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A convenient separation method for di(2-ethylhexyl)phthalate by novel superparamagnetic molecularly imprinted polymers

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Through a surface molecular imprinting technique and coating superparamagnetic Fe_3O_4 nanoparticles with molecularly imprinted polymers (MIPs), a novel magnetic molecularly imprinted polymer (MMIP) was successfully fabricated for the convenient separation of di(2-ethylhexyl)phthalate (DEHP) with methacrylic acid as functional monomer, ethylene glycol dimethacrylate as cross linker and 2,2-azobis(isobutyronitrile) as initiator. A magnetic non-molecularly imprinted polymer (MNIP) was also prepared for comparison purposes. The morphology structure and the magnetic properties of MMIP were characterized by transmission electron microscopy (TEM), Fourier transform infrared spectrometry (FT-IR), X-ray diffraction, vibrating sample magnetometry (VSM) and thermo gravimetric analysis (TGA). The adsorption properties of MMIP and MNIP were investigated by static and dynamic adsorption experiments. The results show that the diameter of the synthesized MMIP microspheres is about 300–500 nm with good dispersibility in solvent. The prepared MMIP shows superparamagnetic properties with the maximum saturation magnetic intensity of 43.97 emu g^{-1} , and it can be conveniently separated using an external magnetic field. Compared with MNIP, MMIP has a higher adsorption capacity and better adsorption selectivity for DEHP, and the imprinting factor reaches 3.012. The regeneration adsorption experiment illuminates that the novel MMIP can be reused with good separation efficiency.

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1 Introduction

Phthalates are a group of organic esters known as estrogens or endocrine disruptors, they are usually used as plasticizers to improve the flexibility and toughness of polyvinyl chloride.^{1,2} As a common phthalate, di(2-ethylhexyl)phthalate (DEHP) is widely used in many consumer products, such as building materials, medical devices, household products, clothes and cleaning materials.³ However, DEHP can easily leach from PVC and other products into the environment, and then transfer into water, air and food. When DEHP reaches a certain concentration, it will harm the health of human organisms, and cause wide adverse effects on the liver, reproductive tract, kidneys, lungs, fetus, and heart.^{4,5} Therefore, the development of new methods for determining and separating DEHP in diverse environmental matrices is highly significant.

Through molecular imprinting technique (MIT), molecularly imprinted polymers (MIPs) can be prepared by copolymerization of functional monomers and templates in the presence of cross-linkers and initiators, and the special recognition ability for a particular substance can be obtained.⁶ The three-dimensional cavities complementary in shape, size and chemical functionality to those of templates are formed on the

surface of prepared MIP, thereby can rebind the templates. MIPs have been widely used in solid phase extraction,^{7,8} chemical sensor,^{9,10} enzyme simulation and catalysis.^{11,12} MIPs are usually prepared by bulk polymerization with complicated process and deep recognition sites, which results in low binding affinity. In the process of preparation and application of MIPs, filtration or centrifugation is often needed, which is time-consuming and difficult to be used widely in practical production. To solve the problems, novel magnetic surface imprinted polymers were reported recently, in which the binding sites were established on the surface of MIPs to facilitate the process of mass transfer and realize controllable movement by the external magnetic field.^{13–15} The magnetic molecular imprinting material can be used as an ideal and convenient separation material. However, there are few literatures of MMIPs for the separation of DEHP.

To realize convenient separation of DEHP with high efficiency, through surface molecular imprinting technique, a novel magnetic molecularly imprinted polymer (MMIP) was successfully synthesized by coating superparamagnetic Fe_3O_4 nanoparticles with MIPs and using methacrylic acid as functional monomer, ethylene glycol dimethacrylate as cross linker and 2,2-azobis(isobutyronitrile) as initiator. The morphology structure and the magnetic property of MMIP were analyzed by TEM, FT-IR, XRD, VSM and TGA, and the adsorption properties were also investigated.

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2 Materials and methods

2.1 Chemicals

DEHP, diallyl phthalate (DAP), dibutyl phthalate (DBP), tetraethoxysilane (TEOS), ferric chloride hexahydrate ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$), ethylene glycol (EG), polyethylene glycol 2000 (PEG2000), sodiumacetate (NaAc), ammonia solution (28 wt%), γ -(methacryloxypropyl) trimethoxy silane (MPS), ethylene glycol dimethylacrylamide (EGDMA), 2,2-azobis(isobutyronitrile) (AIBN), methacrylic acid (MAA), acetonitrile, ethanol, methanol, glacial acetic acid and toluene were all analytical reagent grade and purchased from Sinopharm Chemical Reagent Co., Ltd. (Shanghai, China).

2.2 Preparation of MMIP

First, Fe_3O_4 nanoparticles were synthesized by hydrothermal method. $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ (2.5 g) was dissolved in EG (60 mL) to form a clear solution by stirring. PEG2000 (2 g) and anhydrous NaAc (7.2 g) were added, and the mixture was stirred vigorously for an hour at 50 °C. The mixture was transferred into a Teflon lined stainless-steel autoclave with 100 mL capacity, and the autoclave was heated to 200 °C and maintained for 8 hours. After that, it was cooled to room temperature naturally. The obtained Fe_3O_4 nanoclusters were washed with deionized water and ethanol for three times and separated from the supernatant by external magnetic field. Finally, the Fe_3O_4 nanoparticles were dried in vacuum oven at 50 °C for 6 h.

