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Surface modification of CoCr alloys by electrochemical reduction of diazonium salts†

M. A. Mezour,^{ab} Y. Oweis,^b A. A. El-Hadad,^{bc} S. Algizani,^b F. Tamimi*^b and M. Cerruti*^a

Tailoring the surface chemistry of CoCr alloys is of tremendous interest in many biomedical applications. In this work, we show that CoCr can be modified by diazonium electrografting provided the surface is not homogeneously covered with an oxide layer. Cyclic voltammetry (CV) and X-ray photoelectron spectroscopy (XPS) show the electrografting of a poly(aminophenylene) (PAP) layer on CoCr when treated at a reductive potential (CoCr_{-0.5} v), whereas no PAP film was formed on CoCr_{OCp} and CoCr₁ v, treated at open circuit and anodic potentials respectively. Based on XPS results, we attributed the electrografting to the formation of carbide bonds between PAP and the inhomogeneous thin oxide layer of CoCr_{-0.5} v. We then show an example of application of PAP coatings on CoCr and prove that the presence of a PAP coating on CoCr_{-0.5} v results in a 5-fold increase of the adherence of poly methyl methacrylate (PMMA) to PAP-coated CoCr compared to uncoated samples; this is of prime significance to improving the long-term stability of dental prostheses. These findings support the importance of reducing the oxide layer for effective functionalization of metal oxides with aryl diazonium salts and suggest a promising surface modification approach for biomedical applications.

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1. Introduction

Diazonium-based coatings have been suggested in many applications due to the strong bonds formed between aryl diazonium molecules and the surface of different materials, as well as the possibility of functionalizing the grafted molecules to further tailor surface properties.^{1–3} For instance, aryldiazonium-based compounds have been successfully employed as coupling agents to bind polymers and macromolecules to a variety of surfaces including carbon-based materials, noble metals and oxides.^{4–8}

Aryl diazonium grafting proceeds *via* a reductive electron transfer to the diazonium salt concerted with the cleavage of dinitrogen and generation of aryl radicals that bind to the electrode surface.^{9,10} It is generally accepted that aryl diazonium molecules form a covalent carbon–carbon¹¹ or metal–carbon¹² bond on carbon-based materials and gold substrates, respectively. This has been extensively studied and proved by spectroscopic studies and DFT calculations.^{11,12}

On the other hand, the evidence for a strong bond between aryl diazonium and oxidized surfaces is much scarcer.^{13–16} In

fact, some studies report that diazonium molecules interact only through weak electrostatic interactions with oxide surfaces, which makes the grafting inefficient.^{17–19} Other authors, instead, report the formation of covalent bond between aryl diazonium and the metal oxide.^{9,10,20} For instance, XPS, Raman and time-of-flight secondary ion mass spectrometry (ToF-SIMS) analysis revealed the existence of both M–O–C (metal–oxide–aryl) and M–C (metal–aryl) bonds on Cu,²¹ Al²² and Mn²³ oxides, and only M–O–C on Cr²⁴ and Fe oxides.²⁵ Other authors reported aryl diazonium grafting onto indium tin oxide (ITO),⁸ fluorine tin oxide (FTO),²⁶ and stainless steel,²⁷ but did not discuss the type of bonds involved.

In this work, we aim at understanding the nature of the bonds formed between aryl diazonium compounds and oxide surfaces using CoCr as a model metal oxide. CoCr alloys are commonly used in a variety of applications, including wind turbines,²⁸ engine components²⁹ as well as orthopedic and orthodontic devices,^{30,31} thanks to their high biocompatibility, superior wear and corrosion resistance, and good workability and ductility. These properties are directly related to the spontaneous formation of chromium oxide (Cr₂O₃) and cobalt oxide (CoO) on the CoCr surface. In dentistry, CoCr alloys are bound to acrylic polymers to manufacture dental devices such as removable partial dentures (RPDs)^{32,33} and resin-bonded partial prostheses.³⁴ However, these devices often fail prematurely due to the poor bond between the metal and the polymeric components.^{35,36}

