

Cite this: *RSC Adv.*, 2018, 8, 31874

# Facile synthesis of size-controlled Fe<sub>2</sub>O<sub>3</sub> nanoparticle-decorated carbon nanotubes for highly sensitive H<sub>2</sub>S detection†

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Hydrogen sulfide (H<sub>2</sub>S) is one of the most plentiful toxic gases in a real-life and causes a collapse of the nervous system and a disturbance of the cellular respiration. Therefore, highly sensitive and selective H<sub>2</sub>S gas sensor systems are becoming increasingly important in environmental monitoring and safety. In this report, we suggest the facile synthesis method of the Fe<sub>2</sub>O<sub>3</sub> particles uniformly decorated on carbon nanotubes (Fe<sub>2</sub>O<sub>3</sub>@CNT) to detect H<sub>2</sub>S gas using oxidative co-polymerization (pyrrole and 3-carboxylated pyrrole) and heat treatment. The as prepared Fe<sub>2</sub>O<sub>3</sub>@CNT-based sensor electrode is highly sensitive (as low as 1 ppm), selective and stable to H<sub>2</sub>S gas at 25 °C, which shows promise for operating in medical diagnosis and environment monitoring. Excellent performance of the Fe<sub>2</sub>O<sub>3</sub>@CNT is due to the unique morphology of the nanocomposites made from uniformly dispersed Fe<sub>2</sub>O<sub>3</sub> nanoparticles on the carbon surface without aggregation.

Received 1st August 2018

Accepted 6th September 2018

DOI: 10.1039/c8ra06464d

rsc.li/rsc-advances

## 1. Introduction

The synthesis of innovative nanomaterials with advanced functions is a continuously expanding issue of materials science.<sup>1–5</sup> On this subject, nanocomposites, consisting of organic and inorganic components, present improved properties originating from each component and satisfy economical and environmental challenges of the industry.<sup>6–8</sup> Metal oxide-containing hybrid carbon nanomaterials are of great interest for diverse applications such as energy storage, catalysis, and sensor device systems.<sup>9–13</sup>

Recently, the concern over the environmental protection and the demand to control hazardous chemicals in industry have caused comprehensive interest in developing sensor systems for various toxic gases.<sup>14–18</sup> Hydrogen sulfide (H<sub>2</sub>S), originating from microbial breakdown of plants and animals, is one of the most abundant toxic gases found in coal mines, manholes and semiconducting device industries.<sup>19–23</sup> H<sub>2</sub>S is broadly hazardous in the human body causing collapse of the nervous system and disturbance of the cellular respiration due to binding iron in the mitochondrial cytochrome enzymes.<sup>24,25</sup> A threshold limit value

(TLV) and recommended acceptable ambient levels of H<sub>2</sub>S are, respectively, 10 ppm for 8 h and lower than 0.1 ppm by the Occupational Safety and Health Administration (OSHA).<sup>26</sup> Accordingly, highly sensitive (<10 ppm) and selective detection of H<sub>2</sub>S is emerging importance issue in the field of environmental monitoring, disease diagnosis, and food safety. However, conventional H<sub>2</sub>S gas sensor has several limitations such as slow response/recovery times and high operating temperature that cause complex sensing platform and lead to difficult placement in the confined area.<sup>27</sup>

Metal oxide semiconductors (MOS) have been conducted in developing H<sub>2</sub>S sensor systems due to low operating power, stability, and ease of incorporation into the microelectronic devices.<sup>28–32</sup> Among them, Fe<sub>2</sub>O<sub>3</sub> displays critical advantages (*i.e.* ease synthesis, low cost, nontoxicity and abundance in the earth crust) to apply into the H<sub>2</sub>S sensor system.<sup>33–36</sup> In spite of these benefits, Fe<sub>2</sub>O<sub>3</sub> has several limitations such as a long recovery time and a requirement of heating (200–400 °C) to promote sensing reaction to apply into the real sensing system. Therefore, there are a lot of studies about improving H<sub>2</sub>S detecting performance of the Fe<sub>2</sub>O<sub>3</sub>-based sensor systems through loading with noble metals, doping with catalytic oxides, and composing composites with carbon materials.<sup>37–39</sup> For example, Sun *et al.* fabricated  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> nanotubes on the carbon nanotube templates using a hydrothermal synthetic method and a following calcination.<sup>40</sup> Jiang *et al.* synthesized paper like Fe<sub>2</sub>O<sub>3</sub>/graphene nanosheets by means of a super critical CO<sub>2</sub>-assisted thermal process and a controlled magnetic field step.<sup>41</sup> However, it is difficult to control the size of these materials, particularly on the nanoscale; thus, material fabrication process can be very complicated.

