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# Hierarchical nanostructures of nitrogen-doped molybdenum sulphide for supercapacitors†

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Flower-like nanostructures of molybdenum disulphide (MoS<sub>2</sub>) have been effectively synthesised by the hydrothermal method and further doped with nitrogen using varying concentrations of urea. The formed hierarchical nanostructures are characterised by spectroscopy as well as electrochemical techniques. The structural analysis confirms the formation of a hexagonal MoS<sub>2</sub> crystal structure. The existence of MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> composites is also observed after heating MoS<sub>2</sub> with a lower urea concentration. Surface morphological analysis of all the prepared compositions shows the appearance of flower-like nanostructures formed by the stacking of 20–80 nanosheets to create individual flower petals. Nitrogen doping shows enhancement in the specific capacitance of MoS<sub>2</sub> due to an increase in the electronic conductivity. Furthermore, the specific capacitance is enhanced due to the formation of an MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> composite. The highest specific capacitance calculated from the charge–discharge curve for nitrogen-doped MoS<sub>2</sub> prepared using 1 : 1 (MoS<sub>2</sub> : urea) weight ratio is observed at around 129 (F g<sup>-1</sup>) at 2 (A g<sup>-1</sup>) specific current. The nitrogen-doped MoS<sub>2</sub> demonstrates almost four-fold enhancement in specific capacitance than pristine nano-shaped MoS<sub>2</sub>.

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## Introduction

Supercapacitors are efficient energy storage devices that act as a link between traditional capacitors and batteries. Higher power densities, longer life cycles and fast charging-discharging rates of supercapacitors play important roles in applications in industrial energy management, military devices, transportation, portable devices, memory and power backups, and hybrid vehicles.<sup>1–4</sup> These high-energy capacitors have importance in scientific research as clean energy sources.

There are two different types of charge storage mechanisms for supercapacitors; the first one is double-layer capacitance that arises due to non-faradaic charge separation at the electrolyte/electrode interface, and the second one is pseudocapacitance that arises due to a surface redox reaction of the active materials such as conducting polymers, metal oxides and

hydroxides. Carbon materials such as active carbon, carbon nanotubes, and graphene are the most commonly used electrode materials owing to their high surface area and electrical conductivity, and they come under the category of double layer capacitors.<sup>1,5–7</sup> Metal organic frameworks, metal oxides, and metal sulphides/selenides have also been reported as supercapacitor electrodes because of the presence of pseudocapacitance.<sup>8–17</sup> The common strategy to enhance the capacitance of supercapacitors is to form composites using metal oxides and conducting polymers with carbon materials to obtain pseudocapacitance as well as double-layer capacitance.

Molybdenum disulphide (MoS<sub>2</sub>) is a two-dimensional transition metal dichalcogenide that has a layered sheet-like structure. MoS<sub>2</sub> acts as a good electrode material due to the presence of both pseudocapacitance and double-layer capacitance. The large surface area of layered MoS<sub>2</sub> is responsible for double-layer capacitance, whereas redox reactions at the Mo metal center give rise to pseudocapacitance. Therefore, MoS<sub>2</sub> can replace composites of metal oxides and carbon as it has faster intrinsic ionic conductivity than oxides and higher theoretical capacitance than graphite (~700 F cm<sup>-3</sup>).<sup>18</sup> There are only few investigations on MoS<sub>2</sub> as an electrode material for electrochemical applications.<sup>10,19,20</sup> Some of the reports include doping and composite formation of MoS<sub>2</sub> with graphene/nitrogen-doped graphene for supercapacitor and battery applications.<sup>21–26</sup>

The main difficulty in using MoS<sub>2</sub> electrode in supercapacitor applications is its low rate performance and capacity fading.<sup>25</sup> The use of nanomaterials and doped materials can

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enhance the electrochemical performance of MoS<sub>2</sub> and help overcome these capacitance problems. There are reports demonstrating enhancement in the electronic properties of MoS<sub>2</sub> by doping with transition metals.<sup>27,28</sup> Nevertheless, non-metal doping of MoS<sub>2</sub> that increases supercapacitor performance has not yet been reported. Nitrogen-doped graphene has shown enhanced performance for supercapacitors and other energy storage devices.<sup>29–33</sup> In the case of graphene, defects improve the storage capacity of the graphene electrode but decrease its conductivity; therefore, the overall performance of the graphene electrode decreases. Nitrogen doping helps to improve this decreased performance by enhancing the conductivity.<sup>30–32</sup> MoS<sub>2</sub> nanosheets are similar to graphene sheets; therefore, nitrogen-doped MoS<sub>2</sub> can be expected to show enhanced performance as an electrode material for supercapacitors.

In this regard, for the first time, we report the synthesis of MoS<sub>2</sub> and nitrogen-doped MoS<sub>2</sub> nanosheets as electrode materials for supercapacitor applications using a hydrothermal route. The increase in nitrogen doping in MoS<sub>2</sub> shows marginal increase in specific capacitance due to enhancement in electronic conductivity.

