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# Highly intense room-temperature photoluminescence in V<sub>2</sub>O<sub>5</sub> nanospheres

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Room-temperature photoluminescence (PL) in V<sub>2</sub>O<sub>5</sub> nanospheres with zero-like dimensional structure was investigated. A large number of V<sup>4+</sup> oxidation states in the nanospheres were observed by X-ray photoelectron spectroscopy (XPS) measurements. The nanospheres revealed high PL intensity compared with previous studies. In particular, they showed intense ultraviolet (UV) PL near a wavelength of 396 nm (3.13 eV). The intense UV PL was attributed to the enhanced transition probability due to the large number of V<sup>4+</sup> (3d<sup>1</sup>) oxidation states. The PL properties showed strong dependencies on the oxidation states and their distribution.

## 1. Introduction

Vanadium pentoxide (V<sub>2</sub>O<sub>5</sub>) is very stable and has notable applications in chemical sensing, catalysis, Li-ion intercalation, and visible photoluminescence (PL).<sup>1–4</sup> Much effort has been devoted to improving these promising properties, and the synthesis of nanostructures has drawn much attention. Thin-film V<sub>2</sub>O<sub>5</sub> shows visible PL due to band-edge transitions and defects, but its emission intensity is very weak at room temperature. To enhance the PL intensity, various nanostructures have been grown, such as nanorods, nanowires, and nanoparticles.<sup>5–7</sup> Previous studies have reported the synthesis of V<sub>2</sub>O<sub>5</sub> nanostructures and enhanced visible PL. The reported V<sub>2</sub>O<sub>5</sub> nanostructures show a weak and broad emission centred at ~530 nm (2.34 eV) due to the band edge transition and centred at ~710 nm (1.74 eV), which is attributed to mid-gap states (~1.84 eV) formed by oxygen defects that occurred during growth.<sup>8–10</sup> This result may be due to their growth structure. Most nanoparticles and nanorods are attached to each other, and their growth length is also short. The attachment and small growth size of the nanoparticles and nanorods allows them to form many interfaces between them, and the interfaces can promote the separation of photo-generated carrier pairs, which reduces the radiative recombination rate. However, Hu *et al.* reported intense room-temperature PL near 700 nm in V<sub>2</sub>O<sub>5</sub> nanorods with a length of ~2 μm.<sup>9</sup> This result may indicate that intense room-temperature PL can be obtained in micro-length nanorods. Therefore, The PL intensity in V<sub>2</sub>O<sub>5</sub> nanostructures strongly depends on the synthesis method and microstructure. Enhanced PL properties could make V<sub>2</sub>O<sub>5</sub> a promising material for applications in indicators and light sources.

Micro- and nano-sized V<sub>2</sub>O<sub>5</sub> spheres have also been studied as cathode materials for Li-ion batteries to achieve high capacity and high output voltage.<sup>11,12</sup> However, other properties have received less attention, including catalysis, photochromism, and visible PL in V<sub>2</sub>O<sub>5</sub> with spherical structures. One-dimensional structures such as nanorods and nanowires show superior PL properties to two-dimensional structures (thin films). The investigation of nanospheres (which have zero-like dimensional structure) is very important for the development of optical devices.

We report on the synthesis of V<sub>2</sub>O<sub>5</sub> nanospheres and their PL properties at room temperature. The microstructure and oxidation state of the nanospheres were investigated, and PL measurements were performed at room temperature. Also, the PL mechanism in the nanospheres was discussed.

## 2. Experimental section

### Preparation of V<sub>2</sub>O<sub>5</sub> nanospheres

V<sub>2</sub>O<sub>5</sub> nanospheres were synthesized by a chemical reaction method.<sup>11</sup> NH<sub>4</sub>VO<sub>3</sub> of 0.234 g was dissolved in distilled water of 90 mL and stirred for 1 h. Then, 1 M HCl solution of 1 mL was dropped into the solution and stirred for 30 min before adding hydrazine NH<sub>4</sub>·H<sub>2</sub>O of 3 mL, followed by stirring for 10 min at room temperature. The resulting V(OH)<sub>2</sub>NH<sub>2</sub> nanospheres were filtered and washed several times with deionized water and then dried at 80 °C for 6 h. The precursor was dispersed in deionized water and then spin-coated on a quartz substrate at a speed of 3000 rpm for 1 minute. Finally, to obtain V<sub>2</sub>O<sub>5</sub> nanospheres, the coated precursor was calcined using a tubular furnace at 350 °C in air for 2 h.

