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# Influence on ferric chloride aqueous solution caused by external electrostatic field: a molecular dynamics simulation study

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A detailed analysis of structural properties and dynamic properties of ferric chloride aqueous solution under external electrostatic fields with different intensities was performed by molecular dynamics (MD) simulations. The effects on the ferric chloride aqueous solution caused by using an electrostatic field were examined in terms of the radial distribution function of the  $\text{Fe}^{3+}$  ion and water molecule/ $\text{Cl}^-$ , coordination number of water/ $\text{Cl}^-$  around the  $\text{Fe}^{3+}$  ion, characteristics of hydrogen bonds, solution viscosity, and how these effects influence the hydrolysis process of the  $\text{Fe}^{3+}$  ion. The goal behind the study is to attain additional insights into the mechanism of electrocoagulation when ferric chloride is used as coagulant, and provide a fundamental basis for the practical use of this technology.

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## Introduction

Flocculation is an effective method in many water treatment engineering fields. Adding coagulant is the most important step in the flocculation process. The most widely used coagulants are aluminium salt and ferric salt. Aluminium salt has a certain biological toxicity to human beings, which has a correlation with the occurrence of Alzheimer's disease.<sup>1,2</sup> Therefore the use of aluminium salt coagulant may cause the secondary pollution of raw water. Compared with the aluminium salt coagulant, the ferric salt coagulant has the advantages of no biological toxicity, higher coagulation and settlement rate, better dewatering of sludge and shorter water treatment time.<sup>3</sup> However, in the practical application of ferric salt coagulant, there are still many disadvantages such as the high dosage required, high cost and secondary environmental pollution. Therefore, an enhanced electrocoagulation technique has been developed. The enhanced electrocoagulation technique can effectively improve coagulation efficiency and significantly reduce coagulant dosage. At present, the enhanced electrocoagulation technique has been widely used in the treatment of many kinds of wastewater, such as dye and non-ferrous waste water, oily wastewater and municipal wastewater, which has achieved good results. Many researchers have done a lot of research on the mechanism and optimization method of the enhanced electrocoagulation technique,<sup>4-9</sup> but

most of the research only stays in the stage of observing and analysing experimental phenomena and does not obtain the reliable intrinsic micro mechanism. This is mainly due to the rapid hydrolysis reaction when the coagulant is dissolving in the water,<sup>3</sup> and this process is difficult to observe and measure through experimental methods. Therefore, studying the influence of electrostatic field on the hydrolysis process of ferric salt is important to explore the mechanism of electrocoagulation.

Ferric chloride is the most widely used ferric salt coagulant. It is a kind of  $\text{Fe}(\text{III})$  salt. When the  $\text{Fe}(\text{III})$  salt enters the water, it will generate  $\text{Fe}^{3+}$  through ionization. But the  $\text{Fe}^{3+}$  is not stable in the water, which is quickly combined with the water molecules to form  $\text{Fe}(\text{H}_2\text{O})_6^{3+}$ . Then the  $\text{Fe}(\text{H}_2\text{O})_6^{3+}$  will show a strong tendency to hydrolyse in water and form various hydroxyl iron ions.<sup>10,11</sup> When external electrostatic field is present, the process of  $\text{Fe}^{3+}$  combining with the water molecules will be influenced.

Therefore, in the present article the ferric chloride solution was taken as the study object, and the influence on the ferric chloride solution system caused by external electrostatic field during the time slot before the chemical reaction of  $\text{Fe}^{3+}$  and water molecules was studied. Since molecular dynamics (MD) simulation can cover up a time span of micro-seconds or nano-seconds, we used MD simulation to study the water shell radius of the hydrated  $\text{Fe}^{3+}$ , the coordination number of water molecules, the average residence time of water molecules in water shell of  $\text{Fe}^{3+}$ , the self-diffusion coefficient of  $\text{Fe}^{3+}$  and water molecules, the hydrogen bond structural parameters and the changes of solution viscosity under electrostatic fields with different intensity. The potential effects of these changes on the hydrolysis of  $\text{Fe}^{3+}$  and the condensation process of  $\text{Fe}^{3+}$  complexes were also analyzed.

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## The detailed description of the simulation

The changes of ferric chloride aqueous solution under electrostatic field were investigated by equilibrium molecular dynamics. The simulations were performed in the NPT ensemble (at constant particle number, pressure and temperature) using the GROMACS package.<sup>12</sup> All the simulated temperatures were set at 300 K by using Berendsen temperature coupling method. The pressures were set at 1 atm. Since the chemical reaction of  $\text{Fe}^{3+}$  and water molecules doesn't begin at that initial time slot, the water molecules will not be broken. In addition, SPC/E model is one of the most popular water models for the remarkable accuracies of various properties over wide temperatures and pressures. In particular, this model reproduces the dielectrostatic constant quite well as compared with other models, which is very important for the salt-water mixtures. Therefore, according to the simulation method of salt-water system,<sup>13-15</sup> the SPC/E model was adopted for water molecules in our work. The ferric chloride solution with a concentration of 1.03 mol  $\text{kg}^{-1}$  was selected as the research object. According to the requirement of the composition ratio of ferrous chloride in ferric chloride coagulant in the Chinese national standard GB-4482-2006 "water treatment agent ferric chloride", 2 ferrous chloride molecules were added into the system. According to the concentration of the solution, the whole simulation system consisted of 3765 water molecules, 70  $\text{Fe}^{3+}$  ions, 2  $\text{Fe}^{2+}$  ions and 214  $\text{Cl}^-$  ions (a total of 4051 particles), which were placed in the box of 5 nm  $\times$  5 nm  $\times$  5 nm. The periodic boundary condition was used. In this paper, the GROMOS force field was used to calculate the interaction between iron ions and water molecules.<sup>15</sup> When calculating the Lennard-Jones interaction, the cut-off was set as 1 nm. Fast Particle-Mesh ESD electrostatic method (PME)<sup>16</sup> was used to calculate the electrostatic interaction with a real space cut-off of 0.7 nm. The value of Fourier spacing was 0.12 nm. The time step length used for all simulations was 2 fs,<sup>15</sup> and the simulation time of each electrostatic field intensity was 20 ns. The Lennard-Jones potential parameters of each particle are listed in Table 1, and the potential parameters of  $\text{Fe}^{3+}$  and  $\text{Fe}^{2+}$  are referred in ref. 17. The electrostatic field intensity employed in the simulation process was set as  $10^5 \text{ V m}^{-1}$ ,  $10^6 \text{ V m}^{-1}$ ,  $10^7 \text{ V m}^{-1}$ ,  $10^8 \text{ V m}^{-1}$ , respectively, and the direction of electrostatic field was along the direction of the Y axis. The finished system model is shown in Fig. 1. Since the parameters will change easily during the whole simulation time, it is necessary to evaluate the degree of numerical fluctuation in this process. The results of the parameters were obtained from the last 10 ns of every simulation. The 10 ns was divided into ten segments, and we obtained

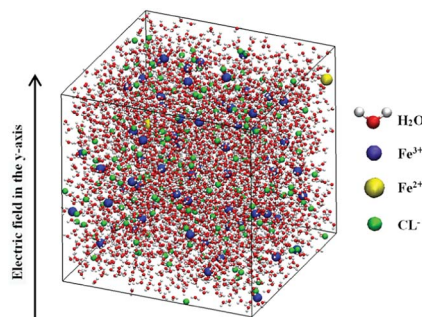


Fig. 1 Figure of setup for simulations.

