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# Preparation, characterization, and luminescence properties of double perovskite $\text{SrLaMgSbO}_6\text{:Mn}^{4+}$ far-red emitting phosphors for indoor plant growth lighting

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$\text{Mn}^{4+}$ -activated  $\text{SrLaMgSbO}_6$  far-red emitting phosphors with double perovskite structure were prepared by traditional solid-state reaction. The research on the crystal structure of the  $\text{SrLaMgSbO}_6\text{:0.8\%Mn}^{4+}$  (SLMS:0.8%Mn<sup>4+</sup>) phosphors showed that the as-prepared sample was made up of two polyhedrons,  $[\text{SbO}_6]$  and  $[\text{MgO}_6]$ . Under the excitation of 333 nm, the SLMS:0.8%Mn<sup>4+</sup> phosphors exhibited an intense far-red emission in the 625–800 nm wavelength range with CIE chromaticity coordinates of (0.733, 0.268), which could match well with the absorption spectrum of phytochrome  $\text{P}_{\text{FR}}$ . The optimal concentration of  $\text{Mn}^{4+}$  ions in the SLMS:Mn<sup>4+</sup> phosphors was 0.8 mol%. Importantly, the as-prepared SLMS:0.8%Mn<sup>4+</sup> phosphors had an internal quantum efficiency of 35%. The thermal stability of SLMS:0.8%Mn<sup>4+</sup> phosphors was also investigated, and the activation energy was found to be 0.3 eV. Thus, the  $\text{Mn}^{4+}$ -activated SLMS phosphors have great potential to serve as far-red emitting phosphors in indoor plant growth lighting.

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## Introduction

Nowadays, an intense research effort has recently been devoted to developing red-emitting phosphors for phosphor-converted light-emitting diodes (pc-LEDs), due to the fact that pc-LEDs, which show many merits including high durability, low energy consumption, short response time, and environmental protection, can be used in indoor plant cultivation, decoration, and medical applications.<sup>1–9</sup> For the indoor plant growth, the growth status can be controlled by adjusting the ratio of the phytochrome  $\text{P}_{\text{R}}$  and  $\text{P}_{\text{FR}}$ , because the  $\text{P}_{\text{FR}}$  can convert to  $\text{P}_{\text{R}}$  by absorbing far-red light centered at 730 nm.<sup>10,11</sup> But in the agricultural industry, especially indoor plant cultivation, there is lack of far-red light to meet the requirement of mutual transformation between the phytochrome  $\text{P}_{\text{FR}}$  and  $\text{P}_{\text{R}}$  due to small percentage of far-red light in the sunlight compared to red light.<sup>10,12</sup> Meanwhile, the traditional lamp cannot match well with the absorption spectrum of the phytochrome  $\text{P}_{\text{FR}}$ . So pc-LEDs play an important role in indoor plant growth, and many research on the phosphors used in indoor plant growth lighting have been implemented such as  $\text{Ca}_{1-x}\text{Y}_x\text{Al}_{12-x}\text{Mg}_x\text{O}_{19}\text{:Mn}^{4+}$ ,<sup>13</sup>  $\text{La}(\text{MgTi})_{1/2}\text{O}_3\text{:Mn}^{4+}$ ,<sup>10</sup>  $\text{Ca}_{14}\text{Al}_{10}\text{Zn}_6\text{O}_{35}\text{:Bi}^{3+},\text{Mn}^{4+}$ ,<sup>14</sup>  $\text{SrMg}_2\text{-La}_2\text{W}_2\text{O}_{12}\text{:Mn}^{4+}$ ,<sup>15</sup> and  $\text{Ca}_3\text{Al}_4\text{ZnO}_{10}\text{:Bi}^{3+},\text{Mn}^{4+}$ .<sup>16</sup> Since the

far-red emitting phosphors are vital to the pc-LEDs and relatively little research has been conducted on that kind of phosphors, thus it is meaningful to study the far-red emitting phosphors for applications in indoor plant cultivation.

It is known that  $\text{Mn}^{4+}$  ions with a  $3d^3$  electronic configuration in a strong crystal fields of octahedral coordination (coordination number (CN) = 6) can give rise to red emission in the wavelength region of 620–800 nm owing to the spin-forbidden  $^2\text{E}_g \rightarrow ^4\text{A}_{2g}$  transition.<sup>17–24</sup>  $\text{Mn}^{4+}$ -doped phosphors contain two categories: (1)  $\text{Mn}^{4+}$ -activated fluorides, such as  $\text{Li}_3\text{Na}_3\text{Ga}_2\text{F}_{12}\text{:Mn}^{4+}$  ( $\lambda_{\text{em}} = 627$  nm),<sup>25</sup>  $\text{Na}_3\text{TaF}_8\text{:Mn}^{4+}$  ( $\lambda_{\text{em}} = 627$  nm),<sup>26</sup> and  $\text{K}_3\text{ScF}_6\text{:Mn}^{4+}$  ( $\lambda_{\text{em}} = 631$  nm);<sup>27</sup> and (2)  $\text{Mn}^{4+}$ -activated oxides, such as  $\text{Ca}_3\text{La}_2\text{W}_2\text{O}_{12}\text{:Mn}^{4+}$  ( $\lambda_{\text{em}} = 711$  nm),<sup>28</sup>  $\text{Li}_2\text{MgZrO}_4\text{:Mn}^{4+}$  ( $\lambda_{\text{em}} = 670$  nm),<sup>29</sup> and  $\text{Ba}_2\text{GdSbO}_6\text{:Mn}^{4+}$  ( $\lambda_{\text{em}} = 687$  nm).<sup>30</sup>  $\text{Mn}^{4+}$ -activated fluorides generally exhibit red emissions in the 600–650 nm wavelength range and they are widely used in white LEDs application.<sup>13,30–33</sup> In sharp contrast,  $\text{Mn}^{4+}$ -activated oxides showed far-red emissions at around 700 nm, which are vital to plant growth in photomorphogenesis process. So  $\text{Mn}^{4+}$ -activated oxides have potential applications as spectral converters for pc-LEDs towards indoor plant growth lighting.<sup>14,34,35</sup>