### Structure characterization

The morphology of the nanospheres was investigated using high-resolution scanning electron microscopy (HR FE-SEM, Mira LMH, Tescan) and transmission electron microscopy

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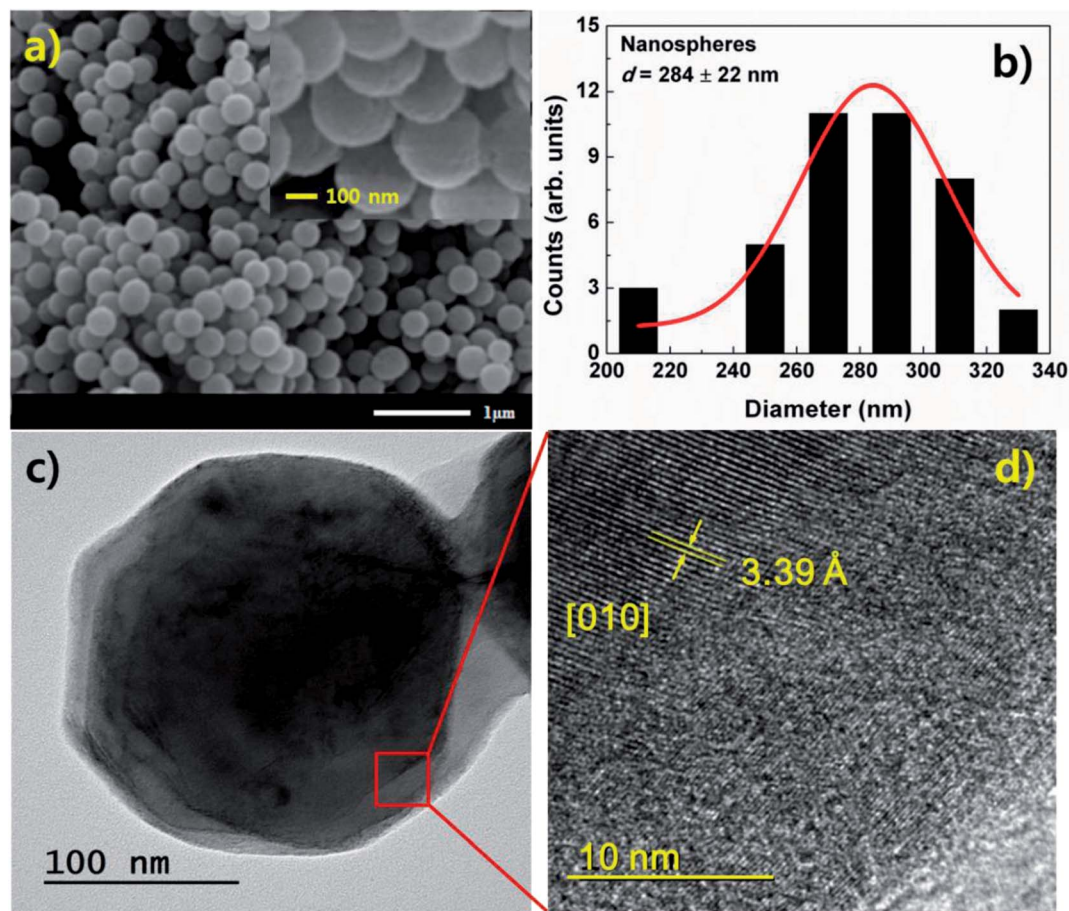


Fig. 1 (a) SEM image and (b) histogram of  $\text{V}_2\text{O}_5$  nanospheres. (c) TEM image of the nanospheres showing the solid sphere structure. (d) High-resolution TEM image with atomic array of  $\text{V}_2\text{O}_5$  observed at center of the nanosphere.

(TEM, H-8100, Hitachi). The crystal structure was characterized using an X-ray diffractometer (XRD, DMax 2000, Rigaku) with  $\text{Cu K}\alpha$  radiation, as well as a Raman/PL spectrometer (Horiba-Jobin Yvon, LabRAM HR) equipped with a He-Ar laser at a wavelength of 514.5 nm.

