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# Comparison of two water oxidation electrocatalysts by copper or zinc supermolecule complexes based on porphyrin ligand†

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Based on the 5,10,15,20-tetra(cyanophenyl)porphyrin (CNTCPP) ligand, two supermolecule complexes formulated as [Cu(CNTCPP)]·2DMF (Cu-CNTCPP) and [Zn(CNTCPP)] (Zn-CNTCPP) have been synthesized and structurally characterized by single-crystal X-ray diffraction. Cu-CNTCPP features a 0 dimensional (0D) supramolecular structure, whereas Zn-CNTCPP is a 2D structure. With coordination unsaturated points, Cu-CNTCPP exhibits electrocatalytic activity toward oxygen evolution reaction (OER) with low potential in 1.0 M KOH, generating a current density of 10 mA cm<sup>-2</sup> at an potential of 1.66 V vs. RHE, which means it an efficient electrocatalytic material for OER.

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## Introduction

With increasing energy consumption and the greenhouse effect due to carbon dioxide emission, seeking renewable and non-carbon energy alternatives is an urgent and important issue.<sup>1,2</sup> The water splitting reaction provides a convenient chemical method for renewable energy storage.<sup>3,4</sup> As the oxidative half reaction of water splitting, the oxygen evolution reaction (OER) is kinetically sluggish and leads to significant overpotential and energy loss.<sup>5</sup> Developing highly efficient electrocatalysts for OER has emerged as one of the most research hot spots in clean energy generation and storage.<sup>6,7</sup> The most reported precious-metal electrocatalysts such as IrO<sub>2</sub> and RuO<sub>2</sub> are the state-of-art OER catalysts, but they are limited by low abundance and high cost for large-scale application.<sup>8,9</sup> Therefore, there is tremendous interest in developing inexpensive and earth-abundant OER catalysts.<sup>10</sup> Keeping the advantages of low cost, high chemical stability, and electrocatalytic ability, the most reported catalysts for OER are 3d transition metal (*e.g.* Fe, Cu, Co, Ni and Mn) derived oxides or hydroxides, phosphates,

perovskites, chalcogenides and so on,<sup>11–15</sup> but there are scarce reports of supermolecule complexes used as direct electrocatalyst for OER.<sup>16</sup>

On the other hand, supermolecule complexes are a class of interesting materials which have been intensively investigated in diverse fields, such as luminescence, gas storage, sensing, magnetics, catalysis, and so on.<sup>17–21</sup> The supermolecule complex structures and properties are mainly dependent on functional ligands and metal nodes.<sup>22–24</sup> As we know, porphyrins are valuable building ligands in supermolecule complex systems; their extended square planar ligand favors the formation of ordered frameworks.<sup>25–27</sup> Besides that, porphyrins can be metalated in their center, and could function as catalysis active centers,<sup>28–30</sup> so porphyrins or metallporphyrin molecules have been used as active constituents for constructing a variety of supermolecule complexes.<sup>31–33</sup> For example, reports on lithium-ion battery electrodes have been recorded based on iron-porphyrin materials.<sup>34–36</sup> Co-based porphyrin and graphitic materials was predicted to be the OER catalyst by density functional theory (DFT) calculation.<sup>37</sup> Analysis research of porphyrines as catalysts for electrochemical reduction of O<sub>2</sub> and oxidation of H<sub>2</sub>O have also been reported.<sup>38</sup> However, only a few supermolecule complexes show sufficient electrocatalytic OER activity at the high electric potential experienced during electrolysis (1.6–2.0 V vs. RHE).

Inspired by the above reports, we choose the 5,10,15,20-tetra(cyanophenyl)porphyrin (CNTCPP) ligand as organic linker for building supermolecule complexes based on the following reasons: (i) the ligand possesses delocalized  $\pi$ -electron system, which can provide an intense electrical conductivity;<sup>39</sup> (ii) four cyanide bridged coordination groups and the rigid phenyl in the ligand can afford more coordination opportunities, which will leading to diverse structures;<sup>40</sup> (iii)

