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## Mesoporous anodic $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometer for organic vapor sensing application

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In this work, we reported the utilization of mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> films as optical sensors for detecting organic vapors. The mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> thin films, which exhibited obvious Fabry–Perot interference fringes in the reflectance spectrum, were successfully fabricated through electrochemical anodization of Fe foils. Through monitoring the optical thickness of the interference fringes, three typical organic species with different vapor pressures and polarities (hexane, acetone and isopropanol) were applied as probes to evaluate the sensitivity of the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> based interferometric sensor. The experiment results showed that the as-synthesized mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometer displayed high reversibility and stability for the three organic vapors, and were especially sensitive to isopropanol, with a detection limit of about 65 ppmv. Moreover, the photocatalytic properties of  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> under visible light are beneficial for degradation of dodecane vapor residues in the nano-pores and refreshment of the sensor, demonstrating good self-cleaning properties of the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>-based interferometric sensor.

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### 1. Introduction

Gas sensors have attracted considerable attention for their wide application in numerous areas such as air monitoring, gas (toxic, flammable and explosive) detection, medical diagnosis and food safety.<sup>1</sup> To date, much effort has been devoted to the design of small, low-cost and fast-response gas sensors with high sensitivity, selectivity and stability.<sup>2</sup> Such research interest has inspired a number of gas sensors based on different sensing mechanisms, such as electrical, thermal, acoustic or optical properties.<sup>3</sup> Among these devices, the label-free sensors based on reflectometric interference<sup>4,5</sup> have proved to be an attractive and promising alternative for detecting chemical and gas vapors.<sup>6–9</sup> Generally, the interferometric gas sensor is dependent on the changes in the thin film refractive index due to the adsorption of the gases in the pores, and the shift of the optical thickness  $nL$  (product of refractive index and film thickness) in the reflected spectrum.<sup>10</sup> In order to achieve porous thin films that are usually preferred as interferometric sensor owing to their unique structure for gas adsorption,<sup>11</sup> tremendous efforts have been dedicated to the fabrication of a variety of porous structures including porous silicon, aluminum oxide (Al<sub>2</sub>O<sub>3</sub>) and titanium dioxide (TiO<sub>2</sub>), which possess large surface area, and tunable surface chemistry.<sup>12–14</sup>

In addition, iron oxide (Fe<sub>2</sub>O<sub>3</sub>) with different kinds of structures is another platform for sensor application owing to its low cost, environmental friendliness and high stability.<sup>15–18</sup> However, the majority of the previously reported Fe<sub>2</sub>O<sub>3</sub> based sensor are mainly based on their electrical properties and usually operate at high temperature. Moreover, complicated nanostructures (*e.g.*, urchin-like  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> nanostructures,<sup>16</sup> core-shell  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>@SnO<sub>2</sub> hollow nanoparticles,<sup>17</sup> PbO loaded double-shell Fe<sub>2</sub>O<sub>3</sub> hollow nanospheres<sup>18</sup>) are commonly needed, which are usually subject to tedious fabrication procedures with relatively high cost. Compared with conventional nanofabrication method, electrochemical anodization is a powerful method owing to its simplicity, reliability, high controllability and low cost.<sup>19,20</sup> Recently, anodic iron oxide nanostructures obtained by directly anodizing iron (Fe) metal have aroused wide interest in lithium-ion battery,<sup>21,22</sup> supercapacitor,<sup>23</sup> water splitting,<sup>24</sup> and fuel cell applications<sup>25</sup> because of its excellent optical, electrochemical and photocatalytic properties. However, to our knowledge, anodic  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> for optical sensing of volatile gas vapors have not been reported. Herein, we report a mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometer that was synthesized by anodization of Fe in ethylene glycol solution. Through varying the anodization condition, the mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> nanostructures with well resolved Fabry–Perot fringes in the reflected spectrum were synthesized and further applied for detecting three organic vapors (hexane, acetone and isopropanol). The experiment results demonstrate that the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometer shows good sensitivity, reversibility and stability to the three organic vapors. Meanwhile, the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> is also found to exhibit efficient photo-catalytic activity for the

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degradation of the dodecane vapor residues in the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> nanopores under visible light irradiation, therefore providing a self-cleaning platform for the future sensor application.