Then, the $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres were prepared by Stöber method. 0.2 g Fe_3O_4 nanoparticles were dispersed by ultrasonic in 200 mL mixed solvent of ethanol and water ($v_{\text{ethanol}} : v_{\text{water}} = 4 : 1$), and 3 mL $\text{NH}_3 \cdot \text{H}_2\text{O}$ (28 wt%) was rapidly added into the solution. After well-dispersed, 2 mL TEOS solvent was added drop by drop, and the reaction was proceeded at room temperature under mechanical stirring for 6 hours. The product was washed with ethanol and distilled water for three times and then dried in vacuum even at 50 °C for 6 h to obtain $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres.

The surface of $\text{SiO}_2\text{-Fe}_3\text{O}_4$ particles was further modified with reactive C=C bonds by reacting with γ -(methacryloxypropyl) trimethoxy silane. The obtained $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres (0.2 g) were dispersed homogeneously in toluene (80 mL) under ultrasonic condition. 6 mL MPS solvent was then added and reacted at 65 °C with mechanical stirring for 24 hours under nitrogen protection. The obtained MPS- $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres were separated with the help of external magnetic field and washed repeatedly with ethanol and water. The MPS- $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres were also dried in vacuum even at 50 °C for 6 h.

Finally, 0.25 mmol DEHP, 1 mmol MAA, 945 μL EGDMA, 40 mg AIBN and 0.1 g MPS- $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres were dissolved in 80 mL acetonitrile with ultrasonic stirring for 15 min. The mixture was polymerized at 65 °C with mechanical stirring for 24 hours under the protection of nitrogen atmosphere. The product was washed with ethanol and distilled water for several times and collected by magnetic separation. The particles were eluted by the solution of methanol and acetic

acid (9 : 1, v : v) to remove the template molecules and dried in vacuum even at 50 °C for 12 h to obtain MMIP. To compare the separation performance, the magnetic non-molecularly imprinted polymer (MNIP) was also prepared in an identical manner but in the absence of template molecule.

2.3 Characterization

A JEM-2100 (HR) microscope was used to obtain transmission electron microscope (TEM) images of Fe_3O_4 nanoparticles, $\text{SiO}_2\text{-Fe}_3\text{O}_4$ microspheres and MMIP microspheres. For TEM analyses, samples were prepared by placing one or two drops of nanoparticle solution onto the carbon-coated copper grid and drying it in air at room temperature.

Fourier transform infrared (FTIR) spectra were recorded using the FTIR spectrophotometer (Nicolet MX-1E, Thermo Fisher Scientific), and the scanning range is from 4000 cm^{-1} to 400 cm^{-1} .

The identification of crystalline phase of the samples was performed by X-ray powder diffractometer with Cu KR radiation. The 2θ angles were from 10° to 80° at a rate $10^\circ \text{ min}^{-1}$.

Vibrating sample magnetometer (VSM) (Squid-VSM, USA) was used to test the magnetic properties of samples within 1 T at room temperature.

Ultraviolet visible spectrophotometer (UV-2800, SHIMADZU, Japan) was used to determine the concentration of DEHP in solution at room temperature.

The content and thermal stability of MPS- $\text{SiO}_2\text{-Fe}_3\text{O}_4$ and MMIP were determined through thermo gravimetric analysis (TGA, STA449C, Germany) in the range from room temperature to 800 °C with a heating rate of $10^\circ \text{ C min}^{-1}$ under nitrogen atmosphere.

2.4 Static adsorption experiments of MMIP and MNIP

20 mg MMIP (or MNIP) was placed in the centrifuge tubes containing 2 mL ethanol solutions of DEHP with the concentrations from 20 mg L^{-1} to 200 mg L^{-1} , and the mixture was shaken at room temperature for 12 hours. After the adsorption reached equilibrium, the adsorbent was separated from the solution by an external magnetic field. The concentration of DEHP in the supernatant solution was measured by UV spectrophotometer, and the detection wavelength of 225 nm was selected. The adsorption capacities (Q_e) of MMIP (or MNIP) for DEHP were calculated according to the following equation:

$$Q_e = \frac{(C_0 - C_e)}{m} V \quad (1)$$

where C_0 (mg L^{-1}) is the initial concentration of DEHP solution, and C_e is the equilibrium concentration in supernatant solution (mg L^{-1}). V (mL) is the volume of the DEHP ethanol solution, and m (mg) is the mass of dry MMIP (or MNIP) that added to the solution.

2.5 Dynamic adsorption experiment

20 mg MMIP (or MNIP) was dispersed in 2 mL of 100 mg L^{-1} DEHP solution to investigate the kinetic adsorption properties. The mixture was shaken at room temperature for 10 min, 20 min, 30 min, 40 min, 50 min, 1 h, 2 h and 3 h, respectively.



The microspheres were then separated by a magnet to determine the residual concentration of DEHP using UV spectrophotometer at 225 nm. The adsorption capacity Q was calculated according to formula (1).

2.6 Selective adsorption experiment

20 mg of dried MMIP and MNIP were placed in mixture solution of DEHP, DAP and DBP with initial individual concentrations of 100 mg L⁻¹, and the mixture was respectively shaken at room temperature for 12 h. The concentrations of DEHP, DAP and DBP in the supernatant solution was also measured by UV spectrophotometer.

