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Solvent-free hydrogenation of cyclohexene over Rh-TUD-1 at ambient conditions

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Rhodium nanoparticles ($\approx 3\text{--}5$ nm) were incorporated into the 3D mesoporous TUD-1 material by using sol-gel technique. The prepared catalyst shows high activity in the liquid phase conversion of cyclohexene to cyclohexane at room temperature (298 K), 1 atm H_2 pressure, and under solvent-free conditions. Rhodium nanoparticles exhibited high stability, reusability and negligible leaching.

Introduction

The hydrogenation of olefins is considered one of the most important reactions in petroleum as well as chemical industries.^{1,2} The hydrogenation of cyclohexene produces cyclohexane, which undergoes oxidation reaction to produce cyclohexanol and cyclohexanone, the precursors of Nylon 6 and Nylon 6,6.^{3,4} Several homogeneous catalysts have been reported to catalyze the liquid phase hydrogenation of cyclohexene, notably Wilkinson type catalysts.⁵ However, due to the industrial problems associated with the use of homogeneous catalysis such as separation and reutilization, the immobilization of homogeneous catalysts on solid supports was targeted.^{6,7} However, the synthesis of these catalysts is normally performed by multi-step procedures which increases the total costs of the overall reaction.

Noble metals attract a lot of interest in heterogeneous hydrogenation reactions due to their high activity.⁸ The catalytic materials which consist of a proper support incorporating highly-dispersed and structurally-defined metallic nanoparticles seem a smart alternative to bulk metal catalysts.⁹ This is because the high surface area of the catalyst and also its low cost compared to the bulk metal catalysts.¹⁰ Several studies to design heterogeneous catalysts that can activate the cyclohexene hydrogenation have been reported.^{11,12}

Rhodium (Rh) is one of the most interesting metal that is used in hydrogenation reactions of olefins either in homogeneous or heterogeneous catalysts.^{13,14} Few attempts were carried out to hydrogenated cyclohexene to cyclohexane by using supported nanoparticles of rhodium. Crus reported the use of silica supported Rh particle to hydrogenate cyclohexene in a gas phase,¹⁵ Dupont and co-workers reported the use of silica

encapsulate Rh by the help of ionic liquids to hydrogenate cyclohexene at 50 and 70 °C.¹⁶ On the other hand, Ronzón reported the hydrogenation of cyclohexene over Rh/SiO₂ at room temperature, but they used H_2 pressure of 869 KPa.¹⁷

In the current work, a new catalyst is introduced. The catalyst is made of Rh nanoparticles incorporated into the three-dimensional mesoporous silica TUD-1. The synthesis was carried out by one-pot synthesis procedure based on sol-gel technique very similar to the synthesis method which was reported earlier by Hamdy and coworkers.^{18–20} The activity of the prepared catalyst was investigated in the liquid phase, room temperature (298 K) and ambient pressure hydrogenation of cyclohexene. The stability and reusability were also investigated.

Experimental

Nanoparticles of Rh were incorporated into TUD-1 mesoporous silica material by using sol-gel technique. A sample with a Si/Rh ratio of 200 was synthesized by using the molar oxide ratio of $1\text{SiO}_2 : 0.005\text{RhO} : 0.5\text{TEAOH} : 1\text{TEA} : 15\text{H}_2\text{O}$. In a typical procedure, triethanolamine (TEA, 97%, ACROS) was diluted by deionized water was added to rhodium(III) nitrate hydrate ($\text{Rh}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$, Sigma-Aldrich) solution under stirring. The obtained solution was added dropwise to tetraethyl orthosilicate (TEOS, +98%, ACROS) under vigorous stirring. Then, tetraethyl ammonium hydroxide (TEAOH, 35%, Aldrich) was added dropwise and the overall mixture was stirred for at least two hours at room temperature. The resulting homogeneous solution/gel was aged at room temperature for 24 h, and then dried at 98 °C for another 24 h. The obtained brownish solid was grounded and thermally treated in a 50 ml Teflon-lined stainless-steel autoclave at 178 °C under autogenous pressure for 4 h. The obtained solid was grounded again, and then calcined in static air at 600 °C for 10 h by applying a heating ramp rate of 1 degree per min. The obtained material was coded as A-Rh-0.5. Finally, A-Rh-0.5 sample was hydrogenated in

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a tube furnace at 300 °C for 3 h to ensure the formation of Rh⁰ in the final product which coded as Rh-0.5.