In recent years, double-perovskite oxides with octahedral crystal sites are considered as good host materials for  $\text{Mn}^{4+}$  activator.<sup>35–37</sup> In a previous paper by Liu and co-workers, the  $\text{Sr}_{1-x}\text{Ca}_x\text{LaMgSbO}_6\text{:Eu}^{3+}$  phosphors have been researched, and the structural analysis shows that the host compound has two kind of octahedral structures,  $[\text{MgO}_6]$  and  $[\text{SbO}_6]$ .<sup>38</sup> Thus, in this

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present paper,  $\text{Mn}^{4+}$ -doped  $\text{SrLaMgSbO}_6$  (SLMS) phosphors have been synthesized and their crystal structure and optical properties were studied. The SLMS host contains two octahedral structures of  $[\text{MgO}_6]$  and  $[\text{SbO}_6]$ , and the  $\text{Mn}^{4+}$  ions prefer to occupy the  $\text{Sb}^{5+}$  sites to form the phase of  $\text{SLMS:Mn}^{4+}$ . Under the excitation of 333 nm, the  $\text{SLMS:Mn}^{4+}$  phosphors emitted far-red light peaking at 705 nm in the spectral range from 625 to 800 nm, which was attributed to the  ${}^2\text{E}_g \rightarrow {}^4\text{A}_{2g}$  transition of  $\text{Mn}^{4+}$  ions. More importantly, the emission spectrum of the  $\text{SLMS:Mn}^{4+}$  phosphors could match well with the absorption curve of phytochrome  $\text{P}_{\text{FR}}$ , indicating the far-red light from the  $\text{SLMS:Mn}^{4+}$  phosphors could be efficiently absorbed by phytochrome  $\text{P}_{\text{FR}}$ . The CIE chromaticity coordinates, internal quantum efficiency (IQE), and thermal stability of  $\text{SLMS:Mn}^{4+}$  phosphors were also investigated in detail. The results demonstrated that the  $\text{Mn}^{4+}$ -activated SLMS phosphors have much potential to serve as far-red emitting phosphors for indoor plant growth lighting.

## Experimental section

$\text{SrLaMgSb}_{1-x}\text{O}_6:\text{xMn}^{4+}$  (abbreviated as:  $\text{SLMS:xMn}^{4+}$ ;  $x = 0.05$ – $1.2\%$ ) phosphors were synthesized by a traditional solid-state reaction method. The starting materials of  $\text{SrCO}_3$  (analytical reagent, AR),  $\text{La}_2\text{O}_3$  (99.99%),  $\text{MgO}$  (AR),  $\text{Sb}_2\text{O}_5$  (99%), and  $\text{MnCO}_3$  (AR), were weighed according to the stoichiometric ratio. Using an agate mortar, the mixed starting materials were ground and then transferred to  $\text{Al}_2\text{O}_3$  crucible to preheat at  $500^\circ\text{C}$  for 3 h in air. After cooling to room temperature, the preheated mixtures were reground again into powders for further calcination at  $1500^\circ\text{C}$  for 6 h in air to get the target  $\text{SLMS:xMn}^{4+}$  phosphors.

The X-ray diffraction (XRD) patterns and morphology properties of  $\text{SLMS:xMn}^{4+}$  phosphors were measured on a X-ray diffractometer (Cu-K $\alpha$  radiation; Bruker D8 Advance) and a field-emission scanning electron microscope (FE-SEM; MAIA3 TES-CAN), respectively. Structure refinements were made using Fullprof program. The room-temperature photoluminescence (PL) spectra and the PL excitation (PLE) spectra were measured by an Edinburgh FS5 spectrofluorometer with a 150 W continued-wavelength xenon lamp. The luminescence decay lifetimes, the temperature-dependent spectra, and IQE were recorded by using the same spectrofluorometer, equipping with a pulsed xenon lamp, a temperature controller, and an integrating sphere, respectively. The prototype far-red light-emitting device was fabricated by coating the  $\text{SLMS:0.8\%Mn}^{4+}$  phosphors onto a 365 nm near-UV InGaN LED chip, and the corresponding electroluminescent (EL) properties were measured by using a spectroradiometer system (HAAS-2000, Everfine).

