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In situ fabrication of a direct Z-scheme photocatalyst by immobilizing CdS quantum dots in the channels of graphene-hybridized and supported mesoporous titanium nanocrystals for high photocatalytic performance under visible light†

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We report the considerable advantages of direct Z-scheme photocatalysts by immobilizing high-quality CdS quantum dots (QDs) in the channels of graphene-hybridized and supported mesoporous titania (GMT) nanocrystals (CdS@GMT/GR) under facile hydrothermal conditions. The photocatalysts have been characterized by XRD, PL, XPS, SEM, DRS, TEM, EIS, and N₂ adsorption. CdS QDs primarily serve as photosensitizers with a unique pore-embedded structure for the effective utilization of the light source. This direct Z-scheme CdS@GMT/GR exhibits higher photocatalytic activity than CdS/GR, GMT/GR, or CdS@MT. In addition, the rate constant of CdS@GMT/GR-2 is approximately twice the sum of those of CdS@MT and GMT/GR, because GR played the role of hole-transporting and collection layer as well as the hybridization level formation in terms of hybridizing MT and serving as a support. Therefore, the GR content tunes the energy band, affects the surface area, and controls the interfacial hole transfer and collection rate of the direct Z-scheme system. Furthermore, CdS@GMT/GR retains its high performance in repeated photocatalytic processes. This can be attributed to the fact that GR prevents QDs from photocorrosion by means of the hole-transporting and collection effect. A possible reaction mechanism is proposed. This work provides a promising strategy for the construction of highly efficient visible-light-driven photocatalysts to reduce the growing menace of environmental pollution.

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1. Introduction

Over the past two decades, considerable efforts have been expended in developing efficient and stable semiconductors for photocatalysis under solar irradiation to reduce the growing menace of environmental pollution.^{1–7} Among sensitized solar materials, quantum dots (QDs) have gained particular interest to serve as photosensitizers due to their combined advantages of low cost, mechanical flexibility, solution processability, bandgap tunability, suitability for large-surface-area film fabrication, and designability.^{4–7} Due to these benefits, QD-sensitized solar catalysts (QDSSCs) have attracted widespread attention. A plethora of architectures for QDSSCs employing inorganic QDs, such as CdS,⁸ CdSe,⁹ PbS,¹⁰ InP,¹¹ PbSe,¹² and InAs,¹³ adsorbed onto wide-bandgap semiconducting oxides like ZnO, TiO₂, and SnO₂, have hitherto been proposed, and

their applicability to visible-light harvesting has been investigated. Among QD materials, CdS has been one of the most representative materials as a stable and efficient photoabsorber because its optical bandgap can be readily tuned from 0.3 to 1.6 eV and its response in the near-infrared region makes CdS QDs suitable for a wide range of applications.^{8,14,15} Meanwhile, TiO₂ has been recognized as one of the most promising semiconductors due to its unique properties, such as strong oxidation ability, biological and chemical inertness, low cost, *etc.*^{1,2} Several groups have reported the use of CdS QDs as stable and efficient photoabsorbers in TiO₂-based heterojunction systems.^{16,17} However, in traditional heterojunction systems, there remains a disadvantage in which the oxidation and reduction abilities derived from photoinduced holes and electrons, respectively, are weakened after charge transfer.¹⁸ However, in a Z-scheme system, the strong redox potential of photogenerated charge carriers can also be retained.^{19–21} However, the photocatalytic ability of Z-scheme photocatalysts is also limited by the fast recombination of the photoinduced electron-hole pairs, and enormous efforts have been made to overcome this issue. One approach to alleviate the problem of

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poor electron transport and charge separation in the crystal and/or on the surface of Z-scheme photocatalysts involves the tethering of graphene (GR) to a semiconducting oxide couple.^{22,23} Some previous reports on GR-based Z-scheme photocatalysts,^{24–27} such as metal sulfides/GR/CoO_x-loaded BiVO₄,²⁴ TiO₂/CdSe QDs/functionalized graphene oxide (GO) nano-sheets,²⁵ TiO₂ nanotube array/GR/CdS QDs,²⁶ and a CdS/reduced GO/TiO₂ core-shell nanostructure,²⁷ have successfully demonstrated the role of GR in enhancing charge transport and transfer. Despite several reports on TiO₂, CdS QDs, and GR composites in Z-scheme systems, GR was selected as a solid electron mediator for promoting electron transport and to serve as an electron sink without a direct Z-scheme.^{26,27} It is well known that a third-generation Z-scheme photocatalytic system, namely, the direct Z-scheme photocatalyst, has more advantages than the previous two generations.^{21,28} However, to the best of our knowledge, the fabrication of a CdS@GMT/GR direct Z-scheme photocatalyst with QDs confined in the channels of GMT, with GR as a hole-transporting and collection layer in addition to hybridizing MT and serving as a support, has not hitherto been reported. Meanwhile, the application of Z-scheme photocatalytic systems in the photocatalytic degradation of pollutants has not yet received widespread attention.

2. Experimental section

2.1. Materials

Butyl titanate, hydrochloric acid (HCl), 1-hexadecyl-3-methyl imidazolium bromide (C₂₀H₃₉N₂Br, >98%) ionic liquid (IL), sodium sulfide (Na₂S), cadmium nitrate (Cd(NO₃)₂), golden orange II (GOII), and terephthalic acid (TA) were purchased from Jishou Chemical Glass Instrument Co., Ltd., as analytically pure reagents. GO and deionized water were prepared in the laboratory.

2.2. Synthesis

Preparation of the direct Z-scheme CdS@GMT/GR photocatalyst is shown in Scheme 1.

Synthesis of CdS QDs and their modification with IL (IL-CdS). CdS QDs were synthesized by an ion precipitation method.¹¹ In a typical synthesis process, Na₂S and Cd(NO₃)₂ were each dispersed in water at 0.01 mol L⁻¹; then, the respective solutions were adjusted to pH 2.0 with HCl. After stirring for more than 30 min at room temperature, the Cd(NO₃)₂ solution was further stirred at 100 °C by immersing

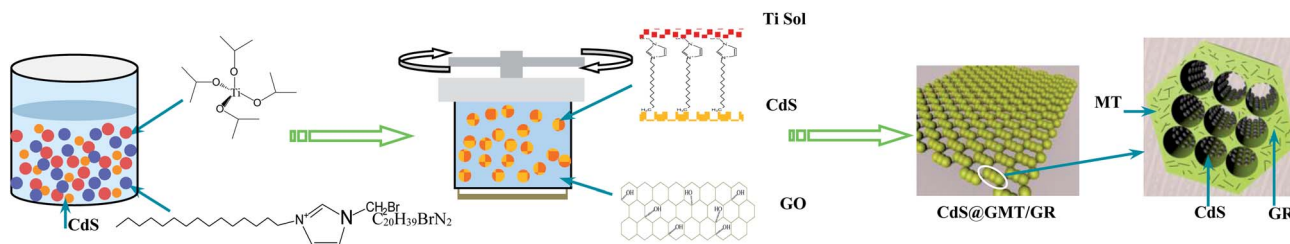
the flask in an oil bath. Thereafter, the Na₂S solution was added dropwise to the Cd(NO₃)₂ solution with stirring over a period of 30 min. After cooling to room temperature, the product was purified by repeated precipitation and centrifugation with ethanol and deionized water. Finally, the obtained powder was dried at 60 °C overnight. Then, 1-hexadecyl-3-methyl imidazolium bromide (C₂₀H₃₉BrN₂) IL is dissolved in water at a concentration of 10 g L⁻¹. Further, the prepared CdS QD water solution was mixed in a 1 : 1 weight ratio and placed in a rectangular glass cell, and the mixture was gently stirred at 0 °C in an ice-water bath to form IL-CdS.

