



Cite this: *RSC Adv.*, 2018, 8, 35226

Tunable transport and optoelectronic properties of monolayer black phosphorus by grafting PdCl₂ quantum dots†

Cuicui Sun,^{id} Yuxiu Wang, Yingjie Jiang, Zhao-Di Yang, Guiling Zhang* and Yangyang Hu^{id}*

The electronic, transport, and optoelectronic properties of monolayer black phosphorus (MLBP) are much influenced by grafting PdCl₂ groups, demonstrated here by using density functional theory (DFT) and non-equilibrium Green's function (NEGF) as well as the Keldysh Nonequilibrium Green's Functions (KNEGF) methods. We find that the PdCl₂ groups prefer to locate over the furrow site of MLBP and form a planar

quadridentate structure of . The PdCl₂ groups serve as quantum dots by introducing discrete

flat levels between the MLBP valence band and the Fermi level (E_F). The conductivity is much lowered after attaching PdCl₂ quantum dots, due to the fact that the scattering effect of PdCl₂ plays a major role in the process of electron transporting. A threshold voltage is found for the functionalized system with a large density of PdCl₂ quantum dots, a valuable clue for exploring current switches. However, no evident threshold voltage is found for the pure MLBP. Electrons permeate easier through the armchair direction compared with the zigzag either in the pure MLBP or in the functionalized composites. More importantly, grafting PdCl₂ quantum dots is very beneficial for enhancing photoresponse. The values of photoresponse for the modified species are about 20 times higher than the free MLBP. A significant photoresponse anisotropy is observed for both MLBP and n PdCl₂-BP ($n = 1, 2$, and 4), contrary to the conductivity, the zigzag direction shows much stronger photoresponse than the armchair. All of the aforementioned unique properties make these new two-dimensional (2D) MLBP based materials especially attractive for both electronic and optoelectronic devices.

Received 23rd August 2018
Accepted 3rd October 2018

DOI: 10.1039/c8ra07053a

rsc.li/rsc-advances

1. Introduction

Inspired by the great achievements of graphene, two-dimensional (2D) nanomaterials have received great attention in the fields of micro-/nano-electronics and optoelectronics due to their distinct confinement effect of electrons in two dimensions,¹ strong in-plane covalent interactions,² large lateral size,³ high exposure of surface,⁴ and so forth. After graphene with zero band gap⁵ and molybdenum disulphide with relatively low carrier mobility,⁶ monolayer black phosphorus (MLBP) has emerged as an important member of 2D materials family⁷ in recent years. In this layer, each P atom bonds to three neighboring P atoms through sp^3 hybridized orbitals forming warped hexagons in a nonplanar fashion. To date, many production methods for MLBP have been developed, including liquid

exfoliation,^{8–12} Ar⁺ plasma thinning process,¹³ photochemical etching,¹⁴ and mechanical exfoliation,^{15,16} etc. Importantly, it has enormous potential to be used in high-performance electronic and optoelectronic devices because of its direct tunable band gap ranging from 0.3 to 2.0 eV,^{17,18} high carrier mobility of $\sim 1000 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ at room temperature,^{7,17,19} and anisotropic optical characteristics^{20,21} that could complement or exceed graphene and molybdenum disulphide in the next generation of novel devices.

In the electronic arrangement, the P atom with a valence shell configuration $3s^2 3p^3$ has five valence electrons available for bonding, and the existence of the lone pair makes MLBP reactive to air. Functionalization is proved to be a controllable and effective method to not only improve the stability but also to enrich the property^{22,23} of 2D materials. Tunability of electronic properties of 2D materials is crucial for their practical applications in electronics and optoelectronics. Up to now, many functionalized 2D black phosphorus (BP) have been reported, such as metal/BP contact systems,^{24–27} alkali metal/nonmetal atom doped BP,^{28–31} carbon nanotube/BP composites,³² as well as hBN/BP^{33–35} and MoS₂/BP^{36,37} heterojunctions.

School of Materials Science and Engineering, College of Chemical and Environmental Engineering, Harbin University of Science and Technology, Harbin 150080, China.
E-mail: guiling-002@163.com

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c8ra07053a



Out of the above mentioned series, the most recent addition to the family are the functionalization of BP^{38,39} by adsorbing transition metal atoms on surface.^{40–44} It has been found that palladium (Pd) adatom adsorption on MLBP presents excellent electronic and optoelectronic behaviors.^{41–43} Up to now, research on functionalization of MLBP is predominantly focused on its physical modification; far less information is available regarding chemical modification, which is determined to be efficient in manipulating electronic and optoelectronic properties.⁴⁵

In general, the Pd atom is a typical transition metal for coordinating P and Cl ligands.^{46–48} These complexes have shown wide spread applications in electronics, photoelectrochemistry, and electrochemistry.^{49–51} Usually, Pd has a 4d¹⁰ configuration and preferentially provides a dsp² hybrid environment to coordinate with four ligands, *i.e.*, forms a planar quadridentate complex. The exposed P atoms of MLBP are perfect ligands by providing lone pair electrons to coordinate with the Pd atom of

PdCl₂ to form a planar quadridentate structure of $\begin{pmatrix} \text{Cl} & \text{P} \\ & \text{Pd} \\ \text{Cl} & \text{P} \end{pmatrix}$. Of

particular interest in this paper is investigation of covalent functionalization of MLBP with PdCl₂ groups including the effects of PdCl₂ on electronic, transport, and optoelectronic properties.

2. Models and computational methods

For the periodic systems, we constructed the supercell as 2 × 4 (*x* × *y*) containing 32 P atoms (Fig. 1). We selected *n* PdCl₂ groups (referred as *n*PdCl₂, where *n* = 1, 2, and 4) to graft MLBP per supercell to study the effects of grafting density *n*. Hereafter, the functioned structures are denoted as *n*PdCl₂-BP (*n* = 1, 2, and 4) for simplification. For the geometrical optimization, the PdCl₂ was put at each adsorption site with all the atomic coordinates in the system being relaxed. The maximum force and maximum stress were set to the same value of 0.02 eV Å^{−1}.

It has been reported that the MLBP displays anisotropic properties due to its anisotropic structure,^{52,53} which triggered us to calculate transport and optoelectronic properties for two mutually perpendicular directions, *i.e.*, the armchair and zigzag directions (Fig. 2(a and c) vs. 2(b and d)). So we curved a nearly square plane (~26.04 Å × 26.50 Å) of MLBP and *n*PdCl₂-BP (*n* = 1, 2, and 4) as the scatter region to be sandwiched between two Au (100)-(9 × 3) electrodes to build two-probe devices. Such scatter region was formed by repeating each optimized supercell 3 times at the armchair direction and 2 times at the zigzag direction. When the scatter region used the zigzag edge to connect with the Au electrodes, the armchair direction properties were calculated (denoted as a-MLBP and a-*n*PdCl₂-BP (*n* = 1, 2, and 4)). Another was just the reverse (nominated as z-MLBP and z-*n*PdCl₂-BP (*n* = 1, 2, and 4)). As a benchmark test, the *E*_f positions of the two-probe systems with two or four buffer layers of Au were calculated at 0.0 V. We found that the *E*_f were almost locate at the same positions for the two cases (−3.533772 and −3.533858 eV for the two and four buffer layers, respectively). So we used two buffer layers of Au to perform further calculations. The nearest P–Au distance is 2.37 Å, corresponding to their covalent bonding.⁵⁴ Transport current was computed by changing the applied bias voltage in the step of 0.2 V in the range of −1.0 to 1.0 V.

Electronic and transport property computations were performed by using the software package Atomistix ToolKit (ATK)^{55–58} based on a basis of the combination of density function theory (DFT) and non-equilibrium Green's function (NEGF) methods. The on-site correlation effect among 4d electrons of the Pd atom was accounted for by using the GGA+U scheme⁵⁸ where the parameter U–J was set to be 6.0 based on the literature (6.0),⁵⁹ which was also close to that (5.77) in the ATK_U database. The Perdew–Burke–Ernzerhof (PBE) exchange–correlation functional was employed. A single-ζ basis with polarization (SZP) was used for all atoms. A (7 × 5 × 1) *K*-point in the Brillouin zone (*x*, *y*, and *z* directions, respectively) was adopted, and 150 Ry cutoff energy was applied to describe the periodic wave function.

