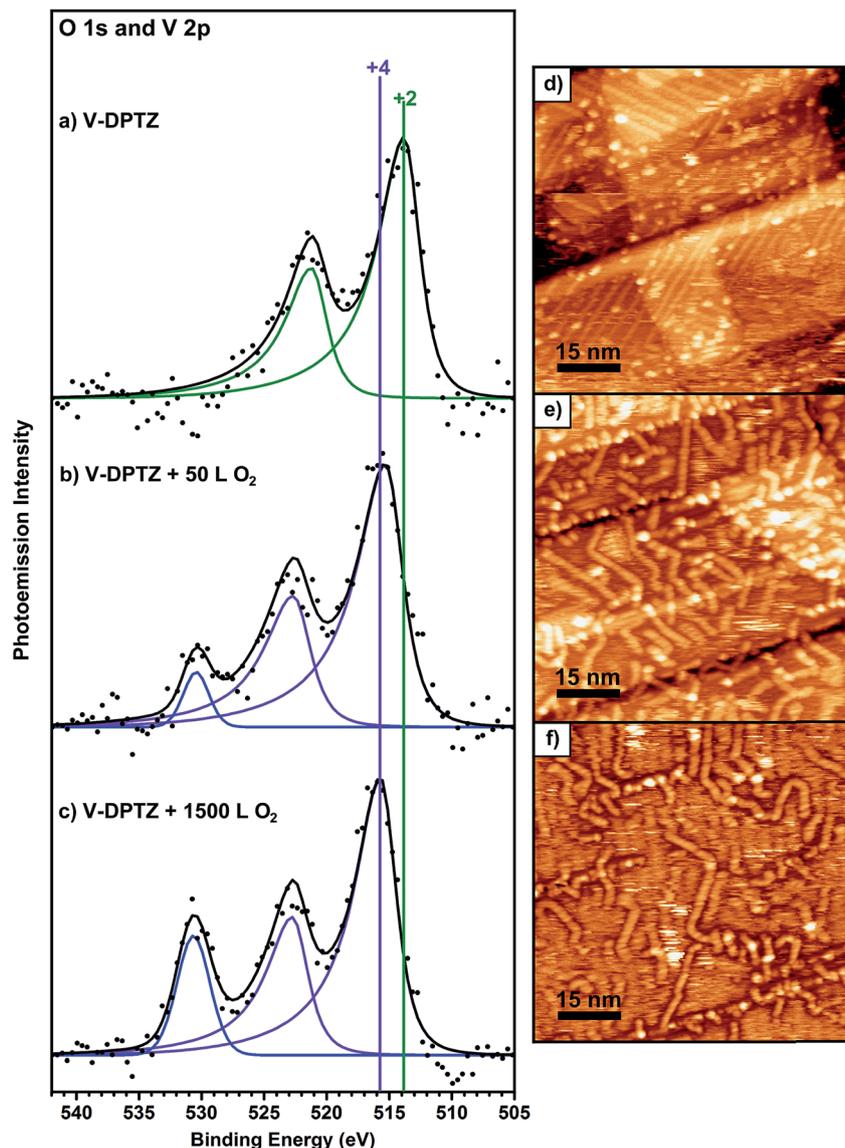


Fig. 3 XPS spectra and STM images showing the reaction of O_2 with V-DPTZ chains on Au(100) while holding the surface at $190^\circ C$. (a, b, c) XPS data shown as black points with fit components in green and purple colors and the sum of the fit components as a black line. V oxidized to +2 is shown in (a) before the addition of O_2 with a typical STM image showing the 1D chains in (d). XPS of O 1s and V 2p after a 50 L O_2 exposure is shown in (b), with the corresponding STM image in (e). In (c), the O 1s and V 2p spectrum after a total of 1500 L O_2 is shown, with the corresponding STM image in (f). The vertical green and purple lines indicate V $2p_{3/2}$ binding energies for the +2 and +4 oxidation states, respectively.

the chain structure is entirely planar with distance between two nearest V atoms of $\sim 6.85 \text{ \AA}$. In this chain structure, each V atom has a magnetic moment of $\sim 3 \mu_B$. The V atoms slightly prefer to be in an antiferromagnetic configuration ($\sim 0.01 \text{ eV}$ lower energy in the antiferromagnetic than the ferromagnetic configuration). Tetrazines in this planar V-DPTZ chain structure have two short C=N and four long intra-ring distances (Fig. 4a), which is different from neutral tetrazine in which there are three long and three short bonds, and is consistent with two localized C=N bonds in the 2-electron reduced ring (see Scheme 2). A second diagnostic of reduced tetrazine is that its V-N distances are shorter by $0.05\text{--}0.1 \text{ \AA}$ than those of pyridine in these structures, since dianionic tetrazine, containing amide N, binds to V more strongly than does pyridine N. Additionally, the

redox character of the interaction between V and DPTZ is confirmed as the V 2p binding energy is calculated to be 2 eV higher than that of bulk V, indicating an oxidation state of +2, consistent with experimental XPS data.

For V-DPTZ chain on the hexagonal reconstructed Au(100) surface, our DFT calculations reveal that the VN_4 unit is not planar: V is drawn towards the surface out of the N_4 plane by 0.43 \AA (for comparison, this displacement is only 0.02 \AA for Pt-DPTZ on the same surface).⁶⁶ This results in a V-Au distance of $\sim 2.59 \text{ \AA}$. This is the analog of the strong metal-support interaction observed between metals and oxide supports, but here it originates from direct metal/metal bonding because of the reduced form of our support. Similar to the gas phase chain, bond lengths in the tetrazine ring (Fig. 4b) show the alternating



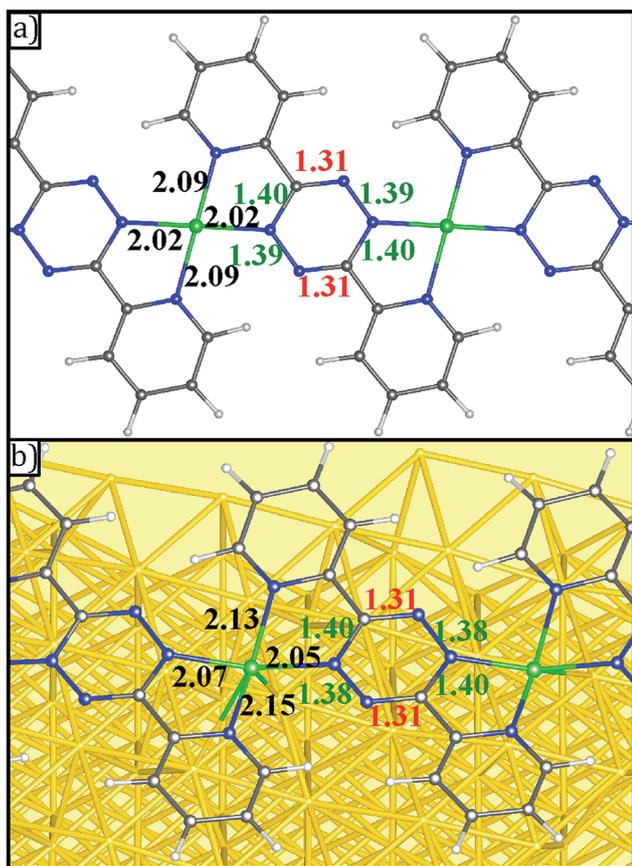
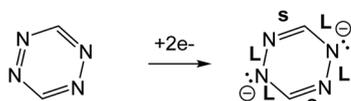


