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# Preparation, characterization and catalytic application of nano- $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{(CH}_2\text{)}_3\text{OCO}_2\text{Na}$ as a novel basic magnetic nanocatalyst for the synthesis of new pyranocoumarin derivatives†

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We present a study on the synthesis, characterization and application of sodium carbonate tag silica-coated nano- $\text{Fe}_3\text{O}_4$  ( $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{(CH}_2\text{)}_3\text{OCO}_2\text{Na}$ ) as a novel and efficient heterogeneous basic catalyst. The described catalyst was fully characterized via FT-IR, X-ray diffraction (XRD), energy dispersive X-ray spectroscopy (EDS), and field emission scanning electron microscopy (FE-SEM). The reported novel magnetic nanocatalyst presents an excellent activity and catalytic performance for the synthesis of a novel series of pyranocoumarins through the reaction of dialkyl acetylenedicarboxylates and 5,7-dihydroxy coumarin derivatives at 100 °C under solvent-free conditions.

## Introduction

In recent years, there has been a pronounced tendency to utilize heterogeneous catalysts.<sup>1</sup> The reason why heterogeneous catalysts are attractive is their unique properties including easy separation, low toxicity, air tolerance, easy handling, and reusability.<sup>2–4</sup>  $\text{Fe}_3\text{O}_4$  magnetic nanoparticles (MNPs) have been widely used in design of environmentally benign heterogeneous catalysts.<sup>5</sup> The reason for the above choice is their large surface areas, good textural properties, supermagnetism, high coercivity and low Curie temperature as well as non-toxicity.<sup>6–8</sup> A striking feature of magnetic nanocatalysts is that they can be readily separated using an external magnet, which achieves a simple separation of the catalyst without filtration.<sup>9</sup> Also, they possess high potential active sites for loading of other functional groups.<sup>10</sup> To prevent  $\text{Fe}_3\text{O}_4$  nanoparticles from undergoing oxidation in an air atmosphere and in order to increase the surface area and simplify the surface functionalization, a protective shell of silica can be formed onto their surface.<sup>11,12</sup>

Pyranocoumarins with excellent chemical and physical characteristics have acquired substantial attention and show great practical values in many fields, such as medicine discovery, dye chemistry, and materials chemistry. Structural diversity associated with pyranocoumarins has resulted in a large number of new molecular entities, which have been found to be useful as anti-cancer, anti-oxidant, anti-inflammatory, anti-allergic, hepatoprotective, anti-viral, anti-

carcinogenic agents, enzyme inhibitor, and precursor of toxic substances.<sup>13–16</sup> In view of the biological, industrial, and synthetic importance of pyranocoumarins, several methods for their synthesis have been reported.<sup>17–19</sup> Although most of these processes have distinct advantages, the use of high temperatures, environmentally harmful catalysts, harsh reaction conditions, long reaction times and large quantity of volatile organic solvents limit the use of these methods. Therefore, the search for the possibility synthesis of these compounds under mild reaction conditions with recoverable effective heterogeneous catalyst is still in high demand.

In continuation of our efforts to design, synthesis and application of novel nanocatalyst,<sup>20,21</sup> we have immobilized sodium carbonate on silica-coated  $\text{Fe}_3\text{O}_4$  magnetic nanoparticles ( $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{(CH}_2\text{)}_3\text{OCO}_2\text{Na}$ ) and then investigated its performance as novel strong, recoverable, and stable basic nanocatalyst for synthesis of new pyranocoumarin derivatives.

## Experimental

All chemical reagents were purchased from Merck and Sigma Aldrich companies and used without further purification. Melting points were determined in open capillaries using an electro-thermal KSB1N-apparatus (Krüss, Germany). FT-IR spectra were obtained with potassium bromide pellets in the range 400–4000  $\text{cm}^{-1}$  with a FT-IR-680 plus spectrometer (JASCO, Japan).  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra were recorded on a FT-NMR Bruker Avance Ultra Shield Spectrometer (Bruker, USA) at 400 and 100 MHz, respectively. X-ray diffraction pattern of the prepared catalyst was obtained using D<sub>8</sub> ADVANCE X-ray diffraction using Co-K $\alpha$  radiation ( $\lambda = 1.7890 \text{ \AA}$ ) (Bruker, Germany). Energy dispersive spectroscopy (EDS) was performed using TESCAN Vega model

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instrument. The morphology of the particles was studied by Field Emission Scanning Electron Microscopy (FE-SEM) in a MIR-A3TESCAN-XMU FE-SEM instrument.