Several methods have been used to increase the bond strength between polymers and CoCr alloys, including the

^aDepartment of Mining and Materials Engineering, McGill University, Montreal, Quebec H3A 2B2, Canada

^bFaculty of Dentistry, McGill University, 3640, Strathcona Anatomy and Dentistry Building, Rue University, Montreal, Quebec H3A 0C7, Canada. E-mail: marta.cerruti@mcgill.ca

^cPhysics Department, Faculty of Science, Al-Azhar University, Nasr City, Cairo, Egypt

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surface modification of the alloy with silane and phosphate molecules.^{37–39} However, the low stability of these molecular layers in the pH and temperature conditions experienced in the oral cavity urged us to explore diazonium coating as potential alternative.

Here, we first study how the presence of an oxidation layer on CoCr affects the electrochemical grafting of PAP, and then we show that a thin, cohesive, and covalently bonded PAP film significantly strengthens the adhesion of PMMA to CoCr.

2. Materials and methods

Materials

Cobalt chromium (CoCr alloy super 6) cylinders with a length of 18.71 mm and a diameter of 7.86 mm were obtained from Dental Depot (Fort Lauderdale, USA) and were cut into shorter disks (4.57 mm length). PPD (*para*-phenylenediamine), sodium nitrite (NaNO_2) and ruthenium hexamine trichloride ($\text{Ru}(\text{NH}_3)_6\text{Cl}_3$) were purchased from Sigma Aldrich (St. Louis, MO). PMMA and MMA were purchased from Great Lakes Orthodontics (Tonawanda, NY). All chemicals were used as received.

Methods

Surface polishing. The CoCr disks were polished with 2000 and 4000 silicon carbide abrasive paper followed by a fine polishing with Buehler 1 μm , 0.3 μm , and 0.05 μm alumina slurry. After each polishing step, CoCr disks were rinsed with deionized water, sonicated in an ultrasonic bath for 10 min, and dried using N_2 .

Electrochemical treatment. Electrochemical experiments were carried out using a CHI 760C workstation (CH Instruments Inc, Austin, TX) and a three-electrode cell. An Ag/AgCl (saturated KCl) electrode was used as a reference, platinum mesh as the

counter electrode, and the CoCr cylinders or disks as working electrode. All solutions were deaerated for 15 min with N_2 gas prior to electrochemical experiments. Open circuit potential (OCP) was measured for 15 min. The potentiodynamic polarization tests were performed by scanning the potential from -1 V to 1.2 V at 0.01 V s^{-1} in aqueous HCl solution (0.5 M). Electrochemical pre-treatment of CoCr surface was carried out by applying the potentials of -0.5 and 1 V (*vs.* Ag/AgCl electrode) for 15 minutes; these pre-treatments were compared with CoCr at OCP. The pre-treatment potentials were selected based on the domains found on the CoCr polarization curves (see Fig. 1 and discussion in the Results section). The oxide reformation was minimized by operating in O_2 free solution and by storing samples under N_2 .

Surface modification by electrochemical reduction of diazonium. The CoCr surface modification with *in situ* generated diazonium cations was performed in a 0.5 M HCl solution containing PPD and NaNO_2 . At first, one equivalent of NaNO_2 (2 mM) was added under stirring to the 0.5 M HCl electrolytic solution (20 ml) containing PPD (2 mM) at room temperature, and left to react for 5 min. The electrografting was then carried out by repetitive cyclic voltammetry (CV) cycles from 1 V to -0.5 V at 0.05 V s^{-1} and/or by holding the potential at -0.4 V for 5 min. The modified CoCr surfaces were then rinsed and sonicated for 5 min in deionized water. The relative electroactivity of the modified CoCr electrode (PAP-CoCr), which is inversely related to its blocking properties, was quantified by CV in 5 mM $\text{Ru}(\text{NH}_3)_6\text{Cl}_3/0.1$ M KBr at 0.05 V s^{-1} .

