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Synthesis, aggregation induced emission and through space conjugation of triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene†

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4-Bromo substituted [2.2]paracyclophane-1,9-diene was synthesized from the corresponding dithia[3.3]paracyclophane in three steps through benzyne Steven rearrangement, oxidation, and a thermal elimination reaction. 4-Triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene was successfully prepared by the Suzuki–Miyaura cross-coupling reaction of 4-bromo substituted [2.2]paracyclophane-1,9-diene and 4,4,5,5-tetramethyl-2-(4-(1,2,2-triphenylvinyl)phenyl)-1,3,2-dioxaborolane using $Pd(OAc)_2$ as a catalyst, S-Phos as a ligand and K_3PO_4 as a base. The structures of bromo substituted [2.2]paracyclophane-1,9-diene and triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene were fully characterized by 1H NMR spectroscopy and X-ray crystallography. 4-Triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene exhibited aggregation-induced emission characteristics when the water fraction was higher than 80% in the THF/water mixtures. 4-Triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene displays much higher fluorescence when the water fraction is 90% compared to that of model compounds due to both through bond and through space conjugation. To the best of our knowledge, we are the first to synthesize triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene with aggregation-induced emission characteristics.

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Introduction

[2.2]Paracyclophanes have received tremendous attention over the past six decades.^{1–5} The structures of [2.2]paracyclophanes which comprise intramolecular face-to-face π -conjugated phenyl rings give rise to unique electronic and optical properties.^{3,6} The introduction of substituents onto the individual phenyl rings of [2.2]paracyclophanes makes them become chiral molecules that were widely used as the building blocks for organic reactions and catalyst applications.^{7–9} Recently, a number of conjugated molecules bonded to the two individual phenyl rings of the [2.2]paracyclophane have been intensively prepared and investigated.^{10–12} These types of substituted [2.2]paracyclophanes exhibited remarkable intramolecular through space conjugation and charge transfer features which lead to potential applications in optoelectronics.

[2.2]Paracyclophane-1,9-dienes were first synthesized by Dewhurst and Cram in the 1960s.¹³ [2.2]Paracyclophane-1,9-dienes are one of the paracyclophane families in which the two bridges

are replaced by *cis*-vinylenes that increase the angles between the bridgehead carbons and the other central carbons of the phenyl rings. Therefore, the highly twisted, ring strained nature with *cis*-vinylene bridges of [2.2]paracyclophane-1,9-dienes makes them enable to be ring-opened by carbene initiator to form phenylenevinylenes.^{14–17} The alkoxy or alkyl substituted paracyclophane-1,9-dienes and their soluble phenylenevinylene homopolymers and block copolymers prepared by direct ring-opening metathesis polymerization (ROMP) have been reported.^{18–22} In addition, a variety of phanediene, phanetriene or phanetetraenes have been synthesized and investigated.^{23–25} One of the advantages for soluble phenylenevinylene polymers by ROMP of alkoxy or alkyl substituted paracyclophane-1,9-dienes is that the precise control of the molecular weights, polydispersity index (PDI) as well as the functional end groups as this give low defect structures of the polymers.

In general, most of the conjugated small molecules showed significant fluorescence quenching in the aggregate state due to strong intermolecular interaction. The aggregation-caused quenching (ACQ) effect therefore limits the uses of conjugated molecules in aqueous solution or in solid state. This issue can be solved after the aggregation induced emission (AIE) phenomenon was first discovered by Tang and co-workers.²⁶ After that, the AIE active molecules have been intensively prepared. In particular, tetraphenylethenes (TPE) with propeller-like structures have attracted much attention possibly due to the facile preparation and purification. The TPE and its

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derivatives have been intensively investigated and widely used as sensors for the detection of a variety of analysts^{27–29} owing to high luminescence efficiency and good photostability.

Herein, we reported the synthesis of 4-triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene (**TPE-PCPDE**) by Suzuki–Miyaura cross-coupling reaction of 4-bromo substituted [2.2]paracyclophane-1,9-diene and 4,4,5,5-tetramethyl-2-(4-(1,2,2-triphenylvinyl)phenyl)-1,3,2-dioxaborolane using $\text{Pd}(\text{OAc})_2$ as a catalyst, *S*-Phos as a ligand and K_3PO_4 as a base. The structural characterization, optical properties as well as the aggregation-induced emission characteristics of **TPE-PCPDE** compared to the model compounds such as para-cyclophaneadiene (**PCPDE**), tetraphenylethene (**TPE**) and phenyl substituted tetraphenylethene (**P-TPE**) were investigated. The structure of triphenylvinylphenyl substituted para-cyclophaneadiene **TPE-PCPDE** and the model compounds such as **PCPDE**, **TPE** and **P-TPE** is shown in Fig. 1.