## 2. Experimental

### 2.1 Fabrication of porous $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> film

The mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> structures were fabricated following a modified anodization process.<sup>26</sup> Firstly, the Fe foils (0.3 mm in thickness, 99.9% purity) were ultra-sonicated sequentially in acetone, ethanol and water. Then, the anodization was carried out in a two-electrode electrochemical cell, with the cleaned Fe foil and Pt as working anode and cathode, respectively (Fig. 1a). The Fe substrate was pressed together with an Al foil against an O-ring, defining a working area of 1.2 cm<sup>2</sup>. The composition of the electrolyte was 0.1 mol L<sup>-1</sup> NH<sub>4</sub>F (96%, Aladdin) and 1 mol L<sup>-1</sup> H<sub>2</sub>O in an ethylene glycol (99%, Macklin) solution. Finally, the anodization experiment was carried out at 40 V for 30, 50, and 80 min respectively. To improve the crystallinity of Fe<sub>2</sub>O<sub>3</sub>, the as-fabricated specimen was further annealed in Ar atmosphere at 450 °C for 2 h prior to sensing test.

### 2.2 Sensor design and testing method

The setup for organic vapor sensing shown in Fig. 1b consists of a mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> thin film mounted right below the transparent glass in a home-made Teflon flow cell. To measure the response of the sensor, the flow cell was assembled with an inlet connected to a glass bubbler. The saturated organic vapors were firstly generated through purging air into the glass bubbler and then diluted to different concentrations by a certain amount of pure air. The flow-rate of the organic vapor was

controlled by mass flow controllers. Tungsten light source focused through the microscope lens was put right above the flow cell and the reflection spectra of the sensor were collected through an Ocean Optics CCD spectrometer (USB 4000, Ocean optics, Inc.) from 400–1000 nm. By applying a fast Fourier transform (FFT) to the reflectance spectrum, the optical thickness of the film was calculated directly from the positions of the peaks in the FFT spectra. The FFT peaks were monitored in real-time to detect the optical thickness changes throughout the gas test. Note that the sensing experiment was all carried out at room temperature and the vapor concentrations were measured independently using a photoionization detector (ppbRAE, RAE Systems). Dodecane with a very low vapor pressure was used as a mode for the photocatalytic test. After the dosing and purging step, the contaminated optical sensor with dodecane was further irradiated by a 250 W metal halide lamp coupled with a UV cutoff filter ( $\lambda > 420$  nm) (Xujiang Machine-Electronic Plant, Nanjing) for about 20 minutes. After irradiation, the optical thickness of the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> was measured by the CCD spectrometer.

## 3. Results and discussion

The Fe anodization was performed under a positive potential which would form a mesoporous iron oxide layer over the Fe surfaces. Typically, the formation of anodic iron oxide in fluoride electrolytes is associated with the two competing processes: (1) the field assisted oxidation of iron to form iron oxide (eqn (1)); (2) the chemical dissolution of iron oxide by F<sup>-</sup> in the solution (eqn (2)).<sup>27,28</sup> Usually, the formation of iron oxide continues to take place until an equilibrium is reached between the two competing reactions, following which the nano-structured iron oxide ceases development.

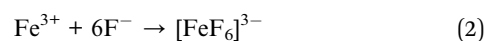
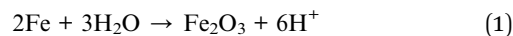


Fig. 2a and b show the SEM images of the anodic Fe<sub>2</sub>O<sub>3</sub> obtained at 30 min, displaying a porous structure with partially opened nanopores at the surface. With the anodic time increasing to 50 min, the top-view SEM images in Fig. 2c present fully open pores, while the cross-sectional SEM images exhibit vertically-oriented nanochannels from the top surface to the metal–oxide interfaces (Fig. 2d). The typical pore diameter is in the range of 30–50 nm, with an average value of 41 nm (Fig. 2h). When the anodization time was further increased to 80 min, the film thickness of the nanostructures keeps decreasing to about 4.5  $\mu\text{m}$  and start to crumble due to the over-etching, as shown in Fig. 2e and f. Moreover, the pore diameter distribution of the as-obtained Fe<sub>2</sub>O<sub>3</sub> is further investigated by top-view SEM analysis. It is observed that the diameters of the pores are in the range from 30 to 50 nm (Fig. 2g). It is noticed that compared with the anodization time at 30 min, the pore diameter anodized at 50 min increases from 23 nm to 41 nm. When the anodization time further increases to 60–80 min, the film thickness nearly keeps unchanged.

