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Insight into the electronic and thermodynamic properties of NbSi₂ from first-principles calculations†

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The electronic and thermodynamic properties of NbSi₂ with four structures (C40, C11_b, C54 and C49) were studied in terms of first-principle calculations. The band structure and density of states of four NbSi₂ are calculated. Those disilicides show electronic properties because the band gap between the conduction band and the valence band is near the Fermi level. Their metallic character is mainly due to the strong metallic interaction between the Nb-4d state and Si-3p state. The formed Si-Si covalent bond is mainly concentrated in the valence band. The valence electron configuration of Nb-Si and Si-Si bonds is also explored. Besides, we study the thermodynamic properties of NbSi₂ as a function of temperature. The results indicate that the C54 structure has the best thermal stability in the obtained phases. In particular, the Debye temperature and heat capacity of the C54 structure are 547.8 K and 142.7 J mol⁻¹ K⁻¹, respectively. The calculated phonon DOS provides the explanation that the nature of the thermodynamic properties is mainly derived from the vibration of Si atoms.

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1. Introduction

Transition metal disilicides (TMSi₂) have spurred significant interest because of their high electrical conductivity, thermal stability, excellent high-temperature oxidation resistance, strength, *etc.*^{1–5} These properties can be exploited in a variety of applications in the microelectronics industry, such as ohmic contacts, interface diffusion barriers and interconnections.^{6–8} In addition to these applications, TMSi₂ also exhibits great advantages in the field of lithium-ion batteries. For example, a number of transition metal disilicides, such as MoSi₂, CrSi₂, TiSi₂, NiSi₂, TaSi₂ and VSi₂, are used as anode materials or anode coating materials for lithium-ion batteries.^{9–12}

Previous studies have reported that there are four typical crystal structures in transition metal disilicides, *i.e.*, hexagonal C11_b (space group *I4/mmm*, no. 139), tetragonal C40 (space group *P6₂22*, no. 180), orthorhombic C54 (space group *Fddd*, no. 70) and orthorhombic C49 (space group *Cmcm*, no. 63) structures.^{13–16} MoSi₂ and WSi₂ have C11_b and C40 structures under different temperature, and the differences between C11_b structure and C40 structure in performance are great.^{17,18} CrSi₂, TaSi₂, NbSi₂ and VSi₂ are the hexagonal C40 structure while ZrSi₂ and HfSi₂ have C54 structure.^{19–22} TiSi₂ has the C49 structure and high pressure can result in phase transition from C49 structure to C54 structure.²³

Among these transition metal disilicides, NbSi₂ is known as the hexagonal C40 structure.^{24,25} In recent years, the physical properties of hexagonal NbSi₂ have been widely investigated.^{26–29} For example, Xu *et al.* have studied the electronic, elastic and thermodynamic properties of NbSi₂ under high temperature and pressure using first-principle calculations.³⁰ Gottlieb *et al.* have investigated some transport properties of NbSi₂ single crystal.³¹ Specifically, our recent studies indicated that vacancy can influence the structural and elastic properties of NbSi₂.³² In addition, Lasjaunias *et al.* have shown that the obtained Debye temperature of hexagonal NbSi₂ was 579 ± 10 K from low-temperature specific heat measurement.³³ As mentioned by Antonov *et al.*, the optical properties of NbSi₂ are studied theoretically and experimentally.³⁴ However, the other phases of NbSi₂ are rarely studied until now. Hence, we attempt to design three new phases (C11_b, C54 and C49) for NbSi₂. Importantly, our recent work has indicated these three phases for NbSi₂ are stables at the ground state. Unfortunately, the electronic and thermodynamic properties of NbSi₂ with these new structures are unknown.

Herein we investigate the electronic and thermodynamic properties of NbSi₂ with four crystal structures (C40, C11_b, C54 and C49) using the first-principle calculation. Firstly, we compare the energy band structure and density of state of these new phases with that of C40-NbSi₂. To better understand the origin of internal bonding, we study the valence electron configuration of Nb-Si bond and Si-Si bond. Then, their thermodynamic properties (including Debye temperature, heat capacity, enthalpy and entropy) are calculated and discussed in detail. Finally, in order to explore the nature of thermodynamic properties of these phases, we also calculate and analyze their phonon DOSs.

