




Cite this: *RSC Adv.*, 2018, 8, 39937

# A novel porous carbon material derived from the byproducts of bean curd stick manufacture for high-performance supercapacitor use†

Lulu Shi,<sup>a</sup> Lihua Jin, \*<sup>a</sup> Zheng Meng,<sup>a</sup> Yanan Sun,<sup>a</sup> Cong Li<sup>b</sup> and Yehua Shen<sup>a</sup>

The exploitation of efficient renewable energy resources and the promotion of added value of agricultural products are always hot topics. In this study, we present a novel porous carbon material for high performance supercapacitor applications made from the byproduct of bean curd stick manufacture. The as-prepared carbon material possesses a hollow interconnected structure with a large specific surface area (2609 m<sup>2</sup> g<sup>-1</sup>), while containing 5.01% oxygen and 1.75% nitrogen heteroatoms. Therefore, besides electrical double-layer capacitance, it can also produce additional pseudocapacitance to enhance the overall capacitance. Benefiting from these advantages of structure and composition, the bean curd stick based porous carbon material demonstrates a high specific capacitance of 405 F g<sup>-1</sup> at 0.5 A g<sup>-1</sup>. Moreover, the presented porous carbon material based symmetric supercapacitor also offers a high energy density (11.35 W h kg<sup>-1</sup> at a power density of 125 W kg<sup>-1</sup>). All the above findings indicate that the byproduct of bean curd stick manufacture is an excellent optional material for preparing high-performance supercapacitor material. Simultaneously, this work also provides an effective strategy for adding value to agricultural products.

Received 19th October 2018  
 Accepted 15th November 2018

DOI: 10.1039/c8ra08664h

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## 1. Introduction

In recent years, owing to the growing demands for energy and the exhausting of traditional fossil energy, the exploitation of efficient renewable energy resources has become extremely urgent and important.<sup>1–4</sup> Scholars and researchers around the world are trying to explore and develop biomass energy to replace fossil energy. Supercapacitors, as electrochemical energy storage devices, have been a hot research topic because of their excellent electrochemical properties, such as high power density, fast charge/discharge rate and excellent cycle stability.<sup>5,6</sup> In the light of the energy storage mechanism of capacitors, supercapacitors can be roughly divided into two categories: pseudocapacitors and electrical double-layer capacitors (EDLCs).<sup>7</sup> Conventionally, the energy storage of pseudocapacitors is primarily dominated by the quasi Faraday capacitance produced by redox reaction,<sup>8,9</sup> while the capacitance generation of EDLCs generally arises from the behavior of electrostatic charge diffusion and accumulation at the

interface of electrode/electrolyte.<sup>10–13</sup> Contrastingly, pseudocapacitors possess high capacitance, and EDLCs exhibit a higher power density and more excellent cyclic performance. There are merits to both the two supercapacitor types. Therefore, the exploration and development of a comprehensive supercapacitor that allows the two types to intermingle without losing their separate merits is a marvelous way to get a higher-performance supercapacitor.

As we all know, the characteristics of electrode materials play an important part in the performance of a supercapacitor. Therefore, a tremendous amount of effort has been focused on developing innovative supercapacitor materials with high capacitance and excellent cycle stability. Biomass-derived porous carbon materials are a typical promising candidate for supercapacitors.<sup>14,15</sup> They not only overcome non-renewable/high cost disadvantages as compared with graphene or carbon nanotubes, but also possess advantages including large specific surface area and high energy storage capability.<sup>10,16</sup> Up to now, various biomass-based porous carbon materials derived from leaves,<sup>17,18</sup> bamboo,<sup>19–22</sup> egg yolk,<sup>23</sup> fungus<sup>24–26</sup> and bean products<sup>27–30</sup> have been widely used in the electrodes of supercapacitors. However, most of the existing biomass-based electrode materials have the characteristics of low capacitance and poor cycle stability. Therefore, it is still important to explore simple synthesis processes to develop different carbon electrode materials with excellent supercapacitor properties.

Bean curd stick is a traditional Chinese food that is made from soya beans. Weight for weight it contains more protein

<sup>a</sup>Key Laboratory of Synthetic and Natural Functional Molecule Chemistry of the Ministry of Education, College of Chemistry & Materials Science, Northwest University, Xi'an 710127, People's Republic of China. E-mail: jinlihua@mwu.edu.cn; Fax: +86-029-88302635

<sup>b</sup>National Demonstration Center for Experimental Chemistry Education, Northwest University, Xi'an 710127, People's Republic of China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c8ra08664h



than soya bean, bean curd, or bean dregs; its protein content can even reach 45–50%. Bean curd stick is usually processed into dried strip products, and during the manufacturing process, byproducts are usually produced. The tons of slag material left over from bean curd stick manufacture are at present discharged as agro-industrial waste or used as animal fodder. Therefore, the exploration of a feasible approach to further improve the added value of bean curd stick manufacture will obviously increase economic efficiency of the processing plant. Referring to previous studies,<sup>23,26</sup> doping porous carbon materials with nitrogen atoms can remarkably improve their capacitive behavior. And an effective and feasible method to fabricate nitrogen-functionalized porous carbon materials is to directly carbonize high protein content biomass, such that the nitrogen is incorporated firmly in the carbon frame. Therefore, all the bean byproducts that are rich in C and N elements are considered to be good candidates for making high added value heteroatom-doped carbon materials. Recently, many active carbon based electrode materials for supercapacitors derived from soya bean, bean curd, or bean dregs have been made. For example, Ma and his co-workers synthesized a nitrogen-doped porous carbon from bean dregs which could be used as electrode material for supercapacitors.<sup>27</sup> Cao and his group also prepared a kind of one-dimensional (1D) carbon nanobelt from bean curd for high specific capacitance (262 F g<sup>-1</sup> at 0.5 A g<sup>-1</sup>) supercapacitors.<sup>28</sup> Recently, Wang and his group also reported functionalized porous carbon from soybean root for supercapacitors.<sup>29</sup> All the mentioned researches clearly demonstrated that bean products are promising carbonaceous raw material for supercapacitor electrodes. However, until now, no relative reports have appeared that study the application of bean curd stick in supercapacitor materials. Judging purely by the content of available components, such as the content of protein, bean curd stick based supercapacitor electrode materials are suspected to possess great performance as supercapacitors because of their abundant protein content, which can provide much more effective heteroatom doping and additional pseudocapacitance from the redox reaction.<sup>15,31</sup>