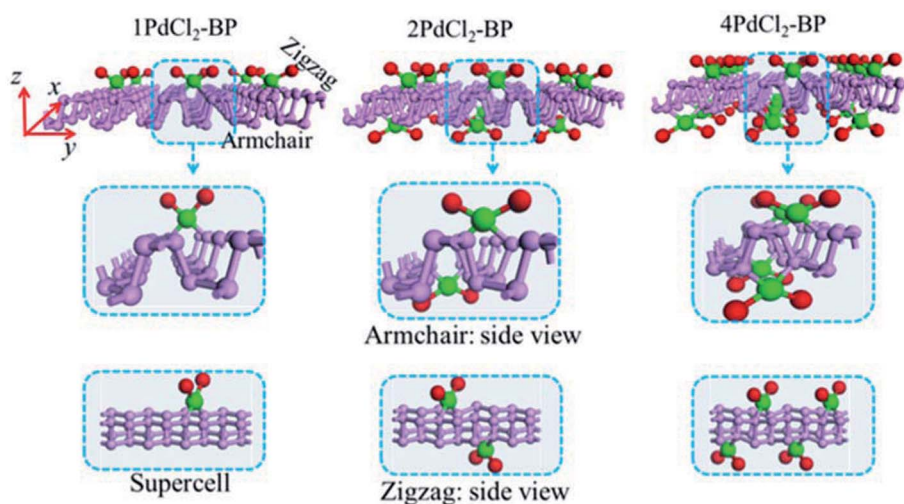


Fig. 1 Optimized structures of three most stable systems of 1PdCl₂-BP, 2PdCl₂-BP, and 4PdCl₂-BP.



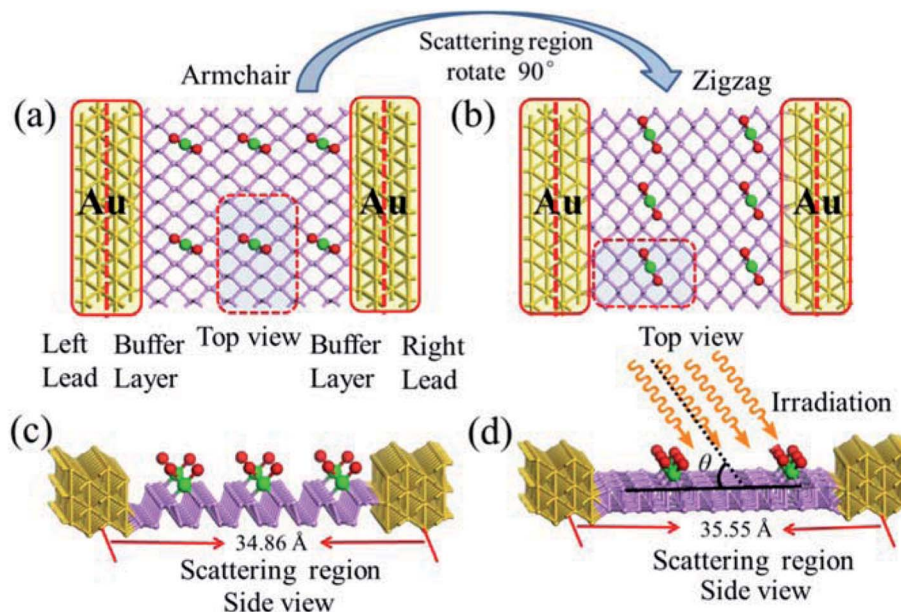


Fig. 2 Exemplified two-probe devices by 1PdCl₂-BP. (a) Top view of a-1PdCl₂-BP; (b) top view of z-1PdCl₂-BP; (c) side view of a-1PdCl₂-BP; (d) side view of z-1PdCl₂-BP. Irradiation on the scatter regions of two-probe devices are schematically shown in (d).

To gain an insight into the optoelectronic properties of MLBP and n PdCl₂-BP ($n = 1, 2$, and 4), we irradiated the scatter regions of above mentioned two-probe systems with linearly polarized light (exemplified as Fig. 2(d)). The polarized angle θ was assigned between the scatter region plane and the irradiation direction. The photo energy was set from 0.0 to 1.0 eV with an interval of 0.1 eV. We calculated the photocurrent varying with θ from 0° to 180° at the 0.0 V bias voltage.

The optoelectronic properties of a-MLBP, z-MLBP, a- n PdCl₂-BP, and z- n PdCl₂-BP ($n = 1, 2$, and 4) were evaluated using the Nanocal⁶⁰ package which carried out DFT within the Keldysh Nonequilibrium Green's Functions (KNEGF).^{60,61} Standard norm-conserving nonlocal pseudopotentials were adopted to define the atomic cores, and SZP linear combination of atomic orbital basis set was used to expand physical quantities. The exchange–correlation potential was treated at the GGA+U with PBE.

3. Results and discussions

3.1 Stability and geometry

The MLBP is not residing in a flatland, instead, it forms a puckered hexagonal structure. The optimized supercell vectors of MLBP and n PdCl₂-BP ($n = 1, 2$, and 4) were supplied in Table S1 in the ESI†. On the surface of this supercell one, two, or four PdCl₂, corresponding to $n = 1, 2$, or 4 , respectively, are considered to coordinate with the P atom for investigating the effect of grafting density. Taking into account all the circumstances, 20 configurations are computed for n PdCl₂-BP ($n = 1, 2$, and 4), *i.e.*, 2 for 1PdCl₂-BP, 10 for 2PdCl₂-BP, and 8 for 4PdCl₂-BP. Table S2 in the ESI† summarizes the optimized total energies of supercells of these 20 configurations. We find that the PdCl₂ groups prefer to locate over the furrow site rather than

upon the ridge position of the MLBP as the former has lower

total energy than the latter. A planar structure of $\begin{pmatrix} \text{Cl} & \text{Pd} \\ | & | \\ \text{Cl} & \text{P} \end{pmatrix}$ is formed in these n PdCl₂-BP complexes, a typical quadridentate structure for Pd. As for 2PdCl₂-BP and 4PdCl₂-BP, the PdCl₂ components tend to evenly and staggerly distributed on both sides of MLBP. Finally, we chose three most stable configurations for $n = 1, 2$, and 4 , respectively, as given in Fig. 1, to further investigate the electronic, transport, and optoelectronic properties.

For the three most stable structures of 1PdCl₂-BP, 2PdCl₂-BP, and 4PdCl₂-BP, we have calculated the binding energies, E_b , by using $E_b = E_{n\text{PdCl}_2\text{-BP}} - E_{\text{MLBP}} - nE_{\text{PdCl}_2}$, where $E_{n\text{PdCl}_2\text{-BP}}$, E_{MLBP} , and E_{PdCl_2} denote the total energies of the functioned systems, pristine MLBP, and PdCl₂ in the unit cell, respectively. For the PdCl₂ calculation, the unit cell was set to be $30 \times 30 \times 30$ Å, so that PdCl₂ can be viewed as an isolated system. The calculated binding energies are all negative with increasing exoenergetic values of -1.95 , -3.83 , and -7.35 eV from $n = 1, 2$ to 4 (Table 1), demonstrating that grafting PdCl₂ groups to the MLBP surface is energetically favorable.

After optimizing, the P–P, Pd–P, Pd–Cl, and Pd–Pd bond lengths as well as the Cl–Cl distances for MLBP and n PdCl₂-BP ($n = 1, 2$, and 4) are collected in Table 1. MLBP consists of double puckered planes; each atom is bound to two in-plane atoms with a bond length of $r_{\text{P-P}}^{\text{in}} = 2.224$ Å and to one out-of-plane atom with a bond length of $r_{\text{P-P}}^{\text{out}} = 2.244$ Å, fully in agreement with experimental observation.⁶² For n PdCl₂-BP ($n = 1, 2$, and 4), the Pd–P bond lengths are $r_{\text{Pd-P}} = 2.363$ – 2.374 Å, in line with genral Pd–P chemical bond lengths calculated by M. A. Carvajal *et al.*⁶³ (2.221–2.476 Å) and experimental data reported by V. I. Bakhmutov *et al.* (2.290–2.520 Å).⁴⁷ This shows that the



Table 1 Calculated binding energies E_b and geometries for MLBP and $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4)

Species	E_b (eV)	$r_{\text{P-P}}^{\text{in}}$ (Å)	$r_{\text{P-P}}^{\text{out}}$ (Å)	$r_{\text{Pd-P}}$ (Å)	$r_{\text{Pd-Cl}}$ (Å)	$D_{\text{Pd-Pd}}$ (Å)	$D_{\text{Cl-Cl}}$ (Å)
MLBP	—	2.224	2.244	—	—	—	—
1PdCl ₂ -BP	−1.95	2.264	2.307	2.363	2.447	—	—
2PdCl ₂ -BP	−3.83	2.268	2.345	2.364	2.442	6.359	7.667
4PdCl ₂ -BP	−7.35	2.271	2.365	2.373	2.438	6.627	6.205

covalent bonds are formed between grafted PdCl₂ groups and the P atoms. These P atoms (bonded to Pd) are pushed slightly inward (toward another plane), consequently, the $r_{\text{P-P}}^{\text{in}}$ and $r_{\text{P-P}}^{\text{out}}$ are elongated to 2.264–2.271 Å and 2.307–2.365 Å, respectively. The Pd–Cl bond lengths in $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) are about $r_{\text{Pd-Cl}} = 2.429\text{--}2.457$ Å, in accordance with literature reported values of 2.333–2.575 Å.⁶³ For $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4), the nearest Cl–Cl distances ($D_{\text{Cl-Cl}}$) or Pd–Pd distances ($D_{\text{Pd-Pd}}$) between two adjacent PdCl₂ are all longer than 6.0 Å, implying negligible PdCl₂–PdCl₂ interactions in these considered systems.

