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Comparative studies on the interaction of anticancer drug irinotecan with dsDNA and ssDNA†

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A systematic comparative study on the binding of anticancer drug irinotecan (Iri) with dsDNA and ssDNA was investigated in phosphate buffer solutions using voltammetric and spectroscopic methods. The voltammetric results show that the Iri molecule, acting as an intercalator, is inserted into the base stacking domain of the DNA double helix and the strength of interaction is independent of the ionic strength. The hyperchromic effect observed in the UV-visible spectra of Iri in the presence of dsDNA provided the evidence for the intercalation of the drug chromophore with dsDNA base. The interaction mode of Iri molecules with ssDNA is electrostatic attraction *via* negative phosphate on the exterior of the ssDNA with Iri. The binding constants, stoichiometric coefficients and thermodynamic parameters of Iri–dsDNA and Iri–ssDNA complexes were evaluated. The magnitude of changes in ΔG° , ΔH° and ΔS° indicated that the binding process of Iri with ssDNA was more affected than that with dsDNA. The decrease of the peak current of Iri was proportional to DNA concentration, which was applied for determination of dsDNA and ssDNA concentration. The achieved limits of detection of dsDNA and ssDNA were 5.49×10^{-7} and 1.87×10^{-7} M, respectively.

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1. Introduction

Transcription and replication are vital to cell survival and proliferation as well as for smooth functioning of all body processes. DNA starts transcribing or replicating only when it receives a signal, which is often in the form of a regulatory protein binding to a particular region of the DNA.¹ Thus, if the binding specificity and strength of this regulatory protein can be mimicked by a small molecule, then DNA function can be artificially modulated, inhibited or activated by binding this molecule instead of the protein. Thus, synthetic/natural small molecules can act as a drug when activation or inhibition of DNA function is required to cure or control a disease. In recent years increased attention has been focused on the ways in which anticancer drugs interact with biological systems, with the goal of understanding the toxic as well as chemotherapeutic effects of those small molecules.² In this context there are primarily three different ways by which anticancer drugs interact with DNA: (i) through control of transcription factors and

polymerases. Here, the drugs interact with proteins that bind to DNA; (ii) through RNA binding to the DNA double helix to form nucleic triple helix structures or RNA hybridization to exposed DNA single strand forming DNA–RNA hybrids that may interfere with transcriptional activity; and (iii) through binding of small aromatic ligand molecules to DNA double helical structures. The binding of small molecules to DNA involves electrostatic interaction, intercalation between base pairs and minor and major DNA grooves binding interaction. The interaction of some anticancer drugs with DNA has been studied by a variety of techniques including electrophoresis,^{3,4} X-ray crystallography,⁵ structural modeling,⁶ luminescence,⁷ NMR,⁸ resonance Rayleigh light scattering,⁹ fluorescence,^{10–13} dynamic viscosity measurements,¹⁴ high-performance liquid chromatography,¹⁵ and UV-visible absorption spectroscopy.^{16–23} In recent years, there has been a growing interest in electrochemical investigation of interactions between anticancer drugs and other DNA-target molecules and DNA.^{24–36} The use of electrochemical techniques to increase the knowledge of drug molecules and their mechanism of action is one of the most important phases of drug discovery.³⁷ On the other hand, the electrochemical DNA-biosensors form a useful model for simulating nucleic acid interactions with cell membranes, potential environmental carcinogenic compounds and clarifying the mechanism of interaction with drugs and chemotherapeutics.^{38–40} However electrochemistry coupled with spectroscopy techniques could provide a relatively easy way to obtain useful insights into the molecular mechanism of drug–DNA interaction, which could be

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an important step in the development of a new anticancer drug. Among anticancer drugs, Irino has shown activity against colorectal, esophageal, gastric, non-small-cell and small-cell lung cancers, leukemia and lymphomas, as well as central nervous system malignant gliomas.⁴¹ Irino interacts specifically with DNA topoisomerase 1 (Topo 1). The formation of a cleavable drug-Topo1-DNA complex results in lethal double-stranded DNA breakage and cell death.⁴² To the best of our knowledge, no attention has been paid to a comparative study on the interaction of anticancer drug Irino with dsDNA and ssDNA. This investigation offers an opportunity to understand how the kind of DNA affects its binding and affinity of binding to the anticancer drug Irino.