Herein, enlightened by this conjecture, we present a facile method to synthesize heteroatom-doped interconnected porous carbon electrode material by using the byproducts of bean curd stick biomass as carbon precursor and nitrogen source. Through a process of hydrothermal carbonization, pre-carbonization and activation with KOH, active carbon materials were synthesized. It was no accident that the selected as-prepared porous carbon material exhibited a well-developed interconnected pore structure with a specific surface area as high as 2609 m<sup>2</sup> g<sup>-1</sup>. Moreover, as an electrode material, it also exhibited fine capacitive performance with relatively high specific capacitance (405 F g<sup>-1</sup> at a current density of 0.5 A g<sup>-1</sup>), favorable rate performance and good cycling stability. Generally, the electrochemical performance of the present porous carbon material is much better than that of the existing commercial porous carbons and most other nitrogen-doped carbon materials.

## 2. Experimental section

### 2.1 Materials

Fresh slag material from bean curd stick was purchased from a local bean curd stick food processing workshop (Yuan Village, Xianyang, China). Acetone, ethanol and potassium hydroxide were bought from Sinopharm Chemical Reagent Co. Ltd (China). Polytetrafluoroethylene (PTFE) solution (60%) and acetylene black were obtained from Daikin, Japan Co., Ltd (Japan). NKK-MPF30AC-100 aqueous separator was provided by NKK, Japan Co., Ltd (Japan). Nickel foam (1.5 cm × 1 cm) was obtained from Changde Lyrun Material Co., Ltd (Hunan, China).

### 2.2 Preparation of bean curd stick derived porous carbon materials

In a typical synthesis, the fresh slag material from bean curd stick was firstly pounded into a fine powder in a mortar. Subsequently, 12 g bean curd stick powder and 60 mL deionized water were put into a Teflon-lined autoclave for hydrothermal treatment at 200 °C for 6 h. A brown hydrothermal product was collected and freeze-dried at -40 °C. After that, the obtained product was activated with KOH through resistance furnace heating. The compound was first pyrolyzed in a tubular furnace (400 °C for 1 hour, 700 °C for another 1 hour; heating rate of 10 °C min<sup>-1</sup>; N<sub>2</sub> atmosphere). Finally, the samples were centrifuged with 1 M hydrochloric acid and distilled water until they reached neutral pH. The residue was collected and dried in a vacuum drying chamber at 80 °C. For control experiments, different carbonization temperatures (600 °C, 700 °C, 800 °C) and different weight ratios between carbonaceous samples and KOH (1 : 0, 1 : 1, 1 : 2, 1 : 3) were investigated, and different products, named 700-0, 700-1, 700-2, 700-3, 600-3, 800-3, were obtained.

### 2.3 Characterization

The micromorphology of materials was characterized by scanning electron microscopy (SEM; SU8010, Hitachi Co.) and transmission electron microscopy (TEM; FEI Tecnai G2 F20). The crystal structure was determined by X-ray diffraction (XRD; AXS-D8, Bruker). The Raman spectra were analyzed using a laser Raman spectrometer (DXR2, Thermo Electron Co.). The surface elemental composition was examined by X-ray photoelectron spectroscopy (XPS; AXIS ULTRA). N<sub>2</sub> adsorption-desorption isotherms were obtained at 77 K to assess specific surface area and porosity (Tri Star-3020, Micromeritics Instrument Co.). Electrochemical testing was performed on a CHI 660D electrochemical workstation (Chenhua).

### 2.4 Electrochemical measurements

The working electrode of the bean curd stick based porous carbon material (C-BCS) for supercapacitor devices was prepared as follows: a weight ratio of 85 : 10 : 5 mixture of carbon material, acetylene black and PTFE binder was pressed on a nickel foam current collector. The loaded electrode was