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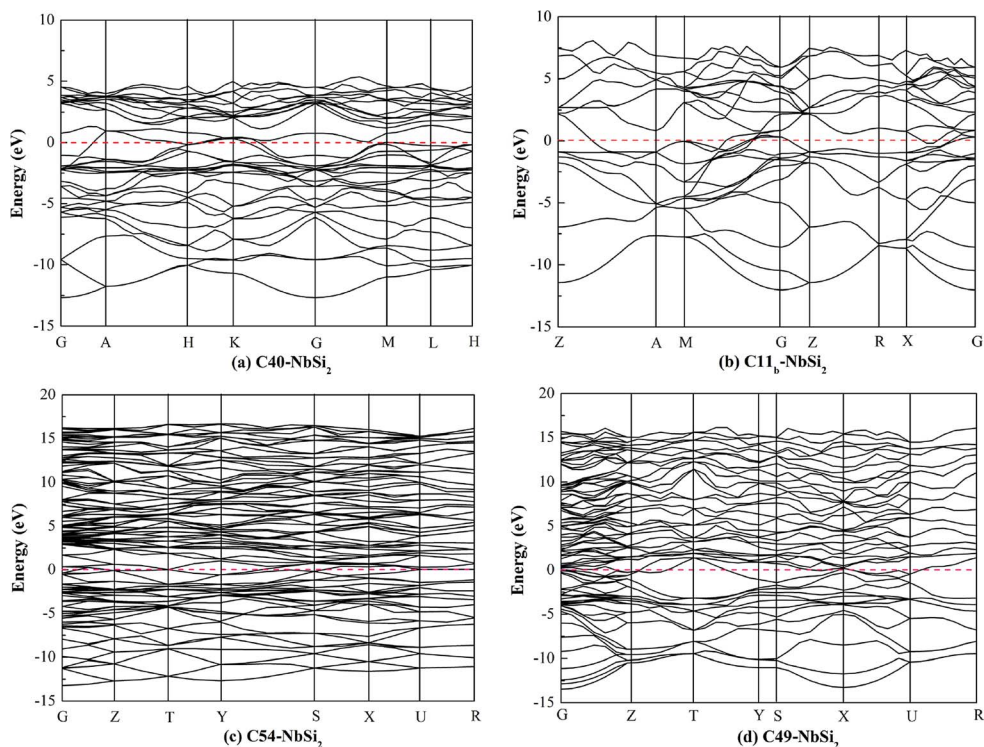


Fig. 1 Calculated band structure of various NbSi₂ phases.

2. Models and theoretical methods

Our calculations were carried out within the density function theory (DFT), as implemented in the CASTEP packages.³⁵ The

exchange–correlation functions were treated by using the Perdew–Burke–Ernzerhof (PBE) generalized gradient approximation (GGA) functional.³⁶ The valence electronic configurations of Nb and Si atoms were selected as 4p⁶4d⁴5s¹ and 3s²3p²,