## Results and discussion

Fig. 1 shows the XRD patterns of the  $\text{SLMS:xMn}^{4+}$  ( $x = 0.05\%$ ,  $0.8\%$ , and  $1.2\%$ ) phosphors together with the standard card of the  $\text{CaLaMgMoO}_6$  (JCPDS #43-0043). Compared to the standard card of the  $\text{CaLaMgMoO}_6$  compound, no other crystalline phase

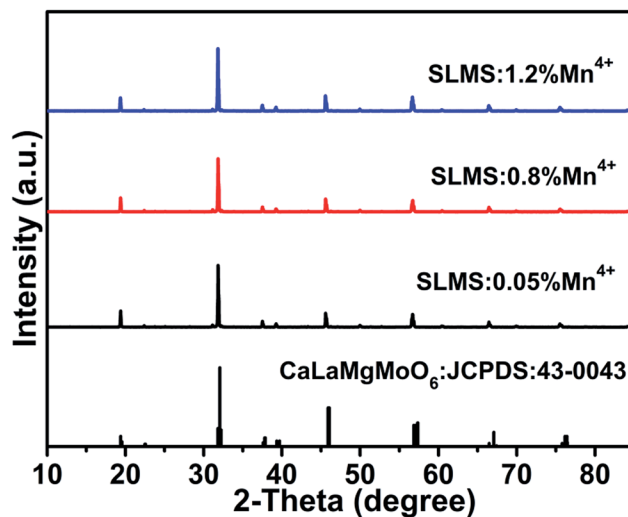


Fig. 1 XRD patterns of the  $\text{SLMS:xMn}^{4+}$  ( $x = 0.05\%$ ,  $0.8\%$ , and  $1.2\%$ ) phosphors together with the stand card of  $\text{CaLaMgMoO}_6$  (JCPDS #43-0043).

was formed in the  $\text{SLMS:xMn}^{4+}$  phosphors. Fig. 2(a) and Fig. 2(b) represent the refinement result and the corresponding crystal structure of the  $\text{SLMS:0.8\%Mn}^{4+}$  phosphors, respectively. Table 1 gives the main parameters of the crystal cell for  $\text{SLMS:0.8\%Mn}^{4+}$  phosphors, in which  $a = 5.63586(16) \text{ \AA}$ ,  $b = 5.62543(13) \text{ \AA}$ ,  $c = 7.95407(19) \text{ \AA}$ , and  $V = 252.177(6) \text{ \AA}^3$ . The  $\text{SLMS:0.8\%Mn}^{4+}$  phosphors with double perovskite structure was belong to monoclinic crystal system and a  $P2_1/n$  space group. The host was composed by two octahedral structures of  $[\text{MgO}_6]$  and  $[\text{SbO}_6]$ . Considering the ion radii and CN of the  $\text{Sb}^{5+}$  ions ( $r = 0.6 \text{ \AA}$ , CN = 6),  $\text{Mg}^{2+}$  ions ( $r = 0.72 \text{ \AA}$ , CN = 6), and  $\text{Mn}^{4+}$  ions ( $r = 0.53 \text{ \AA}$ , CN = 6), the  $\text{Mn}^{4+}$  ions were more inclined to occupy the  $\text{Sb}^{5+}$  sites in SLMS host to form the  $\text{SLMS:Mn}^{4+}$  compound.<sup>39–41</sup> Another parameter to determine which ions were substituted by  $\text{Mn}^{4+}$  ions was the radii percentage difference ( $D_r$ ) between the doping ions (here was  $\text{Mn}^{4+}$  ions) and the possible substituted ions (here were  $\text{Sb}^{5+}$  and  $\text{Mg}^{2+}$  ions). The ratio can be estimated by the following expression:<sup>42,43</sup>

$$D_r = \frac{R_m(\text{CN}) - R_d(\text{CN})}{R_m(\text{CN})} \times 100\% \quad (1)$$

where  $R_m(\text{CN})$  and  $R_d(\text{CN})$  are the ionic radii of the host cations ( $\text{Mg}^{2+}$  and  $\text{Sb}^{5+}$  ions) and the doped  $\text{Mn}^{4+}$  ions, respectively. Herein,  $R_d(\text{CN}) = 0.53 \text{ \AA}$ ; the value of  $R_m(\text{CN})$  is  $0.72 \text{ \AA}$  and  $0.6 \text{ \AA}$  for  $\text{Mg}^{2+}$  and  $\text{Sb}^{5+}$  ions, respectively. In this host, CN = 6 for  $\text{Mn}^{4+}$ ,  $\text{Mg}^{2+}$ , and  $\text{Sb}^{5+}$  ions. Thus, the values of  $D_r$  were 26% and 12% for  $\text{Mg}^{2+}$  host cation and  $\text{Sb}^{5+}$  host cation, respectively. Because when  $\text{Mn}^{4+}$  ions substituted for  $\text{Mg}^{2+}$  ions, the  $D_r$  ratio was higher than  $\text{Mn}^{4+}$  ions substituted for  $\text{Sb}^{5+}$  ions, so the  $\text{Mn}^{4+}$  ions were prefer to substitute for  $\text{Sb}^{5+}$  than  $\text{Mg}^{2+}$  ions. On the other hand, the valence of the  $\text{Sb}^{5+}$  ions was closer to  $\text{Mn}^{4+}$  ions than  $\text{Mg}^{2+}$  ions. Therefore, considering the radii and the valence, the  $\text{Mn}^{4+}$  ions would substitute for  $\text{Sb}^{5+}$  sites to form the  $\text{SLMS:Mn}^{4+}$  phosphors. Fig. 3 shows the typical FE-SEM