Fig. 4 Geometries of (a) gas phase V-DPTZ and (b) V-DPTZ on reconstructed Au(100) as obtained by DFT simulations. Gold, gray, blue, white, and green points represent Au, C, N, H, and V, respectively. The red, green, and black insert numbers (Å) are short, long (as shown in Scheme 2) bond lengths within the tetrazine ring and VN bond lengths, respectively.



Scheme 2 Bond lengthening and shortening of $2e^-$ reduction of tetrazine.

pattern consistent with C=N bonds and the remaining CN and NN bonds longer, consistent with single bonds. The bond lengths from V to pyridyl (2.15 Å) are longer than to tetrazine (2.07 Å) because the reduced tetrazines serve as stronger amide donors. It is worth noting that because of the pyramidal VN_4 structure, the VN bond lengths are slightly longer than that of gas phase V-DPTZ chain.

The influence of redox-active ligand: V-DPTZ reaction with oxygen on Au(100)

Upon exposure of these V-DPTZ chains to 50 L O_2 while holding the surface at 190 °C, the V 2p region undergoes a noticeable change (Fig. 3b). The V 2p region shifts to higher binding

energy, as evidenced by a shift in the V $2p_{3/2}$ peak by 1.6 eV to 515.4 eV. The FWHM also increases slightly, by 0.3 eV, suggesting that at this stage of the reaction, more than one oxidation state of V is present on the surface. An alternative fitting model that holds the width constant, but allows for more fitting components than Fig. 3b is shown in Fig. S5 (ESI†). An oxygen peak also appears. Based on this evidence, we can conclude that the V is being oxidized to a higher binding energy by reacting with O_2 . Further O_2 exposures were performed on this same surface, reaching total exposures of 500 and 1500 L, holding the surface at 190 °C during each exposure. After 500 L total O_2 , the O 1s peak intensity increases significantly and the V $2p_{3/2}$ peak shifts an additional 0.4 eV to 515.8 eV. After a total O_2 exposure of 1500 L, these values remain unchanged. It is important to note that at this point in the experiment that the FWHM has decreased to match the observed value for just V-DPTZ (and for V nanoparticles), suggesting that all of the vanadium has been oxidized, even though the O : V ratio, based on XPS peak areas, is 0.6. It has been discussed previously that sensitivity factors for the 2p lines of transition metals can vary with oxidation state, which makes intensity comparisons difficult.^{67,68} Based on the narrow peak width and low O : V saturation ratio, we conclude that atomic oxygen binds to every vanadium site. A V $2p_{3/2}$ BE of 515.8 eV represents V^{IV} , as this value corresponds to BE values for V^{IV} oxides.^{69–71} BE values for different V oxidation states are compared in a table in the ESI (Table S1†) showing that the trends we report are consistent with previous research. The N 1s XPS peak is not changed by the O_2 reaction, indicating that DPTZ remains dianionic (Fig. S1 and S2†).

Nature of oxygen bound on V-DPTZ chains and reaction mechanism

The observed oxidation of V^{II} to V^{IV} indicates that only one oxygen atom is binding to each V site. This would mean that V dissociates the incoming O_2 molecules, establishing a product stoichiometry of (DPTZ)VO rather than (DPTZ)VO₂. After 50 L O_2 exposure on V-DPTZ, HREEL spectroscopy shows a new peak at 1007 cm^{-1} (Fig. 5). Previous papers on vanadium complexes report a vanadyl ($V=O$) peak within the range of 960–1033 cm^{-1} .^{72–74} As O_2 exposure increases, the $V=O$ peak increases in intensity and dominates the spectrum, but there are also other small changes to the HREEL spectrum, including the appearance of a small shoulder at 845 cm^{-1} and a slight broadening on the high energy side of the peak at 640 cm^{-1} . These small changes indicate a small fraction of V–O–V bending modes at the surface or possibly some change in the ligand out-of-plane modes as a result of the newly formed $V=O$ bond. These are small in population compared to the $V=O$ species.

DFT calculations provide additional information about the post-reaction product. We have considered a variety of structures for products with one O_2 molecule or two O atoms in a supercell (Fig. 6b–j). Of all structures studied, the most stable structure is that with every metal bond being VO (bond length = 1.61 Å), with the O pointing away from gold (Fig. 6j).



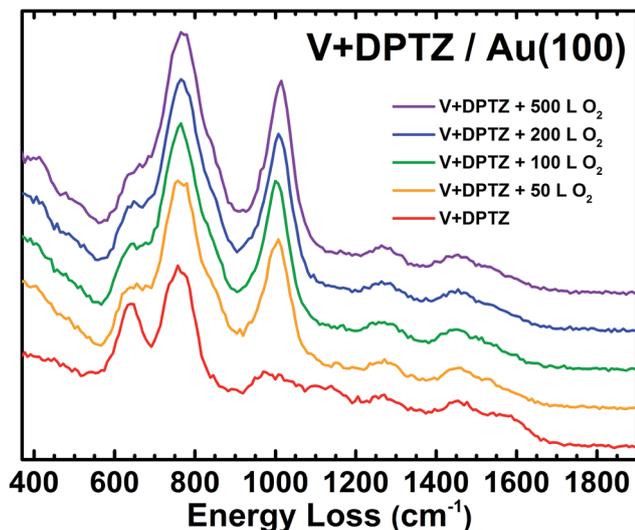


Fig. 5 HREELS of vanadium co-deposited with DPTZ on the Au(100) surface. The V-DPTZ complex was then exposed to increasing amounts of O_2 up to 500 L while the surface was held at 190 °C. HREELS was acquired after the annealing was allowed to cool to RT. A vanadyl peak is observed at 1007 cm^{-1} . No other peaks or bending modes are observed. All other peak intensity changes result from intraligand electronic and geometric differences after oxygen binding.