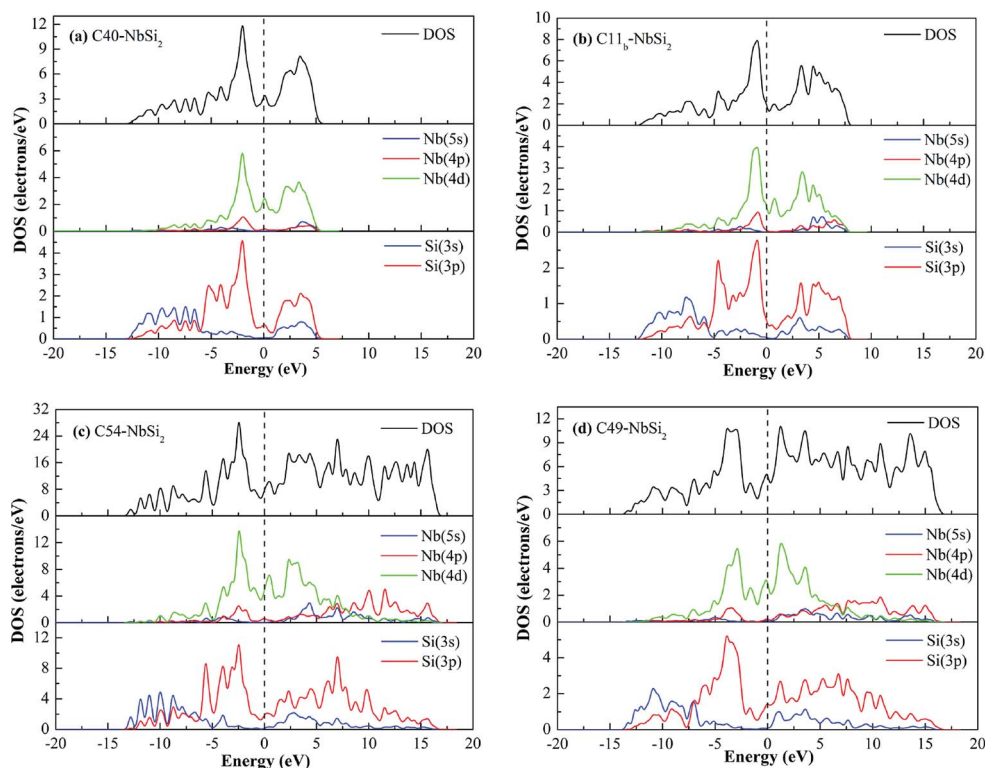


Fig. 2 Calculated total and partial DOS profile of various NbSi₂ phases.



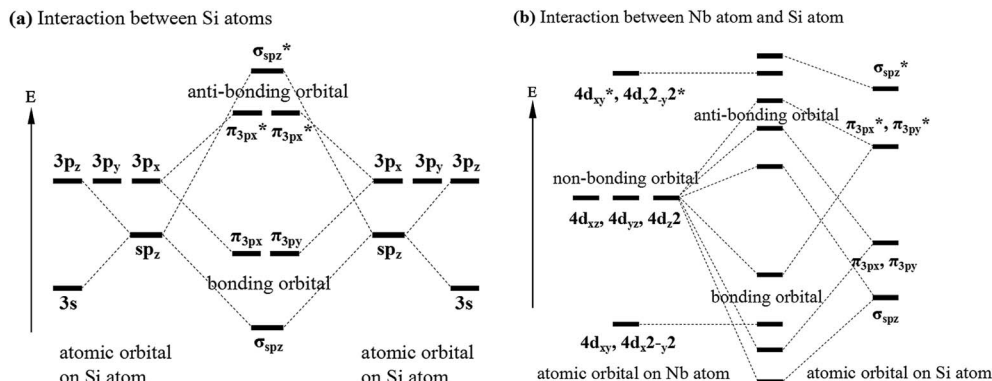


Fig. 3 Valence electron configuration for (a) interaction between Si atoms and (b) interaction between Nb atom and Si atom.

respectively. The cutoff energy with plane-wave basis was 450 eV for all structures. The Monkhorst–Pack k -point grids of $17 \times 17 \times 10$, $20 \times 20 \times 8$, $8 \times 14 \times 8$ and $18 \times 4 \times 18$ were chosen for C40, C11_b, C54 and C49 structures, respectively. In the process of structural optimization, the energies, maximum ionic displacement and maximum stress were converged within 1×10^{-5} eV per atom, 1×10^{-3} Å and 0.05 GPa, respectively.

3. Results and discussion

To investigate the electronic properties of NbSi₂ with four structures, the calculated band structure along the high symmetric direction within the Brillouin zone have been plotted in Fig. 1. The pink dashed line at zero represents the Fermi level (E_F). It is found that the high symmetric direction of C40–NbSi₂ is $G \rightarrow A \rightarrow H \rightarrow K \rightarrow G \rightarrow M \rightarrow L \rightarrow H$ direction while that of C11_b–NbSi₂ is $Z \rightarrow A \rightarrow M \rightarrow G \rightarrow Z \rightarrow R \rightarrow X \rightarrow G$ direction. For orthorhombic C54– and C49–NbSi₂, their high symmetric direction is $G \rightarrow Z \rightarrow T \rightarrow Y \rightarrow S \rightarrow X \rightarrow U \rightarrow R$. Furthermore, as with C40–NbSi₂ (see Fig. 1(a)), although there is no band overlap between the valence band (VB) and the conduction band (CB). In particular, the energy of electrons at the top of the valence band is higher than that of the bottom of the conduction band. Thus, the electron in valence band can enter the conduction band through the Fermi level so as to make C40–NbSi₂ exhibit metallic behavior. For the other three structures, there is the band overlap between the valence band and the conduction band, indicating that they have metallic behavior. Compared with C40–NbSi₂, we can see that the predicted three phases exhibit the better electronic properties.