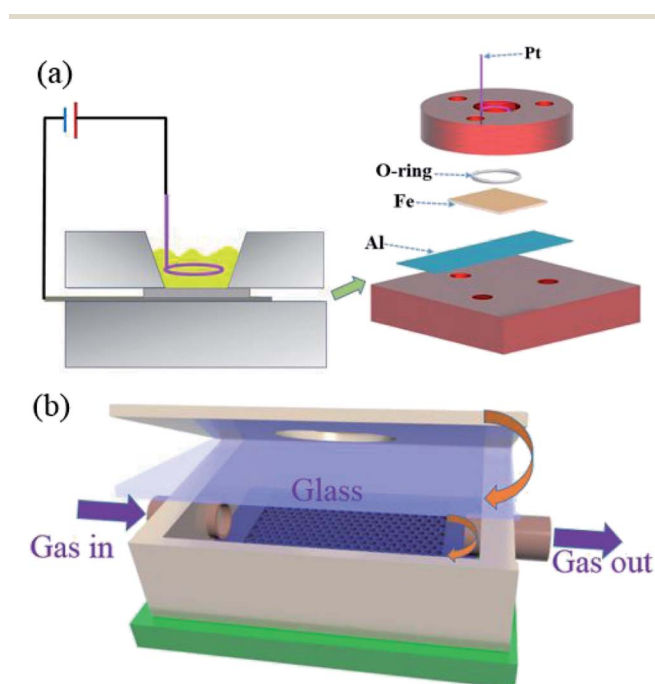


Fig. 1 (a) Schematic illustration of the setup for Fe foil anodization. (b) Schematic diagram of the flow-cell in the sensing experiment.



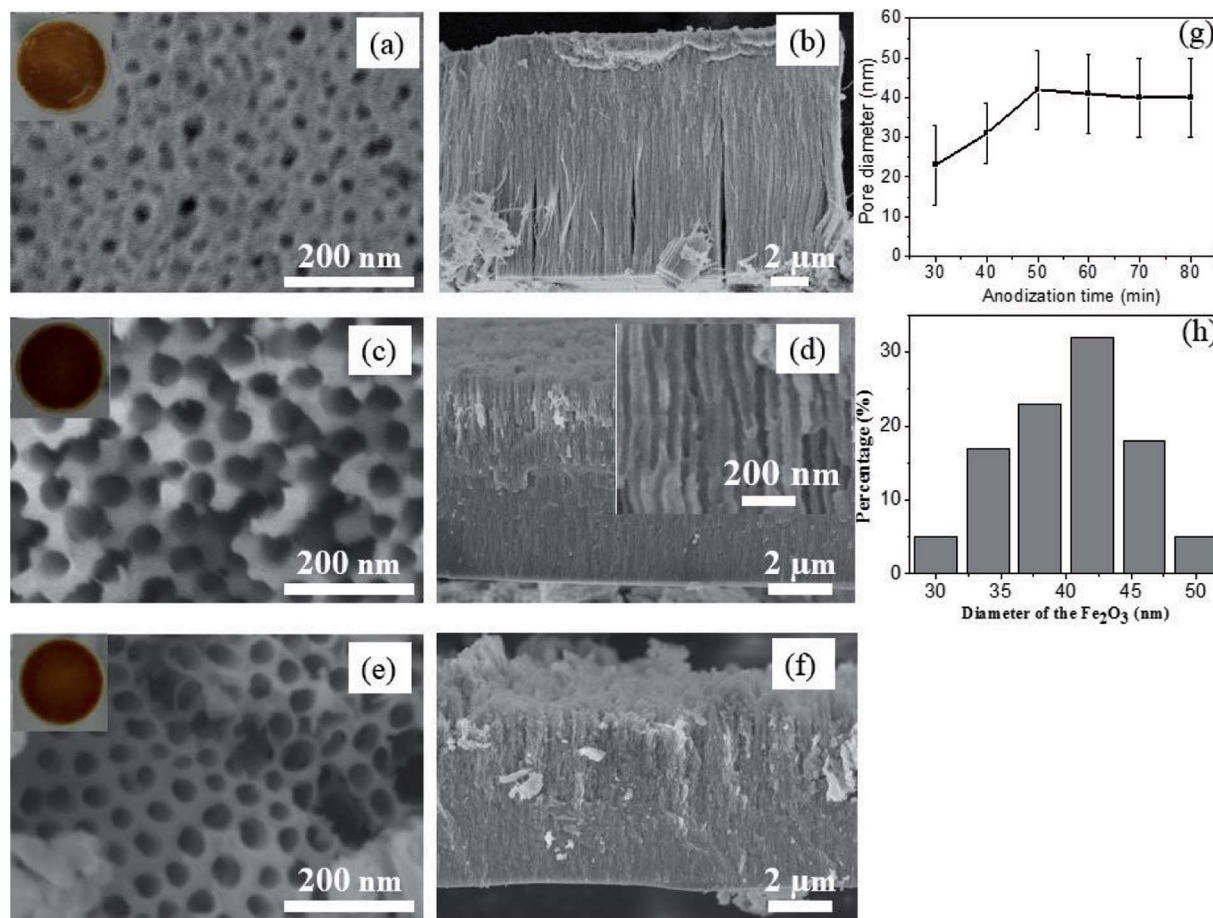


Fig. 2 The top view (a, c and e) and cross-sectional (b, d and f) SEM images of the anodic Fe<sub>2</sub>O<sub>3</sub> under anodization time of 30, 50, and 80 minutes, respectively. (g) Pore diameter and length of the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> anodized at different times. (h) The pore diameter distribution of the anodic  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> nanostructures with anodization time of 50 min.