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† Electronic supplementary information (ESI) available: Raman spectra of nanocomposites; sensing ability of nanocomposites; comparison H<sub>2</sub>S sensing performance of different chemical sensors; stability test to H<sub>2</sub>S gas. See DOI: 10.1039/c8ra06464d



Herein, we describe a facile synthesis method for Fe<sub>2</sub>O<sub>3</sub> particle uniformly decorated carbon nanotubes (Fe<sub>2</sub>O<sub>3</sub>@CNT) using functional group (–COOH) incorporated polypyrrole nanotubes. First, 3-carboxylated polypyrrole nanotubes (CPPyNT) were fabricated using oxidative co-polymerization between monomers (pyrrole and 3-carboxylated pyrrole) and initiator (iron cation (Fe<sup>3+</sup>) and methyl orange complex). Then, Fe<sub>2</sub>O<sub>3</sub> nanoparticles were formed onto the nanotube structure with phase change of polypyrrole to carbon through simple heat treatment. The Fe<sub>2</sub>O<sub>3</sub>@CNT was then applied as a hydrogen sulfide (H<sub>2</sub>S) gas sensor transducer. The Fe<sub>2</sub>O<sub>3</sub>@CNT electrode displays enhanced performances (such as rapid response/recovery time, superb cycle stability, low working temperature and highly sensitive to H<sub>2</sub>S molecule) than other conventional H<sub>2</sub>S sensors owing to uniformly dispersed Fe<sub>2</sub>O<sub>3</sub> nanoparticles on the carbon surface. In particular, a minimum detectable level (MDL) of the sensor system is as low as 1 ppm that is much lower than that of other sensors based on Fe<sub>2</sub>O<sub>3</sub> contained nanomaterials.<sup>38–42</sup>

## 2. Experimental section

### 2.1 Materials

Pyrrole (98%), methyl orange, and iron chloride (FeCl<sub>3</sub>, 97%) were obtained from Aldrich Co. Pyrrole-3-carboxylic acid was purchased from Acros organics and used without further purification.

### 2.2 Fabrication of Fe<sub>2</sub>O<sub>3</sub>@CNT nanocomposites

As a starting material, various carboxylated polypyrrole nanotubes were prepared using a self-degradation method. In detail, a 1.5 mmol of FeCl<sub>3</sub> solution was added into 30 ml of 5 mM methyl orange aqueous solution. Then, 2 mmol of pyrrole and pyrrole-3-carboxylic acid mixed monomers were added and stirred at 10 °C for 12 h. Especially the amount of the pyrrole-3-carboxylic acid in the mixed monomers was controlled from 0.2 wt% to 5.0 wt% to form different amount of Fe<sub>2</sub>O<sub>3</sub> on the carbon surface. The resulting black precipitate was purified by washing with distilled water several times until the filtrate became transparent. The as prepared black powder was then dried under vacuum at 40 °C for 12 h. Iron oxide (Fe<sub>2</sub>O<sub>3</sub>) particle decorated carbon nanotubes (Fe<sub>2</sub>O<sub>3</sub>@CNT) were prepared *via* heating process at 400 °C.

### 2.3 Characterization

A JEOL 6700 was used to obtain field-emission scanning electron microscope (FE-SEM) images. Transmission electron microscope (TEM) and high-resolution transmission electron microscope (HR-TEM) images were achieved with a JEOL JEM-200CX and JEOL-3010, respectively. X-ray diffraction (XRD) patterns and X-ray photoelectron spectroscopy (XPS) spectra were obtained using the M18XHF SRA (MAC Science Co.) and AXIS-HIS (KRATOS) systems.

### 2.4 Electrical measurements of the Fe<sub>2</sub>O<sub>3</sub>@CNT sensor electrode

Fe<sub>2</sub>O<sub>3</sub>@CNT (0.5 wt% in ethanol solution) was prepared by sonication for deposition onto an as-prepared interdigitated array (IDA) electrode. The as-prepared IDA electrode measured resistance changes in the composites with a source-meter connected to a computer. The Fe<sub>2</sub>O<sub>3</sub>@CNT sensors were placed in a vacuum chamber with a vapour inlet/outlet pressure of 10<sup>0</sup> torr. Various concentration of H<sub>2</sub>S gas (1–100 ppm) was injected into the chamber using a mass flow controller (MFC, KNH Instruments). The real-time responses from the Fe<sub>2</sub>O<sub>3</sub>@CNT were systematically evaluated by normalized resistance changes. The normalized resistance changes of the Fe<sub>2</sub>O<sub>3</sub>@CNT based sensor was monitored in a real-time during exposure to various gases at a constant applied current (10<sup>−6</sup> A) until saturation was reached. After the Fe<sub>2</sub>O<sub>3</sub>@CNT was exposed to various concentrations of H<sub>2</sub>S gas for several minutes, the gas vapour was then replaced by compressed nitrogen gas to remove any molecules attached to the nanomaterials. This process was repeated several times. Vapour/air was supplied at various flow rates ranging from 2 to 8 slm and 1 to 5 sccm, as controlled by the MFC.