## Experimental

### Chemicals

Ammonium heptamolybdate tetrahydrate, thiourea, urea, and sulphuric acid were purchased from Sigma Aldrich and used as such without further purification.

### Synthesis of 2H-MoS<sub>2</sub>

The hydrothermal method was used to synthesize semi-conducting hexagonal molybdenum disulphide nanostructures. In a typical experimental procedure, ammonium heptamolybdenate and thiourea were used as molybdenum and sulphur precursors, respectively, to synthesize MoS<sub>2</sub>. A mixture of ammonium heptamolybdenate and thiourea in 60 ml of water was prepared by keeping the molar ratio of Mo : S as 1 : 8. Then, this mixture was transferred to a Teflon-lined stainless steel autoclave. The sealed autoclave was kept at 200 °C for 24 h in a hot air oven. The obtained reaction product was washed with water and subsequently with ethanol. This powder was dried in a vacuum oven at 50 °C for 5 h and used for further process.

### Nitrogen doping of MoS<sub>2</sub>

The synthesised 2H-MoS<sub>2</sub> powder was heated in a tube furnace under a nitrogen environment at 400 °C for 4 h. The obtained product was termed as MoS<sub>2</sub>-400. Furthermore, to dope with nitrogen, MoS<sub>2</sub>-400 was mixed with urea in weight ratios of 1 : 1 and 1 : 5 and subsequently heated in a tube furnace at 400 °C for 30 min under a nitrogen environment; the obtained products were termed as MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5, respectively.

### Characterisation

As-synthesized MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5 nanostructures were characterized by different spectroscopic

techniques. The phase purity of the materials was determined by X-ray diffraction (XRD) using a Bruker AXS D8 Advance powder X-ray diffractometer with Cu K $\alpha$  radiation ( $\lambda = 1.5417$  Å). The diffraction patterns were recorded in the range of  $2\theta$  (10° to 80°) with a step size of 0.01°. Raman spectra were recorded using Jobin Yvon HORIBA LabRAM HR visible micro Raman system, employing a He–Ne laser operating at 632.8 nm as the source in the backscattering mode. The morphology of the materials was studied by scanning electron microscopy with a Zeiss™ Ultra Plus field-emission scanning electron microscope (FESEM) equipped with an energy-dispersive X-ray (EDAX) spectrometer. Field emission transmission electron microscopy (FETEM, JEOL, JEM-2200FS) operated at 200 keV was used to obtain the morphology and SAED pattern for synthesised MoS<sub>2</sub> nanomaterials. X-ray photoelectron spectroscopy (XPS) of the prepared MoS<sub>2</sub> materials was performed with MultiLab 2000, Thermo VG with Mg K $\alpha$  X-ray source to determine the chemical compositions of the samples.

Electrochemical measurements were carried out using three-electrode cell configuration on Autolab PGSTAT302N. Cyclic voltammetry measurements of the MoS<sub>2</sub> electrodes were performed in H<sub>2</sub>SO<sub>4</sub> electrolyte at different scan rates over a potential window from –0.2 to +0.8 V with MoS<sub>2</sub> as the working electrode, platinum as the counter electrode and Ag/AgCl as the reference electrode. The loaded mass of the synthesised MoS<sub>2</sub> on the working electrode was around 3–4 mg for each sample. Charge–discharge characterisations were performed at different specific currents within the potential window from –0.2 to +0.8 V with a similar experimental setup. Electrochemical impedance spectroscopy (EIS) measurements were carried out in the range from 100 MHz to 100 kHz with an AC amplitude of 10 mV in 1 M H<sub>2</sub>SO<sub>4</sub> electrolyte.

## Results and discussion

The XRD patterns of synthesised MoS<sub>2</sub> and nitrogen-doped MoS<sub>2</sub> nanostructures are depicted in Fig. 1(a–d) for  $2\theta$  range from 10° to 80°. The intense and broad reflection peaks confirm the formation of the 2H-MoS<sub>2</sub> nanostructure for all the prepared compositions. The XRD pattern of the synthesised MoS<sub>2</sub>

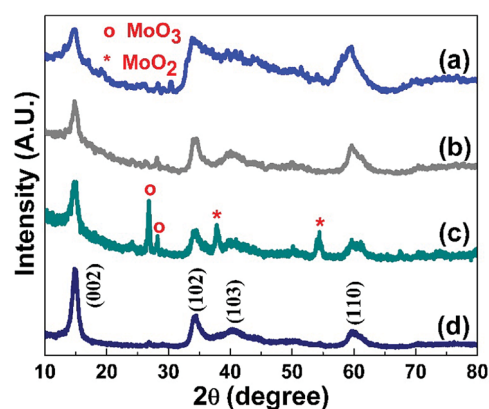


Fig. 1 X-ray diffraction patterns for MoS<sub>2</sub> (a), MoS<sub>2</sub>-400 (b), MoS<sub>2</sub>-N1 (c), and MoS<sub>2</sub>-N5 (d).