10 parameter values from each segment. The final result of parameter was obtained from the average of the 10 parameter values. In addition, the volatility of the data was measured by standard deviation marked as error bar in the figures.

In order to verify the correctness of the simulation method, the self-diffusion coefficient of two kinds of iron ions, the self-diffusion coefficient of water molecules, the first peak position of the radial distribution function of the iron ion-water molecules and the water molecule coordination number of hydrated  $\text{Fe}^{3+}$  ion were calculated without external electrostatic field. The solution temperature was set as 300 K and the solution concentration was set as 0.11 mol  $\text{kg}^{-1}$ . The results were compared with the simulation and experimental results in other literatures. The results are shown in Table 2.

As can be seen from Table 2, the calculation results of  $D_{\text{Fe(III)}}$ ,  $D_{\text{Fe(II)}}$ ,  $n_{\text{Fe(III)-H}_2\text{O}}$ ,  $n_{\text{Fe(II)-H}_2\text{O}}$ , location of the first peak of the  $g_{\text{Fe(III)-H}_2\text{O}}(r)$  and  $g_{\text{Fe(II)H}_2\text{O}}(r)$  are close to the results of other literatures, which proves that the MD simulation method in this paper is more reasonable and can be used to study the change of solution properties of ferric chloride solution under external electrostatic field.

## Results and discussion

### Radial distribution function

In this section, the values of  $P_{\text{Fe(III)-H}_2\text{O}}$  (first peak value),  $P_{\text{Fe(III)-Cl}}$ ,  $R_{\text{Fe(III)-H}_2\text{O}}$ ,  $n_{\text{Fe(III)-H}_2\text{O}}$ ,  $R_{\text{Fe(III)-Cl}}$ ,  $n_{\text{Fe(III)-Cl}}$  were calculated under the external electrostatic field with intensity of  $10^5 \text{ V m}^{-1}$ ,  $10^6 \text{ V m}^{-1}$ ,  $10^7 \text{ V m}^{-1}$  and  $10^8 \text{ V m}^{-1}$ , respectively, and the results were compared with that under no external electrostatic field. The radial distribution function was calculated according eqn (1):<sup>22</sup>

$$g_{\text{AB}}(r) = \frac{\langle \rho_{\text{B}}(r) \rangle}{\langle \rho_{\text{B}} \rangle_{\text{local}}} = \frac{1}{\langle \rho_{\text{B}} \rangle_{\text{local}}} \frac{1}{N_{\text{A}}} \sum_{i \in \text{A}} \sum_{j \in \text{B}} \frac{\delta(r_{ij} - r)}{4\pi r^2} \quad (1)$$

where  $N_{\text{A}}$  is the quantity of particle A, and  $N_{\text{B}}$  is the quantity of particle B;  $\langle \rho_{\text{B}}(r) \rangle$  is the particle density of particle B at a distance  $r$  from particle A, and  $\langle \rho_{\text{B}}(r) \rangle_{\text{local}}$  is the particle density of particle B averaged over all spheres around particle A with radius  $r_{\text{max}}$ . Usually the value of  $r_{\text{max}}$  is half of the minimum box length.

The water molecule coordination number of  $\text{Fe}^{3+}$  is calculated according to eqn (2):<sup>23</sup>

Table 1 Lennard-Jones potential parameters used in the MD simulations

	$\epsilon$ ( $\text{kJ mol}^{-1}$ )	$\sigma$ (nm)
$\text{Fe}^{3+}$	2.166	0.191
$\text{Fe}^{2+}$	2.106	0.191
$\text{Cl}^-$	1.2889	0.3470
O	0.6502	0.3166



Table 2 The simulation results from our work comparing with other simulation method and experiment<sup>a</sup>

	M1	M2	M3
$D_{\text{Fe(III)}} (\text{cm}^2 \text{s}^{-1} \times 10^{-5})$	0.52	0.53 <sup>b</sup> , 0.50 <sup>c</sup>	0.56 <sup>d</sup>
$D_{\text{Fe(II)}} (\text{cm}^2 \text{s}^{-1} \times 10^{-5})$	0.61	0.56 <sup>b</sup> , 0.69 <sup>c</sup>	0.7 <sup>e</sup>
$n_{\text{Fe(III)-H}_2\text{O}}$	6.21	6 <sup>b</sup> , 6.46 <sup>c</sup>	6 <sup>f</sup>
$n_{\text{Fe(II)-H}_2\text{O}}$	5.96	6 <sup>b</sup> , 6.23 <sup>b</sup>	6 <sup>f</sup>
Location of first peak of the $g_{\text{Fe(III)-H}_2\text{O}}(r)$ (nm)	0.202	0.196 <sup>b</sup> , 0.195 <sup>c</sup>	0.198 <sup>f</sup>
Location of first peak of the $g_{\text{Fe(II)-H}_2\text{O}}(r)$ (nm)	0.208	0.209 <sup>b</sup> , 0.212 <sup>c</sup>	0.210 <sup>f</sup>

<sup>a</sup> M1 is from our work, M2 is the result by other simulate method, M3 is from experiments. <sup>b</sup> From ref. 18 using MD simulation. <sup>c</sup> From ref. 17 using MD simulation. <sup>d</sup> From ref. 19 using diaphragm cell method. <sup>e</sup> From ref. 20 using diaphragm cell method. <sup>f</sup> From ref. 21 using X-ray.

Table 3 Values of radial first peak position of  $\text{Fe}^{3+}-\text{H}_2\text{O}$  and  $P_{\text{Fe(III)-H}_2\text{O}}$  under electrostatic field with different intensity

Electrostatic field strength	0 V m <sup>-1</sup>	10 <sup>5</sup> V m <sup>-1</sup>	10 <sup>6</sup> V m <sup>-1</sup>	10 <sup>7</sup> V m <sup>-1</sup>	10 <sup>8</sup> V m <sup>-1</sup>
Radial first peak position of $\text{Fe}^{3+}-\text{H}_2\text{O}$ (nm)	0.206	0.206	0.206	0.206	0.206
$P_{\text{Fe(III)-H}_2\text{O}}$	21.952	21.768	21.667	21.663	21.517

$$n_{\text{Fe-H}_2\text{O}}(r_{s1}) = 4\pi\rho_0 \int_0^{r_{s1}} g_{\text{Fe-H}_2\text{O}}(r)r^2 dr \quad (2)$$

where  $r_{s1}$  is the location of the first local minimum of  $g_{\text{Fe-H}_2\text{O}}(r)$ .