Synthesis of CdS@MT. Titanium sol was prepared by the same method as reported earlier.² Subsequently, the requisite amount of the prepared IL-CdS solution was added dropwise to the stirred titanium sol to form a Ti-coated IL-CdS sol. The mixture was then transferred to a 50 ml Teflon-lined stainless steel autoclave and subjected to solvothermal reaction at 180 °C for 12 h. The product was then cooled to room temperature, separated by centrifugation, and washed three times with acetone and once with absolute ethanol. It was then dried in an electric oven at 80 °C for 12 h and calcined at 400 °C in air for 1 h, yielding CdS@MT nanocomposites with QD weight ratios of 2%.

Synthesis of GO, CdS@GMT/GR, GMT/GR, and CdS/GR. GO was synthesized from natural graphite powder by a modified Hummers method.^{29,30} A certain amount of GO was dispersed in deionized water (30 ml) and stirred for 1 h to obtain a uniform suspension. The Ti-coated IL-CdS sol, with the same CdS content as CdS@MT (2%), was then added to the suspension, and stirring was continued for a further 0.5 h. The obtained mixture was transferred to a 50 ml Teflon-lined autoclave and maintained at 180 °C for 12 h. Following the same steps as those for the CdS@MT synthesis, a series of CdS@GMT/GR samples with theoretical 90 wt%, 85 wt%, and 80 wt% GR were obtained, denoted as CdS@GMT/GR-1, CdS@GMT/GR-2, and CdS@GMT/GR-3, respectively. The synthesis methods for GMT/GR and CdS/GR were the same as that for CdS@GMT/GR, and the ratios of GMT to GR in GMT/GR and of CdS to GR in CdS/GR were also the same as that in CdS@GMT/GR.

2.3. Characterization

Wide-angle X-ray diffraction (WAXRD) and small-angle X-ray scattering (SAXS) patterns were obtained on a Bruker D₂ Phaser diffractometer using Cu-K_α radiation ($\lambda = 1.5418 \text{ \AA}$) and a Nanostar U SAXS system (Bruker, Germany) using Cu-K_α



Scheme 1 Synthesis of a direct Z-scheme CdS@GMT/GR photocatalyst.



radiation (40 kV, 35 mA), respectively. SEM images were obtained with the Hitachi S3400N (Japan) instrument. High-resolution transmission electron microscopy/selected-area electron diffraction (HRTEM/SAED) was performed on a JEOL JEM2100 (Japan) instrument. X-ray photoelectron spectroscopy (XPS) analysis was carried out on a Thermo Fisher Scientific VG Multilab 2000 X (US) instrument. Ultraviolet/visible (UV-vis) diffuse-reflectance spectra (DRS) were recorded on a spectrophotometer (Agilent, Inc., Cary 5000, US). Photoluminescence (PL) spectra were measured on a spectrometer (Edinburgh FLS980, UK). Raman spectra were obtained from a Thermo Fisher Scientific DXR2xi (US) instrument. EIS experiments were conducted from 0.1 Hz to 100 kHz in order to investigate the internal resistance and capacity of the electrode material. Brunauer–Emmett–Teller (BET) technique was used to determine the surface area and pore size distribution by recording nitrogen adsorption–desorption isotherms (ASAP2010, Micromeritics Co., US) at 77 K.

2.4. Photocatalytic measurements

The photocatalytic activities of the prepared photocatalysts were examined through the decomposition of GOII under visible-light irradiation (a 300 W high-pressure Hg lamp with a UV filter). In a typical photocatalytic reaction, GOII (10 mg) was dissolved in deionized water (20 ml) containing 1 g l⁻¹ catalyst. The photocatalytic process and reactor were described in an earlier report.²

2.5. Analysis of hydroxyl radicals

Hydroxyl radicals ($\cdot\text{OH}$) were detected by the PL method using TA as a probe molecule. In a typical procedure, 10 mg of the respective photocatalyst was dispersed in 10 ml of an aqueous solution of 1.5×10^{-4} M TA and 2×10^{-3} M NaOH. The photocatalytic reactor was described in an earlier report.²

3. Results and discussion

3.1. Morphology and structural characterization

The XRD patterns are shown in Fig. 1. Fig. 1a shows three diffraction peaks at 26.5, 44.0, and 52.1°, which are in good agreement with the main (111) reflection and two weaker (220)

and (311) reflections of the standard cubic CdS (JCPDS: 80-0019), confirming the crystalline phase of the QDs in CdS/GR.³¹ In addition, the XRD patterns also reveal that the principal peaks, consistent with the crystalline phases of (101), (004), (200), (105), and (211), were found for both the CdS@GMT/GR-2 and GMT/GR nanocomposites, indicating that the obtained samples primarily consisted of the anatase phase. Meanwhile, there are no the peaks corresponding to CdS for CdS@GMT/GR-2. This phenomenon may be attributed to the small amount of CdS, Cd²⁺ in place of Ti⁴⁺ of the anatase crystal structure, and both peaks overlap for CdS (111) and anatase (101) located at $2\theta = 26.5^\circ$.³² The average crystallite sizes of anatase and QDs, calculated from the main diffraction peaks of anatase (101) and CdS (002) using Scherrer's equation, are listed in Table 1. The crystallite sizes of 19.6 and 17.2 nm were calculated for GMT in the CdS@GMT/GR-2 and GMT/GR nanocomposites, respectively. These results reveal that the incorporation of CdS QDs could effectively promote crystallite growth. All the CdS@GMT/GR samples had almost the same crystallite sizes (about 19.6 nm for anatase; 5.5 nm for QDs), indicating that the amount of GR has no obvious influence on the crystallite size, as revealed by the TEM results shown in Fig. 4 (*vide infra*). The SAXS measurements of the obtained samples are shown in Fig. 1b. Pure MT showed one intense diffraction peak at $2\theta \approx 0.55^\circ$ due to the removal of the organic group from the IL

Table 1 Average crystal size and elemental content of the obtained samples

Samples	Average crystal size ^a (nm)		Elemental content ^b (%)				
	anatase	QDs	Ti	C	Cd	S	O
MT	17.7	—	32.7	1.5	—	—	65.8
GMT/GR	17.2	—	4.3	87.1	—	—	8.6
CdS/GR	—	9.1	—	92.8	3.0	3.2	1.0
CdS@GMT/GR-2	19.6	8.5	11.2	44.1	3.2	3.3	38.2
CdS@GMT/GR-1	19.7	8.4	10.3	47.7	2.9	3.1	36.0
CdS@GMT/GR-3	19.4	8.5	12.7	40.0	3.9	4.1	40.3

^a Average crystal size calculated from the Scherrer's equation.