matches the standard MoS<sub>2</sub> pattern (JCPDS ICDD no. 37-1492). The observed XRD peaks at  $2\theta = 14.53^\circ, 32.67^\circ, 33.50^\circ, 35.87^\circ, 39.53^\circ, 58.33^\circ,$  and  $60.14^\circ$  correspond to (002), (100), (101), (102), (103), (110), and (008) *hkl* planes of hexagonal MoS<sub>2</sub>. The XRD patterns of nitrogen-doped MoS<sub>2</sub> nanostructures (MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5) are compared with the XRD patterns of as-synthesised MoS<sub>2</sub> and MoS<sub>2</sub>-400, as depicted in Fig. 1(a–d). After nitrogen doping, there is slight change in the XRD peak intensity, but the peak positions remain the same. In the case of MoS<sub>2</sub>-N1 sample, the formation of MoO<sub>2</sub> and MoO<sub>3</sub> along with MoS<sub>2</sub> is observed; this clearly indicates the formation of MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> nanocomposites. The reflection peaks due to MoO<sub>2</sub> and MoO<sub>3</sub> match with JCPDS no. 32-0671 and 47-1081, respectively. Composite formation is observed for MoS<sub>2</sub>-N1 sample only. Even though the synthesis methods for MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5 are similar, the conversion of MoS<sub>2</sub> to MoO<sub>2</sub>/MoO<sub>3</sub> is different, as seen in the XRD pattern. In the case of MoS<sub>2</sub>-N1, the amount of carbon available (in the form of urea) to react with oxygen to form carbon dioxide is low as compared to that for MoS<sub>2</sub>-N5. Therefore, for MoS<sub>2</sub>-N1, the remaining oxygen reacts with MoS<sub>2</sub> to form molybdenum oxide. Thus, we observe difference in oxide formations for MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5.

The broadening of the XRD peak decreases, indicating a more crystalline structure for nitrogen-doped MoS<sub>2</sub>. The broad XRD peaks confirm that the prepared structures are still in the nanoscale even after heating at 400 °C for nitrogen doping. Furthermore, structural analysis of prepared MoS<sub>2</sub> and nitrogen-doped MoS<sub>2</sub> are conducted by Raman spectroscopy. Fig. 2(a–d) shows the Raman scattering curves for MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5 in the range from 350 to 450 cm<sup>-1</sup>. The two peaks correspond to in-plane (E<sub>2g</sub><sup>1</sup>) and out-of-plane (A<sub>1g</sub>) vibrations at 378 and 403 cm<sup>-1</sup> for the synthesised MoS<sub>2</sub> materials. These two characteristic Raman peaks confirm the formation of nanoscale MoS<sub>2</sub>. There is negligible change in the peak separation between A<sub>1g</sub> and E<sub>2g</sub><sup>1</sup> for nitrogen-doped MoS<sub>2</sub> compare to that for pristine MoS<sub>2</sub>. The A<sub>1g</sub>/E<sub>2g</sub><sup>1</sup> ratios for MoS<sub>2</sub>, MoS<sub>2</sub> 400, MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5 are 2.3, 2.7, 2.2 and 2.1,

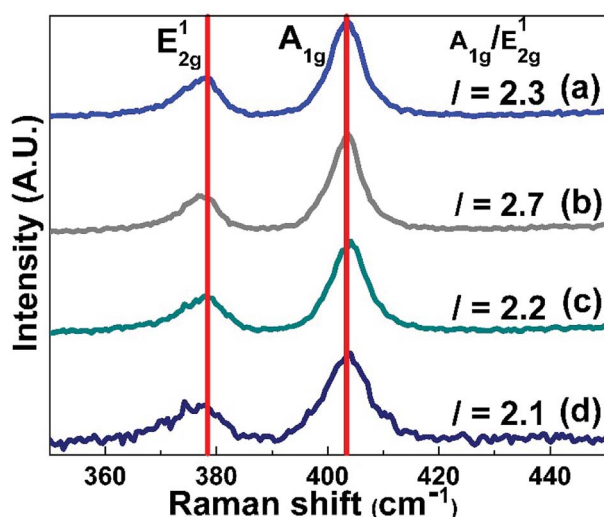


Fig. 2 Raman spectra for MoS<sub>2</sub> (a), MoS<sub>2</sub>-400 (b), MoS<sub>2</sub>-N1 (c), and MoS<sub>2</sub>-N5 (d).

respectively. The A<sub>1g</sub>/E<sub>2g</sub><sup>1</sup> ratio decreases from MoS<sub>2</sub> 400 to MoS<sub>2</sub>-N5, confirming that the defect edge sites of MoS<sub>2</sub> are occupied by nitrogen species.