## 3. Results and discussion

### 3.1 Fabrication of Fe<sub>2</sub>O<sub>3</sub>@CNT

Fig. 1a suggests the overall procedure for a synthesis of the Fe<sub>2</sub>O<sub>3</sub> particle decorated carbon nanotubes (Fe<sub>2</sub>O<sub>3</sub>@CNT), based on an oxidative co-polymerization process. As a first step material, 3-carboxylated polypyrrole coated nanorod complex was formed by a co-polymerization of monomers (pyrrole and 3-carboxylated pyrrole) and nanorod template (composed of methyl orange and iron chloride). Then, the nanorod template was washed using distilled water to generate 3-carboxylated pyrrole nanotubes (CPPyNT) of *ca.* 100 nm-diameter (Fig. 1b). In particular, iron cations (Fe<sup>3+</sup>) from the nanorod template are fastened on the nanotube surface because of the covalent bonding between the Fe<sup>3+</sup> ion and the negative charge of the O atom in the carboxyl group.<sup>43</sup> The CPPyNT was then conducted heat treatment to induce phase transfer of 3-carboxylated polypyrrole to carbon and iron cation (Fe<sup>3+</sup>) to iron oxide (Fe<sub>2</sub>O<sub>3</sub>), respectively. Consequently, Fe<sub>2</sub>O<sub>3</sub>@CNT was formed as shown in Fig. 1c.

During the heating procedure, a carboxyl group (–COOH) acts as a nucleation site to form Fe<sub>2</sub>O<sub>3</sub> nanoparticles on the carbon surface. For the without carboxyl group based composites, a small number of Fe<sub>2</sub>O<sub>3</sub> particles is randomly decorated on the carbon surface with different sizes from 4 nm to 11 nm (Fig. 2a). On the other hand, a size and a population of the Fe<sub>2</sub>O<sub>3</sub> particles are uniformly reduced and enhanced, respectively, with increasing concentration of the carboxyl group on the surface (from 0.2 wt% to 3.0 wt%) due to improve amount of nucleate sites (–COOH) and iron cations (Fe<sup>3+</sup>). The Fe<sub>2</sub>O<sub>3</sub>@CNT with 0.2 wt%, 1.0 wt%, and 3.0 wt% carboxyl group concentrations are denoted as Fe<sub>2</sub>O<sub>3</sub>@CNT\_0.2, Fe<sub>2</sub>O<sub>3</sub>@CNT\_1.0, and Fe<sub>2</sub>O<sub>3</sub>@CNT\_3.0. Fig. 2b–d display that an average diameter of





Fig. 1 (a) Illustrative of the sequential synthesis of the iron oxide decorated carbon nanotube composites. Field-emission scanning electron microscopy (FE-SEM) and transmission electron microscopy (TEM) (inset) images of (b) the carboxylated polypyrrole nanotubes (CPPyNT) and (c) the iron oxide decorated carbon nanotubes ( $\text{Fe}_2\text{O}_3\text{@CNT}$ ).

the particles decreases from 7 nm (0.2 wt%) to 4 nm (3.0 wt%) and a population of the particles dramatically increases up to 3.0 wt%. However, a diameter of the  $\text{Fe}_2\text{O}_3$  particles increases from more than 3.0 wt% concentration of a carboxyl group owing to a large amount of iron cations is caused aggregation of the  $\text{Fe}_2\text{O}_3$  during phase transfer process. Fig. 2e shows  $\text{Fe}_2\text{O}_3\text{@CNT}$  at 5.0 wt% of a carboxyl group (denoted as  $\text{Fe}_2\text{O}_3\text{@CNT}_{5.0}$ ) contained large diameter (more than 10 nm) of the particles on the surface. In addition, a high resolution transmission microscopy (HR-TEM) image of the particles indicates interlayer spacing of 0.25 nm and 0.27 nm for (110) and (104) of hematite  $\text{Fe}_2\text{O}_3$  and confirms that  $\alpha\text{-Fe}_2\text{O}_3$  phase is generated after a heat treatment step (Fig. 2f).