To gain more insight into the origin of electronic properties of various NbSi₂ phases, the total and partial density of states (DOSs) are calculated and displayed in Fig. 2. It is clearly observed that the total DOS of four structures at the Fermi level are positive, indicating that they exhibit metallic behavior in their crystal phases. Furthermore, we find that the DOS of these phases are characterized by a small peak close to the Fermi level. The hybridized orbital of this peak are composed of Nb-4d and Si-3p orbital. This is to say, there is the formation of Nb–Si bond at the Fermi level. Hence, we can conclude that their metallic behavior may be attributed to the strong metallic

interaction between Nb atom and Si atom. Moreover, the existence of a pseudogap, *i.e.*, a deep valley through the Fermi level, differentiates the bonding state from the anti-bonding/non-bonding state.³⁷

In the valence band zone, we can see that the DOS of C40–, C11_b– and C54–NbSi₂ located from about -5 – 0 eV mainly come from Nb-4d and Si-3p orbital. Moreover, the hybridized orbital in the range from about -13 eV to -5 eV is mainly composed of Si-3s and Si-3p orbital, which will drive the formation of Si–Si covalent bond. However, for C49–NbSi₂, its DOS in the valence band zone is almost made up of Nb-4d and Si-3p orbitals. Furthermore, the significant hybridization peak below the Fermi level in the DOS of these phases manifests that the strong chemical bond is formed between Nb atoms and Si atoms. The Mulliken population can reflect the bond strength to some extent.³⁸ Hence, in this article, the calculated bond lengths of Nb–Si bond with the largest Mulliken population are 2.743 Å, 2.706 Å, 2.689 Å and 2.577 Å for C40–, C11_b–, C54– and C49–NbSi₂, respectively. The calculated bond length of Nb–Si bond in C40–NbSi₂ is in good agreement with experimental data (2.770 Å).³³ This result also reflects that the electronic properties of these predicted phases better than C40–NbSi₂. As mentioned above, we conclude that the electronic properties of various NbSi₂ phases are mainly derived from Nb–Si metallic bond, and Si–Si covalent bond makes a little contribution to their electronic properties.

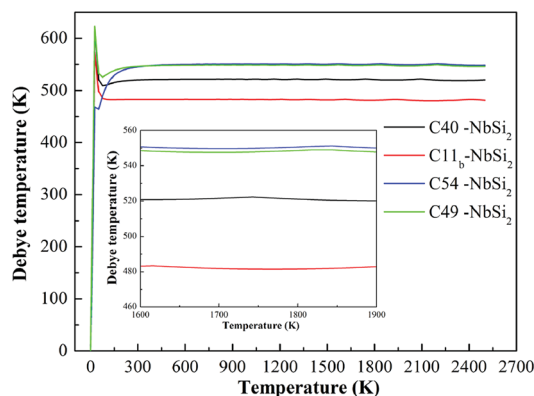


Fig. 4 Calculated the Debye temperature (θ_D) of various NbSi₂.



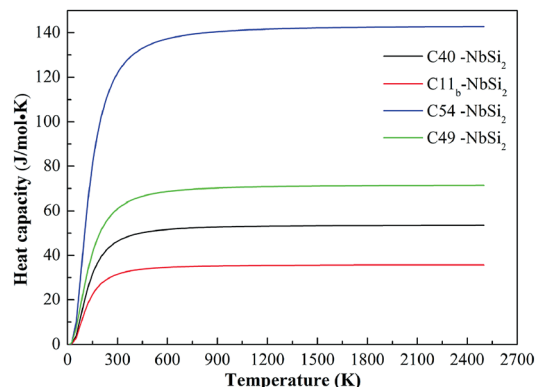


Fig. 5 Calculated the heat capacity (C_V) of various NbSi₂ phases.

Previous studies have shown that there are Nb–Si and Si–Si bonds in C40–NbSi₂ phases.^{24,29,38} To further reveal the nature of these bonds in various NbSi₂, Fig. 3 depicts the valence electron configuration for Si–Si and Nb–Si bonds, and the bond with an asterisk exhibits anti-bonding behavior. From Fig. 3(a), it is clearly found that a electron of Si-3s orbital jump to Si-3p orbital, producing four hybridized orbitals, which forms the normal tetrahedron sp³ hybridization. This sp³ hybridization can be called σ_{sp} bond. The interaction between Si-3p orbitals can form π_{3p} bond. However, the interaction between Nb and Si atoms, *i.e.*, the overlap between sp and d orbital, is very complicated. Here, we simplify their bonding model and just predict that the possible bonding styles. We can see from

Fig. 3(b) that the bonding and antibonding orbitals are formed between the hybridized orbitals of Si atom and the atomic orbital of Nb atom. Besides, there is non-bonding orbital in the atomic orbital on Nb atom. Therefore, we suggest that the structural stability of various NbSi₂ phases is mainly attributed to the Nb–Si bond and Si–Si covalent bond.

