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Green fabrication of nanoporous BiVO₄ films on ITO substrates for photoelectrochemical water-oxidation†

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A new green method was developed to prepare nanoporous BiVO₄ films on ITO substrates for photoelectrochemical (PEC) water-oxidation under visible light irradiation. The films can be prepared by simple drop-casting of a stable aqueous solution of Bi³⁺ and V⁵⁺ complexes with tartaric acid and ethylenediaminetetraacetic acid, followed by drying and calcination in air. Thanks to these ligands, the aqueous precursor solution is remarkably stable over a wide range of pH (pH 4–9). The BiVO₄ films on ITO substrates possess a 3D-network structure comprised of nanoparticles with a scheelite–monoclinic phase and a diameter of ca. <100 nm, after calcination at 450–500 °C for 1 h. The PEC performance clearly depended on the film thickness that can be controlled by coating times, and calcination conditions (temperature and time). The CoPi-loaded BiVO₄ electrodes exhibited relatively high performance for PEC water oxidation (ABPE of 0.35% at 0.8 V vs. RHE) under simulated sunlight irradiation.

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Introduction

Photoelectrochemical (PEC) water splitting using semiconductor photoelectrodes is a promising technology for sustainable energy supply, equal with photocatalytic water splitting using semiconductor particles.^{1–4} Development of photoelectrodes that can harvest a wide range of visible light to split water efficiently is indispensable for achieving practically useful conversion of solar energy to H₂. Among various visible-light-responsive photoelectrode materials, BiVO₄ (ref. 5) has attracted much attention as an efficient photoanode material that can absorb light up to 520 nm.^{6–27} A lot of methods have been reported to fabricate BiVO₄ thin films on conducting electrodes. For example, Kim and Choi have reported that nanoporous BiVO₄ electrodes can be prepared by applying a DMSO solution of VO(acac)₂ onto a surface of electrochemically synthesized BiOI electrodes, and calcining (Fig. S1(a)†).¹² The electrode exhibits high photocurrent (ca. 2.7 and 5.0 mA cm⁻² at 0.60 and 1.23 V vs. RHE, respectively) under simulated solar-light irradiation because the nanoporous BiVO₄ electrodes have appropriate conductive paths for efficient electron transfer as well as controlled pores allowing efficient reaction and

transportation of substances. For practical use, however, it is desired that such a controlled structure is prepared by more facile processes. Various metal oxide thin films are prepared from metal precursors, for example, *via* the sol–gel method.^{6,20} However, the metal precursors are generally unstable in aqueous solutions; therefore, organic solvents are commonly used to dissolve corresponding metal sources. The use of organic solvents is an obstacle for industrial processes from the viewpoint of environmental pollution. Therefore, in order to reduce volatilization of organic solvents, the use of water as a solvent has been strongly demanded to widen the range of application and also to reduce the environmental burdens. Recently, it has been reported that BiVO₄ photoelectrodes can be prepared by using highly acidic aqueous precursor solutions, although it is also not environmentally friendly (Fig. S1(b)†).^{16,17}

We recently found that BiVO₄ nanoparticles can be prepared *via* a newly-developed aqueous-chelate method using an aqueous solution of Bi and V ions containing ethylenediaminetetraacetic acid (EDTA) and tartaric acid (tart).²⁸ In these aqueous solutions, EDTA and tart can serve as good chelators to stabilize Bi and V ions in aqueous solutions at pH 4 to 9. We also found that this new method affords BiVO₄ particles with scheelite–monoclinic (s–m) phase, which is known as the most favourable crystal phase for water oxidation reaction, and small particle sizes (<100 nm) and partly networked structures, even after calcination process. These findings have motivated us to apply the precursor aqueous solution of mild pH to the fabrication of BiVO₄ photoelectrodes that can efficiently act under visible light. In this study, we have successfully fabricated nanoporous BiVO₄ photoanodes *via* the

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environmentally friendly water-based method and applied them for PEC water oxidation reaction under visible light.

Experimental

Materials

Ammonium vanadate(V) (NH_4VO_3 , 99%) and bismuth(III) nitrate pentahydrate ($\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$, 99.9%) were purchased from Kanto Chemical. EDTA, tart and ammonia solution (28.0–30.0%) were purchased from Wako Pure Chemical Industries, Ltd. All reagents were used as received, and all the experiments were carried out under ambient conditions.