### 3.2 Characterization of $\text{Fe}_2\text{O}_3\text{@CNT}$

To confirm crystallinity and chemical composition of the nanocomposites, X-ray diffraction (XRD), Raman and X-ray photoelectron spectroscopy (XPS) are determined as follows. Fig. 3a suggests XRD patterns of the each synthesis state of the nanotubes (CPPyNT and  $\text{Fe}_2\text{O}_3\text{@CNT}$ ). The peaks of the

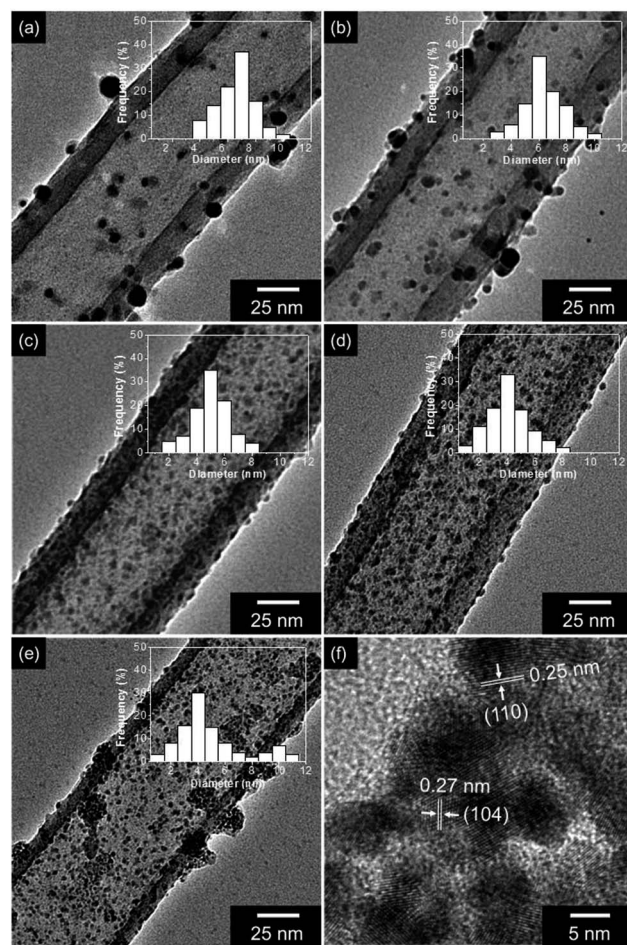


Fig. 2 TEM images and  $\text{Fe}_2\text{O}_3$  size distributions (inset) of the nanocomposites with different amount of carboxyl group on the polypyrrole surface: (a) without; (b) 0.2 wt%; (c) 1.0 wt%; (d) 3.0 wt%; (e) 5.0 wt%. (f) High-resolution transmission electron microscopy (HR-TEM) image of the  $\text{Fe}_2\text{O}_3\text{@CNT}$  with 3.0 wt% of carboxyl group.

inorganic materials in the  $\text{Fe}_2\text{O}_3\text{@CNT}$  are well matched  $\alpha\text{-Fe}_2\text{O}_3$  phase (hematite, JCPDS 33-0664), indicating the formation of  $\alpha\text{-Fe}_2\text{O}_3$  nanoparticles into the carbon structure. However, the broad peaks around  $25^\circ$  are difficult to illustrate phase transform of the CPPy to carbon structure. Therefore, Raman spectroscopy is used to confirm phase change of the CPPy after heat treatment (Fig. S1†). The ratio of D and G peaks ( $I_G/I_D$ ) is higher for the  $\text{Fe}_2\text{O}_3\text{@CNT}$  (*ca.* 1.3) than that for the CPPyNT (*ca.* 0.8), and peaks for  $\text{Fe}_2\text{O}_3\text{@CNT}$  are sharper than CPPyNT. Consequently CPPyNT is converted to amorphous carbon with  $\text{Fe}_2\text{O}_3$  nanoparticle generation through heating process.

Fig. 3b presents complete spectra over the range over of 0–1200 eV. These overview spectra illustrate that the C, N, O and Fe atoms are existed in the  $\text{Fe}_2\text{O}_3\text{@CNT}$ , whereas only C, N and O atoms are suggested in the CPPyNT. The peak of N 1s is attributed to nitrogen atoms from the pyrrole component. High-resolution XPS spectra for the C 1s around 285 eV are shown in Fig. 3c; this peak is deconvoluted into four components. The peaks of the CPPyNT are attributed as follows: the 284.3 eV peak