In the present work, comparison of the mode of interaction of Irino with dsDNA and ssDNA was investigated using voltammetric and spectroscopic methods. The formed anticancer Irino-dsDNA and Irino-ssDNA adducts at different temperatures leads to change in the thermodynamic and structural properties of DNA. The thermodynamic characterization of the drug binding might be used, together with structural studies, in development of new strategies to interface with gene expression in living cells. The present investigation offers also an opportunity to use Irino drug as a new electrochemical probe for determination of dsDNA and ssDNA concentrations.

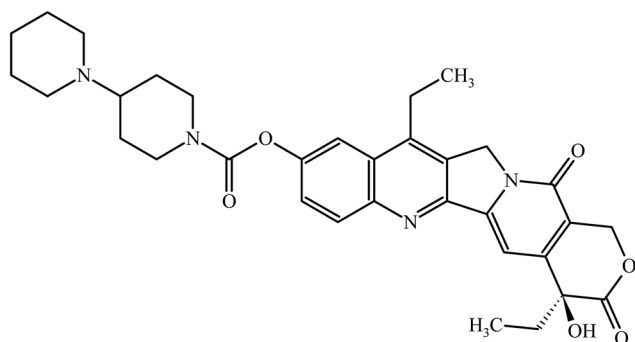
2. Experimental

2.1. Instrumentation

Cyclic voltammetry (CV) and square wave voltammetry (SWV) were performed using an EG&G PAR 384 B (Princeton Applied Research, Oak Ridge, TN, USA) polarographic analyzer controlled by 394 software in conjunction with a PAR Model 303A HMDE (Hanging Mercury Drop Electrode). The three electrode system was completed by an Ag/AgCl (saturated KCl) reference electrode and a Pt wire auxiliary electrode. A PAR Model 305 stirrer was used for the SWV. The ultraviolet and visible absorption spectra measurements were obtained using PerkinElmer (Lambda) spectrophotometer.

2.2. Chemicals and reagents

Calf thymus double stranded deoxyribonucleic acid (dsDNA) and irinotecan (Irino, for chemical structure see Scheme 1) were purchased from Sigma-Aldrich chemicals (St. Louis, Mo, USA)



Scheme 1 Chemical structure of irinotecan.

and were used as received without further purification. A stock standard solution of dsDNA was prepared by dissolving 0.1 g of dsDNA in 100 ml autoclaved deionized water and stored at 4 °C and discarded after no more than four days. The concentration of the stock solution of dsDNA (1.86×10^{-3} M in nucleotide phosphate, NP) was determined by UV absorbance at 260 nm using the molar extinction coefficient (ϵ) at $6600 \text{ M}^{-1} \text{ cm}^{-1}$. Denatured single-stranded DNA was prepared by heating native double-stranded DNA in water bath at a 100 °C for 30 min, followed by promptly cooling in ice bath.⁴³ The formation of ssDNA from dsDNA is checked by gel electrophoresis measurements. Images of gel electrophoresis of ssDNA and dsDNA were provided in Fig. 1. A stock solution of Irino (1×10^{-3} M) was prepared by dissolving an appropriate amount of the drug in deionized water and storing the solution in the dark at 4 °C. The supporting electrolyte was phosphate buffer prepared by mixing stock solution of 0.2 M NaH_2PO_4 and 0.2 M Na_2HPO_4 and adjusting the pH with H_3PO_4 or NaOH to obtain buffers in pH range 2–12. All chemicals were of reagent grade (Merck, Darmstadt, Germany). Double-distilled deionized water was used to prepare the solutions.

2.3. Voltammetric measurements

For voltammetric measurements, the test solution was placed in a polarographic cell of volume 10 ml and deoxygenated by bubbling nitrogen for 15 min. During measurements a stream of nitrogen was passed over the solution. SWV was performed using a pulse amplitude of 50 mVpp; a frequency of 120 Hz; a scan increment of 6 mV; an adsorption time of 60 s and equilibrium time of 15 s. CVs were recorded at a scan rate of 100 mV s^{-1} (unless otherwise stated). The peak potentials were reproducible to at least ± 5 mV in the CVs and ± 2 mV in the SWV. Concentrations of Irino and the total volume solution constant (5 ml) were kept constant while the concentrations of dsDNA and ssDNA were varied. The equilibrium binding constants and thermodynamic parameters for dsDNA-Irino and ssDNA-Irino complexes were determined at different temperatures. All experiments were carried out at 25 ± 0.05 °C.