3.2 Band structure

The band structure and projected density of states (PDOS) as well as the Kohn–Sham orbitals near the Fermi level (E_f) of MLBP and $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) were calculated as shown in Fig. 3 and 4. Clearly, the MLBP shows a semiconductive feature with a direct band gap (E_g) of 0.85 eV, in agreement with the previous report value about 0.90 eV.^{64–66} Decoration of PdCl₂ groups on MLBP surface can tune the transport property by altering the electronic structure. They may carry properties exhibited by each component as well as new properties generated as a result of their combination.

We can evidently find that $n\text{PdCl}_2\text{-BP}$ introduces flat PdCl₂ bands just above the MLBP valence band and below the E_f , indicating that the PdCl₂ groups contribute localized states in the MLBP band gap region. The Kohn–Sham orbitals also clearly demonstrate the localized character of PdCl₂ groups. The

dispersion of some PdCl₂ band structures around the Γ may be attributed to the large interaction between the Pd and Cl atoms. Correspondingly, PdCl₂ groups give sharp PDOS peaks as localized valence states. In this sense, the anchored PdCl₂ groups behave as quantum dots. These quantum dots mainly originate from the Cl-3p state as can be seen from its large PDOS. In general, quantum dots can form ionic like or covalent like interaction.⁶⁷ In the former case, the electrons are localized on the individual dots which exhibit a weak tunnel coupling. For the latter, quantum dot electron states are quantum-mechanically coupled and show strong tunnel coupling. Clearly, $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) systems bear spatially separated PdCl₂ quantum dots with localized and discrete energy levels owing to the large separations between PdCl₂ quantum dots, electron tunneling through direct coupling of PdCl₂ is minor.

For a general planar quadridentate complex, the ligand field splitting of the Pd 4d electrons results in molecular orbitals being formed of principally a character of e_g ($d_{yz,zx}$), a_{1g} (d_{z^2}), b_{2g} (d_{xy}), and b_{1g} ($d_{x^2-y^2}$). The former e_g ($d_{yz,zx}$) and a_{1g} (d_{z^2}) are electron occupied leaving b_{2g} (d_{xy}) and b_{1g} ($d_{x^2-y^2}$) to be unoccupied. We can see from the PDOS, clearly, the e_g ($d_{yz,zx}$) molecular orbitals locate just below the E_f . Usually, the Pd quadridentate compounds belong to inner-orbital complexes with a zero magnetic moment. In $n\text{PdCl}_2\text{-BP}$, the Cl₂P₂ ligands generate an uneven ligand field leading to a so-called Jahn–Teller effect. Pd forms a dsp^2 hybrid environment to interact with the 3p orbitals of the P and Cl atoms. As a result, the

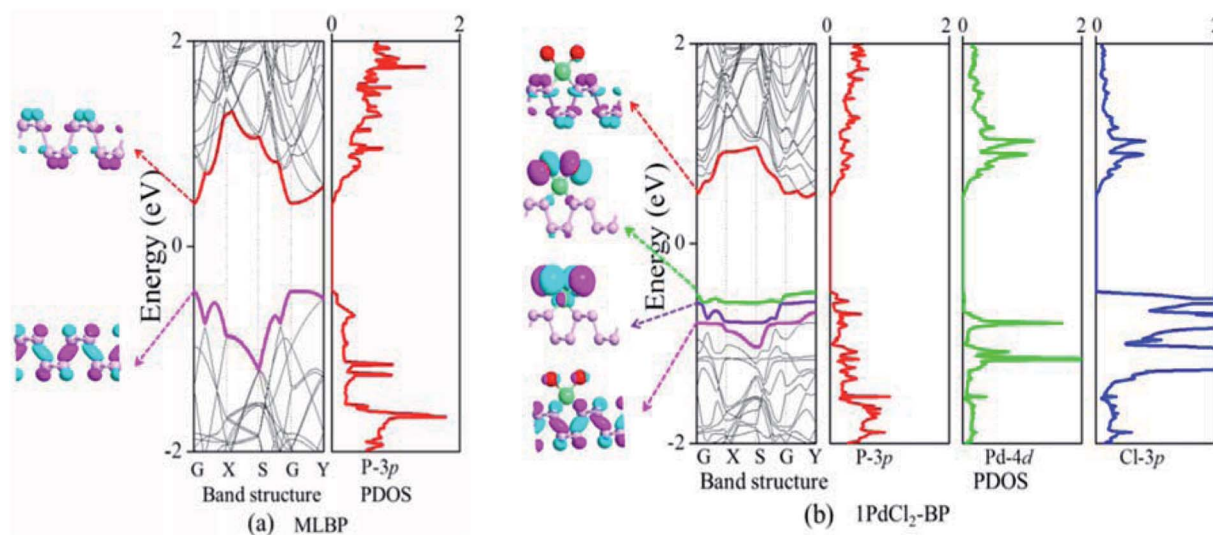


Fig. 3 For (a) MLBP and (b) 1PdCl₂-BP, computed band structures and PDOS as well as the Kohn–Sham orbital corresponding to the energy levels (highlighted in color lines) near E_f at the Γ point. The iso-surface value is 0.028 ($e \text{ Å}^{-3}$).