XRD was performed by using Shimadzu 6000 DX instrument diffractometer equipped with a graphite monochromator using CuK_α radiation ($\lambda = 0.1541$ nm). FTIR spectra were collected by using Agilent Cary 630 FTIR Spectrometer. Raman spectrum was collected at ambient conditions by using a Renishaw Raman Imaging Microscope, system 2000. The green ($\lambda = 514$ nm) polarized radiation of an argon ion laser beam of 20 mW was used for excitation. Nitrogen adsorption/desorption isotherms were recorded on a Quanta Chrome NOVA 2000e instrument. The pore size distribution was calculated from the adsorption branch using the Barrett-Joyner-Halenda (BJH) model. The BET method was used to calculate the surface area (S_{BET}) of the sample. ICP technique was used to quantify the amount of Rh present in the prepared material by using Thermo scientific, ICAP 7000 series, part no: 1340910, Qtegra Soft wear, Germany. Scanning electron microscopy (SEM) Jeol Model 6360 LVSEM, USA, was used to investigate the morphological structure of the prepared sample. Transmission Electron Microscopy (TEM) was carried out on a Philips CM30UT electron microscope with a field emission gun as the source of electrons, operated at 300 kV.

The solvent-free hydrogenation of cyclohexene was performed in a stirred batch Parr reactor with 300 ml capacity. Typically, 0.25 g of the carefully grinded catalyst was placed inside the reactor with 25 ml of cyclohexene ($\geq 99\%$, TEDIA, USA). The autoclave was closed and degassed twice with nitrogen and pressurized to the necessary pressure with hydrogen gas. Then, the reaction was vigorously stirred (up to 1000 rpm) to minimize the external mass transfer limitations. After the reaction time was elapsed, the reactor was flushed with nitrogen and then a liquid sample was withdrawn and filtrated. The concentrations of the product and the reactant were analyzed by using a SHIMADZU GC-17 instrument equipped with RTX-5 capillary column (30 m \times 0.25 mm \times 0.25 μm) and a flame ionization detector (FID).

The conversion of cyclohexene (CXE), was calculated according to the following equations:

$$\text{Conversion} = \frac{[\text{CXE}]_0 - [\text{CXE}]_t}{[\text{CXE}]_0} \times 100$$

where $[\text{CXE}]_0$ is the initial concentration of cyclohexene, $[\text{CXE}]_t$ is the concentration of cyclohexene at time (t).

On the other hand, the turnover frequency (TOF) was calculated according to the equation

$$\text{TOF} = \frac{[\text{CXE}]_0 - [\text{CXE}]_t}{[\text{Rh}] \times \text{time}(s)}$$

where $[\text{Rh}]$ represents the total number of moles of Rh active phase and the reaction time was calculated in seconds.

Results

ICP elemental analysis showed that the loading% of rhodium is 0.493% while the added amount during the synthesis was equivalent to 0.5%. This clearly shows that all the added

rhodium amount was found in the final product. Moreover, it also shows the high predictability of the synthesis method.

XRD patterns of Rh-TUD-1 before hydrogenation (A-Rh-0.5) and after hydrogenation (Rh-0.5) are compared with that of neat TUD-1 silica material in Fig. 1A. All the patterns show a broad band around 23–25° 2 θ which can be characteristic to amorphous phase of silica.²¹ In the A-Rh-0.5 sample, peaks at 24.4°, 32.9°, 34.98°, 44.0°, 48.77°, 53.3°, 57.17°, 61.57°, and 67.54° 2 θ . These peaks can be attributed to the formation of β -Rh₂O₃ phase (JCPDS 76-0148).²² After reduction, Rh-0.5 pattern exhibited only peaks at 41.16°, 47.86°, and 70.17° 2 θ . These peaks are corresponding to the crystal planes of (111), (220), and (220), respectively, which indicating the presence of metallic rhodium (Rh⁰) particles (JCPDS 05-0685).

Fig. 1B showed the FT-IR spectra of neat TUD-1 compared with the as-synthesised Rh-TUD-1 and the final solid product Rh-TUD-1. The bands at 1080 cm⁻¹ and 850 cm⁻¹ are corresponding to the characteristic anti-symmetric vibrations of non-bridging oxygen atoms (Si-O⁻) of Si-O-H bonds and symmetric stretching vibrations (Si-O-Si)_{sym} of tetrahedral SiO₄ structure units. Meanwhile, the band at around 460 cm⁻¹ corresponds to the characteristic tetrahedral bending of Si-O bonds. These bands were also observed in A-Rh-0.5 and Rh-0.5 samples in addition to a shoulder at 960 cm⁻¹. This shoulder can be assigned to the vibration of defected Si-O as a result of Rh presence.^{23,24} The intensity of the shoulder is getting intense in Rh-0.5 due to increase the interaction strength between the Rh nanoparticles and the silica support.