### Characterization

X-ray photoelectron spectroscopy (XPS) measurements were performed to investigate the oxidation states of the nanospheres. The XPS spectra were recorded with a Thermo-Scientific MultiLab ESCA 2000 system equipped with a 300 W  $\text{Mg K}\alpha$  source. The  $\text{C1s}$ ,  $\text{O1s}$ , and  $\text{V2p}$  core levels were measured. When fitting the  $\text{V2p}$  and  $\text{O1s}$  core-level spectra, the Gaussian width was fixed at the instrumental resolution of 0.75 eV. For curve-fitting of the  $\text{V2p}$  peaks, the values of the spin-orbit splitting and the branching ratios  $[I(2p_{3/2})/I(2p_{1/2})]$  were fixed at 7.6 eV and 0.5, respectively.

PL measurements were performed at room temperature using a Raman/PL spectrometer (Horiba-Jobin Yvon, LabRAM HR) equipped with a He-Cd laser with a wavelength of 325 nm and a maximum power of 200 mW. Time-resolved PL (TRPL) measurements were performed using an inverted scanning confocal microscope (MicroTime-200, Picoquant). A single-

mode pulsed diode laser (LDH-P-C-375, Picoquant) with a pulse width of  $\sim 30$  ps was used as an excitation source. The time-correlated single-photon counting (TCSPC) technique was used to count the PL photons. TRPL images with dimensions of  $200 \times 200$  pixels were recorded using the time-tagged time-resolved (TTTR) data acquisition method. The acquisition time of each pixel was 1 ms.

## 3. Results and discussion

Fig. 1(a) and (b) show the morphology and a histogram of  $\text{V}_2\text{O}_5$  nanospheres obtained from SEM measurements. The size distribution of the nanospheres was evaluated using the histogram, which has a Gaussian profile. To obtain the size distribution, we used the ImageJ program. The size of individual nanosphere in SEM image was calculated by using the program. The SEM image and histogram reveal that the nanospheres have a diameter of  $284 \pm 22$  nm, spherical morphology, and numerous defects on their surfaces. These defects are due to the growth feature of  $\alpha\text{-V}_2\text{O}_5$  with a layer-by-layer (plate-like) structure.  $\text{V}_2\text{O}_5$  nanospheres grown in this study have a crystal structure of  $\alpha\text{-V}_2\text{O}_5$ . Initially,  $\text{V}_2\text{O}_5$  is grown to a layer (or plate) structure. As time passes, the number of layer increases. Simultaneously, edges of the plate structure are etched by



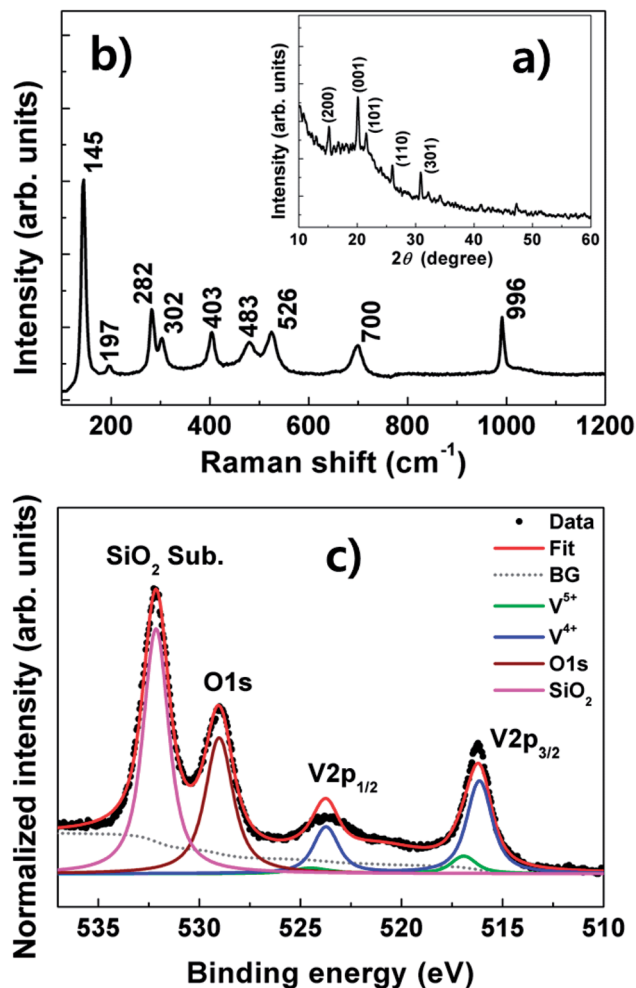


Fig. 2 (a) XRD pattern and (b) Raman spectrum of V<sub>2</sub>O<sub>5</sub> nanospheres. (c) XPS spectrum of the nanospheres. The V2p and O1s core-level spectra shift toward lower binding energy, and the peak at 532.2 eV originated from the SiO<sub>2</sub> substrate.