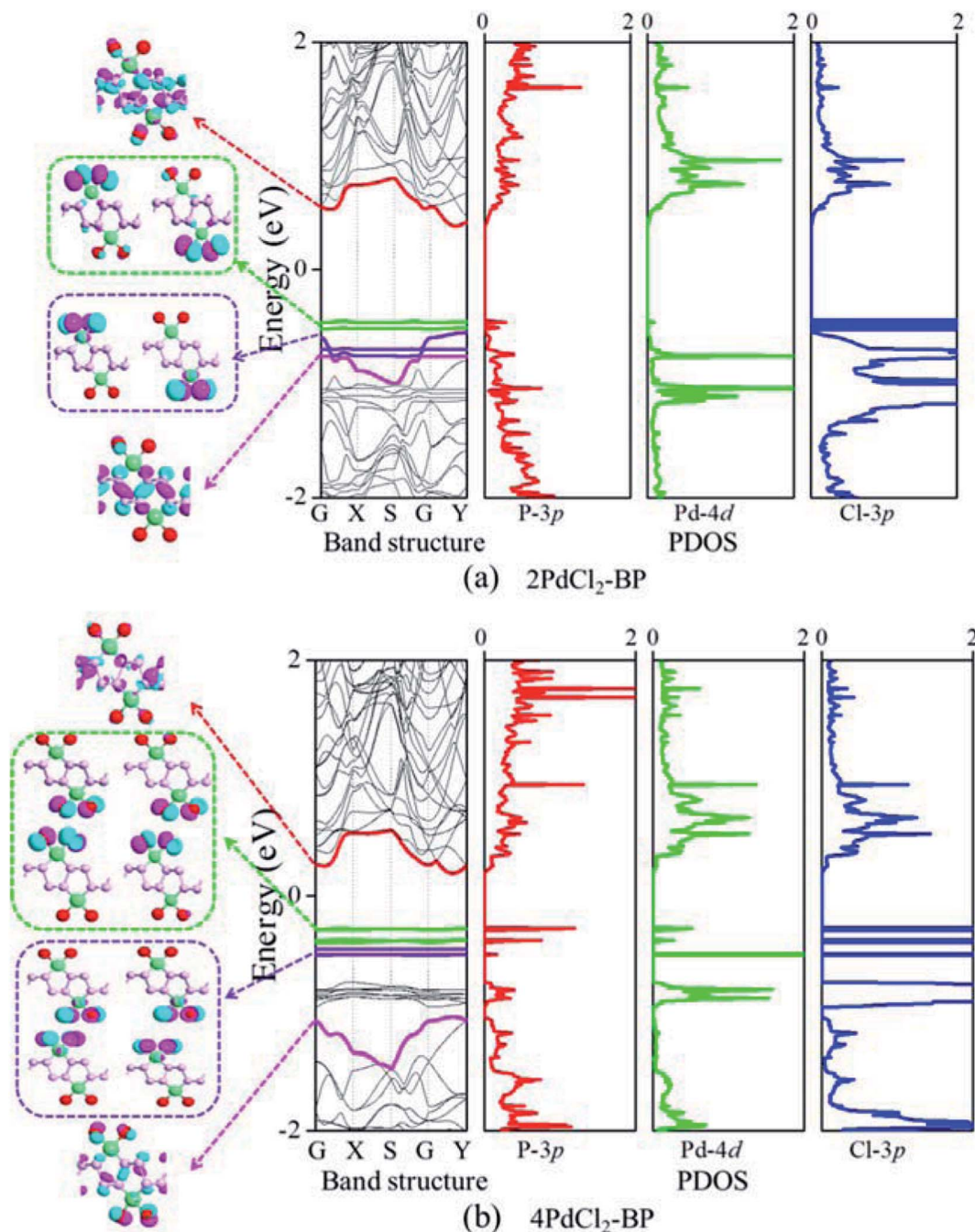


Fig. 4 For (a) $2\text{PdCl}_2\text{-BP}$ and (b) $4\text{PdCl}_2\text{-BP}$, computed band structures and PDOS as well as the Kohn–Sham orbital corresponding to the energy levels (highlighted in color lines) near E_f at the Γ point. The iso-surface value is $0.028 \text{ (e \AA}^{-3}\text{)}$.

degenerate orbitals of e_g ($d_{yz, zx}$) are again split. Therefore, discrete bands appear in the MLBP band gap region of $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4).

To further investigate the interactions between MLBP and PdCl_2 , we calculated the electron difference densities for $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) systems and the results are plotted in Fig. 5. The electron difference density refers to the difference between the self-consistent valence charge density and the superposition of atomic valence density, which indicates the coupling between the atoms in a certain system. The green area means no electron transfer, red part represents obtaining

electrons, while blue region indicates losing electrons. Charge transfers from the MLBP substrate to the PdCl_2 quantum dots can be observed from Fig. 5, suggesting a strong coupling between MLBP and PdCl_2 , which is an important factor in tuning the electronic and optoelectronic properties.

The $n\text{PdCl}_2\text{-BP}$ composites show p-type characteristics accompanied by the appearance of two kinds of midgap energy levels below the E_f (green and purple lines in Fig. 3 and 4). The higher one is close to the E_f and dominated by the Cl-3p state while the lower one is controlled by the Pd-4d state. The MLBP valence band is buried under the PdCl_2 midgap energy levels



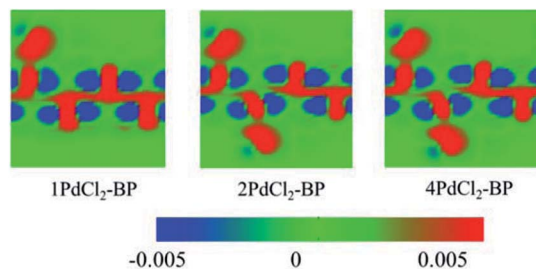


Fig. 5 Computed electron difference densities of 1PdCl₂-BP, 2PdCl₂-BP, and 4PdCl₂-BP. The green area means no electron transfer, red part represents obtaining electrons, while blue region indicates losing electrons.

and hybrids to some extent with Pd-4d/Cl-3p. Therefore, electrons can be transferred from the MLBP substrate to the PdCl₂ quantum dots owing to the strong electronegativity of the Cl atom. The substrate MLBP and the appendant PdCl₂ groups in *n*PdCl₂-BP all contribute to the conduction band. The band gap values of MLBP in *n*PdCl₂-BP do not bring much change relative to the pure MLBP, varying from 1.14 to 1.22 eV. However, the MLBP in *n*PdCl₂-BP becomes to an indirect band gap semiconductor due to the orbital coupling of P with PdCl₂, different from the direct feature in the pristine MLBP. In this case, the PdCl₂ quantum dots play two-fold effects on the transport property. On one hand, they offer extra valence electrons to be excited to the conduction bands to participate in transporting which is beneficial for enhancing the conductivity. On the other hand, the electron excitation of PdCl₂ midgap levels leaves holes for trapping electrons from MLBP, that is, PdCl₂ quantum dots behave as additional scattering centers which degrade the mobility of the charge carriers. This is useful for devices which require fast switch off times.⁶⁸ From the analysis above, it is concluded that the PdCl₂ quantum dots exert an important effect in electron transporting, furthermore, such influence becomes more significant with the increasing grafting density *n*. This can be further confirmed by the calculated transport properties as addressed in the following section.

3.3 Transport property

To compute the electron transport properties of MLBP and *n*PdCl₂-BP (*n* = 1, 2, and 4), the two-probe devices were constructed by sandwiching a curved 2D structure (26.05 Å × 26.51 Å) of MLBP and *n*PdCl₂-BP (*n* = 1, 2, and 4) between two Au electrodes as shown by Fig. 2. The current *I* through the scatter region was calculated based on the formula (1):⁵⁷

$$I = \frac{2e^2}{h} \int_{-\infty}^{\infty} [f(E(V) - \mu_L(V)) - f(E(V) - \mu_R(V))] T(E, V) dE \quad (1)$$

where *f* is the Fermi function; $\mu_{L(R)}$ is the chemical potentials of left (right) electrode; *T*(*E*, *V*) is the transmission function for electrons with energy *E* at certain bias *V*.

The transmissions along the armchair (*a*-*n*PdCl₂-BP) and zigzag (*z*-*n*PdCl₂-BP) directions are both considered. The calculated current-voltage (*I*-*V*) curves are given in Fig. 6. As for MLBP, electrons permeate easier through the armchair direction compared with the zigzag. For instance, at -1.0 V bias voltage, the current magnitude of *a*-MLBP is -28.88 μA, while it goes down to -13.03 μA in the *z*-MLBP device. This anisotropic phenomenon is still preserved after being pinned with PdCl₂ quantum dots. For example, under -1.0 V bias voltage, *a*-4PdCl₂-BP gives a current of -10.01 μA while *z*-4PdCl₂-BP presents a lower data of -2.82 μA. It is noteworthy that grafting PdCl₂ quantum dots to the MLBP surface could significantly influence the transport properties. On one hand, the conductivity is much lowered after anchoring PdCl₂ quantum dots, and the current magnitudes at a certain bias voltage follow the sequence of MLBP > 1PdCl₂-BP > 2PdCl₂-BP > 4PdCl₂-BP, indicating that the scatter effect of PdCl₂ plays a major role in the process of electron transporting. On the other hand, the threshold voltage becomes more and more distinct with the increasing grafting density *n*. Pure MLBP shows continuously rising current from 0.2 to 1.0 V bias voltage, no evident threshold voltage is found. Turn to 4PdCl₂-BP, off-state is kept until the applied bias reaches to ±0.8 V, on-state begins at ±1.0 V bias voltage. Prospectively, grafting MLBP with a large

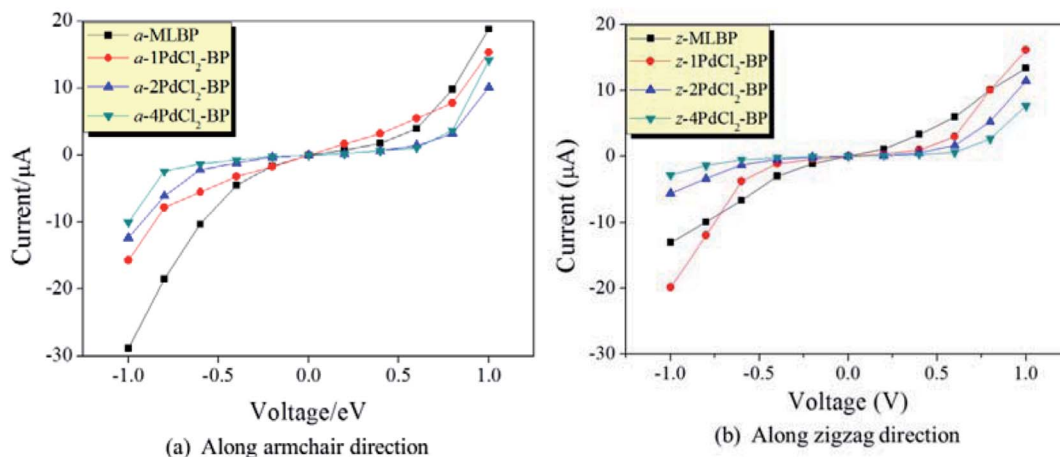


Fig. 6 Computed *I*-*V* curves for MLBP and *n*PdCl₂-BP (*n* = 1, 2, and 4) two-probe devices. (a) Transport along the armchair direction, and (b) transport along the zigzag direction.