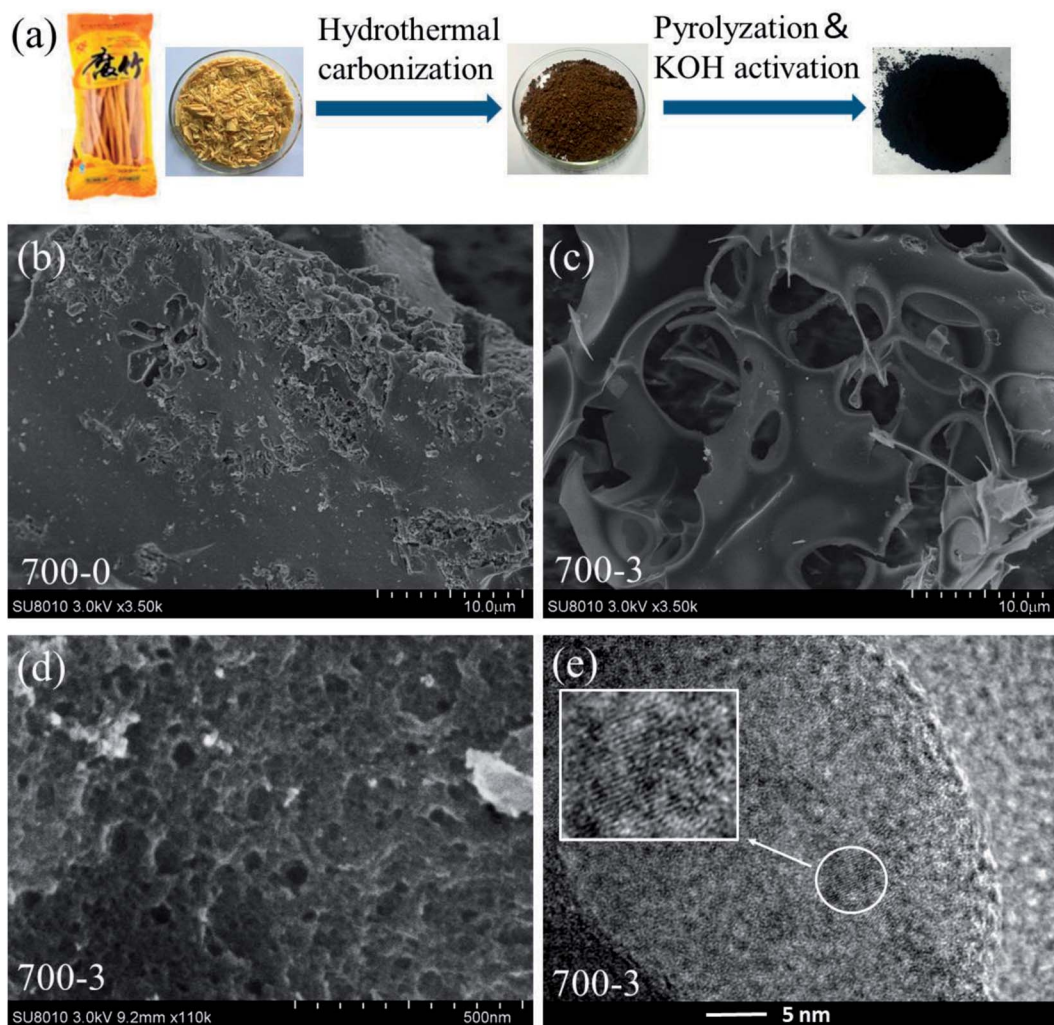


Fig. 1 (a) Schematic diagram of the synthesis route of bean curd stick derived porous carbon; and SEM images of (b) the 700-0 sample, and (c and d) the 700-3 sample at different magnifications; (e) TEM image of the 700-3 sample.

dried at 80 °C overnight and pressed under 10 MPa for 30 seconds; the mass loading of active material was *ca.* 2 mg cm<sup>-1</sup>. The electrochemical measurements, such as cyclic voltammetry tests (CV), galvanostatic charge/discharge measurements (GCD) and electrochemical impedance spectroscopy tests (EIS), were performed in 6 M KOH electrolyte. The specific capacitance ( $C_g$ ) of the present electrode in a three-electrode system was calculated from the charge/discharge curve by using eqn (1). In the 700-3 based symmetrical supercapacitor device, the two electrodes were separated by an MPF30AC-100 membrane which was soaked with 6 M KOH electrolyte. The specific capacitance ( $C_{cell}$ , F g<sup>-1</sup>), specific energy density ( $E$ , W h kg<sup>-1</sup>) and specific power density ( $P$ , W kg<sup>-1</sup>) of the device were calculated by the following eqn (2)–(4), respectively.

$$C_g = \frac{I \times \Delta t}{m \times \Delta V} \quad (1)$$

$$C_{cell} = \frac{I \times \Delta t}{M \times \Delta V} \quad (2)$$

$$E = \frac{1}{7.2} \times C_{cell} \times (\Delta V)^2 \quad (3)$$

$$P = \frac{3600 \times E}{\Delta t} \quad (4)$$

where  $I$  defines the applied current,  $\Delta t$  is the discharge time,  $m$  represents the loading mass of electrode material on the single electrode,  $\Delta V$  refers to the voltage window, and  $M$  is the total mass of electrode material on the two electrodes, respectively.

### 3. Results and discussion

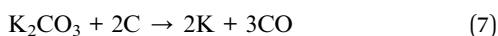
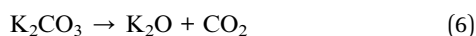
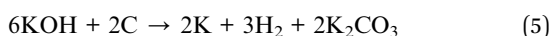
#### 3.1 Material characterization

The detailed preparation process of the bean curd stick based porous carbon materials (C-BCS) is described in Fig. 1a. There were three main steps: hydrothermal carbonization (HTC), pre-carbonization and KOH chemical activation. Bean curd stick is a smooth, dry product, and it is difficult to directly and efficiently make it interact with KOH to obtain a porous carbon



structure. Therefore, we adopted hydrothermal carbonization in the synthesis procedure in order to get a better product with higher throughput and specific surface area. Besides that, owing to the high protein content of bean curd stick, the hydrothermal product exhibited absorbed moisture, and could not be directly dried in a simple oven (as shown in the Fig. S1†). Thus, we chose to use freeze-drying technology. It not only effectively dries the sample, but also leads to the formation of a foamy powder structure, which is expected to be beneficial to the formation of more pore structures and increase the specific surface area. Moreover, it should also be noted that both the KOH activation and the pyrolyzing temperature play a critical role in the generation of the pore structure during the procedure. Therefore, we systematically investigated the influence of the amount of KOH added and the activation temperature on the formation of pores. As shown in Fig. S2,† the C-BCS gradually turns into 3D interconnected pore structure with the increase of KOH activator content. Considering the electrochemical and structural properties of the synthesized samples, the C-BCS that was produced using an activation weight ratio (carbon : KOH) of 1 : 3 with carbonization at 700 °C (700-3) was selected as the optimal sample for further supercapacitor applications.

Comparing the changes of morphology of the prepared C-BCS before and after KOH activation, it could be clearly observed that hydrothermal carbonization treated carbon materials without KOH activation manifested their massive structure, and only some small holes existed on the surface (Fig. 1b). But after being treated with KOH in a weight ratio of 1 : 3, plenty of 3D architecture with interconnected macropores and cavities (in the range of sub-micrometer to a few micrometers) were discovered (Fig. 1c and d). The presence of such structure might be ascribed to pore-forming by KOH etching and the emission of H<sub>2</sub>, CO and CO<sub>2</sub> from the chemical reactions at high temperature (>400 °C) as follows:<sup>14,15,32,33</sup>



The formation of the porous carbon material is suspected to be helpful for full access of electrolyte ions to the charge storage sites and greatly improve its supercapacitor performance.<sup>11</sup> High resolution TEM (HRTEM) imaging further confirmed the existence of partial graphitization on the surface of the 700-3 porous carbon material by the distorted lattice fringe morphology (Fig. 1e). According to previous reports, the degree of graphitization is another crucial indicator for excellent capacitive performance, which suggests that the porous carbon materials possess great electrical conductivity, and is highly beneficial to the charge transfer.<sup>22</sup>