Furthermore, both sp and d orbital can carry the electronic current in NbSi₂ with the overlap between them. To our knowledge, the conductivity of each orbital is a proportional to the velocity of electron on the corresponding orbital at the Fermi level. Hence, the sp orbitals composed of s and p orbital are expected to carry larger current than d orbitals characterized by a lower electron velocity. Besides, the effective mass of sp orbital is higher than d orbital in NbSi₂ phase. As mentioned above, it is essential to take the valence electron configuration into account for analyzing the calculated electronic properties of NbSi₂ phases in the present study.

3.1 Thermodynamic properties

To examine the thermodynamic properties of various NbSi₂ phases, the quasiharmonic Debye model at finite temperature is used.³⁹ In this quasiharmonic Debye model, the Gibbs function (G) is obtained by $G(V, P, T) = E(V) + PV + F_{vib}(\theta(V), T)$, where $E(V)$ represents the total energy of NbSi₂ per unit cell and can be calculated by density function theory (DFT). $\theta(V)$ is the Debye temperature and F_{vib} indicates the vibrational Helmholtz free energy.

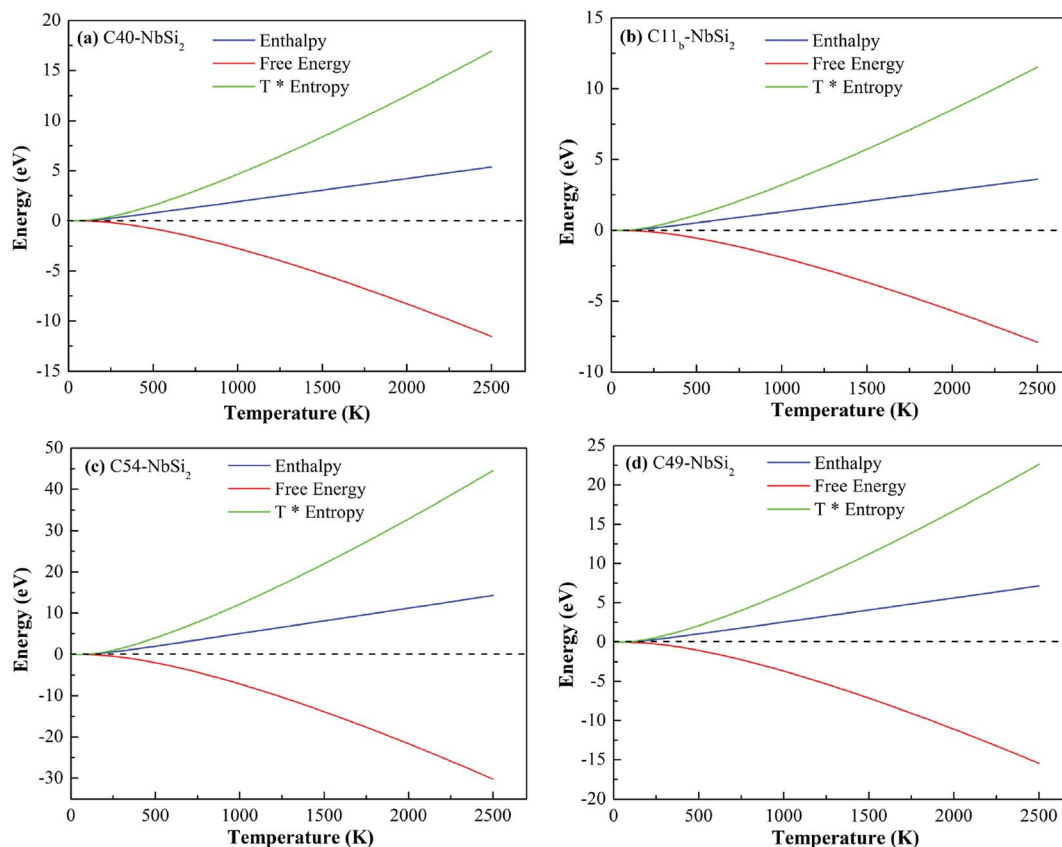


Fig. 6 Calculated enthalpy, free energy and T^* entropy of various NbSi₂ phases.