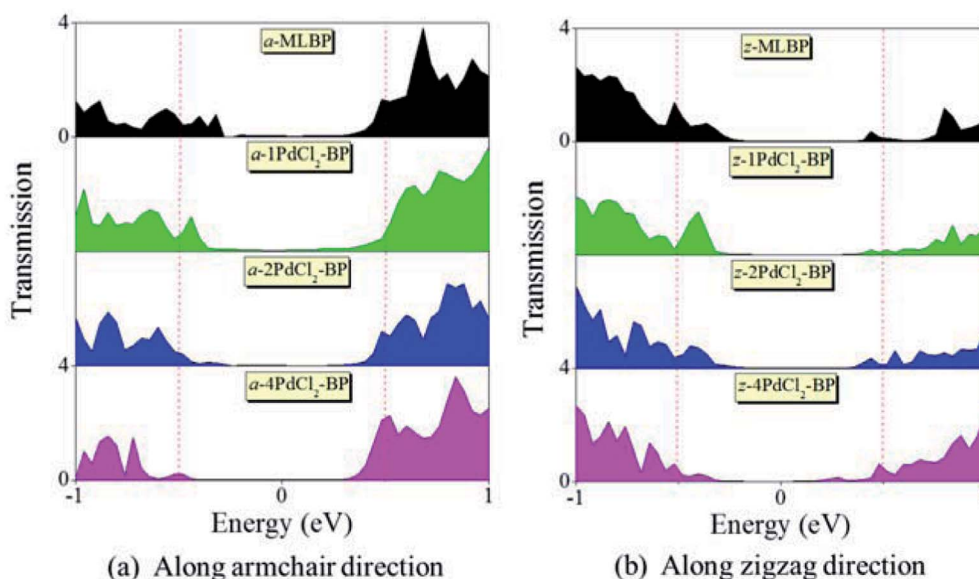


Fig. 7 Transmission spectra of MLBP and n PdCl₂-BP ($n = 1, 2$, and 4) two-probe devices at 1.0 V bias voltage. (a) Transport along the armchair direction, and (b) transport along the zigzag direction. Red dashed lines indicate the bias window.

density of PdCl₂ quantum dots paves a new way toward achieving electronic switch devices.

To further shed light on effects of the PdCl₂ grafting upon the transport properties, the transmission spectra (TS) for MLBP and n PdCl₂-BP ($n = 1, 2$, and 4) at 1.0 V were calculated, as shown in Fig. 7. Usually, resonant peaks in the bias window contribute to the current. Here, the bias window refers to $[-V/2, V/2]$. Clearly, either for a - n PdCl₂-BP device or for z - n PdCl₂-BP device, the valence band resonant peaks (below the E_f) in the

bias window shrink gradually with the growing number n , again indicating the trap effect of PdCl₂ quantum dots. Evidently, the conduction band resonant peaks (above the E_f) in the bias window of a - n PdCl₂-BP are much larger than those of z - n PdCl₂-BP, and thereby the armchair direction gives a higher conductivity than the zigzag, in line with the I - V curves. These features can be further confirmed by the local density of state (LDOS) at the E_f (Fig. 8). Clearly, the LDOS of the scatter regions of both a -

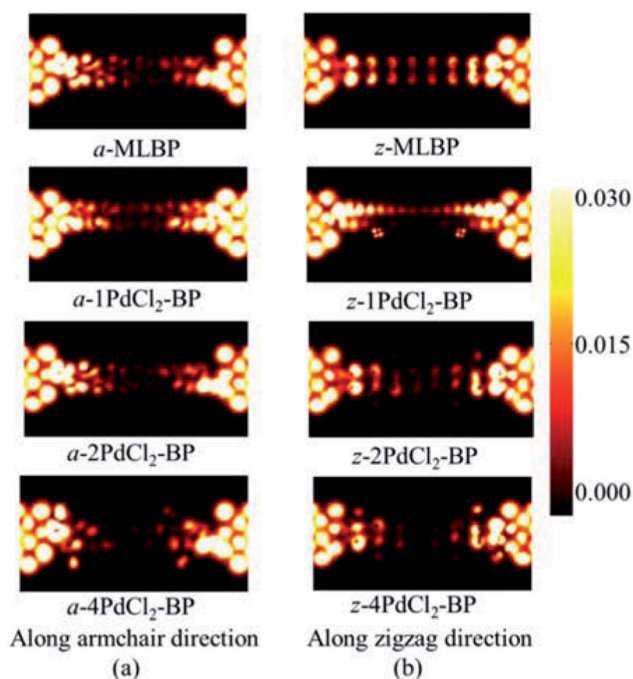


Fig. 8 The computed LDOS at the Fermi level E_f for a -MLBP, z -MLBP, a - n PdCl₂-BP, and z - n PdCl₂-BP ($n = 1, 2$, and 4) two-probe devices at 1.0 V bias voltage.

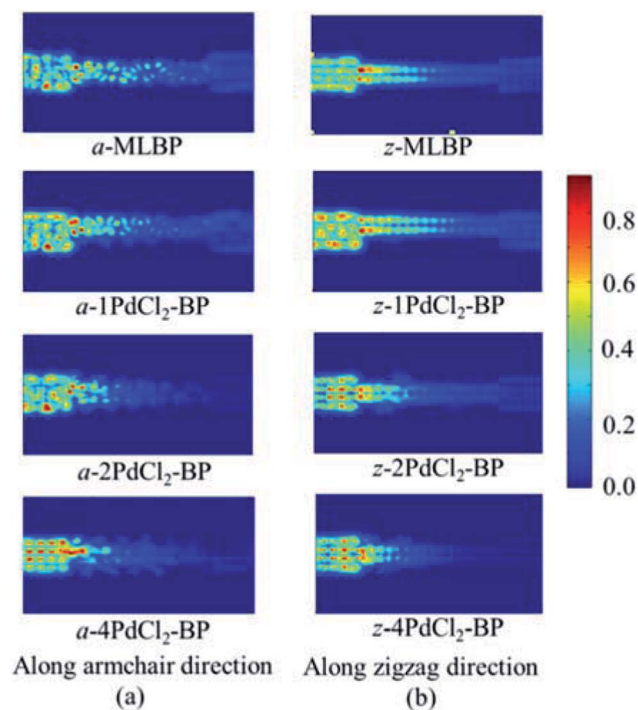


Fig. 9 The real space scattering states at Fermi energy E_f of a -MLBP, z -MLBP, a - n PdCl₂-BP, and z - n PdCl₂-BP ($n = 1, 2$, and 4) two-probe devices.