### Synthesis of $\text{Fe}_3\text{O}_4$ MNPs

A mixture of  $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$  (2.3 g, 8.7 mmol) and  $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$  (0.86 g, 4.3 mmol) was dissolved in deionized water (100 mL). The solution was heated to 90 °C under nitrogen atmosphere and stirred about 30 min. Subsequently, sodium hydroxide solution (10 mL, 25%) was added dropwise to the solution until the brown color solution turned out to the black. After approximately 1 h, the black precipitate isolated in a magnetic field from the reaction mixture, repeatedly washed with deionized water several times to remove the remaining impurities.<sup>22</sup>

### Preparation of $\text{Fe}_3\text{O}_4@\text{SiO}_2$

Dried  $\text{Fe}_3\text{O}_4$  nanoparticles (0.5 g) was suspended in a mixture of ethanol (20 mL) and  $\text{NH}_3 \cdot \text{H}_2\text{O}$  (5 mL, 25%) followed by the addition of tetraethoxysilane (TEOS) (0.3 g) to the solution and the mixture was ultrasonicated for 2 h. Next, the mixture was degassed and stirred for 24 h.<sup>23</sup>

### Procedure for the synthesis of $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{Cl}$

The prepared  $\text{Fe}_3\text{O}_4@\text{SiO}_2$  (0.5 g) was dispersed in dry toluene (80 mL) by sonication for 15 min, then, 3-chloropropyl-triethoxysilane (0.121 g, 0.5 mmol) was added to the mixture and heated to 100 °C under reflux condition and  $\text{N}_2$  atmosphere. After 12 h,  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{Cl}$  was separated magnetically and washed with deionized water and ethanol.<sup>24</sup>

### Preparation of $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$

$\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{Cl}$  (0.5 g) was dissolved in DMSO (50 mL) by sonication. Then  $\text{Na}_2\text{CO}_3$  (1 g) was added to the above mixture and heated to the 90 °C under reflux condition for 24 h. Next, the obtained black precipitate was separated by a normal magnet and washed with distilled water and ethanol and dried overnight under vacuum at 60 °C.

### General procedure for the synthesis of pyranocoumarins 4

In a round bottom flask, a mixture of 5,7-dihydroxy coumarins 2 (1 mmol), dialkyl acetylene dicarboxylate 3 (1.5 mmol) and prepared magnetic nanocatalyst 1 (0.0005 g) was heated to 100 °C. The progress of the reaction was monitored by TLC (*n*-hexane : EtOAc/3 : 1 (v/v)). After completion of the reaction which was distinguished by disappearing of starting materials' spots on TLC, the reaction mixture was diluted by methanol and the catalyst was easily separated magnetically using an external magnet. The desired product was purified using column chromatography (using ethylacetate : *n*-hexane, 3 : 1 as mobile phase).

### General procedure for the recovery of nanocatalyst 1

5,7-Dihydroxy-4-methyl coumarin (1 mmol), dimethyl acetylene-dicarboxylates (1.5 mmol), and  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  (0.0005 g) were stirred at free-solvent condition in 100 °C for 6 h. After completion of the reaction, methanol (5 mL) was added to

the reaction mixture and the catalyst was separated by an external magnet. It was washed three times by EtOH (10 mL) and deionized water (10 mL) and then dried under vacuum at 80 °C.

**Methyl 5-hydroxy-4-methyl-2,8-dioxo-2*H,8H*-pyrano[2,3-*f*]chromene-10-carboxylate (4a).** Yellow crystals (0.278 g), IR (KBr):  $\nu_{\text{max}} = 3444, 2935, 1739, 1675, 1623, 1438, 1235, 623 \text{ cm}^{-1}$ .  $^1\text{H}$  NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 10.40$  (s, 1H), 6.65 (s, 1H), 6.40 (s, 1H), 5.97 (s, 1H), 3.51 (s, 3H), 1.81 (s, 3H).  $^{13}\text{C}$  NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 167.8, 161.6, 160.1, 158.9, 156.0, 152.9, 151.1, 147.1, 120.7, 118.7, 113.2, 111.2, 103.0, 50.4, 27.8$ . Anal. calcd for  $\text{C}_{15}\text{H}_{10}\text{O}_7$ : C, 59.61; H, 3.34. Found: C, 59.69; H, 3.29.

**Methyl 4-(chloromethyl)-5-hydroxy-2,8-dioxo-2*H,8H*-pyrano[2,3-*f*]chromene-10-carboxylate (4b).** Creamy crystals (0.285 g), IR (KBr):  $\nu_{\text{max}} = 3424, 2956, 1723, 1660, 1617, 1434, 1278, 649 \text{ cm}^{-1}$ .  $^1\text{H}$  NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 10.75$  (s, 1H), 6.76 (s, 1H), 6.25 (s, 1H), 6.08 (s, 1H), 4.21 (s, 2H), 3.38 (s, 3H).  $^{13}\text{C}$  NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 167.8, 161.8, 160.8, 160.1, 157.6, 154.5, 152.8, 144.5, 121.5, 118.7, 114.0, 111.7, 106.8, 50.4, 49.0$ . Anal. calcd for  $\text{C}_{15}\text{H}_9\text{ClO}_7$ : C, 53.51; H, 2.69. Found: C, 53.54; H, 2.67.