Experimental

Materials and instrumentation

All reagents were from commercial sources and were used without further purification. Solvents used for spectroscopic measurements were spectrograde. Most of the reactions were monitored by thin-layer chromatography carried out on silica gel plates. Preparative separations were performed by column chromatography on silica gel grade 60 (0.040–0.063 mm) from Merck. The starting material, dithiaparacyclophane **1**, was synthesized by the modification of previously established synthetic methods (more details see ESI†). Nuclear magnetic resonance (NMR) spectra were obtained on a Bruker ultrashield spectrometer operating at 600 MHz for ¹H nuclei and 150 MHz for ¹³C nuclei. Chemical shifts are reported in ppm. All coupling constants are reported in hertz (Hz). The abbreviations are used to indicate multiplicity: s = singlet, d = doublet, dd = doublet of doublets, t = triplet and m = multiplet. UV-vis absorption and photoluminescence spectra were recorded on the Jasco (V–670) UV-Vis-NIR spectrophotometer and on the Jasco (FP-8500) fluorescence spectrophotometer, respectively. X-ray crystallographic data for single crystals were determined using a Bruker Kappa APEX II diffractometer equipped with an Oxford Cryosystems 700 series cryostream cooler using graphite monochromated Mo-K α radiation. Electron impact mass spectra (EI-MS) were recorded on a JEOL JMS-700 spectrometer. The melting point of compounds were recorded on a MEL-TEMP 1001D capillary melting point apparatus without further calibration. In order to model the electron affinity and ionization

potential, the HOMO and LUMO of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** were computed using density functional theory. Becke's three parameter exchange and the Lee–Young–Parry correlation functionals (B3LYP) were used with added polarization basis functions (6-31G*).

Syntheses

Synthesis of bis(sulfoxide) compounds 2. To a boiling solution of **1** (4.22 g, 12 mmol) and anthranilic acid (5.75 g, 42 mmol) in anhydrous 1,2-dichloroethane (300 mL), isoamyl nitrite (6.6 mL, 49 mmol) was added dropwise for at least 30 minutes. The resulting solution was stirred under reflux for an additional 30 minutes. After evaporating the solvent, the residue was transferred to a silica gel column, using hexane following by DCM : hexane (3 : 7) as eluents. The main fraction was collected to give bis(sulfide) compounds as the yellow oils in a yield of 74%. Bis(sulfide) compounds were very difficult to characterize by NMR spectroscopy due to the large numbers of overlapping signals. HRMS (EI, M^+): calculated for $[\text{C}_{28}\text{H}_{23}\text{BrS}_2]^+$: *m/z* 502.0425, found *m/z* 502.0425. Hydrogen peroxide (1.07 mL, 35 wt%) was added dropwise to the solution of bis(sulfide) compounds (3.1 g, 6.16 mmol) in toluene (145 mL) and acetic acid (48 mL) over a period of 30 minutes at 0 °C. The solution was allowed to warm up to room temperature and kept stirring for another 12 hours. The resulting solution was extracted with DCM and brine, dried with anhydrous MgSO_4 and concentrated to give compound **2** as yellow oils in a yield of 93% without further purification. Again, the ¹H NMR spectrum of compound **2** exhibited a complex range of signals and characterization was not possible. HRMS (EI, M^+): calculated for $[\text{C}_{28}\text{H}_{23}\text{BrO}_2\text{S}_2]^+$: *m/z* 534.0323, found *m/z* 534.0320 (see Fig. S21 in ESI†).