To confirm the transformation of Rh₂O₃ in the as-synthesised sample to metallic Rh in the hydrogenated sample, Raman analysis was performed. Raman spectrum of A-Rh-0.5 is compared with that of Rh-0.5 in Fig. 1C. The spectrum of A-Rh-0.5 is dominated by three bands at around 265, 411, 549 cm⁻¹ which is characteristic to the presence of hexagonal β -Rh₂O₃.²⁵ These peaks were totally disappeared after hydrogenation step as an indication for the formation of metallic Rh.

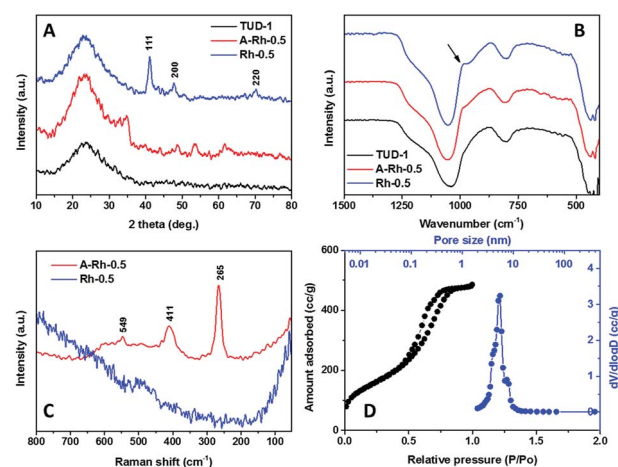


Fig. 1 (A) XRD patterns of the neat TUD-1, as synthesized Rh-TUD-1 (A-Rh-0.5) and Rh-0.5 sample. (B) The corresponding FTIR spectra. (C) Raman spectra of A-Rh-0.5 compared to Rh-0.5 sample. (D) The N₂ sorption isotherms of Rh-0.5 sample together with the pore size distribution.



The N_2 adsorption–desorption isotherm of Rh-0.5 sample is presented in Fig. 1D together with the corresponding pore size distribution. The N_2 sorption isotherms was found to be type IV and exhibited H1 type of hysteresis loop which is characteristic to the mesoporous materials. Moreover, BET surface area was found to be $721 \text{ m}^2 \text{ g}^{-1}$ with an average pore size of 4.8 nm.

In order to investigate the morphological structure of Rh-0.5, SEM analysis was carried out. The SEM micrograph (Fig. 2A) shows the irregular shape of TUD-1 matrix, no other bulky phase was observed. Moreover, EDX analysis exhibited only the peaks of Si, O and Rh, which confirms the formation of the catalysts with high purity. The formation of the Rh nanoparticles was established by HR-TEM analysis (Fig. 2B). The micrograph of Rh-0.5 sample shows the highly dispersed nanoparticles of Rh incorporated in the mesoporous matrix. Moreover, the average nanoparticle size was measured to be between 3 and 5 nm.

The catalytic activity of Rh-TUD-1 was investigated in the solvent-free liquid phase hydrogenation of cyclohexene at ambient conditions. Fig. 3A shows the conversion of cyclohexene and the TOF over Rh-TUD-1 catalyst compared with blank reaction and neat TUD-1. Without catalyst, no conversion at all was obtained after 1 h's reaction, same result was obtained with neat TUD-1. However, when Rh-TUD-1 was used to catalyze

the reaction, 24.8% conversion was obtained after 1 h's reaction with a total TOF of 1.4 S^{-1} . Increasing the reaction time led to increase the cyclohexene conversion% almost linearly; *i.e.* zero order reaction. After 4 h' reaction, cyclohexene was totally converted to cyclohexane (Fig. 3B), with a constant TOF. Increase the H_2 pressure in the autoclave prior the reaction to 3 atm showed positive effect on the conversion, where 60.5% of cyclohexene was converted to cyclohexane with a TOF of 3.4 S^{-1} and to 100% conversion with 5.6 S^{-1} TOF at H_2 pressure of 5 atm (Fig. 3C). The recycling of Rh-TUD-1 was investigated through performing four consecutive reaction without any treatment for the catalyst. The obtained results are plotted in Fig. 3D. The results show minor loss in activity from reaction to another with a constant TOF. Similar behavior was also reported by Leshkov and coworkers about the stability of Rh incorporated hydroxyapatite.²⁶