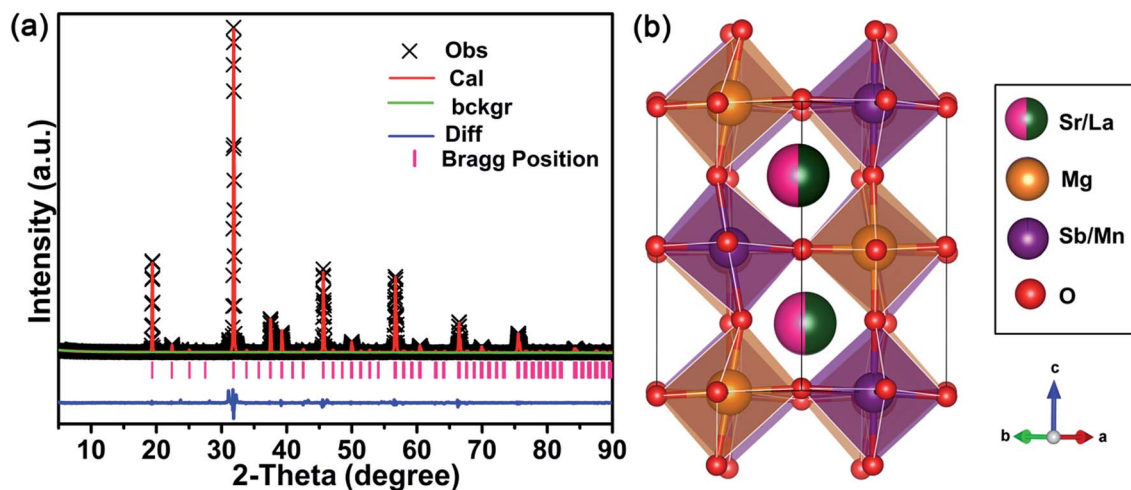


Fig. 2 (a) Rietveld refinement of the XRD pattern for the SLMS:0.8%Mn<sup>4+</sup> phosphors. (b) The crystal structure of the SLMS:0.8%Mn<sup>4+</sup> phosphors.

Table 1 Refined crystallographic parameters of SLMS:0.8%Mn<sup>4+</sup> phosphors

Formula	SLMS:0.8%Mn <sup>4+</sup>
Crystal system	Monoclinic
Space group	<i>P</i> 2 <sub>1</sub> / <i>n</i>
<i>a</i>	5.63586(16) Å
<i>b</i>	5.62543(13) Å
<i>c</i>	7.95407(19) Å
$\alpha$	90°
$\beta$	90.054(4)°
$\gamma$	90°
<i>V</i>	252.177(6) Å <sup>3</sup>
<i>R<sub>p</sub></i>	10.53%
<i>R<sub>wp</sub></i>	13.86%

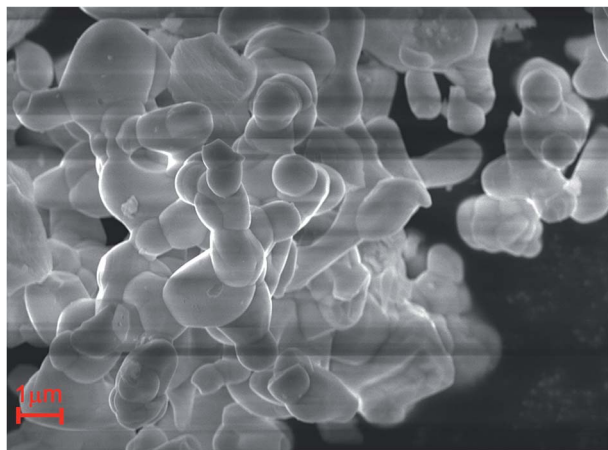


Fig. 3 The typical FE-SEM of SLMS:0.8%Mn<sup>4+</sup> phosphors.

image of the SLMS:0.8%Mn<sup>4+</sup> phosphors. The sizes of the as-prepared sample was found to be about 1–6 μm, and the crystal-surfaces was smooth which was beneficial to get uniform luminescence intensity.