Fig. 4 displays the calculated the Debye temperature (θ_D) as a function of temperature (0–2500 K). As is well-known, the higher Debye temperature of a material, the stronger interactions between atoms. From Fig. 4, it is observed that the calculated Debye temperature of various NbSi₂ increases sharply at low-temperature. We suggest this phenomenon may be due to the fact that the vibration of atoms in crystals is very sensitive to temperature under low-temperature. Furthermore, the Debye temperature of various NbSi₂ tends to be a constant to some degree under high-temperature. The calculated Debye temperature of C40–NbSi₂ is 522 K, which is slightly lower than previous experimental result.³³ The order of obtained Debye temperature for these structures is: C54 > C49 > C40 > C11_b. The calculated Debye temperature of C54 structure is about 548 K. It is worth mentioned that the Debye temperature of C54 and C49 structures is very similar. We can conclude that C54–NbSi₂ exhibits the best thermal stability among considered structures.

In addition to the Debye temperature, the calculated isochoric heat capacity (C_v) curves of various NbSi₂ are illustrated in Fig. 5. It is observed clearly that the C_v of NbSi₂ with four structures is a proportional to T^3 at temperature close to 0 K and then retains a constant when the temperature increases above 500 K. This trend satisfies the Dulong–Petit limit⁴⁰ ($C_v = 3nR$, where n represents the number of atoms per unit cell). The obtained C_v constants of C40, C11_b, C54 and C49 structures are 53.54 J mol^{−1} K^{−1}, 35.70 J mol^{−1} K^{−1}, 142.73 J mol^{−1} K^{−1} and 71.37 J mol^{−1} K^{−1}, respectively. In particular, the calculated C_v value of C40–NbSi₂ is in good consistent with previous theoretical value (52.21 J mol^{−1} K^{−1} at 300 K).³⁰ We suggest the

discrepancies are attributed to the fact that the different structural characteristics affect the vibration of internal atoms under high-temperature. Importantly, C54–NbSi₂ possesses the largest C_v value in comparison to other phases, implying it exhibits best thermal stability in our considered structures. However, the C_v value of C11_b structure is smallest among all phases. As mentioned above, C54 structure is expected to a new phase for NbSi₂ used as electronic material in future high-temperature application.

Fig. 6 depicts the other thermodynamic functions of various NbSi₂ phases, *i.e.*, enthalpy (H), entropy (S) and free energy. Enthalpy of a material can be calculated by $H = F + TS$. From Fig. 6, we can find that the thermal electronic contributions (TEC) can significantly influence enthalpy and entropy. For four NbSi₂ phases, H exhibits an approximately linear relationship with the temperature (>300 K), and their average slopes of C40, C11_b, C54 and C49 structures are 0.22 kJ mol^{−1} K^{−1}, 0.14 kJ mol^{−1} K^{−1}, 0.60 kJ mol^{−1} K^{−1} and 0.30 kJ mol^{−1} K^{−1}. The result also confirms the analysis of heat capacity.

To deeper understand the mechanism of dynamic stability, we carefully calculated the full and partial phonon density of state (DOS) of various NbSi₂ phases, the results are shown in Fig. 7. It is observed that various NbSi₂ are mechanically stable due to there is no negative frequency. With the increase of temperature, the vibration of atoms intensifies. Therefore, it can be seen clearly from Fig. 7 that the low frequency zone (below about 7.0 THz^{−1}) mainly composed of the vibration of Nb atom, indicating that the structural stability of various NbSi₂ under low-temperature are determined by the Nb atoms.