Since the anodic Fe<sub>2</sub>O<sub>3</sub> at low temperature is amorphous, the above as-anodized sample was further crystallized at 450 °C in argon atmosphere for 2 h. Fig. 3a shows the XRD patterns of the anodized sample before and after annealing. Clearly, the as-anodized Fe<sub>2</sub>O<sub>3</sub> is amorphous with only a sharp Fe peak (dark line in Fig. 3a), but it will be converted into  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> (red line in Fig. 3a) after annealing at 450 °C. In addition, very small amount of Fe<sub>3</sub>O<sub>4</sub> is observed, which may probably arise from the incomplete oxidation of the as-obtained  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> in the reducing atmosphere.<sup>23,29</sup> What is more, XPS analysis was carried out for further investigation of the oxidation states of the annealed Fe oxides in argon atmosphere. As shown in XPS survey spectrum (Fig. 3b), there exist three main peaks at around 284.8, 530.2, and 711.0 eV, which are assigned to be the characteristic peaks of C 1s, O 1s and Fe 2p, respectively. From further high resolution XPS spectrum of Fe 2p (Fig. 3c), two peaks at around 711.3 eV and 724.7 eV corresponding to the Fe 2p<sub>3/2</sub> and Fe 2p<sub>1/2</sub> spin orbit peaks of  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> are observed. Moreover, the Fe 2p<sub>2/3</sub> and Fe 2p<sub>1/2</sub> are accompanied by satellite structures on their high binding-energy sides, at around 719.9 eV and 732.9 eV respectively, which can be due to hematite ( $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>).<sup>30</sup> The high resolution spectrum of oxygen region displays two overlapping peaks with binding energies of

530.15 eV and 531.73 eV (Fig. 3d). The higher peak at 530.15 eV is attributed to the O 1s bound to Fe in the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>, while the lower one at 531.73 eV corresponds to the oxygen in the organic impurities exist in the sample surface. These XPS results demonstrate that the main composition of the annealed Fe oxide is  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>, which is in accordance with the XRD results.

Next, the sensing mechanism of the mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> is plotted in Fig. 4a, which is based on white light interference in the thin film. When light crosses an interface where the refractive index of the medium is different, a fraction of light is reflected and the left is transmitted. As shown in Fig. 4a, the incident light is reflected from the air/porous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> and the porous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>/Fe substrate interfaces. As the two reflected waves overlap, they may interfere constructively and the Fabry-Perot fringes in a reflection spectrum are formed given that the extra path traversed by one wave is an integral number of wavelengths as described by the following equation:

$$m\lambda = 2nL \quad (3)$$

where  $m$  is an integer representing the spectral order,  $n$  is the mean refractive index of the porous layer and its contents, and  $L$  is the thickness of the porous layer. When the porous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometer is exposed to the organic vapor, the gas