The surface and morphological studies of prepared MoS<sub>2</sub> nanomaterials were carried out using FESEM and TEM analysis. The magnified images of MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1, and MoS<sub>2</sub>-N5 are shown in Fig. 3(a–d). The formation of flower-like nanostructures is observed in the prepared MoS<sub>2</sub> samples. The flower-like nanostructure is formed by stacking of nanosheets up to 20–50 nm thickness. As exhibited in Fig. 3(b), MoS<sub>2</sub>-400 shows growth and an opening of nanosheets into flower-shaped nanostructures. The nitrogen doping in MoS<sub>2</sub> nanostructures results in rupturing of nanosheets in MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5, as shown in Fig. 3(c and d). In the case of MoS<sub>2</sub>-N1 sample, the existence of rods and rectangular plates is observed, which may be due to the formation of MoO<sub>2</sub> and MoO<sub>3</sub> nanostructures (ESI, Fig. S1†). Further detailed study related to oxide formation is confirmed by TEM analysis.

Microstructural analysis of prepared MoS<sub>2</sub> nanomaterials was conducted using TEM analysis, and the images are shown in Fig. 4. Fig. 4 shows the TEM images of MoS<sub>2</sub> (a and e), MoS<sub>2</sub>-400 (b and f), MoS<sub>2</sub>-N1 (c and g) and MoS<sub>2</sub>-N5 (d and h). The layered structure of prepared MoS<sub>2</sub> is confirmed from TEM analysis. TEM clearly indicates the formation of flower-like morphology consisting of stacked nanosheets as flower petals, and this is consistent with the FESEM results. The size of petals in the as-synthesised MoS<sub>2</sub> is between 20 and 50 nm; for MoS<sub>2</sub>-400, the size is between 20 and 80 nm. However, it is observed that the petal size range is decreased to 20–40 nm in the case of MoS<sub>2</sub>-N1. Every petal is formed by stacking of two-dimensional MoS<sub>2</sub> nanosheets, as can be seen in Fig. 4(e–h). As discussed in the FESEM analysis, an opening of the nanosheets in MoS<sub>2</sub>-400 can be seen in Fig. 4(f) as the interlayer distance increases from ~6.9 Å to ~7.5 Å compared with that observed for pristine MoS<sub>2</sub>. As discussed in XRD and FESEM analysis, the formation of MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> nanocomposites can be seen in Fig. 4(g) for the MoS<sub>2</sub>-N1 nanomaterial. In the case of MoS<sub>2</sub>-N5, such composite formation is not observed. The selected area electron

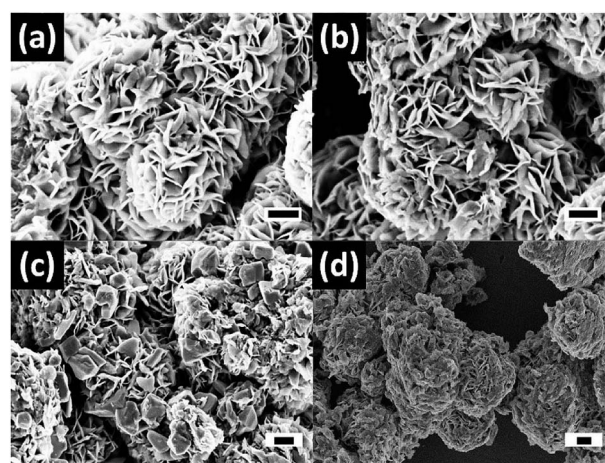


Fig. 3 FESEM images of low and high magnification for MoS<sub>2</sub> (a), MoS<sub>2</sub>-400 (b), MoS<sub>2</sub>-N1 (c), and MoS<sub>2</sub>-N5 (d). Scale bars 400 nm (black).