The chloride ions coordination number of  $\text{Fe}^{3+}$  is calculated according to eqn (3):

$$n_{\text{Fe-Cl}}(r_{s2}) = 4\pi\rho_{\text{Cl}} \int_0^{r_{s2}} g_{\text{Fe-Cl}}(r)r^2 dr \quad (3)$$

where  $r_{s2}$  is the location of the first local minimum of  $g_{\text{Fe-Cl}}(r)$ .

The calculation results of radial first peak position of  $\text{Fe}^{3+}-\text{H}_2\text{O}$  and  $P_{\text{Fe(III)-H}_2\text{O}}$  are shown in Table 3 and Fig. 2. The results of  $R_{\text{Fe(III)-H}_2\text{O}}$  and  $n_{\text{Fe(III)-H}_2\text{O}}$  are shown in Fig. 3 and 4. The  $R_{\text{Fe(III)-H}_2\text{O}}$  was obtained from the location of the first local minimum of  $g_{\text{Fe-H}_2\text{O}}(r)$ .

According to Table 3 and Fig. 2 it can be seen that under electrostatic field with different intensity the position of first peak of  $g_{\text{Fe-H}_2\text{O}}(r)$  are all 0.206 nm, which is consistent with the

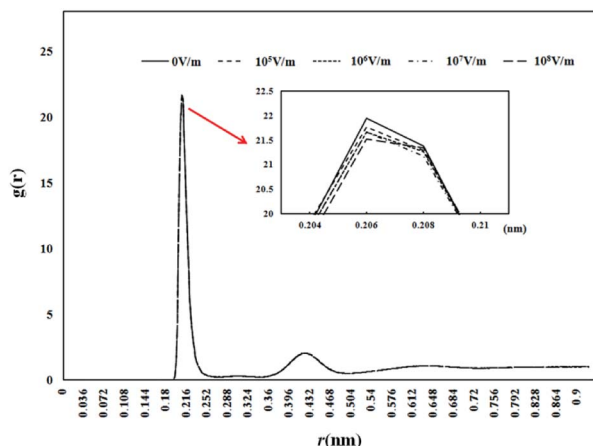


Fig. 2 Radial distribution function of  $\text{Fe}^{3+}-\text{H}_2\text{O}$  under electrostatic field with different intensity.

that under no electrostatic field. This indicates that electrostatic field has no influence on the position of first peak of  $g_{\text{Fe-H}_2\text{O}}(r)$ . But the external electrostatic field decreases the value of  $P_{\text{Fe(III)-H}_2\text{O}}$ , and this value decreases with the increase of electrostatic field intensity. When the electrostatic field intensity are 0 V m<sup>-1</sup>, 10<sup>5</sup> V m<sup>-1</sup>, 10<sup>6</sup> V m<sup>-1</sup>, 10<sup>7</sup> V m<sup>-1</sup> and 10<sup>8</sup> V m<sup>-1</sup>, the  $P_{\text{Fe(III)-H}_2\text{O}}$  values are 21.952, 21.768, 21.667, 21.663, 21.517

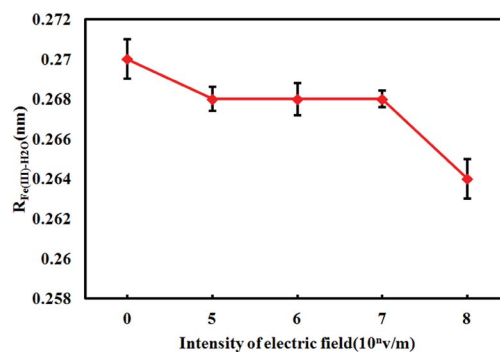


Fig. 3 The radius of hydrated  $\text{Fe}^{3+}-\text{H}_2\text{O}$  under electrostatic field with different intensity.

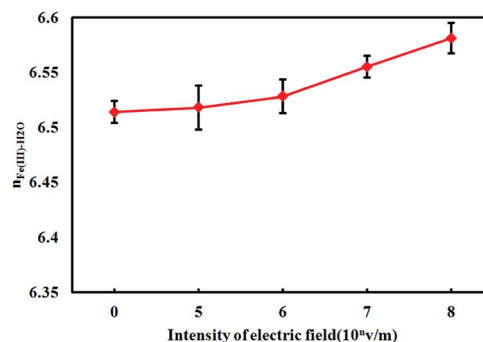
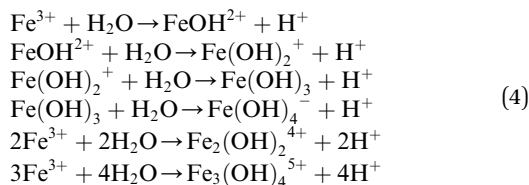


Fig. 4 The number of coordination water around  $\text{Fe}^{3+}$  under electrostatic field with different intensity.



respectively. From Fig. 3, it is known that the value of  $R_{\text{Fe(III)}-\text{H}_2\text{O}}$  decreases notably under electrostatic field. It can be seen that except for the value of  $R_{\text{Fe(III)}-\text{H}_2\text{O}}$  is 0.264 nm when electrostatic field intensity is  $10^8 \text{ V m}^{-1}$ , under other electrostatic fields the  $R_{\text{Fe(III)}-\text{H}_2\text{O}}$  values are all 0.268 nm, which is only less 0.002 nm than that under no electrostatic field. Therefore, the external electrostatic field has little influence on the radius of hydrated  $\text{Fe}^{3+}$  ion. As shown in Fig. 4, compared with no external electrostatic field the value of  $n_{\text{Fe(III)}-\text{H}_2\text{O}}$  under electrostatic field increases, and as electrostatic field intensity increases the value of  $n_{\text{Fe(III)}-\text{H}_2\text{O}}$  is obviously increased.

When external electrostatic field is present the value of  $P_{\text{Fe(III)}-\text{Cl}}$  decreases, which indicates that the number of water molecule in the position of the first peak (0.206 nm) is reduced. But the ionic radius of the hydrated  $\text{Fe}^{3+}$  ion is not affected by electrostatic field, and the number of coordination water molecule of the hydrated  $\text{Fe}^{3+}$  ion is increased. The changes above indicate that the external electrostatic field makes more water molecules become the coordination water of  $\text{Fe}^{3+}$  ions, which means the external electrostatic field can enhance the interaction between  $\text{Fe}^{3+}$  ions and the nearby water molecules. This enhanced interaction is beneficial to the hydrolysis of  $\text{Fe}^{3+}$  ions in water. Hedstrom<sup>10</sup> confirmed the existence of  $\text{FeOH}^{2+}$ ,  $\text{Fe(OH)}_2^+$  and  $\text{Fe}_2(\text{OH})_2^{4+}$  during the hydrolysis process of  $\text{Fe(III)}$  salt, and measured their hydrolysis equilibrium constants. Biederman<sup>11</sup> found that except the above hydrolytic ions,  $\text{Fe}_3(\text{OH})_4^{5+}$  (a kind of trimer) also existed. The equations of possible hydrolytic reaction of  $\text{Fe}^{3+}$  ions in water shows as below:



Under normal conditions the complex form of the ferric monomer produced by hydrolysis intensively tends to be polymerized, and the most likely product is the dipolymer  $\text{Fe}_2(\text{OH})_2^{4+}$ , because dipolymer  $\text{Fe}_2(\text{OH})_2^{4+}$  has significant stability.<sup>10,11,24,25</sup> No matter what forms of hydrolysates produced by hydrolysis of  $\text{Fe}^{3+}$ , water molecules are the main reactants. The electrostatic field enhances the interaction and the reaction probability of the  $\text{Fe}^{3+}$  and surrounding water molecules, which facilitates the polymerization of the complex form of the ferric monomer. This is beneficial to the enhancement of the coagulation effect of the ferric chloride coagulant.