Moreover, the degree of carbonization and graphitization of the 700-3 sample was also characterized by the XRD pattern and Raman spectroscopy. As shown in Fig. 2a, well-developed diffraction of plane (002) and weak diffraction of plane (100)

of the disordered carbon layer can be observed, indicating the formation of a certain degree of graphite crystallite structure.<sup>22,34</sup> In the Raman spectrum (Fig. 2b), two bands of carbon materials ascribed to the D (defects and disorder) band (1348 cm<sup>-1</sup>) and G (graphitic) band (1584 cm<sup>-1</sup>) can also be observed,<sup>35</sup> and the smaller the intensity ratio (*I<sub>D</sub>/I<sub>G</sub>*) of the D and G bands, the higher the degree of graphitization of carbon material. In our case, the intensity ratio (*I<sub>D</sub>/I<sub>G</sub>*) of 700-3 porous carbon materials was calculated to be 0.88, which is comparable to or much lower than many other biomass-derived porous carbon materials, such as those from kiwifruit (0.93),<sup>36</sup> bean dregs (0.97),<sup>27</sup> waste coffee grounds (1.00),<sup>37</sup> bamboo-based byproduct (1.01),<sup>22</sup> auricularia (1.02),<sup>26</sup> Jujun grass (1.03)<sup>38</sup> and human hair (1.23).<sup>39</sup>

In order to investigate the porosity of our prepared porous carbon materials, the nitrogen adsorption-desorption isotherms and pore size distributions were obtained at 77 K. As displayed in Fig. 2c, the curve shows typical type-IV isotherms for the 700-3 porous carbon material, with a type-H2 hysteresis loop, implying the existence of both microporous and mesoporous structures in this material. Furthermore, we also used density function theory (DFT) to analyze the pore size distribution of the porous carbon materials. From Fig. 2d, we find that the 700-3 porous carbon material sample is mainly composed of mesopores with pore size 2–3 nm, while the pore volume is calculated to be about 1.53 cm<sup>3</sup> g<sup>-1</sup>. The specific surface area of that material even reached 2609 m<sup>2</sup> g<sup>-1</sup>, which is much larger than for previously reported porous carbon materials.<sup>22,28,40,41</sup> Considering that both the activation temperature and addition amount of KOH can significantly affect the microstructure of the carbon material, the porosity structures of the obtained 700-1, 700-2, 600-3 and 800-3 samples were also investigated. As shown and listed in Fig. S3† and Table 1, all the samples possessed mesoporous structures with an average pore size of 2–3 nm; when the addition amount of KOH increased from 1 : 1 to 1 : 3, both the specific surface area and pore volume increased gradually. The specific surface areas were 1252, 2609 and 2262 m<sup>2</sup> g<sup>-1</sup> for 600-3, 700-3 and 800-3, respectively, as calculated by the Brunauer-Emmett-Teller (BET) method. The lower surface area at 800 °C might be due to degradation of the carbon skeleton at such a high temperature.<sup>36</sup> The well-developed mesoporous structure of the present carbon materials greatly improved the specific surface area and provided a high density of ion adsorption sites as well as interconnected ion diffusion pathways, which may have an enormous impact on improving charge storage and ion diffusion kinetics.<sup>42</sup> Therefore, the porous carbon material defined as 700-3 is expected to be the ideal electrode material in supercapacitors, with the best capacitive performance.

The surface composition of the 700-3 porous carbon material was subsequently investigated by XPS analysis. As shown in Fig. 3a, three peaks ascribed to C 1s (284.9 eV), N 1s (399.9 eV) and O 1s (531.9 eV) in the XPS survey spectrum can be observed. The XPS elemental analysis data suggest that the material contains 93.24 at% carbon, 5.01 at% oxygen and a small quantity of nitrogen (1.75 at%). From the high-resolution elemental spectrum (Fig. 3b), we find that the C 1s spectrum



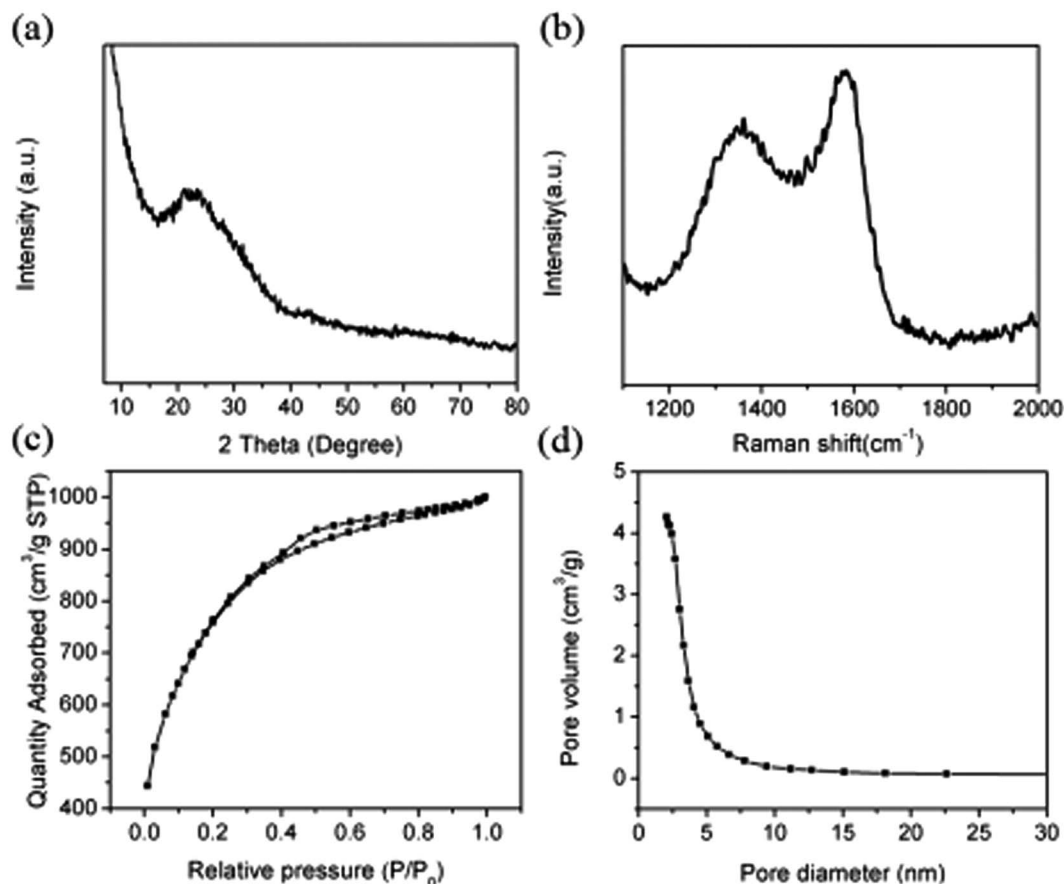


Fig. 2 (a) XRD data, (b) Raman spectrum, (c)  $N_2$  sorption-desorption isotherm and (d) pore size distribution of 700-3 sample.