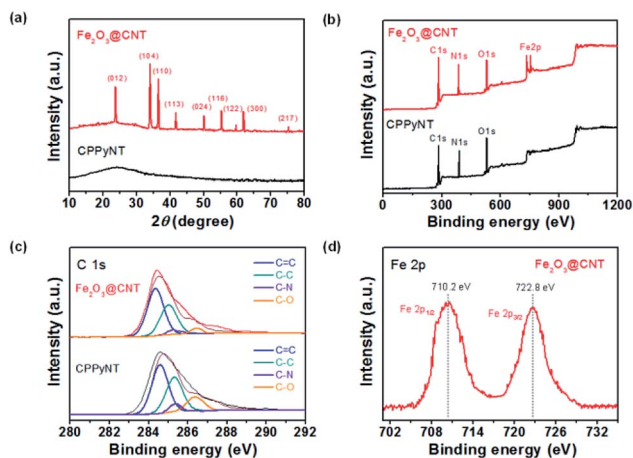


Fig. 3 (a) X-ray diffraction (XRD) spectra and (b) fully scanned X-ray photoelectron spectroscopy (XPS) of the nanotubes. (c) C 1s and (d) Fe 2p XPS spectra of the nanotubes. (black: CPPyNT, red: Fe<sub>2</sub>O<sub>3</sub>@CNT).

to C=C bonds, the 285.3 eV peak to C-C bonds, the 286.6 eV peak to C-O bonds, and the 284.9 eV peak to C-N bonds. But, Fe<sub>2</sub>O<sub>3</sub>@CNT suggests increase peak at 284.6 eV, which is attributed to graphitic sp<sup>2</sup> hybridization, and decrease peaks at 285.5 and 287.9 eV, assigned to C-O and C-N, after heat treatment. A high resolution spectrum for the Fe 2p of the Fe<sub>2</sub>O<sub>3</sub>@CNT is also suggested in Fig. 3d. In the Fe 2p spectrum, spin-orbit components of 2p<sub>3/2</sub> and 2p<sub>1/2</sub> are shown near 710.2 and 722.8 eV, indicating that the valence state of Fe is +3. Consequently, the final product (Fe<sub>2</sub>O<sub>3</sub>@CNT) is composed of Fe<sub>2</sub>O<sub>3</sub> nanoparticle decorated amorphous carbon nanotubes as confirmed by HR-TEM, XRD, and XPS.

### 3.3 Real-time responses for sensing of H<sub>2</sub>S

To characterize electrical properties of the Fe<sub>2</sub>O<sub>3</sub>@CNT-based sensor device, the composite nanotubes are immobilized on an interdigitated array (IDA) sensor electrode using spin-coating method. Fig. 4 displays linear current-voltage (*I*-*V*) curves of the various nanotubes for the voltage range from -0.1 to +0.1 V. A linearity of the *I*-*V* curves indicates that the nanotubes are uniformly immobilized on the sensor electrode with ohmic contact rather than Schottky barriers. In particular, *dI/dV* (=1/*R*) value of the electrodes decreases with an increase in the population of the Fe<sub>2</sub>O<sub>3</sub> particles on the carbon surface due to high resistive metal oxide semiconductor particles enhancing the resistivity of the nanotube composites.

The uniformly dispersed Fe<sub>2</sub>O<sub>3</sub> particles on the carbon surface rapidly detect H<sub>2</sub>S gas at room temperature by a chemical reaction between adsorbed oxygen species and H<sub>2</sub>S gas (Fig. 5). Sensing mechanism of the Fe<sub>2</sub>O<sub>3</sub>@CNT is illustrated as below. Initially, an electron depletion layer, composed of negatively charged adsorbing oxygen species (O<sup>2-</sup>, O<sup>-</sup>, and O<sub>2</sub><sup>-</sup>), is formed on the Fe<sub>2</sub>O<sub>3</sub> particle surface after exposing of the sensor electrode to air.<sup>39</sup> Chemical reactions of adsorbed oxygen species are suggested as follows:

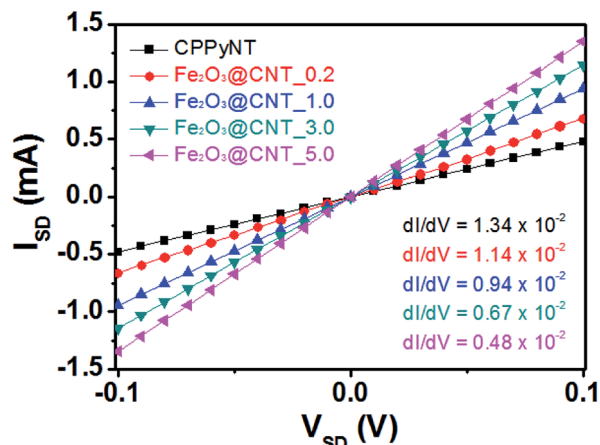


Fig. 4 *I*-*V* curves of the different nanotubes (black: CPPyNT; red: Fe<sub>2</sub>O<sub>3</sub>@CNT\_0.2; blue: Fe<sub>2</sub>O<sub>3</sub>@CNT\_1.0; green: Fe<sub>2</sub>O<sub>3</sub>@CNT\_3.0; pink: Fe<sub>2</sub>O<sub>3</sub>@CNT\_5.0).