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the electron withdrawing group of cyanide and high symmetry framework are favor for electron transferring in the ordered supermolecule structures; (iv) the central flat surface structure would help to increase the active surface area, which can ensure efficient contact between the electrolyte and active catalysts.<sup>41,42</sup> Herein, we demonstrate two supermolecule complexes [Cu(CNTCPP)]·2DMF (**Cu-CNTCPP**) and [Zn(CNTCPP)] (**Zn-CNTCPP**) by assembling the designed CNTCPP ligand with copper (II) or zinc(II) ions, respectively. **Cu-CNTCPP** are zero-dimensional (0D) complex with copper (II) metalated in the porphyrin center, **Zn-CNTCPP** is a 2D structure. Although **Zn-CNTCPP** has been reported by Jiang and coworkers, only its structure was focused upon in their work.<sup>43</sup> The OER properties of **Cu-CNTCPP** and **Zn-CNTCPP** have been investigated. Particularly, due to the different coordination situations, the **Cu-CNTCPP** materials afforded better performance than **Zn-CNTCPP** in 1.0 M KOH, indicating it a promising new class of non-noble catalyst for OER.

## Materials and general methods

All of the chemicals were obtained from commercial sources and were used without further purification, except CNTCPP ligand was synthesized through the condensation of 4-formylbenzonitrile with pyrrole, catalyzed by  $\text{BF}_3 \cdot \text{OEt}_2$ , followed by oxidation with *p*-chloranil (Fig. S1†).<sup>44</sup> C, H, N of elemental analysis were measured on an EA 1110 elemental analyzer. The powder diffraction data were collected on a Cu-K $\alpha$  radiation using an X-Pert PRO MPD diffractometer. Thermogravimetric analyses were performed in the temperature ranged from 40 to 800 °C with 10 °C min<sup>-1</sup> rate of heat using a Mettler Toledo Thermogravimetric analysis (TGA) instrument. The purity of **Cu-CNTCPP** and **Zn-CNTCPP** were confirmed by comparison of their simulated and experimental powder X-ray diffraction (Fig. S2 and S3†). TGA measurement shows that **Cu-CNTCPP** and **Zn-CNTCPP** lose uncoordinated solvents below 200 °C and can be stable up to 580 and 560 °C, respectively. (Fig. S4 and S5†). Method for electrochemical measurements are displayed in ESI.†

The X-ray datum of complexes were obtained on an Agilent Super nova with Cu-K $\alpha$  radiation ( $\lambda = 1.54178$ ) and MoK $\alpha$  ( $\lambda = 0.71073$ ) at room temperature, respectively. The absorption corrections were decided by employing the SADABS program. The structures and hydrogen atoms were refined to utilize the SHELX-97 program 33–34 through the full-matrix least-squares by fitting on F2 and anisotropic thermal parameters, respectively. There are many disordered solvent molecules existed in the cavity of **Cu-CNTCPP**, which can not be achieved through the reasonable modeling. Hence, the diffuse electron was removed by the PLATON/SQUEEZE routine 35. The summary of the structure data is displayed in Tables S11 and S12.† Crystallographic data have been deposited at the Cambridge Crystallographic Data Center (CCDC: 1450485 for **Cu-CNTCPP**, 1449406 for **Zn-CNTCPP**), which can be obtained from the Cambridge Crystallographic Data Center, 12, Union Road, Cambridge CB21EZ, U.K.

## Results and discussion

### Crystal Structure

Complex of **Cu-CNTCPP** crystallizes in the monoclinic space group  $P2_1/c$ , and the asymmetric unit contains two  $\text{Cu}^{2+}$  (Cu1 and Cu2) ions and two CNTCPP ligands. As displayed in Fig. 1a, Cu1 is surrounded by four N atoms from one CNTCPP ligands' central porphyrin ring with the average Cu–N bond length of 2.006 Å. Cu2 adopt similar coordination mode with the average Cu–N bond length of 2.003 Å. The non-coordinated sites of Cu1 and Cu2 in the axial direction mean the possible presence of large number of bridging and terminal hydroxyl or water ligands, which could provide high concentration of potential active sites. The angles between the side benzene rings and the central Cu1-porphyrin ring are 85.828°, 69.302°, 73.563°, and 68.479°, respectively, the angles between the side benzene rings and the central Cu2-porphyrin ring are 84.213°, 67.916°, 67.704°, and 68.709°, respectively (Fig. 1b). The CNTCPP porphyrin ring are metallic by  $\text{Cu}^{2+}$  ion, and all the four cyanide groups in the CNTCPP ligands are non-coordinated. The distance between Cu1 and Cu2 in the two-porphyrin ring is 8.599 Å. There are weak C–H...N (3.695 Å and 3.743 Å) interactions between the 0D Cu-porphyrin molecules to connect it into 3D structure (Fig. 1c and d). There is microspore view along *c* axis and the uncoordinated DMF molecules blocked the pores.