^b Elemental content (%) represents the ratio $A_i/\sum A_i$ (A_i is the area of each peak) obtained by EDS.

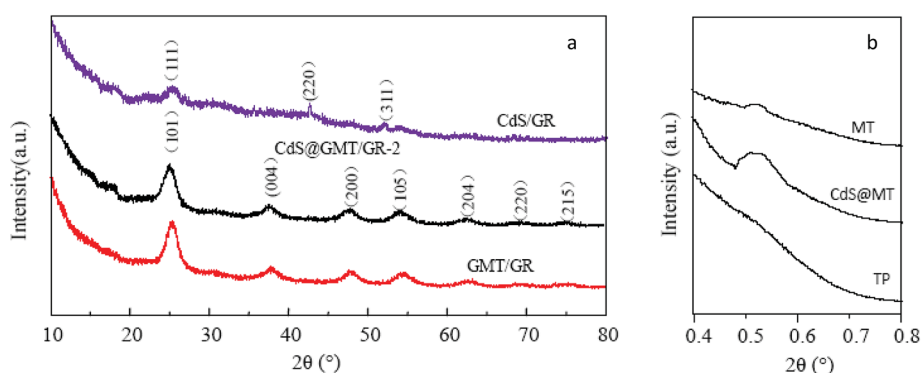


Fig. 1 WAXRD patterns (a) and SAXS patterns (b) of the obtained samples.



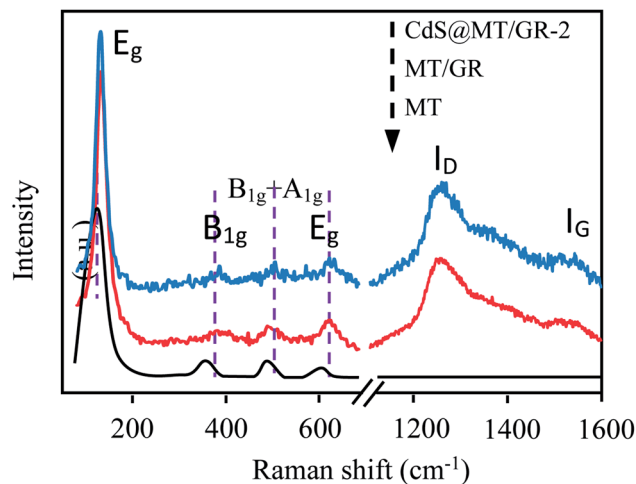


Fig. 2 Raman spectra of the obtained samples.

template (ESI 1†), indicating that a well-organized mesostructure had been formed. CdS@GMT/GR-2 still produced a (100) peak, indicating that the well-organized mesostructure was preserved during the embedding of QDs and the hybridization with GR, although it was broadened, suggesting some degradation of this mesostructure. No SAXS diffraction peak was detectable for TiO₂ nanoparticles (TN) prepared by the same method, but without the IL template. The SAXS results demonstrate that the current approach for introducing GR and CdS QDs into MT is effective, with at least the partial preservation of mesoporosity.

As shown in Fig. 2, the three samples exhibit four prominent Raman bands at *ca.* 140, 393, 513, and 633 cm⁻¹, corresponding to the E_g, B_{1g}, B_{1g} + A_{1g}, and E_g modes, respectively.^{33,34}

These bands can be attributed to the typical anatase phase of MT.^{33,34} Therefore, the Raman spectra are consistent with the XRD results shown in Fig. 1. Moreover, besides the predominant MT features, two prominent D (1282 cm⁻¹) and G (1566 cm⁻¹) peaks are observed for the GMT/GR and CdS@GMT/GR-2 samples, attributable to the breathing mode of the *k*-point phonons of the A_{1g} symmetry and E_{2g} phonons of sp² C atoms, respectively.^{33,35} When compared with the intensity ratio (*I*_D/*I*_G) of bare GO nanosheets (0.81) (ESI 2†), the *I*_D/*I*_G ratios for the GR in GMT/GR and CdS@GMT/GR-2 were evaluated to be about 1.33 and 1.57, respectively. The increase in *I*_D/*I*_G may be attributed to the reduction and restoration of the sp² network of GO after the hydrothermal reduction process.^{33,36} Interestingly, *I*_D/*I*_G for CdS@GMT/GR was found to be greater than that for GMT/GR, confirming the facile regeneration of the conjugated network in the former. Moreover, the four prominent peaks of the E_g, B_{1g}, B_{1g} + A_{1g}, and E_g modes for the GMT component in GR-based samples shifted to higher frequencies when compared with those for bare MT due to the phonon confinement effect between TiO₂ and GR.³⁷ No obvious Raman peak of CdS phase was seen in CdS@GMT/GR due to its quantum size effect and its confinement in the channels of MT. However, the presence of CdS could be verified by TEM and XPS measurements.

As shown in Fig. 3a, the nanoparticles were evenly spread on the GR surface in CdS@GMT/GR-2. From Fig. 3b, it is evident that there was no apparent difference in the morphology

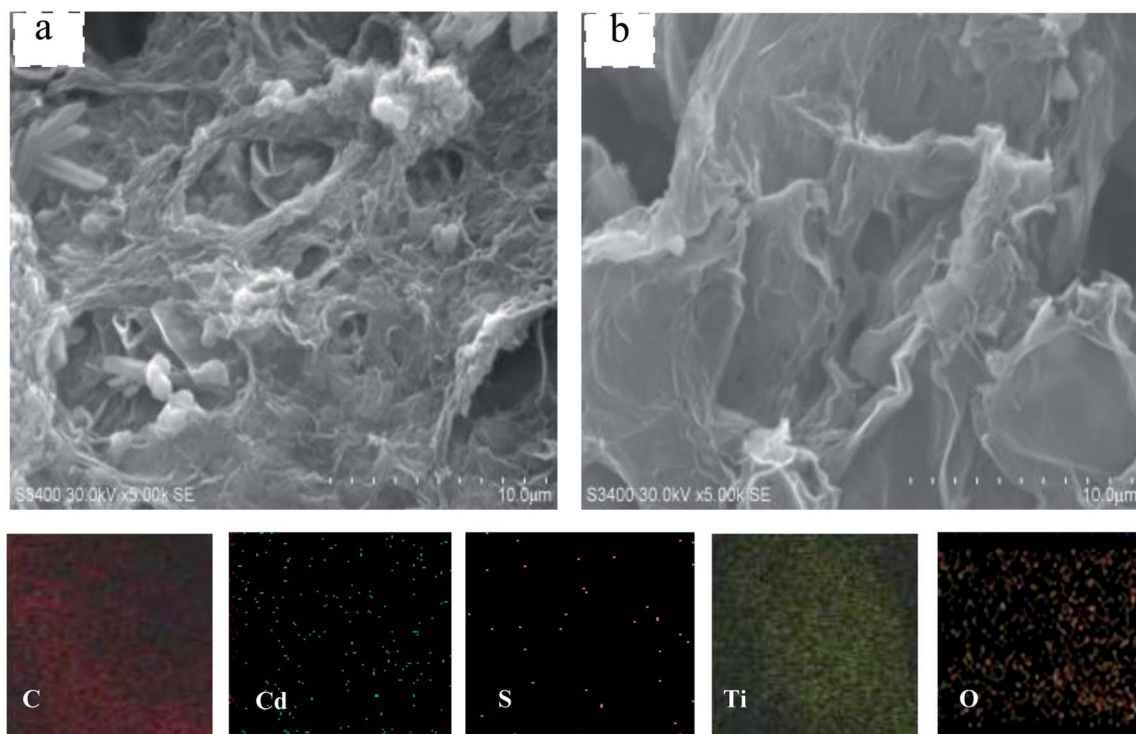


Fig. 3 SEM of the CdS@GMT/GR-2 (a) and pure GR (b).