Tensile bond strength. The strength of the bond achieved between PAP-CoCr disks and PMMA was measured using a tensile test machine (H25K-S, Tinius Olsen Testing Machine Co., Inc Willow Grove, PA) set up at a constant speed of 10 mm min^{-1} . The specimens were loaded in the machine as

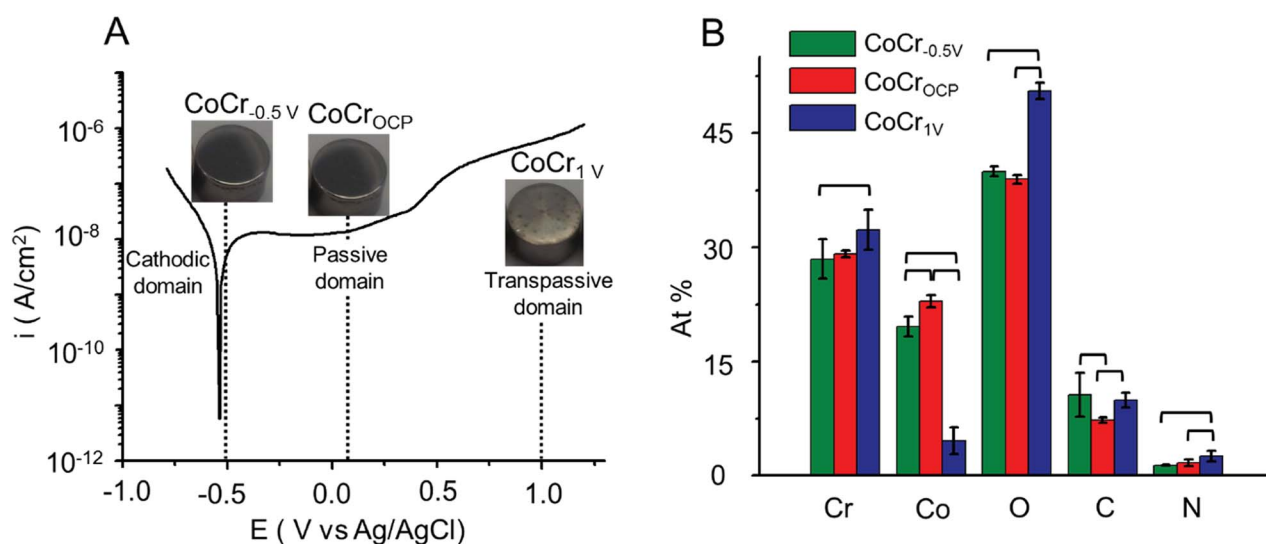


Fig. 1 (A) Potentiodynamic curve of CoCr alloy in HCl (0.5 M). The pictures show the visual appearance of CoCr_{-0.5V}, CoCr_{OCP} and CoCr_{1V} substrates. These samples were obtained by treatment of CoCr at the potentials of -0.5 V, OCP and 1 V, respectively, selected from the cathodic, passive and transpassive domains. (B) Elemental composition of CoCr_{-0.5V}, CoCr_{OCP} and CoCr_{1V} deduced from XPS survey scans. Brackets indicate statistically significant difference between two groups ($p < 0.05$).



shown in Fig. 5A. The upper and lower jaws held PMMA and CoCr respectively. To prepare the specimens, the PAP-CoCr disks were placed in a custom mold prepared from silicone (Exaktosil N 21, Bredent, Germany). PMMA (Biocryl Resin Acrylic, Great Lakes, NY) powder was mixed to MMA liquid monomer with a ratio of 2 : 1 to make a viscous solution that was poured around the PAP-CoCr disks in the mold and was left to set for 3 h at room temperature. The resulting PMMA-PAP-CoCr specimens were cylindrical in shape and measured 37 mm long and 7.86 mm in diameter. PAP free CoCr discs, referred as PMMA-CoCr, served as control.