3 Results and discussion

3.1 Preparation of MMIP

The synthesis process of MMIP was shown in Fig. 1. Fe₃O₄ magnetic nanoparticles were first synthesized through hydrothermal method by partial reduction of FeCl₃ with EG as solvent, PEG as a surfactant against particle agglomeration and NaAc as electrostatic stabilization.¹⁶ A layer of SiO₂ was subsequently

coated on the surface of Fe₃O₄ particles to form Fe₃O₄/SiO₂ core-shell particles by TEOS according to the Stöber method.¹⁷⁻¹⁹ The coating reaction is indispensable, because the silica shell is non-toxic and can provide good biocompatibility and hydrophilicity for Fe₃O₄ nanoparticles. The SiO₂ shell can also effectively prevent the chemical degradation and aggregation of the Fe₃O₄ nanoparticles in a harsh liquid environment, and the silanol groups on the surface of silica can be further modified.²⁰⁻²² Then, the silane coupling agent of MPS was used to modify the surface of silica to obtain reactive double bonds, which could interact with functional monomers and guide the formation of molecularly imprinted polymers on the surface of the magnetic matrix. Finally, the molecularly imprinted shell was coated on the surface of the magnetic substrate to obtain magnetic molecularly imprinted polymers. The polymer not only has specific recognition sites for template molecules, but also can be used for convenient separation by external magnetic field.

3.2 TEM analysis

The morphology structures of Fe₃O₄, SiO₂-Fe₃O₄ and MMIP were observed by TEM (Fig. 2). Fig. 2(a) reveals that the

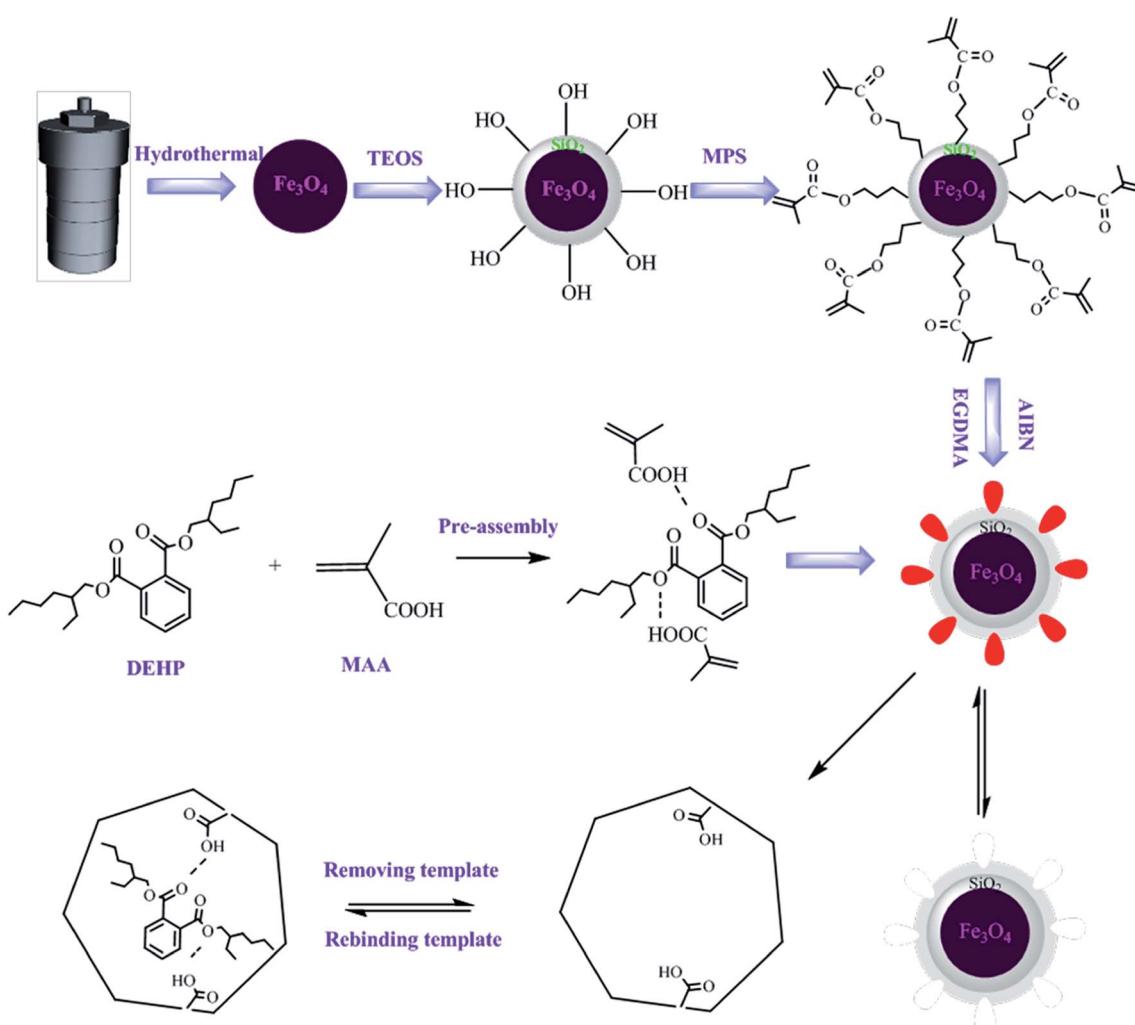


Fig. 1 Schematic diagram of MMIP preparation.