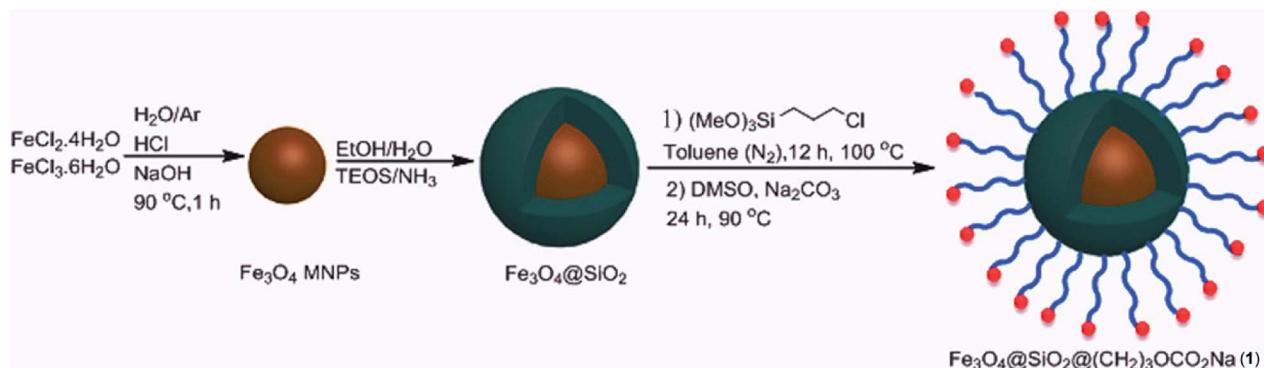
**Ethyl 5-hydroxy-4-methyl-2,8-dioxo-2*H,8H*-pyrano[2,3-*f*]chromene-10-carboxylate (4c).** Pale yellow crystals (0.274 g), IR (KBr):  $\nu_{\text{max}} = 3439, 2972, 1738, 1658, 1622, 1439, 1281, 646 \text{ cm}^{-1}$ .  $^1\text{H}$  NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 10.48$  (s, 1H), 6.87 (s, 1H), 6.57 (s, 1H), 6.10 (s, 1H), 6.07 (q, 2H, *J* = 8 Hz), 2.10 (s, 3H), 1.21 (t, 3H, *J* = 4 Hz).  $^{13}\text{C}$  NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 168.8, 161.9, 160.8, 160.1, 160.0, 155.5, 150.1, 147.1, 140.7, 119.0, 116.0, 112.0, 109.2, 102.6, 62.2, 24.2, 14.1$ . Anal. calcd for  $\text{C}_{16}\text{H}_{12}\text{O}_7$ : C, 60.76; H, 3.82. Found: C, 60.71; H, 3.90.

**Ethyl 4-(chloromethyl)-5-hydroxy-2,8-dioxo-2*H,8H*-pyrano[2,3-*f*]chromene-10-carboxylate (4d).** Yellow crystals (0.280 g), IR (KBr):  $\nu_{\text{max}} = 3440, 2963, 1748, 1661, 1633, 1445, 1279, 639 \text{ cm}^{-1}$ .  $^1\text{H}$  NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 10.76$  (s, 1H), 6.63 (s, 1H), 6.33 (s, 1H), 6.25 (s, 1H), 4.80 (s, 1H), 4.31 (q, 2H, *J* = 7 Hz), 1.32 (t, 3H, *J* = 4 Hz).  $^{13}\text{C}$  NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 169.4, 163.4, 160.6, 157.8, 157.2, 152.8, 151.4, 147.7, 119.3, 117.4, 114.7, 110.7, 105.8, 62.4, 56.2, 15.0$ . Anal. calcd for  $\text{C}_{16}\text{H}_{11}\text{ClO}_7$ : C, 54.80; H, 3.16. Found: C, 54.78; H, 3.20.

**Ethyl 11-hydroxy-2,6-dioxo-7,8,9,10-tetrahydro-2*H,6H*-benzo[c]pyrano[2,3-*h*]chromene-4-carboxylate (4e).** White crystals (0.256 g), IR (KBr):  $\nu_{\text{max}} = 3437, 2977, 1743, 1665, 1641, 1458, 1277, 650 \text{ cm}^{-1}$ .  $^1\text{H}$  NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 10.50$  (s, 1H), 6.63 (s, 1H), 6.16 (s, 1H), 4.07 (q, 2H, *J* = 8 Hz), 1.22 (t, 3H, *J* = 4 Hz).  $^{13}\text{C}$  NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 168.3, 162.5, 159.8, 155.9, 152.3, 150.6, 149.6, 147.3, 122.0, 118.3, 116.1, 111.1, 103.7, 61.6, 28.6, 28.2, 22.0, 21.6, 14.1$ . Anal. calcd for  $\text{C}_{19}\text{H}_{16}\text{O}_7$ : C, 64.04; H, 4.53. Found: C, 64.10; H, 4.55.

