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Synergistic effect of adsorption coupled with catalysis based on graphene-supported MOF hybrid aerogel for promoted removal of dyes†

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A three-dimensional MIL-100(Fe)/graphene hybrid aerogel (MG-HA) was fabricated *via in situ* decoration of graphene oxide with MIL-100(Fe) nanoparticles. The resulting MG-HA with interconnected pore structure was applied as both adsorbent and catalyst for the removal of methylene blue (MB) from aqueous solutions. The result shows that the saturation adsorption capacity of the MG-HA was as high as 333.33 mg g⁻¹, exceeding that of both the corresponding pristine graphene aerogel and MIL-100(Fe) nanoparticles. In the presence of hydrogen peroxide, MG-HA further exhibited catalytic degradation ability. The dual functions achieved a synergistic effect leading to the quick and complete removal of MB. The benefit was revealed in the treatment of high concentration of pollutants without leaving secondary pollution. The merit was intuitively demonstrated in the instant removal of MB through a model separation device in comparison with a series of common adsorbents. A feasible mathematic model was built based on the synergistic adsorption/catalysis process, which perfectly fitted the experimental data. A pseudo-second-order adsorption process and pseudo-first-order catalytic degradation kinetics were revealed. Additionally, the MG-HA was able to retain 93.4% of its initial removal efficiency after 5 cycles of application. The macro-material body can be easily separated and reused without a time-consuming and high-cost recycling process.

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1. Introduction

In the dyestuff industry, increasing water pollution has been a very serious problem and is having an adverse impact on public health and social economy.¹ Nowadays, there are several methods used to treat the polluted water, such as adsorption,² chemical catalytic oxidation,³ biodegradation,⁴ solvent extraction,⁵ photocatalytic degradation⁶ and membrane separation.⁷ Adsorption and chemical catalytic oxidation have been promising approaches. The adsorption technique is one of the effective, economical and easily regenerated ways; however, such a technique will have to deal with the used adsorbents, which will otherwise cause secondary pollution and need an appropriate follow-up process. Chemical catalytic oxidation tends to remedy the problems facing the adsorption techniques *via* the degradation of pollutants. Unfortunately, this technique is not suitable for removing dyes at low concentration level, considering its relatively low oxidation rate and high cost.⁸ In other words, it is difficult for chemical catalytic oxidation to

treat low concentrations of dyes economically. By combining adsorption with the chemical catalytic oxidation technique, in which the adsorbent itself acts as a catalyst for the degradation of dyes, a kind of adsorption/catalysis combined technique has been developed.⁹ The dyes are concentrated on the adsorbent and then degraded by catalytic oxidation. Under this circumstance, the catalytic degradation is accelerated in that the dye is concentrated and more importantly the concentrated dye contacts the catalyst tightly. After the chemical degradation of dyes from the adsorbent, the adsorption equilibrium is broken and more dyes would be adsorbed. The synergistic process will build up in the reciprocal interaction between adsorption and degradation, giving rise to high removal rate without any risk of releasing secondary pollution.

High surface area, appropriate porous structure and high catalytic activity are usually expected to achieve high efficiency in the design of an adsorbent/catalyst.¹⁰ Metal-organic frameworks (MOFs), being composed of organic bridging ligands and metal ions to form three-dimensional (3D) extended networks with uniform pore diameter, are considered superior in both adsorption and catalysis owing to their high intake capacity, better selectivity and also the possibility of tuning the chemical properties by varying organic linkers or metal atoms.^{11,12} In addition, MOFs catalyzed a couple of reactions owing to the underlying building blocks and the frameworks themselves possessing sites of intrinsic catalytic potential (*i.e.* Lewis acidic

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metal sites, organocatalytic moieties, pore size and shape).^{13–15} However, obvious obstacles lie in the difficulty in recycling either as adsorbent or catalyst materials due to the submicron or micron size of MOF particles. Graphene is a one-atom-thick and two-dimensional (2D) layer of sp²-hybridized carbon.¹⁶ Interest in graphene has grown due to its extraordinary physicochemical properties such as high specific surface area, unique morphology, chemical structure, and electrical properties. Graphene has thus been used as an expeditious adsorbent for various pollutants due to its large, delocalized π -electron system, which can form strong interactions with pollutants.^{17,18} Graphene composite materials showed more advantages in terms of the irreversible agglomeration of graphene into graphite, which usually led to poor processability. 2D calcium alginate and graphene oxide (GO) composite materials revealed 114.9–181.8 mg g⁻¹ adsorption capacity for methylene blue (MB).¹⁹ GO/chitosan/silicon dioxide composite fiber adsorbed Congo red at an adsorption capacity of up to 294.15 mg g⁻¹.²⁰ Composite structures of graphene with layered double hydroxides were expected to exhibit high adsorption capacity.²¹

On the other hand, graphene functions as a metal-free catalyst. Researchers listed the active sites, including oxygenated functional groups, carbon vacancies and holes, edge effects, the presence of dopant elements *etc.* Besides the discrete active sites, the catalytic activity also arises from the adsorption and activation of substrates and reagents by charge transfer.²² This implies that the adsorption and catalysis will promote each other.