Preparation of BiVO_4 electrodes by using aqueous solutions containing Bi and V ions

The aqueous precursor solutions containing V and Bi ions were prepared *via* aqueous-chelate method, which was recently developed by our group,²⁸ as follows; A 10 mL aqueous solution that contains $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ (0.17 M, 0.82 g), EDTA (0.34 M, 0.99 g) and 28% ammonia solution (10 mM, 1.0 g) was mixed with a 10 mL aqueous solution of NH_4VO_3 (0.17 M, 0.20 g) and tart (0.17 M, 0.25 g). The simple addition of $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ or NH_4VO_3 to water rapidly form precipitate due to rapid hydrolysis. In contrast, the combinations of $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ with EDTA, and of NH_4VO_3 with tart produce a transparent aqueous solution. As we previously reported,²⁸ we confirmed that EDTA and tart make stable complexes with Bi^{3+} and V^{5+} ions in aqueous solution, respectively. BiVO_4 electrodes were prepared by drop-casting method as follows; the aqueous precursor solution containing V and Bi ions was diluted seven-fold and then dropped on a cleaned tin-doped indium oxide (ITO) transparent conductive glass substrate. The film was dried at 65 °C for 10 min. The as-prepared films were then calcined at 500 °C for 1 h to yield 1-coated BiVO_4 films, which were denoted as 1-coated-500 °C-1 h. Films with more than two coats (*n*-coat) were prepared by repeating drop casting, drying and calcinations. Calcination was conducted basically at 500 °C for 1 hour, while other temperatures (*e.g.* 450 and 550 °C) and calcinations times (*e.g.* 2 and 3 hours) were applied in some cases.

Preparation of CoPi-loaded BiVO_4 electrodes

A CoPi was electrochemically deposited on the BiVO_4 photoelectrodes according to the literature.⁸ The electrolyte was prepared by dissolving 2 mM $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (Wako, 99.5%) in a 0.2 M potassium phosphate buffers solution (pH 7.0). The potential of the working electrode was controlled by a potentiostat (Hokuto Denko, HZ-7000) with a three-electrode cell consisting of the BiVO_4 electrode, a Pt wire and a KCl-saturated Ag/AgCl electrode as the working, counter, and reference electrodes, respectively. The electrodeposition was performed at a constant voltage of 1.1 V *vs.* RHE for 2 min.

Characterization

The resulting BiVO_4 fine particles were characterized by means of an X-ray diffraction (XRD, PANalytical, X'Pert Pro, rotating anode diffractometer, 45 kV, 40 mA) with Cu K_α radiation ($K_\alpha =$

1.5406 Å), a UV-Vis-NIR spectrometer (UV-Vis. DRS, Jasco, V-670), and a scanning electron microscope (SEM, HITACHI, S-4100).

PEC measurements of BiVO_4 electrodes

PEC properties were evaluated in potassium phosphate buffers (0.2 M, pH 7.0, 0.1 M K_2SO_4) using a potentiostat with a three-electrode cell consisting of the BiVO_4 electrode, a Pt electrode, and a KCl-saturated Ag/AgCl electrode as the working, counter, and reference electrodes, respectively. The working electrode was irradiated from the backside (ITO side) with simulated sunlight (AM 1.5G, 100 mW cm^{-2}).

Results and discussion

Characterization of BiVO_4 electrodes fabricated by using aqueous solutions containing Bi and V ions

BiVO_4 electrodes were prepared by drop-casting the precursor solution onto ITO transparent conductive glass substrates, followed by drying and calcining at 450–550 °C for 1–3 h. The prepared electrodes were yellow transparent thin films (see Fig. 1, calcined at 500 °C for 1 h for example). The XRD analysis revealed that the electrodes have pure scheelite phase of BiVO_4 (JCPDS-ICDD: 14-0688) regardless of calcination conditions. No appreciable peak attributed to impurity phases was observed (Fig. 2). The bandgap estimated from the UV-Vis spectra (Fig. S2†) was *ca.* 2.4 eV, which is characteristic of s-m phase BiVO_4 .²⁹ The Raman signal attributed to V–O stretching was observed at around 830 cm^{-1} . The estimated value of V–O bond length (*ca.* 1.69 Å, see Table S1†) is in agreement with the reported value for s-m BiVO_4 (1.72 Å).³⁰ Judging from these results, it can be concluded that all the electrodes consist of s-m BiVO_4 .