**Methyl 11-hydroxy-2,6-dioxo-7,8,9,10-tetrahydro-2*H,6H*-benzo[c]pyrano[2,3-*h*]chromene-4-carboxylate (4f).** Creamy crystals (0.249 g), IR (KBr):  $\nu_{\text{max}} = 3440, 2975, 1738, 1668, 1646, 1462, 1281, 653 \text{ cm}^{-1}$ .  $^1\text{H}$  NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 10.49$  (s, 1H), 6.94 (s, 1H), 6.21 (s, 1H), 4.22 (s, 1H), 1.11–0.98 (m, 8H).  $^{13}\text{C}$  NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta = 165.0, 161.7, 159.4, 159.1, 153.1, 151.7, 147.9, 147.5, 118.3, 117.9, 116.8, 113.7, 99.9, 52.5, 29.5, 24.7, 23.3, 22.3$ . Anal. calcd for  $\text{C}_{19}\text{H}_{16}\text{O}_7$ : C, 64.04; H, 4.53. Found: C, 64.09; H, 4.60.





Scheme 1 Synthesis of  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  (1).

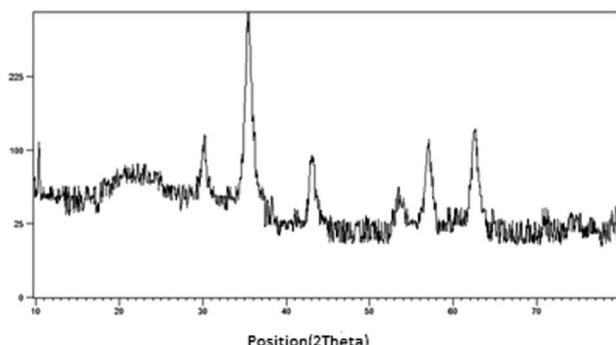


Fig. 1 The XRD patterns of  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$ .

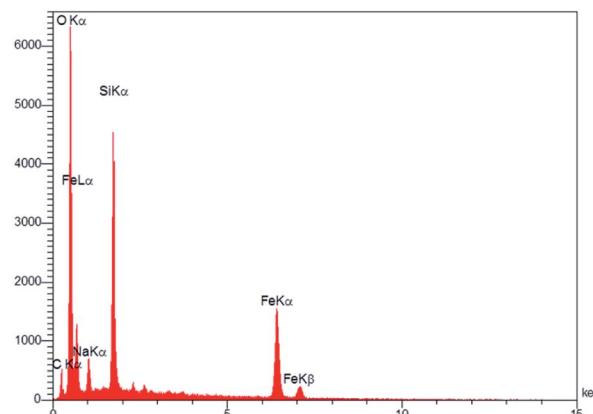


Fig. 3 EDS analysis of  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$ .

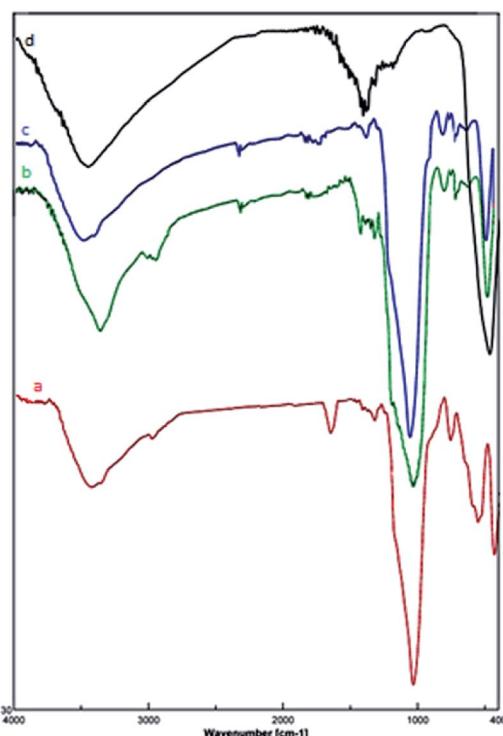


Fig. 2 FT-IR spectra of (a)  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$ , (b)  $\text{Fe}_3\text{O}_4@\text{SiO}_2(\text{CH}_2)_3\text{Cl}$ , (c)  $\text{Fe}_3\text{O}_4@\text{SiO}_2$ , and (d)  $\text{Fe}_3\text{O}_4$  MNPs.

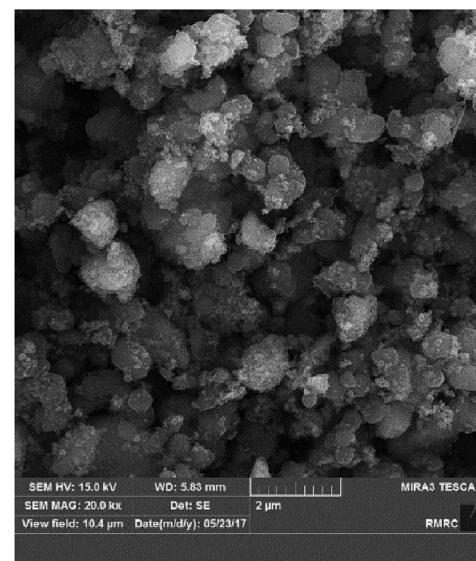


Fig. 4 FE-SEM analysis of nanocatalyst 1.