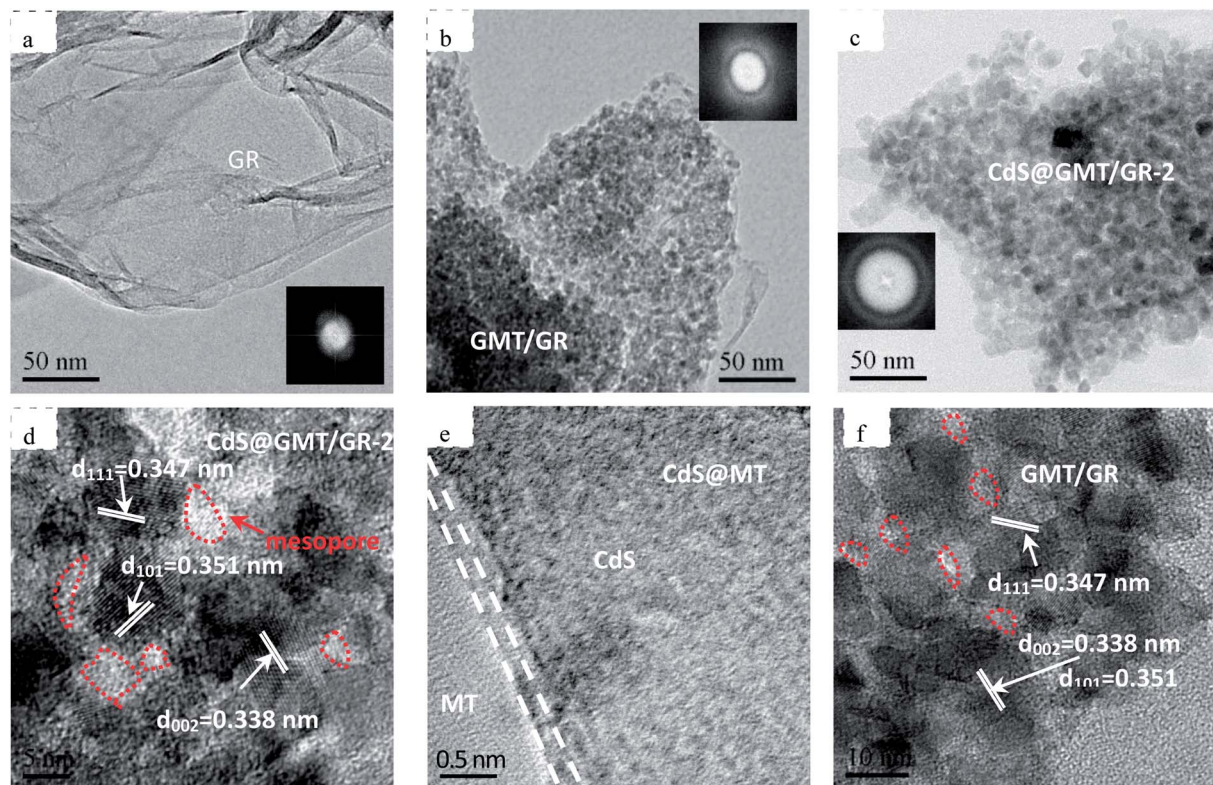


Fig. 4 TEM with SAED of GR (a), GMT/GR (b), and CdS@GMT/GR-2 (c). HRTEM of CdS@GMT/GR-2 (d), CdS@MT (e), and GMT/GR (f).

between pure GR and CdS@GMT/GR-2. The formation of the CdS@GMT nanoparticle phase had a marginal impact on the morphology of GR. Moreover, elemental mapping was carried out to determine the elemental distribution. As shown in the lower panels of Fig. 3, C, Cd, S, Ti, and O were observed to be present in the testing area with even distribution, indicating that CdS and GMT nanoparticles uniformly covered the GR. In addition, it is obvious that the densities of Ti and O were greater than those of Cd and S, in accordance with their contents analyzed by EDS (ESI 3[†]), as listed in Table 1. Fig. 4a shows a low-magnification image of the GR, showing it to be entangled with wrinkles and folds on its surface and edges. The overlapping GR produced crisscross lattice fringes, and the electron diffraction pattern exhibited a dim ball, indicating poor crystalline quality. An image of the GMT/GR also revealed layers of crinkled GR, and the size of the GMT on the GR was estimated to be 17.2 nm (Fig. 4b). CdS@GMT/GR-2 retains the layered structure of GR, and exhibits a smaller size (*ca.* 19 nm) than that of CdS@MT (*ca.* 22 nm) (ESI 4[†]), which is in good agreement with the XRD analyses (see Table 1). The electron diffraction patterns of both GR-supported samples (insets in Fig. 4b and c) reveal that the degree of crystallinity of the GR nanosheets was increased by coating on MT, and CdS@GMT/GR-2 showed bright arrays, confirming the high crystallinity of GR nanosheets. In some regions, distinctive lattice spacings from MT with a periodicity of 0.347 nm (JCPDS: 89-4921) and a fringe separation of 0.351 nm are consistent with the face-centered cubic lattice of CdS (JCPDS: 65-2870). Meanwhile, GR with

dark fringes at 0.338 nm was separated from CdS embedded in the GMT and the latter was anchored on GR nanosheets (Fig. 4d). The direct confinement of QDs in the channels of MT can be attributed to the effects of the IL template and navigation (ESI 1[†]). The pore-embedded structure, being similar to a core-shell structure, showed high light-harvesting efficiency, resulting in a significantly enhanced visible-light photocatalytic performance.^{38,39} Further evidence affirming the attachment of QDs to MT was obtained in the form of a clear interface, as shown in Fig. 4e, wherein the bright and dark fringes originate from MT and CdS, respectively. The change in the contrast of the lattice fringes at the boundary of the two moieties is due to the molecular-level linking between them. To further investigate the effects of CdS QDs on the structure and morphology of GMT/GR, the interaction between the components of GMT/GR is shown in Fig. 4f for comparison. The GMT is heterogeneously flanked by GR nanosheets, and the GR is also grafted onto the MT lattice. There is no obvious difference in the contrast for GMT/GR, which suggests no effect of QDs on the structure of GMT or interaction between MT and GR.