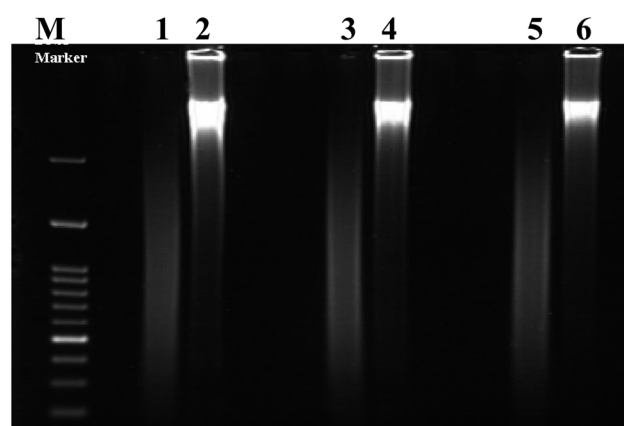


Fig. 1 Images of agarose gel electrophoresis of ssDNA and dsDNA. Lane M: DNA marker; concentrations of ssDNA: 1.0 (Lane1), 0.5 (Lane3) and 0.25 (Lane 5) mg ml^{-1} . Concentrations of dsDNA: 1.0 (Lane 2), 0.5 (Lane4) and 0.25 (Lane 6) mg ml^{-1} .



3. Results and discussion

3.1. Interaction of Irino with dsDNA

Cyclic voltammetric studies of Irino at a HMDE in phosphate buffer solutions showed only one cathodic peak corresponding to direct reduction of the lactone moiety of camptothecin to a lactol ring.³⁰ No anodic peaks were observed at any scan rate, indicating that the reduced form of Irino is electrochemically irreversible. The irreversible nature of the reduction process was confirmed by the shift of the peak potential (E_p) to a more negative value with the increase of the scan rate. The height of the reduction peak reached a maximum and the shape of the curve is better in PBS at pH 5.2. In this context the cathodic peak current of 38.5 μM Irino decreased at pH 5.2 more than at pH 7.2 upon the addition of 3.06 μM dsDNA (Fig. S1†). Therefore pH 5.2 was selected to study the interaction between Irino and dsDNA. Under these conditions the dramatic changes of CV behavior of Irino in the presence of dsDNA should be occurred (Fig. 2). When dsDNA is added to a solution of Irino, the peak current height decreases and peak potential shifts from -900 mV to more negative values -918 mV are observed. The peak potential of CV shifts to more positive or negative values indicating that the action of Irino with dsDNA may be intercalation.^{34,36} Accordingly to these observations it seems that the decrease of peak current of Irino after an addition of excess of dsDNA is caused by the intercalation of Irino to the bulk, slowly diffusing dsDNA, which results in considerable decrease in the apparent diffusion coefficient. This is emphasized from the decrease in the slope of linear $i_p - \nu^{1/2}$ plots ($R \geq 0.994$), where the slope values are 0.075 and 0.035 $\mu\text{A mV}^{-1/2} \text{s}^{1/2}$ in absence and presence of dsDNA respectively. From these values, the diffusion coefficient (D_f) of the free Irino was found to $1.16 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$, whereas $D_b = 2.6 \times 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ for the bound

Irino (in presence of $7.72 \times 10^{-6} \text{ M}$ dsDNA, NP). The changes in current upon addition of dsDNA can be explained in terms of diffusion of an equilibrium mixture free and bound Irino to the electrode and which can be used to quantify the binding of Irino to dsDNA. In this context, current titrations were performed by keeping the concentration of Irino constant while varying the concentration of dsDNA using SWV at pH 5.2. The current titration was described by the following eqn (1).²⁸

$$\log(1/[\text{DNA}]) = \log K + \log(I_{\text{H-G}}/I_{\text{G}} - I_{\text{H-G}}) \quad (1)$$

where K is the apparent binding constant, I_{G} and $I_{\text{H-G}}$ are the peak current of the free guest (G) and the complex (H-G) respectively. Under the assumptions of irreversible, diffusion-controlled electron transfer and a 1 : 1 association complex between the drug and dsDNA (in nucleotide phosphate), then the plot of $\log(1/[\text{DNA}])$ versus $\log(I_{\text{H-G}}/I_{\text{G}} - I_{\text{H-G}})$ becomes linear with the intercept of $\log K$. The binding constant of this complex was evaluated according to eqn (1) and the result is cited in Table 1. The value of K demonstrated that Irino binds to dsDNA with a high association constant.

The change in Gibbs free energy for the Irino–dsDNA adduct was calculated from the well known relationship:

$$\Delta G^\circ = -RT \ln K \quad (2)$$

where the R is universal gas constant ($8.31 \text{ J K}^{-1} \text{ mol}^{-1}$), T the absolute temperature and K the binding constant. The ΔG° values are negative, and increase negatively with increasing temperature, suggesting a thermodynamically favorable binding process.