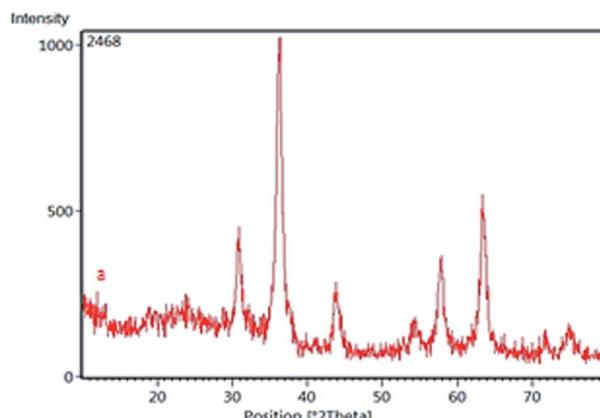
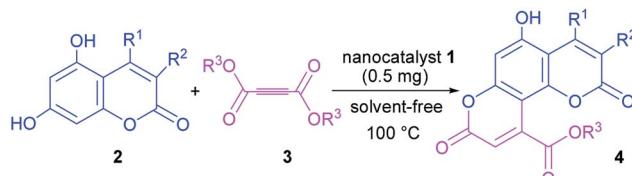


Fig. 5 The XRD patterns of recycled catalyst.

Scheme 2  $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{(CH}_2\text{)}_3\text{OCO}_2\text{Na}$ -catalyzed synthesis of pyranocoumarins 4.

(s, 1H), 3.70 (s, 3H).  $^{13}\text{C}$  NMR (100 MHz,  $\text{DMSO}-d_6$ ):  $\delta$  = 170.2, 159.4, 158.9, 153.8, 152.5, 152.3, 146.5, 142.6, 129.3, 128.3, 127.1, 118.9, 117.7, 113.0, 110.4, 103.6, 52.7. Anal. calcd for  $\text{C}_{20}\text{H}_{12}\text{O}_7$ : C, 65.94; H, 3.32. Found: C, 65.89; H, 3.36.

## Results and discussion

Due to reasonable needs to clean and green heterogeneous basic catalysts, the magnetic nanocatalyst  $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{(CH}_2\text{)}_3\text{OCO}_2\text{Na}$

Table 1 Screening for the model reaction

Entry	Solvent	Catalyst	$T$ (°C)	Time (h)	Yield <sup>a</sup> (%)
1	—	—	25	24	5
2	Toluene	—	25	24	8
3	Toluene	—	Reflux	12	12
4	Toluene	NaOH (5 mol%)	Reflux	12	25
5	Toluene	$\text{Na}_2\text{CO}_3$ (5 mol%)	Reflux	12	20
6	$\text{CHCl}_3$	NaOH (5 mol%)	Reflux	12	15
7	$\text{CHCl}_3$	$\text{Na}_2\text{CO}_3$ (5 mol%)	Reflux	12	15
8	$\text{CH}_2\text{Cl}_2$	$\text{Na}_2\text{CO}_3$ (5 mol%)	Reflux	12	14
9	MeOH	NaOH (5 mol%)	Reflux	12	10
10	MeOH	$\text{Na}_2\text{CO}_3$ (5 mol%)	Reflux	12	12
11	EtOAc	NaOH (5 mol%)	Reflux	12	25
12	EtOAc	$\text{Na}_2\text{CO}_3$ (5 mol%)	Reflux	12	29
13	EtOAc	Catalyst 1 (0.1 mg)	Reflux	6	70
14	EtOAc	Catalyst 1 (0.3 mg)	Reflux	6	80
15	EtOAc	Catalyst 1 (0.5 mg)	Reflux	6	92
16	EtOAc	Catalyst 1 (0.7 mg)	Reflux	6	90
17	Toluene	Catalyst 1 (0.5 mg)	Reflux	6	70
18	$\text{CH}_2\text{Cl}_2$	Catalyst 1 (0.5 mg)	Reflux	6	75
19	$\text{CHCl}_3$	Catalyst 1 (0.5 mg)	Reflux	6	80

<sup>a</sup> Isolated yield.