This introduces a mechanistic question of how one oxygen atom of O_2 arriving at one V reaches the more distant, unoxidized V. The experimentally determined O : V ratio indicates that after the O_2 binding at one V^{II} site (Fig. 6b and c) and $O=O$ dissociation, there must be migration of one of the O atoms from that site. Eventually, there will be oxidation of each V site by one O that would achieve the most stable binding structure shown in Fig. 6j. To achieve this, we propose two potential O-migration mechanisms. The first potential mechanism would involve the oxygen atom moving across the neighboring tetrazine *via* a species with O above the center of the ring with two C–O bonds (Fig. 6h), before moving to the other side and reacting with the neighboring, and unoxidized, V^{II} center (Fig. 6j). The second potential mechanism is that one of the oxygen atoms in the $V(O)_2$ species bends closer to the surface (Fig. 6e). This O then migrates to Au (spillover effect, Fig. 6f) ensuing a weak interaction, then diffusing across the surface until it comes in contact with an unreacted V^{II} center (Fig. 6i and j).

Besides the migration along the Au surface to an empty V center (Fig. 6f and g), O atoms also have two potential desorption pathways available. Previous studies of atomic oxygen on Au(111) and Au(110) have shown that at the temperature used in this study (190 °C), mobile O atoms on Au can react with background CO gas (common in UHV systems) to desorb as CO_2 ,^{75,76} or recombine on the surface to desorb as O_2 .^{76–78}

A detailed investigation of the thermodynamics and kinetics of all competing reactions of O_2 on V-DPTZ chains on the Au(100) surface would provide conclusive information about the mechanism at work. However, the overwhelmingly large system together with the complex nature of the electronic

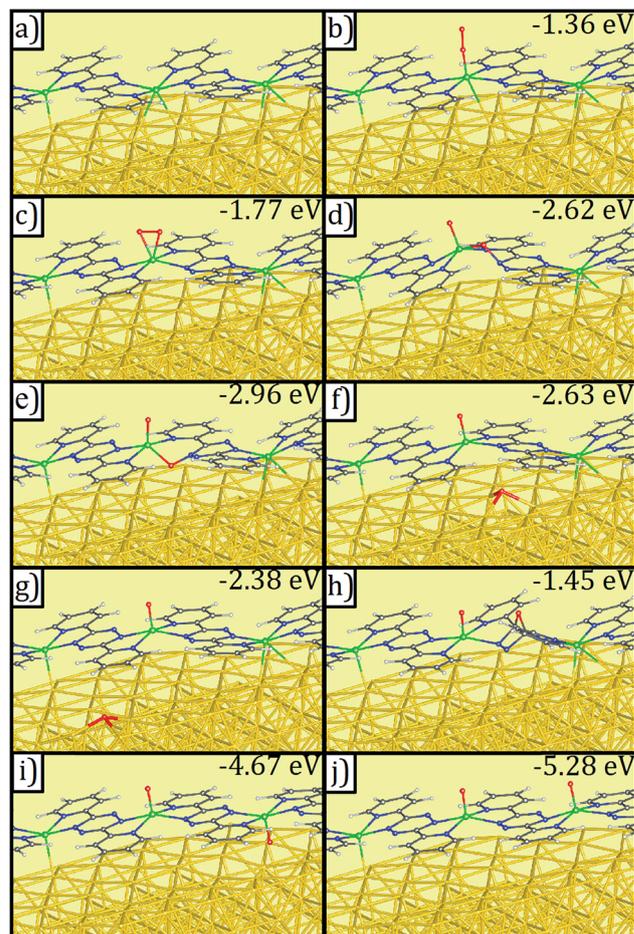


Fig. 6 Geometries of (a) V-DPTZ on reconstructed Au(100) and (b–j) different binding configurations of molecular and dissociated O_2 on V-DPTZ on reconstructed Au(100) as obtained by DFT simulations. Gold, gray, blue, red, white, and green points represent Au, C, N, O, H, and V, respectively. Binding energies of oxygen are listed on the top-right corner of panels (b–j). Molecular O_2 adsorptions are shown in (b) and (c). O_2 dissociation and insertion to VN bonds are shown in (d) and (e). The migration of the O atom across a region of bare Au(100) is shown in (f) and (g). The diffusion of atomic O over tetrazine ring is shown in (h). The most energetically favorable structures, with one O atom per V site, are shown in (i) and (j). Optimized coordinates of these configurations can be found in ESI.†

structure of the VO species makes it a daunting task. Fortunately, comparison of energy between considered states (Fig. 6) suggests a strong preference for the formation of $V=O$ structure as shown in Fig. 6j, *i.e.*, each V binds one O atom, strongly suggesting that O atoms will find unoccupied V centers should they be available. Since the energy of the intermediate with O bridging the two *para* carbons on the tetrazine ring (Fig. 6h) is substantially higher than every other configuration, the transfer of the O in a $[(VO_2)(V)(DPTZ)_2]_n$ chain to the neighboring V by migrating over the *exo* face of the tetrazine (first mechanism) is less favorable than that by migrating across the Au surface (second mechanism). The preferred mechanism occurs through transient insertion of the migrating O into the VN bond (Fig. 6d and e), which illustrates a reverse spillover process, facilitated



by a principle from homogenous catalysis: metal ligand cooperation.⁷⁹

DFT results also indicate that the reaction with O₂ does not significantly alter the distances within the DPTZ ligands and VN bond. The distances within tetrazine ring always show two short C=N and four other longer bonds, consistent with persistence of doubly reduced tetrazine following Scheme 2, as illustrated in Fig. 7 for two configurations shown in Fig. 6b and j. This confirms what was concluded from XP spectroscopic measurements: addition of O₂ to V-DPTZ oxidizes the metal, while leaving the tetrazine reduced. The product also has short VN distances to the tetrazine (~2.09 Å) and longer VN distances to pyridine (~2.13 Å). The VO (bond length 1.61 Å) is square pyramidal and pointing perpendicular to the surface. The major effect of oxygen is the upward movement of V atom (as much as 1 Å) upon VO bond formation, implying an important consequence that shall be discussed next.