The SEM images of BiVO_4 electrodes, which were calcined at 450, 500 and 550 °C for 1–3 hour, are shown in Fig. 3 and S3.† The surface SEM images show nanostructures with particle-networks and pores. The primary particle sizes became larger as increasing calcination temperatures; the samples calcined at 450 and 500 °C consisted of relatively uniform particles with

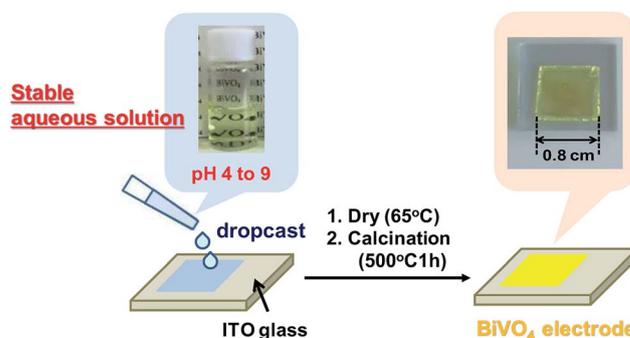


Fig. 1 Fabrication process of a BiVO_4 electrode, which was prepared by coating an aqueous Bi and V solution (pH 7) onto an ITO glass substrate ($1.25 \times 1.25 \text{ cm}^2$), followed by drying (65 °C) and calcination at 500 °C for 1 h.



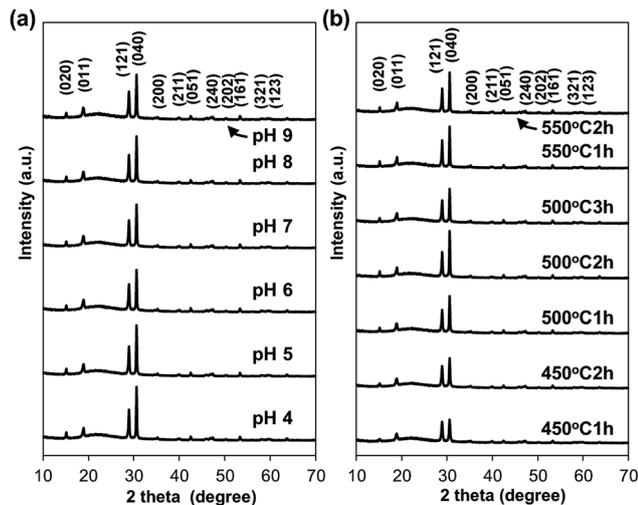


Fig. 2 XRD patterns of BiVO_4 electrodes prepared (a) by coating an aqueous precursor solution of various pH (pH 4 to 9) followed by dryness and calcination at 500°C for 1 h, and (b) by coating an aqueous precursor solution of pH 9 followed by dryness and calcination under different conditions.

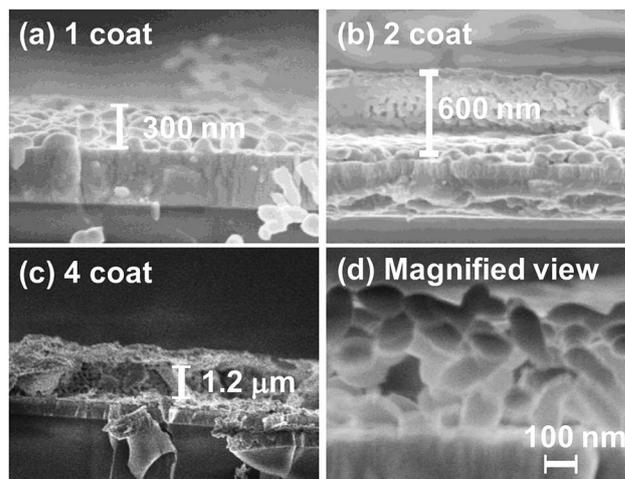


Fig. 4 (a–c) Cross-sectional SEM images of BiVO_4 electrodes with different coating times. (d) Magnified view of cross-sectional SEM images of BiVO_4 electrodes (1 coat- 500°C -1 h).