To confirm the stability of the Rh-TUD-1 during the hydrogenation of cyclohexene, the used Rh-TUD-1 sample (after 4 consecutive reactions) was characterized by using XRD, SEM and Rh mapping. Fig. 4 shows the obtained characterization data. XRD pattern of the used Rh-TUD-1 sample is almost identical to that of the fresh sample. The intensity and location of the peaks did not alter, and more importantly, no new crystalline phase(s) could be observed as an indication of the high stability of Rh-TUD-1 sample. In the background of the left panel, the SEM micrograph of the used Rh-TUD-1 sample showed the surface of the sample in the normal irregular particle shape characteristic of the TUD-1 samples. Left panel of Fig. 4, shows the mapping of Rh nanoparticles over the surface of used sample. No agglomeration was observed and the Rh nanoparticles still well dispersed and nicely spread out through the sample's surface, which is in agreement with XRD result.

A comparison between the TOF of Rh-TUD-1 and other Rh containing different supports in the cyclohexene hydrogenation is listed in Table 1. The comparison clearly shows that Rh-TUD-1 possess high activity under the ambient operating conditions and under liquid phase condition. This can be explained by using the TUD-1 with its open three-dimensional framework which gives high accessibility of the substrate/product to/from the Rh active sites. One more advantage of using Rh-TUD-1 can be related to the ease synthesis procedure, *i.e.* one-step sol-gel technique, more importantly, surface modification and/or ionic liquids were not needed.

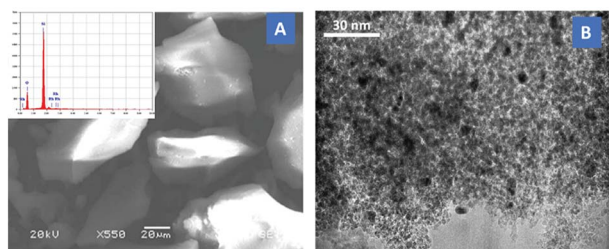


Fig. 2 (A) SEM micrograph of Rh-0.5 sample, inset: the EDX analysis of the sample. (B) HR-TEM micrograph of the Rh-0.5 sample.

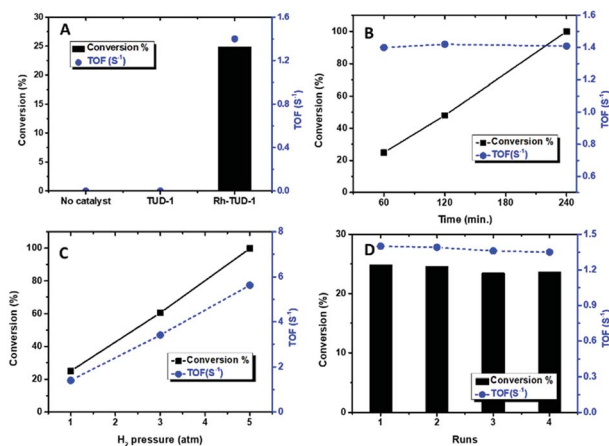


Fig. 3 The conversion% of cyclohexene and the corresponding TOF over Rh-TUD-1 under ambient temperature and H_2 pressure: (A) compared with the blank reaction, (B) at different time up to 4 hours, (C) at different H_2 pressure, and (D) in 4 consecutive reactions without treatment.

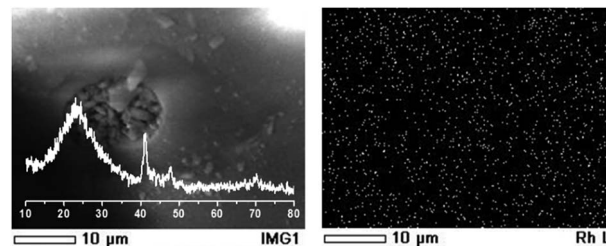


Fig. 4 Left panel: XRD of the used Rh-TUD-1 sample, in the background, a SEM micrograph of the same sample. Right panel: Rh nanoparticles mapping of the used Rh-TUD-1 sample.



Table 1 Comparison between the activity (represented in TOF) of Rh-TUD-1 and the other reported catalysts

Catalyst	TOF (s ⁻¹)	Conditions	Ref.
Rh-TUD-1	1.4	Room T, 1 atm, liquid phase reaction, solvent free	^a
Rh-TUD-1	3.4	Room T, 3 atm, liquid phase reaction, solvent free	^a
Rh/SiO ₂	3.4	Microwave heating, gas phase	27
Rh/SiO ₂	1.7	Conventional heating, gas phase	27
Rh/SiO ₂	1.7	Room T, 14 KPa H ₂ , gas phase	28
Rh/SiO ₂ /HF	0.4	70 °C, 4 atm H ₂ , liquid phase	16
Rh/SiO ₂ /NH ₄ OH	0.13	70 °C, 4 atm H ₂ , liquid phase	16
Rh/hydroxyapatite	0.76	Room T, 3 bar H ₂ , gas phase	26

^a Current study.