X-ray single-crystal diffraction reveals complex of **Zn-CNTCPP** crystallizes in the monoclinic space group  $P2_1/n$ , and the asymmetric unit contains one  $\text{Zn}^{2+}$  ion and one CNTCPP ligand. As displayed in Fig. 2a,  $\text{Zn}^{2+}$  is surrounded by four N atoms from one CNTCPP ligands' central porphyrin ring with the average Zn–N bond length of 2.046 Å and there are weak interactions between the two cyanide groups from two different CNTCPP ligands with the  $\text{Zn}^{2+}$  ions (Zn–N bond length of 2.645 Å). The CNTCPP porphyrin ring central are metallic by zinc ion to form 0D molecule, and the weak Zn–N bond between the two of cyanide groups in the CNTCPP ligands connect another two zinc(II) ions from the metallic porphyrin ring to form 2D

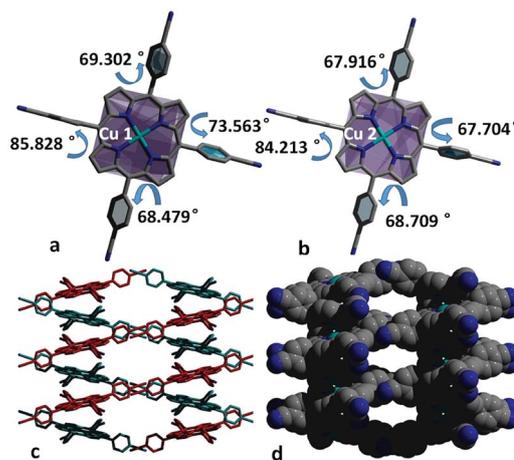


Fig. 1 (a and b) The angles between the side benzene rings and the central Cu1 and Cu2-porphyrin ring. (c and d) 3D framework viewed from *c* axis.



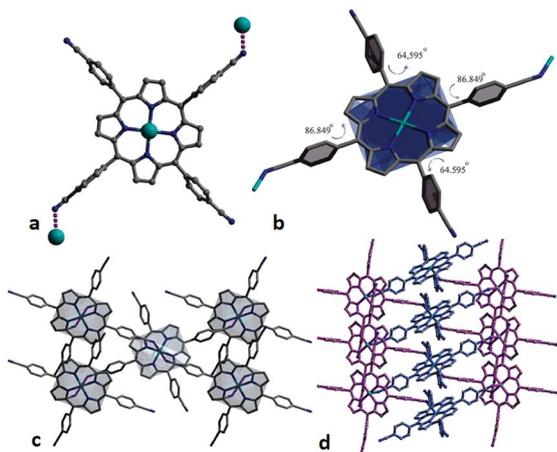


Fig. 2 (a) Coordination environment of the ligand and Zn(II) in complex. (b) Angles between the side benzene rings and the central Zn-porphyrin ring. (c) 2D layer viewed from *b* axis. (d) 3D space filling representation viewed from *a* axis.

supramolecular structure. The angles between the side benzene rings and the central Zn-porphyrin ring are 86.849°, 64.595°, 86.849°, and 64.595°, respectively (Fig. 2b). Each ligand connects three Zn(II) ions, and each Zn(II) attaches three different CNTCPP ligands, generating a 2D network (Fig. 2c). There are weak C1–H...N8 (3.259 Å) supramolecular hydrogen bonds between the 2D layers to connect it into 3D structures. View along *b* axis, there is micro-pores with dimension of 4.5 × 8.2 nm (Fig. 2d). The crystallographic data of the two complexes are summarized in Tables S1 and S2.†