Nevertheless, the above methods usually resulted in the powder state of the adsorbent or catalyst. This is not conducive to recycling, leading to a time-consuming and high-cost recycling process. Graphene aerogels (GA) inherit the properties of graphene, especially having the structure of a macro-material body, high accessible surface areas and tunable macropores, and thus have attracted great attention in adsorption.^{23–26}

An attractive process is to make the structure of an aerogel by the combination of MOF particles with graphene. In this study, an efficient adsorption/catalysis material was fabricated and investigated as both adsorbent and catalyst to evaluate the synergetic effect on the removal of MB. The reusability and recyclability were explored as well.

2. Experimental

2.1. Materials

All chemical reagents were used as received. Iron(III) chloride hexahydrate (FeCl₃·6H₂O, 97%), 1,3,5-benzenetricarboxylic acid (trimesic acid, H₃BTC), 1,2-diaminoethane anhydrous (EDA, 98%), graphite, sulfuric acid (H₂SO₄, 95–98%), nitric acid (HNO₃, 65–68%), phosphoric acid (H₃PO₄, 85%), hydrogen peroxide aqueous solution (H₂O₂, 30%) and potassium permanganate (KMnO₄, 99.5%) were purchased from Sino-pharm Chemical Reagent. MB (70%) was purchased from Aladdin Chemical Reagent. All chemicals were of analytical grade and were used without further purification. Water used in all experiments was distilled and deionized.

2.2. Synthesis of GO

GO was prepared by an improved Hummers' method as described in ref. 27. The as-prepared GO was dispersed in deionized water (2 mg mL⁻¹) and sonicated for 30 min. The resulting brown dispersion was subjected to centrifugation for 20 min at 4000 rpm to remove un-exfoliated GO.

2.3. Synthesis of MIL-100(Fe)

MIL-100(Fe) was synthesized according to a hydrothermal reaction reported by Romain.²⁸ Typically, a mixture of FeCl₃·6H₂O (1 mmol, 162.2 mg) and H₃BTC (0.66 mmol, 138.7 mg) dispersed in 20 mL of deionized water was heated for 72 h at 120 °C in a 50 mL Teflon-lined autoclave. After cooling, the prepared orange powder was collected by centrifugation and washed thoroughly with deionized water and ethanol three times to remove the unreacted H₃BTC. Then, the orange solid was recovered by freeze-drying.

2.4. Synthesis of MIL-100(Fe)/graphene hybrid aerogel (MG-HA)

MG-HA was prepared using the following procedures. Briefly, a certain amount of MIL-100(Fe) was added into 5 mL of GO dispersion (2 mg mL⁻¹) to obtain a homogeneous mixture under ultrasonication. Then 20 μ L of EDA was added dropwise to the mixture. Afterwards 5 mL of the mixture was placed in a 10 mL vial and maintained at 180 °C in a 50 mL Teflon-line autoclave for 12 h. After that, the autoclave was cooled down to room temperature. The as-formed product then was lyophilized to yield the aerogel. Variation of the quantity of MIL-100(Fe) (5, 10, 20 mg) would produce different hybrid aerogels, which reflect the mass ratio of MIL-100(Fe) to GO as 0.5, 1.0 and 2.0, respectively termed as 0.5MG-HA, 1.0MG-HA and 2.0MG-HA. Pure GA was prepared under the same conditions in the absence of MIL-100(Fe) for comparison.

2.5. Characterization

The morphology of MIL-100(Fe), GA and MG-HA was characterized by scanning electron microscopy (SEM, Hitachi, TM3030, Japan) combined with energy dispersive spectroscopy (EDS) and transmission electron microscopy (TEM, JEM-2100USA). Raman spectra were collected by using a Raman spectrometer (Thermo Scientific DXR, USA) equipped with a 633 nm argon ion laser. The chemical structure of the membranes was analyzed with attenuated total reflectance FTIR (PerkinElmer, Spectrum II, USA) in the wavenumber range of 400–4000 cm⁻¹. Samples were ground with KBr and pressed to plates for measurement. The Brunauer–Emmett–Teller (BET) surface area and pore size were determined by nitrogen adsorption (ASAP2460, Micromeritics, USA). The concentration of MB was analyzed with an ultraviolet spectrophotometer (Molecular Devices, USA).