Synthesis of 4-bromo-[2.2]paracyclophane-1,9-diene (Br-PCPDE**).** A solution of **2** (3.08 g, 5.75 mmol) in *p*-xylene (190 mL) was refluxed under an argon steam for 20 hours. After removing the solvent, the residue was chromatographed over silica gel using hexane as an eluent. The compound was further purified by recrystallization to give compound **Br-PCPDE** as white solids with the yield of 16%. ¹H NMR (600 MHz, CDCl_3 , δ): 7.24 (d, J = 10.2 Hz, 1H, ArCH=CH-), 7.19 (d, J = 10.2 Hz, 1H, ArCH=CH-), 7.10 (d, J = 10.2 Hz, 1H, ArCH=CH-), 7.07 (d, J = 7.8 Hz, 1H, ArH), 6.93 (d, J = 10.2 Hz, 1H, ArCH=CH-), 6.66 (s, 1H, ArH), 6.64 (d, J = 7.8 Hz, 1H, ArH), 6.56 (d, J = 7.8 Hz, 1H, ArH), 6.47 (d, J = 7.8 Hz, 2H, ArH), 6.41 (d, J = 7.8 Hz, 1H, ArH). ¹³C NMR (150 MHz, CDCl_3 , δ): δ 139.99, 138.08, 138.05, 137.94, 137.59, 136.48, 136.42, 135.75, 135.69, 133.09, 131.67, 130.75, 130.47, 130.40, 126.71, 124.44. HRMS (EI, M^+): calculated for $[\text{C}_{16}\text{H}_{11}\text{Br}]^+$: *m/z* 282.0044, found *m/z* 282.0044 (see Fig. S22 in ESI†). Mp: 135 °C.

Synthesis of 4-(4-(1,2,2-triphenylvinyl)phenyl)-[2.2]paracyclophane-1,9-diene (TPE-PCPDE**).** A mixture of 4-bromo-[2.2]paracyclophane-1,9-diene **Br-PCPDE** (56.6 mg, 0.2 mmol), 4,4,5,5-tetramethyl-2-(4-(1,2,2-triphenylvinyl)phenyl)-1,3,2-dioxaborolane (110 mg, 0.24 mmol), $\text{Pd}(\text{OAc})_2$ (4.5 mg, 0.02 mmol) and *S*-Phos (16.4 mg, 0.04 mmol) were charged into a Schlenk tube under an argon atmosphere. Deoxygenated THF

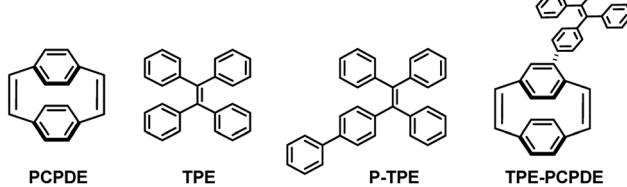


Fig. 1 The structure of PCPDE, TPE, P-PCPDE and TPE-PCPDE.

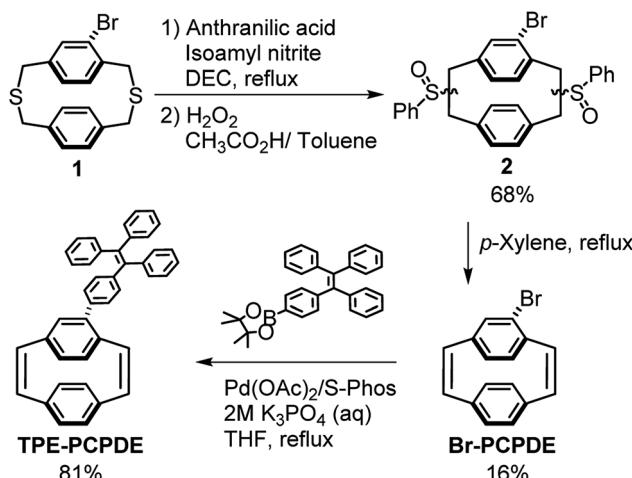


(2.0 mL) and 2.0 M aqueous K_3PO_4 (0.67 mL) were added to the reaction mixture and then the reaction mixture was heated to reflux for 72 hours. After filtrating the insoluble solid and removing most of the solvent, the residue was extracted with DCM and brine several times, dried with anhydrous $MgSO_4$ and removed the solvent. The crude compound was purified by column chromatography (DCM : hexane, 1 : 9 (v/v)) to give compound **TPE-PCPDE** as white solids in a yield of 81%. 1H NMR (600 MHz, $CDCl_3$, δ): 7.25 (d, J = 10.5 Hz, 1H, $ArCH=CH-$), 7.22 (d (AA'BB'), J = 8.4 Hz, 2H, ArH), 7.20 (d, J = 10.5 Hz, 1H, $ArCH=CH-$), 7.14 (d, J = 10.5 Hz, 1H, $ArCH=CH-$), 7.18–7.04 (m, 16H, ArH), 6.90 (d, J = 10.5 Hz, 1H, $ArCH=CH-$), 6.72 (s, 1H, ArH), 6.65 (d, J = 8.1 Hz, 1H, ArH), 6.56 (d, J = 8.1 Hz, 1H, ArH), 6.51 (d, J = 7.8 Hz, 1H, ArH), 6.47 (d (A2), J = 7.8 Hz, 2H, ArH), 6.36 (d, J = 7.8 Hz, 1H, ArH). ^{13}C NMR (150 MHz, $CDCl_3$, δ): 143.92, 143.81, 143.62, 142.63, 141.39, 141.26, 140.88, 138.62, 138.48, 138.22, 137.81, 137.29, 137.00, 136.07, 134.57, 133.31, 131.60, 131.55, 131.51, 131.50, 131.12, 130.74, 130.54, 130.51, 129.33, 128.66, 127.84, 127.79, 127.76, 127.32, 126.63, 126.58, 126.56, 138.24. HRMS (EI, M^+): calculated for $[C_{42}H_{30}]^+$: m/z 534.2348, found m/z 534.2345 (see Fig. S23 in ESI \dagger). Mp: 143 °C.