solution. Due to the process, the layer-by-layer structure changes to sphere structure, and numerous defects are formed on surface of sphere.<sup>13,14</sup>

A TEM image of the nanospheres is shown in Fig. 1(c). The image shows solid sphere structures. Fig. 1(d) shows a high-resolution TEM image. An atomic array of V<sub>2</sub>O<sub>5</sub> is observed at the center of the sphere, indicating high crystallinity. On the surface, the atomic array becomes disordered due to the etching, implying poor crystallinity and the formation of numerous defects. In the observed region, a lattice parameter of  $b = 3.39 \text{ \AA}$  for  $\alpha$ -V<sub>2</sub>O<sub>5</sub> was obtained from the TEM image (Fig. 1(d)). The center of the sphere shows a preferential [010] orientation.

Fig. 2(a) shows the XRD pattern of the nanospheres, where there is a diffraction peak corresponding to the (001) plane of  $\alpha$ -V<sub>2</sub>O<sub>5</sub> at  $20.3^\circ$  (JCPDS 01-077-2418).<sup>8–10</sup> Additional peaks corresponding to the (200), (101), (110), and (301) planes are also observed. This result indicates that the nanosphere has polycrystalline structure. Furthermore, the pattern shows

unidentified peaks. The existence of unidentified peaks implies that the V<sub>2</sub>O<sub>5</sub> nanosphere includes other phases of vanadium oxide. The measured Raman spectra are presented in Fig. 2(b). The spectra reveal the typical Raman-active mode of  $\alpha$ -V<sub>2</sub>O<sub>5</sub>.<sup>15,16</sup> The results in Fig. 2(a) and (b) indicate that the nanospheres contain the  $\alpha$ -V<sub>2</sub>O<sub>5</sub> phase.

Fig. 2(c) shows the XPS spectra of the nanospheres. The V2p and O1s core-level spectra of the nanospheres shift toward lower binding energy with the appearance of an additional oxygen peak at higher binding energy, which are attributed to the SiO<sub>2</sub> substrate. A deconvolution fit indicates that the V2p peak (p<sub>3/2</sub>) of the nanospheres mainly consists of the V<sup>4+</sup> oxidation state (516.1 eV). The energy of the peak is very small (0.8 eV) compared with the V<sup>5+</sup> oxidation state (516.9 eV) corresponding to the  $\alpha$ -V<sub>2</sub>O<sub>5</sub> phase.<sup>17–19</sup> This result implies that the states are oxidized into a lower oxidation state. Furthermore, the peak maxima of the O1s core-level spectra shifts to lower energy for the nanospheres. The stoichiometry ratios ( $S_{ij}$ ) of the nanospheres are calculated from the XPS spectra using the following formula:  $S_{ij} = C_i/C_j = (I_i/ASF_i)/(I_j/ASF_j)$ , where,  $C_i$  and  $C_j$  are the concentrations of the elements,  $I_i$  and  $I_j$  the background corrected intensities of the photoelectron emission lines and ASF<sub>*i*</sub> and ASF<sub>*j*</sub> the atomic sensitivity factors for photoionization of the  $i_{th}$  and  $j_{th}$  elements. For example, ASF<sub>O</sub> and ASF<sub>V</sub> are 2.93 and 9.66, which are supplied with the instrument.<sup>20</sup> The O/V rate of the nanospheres is 3.04, while the amount of V<sup>4+</sup> in the V2p peak is 84%.