### Electrocatalytic activity

Many researchs show promising electrocatalytic properties for a myriad of electrochemical reactions such as OER. Different studies reported various optimal compositions of copper, to achieve maximum OER performance. But slightly literatures reported the positive result got by supramolecular with copper structures. As the Cu<sup>2+</sup> ions are unsaturated coordinated in **Cu-CNTCPP**, there are enriched active sites, we therefore investigate the electrocatalytic activity of the **Cu-CNTCPP**. It is necessary to understand the origination of the OER activity of the as-prepared catalyst, for comparison, we have also measured the electrocatalytic OER activity of the comparative catalysts under the same conditions except **Cu-CNTCPP**: CNTCPP ligand, mixture of CuCl<sub>2</sub> and CNTCPP ligand by manual grinding (named mixture of CNTCPP and Cu<sup>2+</sup>), mixture of CuCl<sub>2</sub> and CNTCPP ligand by reflux in hot DMF solution (named amorphous **Cu-CNTCPP**), and Cu(OH)<sub>2</sub> by conventional precipitin reaction (named Cu(OH)<sub>2</sub>). Fig. 3a shows the polarization curves of **Cu-CNTCPP** and the comparative catalysts. The potential required for a current density of 10 mA cm<sup>-2</sup>, which is a metric associated with solar fuel synthesis.<sup>45</sup> As shown, the **Cu-CNTCPP** delivers a current density of 10 mA cm<sup>-2</sup> at a potential of 1.66 V, whereas a potential as high as 1.74 V, 1.84 V and 1.78 V are required for CNTCPP ligand, mixture of CNTCPP and Cu<sup>2+</sup> and amorphous **Cu-CNTCPP** to achieve the same current density, respectively. And for Cu(OH)<sub>2</sub>, it fails to reach 10 mA

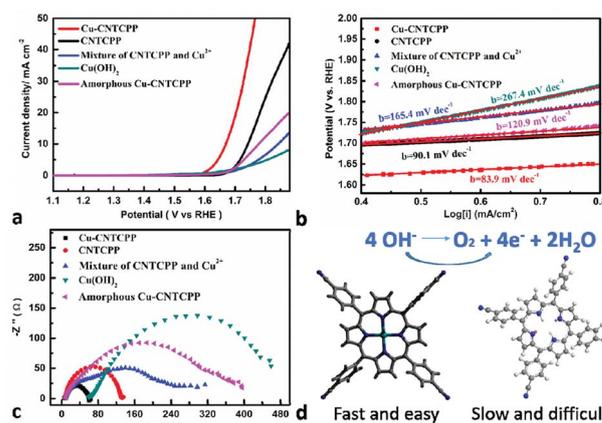


Fig. 3 Electrocatalytic activity of the **Cu-CNTCPP** for OER. (a) Polarization curves. (b) Tafel plots in 1.0 M KOH solution (5 mV s<sup>-1</sup>). (c) Nyquist plots examined at 0.5 V (vs. SCE). (d) Comparison of coordinated and non-coordinated CNTCPP.

cm<sup>-2</sup> in 1.0 M KOH. The OER kinetics of these electrodes is estimated by their corresponding Tafel plots. As shown in Fig. 3b, the Tafel slope of the **Cu-CNTCPP** is 83.9 mV dec<sup>-1</sup>, compared to 90.1 mV dec<sup>-1</sup>, 165.4 mV dec<sup>-1</sup>, 120.9 mV dec<sup>-1</sup> and 267.4 mV dec<sup>-1</sup> for CNTCPP, mixture of CNTCPP and Cu<sup>2+</sup>, amorphous **Cu-CNTCPP** and Cu(OH)<sub>2</sub>, respectively.