2.6. Adsorption experiments

MB was used as the target pollutant. 10 mg of MG-HA was added into 20 mL of MB aqueous solutions of different concentrations



(150, 200, 250, 300 and 400 mg L⁻¹) at 30 °C at 130 rpm in a shaking bath. After a fixed time, the adsorbent was separated. The residual solution concentration was determined by UV-visible absorbance at 664 nm. The adsorption amount was calculated according to the following equation:

$$Q_t = \frac{(C_0 - C_t)V}{m}$$

where Q_t (mg g⁻¹) is the adsorption amount at adsorption time t , C_0 is the initial concentration of dye (mg L⁻¹), C_t is the concentration of dye (mg L⁻¹) at adsorption time t , m is the mass of adsorbent (g), and V is the volume of the solution (L).

2.7. Catalytic degradation

10 mg of MG-HA was added into 20 mL of MB solution (200 mg L⁻¹) containing 100 μL of H₂O₂ (30%) at pH 6.0. After the pH was adjusted to 6.0, the mixture was shaken at 130 rpm at 30 °C. 40 μL samples were taken at predetermined time intervals for the analysis of residual MB concentration in the aqueous solution.

To investigate the synergistic process of adsorption coupled with degradation, 10 mg 1.0MG-HA was added into 20 mL of MB solution (200 mg L⁻¹) at pH 6.0 for adsorption. After shaking for 120 min, the desired amount of H₂O₂ was added to initiate oxidative degradation of MB. At predetermined time intervals, 40 μL of the mixture was taken for analysis of residual MB. After degradation, the adsorbent was collected and reused in the next cycle. All the experiments were performed in triplicate.

3. Results and discussion

3.1. Characterization

A schematic illustration for the fabrication of MG-HA is shown in Fig. 1. At first, MIL-100(Fe) was prepared by a hydrothermal process. Then MIL-100(Fe) and GO were autoclaved to obtain a 3D network of MG-HA. As shown in Fig. S1a,† MG-HA prepared with different ratios of MIL-100(Fe) to GO displayed different appearance. More MIL-100(Fe) resulted in a slacker texture. As an illustration, 1.0MG-HA can be set on dandelion flowers easefully without incurring any shape change of soft villi, indicating that the quality of this material is very light (Fig. S1b†). The hybrid aerogel can also support about 104 times

its own weight, showing good mechanical strength though with ultralight density (Fig. S1c†). Importantly, different shapes can be conveniently fabricated dependent on the shape of the mold in which the aerogel was made (Fig. S1d†). A column shape was obtained, for instance, when a round-bottom vial was used (Fig. S1e and f†). That might meet the requirements in various application fields. Therefore, the macro body of MG-HA with ultralight density and good mechanical strength allows it to be applied efficiently and recovered conveniently as a pollutant adsorbent and catalyst.

The morphology and structural details of the as-prepared MG-HA in different formulations were investigated by SEM (Fig. 2) and TEM (Fig. 3). From the SEM images in Fig. 2a–f, we can see that the as-prepared MG-HA showed well-defined and interconnected porous network with the pore size ranging from submicrometer to micrometers, which was constructed by thin layers of stacked graphene nanosheets. Compared with the smooth surface of pure GA (Fig. S2a†) and with the MIL-100(Fe) nanoparticles (Fig. S2b†), MG-HA exhibited a porous structure and numerous nanoparticles entrapped in a sheet-like structure. The nanoparticles were strongly believed to arise from MIL-100(Fe) and the sheet-like structure originated from graphene. The size of the nanoparticles in MG-HA was in a similar range to that of pure MIL-100(Fe) (Fig. S2b†). More nanoparticles were observed with elevated MOF ratios in the preparation process. EDS mapping analysis of 1.0MG-HA (Fig. 2g and S2c†) showed the distribution of C, Fe and O elements throughout the hybrid aerogel. Fig. 3 clearly shows that the graphene nanosheets in thickness were connected to form the network-like structure and the MG-HA were decorated with clusters of MIL-100(Fe) crystals between the graphene sheets. The interconnected graphene sheets prevented the agglomeration of MIL-100(Fe) and allowed the particles to disperse well in terms of the TEM observation.

The formation of MG-HA was confirmed by XRD examinations, and the XRD patterns of MIL-100(Fe), GA and MG-HA are shown in Fig. 4A. In the XRD pattern of GA, the diffraction peak centered at 25.3° comes from the (001) crystal plane. The diffraction pattern of the 0.5MG-HA composite is similar to that observed for GA; no characteristic diffraction peaks of MIL-100(Fe) are observed, which indicates that the structure of GA is preserved. However, with an increase of the amount of MIL-100 added, the characteristic signals of MIL-100 crystals were also observed in the XRD pattern of 1.0MG-HA, and followed by the presence of XRD peaks of Fe₃O₄, implying that the MIL-100 crystals were gradually assembled in GA microenvironment, which is consistent with SEM and TEM results.