Ionic strength dependence of the interaction between Irino and dsDNA was also studied. When the ionic strength of the solution increased after a salt was added, charge compensation of the nucleotide phosphates caused by cations would hinder the electrostatic binding effect. Consequently, complex formation through electrostatic binding would decrease, where the binding based on an intercalation mechanism should virtually unaffected.⁴⁴ The ionic strength of reaction medium was varied

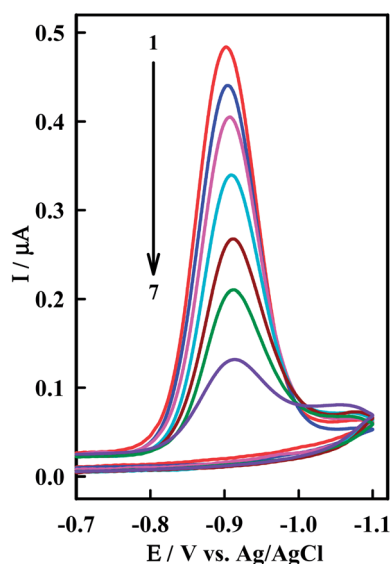


Fig. 2 CV of 38.5 μM Irino at HMDE in phosphate buffer solution (pH = 5.2) in absence (1) and presence of (2) 1.24, (3) 2.06, (4) 3.06, (5) 4.04, (6) 5.0 and (7) 6.23 μM dsDNA. Adsorption potential, -0.2 V ; adsorption time, 60 s; and scan rate, 120 mV s^{-1} .

Table 1 Values of binding constant (K), the standard Gibbs free energy (ΔG°), the enthalpy (ΔH°) and the entropy (ΔS°) for the binding of irinotecan with dsDNA and ssDNA at different temperatures at pH 5.2 using SWV

T (K)	$K \times 10^5$ (M^{-1})	$-\Delta G^\circ$ (kJ mol^{-1})	ΔH° (kJ mol^{-1})	ΔS° ($\text{J K}^{-1} \text{ mol}^{-1}$)
Irino–dsDNA				
278.00 ± 0.05	1.32	27.25	17.93	162.12
288.00 ± 0.05	1.69	28.82		
298.00 ± 0.05	2.10	30.36		
308.00 ± 0.05	2.80	32.12		
Irino–ssDNA				
278.00 ± 0.05	3.30	29.37	27.21	203.59
288.00 ± 0.05	4.89	31.37		
298.00 ± 0.05	6.96	33.33		
308.00 ± 0.05	10.23	35.43		



by changing the concentration of NaCl from 0.05 to 0.3 M. Square wave signals in those solutions with different ionic strength were recorded. No significant change in square wave signals was observed when ionic strength increased. The obtained result indicates that the interaction process of Irino with dsDNA is independent on salt concentration and was mainly intercalation mode. The interaction might be attributed to Irino that intercalates into the DNA double helix between stacked base pairs of dsDNA.

3.2. Spectral profile of the Irino–dsDNA complex

In order to validate the interaction between Irino and dsDNA, the UV-absorption spectra of Irino in the absence and presence of different concentrations of dsDNA were investigated. In the absence of dsDNA Irino has a sharp absorption peak at 255 nm and a broad absorption peak at 368 nm. When Irino and dsDNA are mixed, the spectrum is significantly different from that of Irino or dsDNA (Fig. 3). The absorption band at 255 nm showed increase in the peak intensity (hyperchromic effect) whereas the broad absorption band of Irino at 368 nm decreased slightly with successive additions of dsDNA. The hyperchromic effect may be due to intercalation of the Irino between the dsDNA base pairs. Intercalational phenomena of the Irino with dsDNA could be explained by the presence of two different electronic transitions, $\pi-\pi^*$ and $n-\pi^*$ of Irino spectra. The increase in the concentration of added dsDNA enhanced the absorption peak height *i.e.* hyperchromic effect ($\pi-\pi^*$ region) that could be related with enhanced intercalation of Irino into dsDNA and the complex formed accordingly may become more compact.⁴⁵ To confirm the mode of intercalation between the anticancer Irino and dsDNA, absorption spectra of Irino at a constant

concentration (2×10^{-5} M) in the absence and presence of dsDNA are taken (Fig. 3A). Absorbance maximum of dsDNA is located at 260 nm due to a consequence of the chromophoric groups in purine and pyrimidine moieties responsible for the electronic transitions. When the Irino interacts with dsDNA, it causes hyperchromic effect with increasing concentration of Irino in solution (Fig. 3B). This increase in intensity of absorption maxima (hyperchromic effect) of drug or dsDNA binding molecules is a typical characteristic of dsDNA intercalation.⁴⁶ The increase in the absorbance intensity of dsDNA is attributed also to stacking interaction of drug chromophore with dsDNA bases. This was consistent with the result obtained by the electrochemical method.