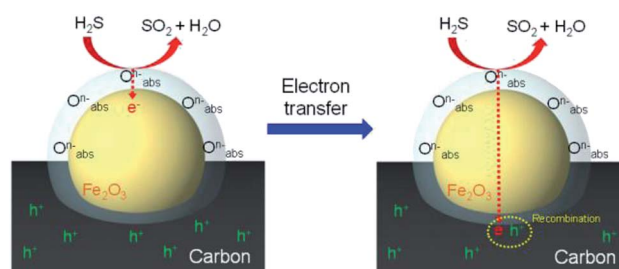
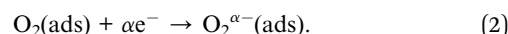
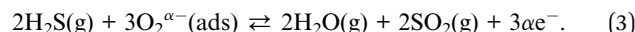


Fig. 5 H<sub>2</sub>S gas detection mechanism of the Fe<sub>2</sub>O<sub>3</sub>@CNT based-sensor electrode at room temperature.



When the sensor electrode is exposed to H<sub>2</sub>S gas, the chemisorbed oxygen species on the Fe<sub>2</sub>O<sub>3</sub> particles oxidize H<sub>2</sub>S gas. In detail, the oxygen species disassemble H<sub>2</sub>S gas into H<sub>2</sub>O and SO<sub>2</sub> gases with transferring electrons to the Fe<sub>2</sub>O<sub>3</sub> and the carbon nanotube structure.<sup>39</sup> Oxidation reaction of H<sub>2</sub>S gas on the Fe<sub>2</sub>O<sub>3</sub> particles is described by



A resistance of the sensor electrode increases with electron transfer due to decrease a population of charge carriers (hole) through recombination of the transferred electron and the hole in the carbon structure. Then, H<sub>2</sub>O and SO<sub>2</sub> gases from the oxidation of H<sub>2</sub>S are desorbed by exposure air contained gas (N<sub>2</sub> 80 vol% and O<sub>2</sub> 20 vol%) again. Consequently, the amount of chemisorbed oxygen species on the Fe<sub>2</sub>O<sub>3</sub> surface effects on the sensing ability of the sensor electrode. Thus, a large population of the small sized-Fe<sub>2</sub>O<sub>3</sub> provides an increase in specific surface area to H<sub>2</sub>S gas and efficient electron transfer to the carbon structure.

To investigate the sensing abilities of the Fe<sub>2</sub>O<sub>3</sub>@CNT-based sensor electrode, real-time responsive resistance changes are evaluated for different concentrations of H<sub>2</sub>S gas at room



temperature. Fig. 6a presents a real time response of the sensor electrodes with different amount of  $\text{Fe}_2\text{O}_3$  decoration as a function of  $\text{H}_2\text{S}$  concentration. Upon various concentrations of  $\text{H}_2\text{S}$  exposure, the  $\text{Fe}_2\text{O}_3/\text{CNT}$  based electrodes display a rapid resistance improvement less than 30 s before reaching a saturated value. A minimum detectable level (MDL) of the each electrode presents as following values: 20 ppm for the

$\text{Fe}_2\text{O}_3/\text{CNT}_{0.2}$ ; 5 ppm for the  $\text{Fe}_2\text{O}_3/\text{CNT}_{1.0}$ ; 1 ppm for the  $\text{Fe}_2\text{O}_3/\text{CNT}_{3.0}$ ; 1 ppm for  $\text{Fe}_2\text{O}_3/\text{CNT}_{5.0}$  (Fig. S2†). The MDL value of the  $\text{Fe}_2\text{O}_3/\text{CNT}$  based electrode suggests much lower than that of other conventional  $\text{Fe}_2\text{O}_3$  sensors at low working temperature (25 °C) (Table S1†). Thus, a better sensitive response is attained concomitant with a high density of the  $\text{Fe}_2\text{O}_3$  active site that causes enhancement of activity to  $\text{H}_2\text{S}$  gas. However, even though the  $\text{Fe}_2\text{O}_3/\text{CNT}_{5.0}$  contains more  $\text{Fe}_2\text{O}_3$  particles than the  $\text{Fe}_2\text{O}_3/\text{CNT}_{3.0}$ , a value of MDL and a sensitive response of the  $\text{Fe}_2\text{O}_3/\text{CNT}_{5.0}$  are lower than that of the  $\text{Fe}_2\text{O}_3/\text{CNT}_{3.0}$  because an excess amount of  $\text{Fe}_2\text{O}_3$  component generates aggregation rather than small particles and then reduces active surface area to  $\text{H}_2\text{S}$  gas. Moreover, the improved active surface area of the  $\text{Fe}_2\text{O}_3$  particles cause rapid diffusion of  $\text{H}_2\text{S}$  and increase sensitivity with providing more active sites (chemisorbed oxygen species) to detect  $\text{H}_2\text{S}$ .