The interaction between the  $\text{Fe}^{3+}$  and the surrounding water molecules is enhanced by the external electrostatic field, and it can be indirectly reflected from the distribution of the anion around the  $\text{Fe}^{3+}$ . Fig. 5 and 7 show the value of  $P_{\text{Fe(III)}-\text{Cl}}$ ,  $R_{\text{Fe(III)}-\text{Cl}}$ ,  $n_{\text{Fe(III)}-\text{Cl}}$  under electrostatic field with different intensity.

From Fig. 5, it can be seen that the first peak position of the radial distribution function of  $\text{Fe}^{3+}-\text{Cl}^-$  is 0.232 nm no matter whether there is external electrostatic field or not. But the electrostatic field reduces the value of  $P_{\text{Fe(III)}-\text{Cl}}$ , and as the

intensity of the electrostatic field increases, the value of  $P_{\text{Fe(III)}-\text{Cl}}$  is getting smaller. When the electrostatic field intensity is  $0 \text{ V m}^{-1}$ ,  $10^5 \text{ V m}^{-1}$ ,  $10^6 \text{ V m}^{-1}$ ,  $10^7 \text{ V m}^{-1}$  and  $10^8 \text{ V m}^{-1}$ , the  $P_{\text{Fe(III)}-\text{Cl}}$  values are 19.332, 18.554, 16.837, 16.545, 15.395 respectively. From Fig. 6 and 7, the electrostatic field has little influence on the value of  $R_{\text{Fe(III)}-\text{Cl}}$ . No matter whether there is electrostatic field or not, the value is about 0.254 nm. But electrostatic field significantly reduces the coordination number of  $\text{Cl}^-$  around

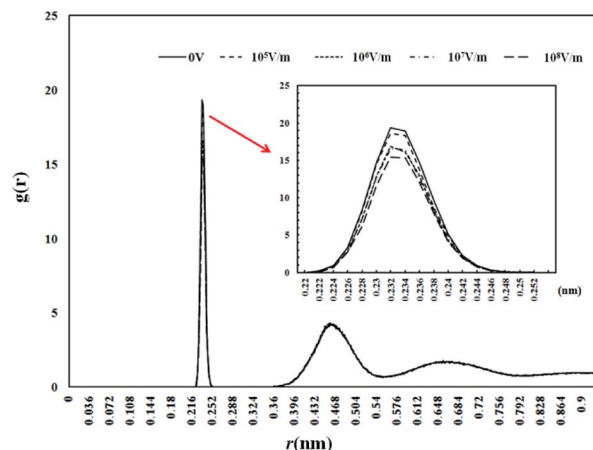


Fig. 5 Radial distribution function of  $\text{Fe}^{3+}-\text{Cl}^-$  under electrostatic field with different intensity.

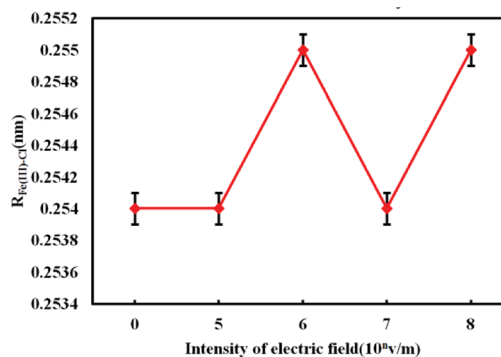


Fig. 6 The radius of hydrated  $\text{Fe}^{3+}-\text{Cl}^-$  under electrostatic field with different intensity.

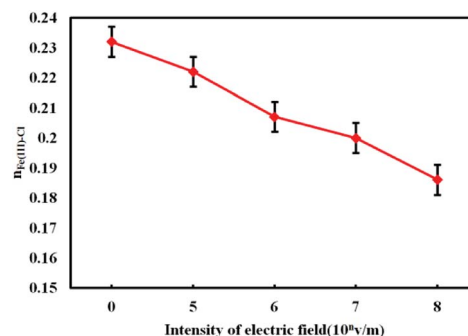


Fig. 7 The number of coordination  $\text{Cl}^-$  around  $\text{Fe}^{3+}$  under electrostatic field with different intensity.



the  $\text{Fe}^{3+}$ , and the coordination number decreases with the increase of the electrostatic field intensity.

According to Debye–Huckel Theory proposed by Debye and Huckel in 1923, on the one hand, the Coulomb force between the  $\text{Fe}^{3+}$  and the  $\text{Cl}^-$  makes the ions tend to be orderly arranged as in the lattice. On the other hand, the Brownian motion will make the ions in a random distribution. Since Brownian motion is not strong enough to counteract the Coulomb force, the result of the interaction between these two forces is bound to form such a situation: around a  $\text{Fe}^{3+}$  (central ion), the  $\text{Cl}^-$  ions are more likely to appear than other  $\text{Fe}^{3+}$  ions. Therefore, it can be considered that the  $\text{Cl}^-$  ions with negative charges are relatively concentrated around the  $\text{Fe}^{3+}$  central ions. Comparing with Fig. 2 the reduction in value of  $P_{\text{Fe(III)}-\text{Cl}}$  is much more than  $P_{\text{Fe(III)}-\text{H}_2\text{O}}$  with increase in the strength of the electrostatic field. This may be because that the interaction between  $\text{Fe}^{3+}$  and  $\text{H}_2\text{O}$  is electrostatic interaction, but the interaction between  $\text{Fe}^{3+}$  and  $\text{Cl}^-$  is composed of electrostatic interaction and ionic polarization interaction. Electrostatic interaction and ionic polarization interaction are all caused by the electric field of electron cloud. Therefore the interaction between  $\text{Fe}^{3+}$  and  $\text{Cl}^-$  is easier to be influenced by external electric field than the interaction between  $\text{Fe}^{3+}$  and  $\text{H}_2\text{O}$ . In addition, the  $P_{\text{Fe(III)}-\text{H}_2\text{O}}$  is higher as compared to the one given in ref. 26. Also in present study, the  $P_{\text{Fe(III)}-\text{H}_2\text{O}}$  is higher as compared to  $P_{\text{Fe(III)}-\text{Cl}}$  whereas in ref. 26, the trend is opposite. We believe that this phenomenon may be caused by the difference of solution concentration. Our solution concentration is  $1.03 \text{ mol kg}^{-1}$ , but the concentration in the reference is  $4.1 \text{ mol kg}^{-1}$ . We also run another simulation with the  $\text{FeCl}_3$  concentration of  $4.1 \text{ M}$  for verification. The result showed that the intensity was higher for  $P_{\text{Fe(III)}-\text{Cl}}$  as compared to  $P_{\text{Fe(III)}-\text{H}_2\text{O}}$ , which met with the result in the reference.