3.3. Interaction of Irino with ssDNA

The interaction of the anticancer drug Irino with ssDNA was investigated by means of CV as shown in Fig. 4. When ssDNA is added to a solution of Irino, marked decreases in the peak current heights are observed. The cathodic peak current of Irino decreased by *ca.* 71.1% in presence of 2.08×10^{-6} M ssDNA and the peak potential shifted to more negative potential. Accordingly to these observations it seems that the decrease of peak current of Irino after an addition of excess ssDNA is caused by the electrostatic attraction of Irino to the bulk, slowly dispersion ssDNA, which results in considerable decrease in the apparent diffusion coefficient. This is emphasized from the decrease in the slope of the linear $i_p - \nu^{1/2}$ plots ($R \geq 0.995$), where the slope values are 0.075 and $0.052 \mu\text{A mV}^{-1/2} \text{s}^{1/2}$ in absence and presence of ssDNA respectively. From these values, the diffusion coefficient (D_f) of the free Irino was found to $1.16 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$, whereas $D_b = 5.74 \times 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ for the bound Irino (in presence of 6.67×10^{-7} M ssDNA). The binding constant of the Irino with ssDNA was determined from the current titration as shown above (eqn (1)). The plot of $\log(1/[\text{ssDNA}])$ versus

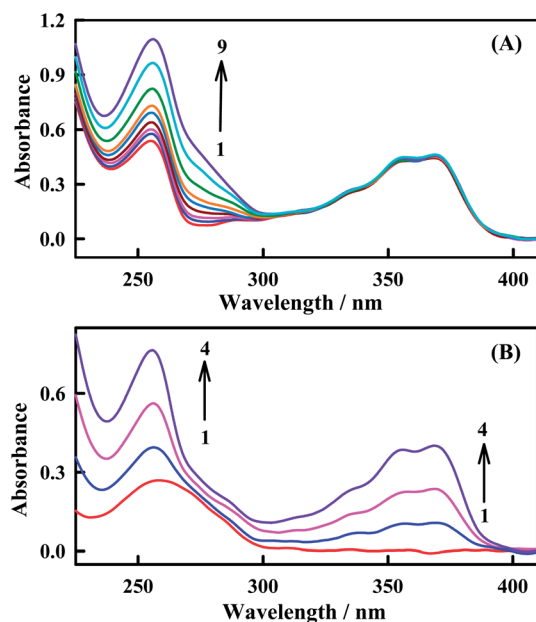


Fig. 3 (A) UV-vis spectra of $20 \mu\text{M}$ Irino (1) with increasing concentrations of dsDNA (2) 2, (3) 4, (4) 7 (5) 10 (6) 13, (7) 20, (8) 30, (9) $40 \mu\text{M}$. (B) UV-vis spectra of $20 \mu\text{M}$ dsDNA (1) with increasing concentrations of Irino (2) 5, (3) 10, (4) $16 \mu\text{M}$ at pH 5.2.

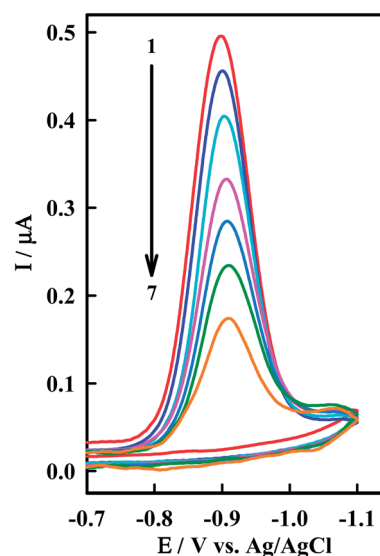


Fig. 4 CV of $38.5 \mu\text{M}$ Irino at HMDE in phosphate buffer solution (pH = 5.2) in absence (1) and presence of (2) 0.33, (3) 0.66, (4) 0.98, (5) 1.14, (6) 1.61 and (7) $2.08 \mu\text{M}$ ssDNA. Other conditions as in Fig. 2.



$\log(I_{H-G}/I_G - I_{H-G})$ is linear with the intercept of $\log K$. The values of K and ΔG° demonstrated that Irino binds to ssDNA with a high association constant (Table 1). The interaction mode of Irino molecules with ssDNA is electrostatic attraction *via* negative phosphate on the exterior of the ssDNA with Irino. In addition, no intercalative interactions are expected if the bind is with single stranded DNA.