Fig. 6b represents sensitivity changes of the electrode as a function of an amount of  $\text{Fe}_2\text{O}_3$  particles, with respect to  $\text{H}_2\text{S}$  concentration. The sensitivity ( $S$ ) is determined from the saturation point of the normalized resistance change after 30 s of  $\text{H}_2\text{S}$  exposure. As lower than 1 ppm, the sensor electrodes present nonlinear changes in sensitivity. On the other hand, linear tendency is observed over a wide range of  $\text{H}_2\text{S}$  concentration (from 1 ppm to 100 ppm). Accordingly, the  $\text{Fe}_2\text{O}_3/\text{CNT}$  based sensor electrodes illustrate reversible and reproducible responses to various concentration of the target analyte ( $\text{H}_2\text{S}$ ), and sensing ability is more pronounced as enhancing concentration of the analyte.

To apply practical application into the sensor systems, splendid cycle stability is desired for sensor transducer materials. Fig. 6c shows the electrical responses of the sensor electrodes upon periodic disclosure to 20 ppm of  $\text{H}_2\text{S}$  at room temperature. The different  $\text{Fe}_2\text{O}_3/\text{CNT}$  nanotubes show similar responses for an each time without retardation of the response and recovery times owing to small size of the  $\text{Fe}_2\text{O}_3$  nanoparticles transfer electrons quickly and generate the uniform electron depletion layer on the surface. Moreover, sensing ability of the  $\text{Fe}_2\text{O}_3/\text{CNT}$  sensor electrode also remains sensitivity more than 95% of initial value after 4 weeks that is higher than that of other conventional metal oxide based sensor electrode (Fig. S3†).<sup>19–23</sup> Therefore, the  $\text{Fe}_2\text{O}_3/\text{CNT}$  based sensor electrode suggests high stability to  $\text{H}_2\text{S}$  gas detection.

The selectivity of the sensor electrodes is also one of the important issues to apply them into practical application. In other words, it is essential to specific detect the target analyte among various chemicals. To evaluate selectivity,  $\text{Fe}_2\text{O}_3/\text{CNT}_{3.0}$  based electrode is exposed to different reducing ( $\text{H}_2\text{S}$  and  $\text{NH}_3$ ) and oxidizing ( $\text{NO}_2$ ,  $\text{MeOH}$  and  $\text{Et-OH}$ ) gases at concentrations of 20 ppm. As shown in Fig. 7, the sensor electrode displays much higher response to  $\text{H}_2\text{S}$  than that of other gases due to strong interaction between  $\text{H}_2\text{S}$  and the chemisorbed oxygen species and low bonding energy (381 kJ mol<sup>-1</sup>) of  $\text{H}_2\text{S}$ .<sup>44</sup> A faster release of trapped electrons in the  $\text{Fe}_2\text{O}_3$  particles is also one of important constituents of the selectivity. Thus,  $\text{H}_2\text{S}$  can be classified from other chemicals based on the extent and direction of the resistance changes upon analyte disclosures.