Control experiments were conducted to expose either a clean Au(100) surface or DPTZ alone on the surface (no V) to O₂. In either case, an O 1s XPS peak was not observed, demonstrating that O₂ does not adsorb to or react with either DPTZ or the

Au(100) surface at room temperature. O₂ does not adsorb or react on clean Au(111) and Au(110) surfaces.^{76–78}

Disruption of chain structure during oxidation

When the O₂ exposure is performed while holding the surface at room temperature, STM visualization shows that the 1D V-DPTZ chains are disrupted (Fig. S9†), leaving only a few short chains, mostly near step edges. Some striated features in the image indicate the presence of mobile species that are moving faster than the STM scan rate.⁸⁰ We can be confident that the DPTZ ligands have not desorbed upon O₂ exposure based on XPS signal intensity showing that carbon and nitrogen are still present on the surface at the same concentration as before O₂ exposure. Because of the +4 oxidation state of the metal, we can conclude that these mobile species on the surface must be units of formula [(DPTZ)VO]_n. A possible explanation for the disruption of the chain structure, and loss of order, is the strong bond that forms between O and V. Through DFT calculations (Fig. 6 and 7), we see that when oxygen binds to the V centers, V moves farther from the surface by 1.0 Å, making the chains generally less fixed, hence less ordered. Shorter chain fragments could form at this step, broken apart locally by the highly exothermic heat of reaction released in this oxidation. Based on the previous discussion, the surface species shown in Scheme 3 (which both contain V^{IV}) are candidates for what is formed, with dimer B more stable since it has a higher, and typical, coordination number for V^{IV}. Such an arrangement has no outwardly directed pyridyl nitrogen lone pairs, so it is not suitable for (and disrupts) chain growth. This could form by binding O₂ to one vanadium in Scheme 1, followed by rotation of pyridyl nitrogen. While DPTZ in a crystal lattice, and in our 1D chains, has the pyridyl nitrogens mutually *anti*,⁸¹ molecular structures with two metals, *i.e.*, [DPTZ(M)]₂, show that the *syn* conformer can be adopted.^{82–86} This efficiently exhausts the full 4 electron oxidizing power of O₂ locally, and without high-energy (*e.g.*, oxygen atom) intermediates.

Two methods can be used to re-form the chains on the surface. First, annealing the surface, after the oxygen exposure, at temperatures between 170–190 °C, brings back chains (Fig. S9†). These chains have a similar structure as before the reaction with O₂, but now each V center has an oxygen atom bound to it. The second method involves heating the surface during the oxygen exposure. In this case, chains are observed by

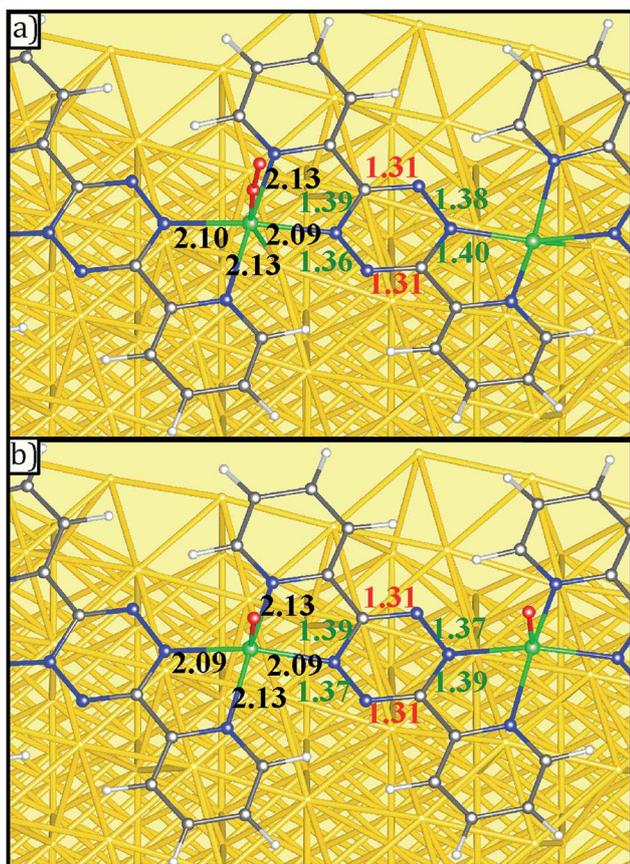
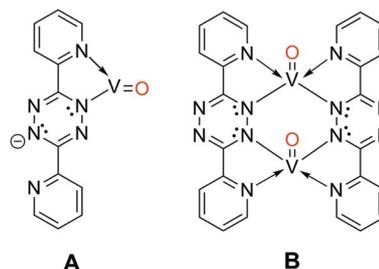


Fig. 7 Details bond length of (a) O₂ molecular adsorption and (b) O atomic adsorption on V-DPTZ on reconstructed Au(100) as obtained by DFT simulations. Gold, gray, blue, red, white, and green points represent Au, C, N, O, H, and V, respectively. The red, green, and black insert numbers (Å) are short, long (as shown in Scheme 2) bond lengths within the tetrazine ring and VN bond lengths, respectively.



Scheme 3 Potential post-O₂ reaction species that form due to the formation of V–O bonds.



STM immediately upon imaging after the exposure. In either case, it is clear that thermal energy is needed for chains to form after the reaction with O₂.

In both cases, there are some clear differences in the chain structure after the reaction with oxygen compared to the V-DPTZ chains before the reaction. The main difference is that the chains show less preferential orientation with respect to the gold surface. The chains do not show a preferred 43° angle of orientation with respect to the reconstruction rows, and instead show many different orientations, including an increased number of chains growing along the rows. Also, the chains are much less straight and uniform, showing a greater number of bends and kinks, which were not present before the reaction with oxygen. DFT calculations can provide some insight into these structural differences. Before the reaction, the V atoms were close to the surface, positioned below the plane of the DPTZ molecules (Fig. 6), showing a relatively strong interaction with the surface. After the reaction, the V atoms rise from the surface by an additional 1.0 Å to 3.60 Å, moving above the plane of the DPTZ molecules and towards the oxo, thus removing the Au/V interaction. This effect of adsorbed gas species weakening the metal-surface bond has been observed before through DFT calculations,³⁰ but the effect is much larger here. V^{IV} complexes are known to prefer square pyramidal geometries over octahedral geometries,^{87,88} so it is reasonable to assume that the oxo bonding on the V sites would weaken bonding at the *trans* position, *i.e.*, weaken the V interaction to the gold surface. This weakened interaction with the surface would account for the increased mobility. It also weakens the stabilization of V atoms in good registry with the surface to make the chains less rigid, allowing for increased bending along the chains, and also allows for less regular registry with the surface.

This study clearly shows that ligand design can have a major influence on selectivity, and opens a broader study of spillover effects and migration mechanisms for substrate atoms between surface-bound single-site metals too far apart to allow concerted M₂ interaction with substrate. Catalysts benefit from not only metal choice, but also ligand design as tools to achieve improved performance.