The PLE spectrum ( $\lambda_{em} = 705$  nm) of the SLMS:0.8%Mn<sup>4+</sup> phosphors with four Gaussian fitting peaks was shown in Fig. 4(a). Four PLE peaks at 298 nm ( $33\,557\text{ cm}^{-1}$ ; Mn<sup>4+</sup>–O<sup>2–</sup> charge-transfer transition), 338 nm ( $29\,586\text{ cm}^{-1}$ ;  $^4A_{2g} \rightarrow ^4T_{1g}$  transition), 423 nm ( $23\,641\text{ cm}^{-1}$ ;  $^4A_{2g} \rightarrow ^2T_{2g}$  transition), and 495 nm ( $20\,202\text{ cm}^{-1}$ ;  $^4A_{2g} \rightarrow ^4T_{2g}$  transition) were observed, and the PLE peaks at 333 nm and 499 nm dominated the PLE spectrum.<sup>44–46</sup> Under excitations at 333 nm and 499 nm, the SLMS:0.8%Mn<sup>4+</sup> phosphors exhibited a far-red emission band peaking at 705 nm ( $^2E_g \rightarrow ^4A_{2g}$  transition) in the range of 625–800 nm,<sup>47–49</sup> as shown in Fig. 4(b). Fig. 4(c) displays the PL spectra of the SLMS:*x*Mn<sup>4+</sup> (*x* = 0.05%, 0.6%, 0.8%, 1.0%, and 1.2%) phosphors with different Mn<sup>4+</sup> doping concentrations under 333 nm excitation. As the Mn<sup>4+</sup> concentration increased, the PL intensity firstly showed an increasing trend, and then decreased when the concentration was higher than 0.8 mol%, indicating the occurrence of the concentration quenching effect. Fig. 4(d) shows the luminescence decay curves of the SLMS:*x*Mn<sup>4+</sup> (*x* = 0.05%, 0.6%, 0.8%, 1.0%, and 1.2%) phosphors. The decay lifetimes of the as-prepared samples were obtained by using the following double exponential expression:<sup>50–53</sup>

$$I = A_1 \exp(-t/\tau_1) + A_2 \exp(-t/\tau_2) \quad (2)$$

herein, *I* refers to the luminescent emission intensity of SLMS:*x*Mn<sup>4+</sup> phosphors at time *t*.  $\tau_1$  and  $\tau_2$  are the lifetimes for the exponential component; *A*<sub>1</sub> and *A*<sub>2</sub> are constants. Accordingly, the decay lifetimes of the SLMS:*x*Mn<sup>4+</sup> phosphors were found to be 1.257, 1.130, 1.068, 1.034, and 0.950 ms when *x* = 0.05%, 0.6%, 0.8%, 1.0%, and 1.2%, respectively. Obviously, the decay lifetimes decreased with the increasing Mn<sup>4+</sup> concentration, due to the non-radiative energy transfer in the SLMS:Mn<sup>4+</sup> phosphors.<sup>54</sup>

The critical distance (*R*<sub>c</sub>) is usually used to determine which kind of interaction contribute to the concentration quenching