Results and discussion

The routes to synthesize 4-triphenylvinylphenyl-[2.2]paracyclophane-1,9-diene are shown in Scheme 1. The starting material, dithiaparacyclophane **1** (see ESI \dagger), was prepared by the cyclization of an equimolar amount of 2-bromo-1,4-bis(bromomethyl)benzene **S2** and 1,4-phenylenedimethanethiol **S4** under a high dilution in the presence of aqueous potassium hydroxide for at least 3 days to avoid the formation of oligomers and polymers.

Compound **2** was prepared by the benzyne induced Stevens rearrangement using anthranilic acid and isoamyl nitrite following by oxidation using H_2O_2 (35 wt%). 30,31 Compound **2** was then obtained as semi-solids in an overall yield of 68%. Compound **2** was then dissolved in *p*-xylene at reflux under an argon atmosphere for 20 hours. The 4-bromo-[2.2]paracyclophane-1,9-diene (**Br-PCPDE**)



Scheme 1 Synthetic route to triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene (**TPE-PCPDE**).

was obtained in a yield of 16% after purification by column chromatography and recrystallization. The 4-triphenylvinylphenyl substituted [2.2]paracyclophane-1,9-diene (**TPE-PCPDE**) was synthesized by the reaction of **Br-PCPDE** and 4,4,5,5-tetramethyl-2-(4-(1,2,2-triphenylvinyl)phenyl)-1,3,2-dioxaborolane in the presence of $Pd(OAc)_2$ and 2-dicyclohexylphosphino-2',6'-dimethoxybiphenyl (*S*-Phos) in THF and aqueous K_3PO_4 at reflux for 72 hours. After purified by column chromatography, the **TPE-PCPDE** was obtained in a yield of 81%.

The 1H NMR spectrum of **Br-PCPDE** in $CDCl_3$ is shown in Fig. 2a. The four doublets (7.24, 7.19, 7.10 and 6.93 ppm, each J = 10.2 Hz, 1H) were assigned to the *cis*-vinylic hydrogens. Two doublets at 7.07 and 6.41 ppm with J = 7.8 Hz were associated with the hydrogens of non-substituted phenyl ring. In addition, the doublet at 6.47 ppm (J = 7.8 Hz) integrating to two hydrogens corresponds to the hydrogens to the non-substituted phenyl ring. A singlet at 6.66 ppm corresponds to the aromatic hydrogen at the *ortho*-position to the bromine atom. The two doublets (6.64 and 6.56 ppm, each J = 7.8 Hz, 1H) were assigned to the hydrogens at the *para*- and *meta*-positions to the bromine atom, respectively. The 1H NMR spectrum of **TPE-PCPDE** in $CDCl_3$ is shown in Fig. 2b. The two doublets (7.25 and 7.20 ppm, each J = 10.5 Hz, 1H) were assigned to the *cis*-vinylic hydrogens. In addition, a doublet at 6.90 ppm (J = 10.5 Hz, 1H)