Conclusion

In this work, we have shown that the V metal centers in 1D V-DPTZ chains on Au(100) are available for reactivity, and that O₂ molecules undergo bond cleavage resulting in one of the oxygen atoms bonded to the V metal center to form V=O and spillover of the other O to the Au surface to migrate and react with another metal center. We have also shown that DPTZ is necessary to fully oxidize the V to V^{IV}. This must be both a structural effect and the influence of charge donation by nitrogen lone pairs; when DPTZ is reduced (dianionic), its nitrogens are strong donors, which increases the reducing power of the attached metal. Utilizing HREELS, the identity of the V–O binding species could be determined. While only a single peak grows with the V-DPTZ for a vanadyl O species, multiple vibrational modes appear when O₂ is exposed to V nanoparticles, showing the formation of a vanadyl V=O species

in addition to V–O–V bending modes. DFT calculations confirmed the preferred bonding structure of one oxygen atom per V site. Additionally, they revealed details regarding the mechanism of the reaction, indicating a likely mechanism involving the insertion of O into a VN bond followed by migration of O across the surface.

These results show that the presence of redox active N-heterocyclic ligands has a noticeable effect on the chemistry of metals on surfaces. In addition, metal atom vaporization here permits exploring ligand-poor (1 : 1 stoichiometry) complexes derived from metal atoms, the latter nearly impossible to access in molecular chemistry in solution. This is a clear demonstration of the uniformity, and high selectivity, that on-surface single-site metal centers have. Studying an increasing variety of metals and ligands is an important next step in understanding the chemistry and potential reactivity of metal–ligand networks on surfaces and their impact for future applications in catalysis and gas capture. Using knowledge from molecular chemistry, including trends in donor power and redox character of both metals and ligands, rational control of selectivity can be envisioned.

Conflicts of interest

There are no conflicts of interest to declare.

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References

- 1 S. D. Pike and A. S. Weller, Organometallic synthesis, reactivity and catalysis in the solid state using well-defined single-site species, *Philos. Trans. R. Soc., A*, 2015, **373**, 20140187.
- 2 J. Lee, O. K. Farha, J. Roberts, K. A. Scheidt, S. T. Nguyen and J. T. Hupp, Metal–organic framework materials as catalysts, *Chem. Soc. Rev.*, 2009, **38**(5), 1450–1459.
- 3 H. Furukawa, K. E. Cordova, M. O’Keeffe and O. M. Yaghi, The Chemistry and Applications of Metal–Organic Frameworks, *Science*, 2013, **341**, 1230444.
- 4 A. S. Poryvaev, A. M. Sheveleva, D. I. Kolokolov, A. G. Stepanov, E. G. Bagryanskaya and M. V. Fedin, Mobility and Reactivity of 4-Substituted TEMPO Derivatives