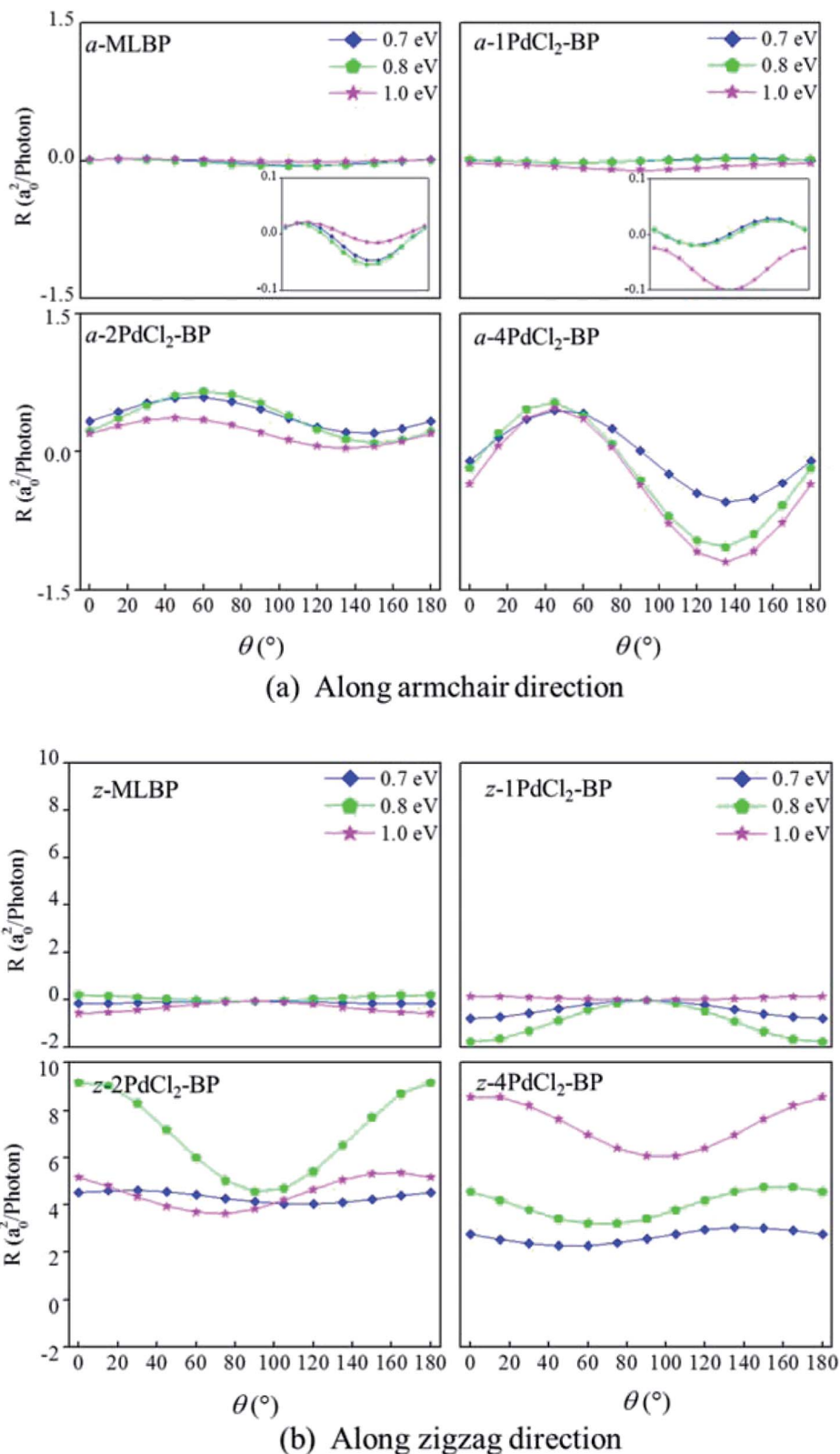


Fig. 10 The calculated photoresponse R for MLBP and $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) two-probe devices irradiated by linearly polarized light. (a) Photoresponse along the armchair direction, and (b) photoresponse along the zigzag direction.

$n\text{PdCl}_2\text{-BP}$ and $z\text{-}n\text{PdCl}_2\text{-BP}$ fade away when adding more PdCl_2 quantum dots, again indicating the block effect of PdCl_2 quantum dots on electron transporting. Compared to $a\text{-}n\text{PdCl}_2\text{-BP}$,

BP, the block phenomenon in $z\text{-}n\text{PdCl}_2\text{-BP}$ is more obvious, correlating well with the higher conductivity for the armchair direction as given by the I - V curves. We also calculated the real-



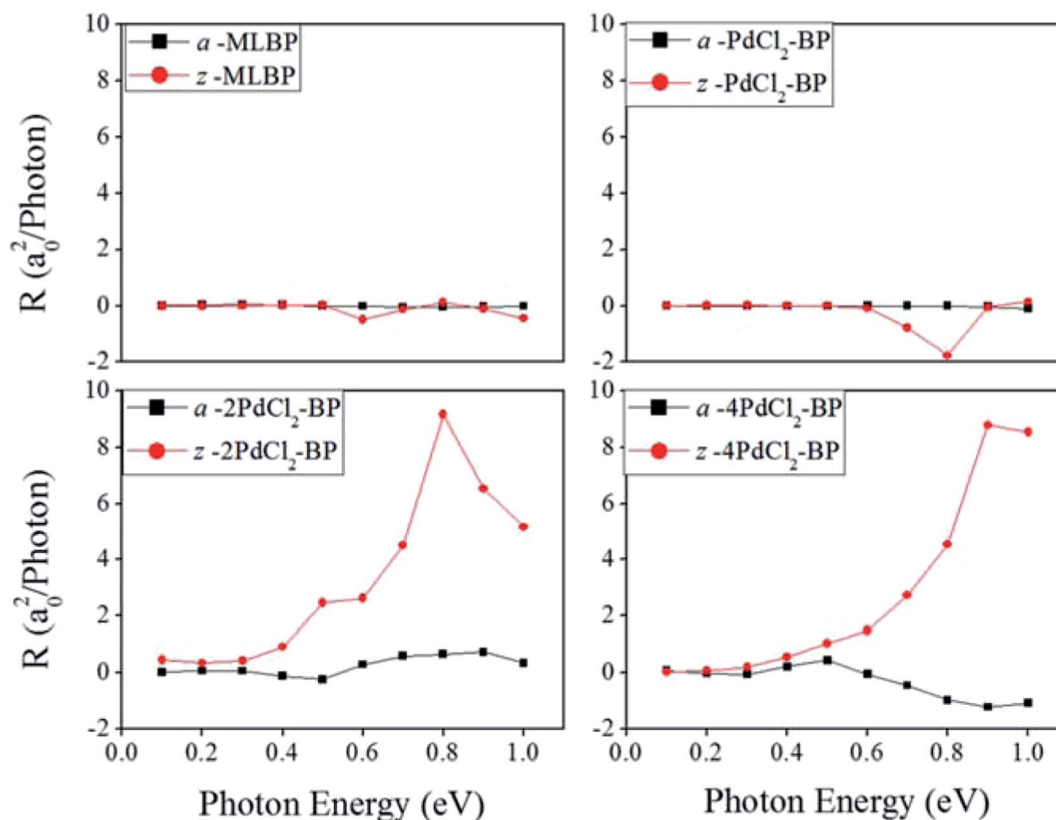


Fig. 11 The calculated maximum photoresponse R for a-MLBP, z-MLBP, a- n PdCl₂-BP, and z- n PdCl₂-BP ($n = 1, 2$, and 4) two-probe devices irradiated by linearly polarized light.

space scattering states of the systems at the E_f for investigating the transport paths. Fourteen transport channels are obtained (cf. Fig. S1–S8 in ESI†), indicating that there exist fourteen subbands in the left electrode along the transport direction. Fig. 9 gives the most effective channels for MLBP and n PdCl₂-BP ($n = 1, 2$, and 4). It can be seen that the transport channels tend to be closed if attaching more PdCl₂ quantum dots due to their scattering effect, also intuitively explaining the reducing of the conductivity. Moreover, the penetrating channel of z- n PdCl₂-BP is weaker at the right side than that of a- n PdCl₂-BP, giving the evidence that charge carriers transport easily along the armchair direction related to the zigzag direction.

3.4 Optoelectronic property

Fig. 2 schematically shows the structures of our devices of n PdCl₂-BP ($n = 1, 2$, and 4) for calculating optoelectronic properties along the armchair and zigzag directions. A linear polarized light was irradiated on the whole scattering region. The applied light has a polarization forming an angle θ with respect to the transport direction of the MLBP plane. The photoresponses were determined at different θ for photon energies ranging from 0 to 1.0 eV with an interval of 0.1 eV. The photocurrent $I^{(\text{ph})}$ can be obtained from the following formula (2):⁶⁹

$$I^{(\text{ph})} = \frac{ie}{h} \int \text{Tr} \{ \Gamma [G^{<(\text{ph})} + f(E)(G^{>(\text{ph})} - G^{<(\text{ph})})] \} dE \quad (2)$$

The photocurrent of eqn (2) can be separated into three terms:⁷⁰

$$I^{(\text{ph})} = \frac{ie}{h} \int \left\{ \cos^2 \theta \text{Tr} \left\{ \Gamma \left[G_1^{<(\text{ph})} + f(E)(G_1^{>(\text{ph})} - G_1^{<(\text{ph})}) \right] \right\} \right. \\ \left. + \sin^2 \theta \text{Tr} \left\{ \Gamma \left[G_2^{<(\text{ph})} + f(E)(G_2^{>(\text{ph})} - G_2^{<(\text{ph})}) \right] \right\} \right. \\ \left. + \sin(2\theta) 2 \text{Tr} \left\{ \Gamma \left[G_3^{<(\text{ph})} + f(E)(G_3^{>(\text{ph})} - G_3^{<(\text{ph})}) \right] \right\} \right\} dE \quad (3)$$

here Γ is the self-energy function, representing the coupling between the central scattering region and the left/right electrode; $G^{<(\text{ph})}$ is the lesser Green's function; $G^{>(\text{ph})}$ is the greater Green's function; $f(E)$ is the Fermi–Dirac distribution function of the left and right electrode. It is easy to check that $I^{(\text{ph})}$ is proportional to the photon flux. Accordingly, the photoresponse function R can be define as

$$R = \frac{I^{(\text{ph})}}{eF} \quad (4)$$

versus photon energy with different polarization direction of the light. Here F is the photon flux.