Fig. 3(a) presents the PL spectra of phase nanospheres measured at room temperature. The nanospheres exhibit highly intense PL with three main peaks. The PL peak centered at  $\sim 530 \text{ nm}$  (2.34 eV) is due to the band edge transition in V<sub>2</sub>O<sub>5</sub> with a bandgap energy of 2.3–2.4 eV.<sup>21,22</sup> Another PL peak centered at  $\sim 710 \text{ nm}$  (1.74 eV) is attributed to mid-gap states ( $\sim 1.84 \text{ eV}$ ) formed by oxygen defects that occurred during growth.<sup>9,10</sup> The nanospheres show an intense UV PL peak centered at  $\sim 396 \text{ nm}$  (3.13 eV), which has not been reported previously. This peak is considered to be due to strong UV absorption resulting from the large amount of V<sup>4+</sup> oxidation states in the nanospheres (Fig. 2(c)), which can be strongly excited by UV light with wavelength in the range of 250–350 nm.<sup>23</sup> Meyn *et al.* investigated the PL properties of V<sup>4+</sup>-doped oxides such as Al<sub>2</sub>O<sub>3</sub> and YAlO<sub>3</sub> and reported that the UV light absorption in the oxides can be enhanced by V<sup>4+</sup>.

The V<sup>5+</sup> and V<sup>4+</sup> oxidation states have valence-electron configurations of V 3d<sup>0</sup> and V 3d<sup>1</sup> in the valence band,<sup>24,25</sup> respectively. The binding energy (516.1 eV) of the nanospheres with a dominant V<sup>4+</sup> oxidation state decreases to 0.8 eV, which is small compared with that of  $\alpha$ -V<sub>2</sub>O<sub>5</sub> (516.9 eV) with a dominant V<sup>5+</sup> oxidation state. The 3d<sup>1</sup> valence-electron configuration and the lower binding energy of the V<sup>4+</sup> oxidation state can significantly increase the probability of transition from the valence band to the V3d t<sub>2g</sub> and V3d e<sub>g</sub> conduction bands *via* UV light absorption (Fig. 4(h)).<sup>24,26</sup> The enhanced transition leads to intense PL in the nanospheres. Thus, it is considered that the PL properties of V<sub>2</sub>O<sub>5</sub> nanospheres strongly depend on the oxidation state.