- in Metal–Organic Framework MIL-53 (Al), *J. Phys. Chem. C*, 2016, **120**(19), 10698–10704.
- 5 S. M. J. Rogge, A. Bavykina, J. Hajek, H. Garcia, A. I. Olivos-Suarez, A. Sepulveda-Escribano, A. Vimont, G. Clet, P. Bazin, F. Kapteijn, M. Daturi, E. V. Ramos-Fernandez, F. X. Llabres i Xamena, V. Van Speybroeck and J. Gascon, Metal–organic and covalent organic frameworks as single-site catalysts, *Chem. Soc. Rev.*, 2017, **46**(11), 3134–3184.
 - 6 J. Liu, F. R. Lucci, M. Yang, S. Lee, M. D. Marcinkowski, A. J. Therrien, C. T. Williams, E. C. H. Sykes and M. Flytzani-Stephanopoulos, Tackling CO Poisoning with Single-Atom Alloy Catalysts, *J. Am. Chem. Soc.*, 2016, **138**(20), 6396–6399.
 - 7 F. R. Lucci, M. T. Darby, M. F. G. Mattera, C. J. Ivimey, A. J. Therrien, A. Michaelides, M. Stamatakis and E. C. H. Sykes, Controlling Hydrogen Activation, Spillover, and Desorption with Pd–Au Single-Atom Alloys, *J. Phys. Chem. Lett.*, 2016, **7**(3), 480–485.
 - 8 C. Y. Wang, G. Garbarino, L. F. Allard, F. Wilson, G. Busca and M. Flytzani-Stephanopoulos, Low-Temperature Dehydrogenation of Ethanol on Atomically Dispersed Gold Supported on ZnZrO_x, *ACS Catal.*, 2016, **6**(1), 210–218.
 - 9 J. Kim, H.-E. Kim and H. Lee, Single-Atom Catalysts of Precious Metals for Electrochemical Reactions, *ChemSusChem*, 2017, DOI: 10.1002/cssc.201701306.
 - 10 J. Liu, B. R. Bunes, L. Zang and C. Wang, Supported single-atom catalysts: synthesis, characterization, properties, and applications, *Environ. Chem. Lett.*, 2017, DOI: 10.1007/s10311-017-0679-2.
 - 11 J. D. E. T. Wilton-Ely, The surface functionalisation of gold nanoparticles with metal complexes, *Dalton Trans.*, 2008, **1**, 25–29.
 - 12 E. M. W. Rumberger, H. S. Ahn, A. T. Bell and T. D. Tilley, Water oxidation catalysis *via* immobilization of the dimanganese complex [Mn-2(μ-O)(2)Cl(μ-O₂CCH₃)-(bpy)(2)(H₂O)](NO₃)(2) onto silica, *Dalton Trans.*, 2013, **42**(34), 12238–12247.
 - 13 C. W. Jones, M. W. McKittrick, J. V. Nguyen and K. Q. Yu, Design of silica-tethered metal complexes for polymerization catalysis, *Top. Catal.*, 2005, **34**(1–4), 67–76.
 - 14 A. W. Moses, C. Raab, R. C. Nelson, H. D. Leifeste, N. A. Ramsahye, S. Chattopadhyay, J. Eckert, B. F. Chmelka and S. L. Scott, Spectroscopically distinct sites present in methyltrioxorhenium grafted onto silica–alumina, and their abilities to initiate olefin metathesis, *J. Am. Chem. Soc.*, 2007, **129**(28), 8912–8920.
 - 15 B. C. Vicente, Z. Huang, M. Brookhart, A. S. Goldman and S. L. Scott, Reactions of phosphinites with oxide surfaces: a new method for anchoring organic and organometallic complexes, *Dalton Trans.*, 2011, **40**(16), 4268–4274.
 - 16 Z. Huang, M. Brookhart, A. S. Goldman, S. Kundu, A. Ray, S. L. Scott and B. C. Vicente, Highly Active and Recyclable Heterogeneous Iridium Pincer Catalysts for Transfer Dehydrogenation of Alkanes, *Adv. Synth. Catal.*, 2009, **351**(1–2), 188–206.
 - 17 J. D. A. Pelletier and J.-M. Basset, Catalysis by Design: Well-Defined Single-Site Heterogeneous Catalysts, *Acc. Chem. Res.*, 2016, **49**(4), 664–677.
 - 18 C. Copéret, A. Comas-Vives, M. P. Conley, D. P. Estes, A. Fedorov, V. Mougel, H. Nagae, F. Núñez-Zarur and P. A. Zhizhko, Surface Organometallic and Coordination Chemistry toward Single-Site Heterogeneous Catalysts: Strategies, Methods, Structures, and Activities, *Chem. Rev.*, 2016, **116**(2), 323–421.
 - 19 J. M. Thomas, R. Raja and D. W. Lewis, Single-Site Heterogeneous Catalysts, *Angew. Chem., Int. Ed.*, 2005, **44**(40), 6456–6482.
 - 20 X.-F. Yang, A. Wang, B. Qiao, J. Li, J. Liu and T. Zhang, Single-Atom Catalysts: A New Frontier in Heterogeneous Catalysis, *Acc. Chem. Res.*, 2013, **46**(8), 1740–1748.
 - 21 W. Auwärter, A. Weber-Bargioni, S. Brink, A. Riemann, A. Schiffrin, M. Ruben and J. V. Barth, Controlled metalation of self-assembled porphyrin nanoarrays in two dimensions, *ChemPhysChem*, 2007, **8**(2), 250–254.
 - 22 S. L. Tait, Y. Wang, G. Costantini, N. Lin, A. Baraldi, F. Esch, L. Petaccia, S. Lizzit and K. Kern, Metal–organic coordination interactions in Fe–Terephthalic acid networks on Cu(100), *J. Am. Chem. Soc.*, 2008, **130**(6), 2108–2113.
 - 23 D. Skomski, C. D. Tempas, G. S. Bukowski, K. A. Smith and S. L. Tait, Redox-active on-surface polymerization of single-site divalent cations from pure metals by a ketone-functionalized phenanthroline, *J. Chem. Phys.*, 2015, **142**, 101913.
 - 24 H. H. Yang, Y. H. Chu, C. I. Lu, T. H. Yang, K. J. Yang, C. C. Kaun, G. Hoffmann and M. T. Lin, Digitized Charge Transfer Magnitude Determined by Metal–Organic Coordination Number, *ACS Nano*, 2013, **7**(3), 2814–2819.
 - 25 N. Abdurakhmanova, A. Floris, T. C. Tseng, A. Comisso, S. Stepanow, A. De Vita and K. Kern, Stereoselectivity and electrostatics in charge-transfer Mn- and Cs-TCNQ(4) networks on Ag(100), *Nat. Commun.*, 2012, **3**, 940.
 - 26 B. R. Takulapalli, G. M. Laws, P. A. Liddell, J. Andreasson, Z. Erno, D. Gust and T. J. Thornton, Electrical detection of amine ligation to a metalloporphyrin *via* a hybrid SOI-MOSFET, *J. Am. Chem. Soc.*, 2008, **130**(7), 2226–2233.
 - 27 D. Skomski, C. D. Tempas, K. A. Smith and S. L. Tait, Redox-Active On-Surface Assembly of Metal–Organic Chains with Single-Site Pt(II), *J. Am. Chem. Soc.*, 2014, **136**(28), 9862–9865.
 - 28 C. G. Williams, M. Wang, D. Skomski, C. D. Tempas, L. L. Kesmodel and S. L. Tait, Metal–Ligand Complexation through Redox Assembly at Surfaces Characterized by Vibrational Spectroscopy, *J. Phys. Chem. C*, 2017, **121**(24), 13183–13190.
 - 29 D. Le and T. S. Rahman, Pt–dipyridyl tetrazine metal–organic network on the Au(100) surface: insights from first principles calculations, *Faraday Discuss.*, 2017, **204**, 83–95.
 - 30 W. Hieringer, K. Flechtner, A. Kretschmann, K. Seufert, W. Auwärter, J. V. Barth, A. Görling, H. P. Steinrück and J. M. Gottfried, The Surface Trans Effect: Influence of Axial Ligands on the Surface Chemical Bonds of Adsorbed Metalloporphyrins, *J. Am. Chem. Soc.*, 2011, **133**(16), 6206–6222.