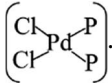
Photoresponses along both armchair and zigzag directions of MLBP and n PdCl₂-BP ($n = 1, 2$, and 4) at photo energies of 0.7, 0.8, and 1.0 eV are given in Fig. 10. According to formula (3), the photoresponses correlate directly to the $\sin^2 \theta$, $\cos^2 \theta$, and $\sin 2\theta$ components. Therefore, the photoresponse curves in Fig. 10 take the sine or cosine shapes with θ , which are



decided by complicated factors such as the geometric symmetry and band structures of the material. The sign of the photoresponse varies with θ and applied photon energy. This is due to the fact that the sign of the photoresponse is determined by the summation of all activated electrons with different velocity distributions. It is noteworthy that bonding PdCl_2 quantum dots is much beneficial for enhancing photoresponses. Regardless of the direction, the maximum photoresponses within the considered photon energies are ordered as $4\text{PdCl}_2\text{-BP} > 2\text{PdCl}_2\text{-BP} > 1\text{PdCl}_2\text{-BP} > \text{MLBP}$. Values for $a\text{-}4\text{PdCl}_2\text{-BP}$ and $z\text{-}4\text{PdCl}_2\text{-BP}$ are up to 1.23 and $8.53 \text{ a}_0^2/\text{photon}$, respectively, 22 and 15 times larger than that of $a\text{-MLBP}$ ($0.0546 \text{ a}_0^2/\text{photon}$) and $z\text{-MLBP}$ ($0.579 \text{ a}_0^2/\text{photon}$). This is due to the fact that electrons can be excited easily by the light from the PdCl_2 quantum dots which can be demonstrated from the band structures aforementioned. Fig. 11 shows the variation of maximum photoresponses with photo energies. Evidently, the $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) nanostructures have produced photoresponse anisotropy: zigzag direction is about one order of magnitude larger than armchair. All of these fascinating photoresponse properties make these new 2D materials especially attractive for optoelectronic devices.

4. Summary

Electronic, transport, and optoelectronic properties of MLBP and $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4) are examined by using DFT and NEGF as well as the KNEGF methods. It is found that the PdCl_2 quantum dots prefer to locate over the furrow site of MLBP and

form a planar quadridentate structure of . A typical

Pd-P coordinate bond is formed. The PdCl_2 quantum dots introduce discrete flat levels just above the MLBP valence band and below the E_F . The conductivity is much lowered by the PdCl_2 grafting, due to the fact that the scatter effect of PdCl_2 quantum dots play a major role in the process of electron transporting. For both parent MLBP and functioned $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4), the armchair direction shows higher conductivities than the zigzag. A threshold voltage is found for $4\text{PdCl}_2\text{-BP}$, a valuable clue for exploring current switches. More importantly, functionalization of PdCl_2 quantum dots is much beneficial for enhancing photoresponse. Values of photoresponse for $a\text{-}4\text{PdCl}_2\text{-BP}$ and $z\text{-}4\text{PdCl}_2\text{-BP}$ are 22 and 15 times larger than that of $a\text{-MLBP}$ and $z\text{-MLBP}$, respectively. A significant photoresponse anisotropy is found for both MLBP and $n\text{PdCl}_2\text{-BP}$ ($n = 1, 2$, and 4), contrary to the conductivity, the zigzag direction shows much stronger photoresponse than the armchair. All of the aforementioned unique properties make this new 2D MLBP based materials especially attractive for both electronic and optoelectronic devices.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by National Natural Science Foundation of China (Grant No. 51473042), Nature Science Foundation of Heilongjiang Province of China (Grant No. B2018007), Harbin foundation for leaders of disciplines (Grant No. 2017RAXXJ002).

References

- 1 A. K. Geim and K. S. Novoselov, *Nat. Mater.*, 2007, **6**, 183–191.
- 2 G. Fiori, F. Bonaccorso, G. Iannaccone, T. Palacios, D. Neumaier, A. Seabaugh, S. K. Banerjee and L. Colombo, *Nat. Nanotechnol.*, 2014, **7**, 768–779.
- 3 M. D. Stoller, S. Park, Y. Zhu, J. An and R. S. Ruoff, *Nano Lett.*, 2008, **8**, 3498–3502.
- 4 D. Voiry, J. Yang and M. Chhowalla, *Adv. Mater.*, 2016, **28**, 6197–6206.
- 5 A. H. Castro Neto, F. Guinea, N. M. R. Peres, K. S. Novoselov and A. K. Geim, *Rev. Mod. Phys.*, 2009, **81**, 109–162.
- 6 B. Radisavljevic, A. Radenovic, J. Brivio, V. Giacometti and A. Kis, *Nat. Nanotechnol.*, 2011, **6**, 147–150.
- 7 Y. Jing, X. Zhang and Z. Zhou, *WIREs Computational Molecular Science*, 2016, **6**, 5–19.
- 8 P. Yasaei, B. Kumar, T. Foroozan, C. Wang, M. Asadi, D. Tuschel, J. E. Indacochea, R. F. Klie and A. Salehi-Khojin, *Adv. Mater.*, 2015, **27**, 1887–1892.
- 9 J. Kang, J. D. Wood, S. A. Wells, J. H. Lee, X. Liu, K. S. Chen and M. C. Hersam, *ACS Nano*, 2015, **9**, 3596–3604.
- 10 Y. Zhang, N. Dong, H. Tao, C. Yan, J. Huang, T. Liu, A. W. Robertson, J. Texter, J. Wang and Z. Sun, *Chem. Mater.*, 2017, **29**, 6445–6456.
- 11 W. Zhao, Z. Xue, J. Wang, X. Zhao and T. Mu, *ACS Appl. Mater. Interfaces*, 2015, **7**, 27608–27612.
- 12 A. H. Woomer, T. W. Farnsworth, J. Hu, R. A. Wells, C. L. Donley and S. C. Warren, *ACS Nano*, 2015, **9**, 8869–8884.
- 13 W. Lu, H. Nan, J. Hong, Y. Chen, C. Zhu, Z. Liang, X. Ma, Z. Ni, C. Jin and Z. Zhang, *Nano Res.*, 2014, **7**, 853–859.
- 14 H. Kwon, S. W. Seo, T. G. Kim, E. S. Lee, P. T. Lanh, S. Yang, S. Ryu and J. W. Kim, *ACS Nano*, 2016, **10**, 8723–8731.
- 15 M. Fortin-Deschênes, P. L. Levesque, R. Martel and O. Moutanabbir, *J. Phys. Chem. Lett.*, 2016, **7**, 1667–1674.
- 16 Y. Chen, G. Jiang, S. Chen, Z. Guo, X. Yu, C. Zhao, H. Zhang, Q. Bao, S. Wen, D. Tang and D. Fan, *Opt. Express*, 2015, **23**, 12823–12833.
- 17 L. Li, Y. Yu, G. J. Ye, Q. Ge, X. Ou, H. Wu, D. Feng, X. H. Chen and Y. Zhang, *Nat. Nanotechnol.*, 2014, **9**, 372–377.
- 18 S. Das, W. Zhang, M. Demarteau, A. Hoffmann, M. Dubey and A. Roelofs, *Nano Lett.*, 2014, **14**, 5733–5739.
- 19 J. Qiao, X. Kong, Z. X. Hu, F. Yang and W. Ji, *Nat. Commun.*, 2014, **5**, 4475.
- 20 F. Xia, H. Wang and Y. Jia, *Nat. Commun.*, 2014, **5**, 4458.
- 21 A. Castellanos-Gomez, *J. Phys. Chem. Lett.*, 2015, **6**, 4280–4291.
- 22 J. S. Kim, Y. Liu, W. Zhu, S. Kim, D. Wu, L. Tao, A. Dodabalapur, K. Lai and D. Akinwande, *Sci. Rep.*, 2015, **5**, 8989.