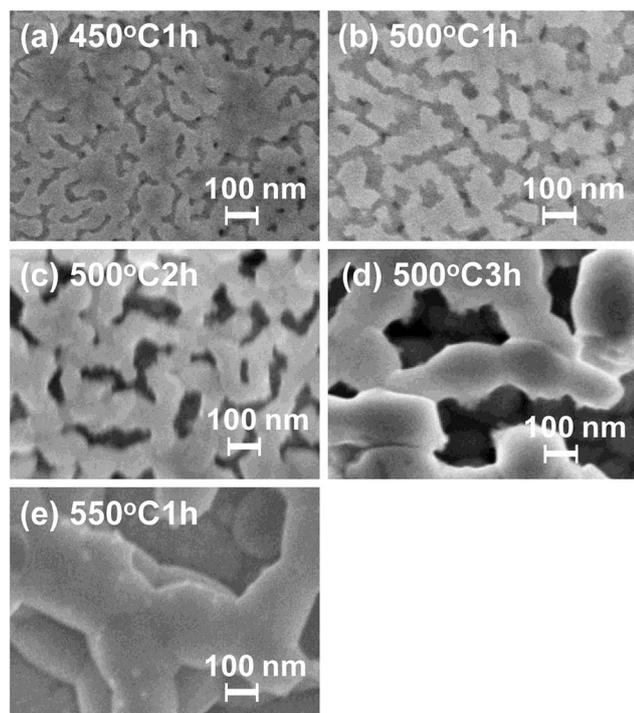


Fig. 3 The surface SEM images of BiVO_4 electrodes (4-coat) prepared by coating an aqueous Bi and V solution (pH 7) onto an ITO glass substrate ($1.25 \times 1.25 \text{ cm}^2$), followed by drying (65°C) and calcination at 450 , 500 and 550°C for 1–3 h.

a diameter of smaller than 100 nm (Fig. 3(a) and (b)), while the sample calcined at 550°C (Fig. 3(e)) was composed of larger particles with a diameter of *ca.* 200 nm .

The photoanode fabricated in this study has 3D-porous network structure as observed with BiVO_4 photoanode reported by Kim and Choi; however, the particle size of our BiVO_4

film (primary particle size: $<100 \text{ nm}$) is smaller than that of their film (primary particle size: $>100 \text{ nm}$).¹² To the best of our knowledge, our report is the first example of BiVO_4 film with porous network structure consisting of such small particles. The particle sizes also increased as the calcinations time increased when compared at the same calcination temperature (see Fig. 3(b)–(d) and S3†). In contrast, BiVO_4 films that were prepared using aqueous precursor solutions of various pH values (pH 4–9) were consisted of BiVO_4 particles with the same s–m structure and similar particle size ($<100 \text{ nm}$) (Fig. 2 and S4†), presumably due to the use of appropriate ligands that effectively stabilize metal ions in the aqueous solution at wide range of pH.

The cross-sectional SEM analysis revealed that the film thickness of *n*-coated- 500°C -1 h ($n = 1$ to 5) were about *ca.* 300 , 600 and 900 nm , and 1.2 and $1.5 \mu\text{m}$, respectively (Fig. 4 and S4†). Thus, the film thickness constantly increased by *ca.* 300 nm with increasing numbers of the coating repetition (see Fig. 4 and S5†). In addition, the magnified view of the cross-sectional images clearly showed that the BiVO_4 film had porous structures with an aggregated network of particles (Fig. 4(d)). As described above, the simple coating of aqueous Bi and V solutions containing EDTA and tart is capable of providing porous nanoparticulate BiVO_4 films with controllable thickness.