- 31 M. Feng, H. Sun, J. Zhao and H. Petek, Self-Catalyzed Carbon Dioxide Adsorption by Metal–Organic Chains on Gold Surfaces, *ACS Nano*, 2014, **8**(8), 8644–8652.
- 32 B. E. Murphy, S. A. Krasnikov, N. N. Sergeeva, A. A. Cafolla, A. B. Preobrajenski, A. N. Chaika, O. Lubben and I. V. Shvets, Homolytic Cleavage of Molecular Oxygen by Manganese Porphyrins Supported on Ag(111), *ACS Nano*, 2014, **8**(5), 5190–5198.
- 33 S. Fabris, S. Stepanow, N. Lin, P. Gambardella, A. Dmitriev, J. Honolka, S. Baroni and K. Kern, Oxygen Dissociation by Concerted Action of Di-Iron Centers in Metal–Organic Coordination Networks at Surfaces: Modeling Non-Heme Iron Enzymes, *Nano Lett.*, 2011, **11**(12), 5414–5420.
- 34 E. Amadio, R. Di Lorenzo, C. Zonta and G. Licini, Vanadium catalyzed aerobic carbon–carbon cleavage, *Coord. Chem. Rev.*, 2015, **301**, 147–162.
- 35 G. Licini, V. Conte, A. Coletti, M. Mba and C. Zonta, Recent advances in vanadium catalyzed oxygen transfer reactions, *Coord. Chem. Rev.*, 2011, **255**(19–20), 2345–2357.
- 36 L. Artiglia, S. Agnoli and G. Granozzi, Vanadium oxide nanostructures on another oxide: the viewpoint from model catalysts studies, *Coord. Chem. Rev.*, 2015, **301**, 106–122.
- 37 S. Beke, A review of the growth of V_2O_5 films from 1885 to 2010, *Thin Solid Films*, 2011, **519**(6), 1761–1771.
- 38 C. A. Carrero, R. Schloegl, I. E. Wachs and R. Schomaecker, Critical Literature Review of the Kinetics for the Oxidative Dehydrogenation of Propane over Well-Defined Supported Vanadium Oxide Catalysts, *ACS Catal.*, 2014, **4**(10), 3357–3380.
- 39 G. C. Bond and S. F. Tahir, Vanadium-Oxide Monolayer Catalysts – Preparation, Characterization and Catalytic Activity, *Appl. Catal.*, 1991, **71**(1), 1–31.
- 40 S. Surnev, G. Kresse, M. G. Ramsey and F. P. Netzer, Novel interface-mediated metastable oxide phases: vanadium oxides on Pd(111), *Phys. Rev. Lett.*, 2001, **87**(8), 086102.
- 41 M. Baumer, J. Biener and R. J. Madix, Growth, electronic properties and reactivity of vanadium deposited onto a thin alumina film, *Surf. Sci.*, 1999, **432**(3), 189–198.
- 42 R. J. Madix, J. Biener, M. Baumer and A. Dinger, The growth of vanadium oxide on alumina and titania single crystal surfaces, *Faraday Discuss.*, 1999, **114**, 67–84.
- 43 H. Niehus, R. P. Blum and D. Ahlbehrendt, Formation of vanadium oxide films on $Cu_3Au(100)$, *Phys. Status Solidi A*, 2001, **187**(1), 151–159.
- 44 K. B. Lewis, S. T. Oyama and G. A. Somorjai, The Preparation and Reactivity of Thin, Ordered Films of Vanadium-Oxide on Au(111), *Surf. Sci.*, 1990, **233**(1–2), 75–83.
- 45 M. Abu Haija, S. Guimond, Y. Romanyshyn, A. Uhl, H. Kuhlenbeck, T. K. Todorova, M. V. Ganduglia-Pirovano, J. Dobler, J. Sauer and H. J. Freund, Low temperature adsorption of oxygen on reduced $V_2O_3(0001)$ surfaces, *Surf. Sci.*, 2006, **600**(7), 1497–1503.
- 46 F. Stavale, H. Niehus and C. A. Achete, Ultra thin V_2O_3 films grown on oxidized Si(111), *Surf. Sci.*, 2009, **603**(17), 2721–2724.
- 47 G. Krenn, J. Schoiswohl, S. Surnev, F. P. Netzer and R. Schennach, Metal-oxide boundary effects in vanadium oxide–Rh(111) inverse model catalysts: a RAIRS, STM and TPD study, *Top. Catal.*, 2007, **46**(1–2), 231–238.
- 48 B. Sedai, C. Diaz-Urrutia, R. T. Baker, R. L. Wu, L. A. Silks and S. K. Hanson, Comparison of Copper and Vanadium Homogeneous Catalysts for Aerobic Oxidation of Lignin Models, *ACS Catal.*, 2011, **1**(7), 794–804.
- 49 I. E. Wachs, Catalysis science of supported vanadium oxide catalysts, *Dalton Trans.*, 2013, **42**(33), 11762–11769.
- 50 T. Z. Huang, J. M. Yu, J. T. Han, Z. L. Zhang, Y. Xing, C. L. Wen, X. Y. Wu and Y. H. Zhang, Oxygen reduction catalytic characteristics of vanadium carbide and nitrogen doped vanadium carbide, *J. Power Sources*, 2015, **300**, 483–490.
- 51 K. Eguchi, T. Nakagawa, Y. Takagi and T. Yokoyama, Direct Synthesis of Vanadium Phthalocyanine and Its Electronic and Magnetic States in Mono layers and Multi layers on Ag(111), *J. Phys. Chem. C*, 2015, **119**(18), 9805–9815.
- 52 L. A. Rochford, I. Hancox and T. S. Jones, Understanding domain symmetry in vanadium oxide phthalocyanine monolayers on Au(111), *Surf. Sci.*, 2014, **628**, 62–65.
- 53 H. Adler, M. Paszkiewicz, J. Uihlein, M. Polek, R. Ovsyannikov, T. V. Basova, T. Chasse and H. Peisert, Interface Properties of VOPc on Ni(111) and Graphene/Ni(111): Orientation-Dependent Charge Transfer, *J. Phys. Chem. C*, 2015, **119**(16), 8755–8762.
- 54 D. Skomski, C. D. Tempas, B. J. Cook, A. V. Polezhaev, K. A. Smith, K. G. Caulton and S. L. Tait, Two- and Three-Electron Oxidation of Single-Site Vanadium Centers at Surfaces by Ligand Design, *J. Am. Chem. Soc.*, 2015, **137**(24), 7898–7902.
- 55 I. L. Botto, M. B. Vassallo, E. J. Baran and G. Minelli, IR spectra of VO_2 and V_2O_3 , *Mater. Chem. Phys.*, 1997, **50**(3), 267–270.
- 56 S. F. Vyboishchikov and J. Sauer, $(V_2O_5)_n$ Gas-Phase Clusters ($n=1-12$) Compared to V_2O_5 Crystal: DFT Calculations, *J. Phys. Chem. A*, 2001, **105**(37), 8588–8598.
- 57 M. A. Vanhove, R. J. Koestner, P. C. Stair, J. P. Bibérian, L. L. Kesmodel, I. Bartoš and G. A. Somorjai, The Surface Reconstructions of the (100) Crystal Faces of Iridium, Platinum and Gold. 1. Experimental-Observations and Possible Structural Models, *Surf. Sci.*, 1981, **103**(1), 189–217.
- 58 C. D. Tempas, D. Skomski and S. L. Tait, Lifting of the Au(100) surface reconstruction by Pt, Cr, Fe, and Cu adsorption, *Surf. Sci.*, 2016, **654**, 33–38.
- 59 G. K. Binnig, H. Rohrer, C. Gerber and E. Stoll, Real-Space Observation of the Reconstruction of Au(100), *Surf. Sci.*, 1984, **144**(2–3), 321–335.
- 60 L. L. Kesmodel, High-resolution Electron Energy Loss Spectroscopy, in *Encyclopedia of Surface and Colloid Science*, ed. P. Somasundaren, Taylor and Francis, New York, 2nd edn, 2006.
- 61 H. Ibach, *Physics of Surfaces and Interfaces*, Springer, Berlin, 2006.
- 62 Y. Shigesato, M. Enomoto and H. Odaka, Thermochromic VO_2 films deposited by RF magnetron sputtering using V_2O_3 or V_2O_5 targets, *Jpn. J. Appl. Phys.*, 2000, **39**(10), 6016–6024.