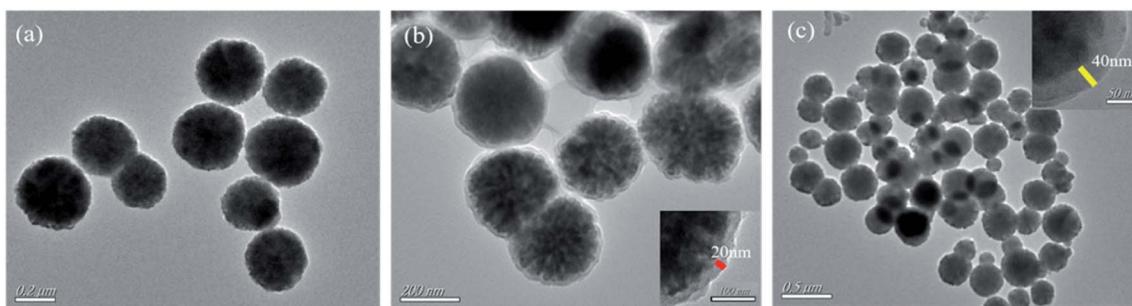


Fig. 2 TEM images of Fe₃O₄ (a), SiO₂-Fe₃O₄ (b) and MMIP (c).

synthesized Fe₃O₄ nanoparticles have regular spherical shape with good dispersibility, and the diameter is in the range of 300–500 nm. As shown in Fig. 2(b), SiO₂ shell is coated on the surface of Fe₃O₄ and displays obvious core–shell structure, the thickness of the silica layer is about 20 nm. Because the density of the silica shell is similar to that of molecularly imprinted shell, the molecularly imprinted shell cannot be distinguished clearly in Fig. 2(c). However, compared with Fig. 2(b), the thickness of the coating layer is further increased, which suggests the formation of imprinting shell.

3.3 Infrared analysis

Fig. 3 shows the infrared characterization of Fe₃O₄, SiO₂-Fe₃O₄, MPS-SiO₂-Fe₃O₄ and MMIP. The peak at 587 cm⁻¹ in curve (a) is the typical vibration of the Fe–O, which also exists in curve (b), (c) and (d). The peaks at 1630 cm⁻¹ in curve (a), (b), (c) and (d) are attributed to the bending vibration of Fe₃O₄ hydration water. The strong peaks at 1090 cm⁻¹, 798 cm⁻¹ of curve (b) are attributed to the anti-stretching vibration of Si–O–Si and the symmetric stretching vibration of Si–O, respectively, indicating the formation of silica layer. The peaks at 2920 cm⁻¹ and 2850 cm⁻¹ in curve (c) correspond to the vibration of C–H, and the peak at 1640 cm⁻¹ in curve (c) is the stretching vibration of C=C, which are attributed to the silica shell modified by MPS.

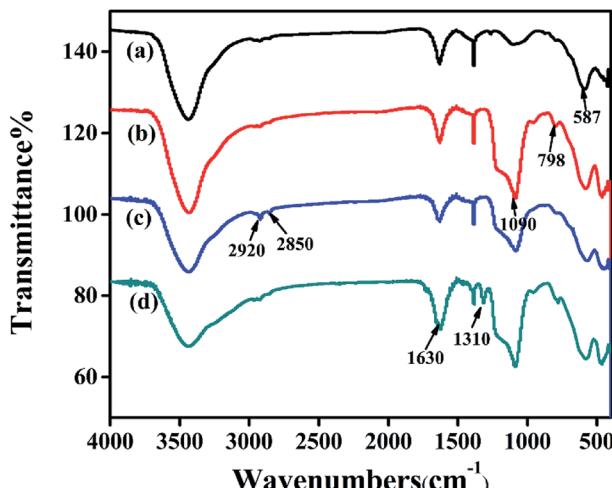


Fig. 3 FTIR spectra of Fe₃O₄ (a), SiO₂-Fe₃O₄ (b), MPS-SiO₂-Fe₃O₄ (c) and MMIP (d).

In curve (d), the peak at 1310 cm⁻¹ is attributed to the stretching of C–O, which shows functional monomers are grafted on the surface of silica. Meanwhile, the peak at 1630 cm⁻¹ is greatly enhanced, which suggests the formation of polymer shells on the surface of the SiO₂-Fe₃O₄. All the results indicate the successful preparation of MMIP.

3.4 TGA analysis

The thermo gravimetric curves of MPS-Fe₃O₄-SiO₂ and MMIP are shown in Fig. 4. Both the samples don't have obvious decline below the 200 °C except for the slight weight loss of physically adsorbed water. From 200 °C to 800 °C, the weight decreased 5.66% in curve (a), which is due to the degradation of silica layer coated on the surface of Fe₃O₄ and the MPS covered on the surface of SiO₂-Fe₃O₄. However, the weight is decreased rapidly in curve (b), especially from 500 °C to 600 °C, which is mainly due to the decomposition of the surface imprinting layer. The results also demonstrate the formation of imprinting layer, which is consistent with the TEM and FTIR analyses.

3.5 XRD analysis

XRD patterns of Fe₃O₄, SiO₂-Fe₃O₄ and MMIP are shown in Fig. 5. Six characteristic peaks appear in the 2θ range of from 20

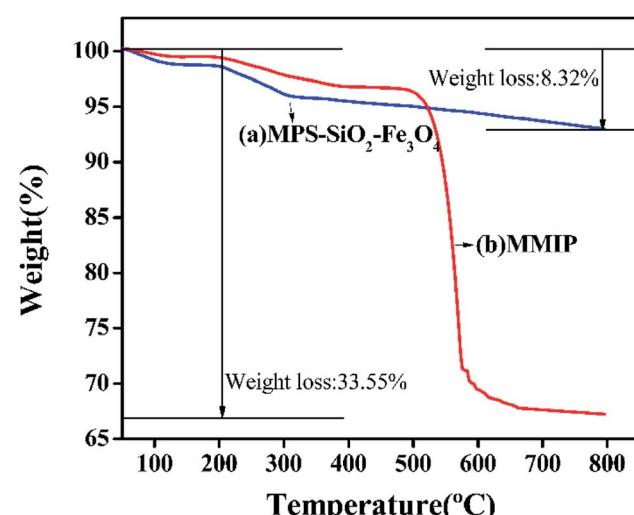


Fig. 4 Thermo gravimetric curves of MPS-Fe₃O₄-SiO₂ (a) and MMIP (b).