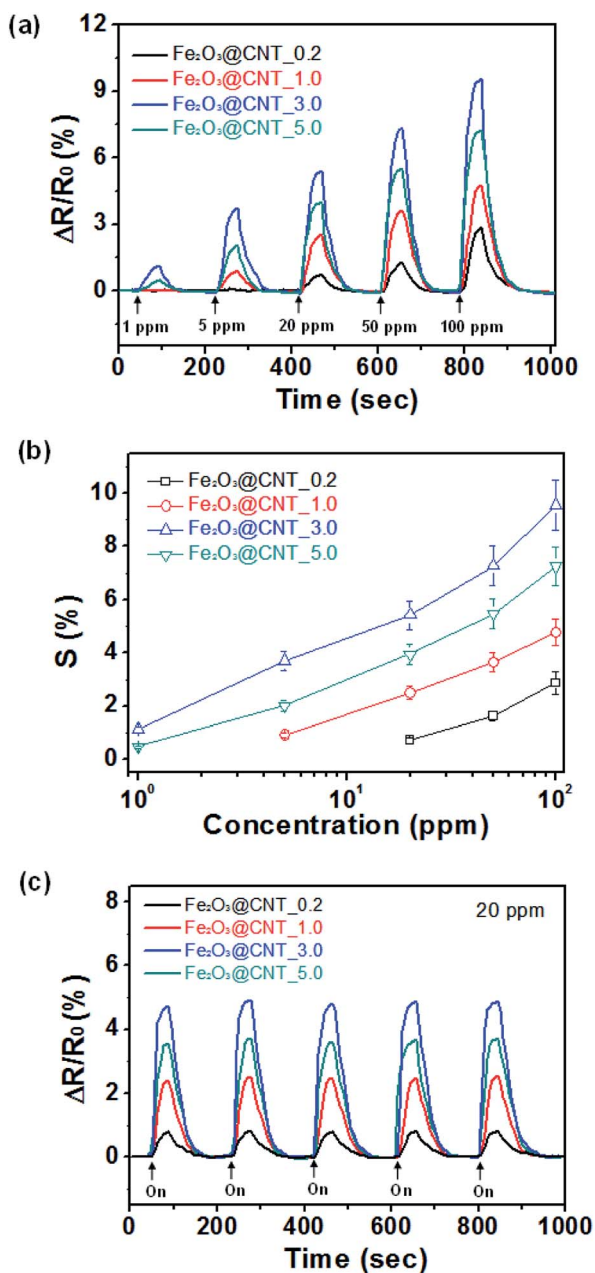


Fig. 6 Reversible and reproducible responses are measured at a current value ( $10^{-6}$  A) with the different nanocomposites. (a) Normalized resistance changes upon sequential exposure to various concentrations of  $\text{H}_2\text{S}$ . (b) Sensitivity ( $S$ ) calibration line as a function of  $\text{H}_2\text{S}$  concentration. (c) Normalized resistance changes of different composites upon sequential periodic exposure to 20 ppm of  $\text{H}_2\text{S}$  gas. (Each nanocomposite represents as following: black:  $\text{Fe}_2\text{O}_3/\text{CNT}_{0.2}$ ; red:  $\text{Fe}_2\text{O}_3/\text{CNT}_{1.0}$ ; blue:  $\text{Fe}_2\text{O}_3/\text{CNT}_{3.0}$ ; green:  $\text{Fe}_2\text{O}_3/\text{CNT}_{5.0}$ .)



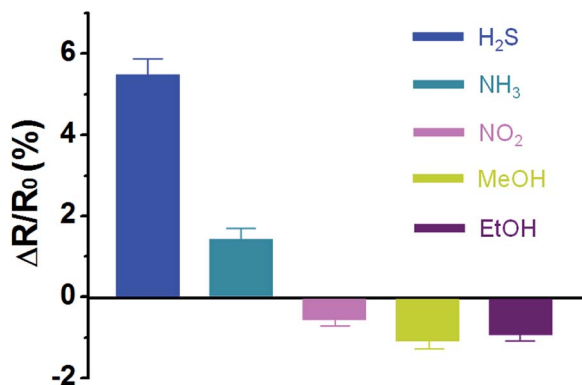


Fig. 7 Normalized resistance changes of Fe<sub>2</sub>O<sub>3</sub>@CNT\_3.0 based sensor electrode to different analytes: concentration of the chemicals is maintained at 20 ppm.

## 4. Conclusions

In summary, a size and an amount of the Fe<sub>2</sub>O<sub>3</sub> decoration controlled carbon nanotubes (Fe<sub>2</sub>O<sub>3</sub>@CNT) are facilely synthesized using co-polymerization of monomers (pyrrole and 3-carboxylated pyrrole) and following heat process. The Fe<sub>2</sub>O<sub>3</sub>@CNT is then applied in an ultra-highly sensitive H<sub>2</sub>S gas sensor electrode at room temperature. Especially, the uniformly decorated Fe<sub>2</sub>O<sub>3</sub> particles conduct critical roles in H<sub>2</sub>S sensing system as below: (1) formation of the chemisorbed oxygen species to cause reaction with H<sub>2</sub>S molecule; (2) maximize of the active surface area to the target analyte; (3) rapid electron transfer from H<sub>2</sub>S to the carbon nanotubes; (4) excellent cycle stability without morphology collapse. As a result, the Fe<sub>2</sub>O<sub>3</sub>@CNT sensor presents excellent cycle stability, selectivity, and highly sensitive (1 ppm of minimum detectable level) to H<sub>2</sub>S gas at room temperature that is much higher performative than other sensor electrodes based on Fe<sub>2</sub>O<sub>3</sub> included nanomaterials.<sup>38–42</sup> Therefore, this work suggests a facile synthesis of metal oxide/carbon nanocomposites for various applications such as energy storage electrode, electrochemical catalyst and sensor electrode.

## Conflicts of interest

There are no conflicts to declare.

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