To get further insight into the activity of as-synthesized **Cu-CNTCPP** modified electrodes toward OER, electrochemical impedance spectroscopy (EIS) analysis was also performed. Fig. 3c describe the obtained Nyquist plots of **Cu-CNTCPP** and the comparative catalysts, the charge-transfer resistance (*R*<sub>ct</sub>) at the surface of the catalysts is determined from the diameter of a semicircle at high frequencies in the Nyquist plot. Generally, *R*<sub>ct</sub> value varies inversely with the electrocatalytic activity. That is, smaller diameter corresponds to faster OER kinetics. The *R*<sub>ct</sub> values of **Cu-CNTCPP** (60 Ω) are much lower than the other catalysts. Thus, such a low *R*<sub>ct</sub> value of **Cu-CNTCPP** indicates that its high electrocatalytic activity for OER could be ascribed to the highly conductive Cu<sup>2+</sup> hybrid CNTCPP ligands improving the charge transfer characteristics of **Cu-CNTCPP**. All the results show that the proprieties of **Cu-CNTCPP** are better than the comparative catalysts. As we know the electrocatalytic reaction is primarily a surface phenomenon, the axial positions of coordinated Cu<sup>2+</sup> ions are naked, so it could offer sufficient active sites for OER. In the long-range ordered crystals of **Cu-CNTCPP**, CNTCPP is supposed to promote electron transport at heterogeneous surface due to its excellent electrical conductive property, which act as an ideal platform for loading Cu<sup>2+</sup>, so it showed higher catalytic activity than pure CNTCPP ligands, mixture of CNTCPP and Cu<sup>2+</sup> and amorphous **Cu-CNTCPP**.

From the crystal structure of **Zn-CNTCPP**, there are weak interactions between the two cyanide groups from two different CNTCPP ligands with the Zn<sup>2+</sup> ions (Zn–N bond length of 2.645 Å), so it will partially block the catalytic active sites of Zn ions, and the OER performance of it also verified our speculation. Similar with **Cu-CNTCPP**, we have also measured the electrocatalytic OER activity of **Zn-CNTCPP** and the comparative catalysts under the same conditions for comparison: ZnCl<sub>2</sub> and



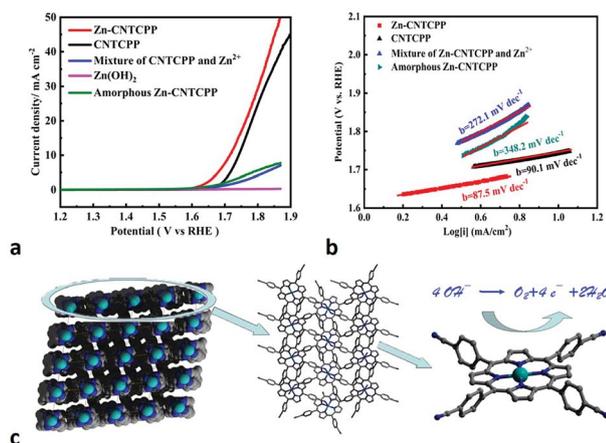


Fig. 4 Electrochemical activity of the Zn-CNTCPP for OER. (a) Polarization curves. (b) Tafel plots in 1.0 M KOH solution ( $5 \text{ mV s}^{-1}$ ). (c) Comparison of coordinated and non-coordinated CNTCPP.

CNTCPP ligand mixture by manual grinding (named mixture of CNTCPP and  $\text{Zn}^{2+}$ ),  $\text{ZnCl}_2$  and CNTCPP ligand mixture by reflux in hot DMF solution (named amorphous Zn-CNTCPP), and  $\text{Zn}(\text{OH})_2$  by conventional precipitation reaction (named  $\text{Zn}(\text{OH})_2$ ). Fig. 4a shows the polarization curves of Zn-CNTCPP and the comparative catalysts. As shown, the Zn-CNTCPP deliver a current density of  $10 \text{ mA cm}^{-2}$  at a potential of 1.71 V, whereas a potential as high as 1.74 V are required for CNTCPP ligand to achieve the same current density. And for mixture of CNTCPP and  $\text{Zn}^{2+}$ , amorphous Zn-CNTCPP and  $\text{Zn}(\text{OH})_2$ , it fails to reach  $10 \text{ mA cm}^{-2}$  in 1.0 M KOH., the Tafel slope of the Zn-CNTCPP is  $87.5 \text{ mV dec}^{-1}$ , compared to  $272.1 \text{ mV dec}^{-1}$  and  $348.2 \text{ mV dec}^{-1}$  for mixture of CNTCPP and  $\text{Zn}^{2+}$ , amorphous Zn-CNTCPP, respectively (Fig. 4b). The  $R_{ct}$  values of Zn-CNTCPP are  $140 \Omega$  (Fig. S6<sup>†</sup>). All the results show the OER properties of Zn-CNTCPP are inferior to Cu-CNTCPP, that coincides with our speculation, that is the weak interactions of Zn–N bond of central Zn(II) ions partially block the catalytic active sites of Zn ions, but the axial positions of coordinated  $\text{Cu}^{2+}$  ions are naked, so it offer sufficient active sites for OER and afford more convenient electrocatalytic reaction.