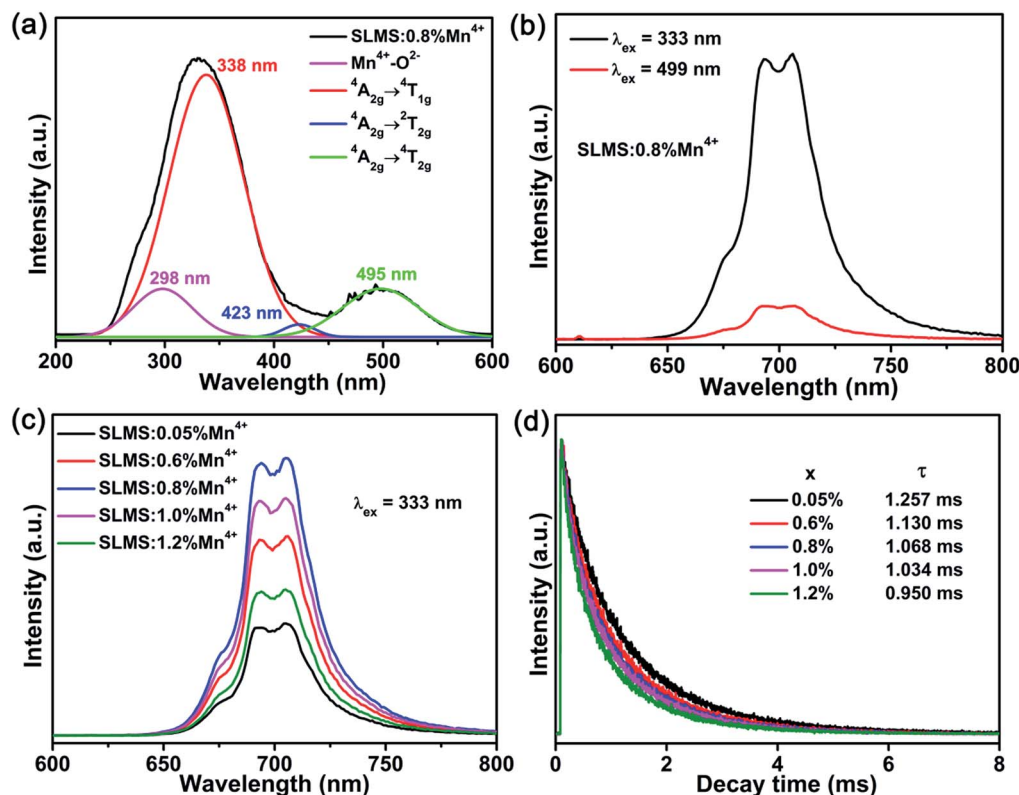


Fig. 4 (a) The PLE spectrum of SLMS:0.8%Mn<sup>4+</sup> phosphors monitored at 705 nm. The four Gaussian fitting curves also displayed in the figure. (b) The PL spectra of SLMS:0.8%Mn<sup>4+</sup> phosphors excited at 333 nm and 499 nm. (c) The PL spectra of SLMS:xMn<sup>4+</sup> (x = 0.05–1.2%) phosphors excited at 333 nm. (d) The luminescence decay curves of SLMS:xMn<sup>4+</sup> (x = 0.05–1.2%) phosphors ( $\lambda_{\text{ex}} = 333$  nm;  $\lambda_{\text{em}} = 705$  nm).

in phosphors. And the  $R_c$  for SLMS:xMn<sup>4+</sup> phosphors can be estimated by the following expression:<sup>55–58</sup>

$$R_c = 2 \left( \frac{3V}{4\pi x_c N} \right)^{1/3} \quad (3)$$

where  $V$  is the volume of the unit cell and  $x_c$  is the optimal doping concentration of Mn<sup>4+</sup> ions in SLMS host;  $N$  is the number of host cations which is available for Mn<sup>4+</sup> ions to be substituted. In SLMS host,  $N = 2$ ,  $V = 252.177(6) \text{ \AA}^3$ , and  $x_c = 0.8\%$ . Thus, the value of  $x_c$  was calculated to be  $31.1 \text{ \AA}$ . The  $R_c$  for SLMS:Mn<sup>4+</sup> phosphors was much larger than  $5 \text{ \AA}$ , and thus the electric multipole interaction was the mechanism responsible for the concentration quenching effect in SLMS:Mn<sup>4+</sup> phosphors.<sup>59</sup>

Fig. 5(a) shows the CIE chromaticity diagram of SLMS:0.8% Mn<sup>4+</sup> phosphors. The corresponding CIE coordinates were determined to be (0.733, 0.268), which were located in the far-red region. Fig. 5(b) compares the PL spectrum of the SLMS:0.8%Mn<sup>4+</sup> phosphors with the absorption spectrum of phytochrome P<sub>FR</sub> in the range of 500–800 nm, in which the phytochrome P<sub>FR</sub> was taken from ref. 10. Apparently, the far-red light emission of the SLMS:0.8%Mn<sup>4+</sup> phosphors was matched well with the absorption spectrum of the phytochrome P<sub>FR</sub>, indicating that the phytochrome P<sub>FR</sub> was sensitive to the far-red light from the SLMS:0.8%Mn<sup>4+</sup> phosphors. Thus, the luminescent properties of the SLMS:0.8%Mn<sup>4+</sup> phosphors

demonstrated that they had bright prospects to be used in far-red LEDs for indoor plant growth.

The IQE of SLMS:0.8%Mn<sup>4+</sup> phosphors was measured to be 35%, as shown in Fig. 6(a). Moreover, the thermal stability of the SLMS:0.8%Mn<sup>4+</sup> phosphors was also investigated. Fig. 6(b) shows the temperature-dependent PL spectra of the SLMS:0.8% Mn<sup>4+</sup> phosphors under 333 nm excitation. The PL intensity was gradually decreased when the temperature increased, and the emission intensity at 363 K remained about 55% of that at 303 K, which could be seen clearly in Fig. 6(b) and (c). Although the thermal stability was not good as the Ca<sub>2</sub>YSbO<sub>6</sub>:Mn<sup>4+</sup> (remained 63% at 373 K),<sup>60</sup> Ca<sub>2</sub>LaNbO<sub>6</sub>:Mn<sup>4+</sup> (remained about 50% at 400 K),<sup>40</sup> and CaLaMgNbO<sub>6</sub> (remained about 45% at 423 K),<sup>41</sup> in the next work, we may improve the thermal behavior and luminescence intensity by co-doping another ions such as Li<sup>+</sup>, Na<sup>+</sup>, Mg<sup>2+</sup>, and Ba<sup>2+</sup>.<sup>49,60–62</sup> Based on the temperature-dependent PL spectra of the SLMS:0.8%Mn<sup>4+</sup> phosphors, the corresponding relationship between  $1/kT$  and  $\ln(I_0/I - 1)$  was shown in Fig. 6(d), and the activation energy can be estimated by using the following expression:<sup>38,63,64</sup>

$$I(T) = I_0 \left[ 1 + c \exp \left( -\frac{\Delta E}{kT} \right) \right]^{-1} \quad (4)$$

where  $I(T)$  and  $I_0$  are the PL intensity at temperature  $T$  and the initial intensity at 303 K, respectively;  $c$  is a constant;  $k$  is the Boltzmann constant. The slope was found to be  $-0.3$ , indicating