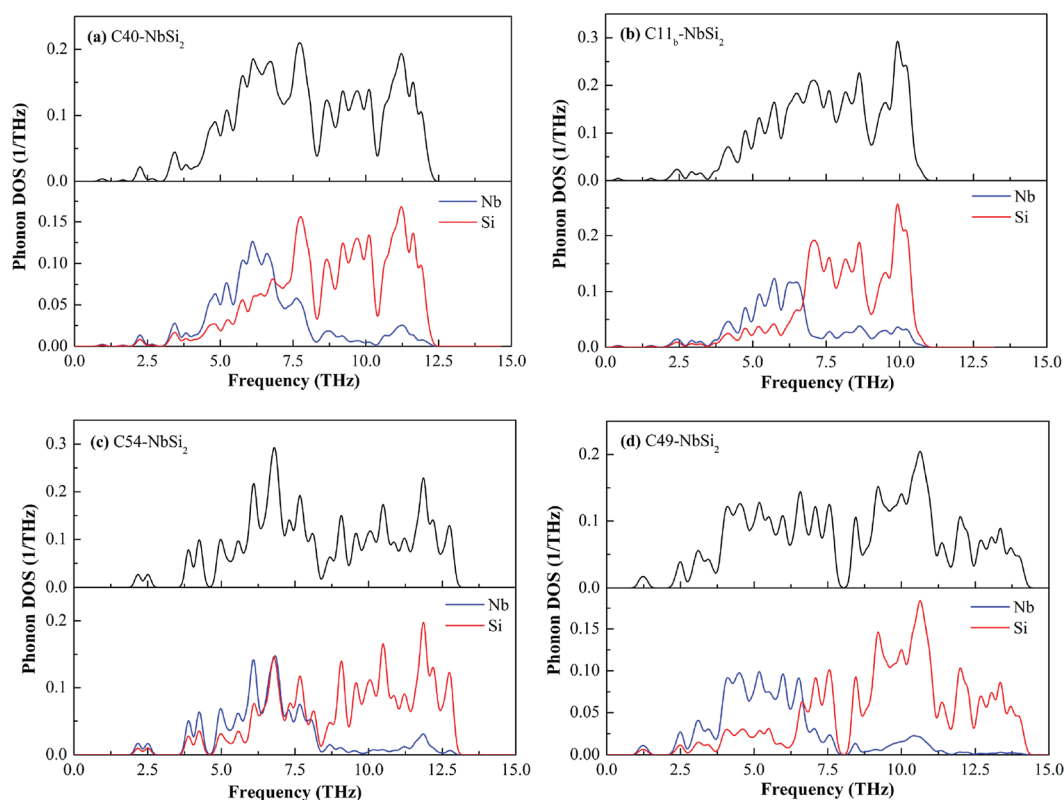


Fig. 7 Calculated full and partial phonon DOS of various NbSi₂ phases.



However, the high frequency zone (about 7.0–15.0 THz⁻¹) is almost all made up of the vibration of Si atoms. Hence, the structural stability of various NbSi₂ under high-temperature is dominated by the Nb–Si bond and Si–Si covalent bond. It is worth pointing out that the vibration of Nb atoms is almost identical to that of Si atoms with the range from 6.30–8.50 THz⁻¹ in Phonon DOS of C54–NbSi₂ (see Fig. 7(c)), which indicates the structural stability in this range are caused of Nb–Si bond.