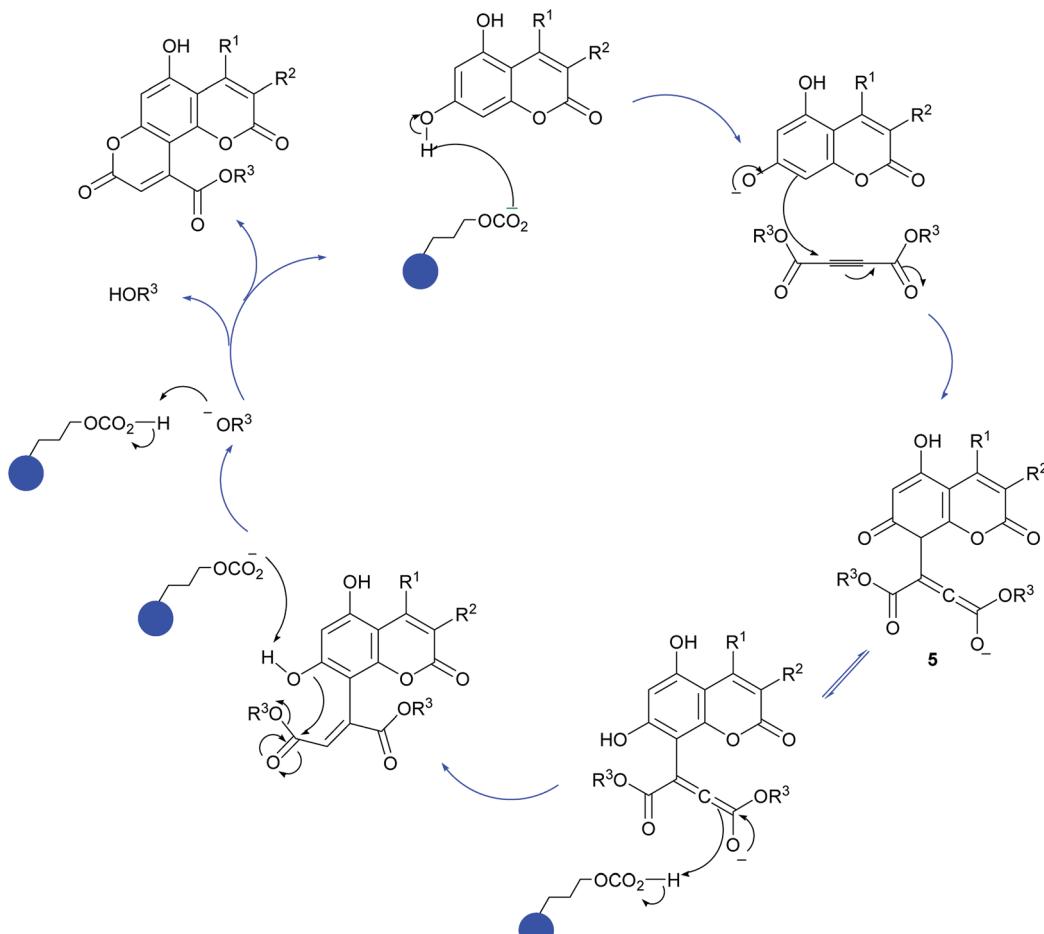
(1) was synthesized following the procedure shown in Scheme 1. The structure of prepared nanocatalyst 1 was studied and fully characterized using FT-IR, EDS, XRD, and FE-SEM analysis. These results provided the evidences that the expected structure was successfully achieved. Fig. 1 presents the X-ray diffraction (XRD) patterns of the magnetic nanocatalyst 1. As can be seen, six

Table 2 Preparation of pyranocoumarins 4 in the presence of  $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{(CH}_2\text{)}_3\text{OCO}_2\text{Na}$ <sup>a</sup>

Entry	Product <sup>b</sup>	Time (h)	Yield <sup>c</sup> (%)	Mp (°C)
4a		6	92	243–244
4b		7	85	278–279
4c		5	87	261–263
4d		7	80	272–273
4e		8	75	237–239
4f		6	70	269–270
4g		9	90	277–278

<sup>a</sup> Reaction conditions: 5,7-dihydroxycoumarins 2 (1 mmol), dialkyl acetylenedicarboxylates 3 (1.5 mmol) and nanocatalyst 1 (0.0005 g) at 100 °C. <sup>b</sup> All products are novel and characterized by IR,  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, and elemental analysis. <sup>c</sup> Isolated yield.





Scheme 3 Proposed mechanism for the synthesis of pyranocoumarin 4.

characteristic peaks over  $2\theta$  range from 10 to  $80^\circ$  at 30.86, 35.36, 43.96, 54.36, 57.26, and 62.56 corresponding to the (220), (311), (400), (422), (511) and (440) crystal planes indicate the cubic spinal structure of  $\text{Fe}_3\text{O}_4$  MNPs.<sup>25</sup> The characteristic peaks confirming the presence of  $-\text{OCO}_2\text{Na}$  is appeared around 10, 34, 36 and 43 ( $2\theta$ ) which some of them is overlapped by the  $\text{Fe}_3\text{O}_4$  peaks. However, the appearance of a peak at 10.53 confirms the connection of  $-\text{OCO}_2\text{Na}$  in nanocatalyst.<sup>26</sup> Additionally, a broad peak at 13–28 in XRD pattern of catalyst indicates the existence of amorphous silica in the structure of the catalyst.<sup>27</sup>