can be divided into four separate peaks respectively attributed to  $sp^2$  C=C (284.3 eV),  $sp^3$  C-C (284.8 eV), C-O&C-N (286 eV) and C=O (289 eV),<sup>43-45</sup> while the O 1s spectrum (Fig. 3c) can be fitted with two peaks that represent C=O (531.6 eV) and C-O (533.0 eV).<sup>34</sup> Moreover, the nitrogen present in the prepared materials can be resolved into three species (Fig. 3d): pyridinic-N (398.4 eV, 6.02%), pyrrolic-N (399.94 eV, 63.17%) and quaternary-N (400.99 eV, 30.81%).<sup>46,47</sup> It has been reported that pyrrolic-N groups in a material can participate in the faradaic reactions and be a major nitrogen configuration contributing to pseudocapacitance. Meanwhile, the quaternary-N is believed to effectively improve the wettability of electrode material by electrolyte.<sup>43,48-50</sup> Comparing the content of different kinds of N species in the C-BCS (Table 2), it is found that the sample 700-3

possesses the highest relative content of pyrrolic-N. In this way, the electrochemical performance of the 700-3 sample based electrode, such as pseudocapacitance, rate performance and cycling performance, are expected to be significantly increased and improved.

### 3.2 Electrochemical performance

Because of its hollow interconnected porous structure and abundant nitrogen- and oxygen-containing functional groups, the prepared 700-3 porous carbon sample is considered to be a suitable candidate for high performance supercapacitor electrode material. Different electrochemical performance measures, such as CV test, GCD measurement and EIS test, were performed in the three-electrode system. Fig. 4a displays the

Table 1 Characteristics of pores in C-BCS samples<sup>a</sup>

Sample	$S_{BET}$ ( $m^2 g^{-1}$ )	$S_{mic}$ ( $m^2 g^{-1}$ )	$S_{meso}$ ( $m^2 g^{-1}$ )	$S_{meso}/S_{BET}$ (%)	$V_T$ ( $cm^3 g^{-1}$ )	$V_{mic}$ ( $cm^3 g^{-1}$ )	$V_{meso}$ ( $cm^3 g^{-1}$ )	$V_{meso}/V_T$ (%)	Pore size (nm)
700-1	1327	580	747	56.29	0.73	0.28	0.45	61.64	2.2
700-2	1666	570	1096	65.79	0.87	0.27	0.60	68.97	2.09
700-3	2609	847	1762	67.54	1.53	0.13	1.40	91.50	2.35
600-3	1252	413	839	67.01	0.71	0.13	0.58	81.69	2.26
800-3	2262	958	1304	57.65	1.84	0.13	1.71	92.93	3.25

<sup>a</sup>  $S_{BET}$ : specific surface area from the BET method;  $V_T$ : the total pore volume (single point adsorption at  $P/P_0 = 0.97$ );  $S_{mic}$  or  $S_{meso}$ : micropore or mesopore surface area;  $V_{mic}$  or  $V_{meso}$ : micropore or mesopore volume.