The electrochemical double-layer capacitance ( $C_{dl}$ ) was investigated under non-faradaic condition to estimate the electrochemical surface area (ECSA). The cyclic voltammograms at various scan rates of the Cu-CNTCPP, Zn-CNTCPP and CNTCPP were shown in Fig. 5a–c, according to which the  $C_{dl}$  was calculated. As shown in Fig. 5d, the  $C_{dl}$  of the Cu-CNTCPP, Zn-CNTCPP and CNTCPP are  $0.531$ ,  $0.254$  and  $0.104 \text{ mF cm}^{-2}$ , suggesting that the Cu-CNTCPP may have an advantage in active surface area and active sites over others for oxygen production. Catalytic activity and stability towards OER are two valid parameters for oxygen production. In order to investigate the stability in alkaline solution of the prepared Cu-CNTCPP and Zn-CNTCPP, continuous CV was performed within the potential range of 1.45 to 1.7 V (vs. RHE) for 2000 cycles at scan rate of  $50 \text{ mV s}^{-1}$ . As shown in Fig. 6a and c, clearly, Cu-CNTCPP exhibits slight loss compared with Zn-CNTCPP. Furthermore, the continuous electrolysis at different current density were

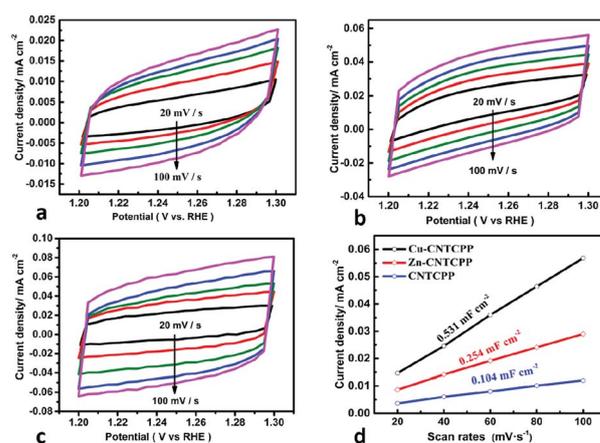


Fig. 5 Cyclic voltammograms (CV) of (a) Cu-CNTCPP. (b) Zn-CNTCPP. (c) CNTCPP. (d) Estimated  $C_{dl}$ .

monitored over 10 h. Moreover, taking further insight from the  $i$ - $t$  curves obtained at  $10 \text{ mA cm}^{-2}$ , it can be seen that the Cu-CNTCPP keeps approximately 83% compared with the Zn-CNTCPP of 54%. Notably, Cu-CNTCPP exhibits a slight variation over the entire time, indicating the outstanding stability. (Fig. 6b and d). In a word, Cu-CNTCPP with lower potential, Tafel slope and higher estimated  $C_{dl}$ , compared to others as depicted in Table 1, has the potential to be developed to be superior catalytic oxygen production from electrolyzed water.

In a word, the OER performance of Zn-CNTCPP with coordination saturation is inferior to the Cu-CNTCPP with coordination unsaturation. The main reason for the difference between Cu-CNTCPP and Zn-CNTCPP in OER performance is due to the presence of weak Zn–N bonds of Zn-CNTCPP, which will partially block the catalytic active sites of Zn ions. Based on the structures and its outstanding OER performance of Cu-