X-ray photoelectron spectroscopy (XPS). XPS spectra were recorded using ThermoFisher Scientific K-alpha instrument equipped with a monochromatic Al K α X-ray source (1486.6 eV). Survey scans and high-resolution scans were collected with an X-ray spot size of 200 μ m and energy steps of 1 and 0.1 eV, respectively. Spectral energies were calibrated by setting the C-C binding energy of C_{1s} at 284.8 eV. Peak fitting was performed using the Thermo Avantage software (version 4.60). The evaluation procedure for peak fitting included Gaussian-Voigt curves functions and an iterated Shirley procedure for background subtraction. For consistency, a similar peak fitting procedure was applied on all samples (Table S1†).^{37,40}

Atomic force microscopy (AFM). AFM was used to examine the changes in surface roughness induced by different treatments of the CoCr surfaces. The AFM experiments were performed at ambient conditions on AFM instrument: a Nanoscope Multimode 8 equipped with a Nanoscope V controller (Bruker, Santa Barbara, CA). The topographies were acquired in ScanAsyst mode and Nanoscope 8.15r3 software. In ScanAsyst mode, silicon nitride cantilevers (ScanAsyst-Air Bruker) with a nominal spring constant of 0.4 N m⁻¹, a nominal resonance frequency of 50–90 kHz, and a nominal tip radius of 2 nm were used for imaging in air. AFM image analysis was performed using WSxM, 69 Nanoscope Analysis 1.4 software.

Scanning electron microscopy (SEM). The morphologies and elemental composition of the CoCr surfaces were examined using an Inspect-50 field emission scanning electron microscope (SEM) (FEI, Japan), at 10 kV operating voltage., equipped with an EDX unit.

Statistical analysis. All results were expressed as mean \pm standard deviation (SD) of at least three samples. Statistical analyses were performed on SPSS software (version 22; SPSS Inc, IBM corporation, Chicago IL). Changes (Δ) in the quantitative data were calculated by comparing pre-treatment and post-treatment measurements. All quantitative data were tested for normality using the Shapiro-Wilk test. For parametric data, a paired-sample *t*-test was used for changes within the same sample, and one-way ANOVA for comparison between more than two groups followed by Tukey HSD test to assess statistical differences. Whereas, Wilcoxon and Mann-Whitney (for comparison between two groups) and Kruskal-Wallis test (for comparison between more than two groups) were used for non-parametric data. The significance level was set at $p < 0.05$.

3. Results and discussion

To understand the nature of the bond formed between aryl diazonium layer and CoCr alloy, we started by testing different electrochemical pre-treatments aimed at controlling the thickness of the oxide layer present on the CoCr surface.

Electrochemical pre-treatment of CoCr surface

A nanometer-thick oxide film grows typically on CoCr upon exposure to oxygen or water at ambient condition.³⁷ This is evidenced by an open circuit potential (OCP) of 0.05 V. The OCP shifts towards more anodic potentials (Fig. S1†) with time indicating an increase in oxide thickness.⁹

Potentiodynamic polarization provides a good qualitative understanding of the oxidation and reduction processes occurring on the CoCr passive film. Fig. 1A represents a typical potentiodynamic curve of CoCr alloy recorded by scanning the potential from -1 V to 1.2 V at 0.01 V s⁻¹ in aqueous HCl solution (0.5 M).