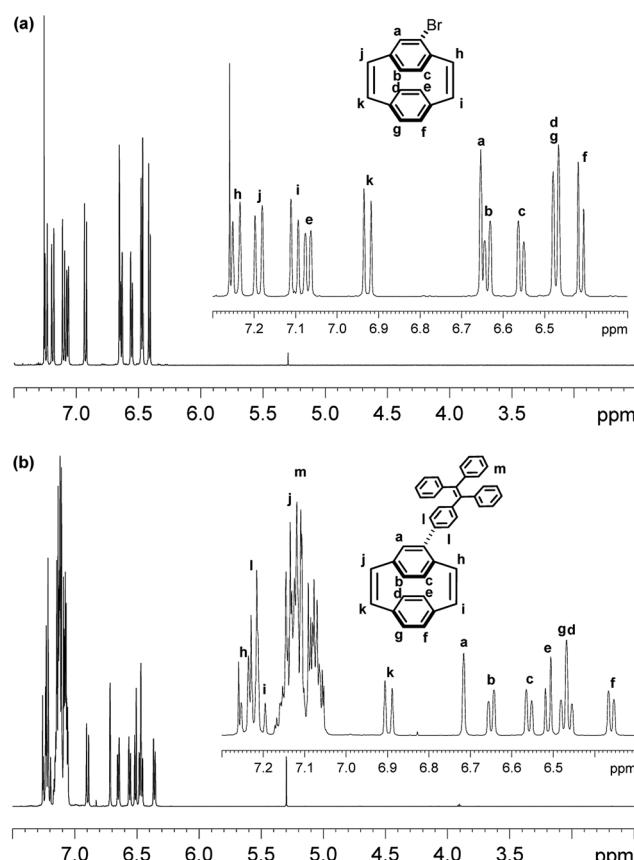


Fig. 2 1H NMR spectra of (a) 4-bromo-[2.2]paracyclophane-1,9-diene **Br-PCPDE** and (b) 4-triphenylvinylphenyl-[2.2]paracyclophane-1,9-diene **TPE-PCPDE** in $CDCl_3$. Signals associated with alkenic and aromatic hydrogens are shown in the insets.



was assigned to the *cis*-vinylic hydrogen. It should be noted that a doublet at around 7.14 ppm ($J = 10.5$ Hz, 1H) can be assigned to the remaining *cis*-vinylic hydrogen which is overlapping with the hydrogens of **TPE** moiety. The doublet at 7.22 ppm (AA'BB' system, $J = 8.4$ Hz, 2H) corresponds to the hydrogens to the *ortho* position bonded to the carbon of the **PCPDE**. The multiple between 7.18 and 7.04 ppm represent for the rest of the hydrogens of **TPE** moiety. A singlet peak at 6.72 ppm and two doublet peaks (6.65 and 6.56 ppm, each $J = 8.1$ Hz, 1H) were assigned to the aromatic hydrogens at the *ortho*-, *para*- and *meta*-positions to the bromine atom, respectively. The two doublets (6.51 and 6.36 ppm, each $J = 7.8$ Hz, 1H) were assigned to hydrogens of the non-substituted phenyl ring. In addition, the two doublet peaks at 6.47 ppm (an A2 system) with $J = 7.8$ Hz are two hydrogens belong to the non-substituted phenyl ring. 2D ^1H - ^1H COSY can confirm the peak assignment and can be found in ESI.†

The single crystal of **Br-PCPDE** and **TPE-PCPDE** can be obtained by slow evaporation from a dilute hexane solution at room temperature and the solid-state structure of **Br-PCPDE** and **TPE-PCPDE** was determined by X-ray crystallography. For solid-state structure of **Br-PCPDE** (Fig. 3a), the average vinylene bond length is 1.34 Å which is in agreement with a standard *cis*-

vinylene bond (1.32 Å). The average intramolecular distance between the two carbons of the phenyl rings bonded to carbons of vinylene is 2.78 Å which is the same value with that of the unsubstituted **PCPDE** reported previously. The average intramolecular separation between the central carbon atoms of the phenyl rings is 3.11 Å. The carbons of the phenyl rings bonded to the carbons of vinylene are significantly out of plane of the four central carbons and the average deviation is of around 13.1°. It should be noted that the largest value (13.43°) of out of plane for the carbon which is in the *ortho*-position relative to the carbon bonded to bromine atom was observed. The average C(aryl)-C(methine) bond length is 1.50 Å. The solid-state structure of **TPE-PCPDE** determined by X-ray crystallography is shown in Fig. 3b. The average bond length of vinylene is 1.32 Å which is the same value compared to a standard *cis*-vinylene bond (1.32 Å). The average intramolecular separation between the central carbon atoms of the phenyl rings is 3.12 Å. The average intramolecular distance between the carbon atoms of the phenyl rings linked to carbons of vinylene is significantly shortened to 2.79 Å. The dihedral angle between the phenyl rings of **PCPDE** and **TPE** moieties is around 39°. The carbons of the phenyl rings connected to the vinylene in **PCPDE** moieties are significantly out of plane of the four central carbons and the average deviation is of around 13.4°. A large distortion in the phenyl rings can be seen which indicates the highly ring strained nature of this molecule. The average C(aryl)-C(methine) bond length is 1.50 Å. Both of the crystal packing structures of compounds **Br-PCPDE** and **TPE-PCPDE** (space group: $P2_1/c$, $Z = 4$) showed four molecules in one unit cells and two pairs of them are racemic mixtures due to the chirality center of substituted phenyl rings of paracyclophane dienes.