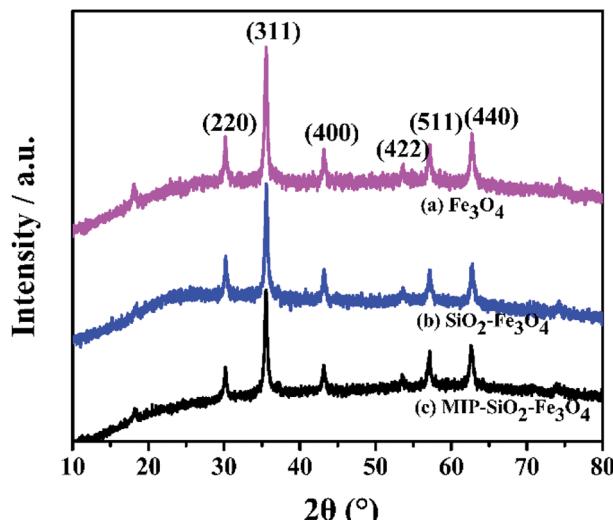


Fig. 5 XRD spectra of Fe_3O_4 (a), $\text{SiO}_2\text{-Fe}_3\text{O}_4$ (b) and MMIP (c).

to 80° , which are indexed as (220), (311), (400), (422), (511) and (440). The peaks are assigned to the diffraction of Fe_3O_4 crystal according to the JCPDS card 19-629, and the planes indicate that Fe_3O_4 crystal is inverse spinel structure.²³ However, it does not exclude the presence of $\gamma\text{-Fe}_2\text{O}_3$, and there may be two types of iron oxide particles of maghemite and magnetite in the synthesized sample.²⁴ The presence of maghemite is due to oxidation of magnetite during the synthesis,²⁵ but both iron oxide particles are magnetic and do not need to be distinguished. After modified with silica, there appears a weaker and broader peak at 2θ values from 21° to 27° , which suggests that the amorphous silica is successfully coated on the surface of Fe_3O_4 . The positions of the diffraction peaks in (a), (b) and (c) are identical, which indicates a sequence of operations do not change the crystal structure of Fe_3O_4 . Compared with (a), the peak strengths of (b) and (c) are decreased significantly, indicating that the silica layer and imprinted layer have a barrier effect for magnetic Fe_3O_4 .

3.6 VSM analysis

The magnetic properties of samples were investigated by VSM at room temperature. Fig. 6 shows the magnetic hysteresis loops of Fe_3O_4 , $\text{SiO}_2\text{-Fe}_3\text{O}_4$ and MMIP. It is apparent that three curves have similar general shape. There is no hysteresis, and the remanence and coercivity are both zero, suggesting that the samples are superparamagnetic. The saturation magnetization values of Fe_3O_4 , $\text{SiO}_2\text{-Fe}_3\text{O}_4$ and MMIP are 72.93 , 60.07 , 43.97 emu g⁻¹, respectively. Because the silica and molecularly imprinted shells on the surface of Fe_3O_4 prevents its magnetic responsiveness, the magnetic property of MMIP is weaker than those of Fe_3O_4 and $\text{SiO}_2\text{-Fe}_3\text{O}_4$. However, MMIP remains strongly magnetic property at room temperature and allows for as effective magnetic separation carrier. The dispersion-agglomeration process of MMIP is also shown in Fig. 6. MMIP is dispersed homogeneously in the solution without external magnetic field. When the magnetic field is applied, the MMIP gathers at right side at around 20 S, and the solution becomes

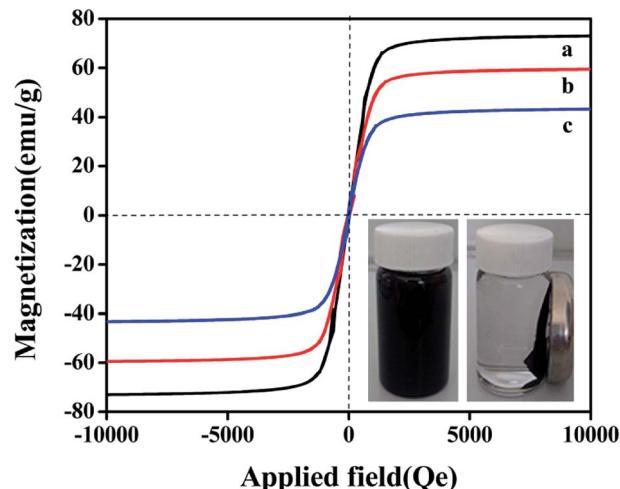


Fig. 6 Magnetization curves of Fe_3O_4 (a), $\text{SiO}_2\text{-Fe}_3\text{O}_4$ (b) and MMIP (c).

clear. The above results show that the synthesized MMIP has strong magnetic properties and can realize convenient magnetic separation.