PEC properties of BiVO_4 photoelectrodes for sulfite oxidation under simulated solar-light

In order to examine the ability of BiVO_4 photoelectrodes prepared under various conditions, the PEC properties of the photoanodes were assessed in an electrolyte solution containing $0.2 \text{ M Na}_2\text{SO}_3$ as a hole scavenger. Typical *J–V* curves for sulfite oxidation over the BiVO_4 films are shown in Fig. 5. All the BiVO_4 electrodes exhibited obvious photo-response under simulated sunlight irradiation. The photocurrents depended on the preparation conditions, while the onset potentials were



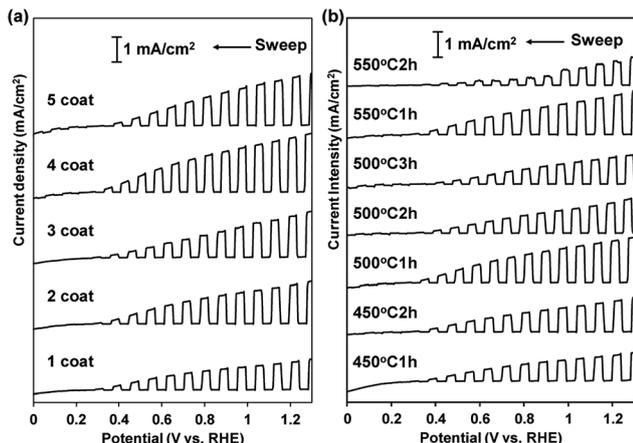


Fig. 5 (a) J - V curves of nanoporous BiVO₄ electrodes for sulfite oxidation. (a) Samples with different coating times. (b) Samples prepared under different calcination conditions (450–550 °C for 1–3 hours). Conditions: in a 0.2 M phosphate buffer (pH 7) containing 0.2 M Na₂SO₃ as hole scavenger under AM 1.5 G, 100 mW cm⁻² illumination (scan rate, 10 mV s⁻¹).

located at about 0.3 V vs. RHE with all samples. In the case of changing numbers of coating repetition, the photocurrent increased with increasing numbers of coating repetition up to fourth cycles (1.4 mA at 1.23 V vs. RHE) and saturated at the fifth coating cycle (Fig. 5(a)).

As described above, the film thickness increased with increasing number of coating. The results suggest that almost all of the particles in the films effectively can absorb the light up to ca. 1 μm in thickness. The photocurrent was also dependent on the calcination conditions (temperature and time); the highest photocurrent was observed with the sample calcined at 500 °C for 1 h (Fig. 5(b)). The highest current density of the sample calcined at 500 °C for 1 h is certainly due to the balance of good crystallinity and high specific surface area of the particles. In contrast, the photocurrent density was not affected by the pH value of the precursor solutions (see Fig. 6, summarized the photocurrent at 0.6 and 1.23 V vs. RHE), which is in agreement with the observation that the film morphology did not depend on the pH value of the precursor solutions (*vide supra*).

PEC properties of BiVO₄ photoelectrodes for water oxidation under simulated solar-light

We next examined water oxidation by the BiVO₄ photoelectrodes. As an effective co-catalyst for water oxidation, small amounts of CoPi were electrochemically deposited on the BiVO₄ photoelectrodes.⁸ The top view SEM elemental mapping images (SEM-EDX) of the surface of CoPi-loaded BiVO₄ photoanode (4 coat-500 °C-1 h) indicate homogeneous loading of Co species (see Fig. S6†). Fig. 7 shows the J - V curves of the CoPi-loaded BiVO₄ photoelectrodes (4-coated-500 °C-1 h) prepared using precursor solutions at pH 3, 7 and 9, which exhibited the best PEC performance for sulfite oxidation. The onset potential of CoPi-loaded BiVO₄ photoelectrodes was more negative than