The model compounds **PCPDE**, **TPE** and **P-TPE** were prepared by the modification of established procedures reported previously.³²⁻³⁴ The initial concentration of compounds **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** is 2×10^{-6} M in dilute THF solution in order to eliminate the potential interference of intermolecular interactions. The UV-vis absorption peak of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** in dilute THF solution exhibited at 242, 309, 321 and 325 nm, respectively (Fig. 4a). The notably bathochromic shift from **PCPDE**, **TPE**, **P-TPE** to **TPE-PCPDE** can be attributed to the extension of the conjugation. It should be noted that the longest absorption wavelength of **TPE-PCPDE** is red-shifted by 4 nm compared to that of **P-TPE** possibly due to the through space conjugation of **TPE-PCPDE**.

The thin films were prepared by spin-casting a solution of the compounds (3 mg mL⁻¹ in toluene) onto glass slides (2.0 cm \times 2.0 cm) at a rotation speed of 1500 rpm for 20 seconds and then 3000 rpm for 20 seconds. The normalized solid state UV-vis absorption spectra of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** were shown in Fig. 4b. The solid state UV-vis absorption maximum of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** is at 292, 326, 331 and 335 nm, respectively which shows bathochromic shifts compared to that of solution due to the relatively strong intermolecular interaction in solid state. The details of optical properties of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** are summarized in Table 1.

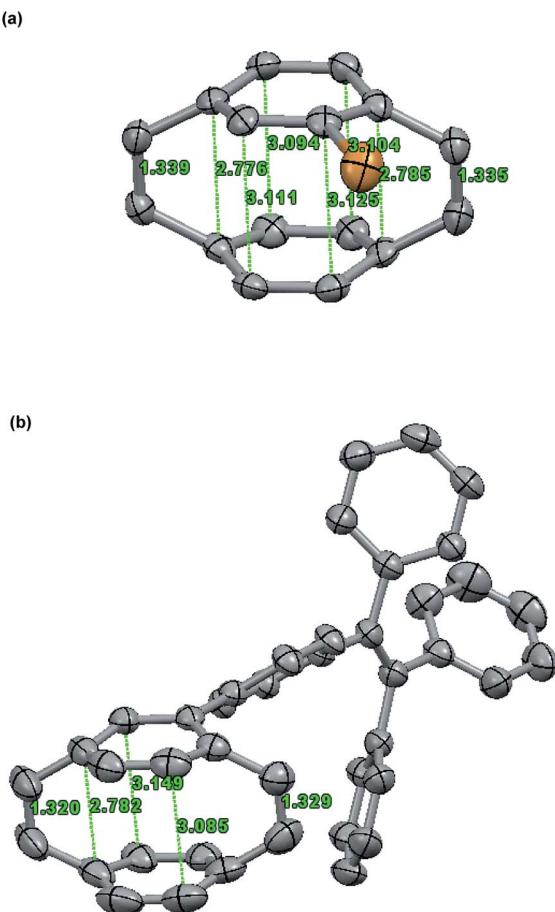


Fig. 3 Solid-state structure of (a) **Br-PCPDE** and (b) **TPE-PCPDE**. Hydrogen atoms are omitted for clarity. Thermal ellipsoids are set at 50% probability.