- 23 J. Pei, G. Xin, J. Yang, X. Wang, Z. Yu, D. Y. Choi, L. D. Barry and Y. Lu, *Nat. Commun.*, 2016, **7**, 10450.
- 24 D. Yuchen, L. Han, D. Yexin and D. Y. Peide, *ACS Nano*, 2014, **8**, 10035–10042.
- 25 L. Li, M. Engel, D. B. Farmer, S. J. Han and H. S. P. Wong, *ACS Nano*, 2016, **10**, 4672–4677.
- 26 K. Gong, L. Zhang, W. Ji and H. Guo, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **90**, 125441.
- 27 Y. Ma, C. Shen, A. Zhang, L. Chen, Y. Liu, J. Chen, Q. Liu, Z. Li, M. R. Amer, T. Nilges, A. N. Abbas and C. Zhou, *ACS Nano*, 2017, **11**, 7126–7133.
- 28 C. Han, Z. Hu, L. C. Gomes, Y. Bao, A. Carvalho, S. J. R. Tan, B. Lei, D. Xiang, J. Wu, D. Qi, L. Wang, F. Huo, W. Huang, K. P. Loh and W. Chen, *Nano Lett.*, 2017, **17**, 4122–4129.
- 29 Q. F. Li, X. G. Wan, C. G. Duan and J. L. Kuo, *J. Phys. D: Appl. Phys.*, 2014, **47**, 465302.
- 30 X. Y. Tan, J. H. Wang, Y. Y. Zhu, Y. Zuo and X. Jin, *Acta Phys. Sin.*, 2014, **63**, 207301.
- 31 S. Zhao, W. Kang and J. Xue, *J. Mater. Chem. A*, 2014, **2**, 19046–19052.
- 32 G. L. Xu, Z. Chen, G. M. Zhong, Y. Liu, Y. Yang, T. Ma, Y. Ren, X. Zuo, X. H. Wu, X. Zhang and K. Amine, *Nano Lett.*, 2016, **16**, 3955–3965.
- 33 G. C. Constantinescu and N. D. Hine, *Nano Lett.*, 2016, **16**, 2586–2594.
- 34 R. A. Doganov, S. P. Koenig, Y. Yeo, K. Watanabe, T. Taniguchi and B. Özyilmaz, *Appl. Phys. Lett.*, 2015, **106**, 083505.
- 35 P. Zhang, J. Wang and X. M. Duan, *Chin. Phys. B*, 2016, **25**, 037302.
- 36 Y. Deng, Z. Luo, N. J. Conrad, H. Liu, Y. Gong, S. Najmaei, P. M. Ajayan, J. Lou, X. Xu and P. D. Ye, *ACS Nano*, 2014, **8**, 8292–8299.
- 37 L. Ye, H. Li, Z. Chen and J. Xu, *ACS Photonics*, 2016, **3**, 692–699.
- 38 V. Kumar, J. R. Brent, M. Shorie, H. Kaur, G. Chadha, A. G. Thomas, E. A. Lewis, A. P. Rooney, L. Nguyen, X. Zhong, M. G. Burke, S. J. Haigh, A. Walton, P. D. McNaughton, A. A. Tedstone, N. Savjani, C. A. Muryn, P. O'Brien, A. K. Ganguli, D. J. Lewis and P. Sabherwal, *ACS Appl. Mater. Interfaces*, 2016, **8**, 22860–22868.
- 39 Q. Li, Q. Zhou, X. Niu, Y. Zhao, Q. Chen and J. Wang, *J. Phys. Chem. Lett.*, 2016, **7**, 4540–4546.
- 40 S. P. Koenig, R. A. Doganov, L. Seixas, A. Carvalho, J. Y. Tan, K. Watanabe, T. Taniguchi, N. Yakovlev, A. H. Castro Neto and B. Özyilmaz, *Nano Lett.*, 2016, **16**, 2145–2551.
- 41 V. V. Kulish, O. I. Mal'yi, C. Persson and P. Wu, *Phys. Chem. Chem. Phys.*, 2015, **17**, 992–1000.
- 42 T. Hu and J. Hong, *J. Phys. Chem. C*, 2015, **119**, 8199–8207.
- 43 N. Suvansinpan, F. Hussain, G. Zhang, C. H. Chiu, Y. Cai and Y. W. Zhang, *Nanotechnology*, 2016, **27**, 065708.
- 44 W. Lei, T. Zhang, P. Liu, A. R. Jose, G. Liu and M. Liu, *ACS Catal.*, 2016, **6**, 8009–8020.
- 45 Y. Zhao, H. Wang, H. Huang, Q. Xiao, Y. Xu, Z. Guo, H. Xie, J. Shao, Z. Sun, W. Han, X. F. Yu, P. Li and P. K. Chu, *Angew. Chem., Int. Ed.*, 2016, **55**, 5003–5007.
- 46 E. Alonso, J. Forniés, C. Fortuño, A. Lledós, A. Martín and A. Nova, *Inorg. Chem.*, 2009, **48**, 7679–7690.
- 47 V. I. Bakmutov, F. Bozoglian, K. Gómez, G. González, V. V. Grushin, S. A. Macgregor, E. Martin, F. M. Miloserdov, M. A. Novikov, J. A. Panetier and L. V. Romashov, *Organometallics*, 2011, **31**, 1315–1328.
- 48 D. E. Berry, P. Carrie, K. L. Fawkes, B. Rebner and Y. Xing, *J. Chem. Educ.*, 2010, **87**, 533–534.
- 49 L. R. Moore, E. C. Western, R. Craciun, J. M. Spruell and D. A. Dixon, *Organometallics*, 2008, **27**, 576–593.
- 50 K. Herbst, P. Zanello, M. Corsini, N. D. Amelio, L. Dahleburg and M. Brorson, *Inorg. Chem.*, 2003, **42**, 974–981.
- 51 J. C. W. Folmer, J. A. Turner and B. A. Parkinson, *J. Solid State Chem.*, 1987, **68**, 28–37.
- 52 V. Tran, R. Soklaski, Y. Liang and L. Yang, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **89**, 235319.
- 53 Y. Du, J. Maassen, W. Wu, Z. Luo, X. Xu and P. D. Ye, *Nano Lett.*, 2016, **16**, 6701–6708.
- 54 R. Fei and L. Yang, *Nano Lett.*, 2014, **14**, 2884–2889.
- 55 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 56 M. Brandbyge, J. L. Mozos, P. Ordejon, J. Taylor and K. Stokbro, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2002, **65**, 165401.
- 57 J. M. Soler, E. Artacho, J. D. Gale, A. Garcia, J. Junquera, P. Ordejon and D. Sanchez-Portal, *J. Phys.: Condens. Matter*, 2002, **14**, 2745–2779.
- 58 ATK, Version 13.8, atomistix a/s, 2013, www.quantumwise.com.
- 59 S. Å. Gustafson and G. Dahlquist, *Chem. Phys. Lett.*, 2011, **508**, 215–218.
- 60 J. Taylor, H. Guo and J. Wang, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2001, **63**, 245407.
- 61 D. Waldron, P. Haney, B. Larade, A. MacDonald and H. Guo, *Phys. Rev. Lett.*, 2006, **96**, 166804.
- 62 T. Takahashi, H. Tokailin, S. Suzuki and I. Shirotni, *J. Phys. C: Solid State Phys.*, 1985, **18**, 825–836.
- 63 M. A. Carvajal, G. P. Miscione, J. J. Novoa and A. Bottoni, *Organometallics*, 2005, **24**, 2086–2096.
- 64 H. Liu, A. T. Neal, Z. Zhu, Z. Luo, X. Xu, D. Tomanek and P. D. Ye, *ACS Nano*, 2014, **8**, 4033–4041.
- 65 S. Das, W. Zhang, M. Demarteau, A. Hoffmann, M. Dubey and A. Roelofs, *Nano Lett.*, 2014, **14**, 5733–5739.
- 66 Y. Du, C. Ouyang, S. Shi and M. Lei, *J. Appl. Phys.*, 2010, **107**, 093718.
- 67 T. H. Oosterkamp, T. Fujisawa, W. G. V. D. Wiel, K. Ishibashi, R. V. Hijman, S. Tarucha and L. P. Kouwenhoven, *Nature*, 1998, **395**, 873–876.
- 68 B. G. Streetman and S. K. Banerjee, *Solid State Electronic Devices, Sixth Edition*, Prentice Hall, 2006.
- 69 L. Zhang, K. Gong, J. Chen, L. Liu, D. Xiao and H. Guo, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **90**, 195428.
- 70 Y. Xie, L. Zhang, Y. Zhu, L. Liu and H. Guo, *Nanotechnology*, 2015, **26**, 455202.