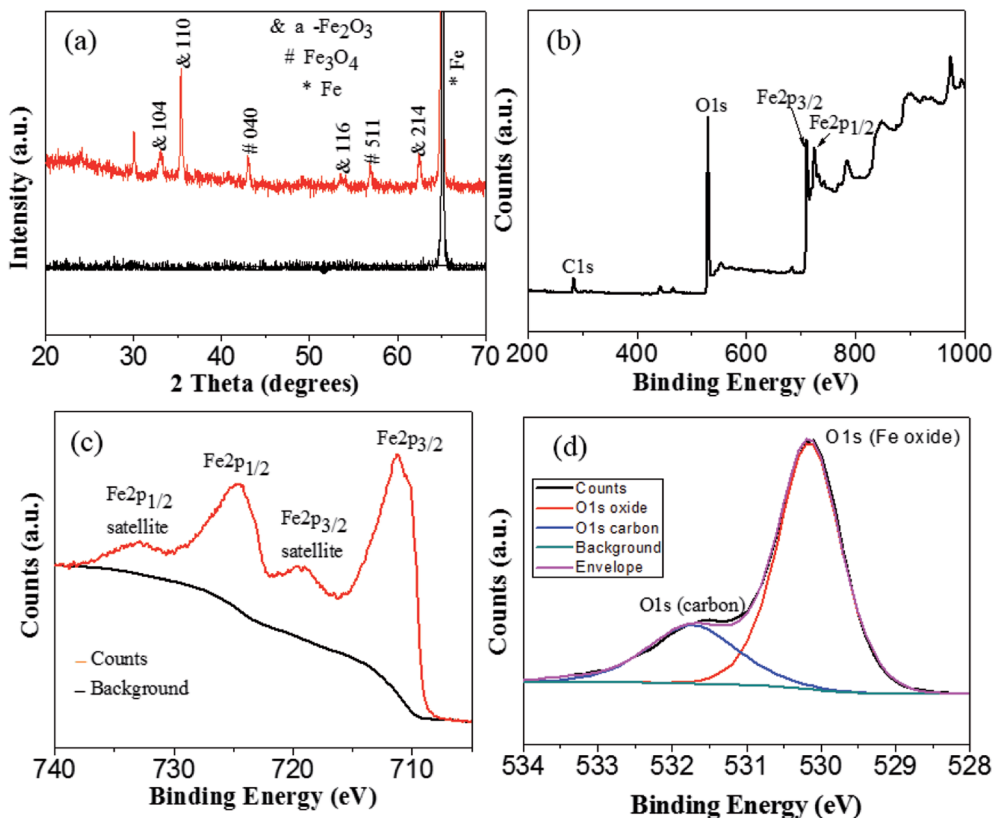


Fig. 3 (a) XRD patterns of the anodic  $\text{Fe}_2\text{O}_3$  (anodization time: 50 min) before and after annealing in argon atmosphere. (b) XPS survey spectrum of the annealed anodic  $\text{Fe}_2\text{O}_3$ . XPS spectrum of Fe 2p (c) and O 1s (d) in high resolution.

molecular could adsorb or condensate into the pores of the mesoporous film promoted by the chemical reaction or capillary condensation effect.<sup>11,14</sup> Organic vapor in the pores is

observed as a shift of the fringes to longer wavelengths, which is caused by the refractive indexes difference between the organic vapor and air. By applying a fast Fourier transform (FFT) to the

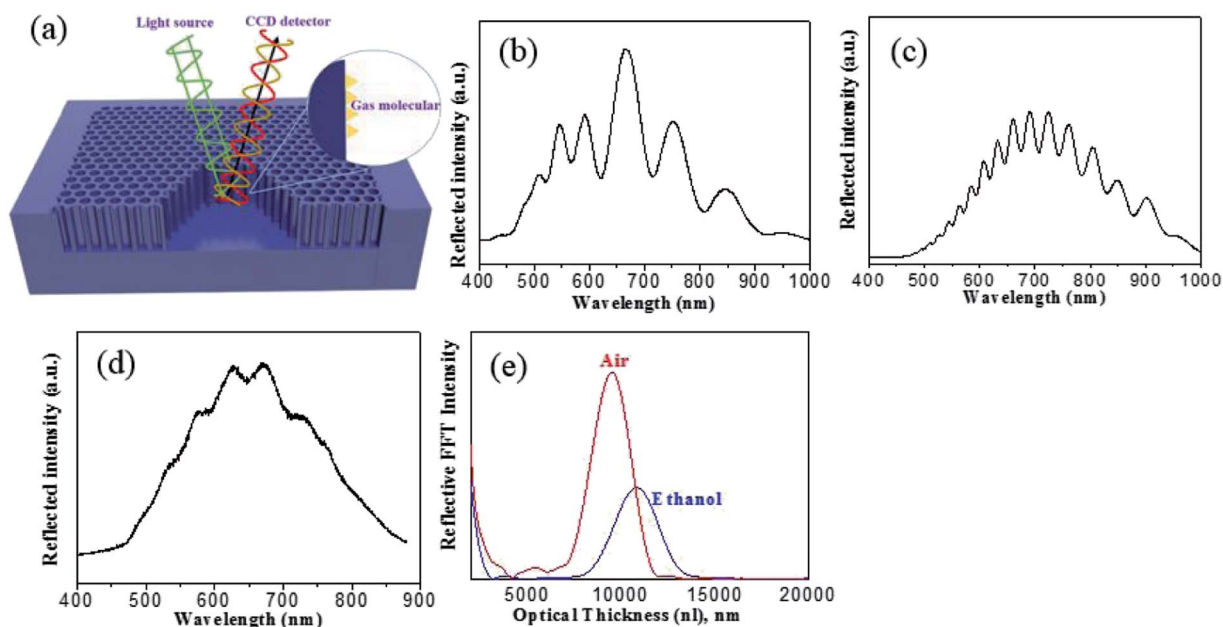


Fig. 4 (a) Schematic illustration of the  $\alpha\text{-Fe}_2\text{O}_3$  based interferometric vapor sensor. (b–d) Reflected spectrum of the  $\alpha\text{-Fe}_2\text{O}_3$  with anodization time of 30 min, 50 min and 80 min. (e) Reflective interferometric Fourier transform spectroscopy in air and ethanol of the  $\alpha\text{-Fe}_2\text{O}_3$  (anodization time: 50 min).