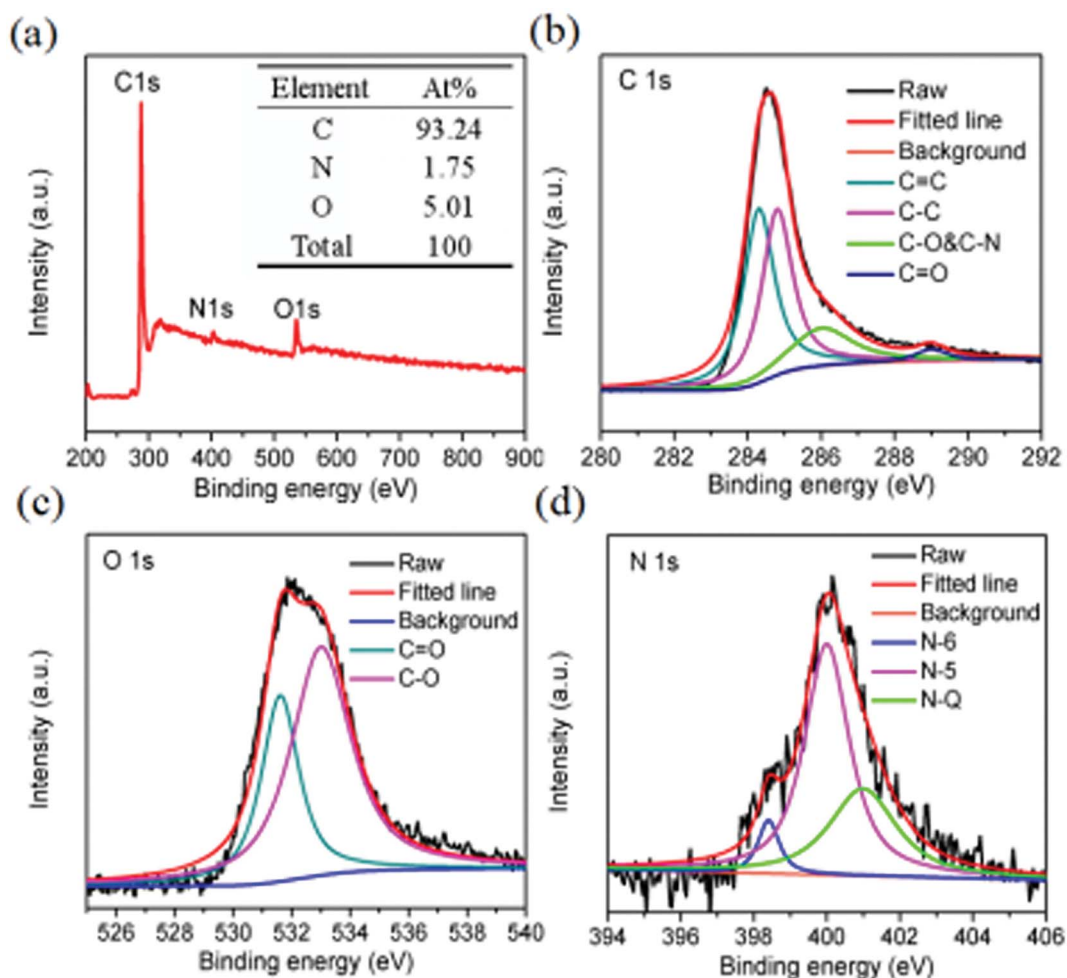


Fig. 3 (a) XPS survey spectrum and element ratio of 700-3 porous carbon material. (b)–(d) High resolution XPS spectra of C 1s, O 1s and N 1s.

Table 2 The content of different kinds of N species in C-BCS calculated by XPS analysis

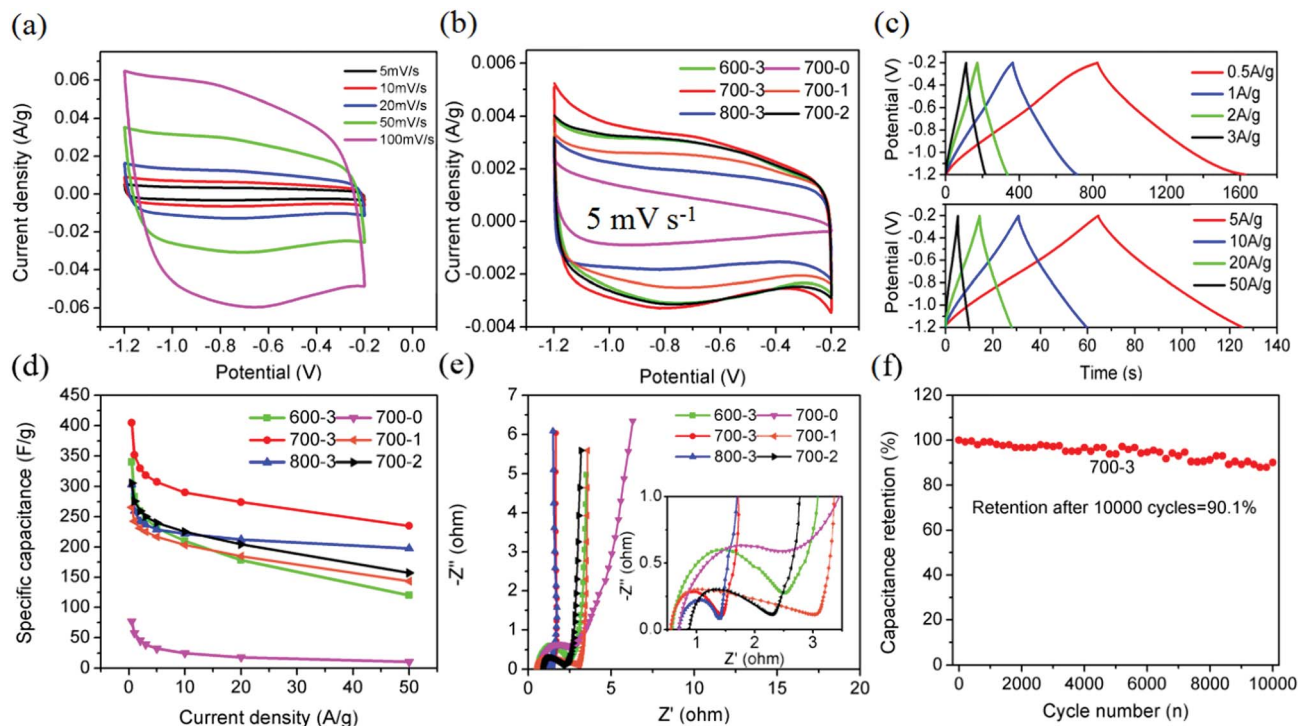
Sample	Pyridinic-N at%	Pyrrolic-N at%	Quaternary-N at%
700-0	36.95	30.21	32.84
700-1	14.34	48.84	36.82
700-2	12.23	55.86	31.91
700-3	6.02	63.17	30.81
600-3	20.38	47.55	32.07
800-3	—	—	—

variation of the CV curves of the 700-3 sample with gradually changing scan rates. The shape of the CV curves maintain almost symmetric rectangular shape over a wide range of scan rates, indicating the prepared 700-3 porous carbon materials exhibited excellent electrochemical capacitive behavior. At relatively high scan rates, a small pair of redox peaks on the CV curve of the 700-3 porous carbon material can be observed, demonstrating that the total capacitance of the material is the result of a combination of double layer capacitance and pseudocapacitance. As reported in previous literature,<sup>18,26,27</sup> the

functional groups present on the 700-3 sample surface, such as C–N, C=N and N–O, can effectively enhance the pseudocapacitance in aqueous electrolytes. Significantly, as compared with representative CV curves of other prepared carbon materials at the same scan rate of 50 mV s<sup>-1</sup>, the 700-3 porous carbon material exhibits the biggest CV area among them (Fig. 4b). This phenomenon corresponds to the largest specific capacitance, which is associated with the high specific surface area, perfect mesoporous structure and suitable heteroatom doping factors.<sup>51</sup>

Moreover, the charge–discharge behavior of the 700-3 porous carbon material was also investigated at various current densities. As shown in Fig. 4c, many highly linear and triangular-shaped charge–discharge profiles can be obtained, and almost no obvious IR drop in the GCD curve can be observed, which confirms that this material possesses high electrochemical reversibility and coulombic efficiency. The gravimetric specific capacitance of the 700-3 sample calculated from the discharge curves was 405, 351.5, 330 and 318.6 F g<sup>-1</sup>, corresponding to a current density of 0.5, 1, 2 and 3 A g<sup>-1</sup>, respectively. Also, even at the relatively high current density of 20 A g<sup>-1</sup>, the specific capacitance was maintained at 274 F g<sup>-1</sup> (retained 67.65% of its initial capacitance, Fig. 4d), which is much better than most of





**Fig. 4** Electrochemical tests on C-BCS measured in 6 M KOH. (a) CV curves of 700-3 sample at different scan rates. (b) CV curves of different C-BCS at a scan rate of  $5 \text{ mV s}^{-1}$ . (c) GCD curves of 700-3 sample at different current densities. (d) Rate performance at different current densities. (e) Electrochemical impedance spectra of C-BCS. The inset shows the magnified region of  $0.5\text{--}3.5 \Omega$ . (f) Cyclic stability of 700-3 at a charge-discharge current density of  $5 \text{ A g}^{-1}$  for 10 000 cycles.

previous reported porous carbon materials derived from biomass (Table 3).