3.7 Adsorption properties of MMIP and MNIP

3.7.1 Static adsorption experiment. Fig. 7 shows the UV absorption spectrum of DEHP ethanol solution with different concentrations from 5 mg L^{-1} to 50 mg L^{-1} . There are three absorption peaks in the curves. However, there is no red shift at the peak of 225 nm with the increasing of concentration, and the peak strength increases obviously. Hence, the UV absorbance of DEHP is measured at 225 nm . The linear equation of concentration-absorbance is fitted by $y = 0.0225x - 0.0118$ according to the absorbance measured at 225 nm with different concentrations, thereby the conversion between absorbance and concentration is obtained.

Static adsorption experiment was conducted to estimate the maximum adsorption capacity of MMIP for DEHP, and the

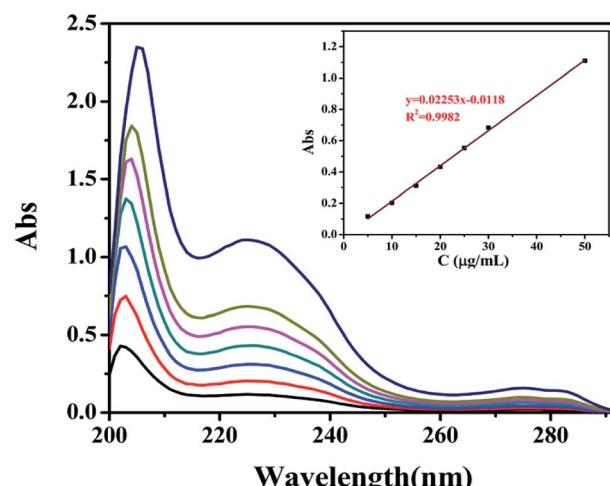


Fig. 7 UV spectra of DEHP ethanol solution with different concentrations and the linear fitting of concentration and absorbance.



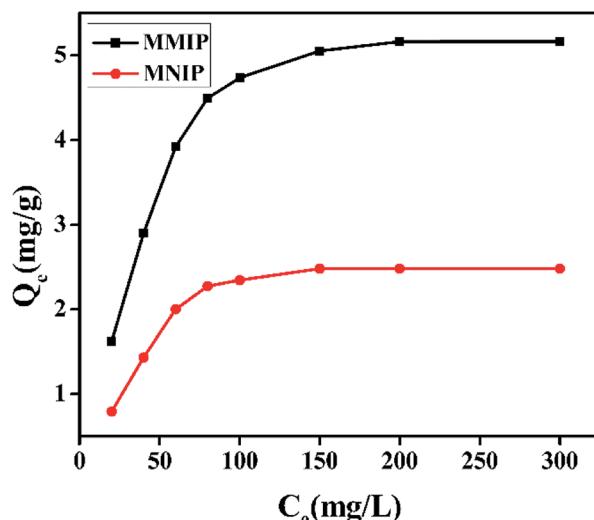


Fig. 8 The adsorption isotherms of MMIP and MNIP for DEHP.

results are displayed in Fig. 8. The binding amounts of MMIP and MNIP for DEHP increase quickly with the increasing of the initial concentration of DEHP, and at higher initial concentration over 200 mg L^{-1} , the adsorption capacity becomes stable. The adsorption capacity of MMIP is higher than that of MNIP at the same concentration due to the obvious difference in the spatial structure of two polymers. The adsorption of DEHP onto MNIP is mainly dependent on the polar groups on the surface of MNIP, which is nonspecific adsorption, and the adsorption ability is relatively weak. The adsorption of DEHP onto MMIP is due to the existence of three-dimension cavities, which are complement with the spatial structure of DEHP, thus the MMIP has “memory function” to template and has strong selective adsorption ability. The difference of adsorption capacity between the two polymers is mainly due to the specific adsorption of the cavities.

Langmuir and Freundlich models are often used to describe the static adsorption process. Linear equation of Langmuir model is listed as follows:

$$\frac{C_e}{Q_e} = \frac{1}{K_L Q_m} + \frac{C_e}{Q_m} \quad (2)$$

where $K_L (\text{L mg}^{-1})$ is the Langmuir constant which relates to the enthalpy of sorption, $C_e (\text{mg L}^{-1})$ is the equilibrium concentration of DEHP, and $Q_m (\text{mg g}^{-1})$ is the maximum binding amount of MMIP or MNIP for DEHP.

Freundlich linear equation is described as follows:

$$\ln Q_e = \ln K_F + \left(\frac{1}{n}\right) \ln C_e \quad (3)$$

Table 1 Langmuir and Freundlich model fitting data onto MMIP and MNIP

Langmuir model				Freundlich model				
K_L	Q_m	R^2	Linear fitting equation	K_F	n_F	R^2	Linear fitting equation	
MMIP	0.1320	5.3591	0.9997	$y = 0.1866x + 1.4133$	1.0002	2.3919	0.9490	$y = 0.4181x + 0.0002$
MNIP	0.0581	2.6960	0.9929	$y = 0.3709x + 6.3796$	0.1866	1.6360	0.9387	$y = 0.6112x - 1.67903$

where $Q_e (\text{mg g}^{-1})$ is the equilibrium binding amount of MMIP or MNIP for DEHP, and $K_F (\text{mg}^{1-n} \text{L}^n \text{g}^{-1})$ and n are the constants of Freundlich sorption. Table 1 shows the simulation data of MMIP and MNIP using Langmuir and Freundlich models. It is observed that the correlation coefficient of Langmuir model fitting for MMIP ($R^2 = 0.9997$) is greater than 0.99 and much larger than that of Freundlich model ($R^2 = 0.9490$). Langmuir fitting model is more suitable for the experimental data of MMIP, which indicates that the adsorption of template molecule on MMIP belongs to monolayer adsorption.