Table 1 Physical properties at room temperature (298 K)<sup>33</sup>

	Refractive index	Saturation vapor pressure (torr)
Acetone	1.359	233.3
Hexane	1.375	124
Isopropanol	1.377	45
Dodecane	1.422	0.134

as-obtained reflective spectrum which is referred as reflective interferometric fourier transform spectroscopy (RIFTS),<sup>31,32</sup> the detection of the vapor is achieved through the analysis of the optical thickness ( $nL$ ) calculated by the RIFTS.

The reflectance spectra were measured at normal incidence using a bifurcated optic fiber, with one end connected to a tungsten lamp while the other to a CCD spectrometer as previously reported.<sup>14</sup> The tungsten light source was focused by a microscope objective lens onto the surface of the film and the corresponding reflected light from the film was collected by the same microscope and recorded by the CCD detector. Fig. 4b–d compares the reflection spectra with Fabry–Perot fringes obtained from anodic  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> with different anodization time. It is seen that the mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> with smooth and fully opened pores anodized at 50 min (Fig. 4c) displays the best interference fringes in comparison with other two samples obtained at 30 min (partially opened pores) and 80 min (partially destroyed structure). In light of this, the anodic  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> from

50 min anodization are chosen for the following sensing measurement. Through immersing the  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> thin film into ethanol, the optical thickness and porosity could be determined by the spectroscopic liquid infiltration method (SLIM).<sup>32</sup> In this method, the reflection spectrum of  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> was collected in both air and ethanol at the same spot. By applying FFT to the reflectance spectrum, the FFT intensity *versus* optical thickness could be obtained in Fig. 4e, from which an apparent red shift of optical thickness is observed upon submerging the film in ethanol. The change of optical thickness arises from the difference between the reflective indexes of air ( $n = 1$ ) and ethanol ( $n = 1.3611$ ), and is highly dependent on the porosity and film thickness of the materials. Therefore through a least-squares fit of the two optical thicknesses (in air and ethanol) to a two component Bruggeman effective medium approximation, the porosity and film thickness of  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometer are finally obtained to be  $38.3 \pm 1.8\%$  and  $6230 \pm 56$  nm, respectively, which agree reasonably well with the value measured from SEM study.

In order to evaluate the sensing ability of the mesoporous  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub> interferometers for volatile organic vapors, three organics (hexane, acetone, and isopropanol (IPA)) with different refractive index and vapor pressure (Table 1) were tested.<sup>33</sup> It is worth noting that even though the three organic vapors have nearly close reflective index, yet, their vapor pressures are completely different, which also plays an important role in determining the sensitivity. During the vapor sensing process, due to the capillary condensation effect,<sup>11</sup> gas vapor with a lower vapor pressure