3.2. Optical absorption and chemical properties

As shown in Fig. 5a, the absorption of MT (ESI 5[†]) was red-shifted by 27 nm following the incorporation of GR, indicating the interaction between TiO₂ and GR to form GMT.^{40–42} As compared to the UV-vis spectrum of CdS QDs (ESI 5[†]), CdS/GR shows strong absorbance in the visible-light region due to GR acting as an electron collector/transporter and a light harvester.



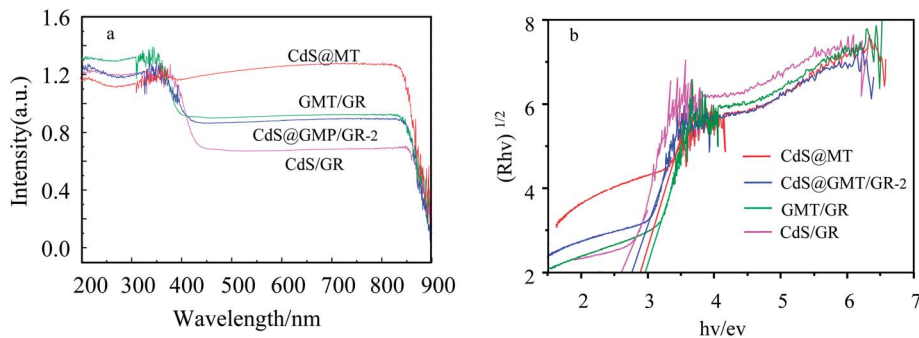


Fig. 5 DRS spectra (a) and corresponding binding energy (b) of the obtained samples.

Table 2 Selected properties of the obtained samples

Samples	BGE ^a (eV)	O 1s content ^b (%)		BET ^c (m ² g ⁻¹)	Average pore size ^d (nm)	Rate constant ^e (× 10 ³ min ⁻¹)
		O-Ti	O-H			
MT	3.17	91	9	136	3.05	—
GMT/GR	2.95	77	23	213	3.48	2.1
CdS@GMT/GR-2	2.67	75	25	179	3.39	12.9
CdS@GMT/GR-1	2.65	69	31	200	3.24	8.3
CdS@GMT/GR-3	2.69	79	21	167	3.42	6.5
CdS@MT	2.90	90	10	127	3.17	4.2

^a BGE represents bandgap energies estimated by UV-vis measurement. ^b Content (%) represents the ratio $A_i/\sum A_i$ (A_i is the area of each peak) measured by XPS. ^c BET surface area calculated from the linear part of the BET plot. ^d BJH desorption average pore diameter. ^e Rate constant calculated from the first-order kinetics.

Owing to the sensitizing effect of CdS on GMT/GR and MT, a new optical absorption edge appears, which is characteristic of QDs. Moreover, when compared with those of CdS@MT, CdS/GR, and GMT/GR, the characteristic absorption edge of CdS@GMT/GR-2 is red-shifted, which can mainly be attributed to synergistic effects among the three different components. The bandgap energies of the samples can be estimated from the plots in Fig. 5b by the Kubelka-Munk function ($F(R_\infty) = (1 - R)^2/(2R)$, where R is the reflectance).³¹ The bandgap energies (E_g) of CdS@MT, CdS/GR, and GMT/GR were determined as ca. 2.90, 2.60, and 2.95 eV, respectively, which suggests a light-responsive transition from the UV to the visible-light region. CdS@GMT/GR showed a narrower bandgap energy in the range 2.65 to 2.69 eV when compared with GMT/GR or CdS@MT (Table 2). In addition, along with the GR content the visible-light absorption intensity of CdS@GMT/GR also increased (ESI 6[†]), although the amount of GR has a marginal impact on the bandgap energy. As shown in Fig. 6, for the MT hybridized with GR, both the PL peaks of GMT/GR clearly decrease; therefore, the recombination of the photogenerated electron-hole pairs is restricted.⁴³ When the MT is embedded with CdS, the PL peaks of CdS@MT also decrease because hole transfer from the valence band (VB) of TiO₂ to that of CdS restricts this recombination.^{32,44} The PL intensity of CdS@GMT/GR-2 is lower than those of CdS@MT and GMP/GR, which can be mainly attributed to the large contact interface between the CdS confined in the channels of GMT anchored on GR. This can

considerably accelerate the migration of charge carriers across the interface between the GMT and CdS QDs. In addition, CdS@GMT/GR-2 shows the lowest PL intensity among these composites because GR has a shielding effect on visible light, while also promoting charge separation (ESI 7[†]).

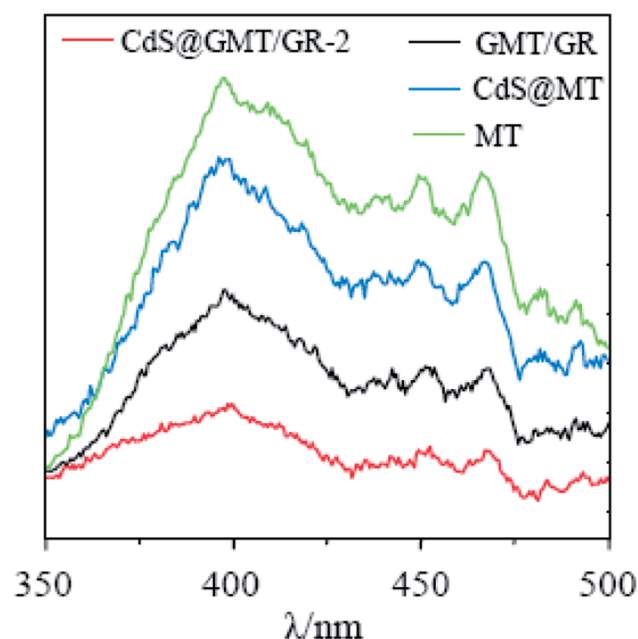


Fig. 6 Photoluminescence (PL) spectra of the obtained samples.



The chemical compositions and valences of the samples were studied by XPS analysis, as shown in Fig. 7. The full spectra of CdS/GR, GMT/GR, and CdS@GMT/GR-2 are shown in Fig. 7a, featuring the expected peaks of C, Ti, S, O, and Cd. Fig. 7b shows the fine spectrum of Ti 2p, in which it can be seen that Ti 2p of MT displays two typical values of TiO₂ at 459.5 and 465.3 eV corresponding to the binding energies of Ti 2p_{3/2} and Ti 2p_{1/2}, respectively.^{33,45,46} In the GR-based sample spectrum, the Ti 2p peaks indicate higher binding energies due to the chemical interaction between MT and GR sheets, leading to facilitated charge separation.^{47,48} These results suggest that GR plays the roles of both hole-transporting and collection layer and the MT-hybridization source.^{49,50} Correspondingly, CdS@MT shows a Ti 2p binding energy without a positive shift (ESI 8†). As shown in Fig. 7c, the obtained O 1s spectrum features two chemical states for oxygen, corresponding to Ti–O bonds ($E_b = 531.4$ eV) and hydroxyl groups ($E_b = 532.5$ eV).⁴⁵ XPS curve-fitting results are also listed in Table 1. The OH content of the GR-based composites was higher than those of MT and CdS@MT, indicating that the water molecule adsorption on the GMT and CdS@GMT surface increased in the presence of the large-surface-area GR. This increase in surface OH content could enhance the photocatalytic activity due to its combining with surface-trapped holes to form ·OH radicals.²¹ The OH