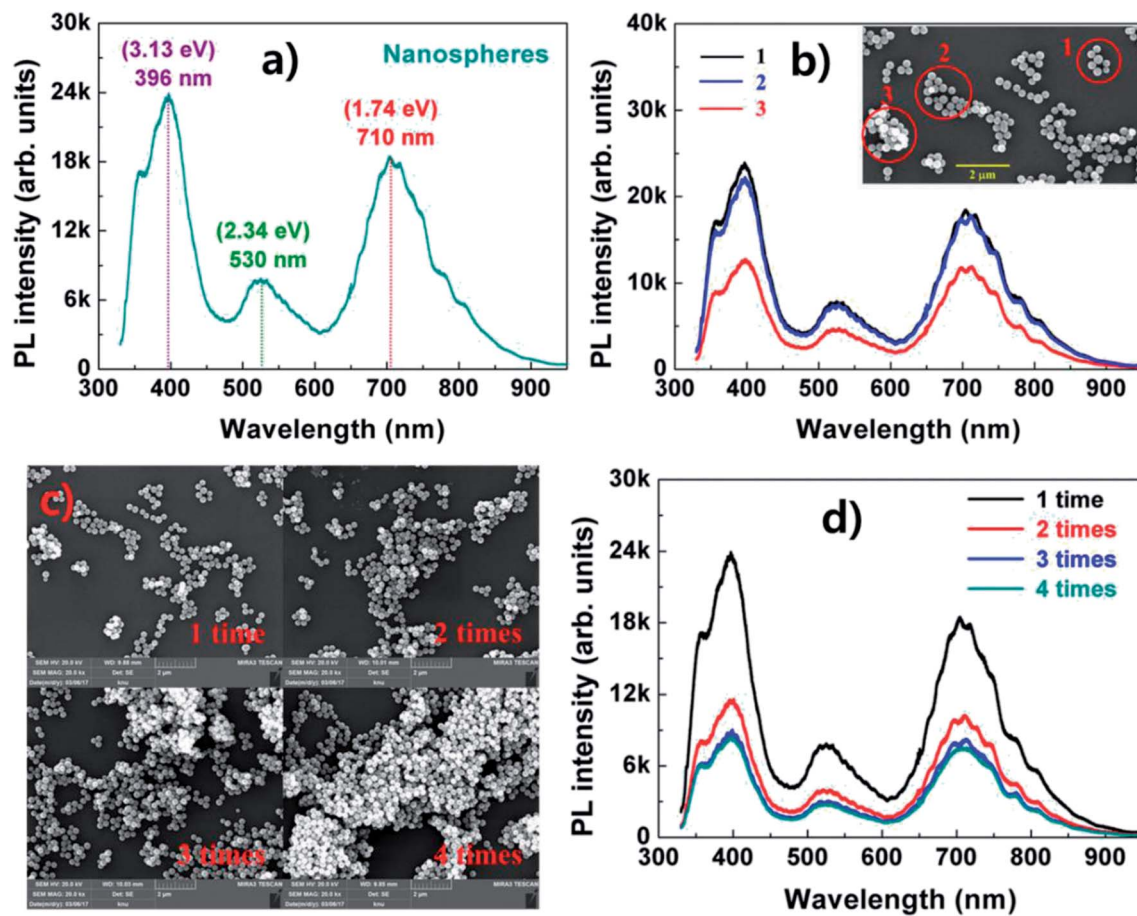


Fig. 3 (a) Room-temperature PL spectrum in  $V_2O_5$  nanospheres. (b) PL spectra with distribution of the nanospheres and (Inset) SEM image. (c) SEM images and (d) PL spectra with respect to spin-coating frequency of the nanospheres.

Fig. 3(b) reveals the PL spectra with the distribution of the nanospheres. As shown in the inset of Fig. 3(b), the distribution of  $V_2O_5$  nanospheres is not uniform. Therefore, we measured the PL spectra with different distributions of the nanospheres. The PL intensities in locations 1 and 2 with an individual distribution do not change much. However, the intensity at location 3, which has a cluster distribution, significantly decreases over the entire spectral range without variation of the emission shape. This result is attributed to the reduction of the absorption due to light scattering in the nanospheres with a cluster distribution. The result in Fig. 3(b) is confirmed in Fig. 3(c) and (d). As the spin-coating frequency increases, the nanosphere distribution changes from clusters to many stacks, and the PL intensity decreases considerably without variation of the emission shape.

The TRPL spectra for the three PL peaks of the  $V_2O_5$  nanospheres are shown in Fig. 4(a)–(d). An exponential fit was performed using Symphotime-64 software (Ver. 2.2) with a bi-exponential decay model,  $I(t) = \sum A_i e^{-t/\tau_i}$ , where  $I(t)$  is the time-dependent PL intensity,  $A$  is the amplitude,  $\tau$  is the PL lifetime, and the ordinal  $i$  is 1 or 2 in this study. The lifetimes for the PL peaks at 396, 530, and 710 nm were obtained from the bi-exponential decay fit.

As shown in Fig. 4(b)–(d), the TRPL spectra of the three PL peaks reveal fast decay with a lifetime of less than 1 ns. The fast component ( $\tau_1$ ) and slow component ( $\tau_2$ ) of the lifetimes for the three PL peaks are 0.6 and 1.90 ns at 396 nm, while they are 0.54 and 3.00 ns at 530 nm and 0.59 and 1.79 ns at 710 nm. This decay feature of the nanospheres may indicate a lack of non-radiative channels. Fig. 4(e)–(g) show the TRPL images obtained using the TTTR data acquisition method. The images of the PL peaks at 396 and 710 nm reveal similar aspects, whereas the image of the PL peak at 530 nm shows a slow decay feature compared with those at 396 and 710 nm. The increase of the red spot in Fig. 4(f) indicates an increase of the slow component in lifetime.

Fig. 4(h) shows a band diagram for the PL process in the  $V_2O_5$  nanospheres. The electronic band structure of  $V_2O_5$  is formed by the associated hybridization of the O 2p orbitals with V 3d orbitals (p–d hybridization) involved in chemical bonding.<sup>24–26</sup> The V 3d state in the conduction band is split into a low  $t_{2g}$  state and a high  $e_g$  state. In the case of the  $V^{4+}$  ( $3d^1$ )-dominant nanospheres, carriers are strongly excited due to the  $V^{4+}$ , and the PL peaks originating from the  $t_{2g}$  state and mid-gap state are highly enhanced. Furthermore, the highly intense UV PL ( $\sim 396$  nm) is due to carrier recombination in the  $e_g$  state.