The FTIR spectra of MIL-100(Fe) as well as MG-HA are shown in Fig. 4B. The MG-HA spectrum exhibited the vibrations of MIL-100(Fe) at identical bands of 1443, 1383, 757, and 708 cm⁻¹,³⁰ which confirmed the successful incorporation of MIL-100(Fe) into GA. Raman spectroscopy is highly sensitive to electronic structure and has proven to be an essential tool for the characterization of carbon-based materials (especially for C=C bonds).³¹ The *in situ* deposition of MIL-100(Fe) particles and the self-assembly of GO sheets into the 3D interconnected aerogel structure are also driven by the combined hydrophobic and π–π

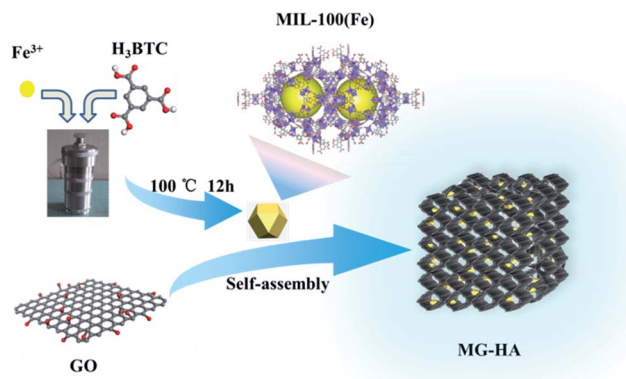


Fig. 1 Schematic illustration of the preparation of MG-HA.



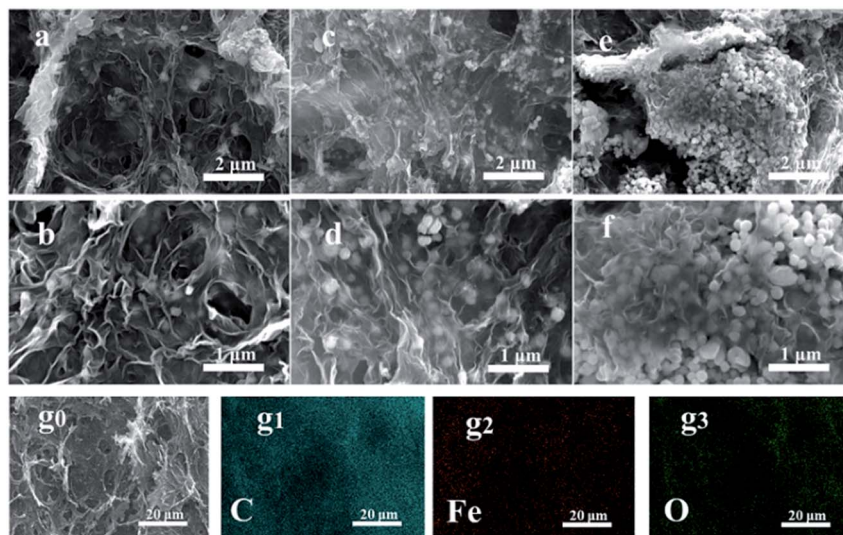


Fig. 2 SEM images of (a and b) 0.5MG-HA, (c and d) 1.0MG-HA, (e and f) 2.0MG-HA, and (g1–g3) EDS mapping corresponding to SEM of 1.0MG-HA (g0).

stacking interaction of molecules due to the reduction of oxygenated moieties on the graphene surface. In Fig. 4C, the Raman spectra of the hybrid aerogels preserved the characteristic bands of GA as reported elsewhere.³² Two strong bands at 1350 and 1590 cm^{-1} of GA can be clearly seen, corresponding to the well-documented D and G bands of carbon materials, respectively.³³ However, these Raman bands of the MIL-100(Fe) phase cannot be observed in the MG-HA spectra. The reason may be that the MIL-100(Fe) crystals of MG-HA were sheltered by GA, implying the good dispersion of MIL-100(Fe) nanoparticles on GA layers without local aggregation. The intensity ratio of D to G bands (I_D/I_G) is generally indicative of the disorder degree and healing of defects. The I_D/I_G intensity ratios of 0.5MG-HA, 1.0MG-HA and 2.0MG-HA were 1.41, 1.39 and 1.42, respectively, all higher than that of GA (1.37). This slight increase in I_D/I_G values was ascribed to the doped MIL-100(Fe) particles and wrinkles in the interconnected frameworks.³² Or, the aerogel was not just simply the physical mixture but involved a structure alteration.

The defective structure may provide a large number of active sites for binding with other molecules.