EIS measurements were also employed to evaluate the internal resistance of the material as well as the resistance in charge transfer and ion diffusion processes of supercapacitors.<sup>52,53</sup> Fig. 4e shows the Nyquist plots collected in 6 M KOH in the frequency range of 0.1 Hz to 100 kHz with an amplitude of 5 mV. It is found that the slope curves of all the remaining samples in the low frequency region are near to  $90^\circ$ , except for the 700-0 sample, indicating that the activated carbon electrodes are porous and have the capacitive characteristics of an ideal capacitor. As the weight ratio between carbonaceous samples and KOH increases, the capacitance characteristics of

the activated carbon electrode become better. The semicircular arc in the high-frequency region simultaneously reflects the bulk properties of the electrode and electrolyte and the charge transfer at the electrode/electrolyte interface. Comparing the semicircular diameters of the high-frequency portions of the six impedance curves in the inset figure, it can be found that the  $R_{ct}$  increases gradually as the carbonization temperature or the weight ratio between carbonaceous samples and KOH increases. According to the previous studies<sup>42</sup> and the obtained data listed in Table 1, the mesopore structure obtained by activation is at its most when the ratio of alkali to carbon is 3 : 1, which can give rise to a small charge transfer resistance. Similarly, when the ratio of alkali to carbon is lower, the micropore

**Table 3** Comparison of various biomass-derived porous carbon materials in supercapacitor applications

Precursor	$S_{\text{BET}}$ ( $\text{m}^2 \text{g}^{-1}$ )	Current density ( $\text{A g}^{-1}$ )	$C_g$ ( $\text{F g}^{-1}$ )	Electrolyte	Ref.
Yogurt	1300	2	225	1 M $\text{H}_2\text{SO}_4$	11
Bamboo byproduct	1472	0.1	301	6 M KOH	22
Auricularia	1607	1	347	6 M KOH	26
Soybean	350	$10 \text{ mV s}^{-1}$	315	6 M KOH	27
Tofu	1208	0.5	262	1 M $\text{Na}_2\text{SO}_4$	28
Soybean root	2143	0.5	276	6 M KOH	29
Soybean stalk	1709	0.5	384.9	1 M $\text{Na}_2\text{SO}_4$	30
Kiwifruit	379.0	0.5	337.4	6 M KOH	34
Jujun grass	2532	1	336	6 M KOH	36
Wheat straw	1201	$5 \text{ mV s}^{-1}$	223.9	6 M KOH	38
Dead pine needles	783	0.5	223	1 M $\text{H}_2\text{SO}_4$	39
Bean curd stick	2609	0.5	405	6 M KOH	This work