Comparison of the interaction of Irino upon the addition of dsDNA and ssDNA at pH 5.2 indicated that the decrease in peak current (ΔI) decreases more sharply with the addition of ssDNA than in the presence of dsDNA as shown in Fig. 5. In this context the apparent diffusion coefficient of the Irino–dsDNA adduct is smaller than that of the Irino–ssDNA complex. Furthermore, the binding constant of Irino–dsDNA adduct is lower than of the Irino–ssDNA adduct (Table 1). As the same the calculated ΔG° values for the binding process of Irino with ssDNA is more negative than of the interaction of Irino with dsDNA. This result indicates a different interaction mechanism of Irino with ssDNA than with dsDNA. The interaction of Irino with dsDNA might be attributed to its intercalation into base stacking domain of DNA double helix. The interaction of Irino with ssDNA indicated that the backbone of ssDNA is negatively charged and can easily attract to the investigated drug *via*

electrostatic attraction. The results offer an opportunity to understand how the kind of DNA affects its binding and affinity of binding to anticancer drug Irino.

3.4. Thermodynamic parameters of Irino–dsDNA and Irino–ssDNA complexes

To have a better understanding of thermodynamics of the reaction between Irino and dsDNA, it is useful to determine the contributions of enthalpy and entropy of the reaction. In this context the integrated form of van't Hoff eqn (3) permits to calculate the values of enthalpy (ΔH°) and entropy (ΔS°) depending on variation of the stability constant with temperature.⁴⁷

$$\ln K = -\Delta H^\circ/RT + \Delta S^\circ/R \quad (3)$$

The van't Hoff plots (Fig. 6) for the Irino–dsDNA adduct shows a linear behavior and allows to derive the thermodynamic parameters (Table 1). The positive value of enthalpy change indicates that the interaction process of Irino with dsDNA is endothermic. Also the entropy associated with the complexation process is positive which indicates an increase in transitional and rotational degrees of freedom of the complexed molecular.²⁸

Accordingly to the thermodynamic data, interpreted as follows, the model of interaction between a drug and biomolecule can be:⁴⁸ (1) when $\Delta H^\circ < 0$ or $\Delta H^\circ \approx 0$ and $\Delta S^\circ > 0$, the electrostatic force dominates the interaction; (2) when $\Delta H^\circ < 0$ and $\Delta S^\circ < 0$, van der Waals interactions or hydrogen bonds dominate the reaction; (3) when $\Delta H^\circ > 0$ and $\Delta S^\circ > 0$, hydrophobic interactions dominate the binding process. By applying this analysis to the binding system of Irino–dsDNA, we determine that non-bonded (hydrophobic) interaction was the most important factor contributing to the observed positive ΔH° and ΔS° and hence to the stability of dsDNA association complex.

The temperature dependence of the binding constant for the interaction of Irino with ssDNA was studied. The interaction of Irino with ssDNA exhibits a positive ΔH° value and also a positive ΔS° value (Table 1). Comparing the thermodynamic parameters for the interaction of anticancer Irino with dsDNA and ssDNA it is obvious that the interaction of Irino with ssDNA displaced high affinity in solution than that with dsDNA. The

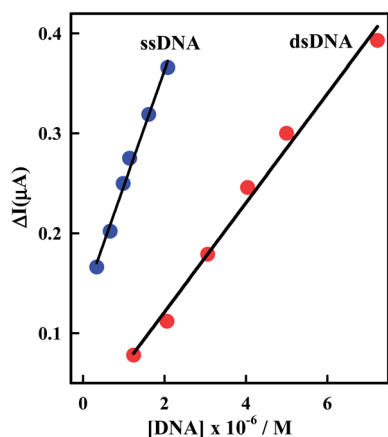


Fig. 5 Plot of ΔI (μA) versus $[\text{DNA}] \times 10^{-6} / \text{M}$ using CV. Other conditions as in Fig. 2.