Pore structure and surface area are vital for adsorption and catalysis, so the BET specific surface area and pore structure parameters of MIL-100(Fe), GA and MG-HA were studied by N_2 adsorption–desorption analysis. As can be seen in Fig. 4D, the N_2 adsorption–desorption curves of MIL-100(Fe), GA and MG-HA displayed a type IV isotherm with an H_3 -type hysteresis loop, indicating a porous structure of all the samples.²⁹ The values of BET surface area and total pore volume are listed in Table 1. MIL-100(Fe) exhibited much larger BET surface area than GA. After the combination, 1.0MG-HA possessed a surface area of 164.94 $\text{m}^2 \text{g}^{-1}$, significantly larger than that of both parent structures. This result can be attributed to the fact that the MIL-100(Fe) particles widened the distance between the GA nanosheets and enlarged the GA pores. Besides, the pores of

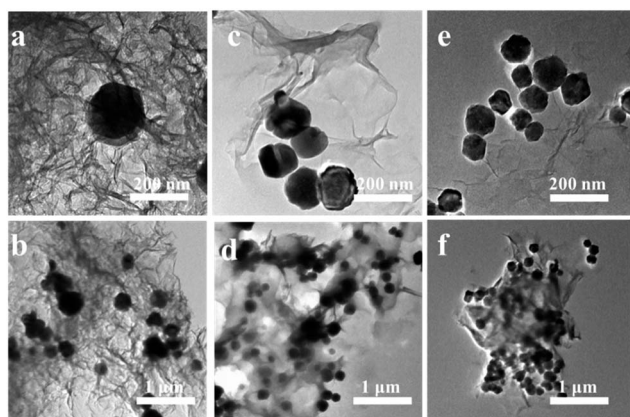


Fig. 3 TEM images of (a and b) 0.5MG-HA, (c and d) 1.0MG-HA and (e and f) 2.0MG-HA.

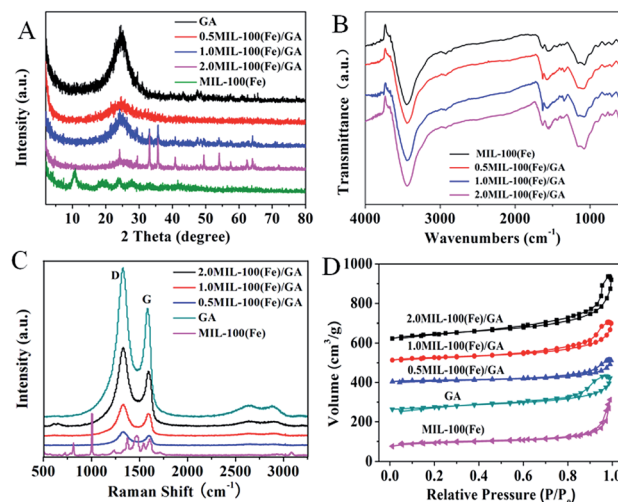


Fig. 4 (A) XRD patterns, (B) FTIR spectra, (C) Raman spectra and (D) N_2 adsorption/desorption isotherms.



Table 1 BET surface area, total pore volume, and pore diameter of MG-HA samples

| | GA | MIL-100(Fe) | 0.5MG-HA | 1.0MG-HA | 2.0MG-HA |
|---|-------|-------------|----------|----------|----------|
| BET surface area ($\text{m}^2 \text{g}^{-1}$) | 88.25 | 149.51 | 96.78 | 164.94 | 40.82 |
| Total pore volume ($\text{cm}^3 \text{g}^{-1}$) | 0.27 | 0.41 | 0.31 | 0.5 | 0.18 |

MIL-100(Fe) were less blocked at the lower amount of MOF (1.0MG-HA). Further increasing the amount of MOF particles will fill the pores of GA. Also, MOF layers stacked on each other and thus the MOF pores were sealed to some degree. Fig. 2 verified the sealing effect. The BET surface area and the pore volume thus decreased with the amount of MOF.

3.2. Adsorption kinetics and adsorption isotherms

In order to highlight the superior performance of the prepared hybrid aerogels, GA and MIL-100(Fe) particles were involved for comparison. Fig. 5 shows the kinetic adsorption, where the adsorption amount reached 70–80% of saturation capacity in the initial 10 minutes for all these adsorbents. The quick adsorption ability was ascribed to the highly interconnected pore structure. Afterward, the adsorption became much slower. GA and MIL-100(Fe) displayed 263.16 and 312.50 mg g^{-1} adsorption capacity, respectively. 0.5MG-HA and 2.0MG-HA did not show as good an adsorption capacity as 1.0MG-HA. 1.0MG-HA showed the highest adsorption capacity (333.33 mg g^{-1}). As discussed above, 1.0MG-HA possesses the largest pore volume and surface area with interconnected pore structure (Table 1). These structural features endowed the aerogel with high capacity as an adsorbent. The 3D interconnected network provides adequate molecular transportation channels and allows easy diffusion of dye molecules from the exterior to interior sites in the adsorbent. A large amount of negatively charged oxygen-containing functional groups (such as hydroxyl, carboxyl and carbonyl groups) available on graphene nano-sheets provide additional opportunity for the adsorption of dye molecules through electrostatic attraction. Even more, the existence of π - π conjugation between MB molecules and graphene will augment the adsorption process. Fe Lewis acid sites in MIL-100(Fe) and Lewis base $-\text{N}(\text{CH}_3)_2$ in MB strengthened

the adsorption as well. These structural and chemical mechanisms contributed to the elevated adsorption. But the aerogel with high proportion of MOF demonstrated lower adsorption capacity probably due to the stacking of MOF, mentioned above.