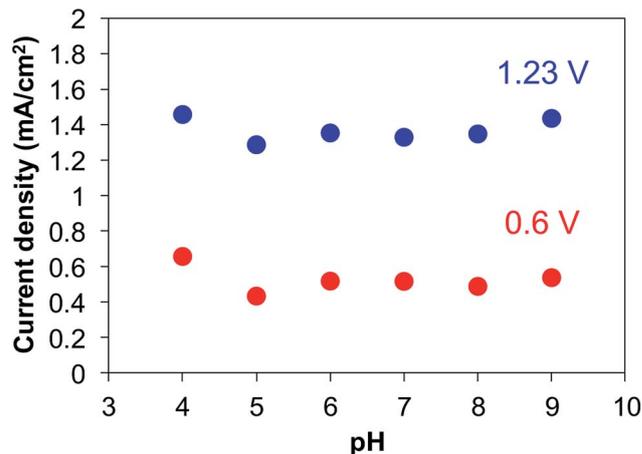


Fig. 6 Effect of the pH value of precursor solutions on the anodic photocurrent density of BiVO₄ electrodes. PEC measurement was carried out in a 0.2 M phosphate buffer (pH 7) containing 0.2 M Na₂SO₃ as hole scavenger under AM 1.5 G, 100 mW cm⁻² illumination (scan rate, 10 mV s⁻¹).

that observed with bare BiVO₄ sample, which is in agreement with the report with CoPi-loaded BiVO₄ photoelectrodes by van de Krol and coworkers.¹⁰ The CoPi-loaded BiVO₄ photoelectrodes showed a photocurrent density of ca. 1.5 mA cm⁻² at 1.23 V vs. RHE for water oxidation, which was higher anodic photocurrent than that of bare BiVO₄ sample (Fig. 7). The value of photocurrent on present CoPi-loaded BiVO₄ photoanode does not reach the high value reported by Kim and Choi (ca. 5.0 mA cm⁻² at 1.23 V vs. RHE);¹² however, the value was higher than that reported with BiVO₄ anodes prepared by using highly-acidic aqueous precursor solutions (<0.5 mA cm⁻² at 1.23 V vs.

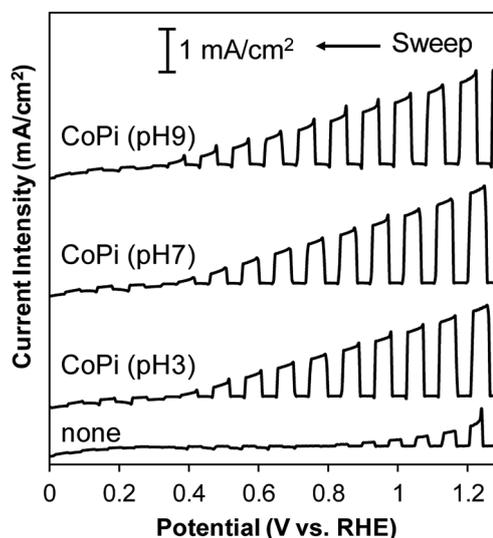


Fig. 7 J - V curves of CoPi-loaded BiVO₄ electrode (4-coated-500 °C-1 h) prepared from an aqueous Bi and V precursor solution of different pH, and BiVO₄ electrode without CoPi loading (4-coated-500 °C-1 h) prepared from an aqueous Bi and V precursor solution of pH 7 for water oxidation measured in a 0.2 M phosphate buffer (pH 7) under AM 1.5 G, 100 mW cm⁻² illumination (scan rate, 10 mV s⁻¹).



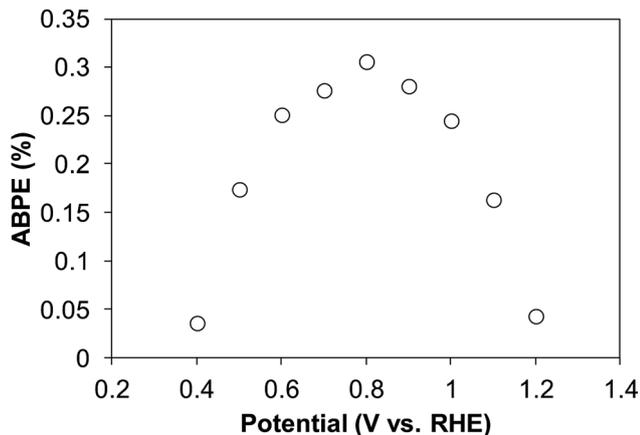


Fig. 8 ABPE plots calculated by using CoPi-loaded BiVO₄ photoanode (4-coated-500 °C-1 h) prepared from a precursor solution of pH 7 under simulated sunlight irradiation.