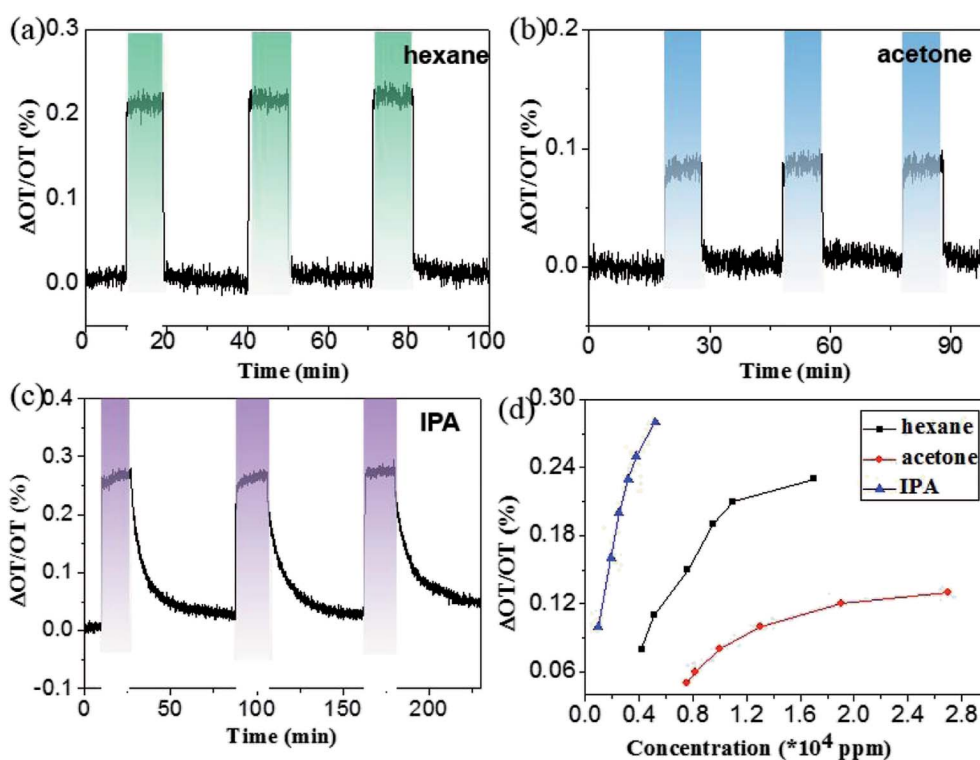


Fig. 5 Dose–response curves of the Fe<sub>2</sub>O<sub>3</sub> interferometer to hexane (a), acetone (b) and isopropanol (c) respectively. (d) Response curves for hexane, acetone and IPA at different concentrations. The quantity  $\Delta OT/OT$  (%) is equal to  $(nL_{\text{final}} - nL_{\text{initial}})/nL_{\text{initial}} \times 100\%$ ,  $nL_{\text{final}}$  and  $nL_{\text{initial}}$  represent the optical thickness at time 0 and  $T$ , respectively.



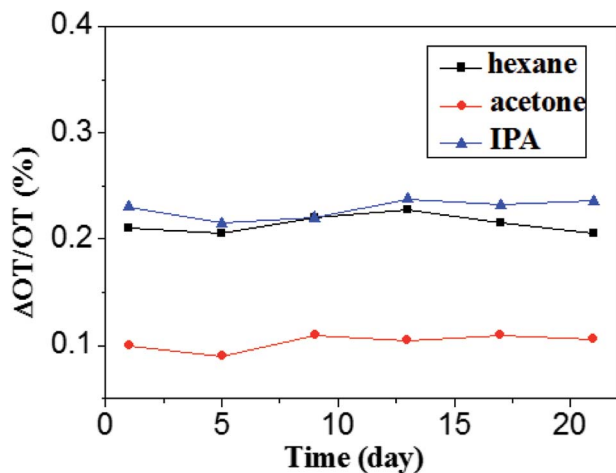


Fig. 6 Long-term stability of the  $\text{Fe}_2\text{O}_3$  interferometer to hexane (11 000 ppm), acetone (13 000 ppm) and isopropanol (2000 ppm) respectively.

will induce more condensed vapors in the inner pores<sup>14</sup> and should translate to larger changes in optical thickness ( $nL$ ) upon exposure to the analytes.

Fig. 5a–c shows the dose–response curves for hexane, acetone and IPA, respectively, from which one can readily find that our device displays different sensitivity to the three gases. As expected, it has the most sensitive response to IPA with the lowest vapor pressure, only 2800 ppmv will generate about 0.25% change in optical thickness. For hexane, 11 000 ppmv is able to induce about 0.2% change in optical thickness. The acetone is the most insensitive, a concentration of 13 000 ppmv only results in about 0.1% change in optical thickness. Further observation of the curves after the dosing found that the signal dropped quickly to the original baseline for hexane and acetone, while a purging time of about 1 h was needed for IPA to return to its original baseline level. The difficulty for desorption of IPA in the mesoporous film is probably related to either the lowest vapor pressure or the hydrophilic properties of  $\text{Fe}_2\text{O}_3$  and IPA that may cause the formation of hydrogen bonds between

them.<sup>14</sup> Fig. 5d plots the relationship between optical thickness change and concentrations for three organic vapors. Obviously, to generate the same  $\Delta\text{OT}/\text{OT}$ , acetone needs the largest concentration, while IPA the least. To evaluate the detection limit of the  $\alpha\text{-Fe}_2\text{O}_3$  based interferometer, an S/N ratio of 2 was used. The detection limits of the present porous  $\text{Fe}_2\text{O}_3$  interferometer for hexane, acetone and IPA were determined to be 1100, 3600 and 65 ppmv, respectively. The sensitivity for IPA is comparable to other well-established interferometer-based optical sensors, e.g., porous silicon and anodic  $\text{TiO}_2$ .<sup>14,34</sup> The long-term stability of the  $\alpha\text{-Fe}_2\text{O}_3$  based interferometer was also measured and shown in Fig. 6. It can be easily seen that the sensor exhibits good stability even the sample was stored in ambient condition for 3 weeks.