- 63 W. G. Menezes, D. M. Reis, T. M. Benedetti, M. M. Oliveira, J. F. Soares, R. M. Torresi and A. J. Zarbin, V_2O_5 nanoparticles obtained from a synthetic bariandite-like vanadium oxide: synthesis, characterization and electrochemical behavior in an ionic liquid, *J. Colloid Interface Sci.*, 2009, **337**(2), 586–593.
- 64 T. W. Hansen, A. T. Delariva, S. R. Challa and A. K. Datye, Sintering of Catalytic Nanoparticles: Particle Migration or Ostwald Ripening?, *Acc. Chem. Res.*, 2013, **46**(8), 1720–1730.
- 65 F. P. Leisenberger, S. Surnev, L. Vitali, M. G. Ramsey and F. P. Netzer, Nature, growth, and stability of vanadium oxides on Pd(111), *J. Vac. Sci. Technol., A*, 1999, **17**(4), 1743–1749.
- 66 D. Le and T. S. Rahman, Pt–Dipyridyl Tetrazine Metal–Organic Network on the Au(100) Surface: Insights from First Principles Calculations, *Faraday Discuss.*, 2017, **204**, 83–95.
- 67 G. Silversmit, D. Depla, H. Poelman, G. B. Marin and R. De Gryse, Determination of the V 2p XPS binding energies for different vanadium oxidation states (V^{5+} to V^{0+}), *J. Electron Spectrosc. Relat. Phenom.*, 2004, **135**(2–3), 167–175.
- 68 C. D. Wagner, L. E. Davis, M. V. Zeller, J. A. Taylor, R. H. Raymond and L. H. Gale, Empirical Atomic Sensitivity Factors for Quantitative-Analysis by Electron-Spectroscopy for Chemical-Analysis, *Surf. Interface Anal.*, 1981, **3**(5), 211–225.
- 69 R. J. Colton, A. M. Guzman and J. W. Rabalais, Electrochromism in Some Thin-Film Transition-Metal Oxides Characterized by X-Ray Electron-Spectroscopy, *J. Appl. Phys.*, 1978, **49**(1), 409–416.
- 70 T. K. A. Hoang, M. I. Webb, H. V. Mai, A. Hamaed, C. J. Walsby, M. Trudeau and D. M. Antonelli, Design and Synthesis of Vanadium Hydrazide Gels for Kubas-Type Hydrogen Adsorption: A New Class of Hydrogen Storage Materials, *J. Am. Chem. Soc.*, 2010, **132**(33), 11792–11798.
- 71 X. Feng, M. K. Abdel-Rahman, C. M. Kruppe and M. Trenary, Deposition and characterization of stoichiometric films of V_2O_5 on Pd(111), *Surf. Sci.*, 2017, **664**, 1–7.
- 72 M. Sutradhar, T. R. Barman, G. Mukherjee, M. G. B. Drew and S. Ghosh, Synthesis, structural characterization and electrochemical activity of oxidovanadium(IV/V) complexes of a diprotic ONS chelating ligand, *Inorg. Chim. Acta*, 2010, **363**(13), 3376–3383.
- 73 S. B. Gajera, J. V. Mehta, D. N. Kanthecha, R. R. Patel and M. N. Patel, Novel cytotoxic oxovanadium (IV) complexes: Influence of pyrazole-incorporated heterocyclic scaffolds on their biological response, *Appl. Organomet. Chem.*, 2017, **31**, e3767.
- 74 K. Yamashita, S. Tazawa and K. Sugiura, Oxo(porphyrinato) vanadium(IV) as a standard for geoporphyrins, *Inorg. Chim. Acta*, 2016, **439**, 173–177.
- 75 B. K. Min, A. R. Alemozafar, D. Pinnaduwege, X. Deng and C. M. Friend, Efficient CO oxidation at low temperature on Au(111), *J. Phys. Chem. B*, 2006, **110**(40), 19833–19838.
- 76 F. Hiebel, M. M. Montemore, E. Kaxiras and C. M. Friend, Direct visualization of quasi-ordered oxygen chain structures on Au(110)-(1 × 2), *Surf. Sci.*, 2016, **650**, 5–10.
- 77 X. Y. Deng, B. K. Min, A. Guloy and C. M. Friend, Enhancement of O₂ dissociation on Au(111) by adsorbed oxygen: implications for oxidation catalysis, *J. Am. Chem. Soc.*, 2005, **127**(25), 9267–9270.
- 78 B. K. Min, A. R. Alemozafar, M. M. Biener, J. Biener and C. M. Friend, Reaction of Au(111) with sulfur and oxygen: scanning tunneling microscopic study, *Top. Catal.*, 2005, **36**(1–4), 77–90.
- 79 J. R. Khusnutdinova and D. Milstein, Metal–Ligand Cooperation, *Angew. Chem., Int. Ed.*, 2015, **54**(42), 12236–12273.
- 80 G. Antczak, K. Boom and K. Morgenstern, Revealing the Presence of Mobile Molecules on the Surface, *J. Phys. Chem. C*, 2017, **121**(1), 542–549.
- 81 A. Klein, E. J. L. McInnes, T. Scheiring and S. Zalis, Electronic structure of radical anionic binuclear organoplatinum complexes. a multiple frequency EPR investigation, *J. Chem. Soc., Faraday Trans.*, 1998, **94**(19), 2979–2984.
- 82 E. C. Constable, C. E. Housecroft, B. M. Kariuki, N. Kelly and C. B. Smith, A near planar disilver complex of 3,6-bis(2-pyridyl)-1,2,4,5-tetrazine, *Inorg. Chem. Commun.*, 2002, **5**(3), 199–202.
- 83 F. R. Dai, H. Y. Ye, B. Li, L. Y. Zhang and Z. N. Chen, Low-valence oxo-centred triruthenium complexes by bridging acetate substitution with pyrazolyldiazine or pyridinyltetrazine ligands, *Dalton Trans.*, 2009, **40**, 8696–8703.
- 84 B. L. Schottel, H. T. Chifotides, M. Shatruck, A. Chouai, L. M. Perez, J. Bacsá and K. R. Dunbar, Anion– π interactions as controlling elements in self-assembly reactions of Ag(I) complexes with π -acidic aromatic rings, *J. Am. Chem. Soc.*, 2006, **128**(17), 5895–5912.
- 85 E. C. Constable, C. E. Housecroft, M. Neuburger, S. Reymann and S. Schaffner, A new twist to 3,6-bis(2-pyridyl)-1,2,4,5-tetrazine complexes of silver(I), *CrystEngComm*, 2008, **10**(8), 991–995.
- 86 M. Maekawa, T. Miyazaki, K. Sugimoto, T. Okubo, T. Kuroda-Sowa, M. Munakata and S. Kitagawa, Structural diversity among copper(I) ethylene adducts of 3,6-bis(2-pyridyl)-1,2,4,5-tetrazine, *Dalton Trans.*, 2013, **42**(12), 4258–4266.
- 87 K. H. Thompson and C. Orvig, Coordination chemistry of vanadium in metallopharmaceutical candidate compounds, *Coord. Chem. Rev.*, 2001, **219**, 1033–1053.
- 88 D. Nicholls, The Coordination Chemistry of Vanadium, *Coord. Chem. Rev.*, 1966, **1**, 379–414.