RHE).^{18,24} The observed relatively high performance should be attributed to the appropriate porous nanoparticulate structure and contact interface. The 3D structure should be attained by direct particle growth from precursor solution on the ITO substrate.

We also examined the stability of BiVO₄ photoanodes with/without CoPi loading. It has been reported that CoPi is stable at pH > 6.³¹ Therefore, we conducted the stability test at pH 7 and 8. Fig. S7† shows the amperometric *J*-*t* curves of BiVO₄ photoanodes with/without CoPi loading at pH 7 and 8 under continuous illumination at 0.8 V vs. RHE. CoPi-loaded BiVO₄ photoanode exhibited better stability at pH 8 than at pH 7; however, about 50% of the initial photocurrent was retained for CoPi-loaded BiVO₄ in the phosphate electrolyte of pH 8 after 2 h illumination (Fig. S7(a)†). In contrast, the BiVO₄ photoanode without co-catalyst loading (Fig. S7(b)†) did not show significant pH dependence of photocurrent decay, and more than 80% of the beginning current was retained after 2 h light irradiation. Thus, the photocurrent decay of CoPi-loaded BiVO₄ photoanode suggests the instability of CoPi, rather than BiVO₄ itself. The stability of the present CoPi-loaded BiVO₄ photoanode is not sufficiently high; however, there are some possibilities to improve the stability by loading other co-catalysts.^{32–34}

The applied bias photon-to-current efficiency (ABPE) was calculated for the CoPi-loaded BiVO₄ photoelectrodes (4-coated-500 °C-1 h) under simulated sunlight irradiation (100 mW cm⁻²) using the following equation.

$$\text{ABPE (\%)} = 100 \times [\text{photocurrent}/\text{mA cm}^{-2}] \times [(1.23 - V_{\text{RHE}})/V]/[\text{photon flux}/\text{mW cm}^{-2}]$$

In the range of 0.4 V to 1.2 V vs. RHE, the CoPi-loaded BiVO₄ photoelectrode gave the highest ABPE of 0.31% at 0.8 V vs. RHE under simulated sunlight irradiation (Fig. 8). Thus, highly efficient photoelectrodes for water oxidation can be

fabricated by drop-casting aqueous metal precursors and calcination under environmentally friendly conditions.

Conclusions

In summary, BiVO₄ photoanodes were directly fabricated onto ITO substrates by drop-casting using aqueous Bi and V solutions containing EDTA and tart. The chelators stabilize the precursor solutions in a wide range of pH. It should be noted that the procedure described here possesses numbers of desirable properties; no use of organic solvents, neutral pH, and easy process to yield nanoporous BiVO₄ films. Especially, it is possible to prepare BiVO₄ films with the same structure at a wide range of pH if our aqueous precursor solutions are employed, which would be a great benefit to many researchers in this field. The essence of our present fabrication approach using aqueous chelating solutions is expected to be widely applicable to the preparation of other metal oxide films.

Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 A. Fujishima and K. Honda, *Nature*, 1972, **238**, 37–38.
- 2 K. Maeda and K. Domen, *J. Phys. Chem. C*, 2007, **111**, 7851–7861.
- 3 R. Abe, *J. Photochem. Photobiol., C*, 2010, **11**, 179–209.
- 4 F. E. Osterloh, *Chem. Soc. Rev.*, 2013, **42**, 2294–2320.
- 5 A. Kudo, K. Ueda, H. Kato and I. Mikami, *Catal. Lett.*, 1998, **53**, 229–230.
- 6 H. Liu, R. Nakamura and Y. Nakato, *J. Electrochem. Soc.*, 2005, **152**, G856–G861.
- 7 K. Sayama, A. Nomura, T. Arai, T. Sugita, R. Abe, M. Yanagida, T. Oi, Y. Iwasaki, Y. Abe and H. Sugihara, *J. Phys. Chem. B*, 2006, **110**, 11352–11360.
- 8 D. K. Zhong, S. Choi and D. R. Gamelin, *J. Am. Chem. Soc.*, 2011, **133**, 18370–18377.
- 9 Q. Jia, K. Iwashina and A. Kudo, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 11564–11569.
- 10 F. F. Abdi, N. Firet and R. van de Krol, *ChemCatChem*, 2013, **5**, 490–496.
- 11 Y. Park, K. J. McDonald and K. S. Choi, *Chem. Soc. Rev.*, 2013, **42**, 2321–2337.
- 12 T. Kim and K. S. Choi, *Science*, 2014, **343**, 990–994.
- 13 W. He, R. Wang, L. Zhang, J. Zhu, X. Xiang and F. Li, *J. Mater. Chem. A*, 2015, **3**, 17977–17982.
- 14 W. He, R. Wang, C. Zhou, J. Yang, F. Li and X. Xiang, *Ind. Eng. Chem. Res.*, 2015, **54**, 10723–10730.
- 15 Y. Tang, R. Wang, Y. Yang, D. Yan and X. Xiang, *ACS Appl. Mater. Interfaces*, 2016, **8**, 19446–19455.
- 16 J. H. Kim, Y. H. Jo, J. H. Kim and J. S. Lee, *Nanoscale*, 2016, **8**, 17623–17631.
- 17 Y. Kuang, Q. Jia, H. Nishiyama, T. Yamada, A. Kudo and K. Domen, *Adv. Energy Mater.*, 2016, **6**, 1501645.