From Fig. 5 and 7, it is found that  $P_{\text{Fe(III)}-\text{Cl}}$  and  $n_{\text{Fe(III)}-\text{Cl}}$  are reduced simultaneously under the electrostatic field, which indicates that the electrostatic field weakens the Coulomb force between the central  $\text{Fe}^{3+}$  and the surrounding  $\text{Cl}^-$  ions and reduces the number of  $\text{Cl}^-$  ions around the  $\text{Fe}^{3+}$ . In the meantime, since there are only  $\text{Cl}^-$  and water molecules around  $\text{Fe}^{3+}$  ions, the vacancies produced by reduced  $\text{Cl}^-$  ions are replaced by water molecules. This also indirectly confirmed the increase of number of coordination water around  $\text{Fe}^{3+}$ .

### Diffusivity

In this section, the values of the self-diffusion coefficient of  $\text{Fe}^{3+}$   $D_{\text{Fe(III)}}$  and water molecules  $D_{\text{H}_2\text{O}}$  are calculated when the external electrostatic field intensity is  $10^5 \text{ V m}^{-1}$ ,  $10^6 \text{ V m}^{-1}$ ,  $10^7 \text{ V m}^{-1}$  and  $10^8 \text{ V m}^{-1}$  respectively, and these values are compared with that under no external electrostatic field. The self-diffusion coefficient is calculated according to eqn (5) ‘Einstein relation’:<sup>27</sup>

$$D = \lim_{t \rightarrow \infty} \frac{1}{6t} \langle [r_i(t) - r_i(0)]^2 \rangle \quad (5)$$

where  $r_i(t)$  is the position of molecule  $i$  at  $t$  time, and  $r_i(0)$  is the initial position of molecular  $i$ . The bracket represents the

ensemble average. The calculated results of  $D_{\text{Fe(III)}}$  and  $D_{\text{H}_2\text{O}}$  under different electrostatic field are shown in Fig. 8 and 9.

As shown in Fig. 8, compared with the value of  $D_{\text{Fe(III)}}$  ( $0.0895 \times 10^{-5} \text{ cm}^2 \text{ s}^{-2}$ ) under no electrostatic field, the  $D_{\text{Fe(III)}}$  increases under external electrostatic field. As the intensity of the electrostatic field increases, the value of  $D_{\text{Fe(III)}}$  is bigger. However, it can be seen from Fig. 9 that, contrary to the change trend of  $D_{\text{Fe(III)}}$ , the self-diffusion coefficient of water molecules  $D_{\text{H}_2\text{O}}$  becomes smaller with the intensity of the electrostatic field, which means the value of  $D_{\text{H}_2\text{O}}$  is inversely proportional to the electrostatic field intensity.

The diffusion coefficient is an important parameter to study the thermodynamic properties of ferric chloride solution, and its change reflects the activity change of particles in solution, which will directly affect the equilibrium of the hydrolysis reaction. The results in Fig. 8 show that the electrostatic field enhances the activity of  $\text{Fe}^{3+}$  ions in the solution, which contributes to increase the collision probability of the  $\text{Fe}^{3+}$  with the water molecules, thus increases the probability of the  $\text{Fe}^{3+}$  to capture the coordination water and promotes the initial hydrolysis of the  $\text{Fe}^{3+}$ . In addition, while the  $\text{Fe}^{3+}$  is hydrolyzing, the polymerization between different  $\text{Fe}^{3+}$  ions will also occur, which can produce the ferric complexes in the form of polymer.<sup>10,11</sup> The increase of the self-diffusion coefficient of  $\text{Fe}^{3+}$  also increases the collision probability between different  $\text{Fe}^{3+}$  and complex monomers, thus promotes the formation of polymers. The positive charge of the polymer is generally higher such as  $\text{Fe}_2(\text{OH})_2^{4+}$  (dipolymer) with 4 positive charges and  $\text{Fe}_3(\text{OH})_4^{5+}$

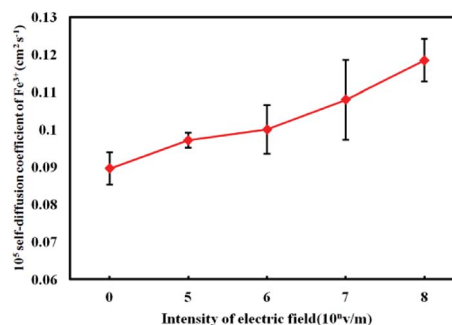


Fig. 8 Self-diffusion coefficient of  $\text{Fe}^{3+}$  under electrostatic field with different intensity.

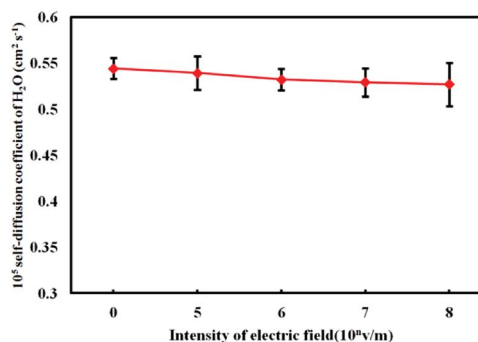


Fig. 9 Self-diffusion coefficient of  $\text{H}_2\text{O}$  under electrostatic field with different intensity.



(tripolymer) with 5 positive charges. The more charges can make stronger adsorption capacity. In addition, it is known from Fig. 9 that the effect of electrostatic field reduces the self-diffusion coefficient of water molecules, mainly because the interaction between  $\text{Fe}^{3+}$  and water molecules is enhanced by the electrostatic field, which means that the binding effect of  $\text{Fe}^{3+}$  on water molecules is enhanced.

### Residence time of water molecules in the first water shell of $\text{Fe}^{3+}$

The dynamical residence time of water molecules is an important parameter describing the effect of solute on water solvent. It can reflect the migration characteristics and stability of water molecules in the presence of solute.<sup>28,29</sup> Residence time of water molecules is the average time for water molecules to stay in the first water shell of solute ions.<sup>30</sup> In this paper, the residence time of water molecules was calculated by eqn (6):<sup>31</sup>

$$\gamma_{\text{res}} = \frac{1}{N_{\text{h}}} \sum_{i=1}^{N_{\text{h}}} \gamma_i \quad (6)$$

where  $N_{\text{h}}$  is the number of coordination water of  $\text{Fe}^{3+}$ . The residence time  $\gamma_i$  of each water molecule can be calculated by eqn (7):

$$\gamma_i = \frac{t_i}{n_i} \quad (7)$$

Where  $t_i$  is the total existence time of the molecule  $i$  in the first water shell,  $n_i$  is the number of water molecules entering and leaving the first water shell of  $\text{Fe}^{3+}$  ion in time  $t_i$ .