The polarization curve can be divided into three different zones.³⁷ The cathodic zone ($E < -0.55$ V) is characterized by a negative current mainly dominated by the reduction of water and hydrogen evolution. In the passive zone (-0.3 to 0.6 V), the current remained relatively constant while the thickness of the passivating film grew proportionally with the potential. Finally, in the transpassive zone ($E > 0.6$ V) a marked increase of current is observed and associated with the dissolution of the oxide layer as well as water oxidation.³⁸

Based on the polarization curve, we selected two potentials in the active and transpassive domains, *i.e.* -0.5 V and 1 V, respectively, and we applied each potential on the CoCr disks for 15 min (samples labelled CoCr_{-0.5 v} and CoCr_{1 v}, respectively). We used CoCr at OCP, which is located in the passive domain of CoCr, as control (sample labelled CoCr_{OCP}). The visual appearance of the discs after this treatment is shown in the pictures reported in Fig. 1. While CoCr_{-0.5 v} and CoCr_{OCP} showed a smooth and mirror shiny surface similar to CoCr surface before the treatment (Fig. S2†), CoCr_{1 v} showed a rough and brighter appearance, indicative of accelerated surface corrosion.

We assessed the elemental composition of CoCr_{-0.5 v}, CoCr_{1 v}, and CoCr_{OCP} using XPS (Fig. 1B). All samples contained Co, Cr, O and C atoms, in addition to a small amount of N (<2.5%), possibly added during the casting process to improve the mechanical properties of CoCr alloy.³⁹ CoCr_{-0.5 v} and CoCr_{OCP} showed similar amounts of Cr and Co, with just approximately 7% more Cr than Co on average. In contrast, the surface layer on CoCr_{1 v} contained mainly Cr ($32 \pm 3\%$) and a lower amount of Co ($5 \pm 2\%$). This can be explained by the increasing thickness of the Cr rich oxide layer and the preferential dissolution of Co atoms in the transpassive domain.³⁷ The concentration of O significantly increased going from CoCr_{-0.5 v} to CoCr_{1 v}, indicating a proportional growth of oxide layer at anodic potential. Less C ($7 \pm 1\%$) is present on CoCr_{OCP} compared to CoCr_{-0.5 v} ($10 \pm 1\%$) and CoCr_{1 v} ($11 \pm 3\%$). This might originate from the carbon component of the CoCr alloy and adventitious carbon-based contaminants.