The structure of prepared nanocatalyst **1** was confirmed by the FT-IR spectra (Fig. 2). The sample exhibits a peak at  $632\text{ cm}^{-1}$  band that is due to the stretching vibration mode associated to the Fe–O absorption band. The stretch found at  $466\text{ cm}^{-1}$  is related to the presence of Fe–O–Si bond in the sample.<sup>28</sup> Furthermore, the spectra present the O–H stretching vibrational around  $3431\text{ cm}^{-1}$  and two absorption peaks at  $798\text{ cm}^{-1}$  and  $1084\text{ cm}^{-1}$  which corresponds to the symmetric and asymmetric stretching vibration of Si–O.<sup>29</sup> The band at  $2928\text{ cm}^{-1}$  is attributed to the alkyl chain  $-\text{CH}_2$ .<sup>30</sup> The presence of the  $-\text{OCO}_2\text{Na}$  group is confirmed by the bands at  $1636\text{ cm}^{-1}$  and  $1407\text{ cm}^{-1}$ .<sup>26</sup>

The prepared magnetic nanocatalyst was analyzed by using an energy dispersive spectrometer (EDS). According to the

Fig. 3, it is seen that  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  contains all expected elemental cases including Si, O, Fe, C and Na.

The FE-SEM image of  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  was shown in Fig. 4. The image demonstrates uniform-size particles with near spherical morphology. As it comes from FE-SEM analysis the average diameter of obtained nanoparticles is around 25 nm.

To test the stability of the catalyst structure, the recycled nano catalyst was examined by XRD analysis; the diffraction patterns and relative intensities of all peaks matched well with those of the primary catalyst (Fig. 5).

Having successfully prepared the new nanocatalyst **1**, further studies were performed for the synthesis of a novel class of pyranocoumarins **4** via the condensation of 5,7-dihydroxycoumarins **2** and dialkyl acetylenedicarboxylates **3** (Scheme 2). The structure of product was confirmed by FT-IR,  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, and elemental analysis. Initially, 5,7-dihydroxycoumarins (**2**) was synthesized according to the reported procedure.<sup>31</sup>

In order to obtain the optimal experimental conditions, we set up a model reaction between 5,7-dihydroxy-4-methyl coumarin and dimethyl acetylenedicarboxylate (DMAD). The model reaction was conducted in the presence of several homogeneous and heterogeneous catalysts at various conditions. It was observed that the best result was achieved using



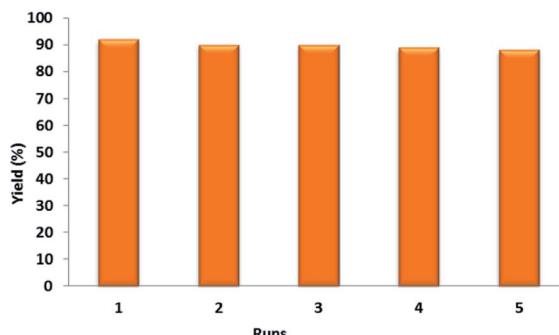


Fig. 6 Reusability study of nanocatalyst 1 in the synthesis of 4a at 100 °C under solvent-free conditions.

0.5 mg nanocatalyst 1 at 100 °C under solvent-free conditions. The results are summarized in Table 1.

After establishing the optimal conditions, the efficiency of the catalyst was further evaluated with a series of 5,7-dihydroxy coumarins. Furthermore, the procedure worked well for diethyl acetylenedicarboxylate as well as DMAD (Table 2).

A plausible mechanism for the formation of pyranocoumarin derivatives catalyzed by  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  is shown in Scheme 3. According to the reaction pathway, in the beginning, the OH group on the C-7 of coumarin attacks the dialkyl acetylenedicarboxylate to give intermediate 5. The role of catalyst comes in this stage where it removes the hydrogen of OH group and facilitates the attack of coumarin to the dialkyl acetylenedicarboxylate. Unstable intermediate 5 regains the previous stable aromatic form *via* keto-enol equilibrium. Then the dehydrogenated oxygen of position 7 attack the carbonyl group of dialkyl acetylenedicarboxylate portion and by removing of the HOR<sup>3</sup> forms the desired product.

To examine the reusability of the prepared nanocatalyst, after completion of the model reaction, the reaction mixture was diluted with methanol and the catalyst was magnetically removed from it. Afterwards, the isolated catalyst was washed by ethanol and deionized water and dried. The separated catalyst was directly used for the next run under the same condition. The results showed that the catalyst can be reused, without deactivation, even after five cycles (Fig. 6).

## Conclusions

In summary, we have reported  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  as a new and efficient basic nanocatalyst. It was fully characterized by FT-IR, FE-SEM, EDS and XRD analysis. Furthermore, herein, the first application of  $\text{Fe}_3\text{O}_4@\text{SiO}_2@(\text{CH}_2)_3\text{OCO}_2\text{Na}$  as a green, convenient and recyclable catalyst to the synthesis of new pyranocoumarins is successfully examined. The introduced reaction reveals broad substrate scope, good functional group tolerance, solvent-free condition and good to excellent yields. The structure of products was confirmed by IR, <sup>1</sup>H NMR, <sup>13</sup>C NMR, and elemental analysis.