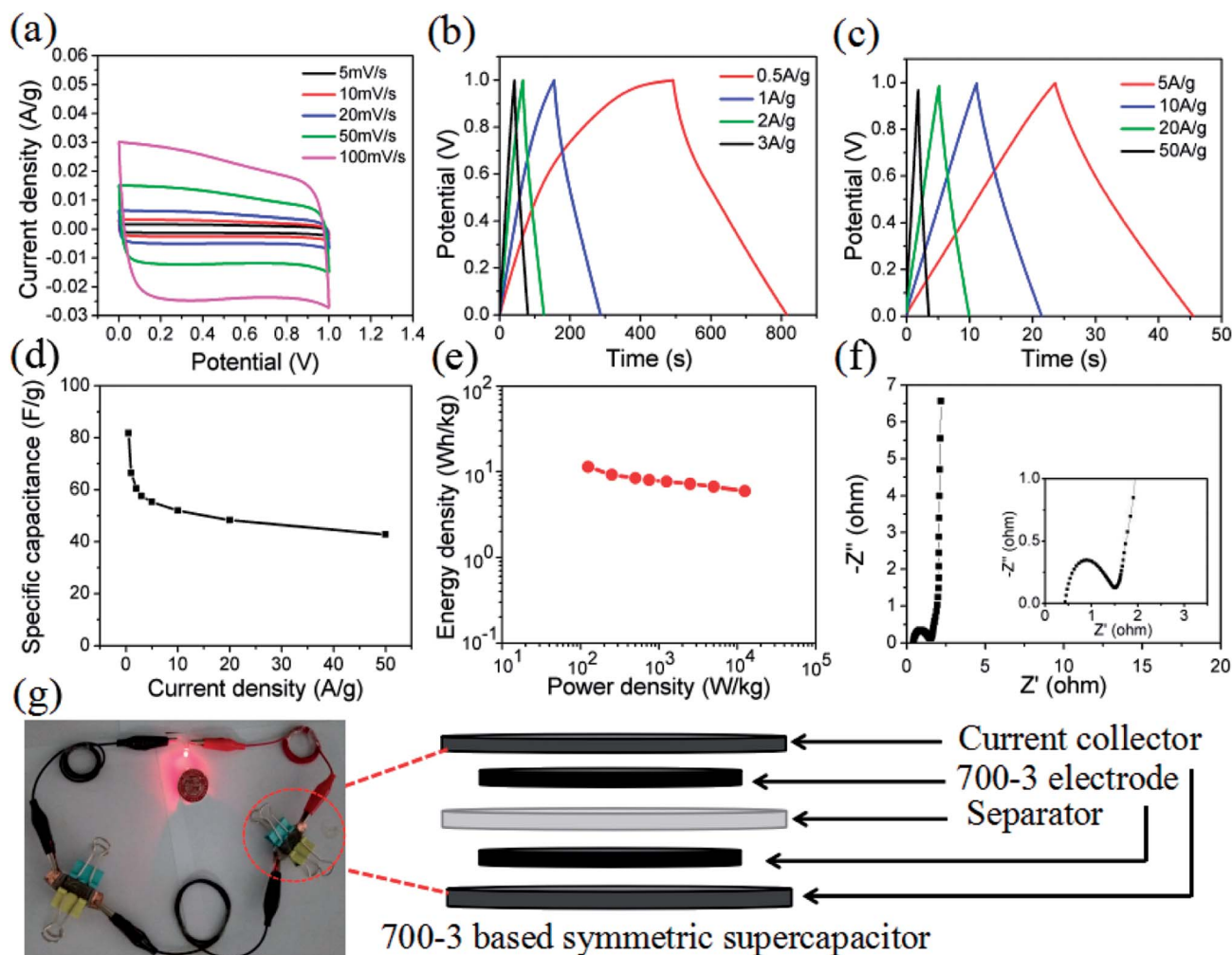


Fig. 5 Electrochemical experiments using the 700-3//700-3 symmetric supercapacitor. (a) CV test at different scan rates. (b) and (c) GCD test at different current densities. (d) Rate performance with different current densities. (e) Ragone plot of 700-3 based symmetrical supercapacitor. (f) EIS measurement of the symmetric supercapacitor. The inset shows the magnified region of 0.5–3.5  $\Omega$ . (g) Red LED powered by the fabricated device and schematic illustration of the fabricated symmetric two-electrode supercapacitor.

structure of activated carbon increases gradually, so the charge transfer resistance is greater. The 700-3 sample possesses the smallest charge transfer resistance and the fastest charge/discharge rate, and therefore has the better capacitive characteristics. Furthermore, we also investigated the long-term cyclic stability of the 700-3 porous carbon materials based electrode. By using GCD measurement at a current density of 5  $A g^{-1}$  within the potential range of  $-1.2$  to  $-0.2$  V vs. SCE, we observed only a slight variation of specific capacitance with the increase of cycle number and the remaining capacitance was 90.1% after 10 000 cycles (Fig. 4f), which indicated that the electrode possesses an excellent charge/discharge stability.

In order to further explore the 700-3 porous carbon material for practical supercapacitor applications, the capacitive performance of a two-electrode symmetric supercapacitor containing the 700-3 sample was also evaluated. As displayed in Fig. 5a, the CV curves exhibited typical capacitive behavior with rectangular-like voltammetry characteristics, and possessed an ideal capacitive behavior consistent with the result for the three-electrode system. Besides, for GCD measurement, no obvious IR

drop could be observed with the linear and nearly symmetric triangular curves, suggesting good capacitive behavior of the 700-3 sample based supercapacitor (Fig. 5b and c). The specific capacitance of the supercapacitor was calculated to be 81.7  $F g^{-1}$  at 0.5  $A g^{-1}$  and 42.8  $F g^{-1}$  at 50  $A g^{-1}$  (Fig. 5d), while the capacitance retention was 52.4%. Fig. 5e displays the Ragone plot of the assembled 700-3 sample based symmetric supercapacitor. As shown, the present symmetric supercapacitor delivered a maximum energy density of 11.35  $W h kg^{-1}$  at a power density of 125  $W kg^{-1}$ , and 5.94  $W h kg^{-1}$  still remained even at a high power density of 12 500  $W kg^{-1}$ . This obtained maximum energy density value is comparable to or even better than the commercial devices ( $<5$   $W h kg^{-1}$ )<sup>54,55</sup> and other previously reported carbon-based symmetric supercapacitors in aqueous electrolyte, such as those based on yogurt (7  $W h kg^{-1}$ ),<sup>11</sup> lignin (8.34  $W h kg^{-1}$ ),<sup>56</sup> wild jujube pit (8.95  $W h kg^{-1}$ ),<sup>57</sup> pomelo peel (9.4  $W h kg^{-1}$ ),<sup>34</sup> bamboo byproduct (9.5  $W h kg^{-1}$ )<sup>22</sup> and soybean curd residue (9.95  $W h kg^{-1}$ ).<sup>27</sup> Such excellent performance of the capacitor is due mainly to the large specific surface area, interconnected





porous structure and efficient N-doped structure of 700-3. On the one hand, the doped nitrogen heteroatom in the carbon skeleton can effectively increase its wettability and induce pseudocapacitance behavior;<sup>48</sup> on the other, the unique pore size distribution in the material would provide a well-defined diffusion pathway for ions and electrons.<sup>51</sup> The EIS measurement also revealed that the symmetric supercapacitor has small ionic resistance of the electrolyte (1.07  $\Omega$ ), and the near-vertical slope curve in the low frequency region also indicates that the supercapacitor has excellent capacitance characteristics (Fig. 5f). Furthermore, two devices connected in tandem can also power a 2.5 V red light-emitting diode (LED) (Fig. 5g), indicating that 700-3 porous carbon is a promising electrode material for supercapacitors.

## 4. Conclusions

In summary, we successfully synthesized a heteroatom-doped porous carbon structure from the byproduct (slag material) of the bean curd stick manufacturing process. The present carbon material not only possesses a large specific surface area (2609  $\text{m}^2 \text{g}^{-1}$ ) but also has an outstanding capacitive performance with high capacitance and superior cycle stability. Remarkably, the 700-3 porous carbon materials based symmetric capacitor also manifests a high energy density of 11.35  $\text{W h kg}^{-1}$ , which is comparable to or much better than the commercially available devices and other previously reported carbon-based symmetric supercapacitors in aqueous electrolyte. The present work not only provides a novel biomass-derived porous carbon material for future high-performance supercapacitors applications, but also can add great secondary-utilization value to the byproduct of bean curd stick manufacture.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This work was supported by the National Natural Science Foundation of China (21305110), the Young Talent Lifting Plan of Association for Science and Technology of Shaanxi (20160215), Shaanxi Provincial Department of Education Natural Science Special Project (17JK0778) and Scholar Tang.

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