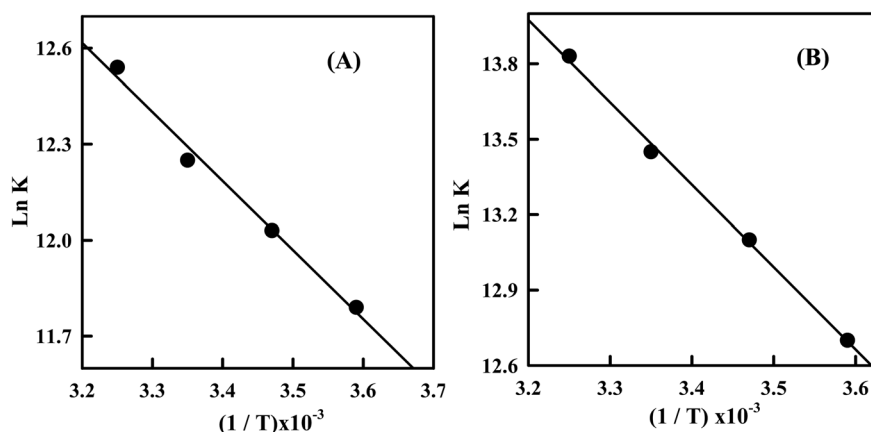


Fig. 6 (A) Plot of $\ln K$ versus $1/T$ for the Irino–dsDNA complex (B) plot of $\ln K$ versus $1/T$ for the Irino–ssDNA complex.



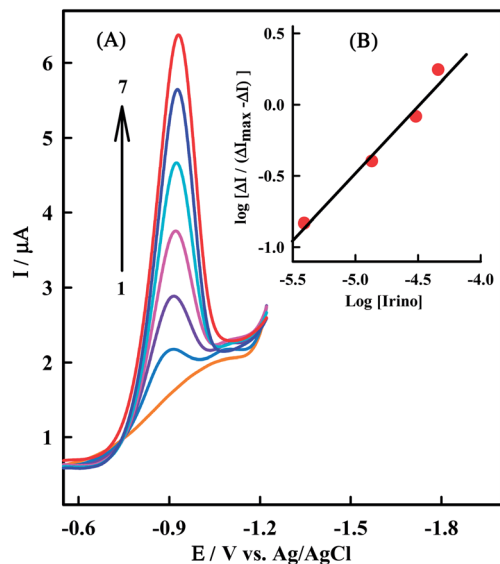


Fig. 7 (A) SWV peak of Irino in presence of 5.19 μM dsDNA at pH 5.2. Adsorption potential, -0.20 V ; adsorption time, 60 s; pulse height, 50 mVpp; scan increment, 6 mV; and frequency, 120 Hz. (1) 0.0, (2) 3.91, (3) 13.5, (4) 30.4, (5) 41.3, (6) 53.8 and (7) 67.6 μM Irino (B) relationship between $\log[\Delta I/(\Delta I_{\text{max}} - \Delta I)]$ and $\log[\text{Irino}]$.

magnitude of changes in ΔG° , ΔH° and ΔS° indicated that the binding process of Irino with ssDNA is more favorable and spontaneous than that with dsDNA. The obtained results indicate a different interaction mechanism of the investigated drug with ssDNA than with dsDNA. The interaction of Irino with dsDNA is attributed to its intercalating into base stacking domain of DNA double helix whereas the interaction of Irino with ssDNA is due to electrostatic attraction.

3.5. Determination of stoichiometric coefficient

As no new electrochemical signals appeared after dsDNA Irino interaction, it would be possible to assume that only one complex is formed $[\text{dsDNA}(\text{Irino})_n]$, according to.⁴⁹



Thus, the binding constant is as follows:

$$\beta_s^n = [\text{dsDNA}(\text{Irino})_n]/[\text{dsDNA}][\text{Irino}]^n \quad (5)$$

where “ n ” is the molar relation of Irino that interacts with a molar quantity of dsDNA bases. Thus, on assessing the change in the current due to the presence of a constant concentration of dsDNA over different concentrations of Irino, the following eqn was used:

$$\log[\Delta I/(\Delta I_{\text{max}} - \Delta I)] = n \log(\beta_s/M^{-1}) + n \log[\text{Irino}] \quad (6)$$

The plot of $\log[\Delta I/(\Delta I_{\text{max}} - \Delta I)]$ vs. $\log[\text{Irino}]$ becomes linear with the slope of “ n ” and the intercept of “ $n \log \beta_s$ ”. Fig. 7A shows the effect of a constant dsDNA concentration on the electrochemical response of a series of Irino concentration solutions and Fig. 7B shows a linear logarithmic relation. The fact that there is one slope in plot represented in Fig. 7B confirms the single complex formation, thus the molar value of “ n ” and binding constant for dsDNA–Irino complex are determined. There is good agreement between K values of Irino–dsDNA obtained from eqn (1) and (6) which are $2.1 \times 10^5\text{ M}^{-1}$ and $2.29 \times 10^5\text{ M}^{-1}$ respectively. The value obtained for “ n ” is 1.34 and smaller than those obtained for bigger molecules such as metallic complexes with reported values up to 25.⁵⁰ This is probably due to the smaller size of Irino compared with metallic complexes.