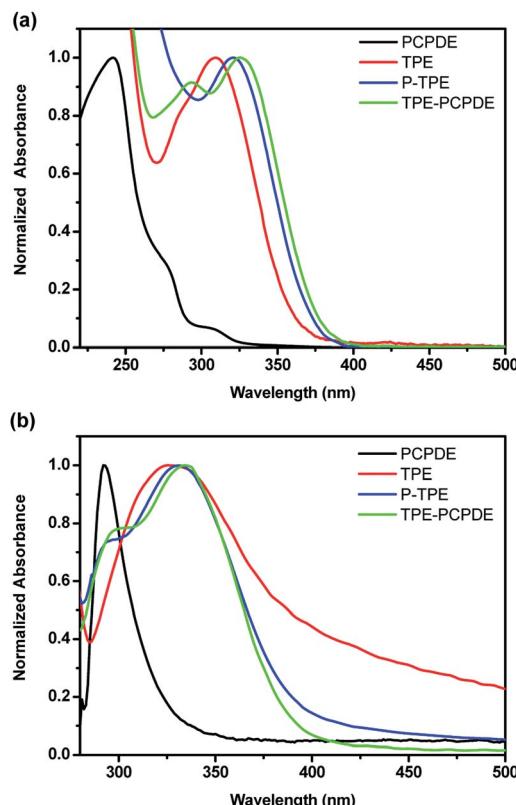


Fig. 4 Normalized (a) solution and (b) solid state UV-vis absorption spectra of PCPDE, TPE, P-TPE and TPE-PCPDE. (The initial concentration for solution is 2×10^{-6} M).

Again, an absorption maximum of **TPE-PCPDE** in solid state exhibited a bathochromic shift compared to the model compounds due to the through bond and through space conjugation. It should be noted that the solid state UV-vis absorption spectrum of **TPE** exhibited a very long tail extending through the whole visible region possibly due to the scattering effect as the quality of films of **TPE** is poor. In addition, the comparative HOMO and LUMO structures and levels of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** using density functional theory calculation (B3LYP functional, 6-31G* basis set) were shown in ESI (Fig. S27 and Table S1†). The HOMO levels and LUMO levels became upper and lower, respectively on going from **PCPDE**, **TPE**, **P-TPE** to **TPE-PCPDE**. Based on the experimental measurements and theoretical calculations, the through bond and through space conjugation make significant

Table 1 Optical properties of PCPDE, TPE, P-TPE and TPE-PCPDE

Compound	λ_{max}^a (nm)	PL_{max}^a (nm)	λ_{max}^b (nm)	PL_{max}^b (nm)
PCPDE	242	— ^c	292	379
TPE	309	375	326	440
P-TPE	321	384	331	475
TPE-PCPDE	325	416	335	477

^a In dilute THF (2×10^{-6} M). ^b In solid state. ^c Intensity of the signals comparable to spectral noise.

contributions to the absorption wavelength and bandgap in **TPE-PCPDE** when compared with **PCPDE**, **TPE** and **P-TPE**.

The emission spectra of **TPE-PCPDE** with different water fractions (f_w) in THF are shown in Fig. 5a. The maximum emission intensity of **TPE-PCPDE** is very low with little change when the water fraction is lower than 80%. However, the maximum emission intensity of **TPE-PCPDE** shows significant increase when the water fraction is higher than 80%. This indicates that the **TPE-PCPDE** exhibited aggregation-induced emission characteristics in THF/water mixtures when the water fraction is higher than 80%. The maximum emission wavelength of **TPE-PCPDE** is at 475 nm in its most aggregated state. The pictures of **TPE-PCPDE** in THF with different water contents under the UV lamp irradiation with a wavelength of 365 nm are shown in Fig. 5b. Clearly, the **TPE-PCPDE** in pure THF solution and lower water contents (<80% of water) showed very weak fluorescence. In addition, the fluorescence can be seen by naked eyes for **TPE-PCPDE** when the water fraction is higher than 80% in the THF/water mixture.

The photoluminescence quantum yields of **TPE**, **P-TPE** to **TPE-PCPDE** with different water fractions are shown in Fig. 6a. The maximum photoluminescence quantum yield of **TPE**, **P-TPE** and **TPE-PCPDE** is 0.049, 0.140 and 0.266, respectively when the water fraction is 90% compared to quinine sulfate (0.546) in 0.1 M H_2SO_4 . **TPE** is a typical AIE molecule. However, the quantum yield of **TPE** in 90% water content reported here was only about 5% compared to 14–18% reported previously. This is due to the low initial concentration used in this experiment (2×10^{-6} M) which is more than 5 times diluted than reported in previous literatures.^{35,36} Clearly, compound **TPE-PCPDE** displays much higher fluorescence when the water