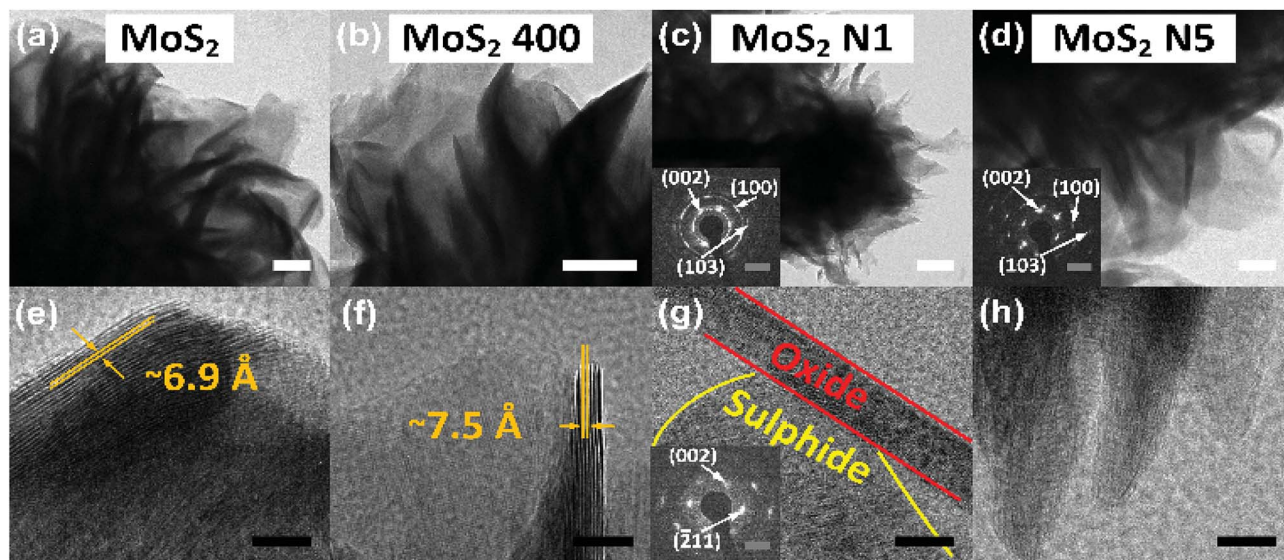


Fig. 4 TEM images for flower petals and stacked nanosheets for (a and b)  $\text{MoS}_2$ , (c and d)  $\text{MoS}_2$ -400, (e and f)  $\text{MoS}_2$ -N1, and (g and h)  $\text{MoS}_2$ -N5. Scale bars: 100 nm (white) and 5 nm (black). Insets show the SAED ring patterns for (e)  $\text{MoS}_2$ -N1 flower petal, (f)  $\text{MoS}_2$ -N1 rod/plates and (g)  $\text{MoS}_2$ -N5 flower. Inset scale bars,  $5 \text{ nm}^{-1}$  (grey).

diffraction (SAED) pattern confirms that the rod- and plate-like structure shows the oxide phase, whereas the layered flowers exhibit the sulphide phase of nanomaterials. The insets in Fig. 4(c and g) show the SAED patterns for flowers and rods of  $\text{MoS}_2$ -N1 nanomaterials, respectively, whereas the inset in Fig. 4(d) shows the SAED pattern for the flowers of  $\text{MoS}_2$ -N5. The SAED pattern demonstrates intense patterns for different crystallographic planes. The  $hkl$  planes (002), (100), and (103) show the sulphide phase for the SAED patterns of  $\text{MoS}_2$ -N1 and  $\text{MoS}_2$ -N5 and are marked in the insets of Fig. 4(c and d). The (002) and (211)  $hkl$  planes of the oxide phase of  $\text{MoS}_2$ -N1 are shown in Fig. 4(g). The  $hkl$  planes and their  $d$  spacing exactly match with those calculated from the XRD pattern. To confirm the formation of oxide/sulphide nanocomposites, elemental mapping is carried out, and the results are shown in the ESI (Fig. S2†). The elemental mapping clearly indicates the presence of oxide along with sulphides. Overall, the TEM analysis shows that the results are consistent with the XRD and FESEM results.

X-ray photoelectron spectroscopy (XPS) of the prepared  $\text{MoS}_2$  samples was conducted to check the elemental composition of the materials. Fig. 5 shows the XPS spectra for  $\text{MoS}_2$  (a),  $\text{MoS}_2$ -400 (b),  $\text{MoS}_2$ -N1 (c) and  $\text{MoS}_2$ -N5 (d) for binding energies ranging from 394 eV to 408 eV. The N 1s peak around 399.1 eV confirms that nitrogen is present in  $\text{MoS}_2$ -N1 and  $\text{MoS}_2$ -N5 nanomaterials. Mo  $3p_{3/2}$  can be observed at 395.5 eV in all synthesised materials. The XPS spectra for the Mo 3d peak and S 2p peak and survey scans of the prepared  $\text{MoS}_2$  nanomaterials are provided in the ESI (Fig. S3†). The sulphur (S) to molybdenum (Mo) atomic percentage ratios for  $\text{MoS}_2$ -N1 and  $\text{MoS}_2$ -N5 are found to be 1.16 and 1.35, respectively. The deviation from the ideal stoichiometric ratio (S : Mo) indicates the amounts of oxygen and nitrogen present in  $\text{MoS}_2$ -N1 and  $\text{MoS}_2$ -N5 nanomaterials. The nitrogen-to-sulphur atomic percentage ratios are found to be 1.83 and 2.19 for  $\text{MoS}_2$ -N1 and  $\text{MoS}_2$ -N5, respectively. The ratio

confirms that the amount of nitrogen present in the  $\text{MoS}_2$ -N5 nanomaterial is higher than that in  $\text{MoS}_2$ -N1 nanomaterials.

The electrochemical performances of prepared  $\text{MoS}_2$  nanostructures are studied by cyclic voltammetry (CV) and galvanostatic charge–discharge (GCD) curves using a three-electrode system. The CV curves of the prepared  $\text{MoS}_2$  nanomaterials were recorded in 1 M  $\text{H}_2\text{SO}_4$  electrolyte for scan rates of 5, 20, 100, and 200  $\text{mV s}^{-1}$ . The nearly rectangular and symmetric shape of the CV curves can be seen in Fig. 6(a) for as-synthesised  $\text{MoS}_2$ . In Fig. 6(a), the oxidation and reduction peaks for Mo metal centre can be seen at lower step sizes (5  $\text{mV s}^{-1}$  and 20  $\text{mV s}^{-1}$ ). These peaks lead to pseudocapacitance in the  $\text{MoS}_2$  materials. The non-regular rectangular CV curve shows electrochemical double layer capacitance in the synthesised  $\text{MoS}_2$  material. Fig. 6(b) shows a comparison of the CV curves at