The calculation results of the residence time of water molecules under electrostatic field with different intensity are shown in Fig. 10.

As shown in Fig. 10, the residence time  $\gamma_{\text{res}}$  of the water molecules under different electrostatic field is larger than that without electrostatic field. And with the increase of electrostatic field intensity,  $\gamma_{\text{res}}$  shows an upward trend. The external electrostatic field prolongs the residence time of water molecule, which is also related to the enhanced interaction between  $\text{Fe}^{3+}$  and water molecules. The results obtained above show that the electrostatic field can effectively enhance the electrostatic interaction between  $\text{Fe}^{3+}$  and water molecules, which makes

$\text{Fe}^{3+}$  to attract more water as the coordination water molecules. Then more water molecules' mobility is limited, thus the diffusivity of water molecules is reduced. The diffusion of water molecules has direct relation with the residence time of water molecules,<sup>32</sup> the decrease of the diffusion of water molecules will reduce the frequency of water molecules entering and leaving the first water shell of  $\text{Fe}^{3+}$ . According to eqn (7) it is known that the reduction of the frequency of entering and leaving the first water shell will prolong the residence time of each water molecule, which makes the mean residence time of water molecules prolonged.

### Hydrogen bond

Hydrogen bond plays a important part during the process of hydrolysis, polymerization and adsorption. The hydrogen bond network structure in the solution directly affects the adsorption properties of polymeric ferric chloride.<sup>33</sup> Therefore, in this section the strength characteristics of hydrogen bond network are studied, including the number of hydrogen bonds ( $n_{\text{HB}}$ ) and the average lifetime of hydrogen bonds ( $\tau_{\text{HB}}$ ), which is due to the positive relationship between these two parameters and the strength of the hydrogen bond network.<sup>34</sup>

In the Gromacs software, hydrogen bonding is defined from two aspects: the distance between donor (D) and receptor (A), and the angle of hydrogen-donor-acceptor. The specific definition is as follows:

- (1) The distance between donor and acceptor is less than van der Waals distance (3.5 Å).
- (2) The angle of hydrogen-donor-acceptor is less than  $30^\circ$ .

In this simulation, the donor and acceptor are oxygen atoms in different water molecules. The angle of the hydrogen-donor-acceptor is expressed by  $\angle \text{HD-OD-OA}$ .

Hydrogen bonds exist widely among water molecules in water. Due to the dynamic characteristics of water molecules, the hydrogen bonds between the water molecules are constantly forming and breaking at the same time.<sup>35</sup> Therefore, hydrogen bond lifetime represents the average time of hydrogen bond in this dynamic process. The hydrogen bond lifetime  $\tau_{\text{HB}}$  can be calculated by eqn (8).<sup>36</sup>

$$\tau_{\text{HB}} = \int_0^\infty C(\tau) d\tau \quad (8)$$

where  $C(\tau)$  is a hydrogen bond correlation function.  $C(\tau)$  is defined as:

$$C(\tau) = \langle h_i(t)h_i(t + \tau) \rangle \quad (9)$$

where  $h_i(t) = \{0,1\}$  is the Heaviside unit step function, which means that if two molecules formed hydrogen bonds at time 0 and  $t$ , the function value is 1, and otherwise 0.

The number of hydrogen bonds ( $n_{\text{HB}}$ ) and the percentage of water molecules with  $n$  hydrogen bonds ( $f_n$ ) were investigated, where  $n = 0, 1, 2, 3$ . The lifetime and the number of hydrogen bonds in solutions under different intensities of electrostatic field are shown in Tables 4 and 5.

It can be seen from Table 4 that the average lifetime of hydrogen bond in the solution decreases under the effect of

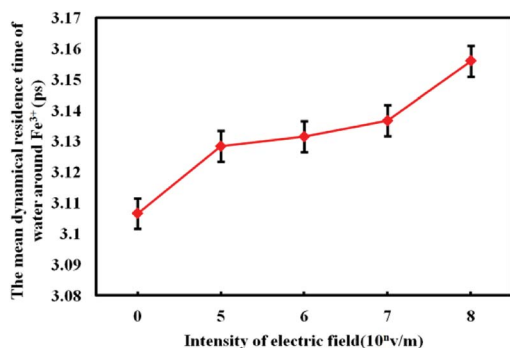


Fig. 10 Average residence time of  $\text{H}_2\text{O}$  molecules around  $\text{Fe}^{3+}$  under electrostatic field with different intensity.



**Table 4** Hydrogen bond life in solution under electrostatic field with different intensity

	Electrostatic field intensity				
	0 V m <sup>-1</sup>	10 <sup>5</sup> V m <sup>-1</sup>	10 <sup>6</sup> V m <sup>-1</sup>	10 <sup>7</sup> V m <sup>-1</sup>	10 <sup>8</sup> V m <sup>-1</sup>
Hydrogen bond life (ps)	7.53	7.53	7.53	7.52	7.50

**Table 5** The percentage of water molecules with *n* hydrogen bonds (*f<sub>n</sub>*) under electrostatic field with different intensity. (*n* = 0, 1, 2, 3)

	Electrostatic field intensity				
	0 V m <sup>-1</sup>	10 <sup>5</sup> V m <sup>-1</sup>	10 <sup>6</sup> V m <sup>-1</sup>	10 <sup>7</sup> V m <sup>-1</sup>	10 <sup>8</sup> V m <sup>-1</sup>
<i>f</i> <sub>0</sub> (%)	47.40	47.43	47.49	47.52	47.53
<i>f</i> <sub>1</sub> (%)	35.54	35.56	35.56	35.59	35.62
<i>f</i> <sub>2</sub> (%)	14.20	14.16	14.16	14.14	14.11
<i>f</i> <sub>3</sub> (%)	2.81	2.81	2.81	2.80	2.78
<i>n</i> <sub>HB</sub>	2722.38	2721.14	2718.70	2716.35	2713.53

electrostatic field, but the decrease is not obvious. Only when the intensity of the electrostatic field is large (10<sup>8</sup> V m<sup>-1</sup>), the average lifetime of hydrogen bond decreases from 7.53 ps to 7.50 ps. Therefore, the external electrostatic field has little effect on the average life of hydrogen bonds. However, according to Table 5 the number of hydrogen bonds decreases with the increase of electrostatic field intensity. The total number of water molecules involved in the bond formation decreases from 2722.38 to 2713.53. As the electrostatic field intensity rises from 0 V m<sup>-1</sup> to 10<sup>8</sup> V m<sup>-1</sup>, the proportion of the water molecules which do not participate in forming hydrogen bond rises from 47.40% to 47.53%, the proportion of water molecules involved in forming 1 hydrogen bonds rises from 35.54% to 35.62%, the proportion of water molecules involved in forming of 2 hydrogen bonds decreased from 14.20% to 14.11%, and the proportion of water molecules involved in forming 3 hydrogen bonds declined from 2.81% to 2.78%.