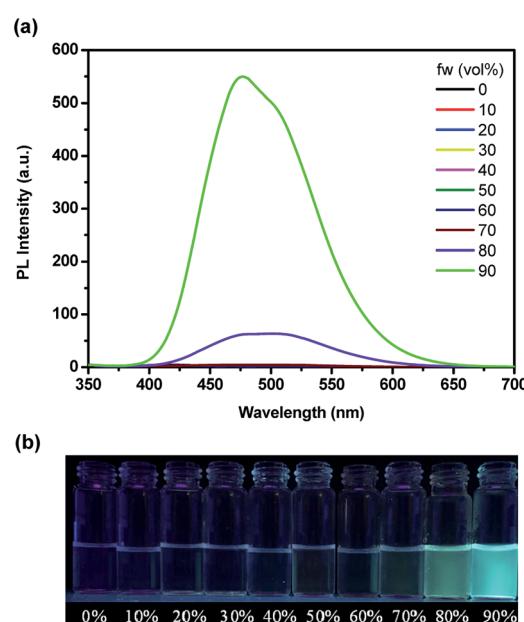


Fig. 5 (a) Photoluminescence spectra of **TPE-PCPDE** in THF/water mixtures with different water fractions (initial concentration: 2×10^{-6} M) and (b) the pictures of **TPE-PCPDE** in THF with different water contents under the UV lamp at a wavelength of 365 nm.

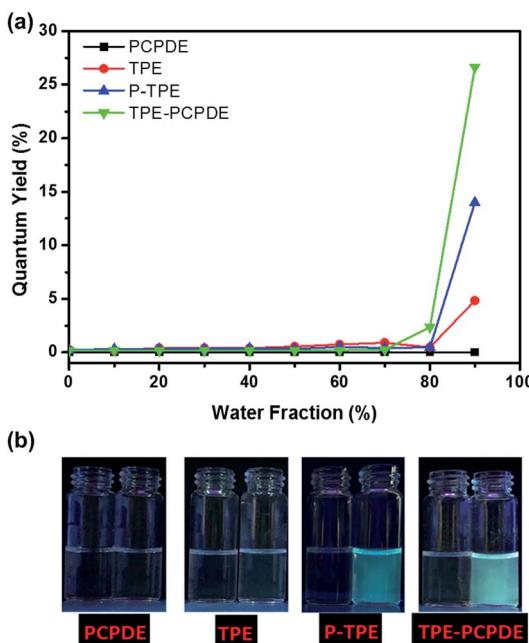


Fig. 6 (a) Plots of photoluminescence quantum yields of PCPDE, TPE, P-TPE and TPE-PCPDE versus water fraction and (b) the pictures of PCPDE, TPE, P-TPE and TPE-PCPDE in THF and in 90% water/THF mixture under the UV lamp at a wavelength at 365 nm.

fraction is 90% compared to that of **TPE** and **P-TPE** molecules. This indicates that through bond and through space conjugation of **TPE-PCPDE** which leads to a higher fluorescence than that of **TPE** and **P-TPE**. The pictures of **PCPDE**, **TPE**, **P-TPE** and **TPE-PCPDE** in THF and in 90% water fraction of the THF/water mixture under the UV lamp at a wavelength at 365 nm are shown in Fig. 6b. The notable fluorescence can be seen by naked eyes for **P-TPE** and **TPE-PCPDE** when the water fraction is 90% in the THF/water mixture. No fluorescence for **PCPDE** and very little fluorescence for **TPE** can be seen by naked eyes.

Conclusions

Br-PCPDE was synthesized from the corresponding dithia[3.3]paracyclophe in three steps through benzene Steven rearrangement, oxidation and thermal elimination reaction. **TPE-PCPDE** was successfully synthesized by the reaction of **Br-PCPDE** and monoboronic ester substituted tetraphenylethene using $\text{Pd}(\text{OAc})_2$ as a catalyst, *S*-Phos as a ligand and K_3PO_4 as a base. Functionalized paracyclophane dienes were fully characterized by nuclear magnetic resonance spectroscopy, high resolution mass spectrometry and X-ray crystallography. The longest absorption wavelength of **TPE-PCPDE** showed bathochromic shift compared to that of **P-TPE** due to the through space conjugation. **TPE-PCPDE** exhibited aggregation-induced emission characteristics when the water fraction was higher than 80% in the THF/water mixtures. **TPE-PCPDE** exhibited high fluorescence when the water fraction is 90% compared to that of **TPE** and **P-TPE** due to the through bond and through space conjugation. The ROMP of **TPE-PCPDE** by ruthenium carbene initiators is currently under investigation.

Conflicts of interest

There are no conflicts to declare.

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