4. Conclusion

In summary, on the basis of density function theory (DFT), the electronic and thermodynamic properties of various NbSi₂ are systematically investigated. Firstly, the analyses of band structure and density of states show that NbSi₂ with four structures exhibits metallic behavior, which is mainly derived from the Nb–Si metallic bond and some Si–Si covalent bonding in these phases. The three new phases exhibit better electronic properties in comparison to C40–NbSi₂. Based on the molecular hybridization theory, we also explore the mechanism of Si–Si and Nb–Si bonds using the valence electron configuration. Furthermore, it is found that the sp orbital composed of s and p orbitals are expected to carry more current than the d orbital. Finally, we calculate the thermodynamic properties of these phases, including Debye temperature, heat capacity, enthalpy and entropy. The calculated Debye temperature and heat capacity of C54–NbSi₂ are 547.8 K and 142.7 J mol⁻¹ K⁻¹ under high-temperature, respectively, which indicate C54–NbSi₂ exhibits better thermal stability than the other three phases. The thermal stability of C11b–NbSi₂ is the worst in the considered phases.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- 1 F. Nava, E. Mazzega and M. Michelini, *J. Appl. Phys.*, 1989, **65**, 1584–1590.
- 2 Y. Pan, Y. Lin, Q. Xue, C. Ren and H. Wang, *Mater. Des.*, 2016, **89**, 676–683.
- 3 R. Hao, X. Y. Zhang, J. Q. Qin, S. H. Zhang, J. L. Ning, N. Sun, M. Z. Ma and R. P. Liu, *RSC Adv.*, 2015, **5**, 36779–36786.
- 4 Y. Pan, Y. Lin, H. Wang and C. Zhang, *Mater. Des.*, 2015, **86**, 259–265.
- 5 Y. Pan, W. M. Guan and Y. Q. Li, *Phys. Chem. Chem. Phys.*, 2018, **20**, 15863–15870.
- 6 A. Anani and R. A. Huggins, *J. Power Sources*, 1992, **38**, 351–362.
- 7 O. P. Balkashin, A. G. M. Jansen, U. Gottlieb, O. Laborde and R. Madar, *Solid State Commun.*, 1996, **100**, 293–296.
- 8 Y. Nakamura, S. Amari, N. Naruse, Y. Mera, K. Maeda and M. Lchikawa, *Cryst. Growth Des.*, 2008, **8**, 3019–3023.
- 9 F. Nava, K. N. Tu, O. Thomas, J. P. Senateur, R. Madar, A. Borghesi, G. Guizzetti, U. Gottlieb, O. Laborde and O. Bisi, *Mater. Sci. Rep.*, 1993, **4–5**, 141–200.
- 10 F. M. Courtel, D. Duguay, Y. Abu-Lebdeh and I. J. Davidson, *J. Power Sources*, 2012, **202**, 269–275.
- 11 Y. Pan, *Int. J. Hydrogen Energy*, 2018, **43**, 3087–3091.
- 12 M. S. Shin, T. W. Lee, J. B. Park, S. H. Lim and S. M. Lee, *J. Power Sources*, 2017, **344**, 152–159.
- 13 Y. Pan and W. M. Guan, *Phys. Chem. Chem. Phys.*, 2017, **19**, 19427–19433.
- 14 K. Hagihara, H. Araki, T. Ikenish and T. Nakano, *Acta Mater.*, 2016, **107**, 196–212.
- 15 K. Tanaka, K. Nawata, H. Inui, M. Yanmaguchi and M. Koiwa, *Intermetallics*, 2001, **9**, 603–607.
- 16 Y. Pan, S. Wang, X. Zhang and L. Jia, *Ceram. Int.*, 2018, **44**, 1744–1750.
- 17 X. Zhang, Z. Wang and Y. Qiao, *Acta Mater.*, 2011, **59**, 5584–5592.
- 18 L. G. Briquet and P. Philipp, *J. Alloys Compd.*, 2013, **553**, 93–98.
- 19 L. F. Mattheiss, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1992, **45**, 3252.
- 20 X. Chen, J. C. Guan, G. Y. Sha, Z. M. Gao, C. T. Williams and C. H. Liang, *RSC Adv.*, 2014, **4**, 653–659.
- 21 B. Wan, F. Xiao, Y. Zhang, Y. Zhao, L. Wu, J. Zhang and H. Gou, *J. Alloys Compd.*, 2016, **681**, 412–420.
- 22 Y. Imai, M. Mukaida and T. Tsunoda, *Intermetallics*, 2000, **8**, 381–390.
- 23 J. A. Kittl, D. A. Prinslow, P. P. Apte and M. F. Pas, *Appl. Phys. Lett.*, 1995, **67**, 2308–2310.
- 24 Y. Pan, P. Mao, H. Jiang, Y. Wan and W. Guan, *Ceram. Int.*, 2017, **43**, 5274–5282.
- 25 Y. Umakoshi, T. Nakano, E. Yanagisawa, T. Takezoe and A. Negishi, *Mater. Sci. Eng., A*, 1997, **239**, 102–108.
- 26 T. Nakano and K. Hagihara, *Scr. Mater.*, 2013, **68**, 313–316.
- 27 Y. J. Choi, J. K. Yoon, G. H. Kim, W. Y. Yoon and J. M. Doh, *Corros. Sci.*, 2017, **129**, 102–114.
- 28 Y. Pan, *Adv. Eng. Mater.*, 2017, **19**, 1700099.
- 29 I. Papadimitriou, C. Utton, A. Scott and P. Tsakirooulos, *Intermetallics*, 2014, **54**, 125–132.
- 30 N. Xu, Y. Xu and J. Ma, *Mater. Sci. Semicond. Process.*, 2015, **30**, 636–644.
- 31 U. Gottlieb, O. Laborde, O. Thomas, A. Rouault, J. P. Senateur and R. Madar, *Appl. Surf. Sci.*, 1991, **53**, 247–253.
- 32 Y. Pan, J. Zhang, C. Jin and X. Chen, *Mater. Des.*, 2016, **108**, 13–18.
- 33 J. C. Lasjaunias, O. Laborde, U. Gottlieb, R. Madar and O. Thomas, *J. Low Temp. Phys.*, 1993, **92**, 335–351.



- 34 V. N. Antonov, B. Yu. Yavorsky and A. P. Shpak, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **53**, 15631–15637.
- 35 M. D. Segall, P. J. D. Lindan, M. J. Probert, C. J. Pickard, P. J. Hasnip, S. J. Clark and M. C. Payne, *J. Phys.: Condens. Matter*, 2002, **14**, 2717–2744.
- 36 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 37 Y. Pan and B. Zhou, *Ceram. Int.*, 2017, **43**, 8763–8768.
- 38 Y. Pan, *Mater. Res. Bull.*, 2017, **93**, 56–62.
- 39 Y. Pan, P. Wang and C. Zhang, *Ceram. Int.*, 2018, **44**, 12357–12362.
- 40 M. A. Blanco, E. Francisco and V. Luana, *Comput. Phys. Commun.*, 2004, **158**, 57–72.