The change of the hydrogen bond structure parameters indicates that the electrostatic field has little influence on the hydrogen bond average life, but has a significant influence on the number of hydrogen bonds. The electrostatic field can effectively destroy the hydrogen bond inside the water cluster in the solution, and reduce the number of water molecules involved in the formation of multiple hydrogen bonds. In addition, the rise of the proportion of water molecules involved in forming 0 and 1 hydrogen bonds means that the number of monomer water molecules and dipolymers water molecules increased under electrostatic field, and this effect tends to be more significant with the increase of electrostatic field intensity. The increase in the number of monomer water molecules and dipolymers water molecules is beneficial to increase the diffusivity and permeability of water molecules. However, the result of the previous study shows that the self-diffusion coefficient of water molecules in the solution decreases under electrostatic field, which is mainly ascribed to the enhanced

binding to water molecules caused by Fe<sup>3+</sup>. Therefore, the change of hydrogen bond structure in the solution also proves the enhancement of the binding effect of Fe<sup>3+</sup> ions on water molecules under the electrostatic field.

## Viscosity

The dynamic viscosity of the solution is an important parameter for the conformation change of the solution. The change of viscosity of ferric chloride solution will affect the Brown motion characteristics of the ferric complex formed by hydrolysis and the rate of particles collision probability, thus affect the polymerization. It has been confirmed that electrostatic field can influence the dynamic viscosity of solution.<sup>37</sup> The dynamic viscosity of the solution can be obtained by eqn (10):

$$\eta = \frac{A \rho}{V k^2} \quad (10)$$

where  $\rho$  and  $V$  are the density and speed of solution,  $A$  is constant, and the two relations below must be satisfied:

$$a_x = A \cos(kz), k = 2\pi/l_z \quad (11)$$

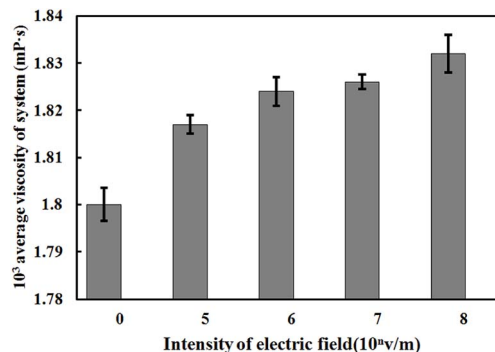
where  $l_z$  is the height of the box. The transient velocity  $V$  can be defined by the Fourier coefficients:

$$V(t) = \frac{2 \sum_{i=1}^N m_i v_{i,x}(t) \cos[kr_{i,z}(t)]}{\sum_{i=1}^N m_i} \quad (12)$$

where  $v_{i,x}$  is the component of velocity  $V$  in the  $X$  axis,  $r_{i,z}$  is the coordinate of the atom in  $Z$  axis, and  $m_i$  is the atomic mass.

The viscosity value of solution under electrostatic field with different intensity is shown in Fig. 11.

From the calculation results of the viscosity of Fig. 11, it is found that the viscosity of ferric chloride solution increases under the action of electrostatic field, and it is positive correlation between viscosity of the solution and electrostatic field intensity. As the electrostatic field intensity rises from 0 V m<sup>-1</sup> to 10<sup>8</sup> V m<sup>-1</sup>, the viscosity of the solution rises from 1.8 mPa to 1.832 mPa s. The viscosity of the solution is a parameter to measure the friction between the fluid layers in the solution system. The viscosity is a macroscopic property to reflect the

**Fig. 11** Viscosity of solution under different intensities of electrostatic field.

friction effect of all the particles in the solution, which is related to the number of hydrocarbon bonds in the solution and the molecular weight of the particles. Since there is no hydrocarbon bond in the solution system, the change of solution viscosity directly reflects the change of internal structure and kinetic parameters of the water solvent. The Bakker's study result<sup>38</sup> shows that the change of the viscosity of the ionic solution is directly related to the residence time of water molecules in the first water shell of the hydrated ion. If the residence time is prolonged, the viscosity of the solution will correspondingly increase.

According to the results in Fig. 10, the external electrostatic field can significantly prolong the residence time of water molecules, thus increasing the viscosity of the solution. In addition, the results of Fig. 11 are consistent with the calculation results of the diffusion coefficient of the water molecules under the electrostatic field in Fig. 9. There is a negative correlation between the diffusion of water molecules and the viscosity of the solution, which means the decrease of the diffusion of water molecules can lead to the increase of solution viscosity. The increase of solution viscosity is not conducive to the destabilization and collision aggregation of colloid particles during the later reaction. Therefore, viscosity increase should be relieved during the application of the technique of electrocoagulation, such as using the method of adjusting the reaction temperature.

In our previous work,<sup>32</sup> we discussed the impact of an electrostatic field and impurity  $Mg^{2+}$  ions on  $CaCl_2$  solution. The results in the previous paper shows that under electrostatic field the  $D_{Ca}$  increased, which meets with the  $D_{Fe}$  in the present article. But the viscosity trend in the two studies is opposite. The experimental results from previous ref. 39 showed that the electric field (magnetic field) can reduce surface tension of pure water, which means the electric field can reduce the viscosity of pure water. But the presence of metal cations will increase the viscosity of solution because of the bound effect to the water molecules. So there will be two opposite effects to the water molecules, one (from electric field) can increase the activity and the other (from metal cations) can impair the activity. The  $D_{H_2O}$  in the previous paper increased under electrostatic field, but decreases in the present article. This may because  $Fe^{3+}$  has a much bigger attraction to water molecule than  $Ca^{2+}$ , and electrostatic field enhances this attraction of  $Fe^{3+}$ , which means the inhibition effect gains the upper hand. Therefore, the lower  $D_{H_2O}$  under electrostatic field in the present study leads to a longer average residence time of  $H_2O$  molecules around  $Fe^{3+}$ , and the enhanced attraction from  $Fe^{3+}$  leads to a less number of hydrogen bonds between water molecules.