To investigate the mechanism of adsorption, kinetic studies were performed at 25 °C. The pseudo-first-order and pseudo-second-order kinetic models were applied to fit the experimental data. The pseudo-first-order model can be described as follows:³⁴

$$\ln(Q_e - Q_t) = \ln Q_e - k_a t$$

where Q_e (mg g^{-1}) and Q_t (mg g^{-1}) correspond to the amount of MB adsorbed at equilibrium and at time t , respectively; k_a is the rate constant of adsorption (min^{-1}). The pseudo-first-order model fitted kinetic data are shown in Fig. S3.† The values of Q_e and k_a are calculated by the slope and intercept of the plot.

The pseudo-second-order kinetic expression is described in the following form:³⁵

$$\frac{t}{Q_t} = \frac{1}{k_a Q_e^2} + \frac{t}{Q_e}$$

where k_a ($\text{g mg}^{-1} \text{min}^{-1}$) is the pseudo-second-order rate constant. The adsorption rate constant k_a was calculated from a plot of t/Q_t versus t . Fig. S3† shows the plots of the pseudo-first-order and second-order kinetics. The corresponding kinetic parameters and the correlation coefficients are summarized in Table S1.†

As can be seen from Table S1,† all the pseudo-second-order correlation coefficients (R^2) were close to 0.99, much higher than those of the pseudo-first-order model. The calculated Q_e values fully agreed with the experimental data in the case of pseudo-second-order kinetics. Therefore, the adsorption behavior can be perfectly described by the pseudo-second-order model, which means that chemical adsorption is the dominating step.

In order to investigate the adsorption capacity, adsorption isotherm experiments were carried out by varying the initial MB concentration from 150 to 400 mg L^{-1} . The Langmuir and Freundlich isotherm models were applied to fit the experimental data. The Langmuir model is based on the assumption of monolayer adsorption on a homogeneous surface³⁶ and is defined by the following formula:

$$\frac{C_e}{Q_e} = \frac{1}{Q_{\max} a_L} + \frac{C_e}{Q_{\max}}$$

The Freundlich model is an empirical expression which describes multilayer adsorption at heterogeneous surfaces³⁶ and is expressed as follows:

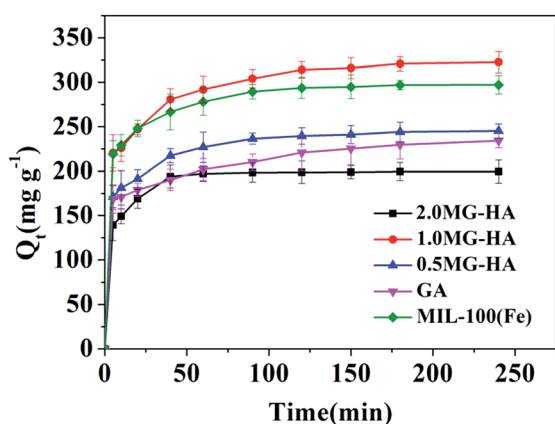


Fig. 5 Adsorption of MB on GA, MIL-100(Fe) and MG-HA (initial MB concentration: 150 mg L^{-1}).



Table 2 Langmuir and Freundlich isotherm parameters for MB adsorption

| | Langmuir | | | Freundlich | | |
|-------------|----------------------------------|-----------------------------|-------|---|------|-------|
| | Q_{\max} (mg g ⁻¹) | a_L (L mg ⁻¹) | R^2 | K_f (mg ^{1-1/n} L ^{1/n} g ⁻¹) | n | R^2 |
| 0.5MG-HA | 322.58 | 0.500 | 0.999 | 132.09 | 4.75 | 0.809 |
| 1.0MG-HA | 333.33 | 0.280 | 0.999 | 144.85 | 5.90 | 0.809 |
| 2.0MG-HA | 263.16 | 0.253 | 0.999 | 83.61 | 3.54 | 0.905 |
| MIL-100(Fe) | 312.50 | 0.593 | 0.999 | 150.85 | 6.31 | 0.829 |
| GA | 263.16 | 0.105 | 0.998 | 110.56 | 6.55 | 0.916 |

$$\ln Q_e = \ln K_f + \frac{1}{n} \ln C_e$$

where Q_e (mg g⁻¹) is the equilibrium adsorption amount, C_e (mg L⁻¹) is the MB concentration after the adsorption reached equilibrium, Q_{\max} (mg g⁻¹) is the maximum adsorption amount calculated by Langmuir simulation curve, which reflects the adsorption ability by theoretical calculation, and a_L (L g⁻¹) is the Langmuir adsorption constant. The Freundlich constants K_f and n respectively represent the adsorption capacity and the adsorption strength.