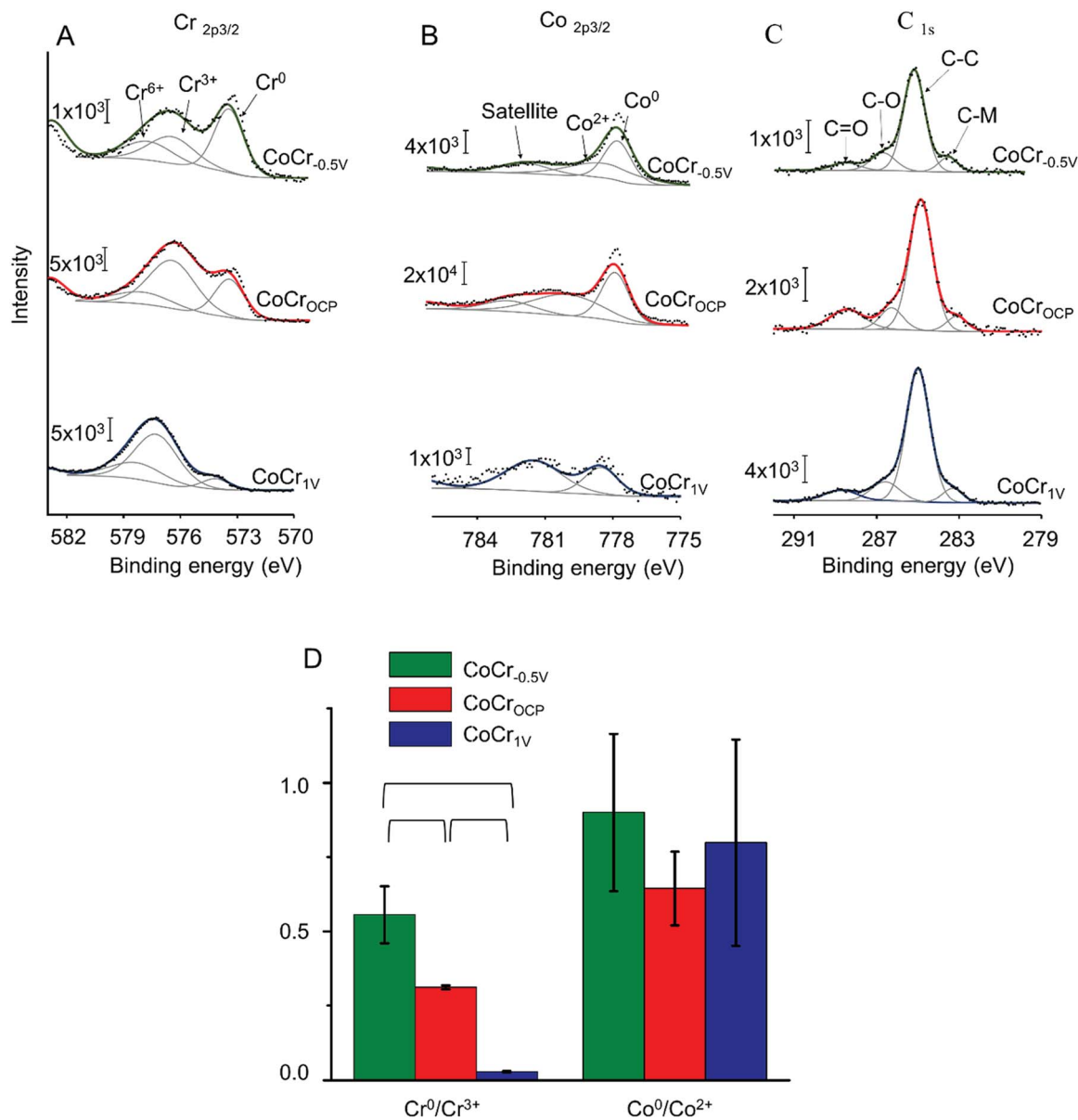


Fig. 2 Representative high resolution (A) Cr $_{2p_{3/2}}$, (B) Co $_{2p_{3/2}}$ and (C) C $_{1s}$ XPS spectra recorded on CoCr $_{-0.5V}$, CoCr $_{OCP}$, and CoCr $_{1V}$. Cr $_{2p_{1/2}}$ and Co $_{2p_{1/2}}$ peaks were omitted for clarity. In all panels (A–C), experimental data are shown in black, peak fits in gray and the overall fit in green (CoCr $_{1V}$), red (CoCr $_{OCP}$), or blue (CoCr $_{-0.5V}$). (D) Ratio of the areas of the Cr 0 /Cr $^{3+}$ and Co 0 /Co $^{2+}$ components measured on the high-resolution spectra of Cr and Co for CoCr $_{-0.5V}$, CoCr $_{OCP}$ and CoCr $_{1V}$. Brackets indicate statistically significant differences between two groups ($p < 0.05$).

To explore further the chemical environment of surface atoms, we recorded high resolution Cr $_{2p}$, Co $_{2p}$, and C $_{1s}$ spectra (Fig. 2 and Table S1 †). On all samples, the Cr $_{2p_{3/2}}$ and Co $_{2p_{3/2}}$ peaks showed a metallic component at low binding energy and a broader envelope of peaks attributed to multiplet splitting of the oxides and hydroxides peaks.⁴¹ In this study, the Cr $_{2p_{3/2}}$ spectra was fitted with a single peak at approximately 573.8 eV attributed to Cr 0 and two oxidized components at approximately 577 eV and 578.6 eV relative to Cr $^{3+}$ and Cr $^{6+}$ respectively. The Co $_{2p_{3/2}}$ envelope was deconvoluted with a metallic Co 0 peak at approximately 778 eV, an oxidized component at approximately 779.5 eV corresponding to Co $^{2+}$, and a satellite peak at approximately 782 eV related to Co $^{2+}$.⁴¹ The carbon peak can be

deconvoluted with four components: a low binding energy peak at 283 eV assigned to carbide bonds (C–M),¹² possibly due to the inter-dendritic precipitation of carbon during CoCr fabrication,⁴² a peak at 284.8 eV related to sp 3 carbon (C–C), and two peaks at higher binding energies (285 eV and 288 eV) attributed to C–O and C=O bonds respectively.⁴³ The latter three peaks are likely due to adventitious carbon-based contaminants usually found on surfaces exposed to atmosphere.²⁵