concentration obviously decreased with decreasing GR content due to the decrease in GR-based composite surface area (Table 2). As indicated in Fig. 7d, the XPS spectrum of C 1s in GMT/GR could be deconvoluted into three peaks at 284.5, 286.6, and 288.2 eV, attributable to sp²-bonded carbon (C–C), carbon–oxygen single bonds (C–O), and carbonyl groups (C=O). In contrast, in the XPS spectrum of C 1s in CdS@GMT/GR-2, the peak attributed to sp²-bonded carbon became predominant and those due to C–O and C=O groups considerably decreased, indicating the removal of most oxygen functional groups and the regeneration of the conjugated network by hydrazine reduction. Fig. 7e shows the XPS spectra of Cd 3d_{7/2}, Cd 3d_{5/2}, and Cd 3d_{3/2}. The corresponding binding energies are 401.5, 405.2, and 411.9 eV, indicating that Cd exists in the form of Cd²⁺ in CdS/GR.^{51,52} The spin orbital photoelectron peaks of S²⁻ (2p_{3/2} and 2p_{1/2}), S⁰ (2p_{1/2}), and S⁶⁺ (2p_{1/2}) for CdS/GR in Fig. 7f correspond to binding energies of 161.6 eV (S²⁻), 162.7 eV (S²⁻), 163.6 eV (S⁰), and 164.7 eV (S⁶⁺), respectively. These may be attributed to the photocorrosion processes of CdS in the presence and absence of H₂O and O₂ described by eqn (1) and (2), respectively:^{53,54}

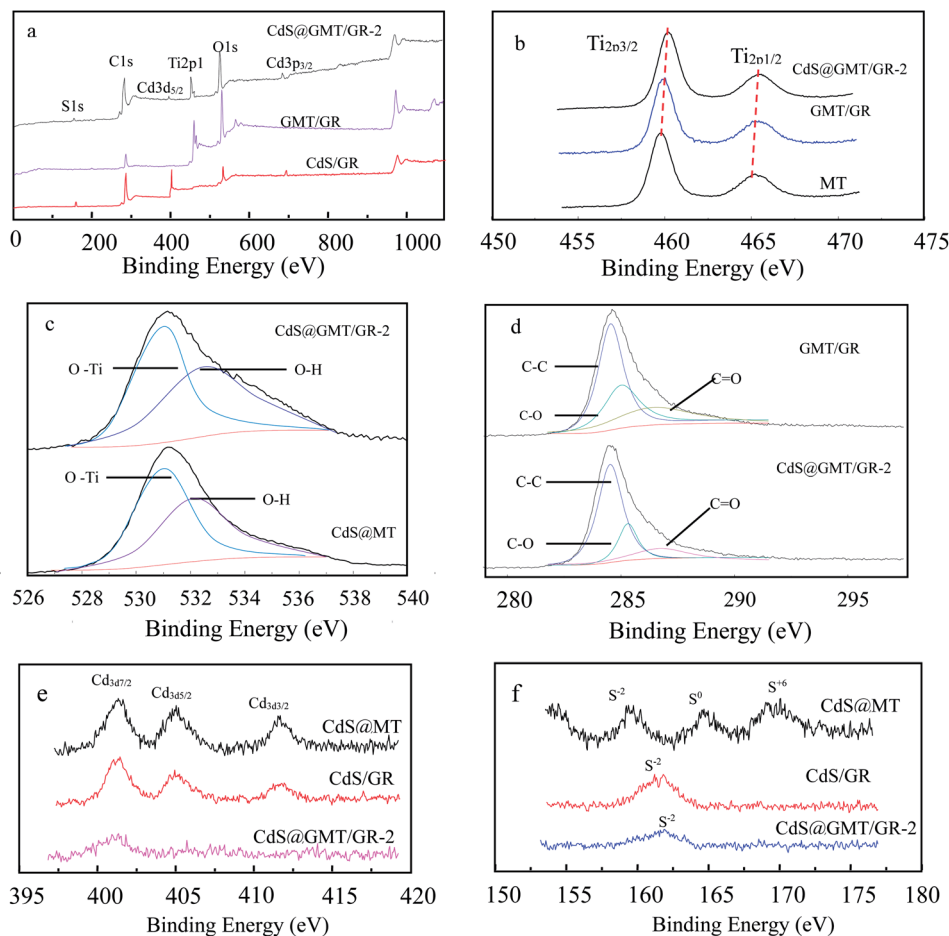
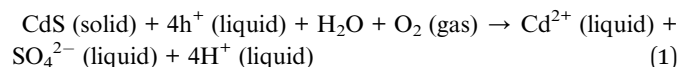


Fig. 7 Surface XPS spectra of the samples (a) and the enlarged spectra of Ti 2p (b), C 1s (c), O 1s (d), Cd 3d (e), and S 2p (f).



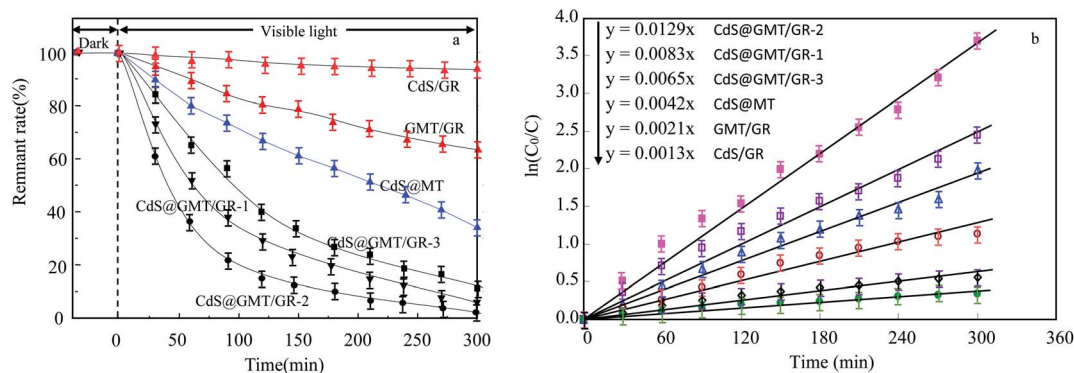
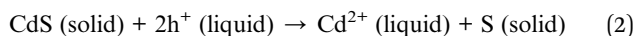


Fig. 8 Photocatalytic activities (a) and linear fittings of the pseudo-first-order kinetics in degrading the GOII solution (b) by different catalysts under visible-light irradiation.



where h^+ represents the photogenerated holes of CdS.