3.7.2 Dynamic adsorption experiment. Dynamic adsorption experiment was carried out to investigate the mass transfer properties of MMIP. Fig. 9(a) shows the adsorption kinetics curves of MMIP and MNIP for 100 mg L^{-1} DEHP ethanol solution. The adsorption capacity of MMIP for DEHP increases rapidly in the first 30 min, then, the curve becomes relatively flat and reaches the adsorption equilibrium at about 40 min. A similar behavior is observed for the adsorption onto MNIP. The results can be explained from the structure of MMIP and MNIP. The shallow cavities on the surface of MMIP and MNIP are favorable to the rapid adsorption for DEHP. However, when the cavities reach saturation, the process of mass transfer is subjected to resistance, and the adsorption rate decreases. Finally, the adsorption reaches adsorption equilibrium, and the adsorption capacity of MNIP increases slowly. The maximum adsorption capacity of MNIP is lower than that of MMIP owing to the absence of specific recognition sites.

Two adsorption kinetic model of pseudo-first-order kinetic model and pseudo-second-order kinetic model are employed to evaluate the dynamic binding processes of MMIP and MNIP,^{26,27} and the linear equation of pseudo-first-order model is expressed as:

$$\ln (Q_e - Q_t) = \ln Q_e - k_1 t \quad (4)$$

where $k_1 (1/h)$ represents the pseudo-first-order kinetic rate constant, $Q_t (\text{mg g}^{-1})$ is the binding amount of MMIP and MNIP for DEHP at contact time t (h).

The linear equation of pseudo-second-order model can be described as:

$$\frac{t}{Q_t} = \frac{1}{k_2 Q_e^2} + \frac{t}{Q_e} \quad (5)$$

where $k_2 (\text{g} (\text{mg}^{-1} \text{h}^{-1}))$ is the pseudo-second-order kinetic rate constant, and $Q_e (\text{mg g}^{-1})$ is the equilibrium adsorption capacity of MMIP and MNIP. The kinetic parameters are displayed in Table 2. The results indicate that the pseudo-second-order model yields better fitting for MMIP and MNIP between



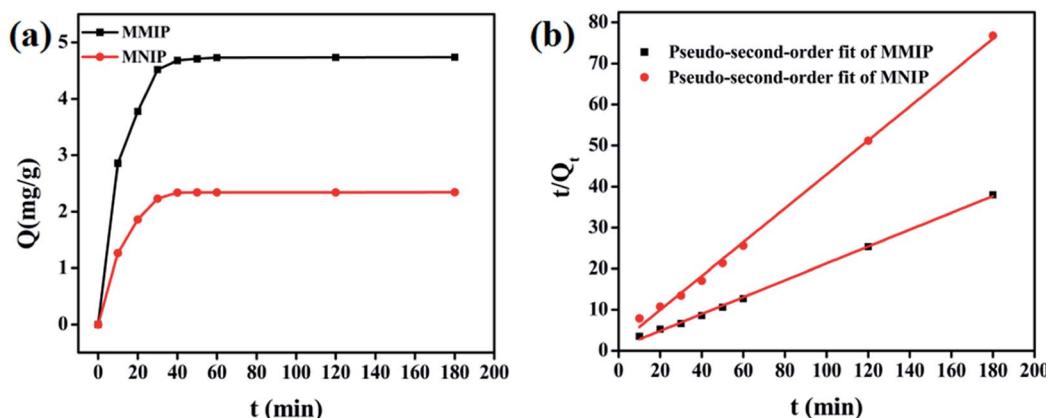


Fig. 9 Dynamic adsorption curves (a) and pseudo-second-order fitting (b) for MMIP and MNIP.

Table 2 Pseudo-first-order and pseudo-second-order kinetic model fitting data for MMIP and MNIP

	Pseudo-first-order model				Pseudo-second-order model			
	k_1	Q_e	R^2	Linear fitting equation	k_2	Q_e	R^2	Linear fitting equation
MMIP	0.1198	7.7967	0.9876	$y = -0.1198x + 2.0537$	0.0576	4.8674	0.9984	$y = 0.2055x + 0.7329$
MNIP	0.1400	4.9105	0.9158	$y = -0.1400x + 1.5914$	0.1010	2.4216	0.9974	$y = 0.4130x + 1.6900$

the two models with correlation coefficients of 0.9984 and 0.9974. The maximum adsorption capacities of 4.87 mg g^{-1} for MMIP and 2.42 mg g^{-1} for MNIP calculated by pseudo-second-order model are close to the experimental maximum adsorption amounts of 4.74 mg g^{-1} for MMIP and 2.35 mg g^{-1} for MNIP. Fig. 9(b) shows the simulations of pseudo-second-order models. The curves illuminate that the adsorption rate of MMIP for DEHP (Q_t/t) is much larger than that of MNIP, which is due to the existence of a large number of imprinted sites on the surface of MMIP.