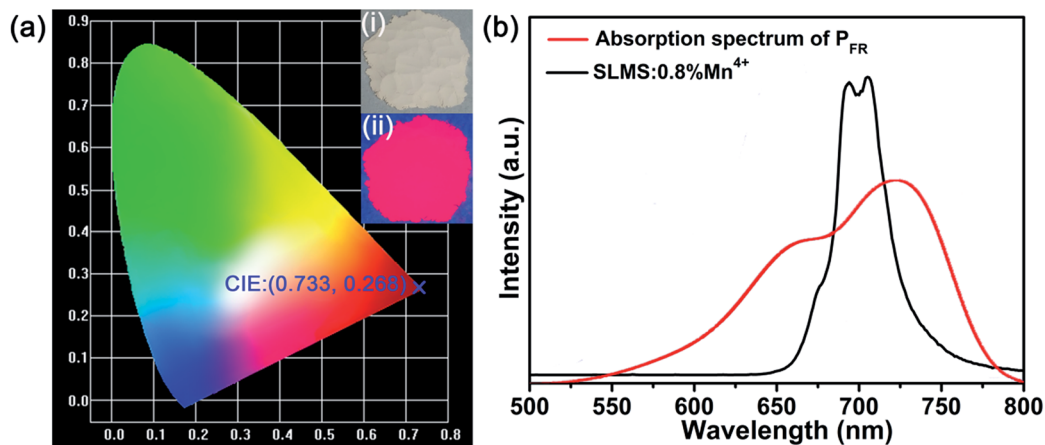


Fig. 5 (a) The CIE chromaticity diagram of the SLMS:0.8%Mn<sup>4+</sup> phosphors. Insets represent the digital photographs under daylight (i) and a 365 nm UV lamp (ii). (b) The PL spectrum of the SLMS:0.8%Mn<sup>4+</sup> phosphors excited at 333 nm and the absorption spectrum of the phytochrome P<sub>FR</sub>.

the activation energy was 0.3 eV. The possible luminescence thermal quenching process could be expressed by the inset in Fig. 6(d). Generally, under the excitation of 333 nm or 499 nm, the electrons in the <sup>4</sup>A<sub>2g</sub> state were firstly excited to the higher <sup>4</sup>T<sub>1g</sub> or <sup>4</sup>T<sub>2g</sub> level, then relax to the lowest <sup>2</sup>E<sub>g</sub> excited state by non-radiative process.<sup>65</sup> After that, the electrons in the <sup>2</sup>E<sub>g</sub> level

may return to the ground state in two ways. The one way was taken place through <sup>2</sup>E<sub>g</sub> → <sup>4</sup>A<sub>2g</sub> radiation transition (path: 1 → 2) with 705 nm far-red emission, and such process was occurred at room temperature. The other way was that the electrons in the <sup>2</sup>E<sub>g</sub> level may absorb enough energy to reach the intersection point at higher temperature, then return to the ground state