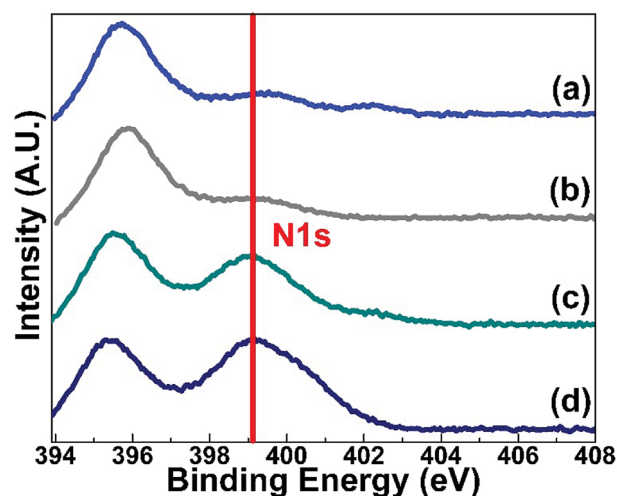


Fig. 5 XPS spectra for  $\text{MoS}_2$  (a),  $\text{MoS}_2$ -400 (b),  $\text{MoS}_2$ -N1 (c) and  $\text{MoS}_2$ -N5 (d) for binding energies ranging from 392 eV to 409 eV.



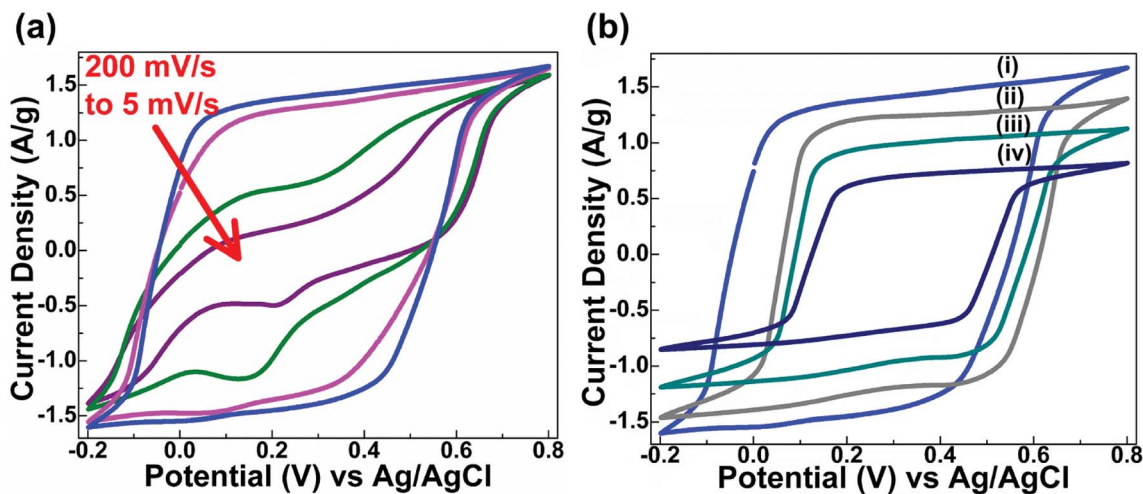


Fig. 6 (a) Cyclic voltammetry (CV) curves for as-synthesised MoS<sub>2</sub> nanomaterial for scan rates of 5 mV s<sup>-1</sup>, 20 mV s<sup>-1</sup>, 100 mV s<sup>-1</sup>, and 200 mV s<sup>-1</sup> (b) CV curves for (i) MoS<sub>2</sub>, (ii) MoS<sub>2</sub>-400, (iii) MoS<sub>2</sub>-N1, and (iv) MoS<sub>2</sub>-N5 at scan rate of 200 mV s<sup>-1</sup>.

200 mV s<sup>-1</sup> for MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1, and MoS<sub>2</sub>-N5. Furthermore, the capacitive performances of prepared MoS<sub>2</sub> nanostructures are evaluated by galvanostatic charge–discharge curves measured at different current densities (2, 3, 4, and 5 A g<sup>-1</sup>) in 1 M H<sub>2</sub>SO<sub>4</sub> electrolyte. The specific capacitance ( $C_s$ ) values of the prepared MoS<sub>2</sub> nanomaterials are calculated from the GCD curves using the following equation:

$$C_s = (I \times \Delta t) / (m \times \Delta V)$$

Here,  $I$  is the discharge current (A),  $\Delta t$  is the discharge time (s),  $m$  is mass of active material (g), and  $\Delta V$  is the potential window (V) for discharge time. The calculated values of specific capacitance are higher for lower current densities and subsequently decrease for higher current densities. This is a common observation in electrode materials due to charge resistive behaviour at higher current values. The specific capacitance values are higher for MoS<sub>2</sub>-N1 in comparison with those for other prepared MoS<sub>2</sub> materials from the GCD curves. The specific capacitance values for MoS<sub>2</sub>-N1 are 129, 99, 88, and 76 F g<sup>-1</sup> in 1 M H<sub>2</sub>SO<sub>4</sub> electrolyte at specific currents of 2, 3, 4, and 5 A g<sup>-1</sup>, respectively. The specific capacitance values at 2 A g<sup>-1</sup> for MoS<sub>2</sub>, MoS<sub>2</sub>-400, and MoS<sub>2</sub>-N5 are 34.9, 35.6, and 74.4 F g<sup>-1</sup>, respectively. Fig. 7 shows the comparison between the calculated specific capacitance values at different specific currents for MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1, and MoS<sub>2</sub>-N5.