Hence, the photocorrosion resistance of CdS is considerably improved by GR. Interestingly, the GR content also affected the antiphotocorrosion performance of CdS (ESI 9[†]), which is consistent with an earlier report.⁵⁴

3.3. Photocatalytic activity

As shown in Fig. 8a, it was observed that the concentrations of GOII remained unchanged for all the samples without light irradiation. In addition, pure light exhibited a negligible effect on GOII degradation, verifying that the decomposition of GOII was triggered by photocatalytic reactions. CdS@MT and MT/GR showed some activity toward GOII decomposition and CdS@GMT/GR showed high photocatalytic activity, implying that the combination of GR and CdS QDs could synergistically enhance the photocatalytic performance. Moreover, the photocatalytic activity of CdS@GMT/GR was found to be highly dependent on the amount of incorporated GR. With increasing GR content, the photocatalytic activity of CdS@GMT/GR was firstly enhanced and then decreased. CdS@GMT/GR-2 showed the highest degree of GOII decomposition of 94%, which may be

attributed to the shielding effect of GR on the active sites of CdS@GMT and the promotion of charge separation. The photocatalytic reaction rate described by linear fitting of quasi-first-order kinetics is shown in Fig. 8b. Evidently, the catalytic rate constant of CdS@GMT/GR-2 is approximately twice the sum of those of CdS@MT and GMT/GR. This can mainly be attributed to the synergistic effects of several factors, including the composite of a GR-hybridized anatase phase with a mesoporous structure, embedded QDs with good crystallinity, and highly conductive GR with a large surface area. Moreover, CdS@GMT/GR showed better activity retention (Fig. 9a) than CdS@MT after repeated use, because the residual concentration of Cd (Fig. 9b) in the reaction solution using CdS@GMT/GR was about twelve-fold lower than that when using CdS@MT. The photocorrosion of CdS can release Cd^{2+} ions into the solution. Hence, the present results demonstrate enhanced chemical stability and improved photostability of CdS@GMT/GR due to GR serving as a hole-transporting and collection layer, restricting the photocorrosion of CdS.³³ The stable XRD pattern (ESI 10[†]), XPS survey spectrum (ESI 11[†]), and valence state of Cd (ESI 12[†]) also prove the chemical stability of CdS@GMT/GR in the photocatalytic process. To understand the photocatalytic mechanism and to investigate the active species involved in the photocatalytic

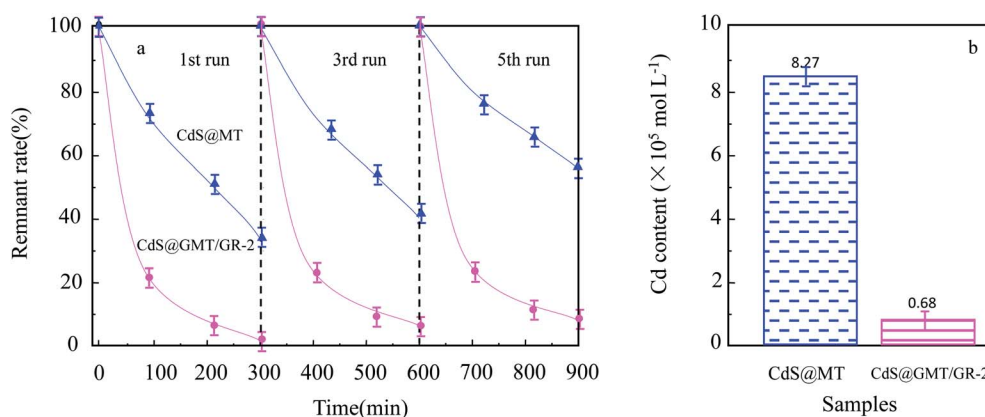


Fig. 9 Photocatalytic performance stability (a) and residue concentration of Cd element in the reaction GOII solution by different catalysts under visible-light irradiation (b).



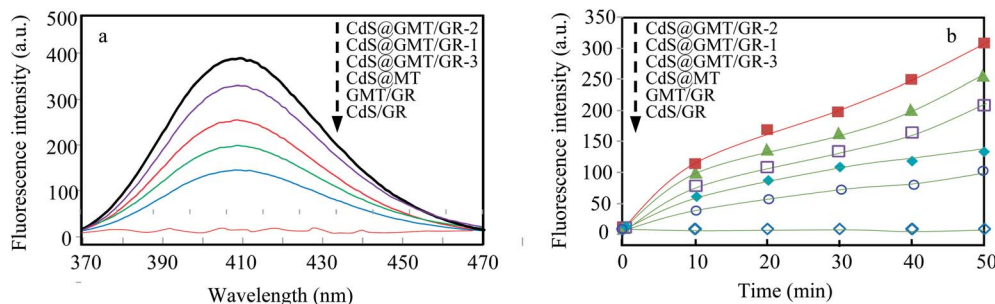
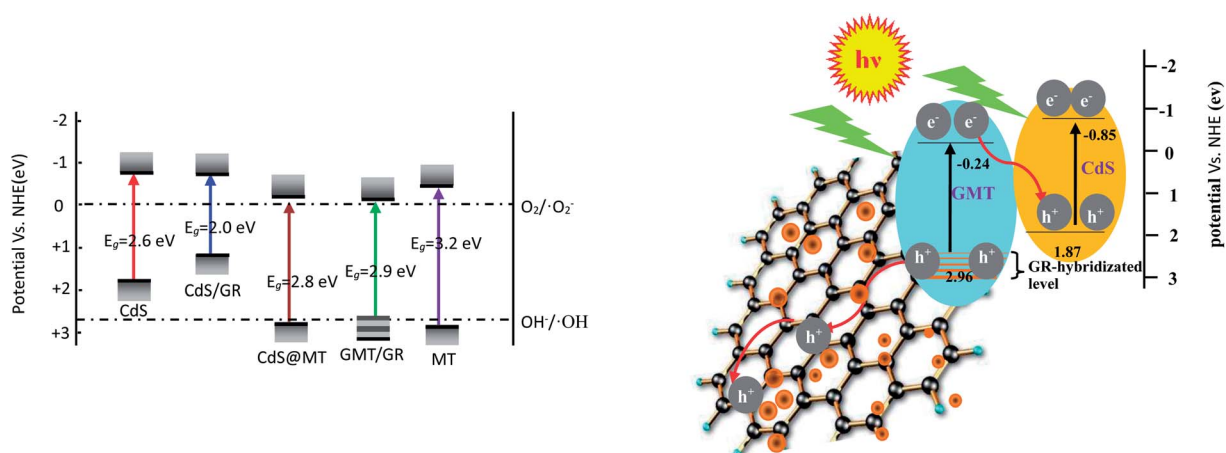


Fig. 10 PL spectral changes for the different samples at 50 min (a) and comparison of PL intensity at 425 nm against irradiation time for the different samples (b).