Conclusions

To conclude, Rh nanoparticles were successfully incorporated into the framework of mesoporous TUD-1 siliceous material. Rh nanoparticles exhibited high activity in the liquid phase, solvent-free hydrogenation of cyclohexene at ambient conditions.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- G. C. Bond and P. B. Wells, *Adv. Catal.*, 1964, **15**, 91–226.
- J. N. Wilson, J. W. Otvos, D. P. Stevenson and C. D. Wagner, *Ind. Eng. Chem.*, 1953, **45**, 1480–1487.
- B. Hereijgers and B. Weckhuysen, *J. Catal.*, 2010, **270**, 16–25.
- I. Berezin, E. Denisov and N. Emanuel, *The Oxidation of Cyclohexane*, Moscow University Press, Pergamon Press, Oxford, English edition, 1966.
- J. A. Osborn, F. H. Jardine, J. F. Young and G. Wilkinson, *J. Chem. Soc. A*, 1966, 1711–1732.
- E. A. Cagnola, M. E. Quiroga, D. A. Liprandi and P. C. L'Argentièrre, *Appl. Catal., A*, 2004, **274**, 205–212.
- M. Pérez-Cadenas, L. J. Lemus-Yegres, M. C. Román-Martínez and C. Salinas-Martínez de Lecea, *Appl. Catal., A*, 2011, **402**, 132–138.
- R. L. Augustine, F. Yaghmaie and J. F. Van Peppen, *J. Org. Chem.*, 1984, **49**, 1865–1870.
- Nanoparticles and catalysis*, ed. D. Astruc, John Wiley & Sons, 2008.
- M. S. Hamdy, *Microporous Mesoporous Mater.*, 2016, **220**, 81–87.
- L. J. Lemus-Yegres, M. Pérez-Cadenas, M. C. Román-Martínez and C. Salinas-Martínez De Lecea, *Microporous Mesoporous Mater.*, 2011, **139**, 164–172.
- S. García, R. M. Anderson, H. Celio, N. Dahal, A. Dolocan, J. Zhou and S. M. Humphrey, *Chem. Commun.*, 2013, **49**, 4241–4243.
- H. Hirai, Y. Nakao, N. Toshima and K. Adachi, *Chem. Lett.*, 1976, **5**, 905–910.
- R. R. Schrock and J. A. Osborn, *J. Am. Chem. Soc.*, 1976, **98**(8), 2134–2143.
- G. M. Da Cruz, G. Bugli and G. Djega-Mariadassou, *Appl. Catal.*, 1989, **46**, 131–144.
- M. A. Gelesky, S. S. X. Chiaro, F. A. Pavan, J. H. Z. dos Santos and J. Dupont, *Dalton Trans.*, 2007, 5549–5553.
- G. Del Angel and E. Ronzón, *J. Mol. Catal. A: Chem.*, 1999, **148**, 105–115.
- M. S. Hamdy and G. Mul, *Appl. Catal., B*, 2015, **174**, 413–420.
- M. S. Hamdy, M. A. Eissa and S. M. A. S. Keshk, *Green Chem.*, 2017, **19**, 5144–5151.
- M. Benaissa, A. M. Alhanash, M. Eissa and M. S. Hamdy, *Catal. Commun.*, 2017, **101**, 62–65.
- F. Adam, T. S. Chew and J. J. Andas, *J. Sol-Gel Sci. Technol.*, 2011, **59**, 580.
- J. C. Park, *Bull. Korean Chem. Soc.*, 2008, **29**, 1787–1793.
- E. M. Rivera-Muñoz and R. Huirache-Acuña, *Int. J. Mol. Sci.*, 2010, **11**, 3069–3086.
- H. Sun, Q. Tang, Y. Du, X. Liu, Y. Chen and Y. Yang, *J. Colloid Interface Sci.*, 2009, **333**, 317–323.
- W. H. Weber, R. J. Baird and G. W. Graham, *J. Raman Spectrosc.*, 1998, **19**, 239–244.
- M. Zahmakıran, Y. Román-Leshkov and Y. Zhang, *Langmuir*, 2012, **28**, 60–64.
- N. Dahal, S. García, J. Zhou and S. M. Humphrey, *ACS Nano*, 2012, **6**, 9433–9446.
- G. M. Da Cruz, *Appl. Catal.*, 1989, **46**, 131–144.