The corresponding parameters of the two models are listed in Table 2. According to the two correlation coefficients (R^2), the adsorption fits better the Langmuir isotherm model in contrast to the Freundlich model, reflecting that the process is monolayer adsorption on the surface of the adsorbent.

Fig. S4† proposes a possible adsorption mechanism of MG-HA for MB. The adsorption process was composed of surface adsorption, interparticle diffusion adsorption, and equilibrium. Surface adsorption was the rate-determining step achieved by three types of interactions: (1) electrostatic interaction; (2) hydrogen bonding; and (3) hydrophobic interaction. Electrostatic interaction occurred between deprotonated groups (*i.e.*, -COO⁻) on the surface of graphene and MIL-100(Fe) particles, and positively charged quaternary ammonium groups in MB. Hydrogen bonding resulted from the interaction between -OH on the surface of graphene and amine in MB. Hydrophobic interaction was a tendency of nonpolar groups to associate in aqueous solution. Because of the aliphatic branches on its pendant groups, the grafted graphene had a hydrophobic character and could react with benzene rings in MB through hydrophobic interaction.

3.3. Catalytic degradation

The approach of adsorption coupled with catalytic degradation was applied for the highly efficient removal of MB by using MG-HA. Fig. 6 illustrates the synergistic process of adsorption coupled with catalysis. MG-HA initially adsorbed MB from aqueous solution and then catalyzed the degradation of MB *via* the activation of hydrogen peroxide (H₂O₂).

Fig. 7 demonstrates the results based on the synergistic effect of adsorption coupled with catalysis. In Fig. 7a, only 3% of MB was cleared within 80 minutes solely in the presence of H₂O₂. This result indicated that MB would not be oxidized by H₂O₂ without the involvement of catalysts. When 1.0MG-HA was applied in the absence of H₂O₂, obviously, only adsorption

happened under these conditions. The adsorption resulted in 73.7% removal of MB within 60 minutes. Prolonged adsorption time brought a very limited increase of adsorption. The adsorption approached equilibrium. Afterward, H₂O₂ was added to initiate the catalytic oxidation process. The removal resumed. An accelerated degradation rate was observed. The removal percentage reached 99.3% in subsequent 20 minutes. The stepwise treatment apparently demonstrated the dual function of the obtained MG-HA. Especially, the removal rate was accelerated by the application of H₂O₂ though the solution concentration was pretty low at the moment that H₂O₂ was added. This is assumedly attributed to the fact that high MB concentration which has been already adsorbed on the surface of the adsorbent/catalyst brought about the quick oxidative degradation rate. The degradation of MB from the surface will in turn accelerate the adsorption. The synergistic mechanism finally led to the easy removal of MB. When H₂O₂ and 1.0MG-HA

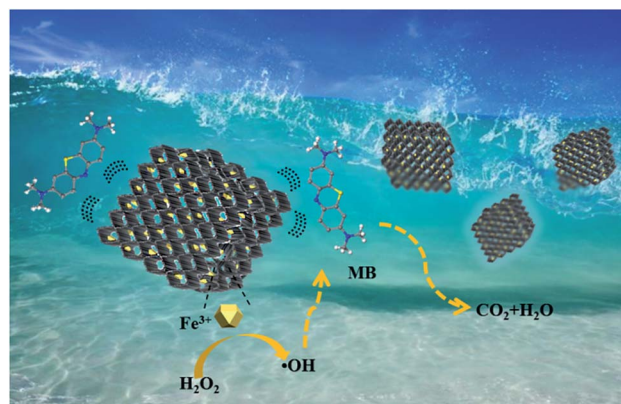


Fig. 6 Schematic diagram of the adsorption/catalysis coupled process of MG-HA to remove MB.