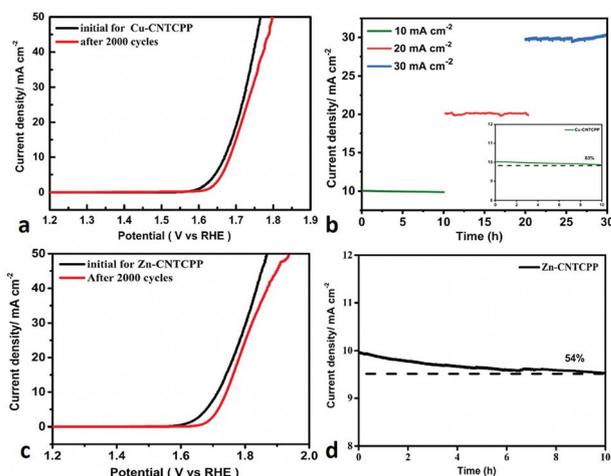


Fig. 6 (a) Polarization curves of Cu-CNTCPP, for initially and after 2000 potential sweeps. (b) Long-term stability test carried out at constant current densities of 10, 20,  $30 \text{ mA cm}^{-2}$  for Cu-CNTCPP (inset: time dependent current density curves at current densities of  $10 \text{ mA cm}^{-2}$ ). (c) Polarization curves of Zn-CNTCPP, for initially and after 2000 potential sweeps. (d) Time dependent current density curves for Zn-CNTCPP.

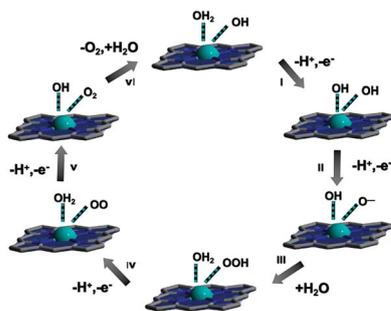


Table 1 Comparison for different non-precious metal catalysts in 1.0 M KOH for OER

Electrocatalyst	Electrolytes	$\eta_{10}$ (mV)	Tafel slope [mV dec <sup>-1</sup> ]	References
<b>Cu-CNTCPP</b>	<b>1 M KOH</b>	430	83.9	<b>This work</b>
<b>Zn-CNTCPP</b>	<b>1 M KOH</b>	480	87.5	<b>This work</b>
CNTCPP	1 M KOH	510	90.1	<b>This work</b>
Mn <sub>3</sub> O <sub>4</sub> /CoSe <sub>2</sub>	1 M KOH	450	49	1 (ref. 46)
Zn-Co-LDH	0.1 M KOH	460	101	2 (ref. 47)
Fe doped Co <sub>3</sub> O <sub>4</sub>	1 M KOH	370	60	3 (ref. 48)
Fe <sub>3</sub> O <sub>4</sub> @Co <sub>9</sub> S <sub>8</sub> /rGO-2	1 M KOH	400	54.5	4 (ref. 49)

CNTCPP, we supposed its reaction mechanism in the OER process, as the axial direction of Cu<sup>2+</sup> ions are unsaturated-coordinated, which provides mainly catalytically active sites used for fast dissipation of the electrons generated during OER (Scheme 1): a Cu atom with an oxidation state of 2+ serves as the center, it could be connected by solvent -OH and H<sub>2</sub>O ligands in 1.0 M KOH solution. As H<sub>2</sub>O adsorption on the surface is favorable thermodynamically, the OER directly splits H<sub>2</sub>O on a Cu<sup>2+</sup> site to produce an adsorbed -OH. -OH then loses a proton to form an O atom, so a Cu-O<sup>-</sup> group would form (steps I and II). After the nucleophilic attack of a solvent molecule on the Cu-O<sup>-</sup> group, a hydroperoxo CuOOH group was yielded (step III). Then a superoxo (step IV) and molecular O<sub>2</sub> (step V) could be formed. Finally, O<sub>2</sub> is evolved, and the catalyst is recovered when a H<sub>2</sub>O molecule connects to the Cu site (step VI).

In summary, complexes of Cu-CNTCPP and Zn-CNTCPP have been successfully synthesized and used as OER catalyst by utilizing of conductive CNTCPP porphyrin ligands. As an advanced OER catalyst, Cu-CNTCPP and Zn-CNTCPP catalyst exhibit a low potential of 1.66 V and 1.71 V. This designed 0D Cu-CNTCPP supramolecular structure possesses plenty of catalytic active sites and high electrical conductivity, which can lead to better OER activity. Compared with those non-noble-metal electrocatalytic materials reported recently, Cu-CNTCPP electrode presents a quite large exchange current density and a relatively small Tafel slope, suggesting it is a promising low-cost and earth-abundant metallic electrocatalyst for OER. We anticipate that the design of the OER electrocatalyst will lead improved strategies and broad the scope for future exploration in MOFs fields.



Scheme 1 Proposed reaction mechanism of Cu-CNTCPP for OER.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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