As often in other gas sensors, the present  $\alpha\text{-Fe}_2\text{O}_3$  interferometer also suffer from contamination due to gas adsorption after repeated sensing, which constitutes main bottleneck for further application.<sup>34,35</sup> Fig. 7a shows the dose–response curve to dodecane with a low vapor pressure of 0.134 torr. A remarkable increase in optical thickness was immediately observed after dodecane vapor introduced into the nonporous  $\text{Fe}_2\text{O}_3$  chamber. However, due to the strong tendency of condensation, it is difficult to remove the dodecane even by extended air purging. Considering the fact that the  $\alpha\text{-Fe}_2\text{O}_3$  is a typical semiconductor with a bandgap of 2.0 eV and it has proved to be a good photo-catalyst for RhB and other organic chemicals degradation under visible light illumination,<sup>36,37</sup> we tried to refresh the sensing ability by shining the porous  $\alpha\text{-Fe}_2\text{O}_3$  interferometer with visible light. Fig. 7b depicts the optical thickness shift of 20 cycles filling and visible light cleaning. Notably, when the above contaminated sample after air purging were further illuminated under a 250 W metal halide lamp (with the UV cutoff) for about 20 minutes, the optical thickness was substantially decreased and virtually recovered to its original value. This means our optical sensor is almost recovered to its original state after visible light refreshment, demonstrating good self-cleaning reversibility and repeatability of the  $\alpha\text{-Fe}_2\text{O}_3$  based interferometer. The self-cleaning mechanism of the  $\alpha\text{-}$

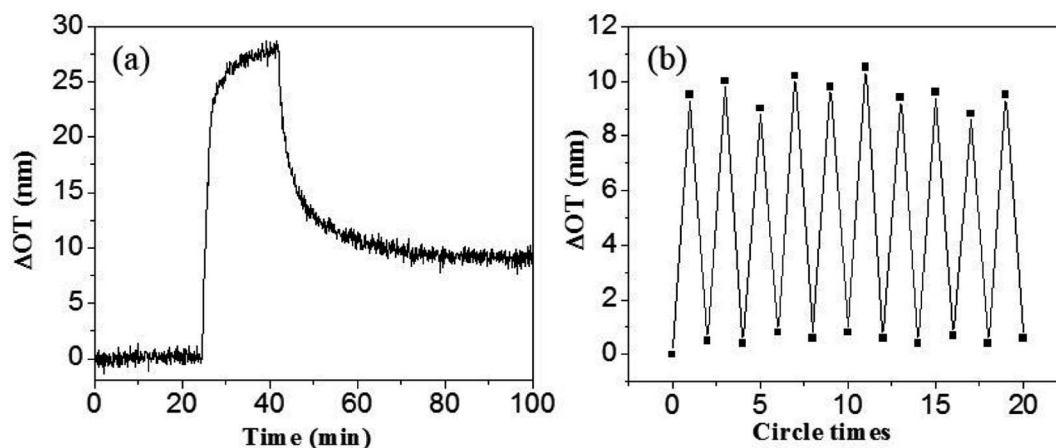
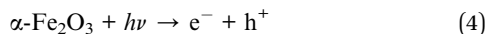


Fig. 7 (a) Dose–response curves of the  $\text{Fe}_2\text{O}_3$  interferometer to dodecane. (b) Optical thickness changes during the 10 cycles filling and cleaning. The quantity  $\Delta\text{OT}$  is equal to  $nL_{\text{final}} - nL_{\text{initial}}$ ,  $nL_{\text{final}}$  and  $nL_{\text{initial}}$  represent the optical thickness at time 0 and  $T$ , respectively in (a). In (b),  $\Delta\text{OT}$  represents two states:  $nL$  changes after dodecane dosing and air purging,  $nL$  changes after visible light cleaning.



Fe<sub>2</sub>O<sub>3</sub> interferometer in the air can be explained as follows: as α-Fe<sub>2</sub>O<sub>3</sub> is irradiated by visible light, holes and electron pairs are created and electron could react with O<sub>2</sub> to generate highly reactive superoxide anion radicals ( $\cdot\text{O}_2^-$ ), (eqn (4) and (5)). The  $\cdot\text{O}_2^-$  are extremely reactive and are able to decompose the organic vapors in the pores of the α-Fe<sub>2</sub>O<sub>3</sub> interferometer.<sup>33</sup>



## 4. Conclusions

In summary, mesoporous α-Fe<sub>2</sub>O<sub>3</sub> interferometers with well-resolved optical fringes were fabricated by an anodization method. The as-fabricated α-Fe<sub>2</sub>O<sub>3</sub> interferometers were applied as optical sensor for various gases detection. The response of the α-Fe<sub>2</sub>O<sub>3</sub> interferometer to different analyte was found to strongly depend on the physical properties of the analyte. The lower the vapor pressure of the analyte, the better the sensitivity will be. Moreover, we have also investigated the self-cleaning ability of the fabricated α-Fe<sub>2</sub>O<sub>3</sub> sensor. With the aid of visible light illumination, the gas sensor was able to recover effectively with good reproducibility and reversibility. These results suggests that the as-fabricated mesoporous α-Fe<sub>2</sub>O<sub>3</sub> interferometer is an attractive sensing material for effective detection of organic vapors with good self-recovery ability.

## Conflicts of interest

There are no conflicts to declare.

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