process, hydroxyl radicals ($\cdot\text{OH}$) were detected on the surface of visible-light-illuminated samples by the PL method using TA as the probe molecule. Fig. 10a shows the PL spectral changes observed in a 5×10^{-4} M basic solution of TA (excitation: 315 nm). No PL signal was observed upon the irradiation of the CdS/GR sample, indicating that no $\cdot\text{OH}$ radicals were produced. In the cases of the CdS@GMT/GR-2, CdS@GMT/GR-1, CdS@GMT/GR-3, CdS@MT, and GMT/GR samples, gradual decreases in PL intensities at about 425 nm were observed after an irradiation time of 50 min. Moreover, except for CdS/GR, the PL signal with all the samples increased with increasing irradiation time (Fig. 10b), reflecting the production of $\cdot\text{OH}$ radicals. The above results can be readily explained by the locations of the VB edges (E_v) of CdS/GR, CdS@MT, GMT/GR, and CdS@GMT/GR, and the normal potential of the $\text{OH}^-/\cdot\text{OH}$ couple (Scheme 2a). Their potentials are 1.87 and 2.7 V *versus* NHE, respectively, which suggests that the photogenerated holes on the surfaces of all the samples can react with $\text{OH}^-/\text{H}_2\text{O}$ to form $\cdot\text{OH}$, except for CdS and its GR composites. On the basis of the above results and discussion, the photocatalytic mechanism of CdS@GMT/GR is proposed and schematically illustrated in Scheme 2b. The conduction band (CB) and VB positions of bulk CdS are *ca.* -0.85 and 1.87 eV, respectively.²⁶ Its VB position is more negative than the standard redox potential of $\text{OH}^-/\cdot\text{OH}$, which

suggests that the holes photoinduced on the VB of QDs are incapable of oxidizing OH^- to $\cdot\text{OH}$ radicals. On the other hand, GR hybridization would introduce an intermediate bandgap state close to anatase VB and extend light absorption into a longer wavelength region.³⁷ Therefore, under visible-light irradiation, electrons can be promoted from the GR-hybridized level to the CB of MT to transfer to the VB of CdS, while the photoinduced holes tend to remain in the VB of GMT. This results in an efficient spatial separation of the photoinduced charge carriers. Electrons stored in the CB of CdS are then trapped by O_2 in air near the surface of CdS to form reactive superoxide radical ions ($\cdot\text{O}_2^-$), while the holes in the VB of GMT transfer to the surface of GR to react with adsorbed water molecules (or surface hydroxyl groups) to form hydroxyl radicals ($\cdot\text{OH}$). Finally, GOII is oxidized by these highly active $\cdot\text{OH}$ radical species to form carbon dioxide and water. The above hydroxyl radical experiments further validate the direct CdS@GMT/GR Z-scheme photocatalytic mechanism. First, if GMT and CdS were to form a conventional heterojunction without the Z-scheme structure, then, under visible-light irradiation, the photoinduced holes would tend to transfer from the VB of GMT to the VB of CdS, while electrons would transfer from the CB of CdS to the CB of GMT. Although this would result in an efficient spatial separation of the photoinduced charge



Scheme 2 Illustrating band positions of different samples together with $\text{OH}^-/\cdot\text{OH}$ and O_2/O_2^- redox potentials (a) and proposed mechanism diagram of the enhancement of the photocatalytic activity of the direct Z-scheme CdS@GMT/GR (b).



carriers, the holes in the VB of CdS could not react with the adsorbed water molecules (or surface hydroxyl groups) to form hydroxyl radicals ($\cdot\text{OH}$) due to the VB position of CdS being higher than the potential of the $\text{H}_2\text{O}/\cdot\text{OH}$ couple. This would result in CdS@GMT/GR having a lower oxidizing ability and photocatalytic activity. Therefore, CdS@GMT/GR forms a direct Z-scheme structure according to the above photocatalytic activity and hydroxyl radical experiments. On the contrary, if CdS@MT were to form a Z-scheme structure, photogenerated electrons could transfer to the surface of CdS from MT only under UV irradiation with $h\nu \geq E_g$ (MT). However, photogenerated holes would be accumulated on the MT and could not transfer to the external support, *i.e.*, GR, to contact OH^- , H_2O , and GOII. Hence, MT would become a recombination center for the photogenerated electrons and holes. Moreover, the accumulated holes would lead to severe photocorrosion of QDs (XPS analysis). Therefore, the photocatalytic mechanism of CdS@MT sample cannot be mainly interpreted in terms of a direct Z-scheme structure, at variance with a previous report.⁵² In addition, an excessively high content of GR would tend to exert a negative influence on the corresponding photocatalytic activity.^{55,56} Once the content of GR exceeds a certain level, the GMT surface will be completely covered by GR. The intensity of light impinging on the surface of the GMT will be decreased due to the shielding effect of the GR, thereby leading to the production of fewer photogenerated electrons and holes in the GMT and poor photocatalytic activity. Hence, CdS@GMT/GR-2 shows high oxidation performance for GOII decomposition under visible light.

To further verify the improvement of charge-transfer efficiency in CdS@GMT/GR, electrochemical impedance spectroscopy (EIS) of the obtained samples was also determined. As shown in Fig. 11, the diameter of the semicircular arc of the GR-based sample was much smaller than those of the MT/GR and CdS@MT samples. CdS@GMT/GR-2 had the smallest radius of the arc in the EIS pattern among all the samples. In general, a smaller semicircle arc in the high-frequency region indicates faster interfacial charge transfer and more effective separation

of the photogenerated electron-hole pairs.^{26,43,57} Therefore, the above photoelectrochemical activity results suggest that a remarkable enhancement of the charge-transport properties of CdS@GMT/GR could contribute toward the suppression of charge recombination and improving the photocatalytic performance.

4. Conclusions

In summary, direct Z-scheme CdS@GMT/GR photocatalysts have been successfully fabricated by a facile hydrothermal method. The as-prepared CdS@GMT/GR photocatalysts exhibit high photocatalytic activity for GOII degradation under visible-light irradiation. This can be mainly attributed to the following facts. GR plays an important role in enhancing the photocatalytic activity, acting as a hole-transporting and collection layer in the Z-scheme system, and remarkably promoting the separation and transfer of the photogenerated electron-hole pairs. Meanwhile, GR hybridized with MT to form GMT through a mixing of energy levels, endowing the charge carriers of this Z-scheme system with strong redox potentials under visible light. Therefore, the GR content affects the photocatalytic activity of CdS@GMT/GR and CdS@GMT/GR-2 exhibited the highest photocatalytic performance. CdS QDs not only serve primarily as photosensitizers in the direct Z-scheme system, but the unique pore-embedded structure of CdS@GMT also effectively enhances the utilization efficiency of the light source. This work may provide a promising strategy for the construction of visible-light-driven Z-scheme photocatalysts for reducing the growing problem of environmental pollution.

Conflicts of interest

There are no conflicts to declare.

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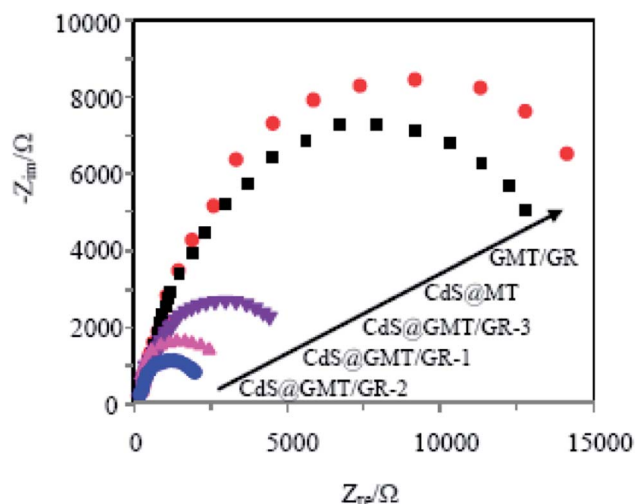


Fig. 11 EIS of the obtained sample electrodes.



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