The electrochemical results show the enhanced specific capacitance for MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5 as compared with that for MoS<sub>2</sub>. This confirms that nitrogen doping increases the capacitance of MoS<sub>2</sub> material. MoS<sub>2</sub>-N1 shows almost 3.5 times improvement in capacitance value relative to MoS<sub>2</sub>. MoS<sub>2</sub>-N1 has higher capacitance than MoS<sub>2</sub>-N5 due to the presence of N-doped MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> composite material. The main reason behind the increase in specific capacitance can be correlated with the structure of the prepared nanomaterials, the composition and their electronic conductivities.

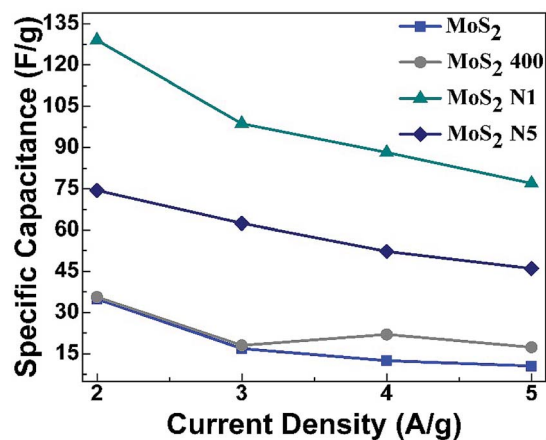


Fig. 7 The comparison between specific capacitances (F g<sup>-1</sup>) of MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1, and MoS<sub>2</sub>-N5 vs. specific current (A g<sup>-1</sup>).

Furthermore, electrochemical impedance spectroscopy (EIS) is performed to understand the charge kinetics and supercapacitor behaviour of prepared MoS<sub>2</sub> nanomaterials. Fig. 8(a) shows the Nyquist plots of MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5. The prepared MoS<sub>2</sub> nanostructures show the supercapacitor feature, *i.e.*, a semicircle at higher frequencies and a straight line at the lower frequency region. The intercept on the real axis is ascribed to the equivalent series resistance ( $R_s$ ), and the diameter of the semicircle corresponds to the charge transfer resistance ( $R_{ct}$ ) in the electrode/electrolyte system. The  $R_s$  values for MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5 are 2.16, 2.63, 1.89 and 1.16  $\Omega$ , respectively. Nitrogen doping decreases the  $R_s$  values for MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5. This confirms that nitrogen doping has enhanced the electrical conductivity of MoS<sub>2</sub>. The oxide/sulphide composite in MoS<sub>2</sub>-N1 may be responsible for the higher  $R_{ct}$  value (2.2  $\Omega$ ) in the Nyquist plot compared to that for other MoS<sub>2</sub> nanomaterials (<1.5  $\Omega$ ). The low frequency straight line parallel to the imaginary axis shows the capacitive behaviour of MoS<sub>2</sub> nanomaterials (Fig. 8(a)). The highest slope



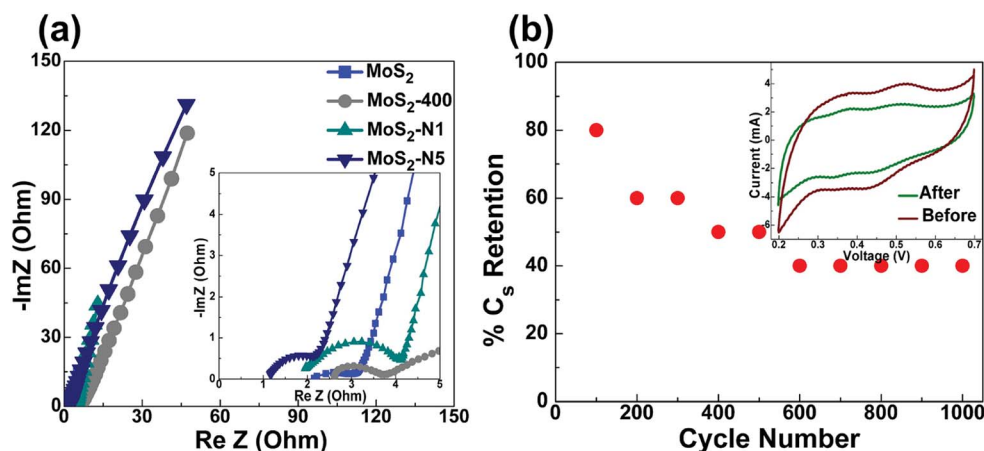


Fig. 8 (a) Nyquist Plots of MoS<sub>2</sub>, MoS<sub>2</sub>-400, MoS<sub>2</sub>-N1 and MoS<sub>2</sub>-N5; (inset shows the high frequency region of (a)). (b) % C<sub>s</sub> retention vs. cycle number for MoS<sub>2</sub>-N1 electrode; (inset shows CV at 100 mV scan rate before and after 1000 cycles).