- 18 A. Iwase, S. Ikeda and A. Kudo, *Chem. Lett.*, 2017, **46**, 651–654.
- 19 Y. Kuang, T. Yamada and K. Domen, *Joule*, 2017, **1**, 290–305.
- 20 M. Rohloff, B. Anke, S. Zhang, U. Gernert, C. Scheu, M. Lerch and A. Fischer, *Sustainable Energy Fuels*, 2017, **1**, 1830–1846.
- 21 K. R. Tolod, S. Hernández and N. Russo, *Catalysts*, 2017, **7**, 13, DOI: 10.3390/catal7010013.
- 22 X. Zhang, R. Wang, F. Li, Z. An, M. Pu and X. Xiang, *Ind. Eng. Chem. Res.*, 2017, **56**, 10711–10719.
- 23 J. H. Baek, B. J. Kim, G. S. Han, S. W. Hwang, D. R. Kim, I. S. Cho and H. S. Jung, *ACS Appl. Mater. Interfaces*, 2017, **9**, 1479–1487.
- 24 A. Iwase, S. Nozawa, S. I. Adachi and A. Kudo, *J. Photochem. Photobiol., A*, 2018, **353**, 284–291.
- 25 G. Park, J. Y. Park, J. H. Seo, K. H. Oh, A. Ma and K. M. Nam, *Chem. Commun.*, 2018, **54**, 5570–5573.
- 26 R. Wang, L. Luo, X. Zhu, Y. Yan, B. Zhang, X. Xiang and J. He, *ACS Appl. Energy Mater.*, 2018, **1**, 3577–3586.
- 27 H. S. Han, S. Shin, D. H. Kim, I. J. Park, J. S. Kim, P. S. Huang, J. K. Lee, I. S. Cho and X. Zheng, *Energy Environ. Sci.*, 2018, **11**, 1299–1306.
- 28 S. Okunaka, H. Tokudome, Y. Hitomi and R. Abe, *J. Mater. Chem. A*, 2016, **4**, 3926–3932.
- 29 A. Kudo, K. Omori and H. Kato, *J. Am. Chem. Soc.*, 1999, **121**, 11459–11467.
- 30 S. M. Thalluri, C. Martinez Suarez, M. Hussain, S. Hernandez, A. Virga, G. Saracco and N. Russo, *Ind. Eng. Chem. Res.*, 2013, **52**, 17414–17418.
- 31 Y. Surendranath, D. A. Lutterman, Y. Liu and D. G. Nocera, *J. Am. Chem. Soc.*, 2012, **134**, 6326–6336.
- 32 D. Wang, R. Li, J. Zhu, J. Shi, J. Han, X. Zong and C. Li, *J. Phys. Chem. C*, 2012, **116**, 5082–5089.
- 33 J. Yang, D. Wang, H. Han and C. Li, *Acc. Chem. Res.*, 2013, **46**, 1900–1909.
- 34 D. K. Lee and K. S. Choi, *Nat. Energy*, 2018, **3**, 53–60.