## Conclusions

In the present article, the change of microscopic properties of ferric chloride solution under electrostatic fields with different intensity was studied by equilibrium molecular dynamics simulations for investigating the mechanism of enhanced electrocoagulation technique. The results show that external electrostatic field can change the distribution characteristics of

water molecules around  $Fe^{3+}$ . When electrostatic field is present, the first peak position of  $Fe^{3+}-H_2O$  radial distribution function almost has no change, but the first peak value decreases, and the number of coordinated water molecules increases. The electrostatic field can change the diffusion characteristics of  $Fe^{3+}$  and water molecules in solution. Under the effect of electrostatic field, the self-diffusion coefficient of  $Fe^{3+}$  increases, but the self-diffusion coefficient of water molecules decreases, and the average residence time of water molecules in the first water shell of  $Fe^{3+}$  increases. The electrostatic field can change the structure of hydrogen bond network in solution. Under the effect of electrostatic field, the average lifetime of hydrogen bond in the solution has no obvious change, but the number of hydrogen bonds decreases, and the proportion of water molecules that do not participate in bonding and to participate in the formation of 1 hydrogen bond has increased. The change amplitude of the parameters above is bigger with the increase of electrostatic field intensity. The changes of these parameters show that the electrostatic field can enhance the interaction between  $Fe^{3+}$  and water molecules, and promote the hydrolysis of  $Fe^{3+}$  and the polymerization of ferric ion complex, which is beneficial to the enhancement of the effect of ferric chloride coagulant. In addition, the external electrostatic field increases the viscosity of the solution. This change is not conducive to the interaction of colloid particles in the solution. Therefore, the viscosity of solution should be controlled during the practical application of the technique of electrocoagulation.

## Conflicts of interest

There are no conflicts to declare.

## Greek symbols and nomenclature

$g_{A-B}$	Radial distribution function, nm
$n_{A-B}$	Particle B coordination number around particle A
$D$	Diffusion coefficient, $cm^2 s^{-1}$
$P$	First peak intensity of $g_{AB}$
$\gamma$	Residence time of water molecules, ps
$\tau_{HB}$	Hydrogen bond lifetime, ps
$C(\tau)$	Hydrogen bond correlation function
$h_i(t)$	Heaviside unit step function
$\rho$	Density of solution
$v$	Velocity of solution
$\eta$	Viscosity value of solution, $mP s^{-1}$

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## References

- 1 Y. Cui and J. Xiao, *Environ. Sci. Pollut. Res.*, 1998, **20**(3), 39.



- 2 G. Q. Fan, *China Water & Wastewater*, 1998, **14**(3), 42.
- 3 H. Zheng, T. Long and X. Shu, *Chongqing Environment Science*, 2000, **22**(3), 51.
- 4 O. T. Can, M. Bayramoglu and M. Kobya, *Ind. Eng. Chem. Res.*, 2002, **42**, 3391.
- 5 N. Daneshvar, A. H. Sorkhabi and M. B. Kasiri, *J. Hazard. Mater.*, 2004, **B1122**, 55.
- 6 L. S. Calvo, J. P. Leclerc and G. Tanguy, *Environ. Prog.*, 2003, **22**(1), 57.
- 7 M. F. Pouet and F. Persin, *Water Sci. Technol.*, 1992, **25**(12), 247.
- 8 N. V. Nikolaev, A. S. Kozlovskii and I. I. Utkin, *Sov. J. Water Chem. Technol.*, 1983, **4**(3), 70.
- 9 N. S. Abuzaid, A. A. Bukhari and Z. M. Hamouz, *Adv. Environ. Res.*, 2002, **6**(3), 325.
- 10 B. O. A. Hedstrom, *Ark. Kemi*, 1953, **5**, 1.
- 11 G. Biederman and P. Schindler, *Acta Chem. Scand. Acta Chem. Scand.*, 1957, **11**, 731.
- 12 V. D. Spoel, D. Lindahl, E. Hess, B. Groenhof, G. Mark and A. E. Berendsen, *J. Comput. Chem.*, 2005, **26**(16), 1701.
- 13 H. J. C. Berendsen, J. R. Grigera and T. P. Straatsma, *J. Phys. Chem.*, 1987, **91**, 6269.
- 14 Z. Zhang and Z. Duan, *Phys. Earth Planet. Inter.*, 2005, **149**(4), 335.
- 15 M. Li, Z. Duan and Z. Zhang, *Mol. Phys.*, 2008, **106**(24), 26857.
- 16 M. P. Allen and D. J. Tildesley, *Computer Simulation of Liquids*, Oxford Clarendon Press, 1989.
- 17 X. Li and Z. Yang, *J. Theor. Comput. Chem.*, 2006, **5**, 341.
- 18 S. Amira, D. Spangberg, M. Probst and K. Hermansson, *J. Phys. Chem. B*, 2004, **108**, 496.
- 19 A. J. Easteal, R. Malhotra, W. E. Price and L. A. Woolf, *J. Solution Chem.*, 1991, **20**(3), 319.
- 20 W. E. Price and L. A. Woolf, *J. Solution Chem.*, 1992, **21**(3), 239.
- 21 B. S. Brunschwig, C. Creutz, D. H. Macartney, T. K. Sham and N. Sutin, *Faraday Discuss. Chem. Soc.*, 1982, **74**, 113.
- 22 D. M. York, T. A. Darden and L. G. Pedersen, *J. Chem. Phys.*, 1993, **99**, 8345.
- 23 A. A. Chialvo and J. M. Simonson, *J. Chem. Phys.*, 2003, **119**, 8052.
- 24 C. M. Flynn, *Chem. Rev.*, 1984, **84**(1), 31.
- 25 J. Y. Bottero, A. Manceau, F. Villieras and D. Tchoubar, *Langmuir*, 1994, **10**(1), 316.
- 26 H. Zhang, G. A. Waychunas and J. F. Banfield, *J. Phys. Chem. B*, 2015, **119**, 10630.
- 27 T. Darden and D. J. York, *Chem. Phys.*, 1993, **98**, 1008.
- 28 H. J. Bakker, *Chem. Rev.*, 2008, **108**, 1456.
- 29 Y. Zhu, X. Lu and H. Ding, *Mol. Simul.*, 2003, **29**, 767.
- 30 M. Li, Z. Duan and Z. Zhang, *Mol. Phys.*, 2008, **106**, 2685.
- 31 M. J. Wei, L. Zhang and L. Lu, *Phys. Chem. Chem. Phys.*, 2012, **14**, 16536.
- 32 L. Zhu, Y. Han, C. Zhang, R. Zhao and S. Tang, *RSC Adv.*, 2017, **7**, 47583.
- 33 W. Hu, *Chemical engineering thermodynamics*, Chemical Industry Press, Beijing, 2001, p. 61.
- 34 X. Zhang, Q. Zhang and D. X. Zhao, *Acta Phys.-Chim. Sin.*, 2011, **27**(11), 2547.
- 35 A. Luzar and D. Chandler, *Nature*, 1996, **379**, 55.
- 36 W. F. V. Gunsteren, *Biomolecular Simulation: The GROMOS96 Manual and User Guide*, Groningen, 1996.
- 37 B. Hess, *J. Chem. Phys.*, 2002, **116**, 209.
- 38 H. J. Bakker, *Chem. Rev.*, 2008, **108**(4), 1456.
- 39 Y. I. Cho and S. H. Lee, *Int. J. Heat Mass Transfer*, 2005, **32**, 1.