3.7.3 Adsorption selectivity experiment. The adsorption selectivity of MMIP for DEHP was investigated with the template structural analogues of DAP and DBP (Fig. 10) as competitive substrates. The adsorption capacities of MMIP and MNIP to the mixture of DEHP, DAP and DBP with respective concentration of 100 mg L^{-1} in 2 mL ethanol solution are displayed in Fig. 11. The imprinting factor α is usually used to evaluate the recognition ability of MMIP,¹⁴ and it is defined in the following equation: $\alpha = Q_{\text{MIP}}/Q_{\text{NIP}}$. The Q_{MIP} and Q_{NIP} are the adsorption capacities of template or structural analogues onto MMIP and MNIP. The imprinting factors of DEHP, DAP and DBP are 3.012,

1.547 and 1.788, respectively, which suggests that the MMIP has specific recognition performance to template molecule of DEHP. MMIP creates tailor-made binding sites with the memory of shape, size and functional groups of template molecules.²⁸ In addition to the hydrogen bond formed between the carbonyl group in target molecules and carboxyl group embedded in MMIP, MMIP recognizing targets also depends on the shape and size. From Fig. 10, we can know DBP shares similar shape and functional groups with DEHP. However, the size of DBP is smaller, so MMIP shows lower binding capacity for DBP. The size of DAP is smaller than DBP, the binding capacity is lower than DBP. In the meanwhile, the adsorption capacities of MNIP are very close for those compounds, since there are no specific recognition sites in MNIP and the adsorption is non-selective.

3.8 Regeneration experiment

Reusability is an important index to evaluate the practical value of materials. The MMIP and MNIP were recycled for three times with the mixed solution of methanol and acetic acid ($v:v, 9:1$)

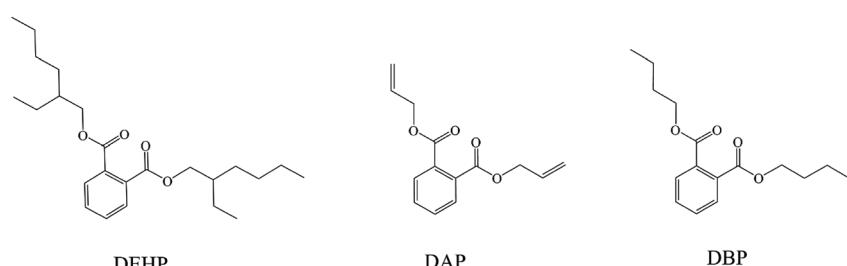


Fig. 10 Structural formula of DEHP and its structural analogues of DAP and DBP.

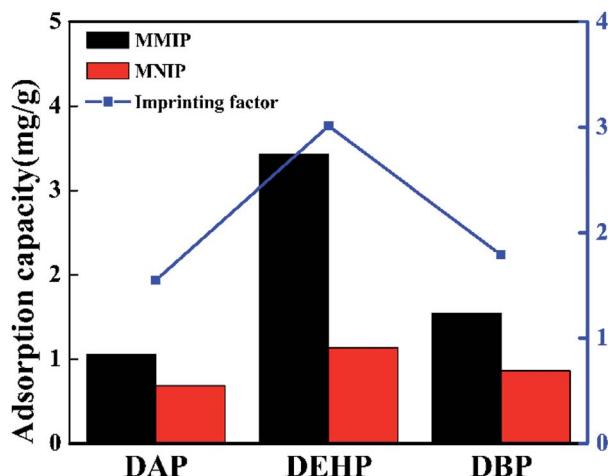


Fig. 11 Adsorption selectivities for DEHP and competing compounds onto MMIP and MNIP in the mixture solution.

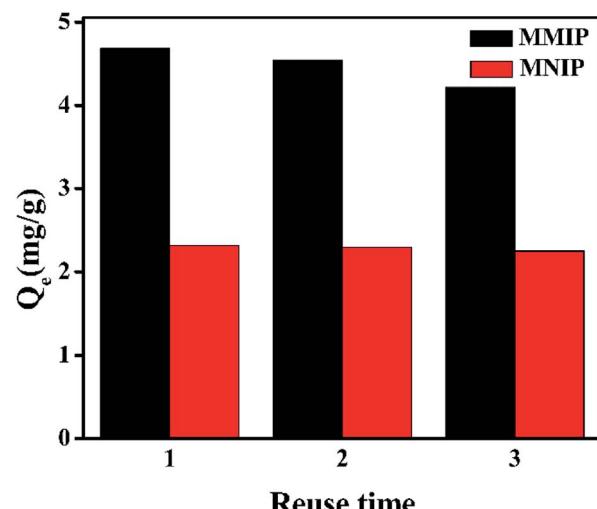


Fig. 12 Regeneration of MMIP and MNIP.

as eluent. The results are shown in Fig. 12, the adsorption capacities of DEHP on MMIP lost only 10% after three cycles, which indicates that MMIP has good and stable regeneration performance. However, the adsorption capacity of MNIP has little change. This is due to the fact that some recognition sites on the surface of MMIP are destroyed or occupied by substrate molecules during the repeated elution. Instead, MNIP is relied on nonspecific adsorption in the absence of specific recognition sites.

4 Conclusions

In this paper, a novel MMIP was successfully fabricated with double core-shell structure for convenient and efficient separation of DEHP from the solution by external magnetic field. The adsorption properties of MMIP and MNIP were compared by a series of static and dynamic adsorption experiments, and the adsorption selectivity of MMIP was investigated. The adsorption capacity of MMIP is always higher than that of MNIP

at the same conditions. Also, Langmuir model is more suitable for the static adsorption of MMIP, which belongs to the monolayer adsorption, and the dynamic adsorption process fits the pseudo-second-order kinetic model well. Compared with DAP and DBP, the adsorption capacity of DEHP on MMIP is higher. The imprinting factors of DEHP, DAP and DBP are 3.012, 1.547 and 1.788, respectively, which indicates that MMIP has good adsorption selectivity for template molecules. The regeneration experiment shows that the adsorption capacity of MMIP has little change after three cycles, which illuminates that MMIP has good application prospect for convenient separation for DEHP.

Conflicts of interest

There are no conflicts to declare.

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