Table 1 The comparison of specific capacitance of reported MoS<sub>2</sub> only materials at specific current and electrolyte. (The contents in the table are arranged in decreasing order of specific current for comparison.)

| Sr. no. | Method                  | Phase | Morphology  | Specific capacitance (F g <sup>-1</sup> )   | Reference |
|---------|-------------------------|-------|-------------|---|-----------|
| 1       | Hydrothermal            | 2-H   | Flowers     | 129 at 2 A g <sup>-1</sup> in 1 M H <sub>2</sub> SO <sub>4</sub>                      | This work |
| 2       | Two step hydrothermal   | 2-H   | 3D flower   | 168 at 1 A g <sup>-1</sup> in 1 M KCl   | 34        |
| 3       | DC magnetron sputtering | 2-H   | Nanoworms   | 138 F g <sup>-1</sup> at 1 A g <sup>-1</sup> in 1 M Na <sub>2</sub> SO <sub>4</sub>   | 35        |
| 4       | Hydrothermal            | 2-H   | Flower-like | 129.2 F g <sup>-1</sup> at 1 A g <sup>-1</sup> in 1 M Na <sub>2</sub> SO <sub>4</sub> | 36        |
| 5       | CVD                     | 2-H   | Nanowall    | 100 F g <sup>-1</sup> at 1 mV s <sup>-1</sup> in 0.5 M H <sub>2</sub> SO <sub>4</sub> | 37        |
| 6       | Hydrothermal            | 2-H   | Nanosphere  | 122 at 0.5 A g <sup>-1</sup> in 1 M Na <sub>2</sub> SO <sub>4</sub>                   | 18        |

of the straight line is obtained for MoS<sub>2</sub>-N1, suggesting that MoS<sub>2</sub>-N1 has optimal performance among the prepared MoS<sub>2</sub> nanostructures. The EIS measurements also support that nitrogen doping enhances the conductive performance of prepared MoS<sub>2</sub> nanostructures.

The cycling stability of MoS<sub>2</sub>-N1 material is studied by cyclic charge–discharge at 1 A g<sup>-1</sup> current density in 1 M H<sub>2</sub>SO<sub>4</sub> electrolyte. Fig. 8(b) shows the stability of the prepared MoS<sub>2</sub>-N1 electrode over 1000 cycles. The CV measurements at 100 mV s<sup>-1</sup> scan rate are also recorded before and after 1000 cycles. The CV analysis shows almost 70% capacitance retention after 1000 cycles. The charge–discharge cycles show 40% retention in specific capacitance for MoS<sub>2</sub>-N1 after 1000 cycles.

The synthesised MoS<sub>2</sub>-N1 shows similar crystallographic phase and structure to that of reported MoS<sub>2</sub> nanostructures. The comparison of specific capacitances of prepared MoS<sub>2</sub>-N1 and reported MoS<sub>2</sub> nanostructures at a specific current and electrolyte is depicted in Table 1. In comparison with the reported MoS<sub>2</sub> nanostructures, the synthesised MoS<sub>2</sub>-N1 shows higher specific capacitance at a higher current density, *i.e.*, 129 F g<sup>-1</sup> at 2 A g<sup>-1</sup> in 1 M of H<sub>2</sub>SO<sub>4</sub>. Fig. 7 confirms that the specific capacitance decreases as the specific current increases. Thus, higher specific capacitance at high current density shows the superiority of MoS<sub>2</sub>-N1 over other reported materials. Therefore, MoS<sub>2</sub>-N1 is a better electrode material because it exhibits higher specific capacitance at higher specific currents compared to reported 2H-MoS<sub>2</sub> materials.

## Conclusion

The successful synthesis of nitrogen-doped MoS<sub>2</sub> hierarchical nanostructures was demonstrated by a simple low-cost synthesis method. The structural and morphological characterisation confirmed the formation of flower-like nanostructures of hexagonal MoS<sub>2</sub> along with the formation of oxide/sulphide nanocomposites. MoS<sub>2</sub>-N1 showed specific capacitance of 129 F g<sup>-1</sup> at a specific current of 2 A g<sup>-1</sup>. The formed nitrogen-doped MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> nanocomposites showed four-fold enhancement in specific capacitance than the as-synthesised MoS<sub>2</sub> nanostructures. These results suggest that nitrogen-doped MoO<sub>2</sub>/MoO<sub>3</sub>/MoS<sub>2</sub> nanocomposites can be promising electrode materials for supercapacitor applications.

## Conflicts of interest

There are no conflicts to declare.

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