The effect of a constant ssDNA concentration on the SWV response of a series of concentration of Irino was also studied. In this context the plot of $\log[\Delta I/(\Delta I_{\text{max}} - \Delta I)]$ versus $\log[\text{Irino}]$ is linear and confirms the single complex formation. The value obtained for “ β ” and “ n ” are $5.3 \times 10^5\text{ M}^{-1}$ and 1.25 respectively.

3.6. Analytical aspects of dsDNA–Irino and ssDNA–Irino interactions

The dependence of the peak current of Irino upon successive additions of dsDNA or ssDNA can be employed to determine the concentration of dsDNA and ssDNA. So Irino can be used as a new indicator measuring DNA concentration. Under the

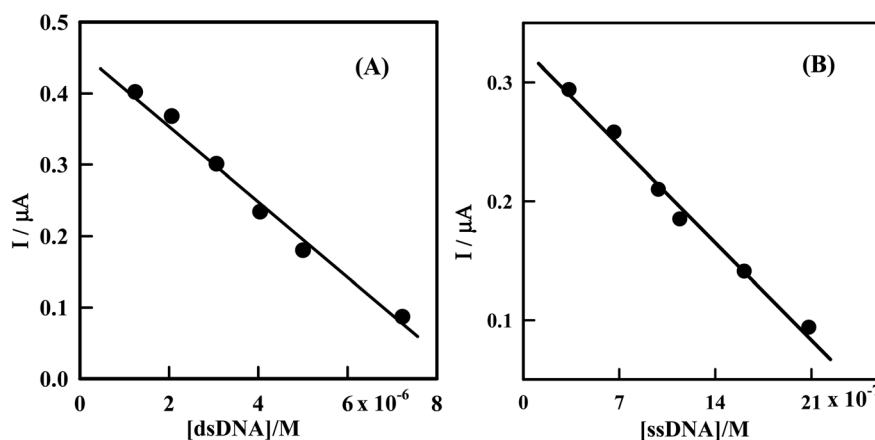


Fig. 8 (A) Plot of I_p (μA) as a function of $[\text{dsDNA}]$ from the voltammograms in Fig. 2. (B) Plot of I_p (μA) as a function of $[\text{ssDNA}]$ from the voltammograms in Fig. 4.



optimum experimental conditions the decreases in peak current of Irino were linearly related to dsDNA or ssDNA concentrations in the range of 1.24×10^{-6} M to 3.33×10^{-7} M when the Irino concentration fixed as at 3.85×10^{-5} M (Fig. 8).

The variation of decrease in peak current *versus* the dsDNA or ssDNA concentration was represented by a straight line followed the equations $I_p (\mu\text{A}) = 0.46 - 5.46 \times 10^4 C (\text{M})$ (dsDNA) and $I_p (\mu\text{A}) = 0.33 - 1.6 \times 10^5 C (\text{M})$ (ssDNA), with regression coefficients of 0.994 and 0.990 respectively. The limits of detection (LOD) and quantitation (LOQ) were computed and the following results obtained: LOD (dsDNA) = 5.49×10^{-7} M, LOQ (dsDNA) = 1.83×10^{-6} M, LOD (ssDNA) = 1.87×10^{-7} M, LOQ (ssDNA) = 6.25×10^{-7} M. The proposed method is simple, sensitive and rapid, and hence can be applied for determination of many kinds of DNA.

4. Conclusion

In this paper a systematic comparative study of the binding of anticancer Irino with dsDNA and ssDNA has been investigated. Based on our investigation, although the anticancer Irino can bind to dsDNA and ssDNA, the nature of binding was found to be different from each them. Irino molecule, acting as intercalator, is inserted into the base stacking domain of the DNA double helix. The thermodynamic parameters ($\Delta H^\circ > 0$ and $\Delta S^\circ > 0$) indicate that the complexation between the anticancer Irino and dsDNA is stabilized by hydrophobic forces. The interaction mode of Irino molecules with ssDNA is electrostatic attraction *via* negative phosphate on the exterior of the ssDNA with Irino. The binding constants and thermodynamic parameters of the binding of Irino with ssDNA indicated a higher affinity of Irino to ssDNA than that with dsDNA. The results also offer an opportunity to use Irino drug as a new electrochemical probe for determination of dsDNA and ssDNA concentrations.

Conflicts of interest

There are no conflicts to declare.

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