## Conflicts of interest

There are no conflicts to declare.

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## Notes and references

- J. Zhao and R. Jin, *Nano Today*, 2018, **18**, 86, <https://www.sciencedirect.com/science/article/pii/S1748013217304395-!>.
- A. D. Murkute, J. E. Jackson and D. J. Miller, *J. Catal.*, 2011, **278**, 189.
- M. Farahi, M. Davoodi and M. Tahmasebi, *Tetrahedron Lett.*, 2016, **57**, 1582.
- M. L. Kantam, M. Roy, S. Roya, B. Sreedhara and R. L. De, *Catal. Commun.*, 2008, **9**, 2226.
- J. Safari and Z. Zarnegar, *RSC Adv.*, 2015, **5**, 17738.
- M. S. Najafinejad, P. Mohammadi, M. M. Afsahi and H. Sheibani, *J. Mol. Liq.*, 2018, **262**, 248.
- M. Nasr-Esfahani, Z. Rafiee, M. Montazerozohori and H. Kashi, *RSC Adv.*, 2016, **6**, 47298.
- R. Sharma, S. Dutta, S. Sharma, R. Zboril, R. Varma and M. Gawande, *Green Chem.*, 2016, **18**, 3184.
- M. Nasr-Esfahani, S. J. Hoseini, M. Montazerozohori, R. Mehrabi and H. Nasrabadi, *J. Mol. Catal. A: Chem.*, 2014, **382**, 99.
- A. H. Lu, E. L. Salabas and F. Schuth, *Angew. Chem., Int. Ed.*, 2007, **46**, 1222.
- R. G. Chaudhuri and S. Paria, *Chem. Rev.*, 2012, **112**, 2373.
- L. Tan, X. Zhang, Q. Liu, J. Wang, Y. Sun, X. Jing, J. Liu, D. Songa and L. Liuc, *Dalton Trans.*, 2015, **44**, 6909.
- E. C. Gaudino, S. Tagliapietra, K. Martina, G. Palmisano and G. Cravotto, *RSC Adv.*, 2016, **6**, 46394.
- I. Kostova, *Curr. Med. Chem.: Anti-Cancer Agents*, 2005, **5**, 29.
- Y. Liu, J. Zhu, J. Qian, B. Jiang and Z. Xu, *J. Org. Chem.*, 2011, **76**, 9096.
- P. Magiatis, E. Melliou, A. Skaltsounis, S. Mitaku, S. Léonce, P. Renard, A. Pierré and G. Atassi, *J. Nat. Prod.*, 1998, **61**, 982.
- A. Kumar, R. Maurya, S. Sharma, P. Ahmad, A. B. Singh, G. Bhatia and A. K. Srivastava, *Bioorg. Med. Chem. Lett.*, 2009, **19**, 6447.
- M. A. Terzidis, E. Dimitriadou, C. A. Tsoleridis and J. Stephanidou-Stephanatou, *Tetrahedron Lett.*, 2009, **50**, 2174.
- S. Yaragorla, P. L. Saini and G. Singh, *Tetrahedron Lett.*, 2015, **56**, 1649.
- M. Farahi, B. Karami, R. Keshavarz and F. Khosravian, *RSC Adv.*, 2017, **7**, 46644.
- S. Akrami, B. Karami and M. Farahi, *RSC Adv.*, 2017, **7**, 34315.
- M. C. Mascolo, Y. Pei and T. A. Ring, *Materials*, 2013, **6**, 5549.
- Q. Zhang, H. Su, J. Luo and Y. Wei, *Catal. Sci. Technol.*, 2013, **3**, 235.
- A. A. Elmekawy, J. B. Sweeney and D. R. Brown, *Catal. Sci. Technol.*, 2015, **5**, 690.

25 Z. Rashid, H. Naeimi, A. H. Zarnani, M. Nazari, M. Nejadmoqaddam and R. Ghahremanzadeh, *RSC Adv.*, 2016, **6**, 36840.

26 K. Eskandari, B. Karami and S. Khodabakhshi, *Catal. Commun.*, 2014, **54**, 124.

27 M. A. Nasser, B. Zakerinasab and M. M. Samieadel, *RSC Adv.*, 2014, **4**, 41753.

28 F. Ahangaran, A. Hassanzadehand and S. Nouri, *Int. Nano Lett.*, 2013, **3**, 23.

29 P. Innocenzi, *J. Non-Cryst. Solids*, 2003, **316**, 309.

30 M. A. Zolfigol and R. Ayazi-Nasrabadi, *RSC Adv.*, 2016, **6**, 69595.

31 B. Karami and M. Kiani, *Catal. Commun.*, 2011, **14**, 62.