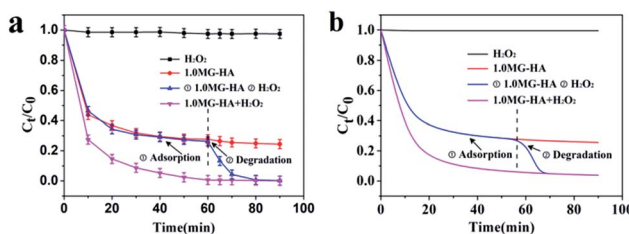


Fig. 7 (a) Adsorption–degradation processes in the presence of H₂O₂, initial 1.0MG-HA and subsequent H₂O₂, and both 1.0MG-HA and H₂O₂. (b) Simulation curves of adsorption–degradation processes.



were added together, 99% of MB was removed within 60 minutes. Apparently, the synergistic effect took place under these conditions. Obviously, the benefits will be revealed fully in the treatment of high concentration of pollutants. In the case of high initial pollutant concentration, the equilibrium concentration of pollutants in solutions would still remain high with the use of only adsorption due to the restriction of the adsorption equilibrium limit. However, the synergistic effect will overcome the equilibrium and initiate the degradation process in the presence of H₂O₂, finally leading to the complete removal of pollutants without the requirement of more adsorbents. More interestingly, the removed pollutants were *in situ* degraded catalytically, avoiding secondary pollution.

The catalytic degradation is further verified in Fig. S5.† There the desorbed sample through adsorption in the absence of H₂O₂ kept the same spectrum as MB itself. However, the sample through the treatment of MG-HA in the presence of H₂O₂ displayed a different spectrum from MB itself. The experimental results indicated the occurrence of catalytic degradation. Meanwhile, the corresponding total organic carbon values were measured (Fig. S6†). The final mineralization rate of MB (200 mg L⁻¹) after H₂O₂ oxidation was 85%. However, the mineralization rate should be higher than the measured value, because the carbon of the catalyst used in the degradation process led to measurement deviation.

To further elucidate the synergistic effect, an adsorption/catalysis coupled mathematic model was established based on the above pseudo-second-order adsorption and pseudo-first-order catalysis behavior. The specific derivation process is presented in the ESI.† The composite model is described as follows:

$$\frac{C_t}{C_0} = 1 - \frac{t}{\frac{mC_0}{k_a V(C_0 - C_{e0}e^{-k_c t})^2} + \frac{tC_0}{C_0 - C_{e0}e^{-k_c t}}}$$

where k_a (g mg⁻¹ min⁻¹) is the adsorption pseudo-second-order rate constant and k_c (min⁻¹) is the degradation pseudo-first-order rate constant. C_{e0} (mg L⁻¹) is the dye concentration after the adsorption reached equilibrium without catalysis effect, C_0 is the initial concentration of dye (mg L⁻¹), and t (min) is the adsorption/reaction time.

To verify the validity of the formula, four different cases (H₂O₂ only, 1.0MG-HA only, first 1.0MG-HA then H₂O₂, and 1.0 MG-HA + H₂O₂) were explored. The constants of pseudo-second-order kinetic plots and degradation pseudo-first-order rate are summarized in Table S2.† The simulation curves are shown in Fig. 7b. It was found that the simulation curves were in good agreement with the experimental curves in Fig. 7a, which reflected the feasibility of this composite model. Therefore, successful agreement of this model further validated the adsorption/catalysis coupled behavior.

The use of recycling is significant in industry. Though MIL-100(Fe) displays good adsorption capacity, it is very inconvenient to recycle in practical applications. In contrast, macro-size materials can overcome this drawback. This is why the hybrid aerogels were fabricated. To further assess the reusability of 1.0MG-HA for dye degradation (Fig. S7†), the MG-HA was

washed with water after one catalytic run and reintroduced into a fresh dye solution containing 100 μL of H₂O₂ for another degradation run at 30 °C. The result showed that the removal ability in the fifth run reached 93.4% of that in the first run, indicating an expected practicability for industrial applications.

To demonstrate potential application in practice, a direct observation experiment was conducted *via* a syringe filter. 10 mg of sample was loaded into the filter first. 5 mL MB solution (80 mg L⁻¹) was put in a 10 mL syringe, and injected through the adsorbent bed at a speed of 100.0 mL h⁻¹ (Fig. S8†). Quantitative analysis indicated that 99.0% MB was removed when 1.0MG-HA was tested. The comparison study showed clearly that the hybrid aerogel gave the most obvious effect. This observation indicated potential industrial applications.

4. Conclusion

In conclusion, a type of MG-HA was fabricated giving a bimodal synergistic effect of adsorption coupled with catalysis and applied in the removal of MB from aqueous solution. The hybrid aerogel had a large specific surface area and large pore volume. The dual functions of adsorption coupled with catalysis and the synergistic effect were fully exhibited in terms of the removal curves in the presence of H₂O₂, initial 1.0MG-HA and subsequent H₂O₂, and both 1.0MG-HA and H₂O₂. Moreover, a feasible mathematical model was built to be used to simulate the adsorption/catalytic behavior. Finally, the reusability of the adsorbent/catalyst was studied and the results showed that MG-HA can be reused without obvious reduction in the removal capacity. The hybrid aerogel of MOF coupled with graphene provides a feasible way to achieve efficient removal of pollutants such as MB.

Conflicts of interest

There are no conflicts to declare.

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