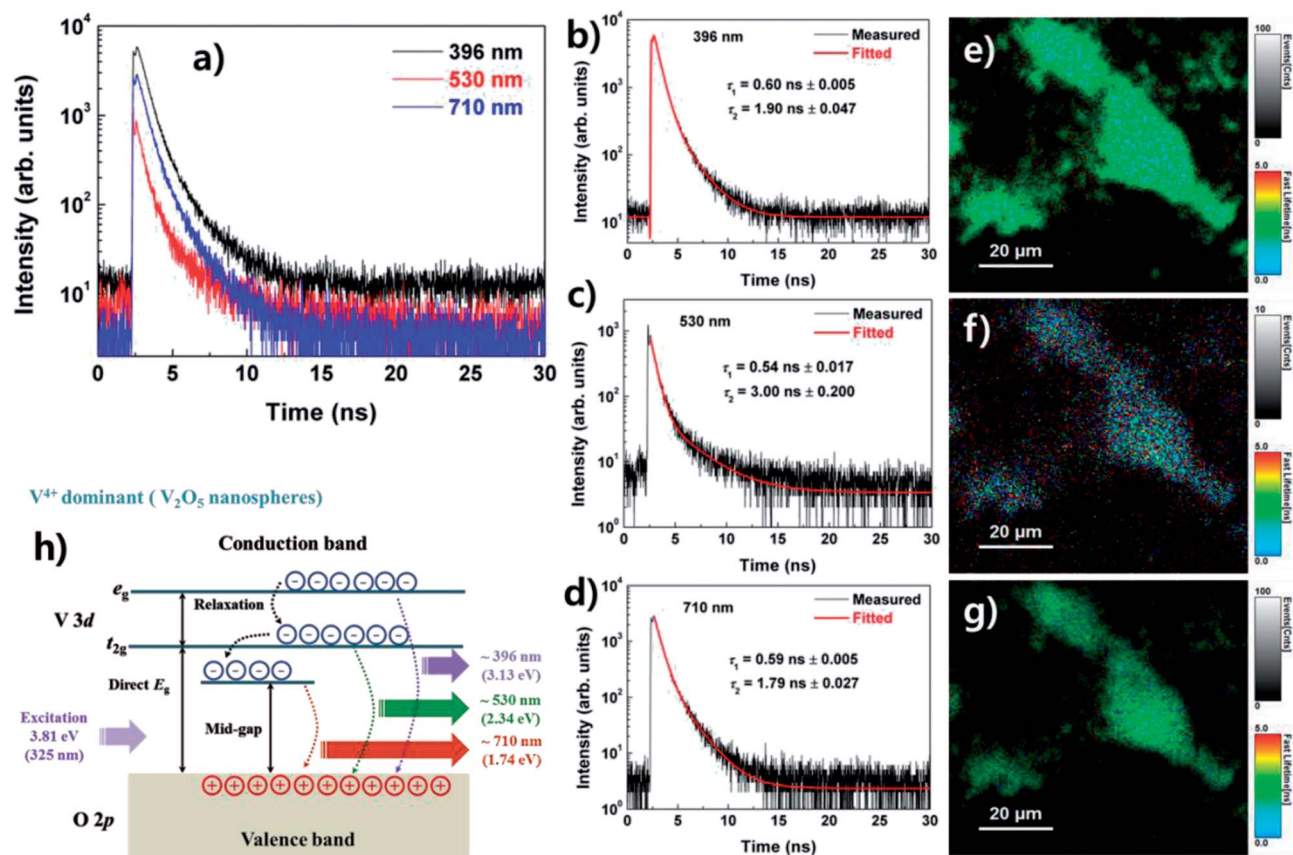


Fig. 4 (a) TRPL spectra in  $V_2O_5$  nanospheres. The lifetimes for the PL peaks at (b) 396, (c) 530, and (d) 710 nm were obtained from the bi-exponential decay fit. (e)–(g) TRPL images obtained using the time-tagged time-resolved (TTTR) data acquisition method. (h) Band diagram of the PL process of  $V_2O_5$  nanospheres.

## 4. Conclusion

The  $V_2O_5$  nanospheres prepared by a chemical reaction method showed a large amount of  $V^{4+}$  oxidation states (84%). Furthermore, the nanospheres revealed very high PL intensity. In particular, the nanospheres showed intense UV PL near 396 nm (3.13 eV), which was attributed to the enhanced transition probability in terms of the large amount of the  $V^{4+}$  ( $3d^1$ ) oxidation states. The PL properties of the nanospheres showed strong dependencies on the oxidation state and their distribution on the substrate.