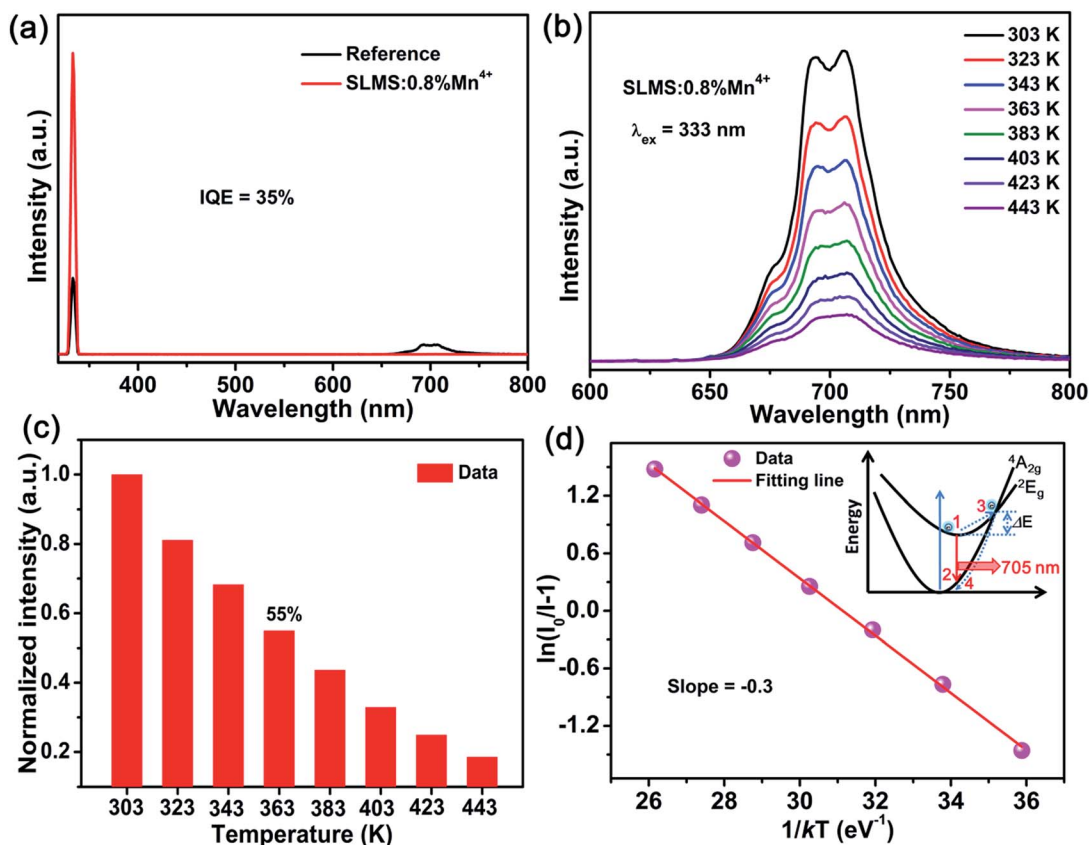
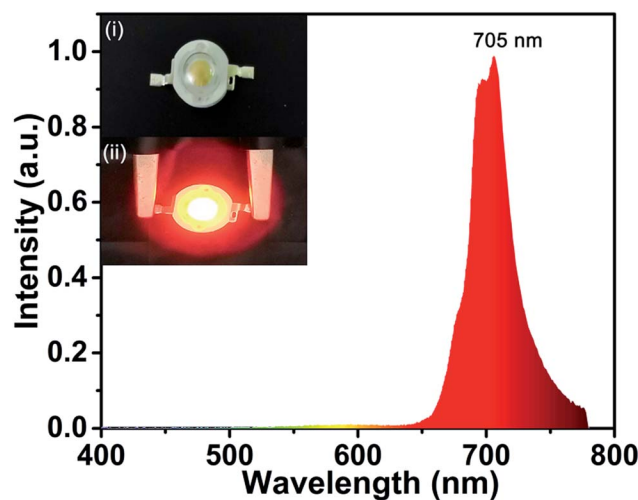


Fig. 6 (a) Excitation profile (λ<sub>ex</sub> = 333 nm) of the reference BaSO<sub>4</sub> and the emission spectrum of the SLMS:0.8%Mn<sup>4+</sup> phosphors obtained by using an integrating sphere. (b) The temperature-dependent spectra of the SLMS:0.8%Mn<sup>4+</sup> phosphors in the temperature range of 303–443 K. (c) The normalized intensity of the PL intensity as a function of the sample temperature. (d) The plot of 1/kT vs. ln(I<sub>0</sub>/I - 1) for the SLMS:0.8%Mn<sup>4+</sup> phosphors. Inset shows the simple configuration coordinate diagram of the <sup>2</sup>E<sub>g</sub> and <sup>4</sup>A<sub>2g</sub> levels of Mn<sup>4+</sup>.





**Fig. 8** EL spectrum of the fabricated far-red-emitting LED device by using SLMS:0.8%Mn<sup>4+</sup> phosphors and a 365 nm near-UV LED chip under a bias current of 240 mA. Insets show the fabricated LED device without (i) and with (ii) a 240 mA bias current.

The external quantum efficiency of the as-prepared far-red LED device was found to be 0.34%. The far-red light from the as-prepared prototype LED device could be easily seen in the inset in Fig. 8. All the results indicated the as-prepared SLMS:Mn<sup>4+</sup> phosphors have a promising prospect to be used in plant growth LEDs.

## Conclusions

In summary, double perovskite SLMS:Mn<sup>4+</sup> phosphors were synthesized by using the high-temperature solid-state reaction. The excitation band of the SLMS:Mn<sup>4+</sup> phosphors could be fitted into four Gaussian peaks at 298, 338, 423, and 495 nm. The PL spectrum of the SLMS:Mn<sup>4+</sup> phosphors showed a far-red emission band peaking at 705 nm, which could match well with the absorption spectrum of the phytochrome P<sub>FR</sub>. The optimal doping concentration of Mn<sup>4+</sup> ions in SLMS:xMn<sup>4+</sup> phosphors was  $x = 0.8\%$ . It was worth noting that the emission intensity of the SLMS:0.8%Mn<sup>4+</sup> phosphors was remained about 55% at 363 K compared with that at 303 K, indicating that the as-prepared phosphors showed good thermal-stability behavior. Furthermore, the IQE of the SLMS:0.8%Mn<sup>4+</sup> phosphors reached as high as 35%. All the above results suggest that the SLMS:xMn<sup>4+</sup> phosphors with far-red emission possess an outstanding potential to be used as far-red emitting phosphors in the application of far-red LEDs for indoor plant growth.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

Fig. 8 shows the EL spectrum of the as-prepared far-red LED device. A far-red emission band in the 600–780 nm wavelength range was observed, attributing to the  ${}^2\text{E}_g \rightarrow {}^4\text{A}_{2g}$  transition.

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