## Conflicts of interest

There are no conflicts to declare.

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## References

- 1 A. Z. Moshfegh and A. Ignatiev, *Thin Solid Films*, 1991, **198**, 251.
- 2 A. Legrouiri, T. Baird and J. R. Fryer, *J. Catal.*, 1993, **140**, 173.
- 3 S. L. Chou, J. Z. Wang, J. Z. Sun, D. Wexler, M. Forsyth, H. K. Liu, D. R. MacFarlane and S. X. Dou, *Chem. Mater.*, 2008, **20**, 7044.
- 4 M. Kang, S. W. Kim, Y. Hwang, Y. Um and J. W. Ryu, *AIP Adv.*, 2013, **3**, 052129.
- 5 Y. Wang, Z. Li, X. Sheng and Z. Zhang, *J. Chem. Phys.*, 2007, **126**, 164701.
- 6 A. Othonos, C. Christofides and M. Zervos, *Appl. Phys. Lett.*, 2013, **103**, 133112.
- 7 D. W. Su, S. X. Dou and G. X. Wang, *J. Mater. Chem. A*, 2014, **2**, 11185.
- 8 C. W. Zou, X. D. Yan, J. M. Bian and W. Gao, *Opt. Lett.*, 2010, **35**, 1145.
- 9 Y. Hu, Z. Li, Z. Zhang and D. Meng, *Appl. Phys. Lett.*, 2009, **94**, 103107.
- 10 M. Kang, M. Chu, S. W. Kim and J. W. Ryu, *Thin Solid Films*, 2013, **547**, 198.
- 11 J. Shao, X. Li, Z. Wan, L. Zhang, Y. Ding, L. Zhang, Q. Qu and H. Zheng, *ACS Appl. Mater. Interfaces*, 2013, **5**, 7671.



- 12 X. F. Zhang, K. X. Wang, X. Wei and J. S. Chen, *Chem. Mater.*, 2011, **23**, 5290.
- 13 L. Abello, E. Husson, Y. Repelin and G. Lucazeau, *Spectrochim. Acta, Part A*, 1983, **39**, 641.
- 14 D. W. Murphy, P. A. Christian, F. J. DiSalvo and J. V. Waszczak, *Inorg. Chem.*, 1979, **18**, 2800.
- 15 C. V. Ramana, R. J. Smith, O. M. Hussian, M. Massor and C. M. Julien, *Surf. Interface Anal.*, 2005, **37**, 406.
- 16 L. J. Meng, R. A. Silva, H. N. Cui, V. Teixeira, M. P. Dos Santos and Z. Xu, *Thin Solid Films*, 2006, **515**, 195.
- 17 C. A. Sawatzky and D. Post, *Phys. Rev. B*, 1979, **20**, 1546.
- 18 M. Demeter, M. Neumann and W. Reichelt, *Surf. Sci.*, 2000, **41**, 454.
- 19 G. Silversmit, D. Depla, H. Poelman, G. B. Marin and R. De Gryse, *J. Electron Spectrosc. Relat. Phenom.*, 2004, **135**, 167.
- 20 Q. H. Wu, A. Thissen, W. Jaegermann and M. Liu, *Appl. Surf. Sci.*, 2004, **236**, 473.
- 21 N. Kenny, *J. Phys. Chem. Solids*, 1996, **27**, 1237.
- 22 C. Lamsal and N. M. Ravindra, *J. Mater. Sci.*, 2013, **48**, 6341.
- 23 J. P. Meyn, T. Danger, K. Petermann and G. Huber, *J. Lumin.*, 1993, **55**, 55.
- 24 T. M. Tolhurst, B. Leedahl, J. L. Andrews, P. M. Marley, S. Banerjee and A. Moewes, *Phys. Chem. Chem. Phys.*, 2016, **18**, 15798.
- 25 R. J. O. Mossaneck, A. Mocellin, M. Abbate, B. G. Searle, P. T. Fonseca and E. Morikawa, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2008, **77**, 075118.
- 26 A. Chakrabarti, K. Hermann, R. Druzinic, M. Witko, F. Wagner and M. Petersen, *Phys. Rev. B*, 1999, **59**, 10583.