The Cr 0 /Cr $^{3+}$ ratio decreased significantly from 0.6 ± 0.1 for CoCr $_{-0.5V}$ to 0.3 ± 0.05 , and to 0.03 ± 0.01 for CoCr $_{OCP}$ and CoCr $_{1V}$ respectively (Fig. 2D). The high Cr 0 /Cr $^{3+}$ ratio observed on CoCr $_{-0.5V}$ might be explained by the partial elimination of the chromium rich oxide layer that generally forms at ambient



conditions (*i.e.* at OCP). *Vice versa*, anodic potentials accelerate the growth of the oxide layer, thus decreasing the $\text{Cr}^0/\text{Cr}^{3+}$ ratio. Differently from Cr, Co showed a higher $\text{Co}^0/\text{Co}^{2+}$ ratio that does not change significantly with surface treatment. This can be attributed to the lower susceptibility to oxidation of Co relative to Cr.

Overall, these results indicate that an oxide layer is formed on CoCr when treated at anodic potentials. On CoCr_{OCP} the oxide film is composed of a mixture of both Co and Cr, whereas on $\text{CoCr}_{1\text{V}}$, it is mainly dominated by chromium oxides. On the contrary, the oxide layer is partially dissolved on CoCr treated at cathodic potential, *i.e.* $\text{CoCr}_{-0.5\text{V}}$.

PAP binding as a function of the CoCr surface electrochemical pretreatment

The efficacy of PAP binding on the pretreated CoCr samples can be compared by performing cyclic voltammetry (CV) on these substrates in the presence of PPD and NaNO_2 (Fig. 3A).

The first cycle of the CV recorded for PAP binding on $\text{CoCr}_{-0.5\text{V}}$ (Fig. 3A, first row) showed a broad irreversible peak at -0.05 ± 0.1 V. In the subsequent CV scans, the current of this reductive peak is drastically diminished and vanishes in the fifth cycle, indicating the passivation of the $\text{CoCr}_{-0.5\text{V}}$ electrode by an electrically insulating layer of PAP molecules. This CV behavior is characteristic of a covalent bonding between the *in situ* generated diazonium radical and the CoCr surface.^{44–46} No oxidative peak was found in the CV, indicating no polyaniline

formation. The control experiments of PAP binding on glassy carbon and gold electrodes showed similar CV behaviour with an irreversible reduction wave at -0.3 V and -0.25 V respectively (Fig. S3†). On the other hand, CoCr_{OCP} and $\text{CoCr}_{1\text{V}}$ show a featureless CV suggesting the failure of PAP binding (Fig. 3A, second and third row).

We tested this hypothesis by investigating the blocking properties toward the $\text{Ru}(\text{NH}_3)_6^{3+/2+}$ redox couple of the substrates before and after PAP binding (Fig. 3B). The presence of a PAP layer is expected to slow the voltammetric response, causing an increase in the separation between the anodic and cathodic peak potential (ΔE_p) and a decrease in the intensity of the cathodic and anodic currents.

Even before electrochemical reduction of PAP (red lines in Fig. 3B), the three surfaces showed different barrier properties toward $\text{Ru}(\text{NH}_3)_6^{3+/2+}$. CoCr_{OCP} (Fig. 3B, second row) revealed a large ΔE_p (>0.9 V) confirming the presence of a passivating oxide layer on CoCr at OCP. This value is much higher than that measured on $\text{CoCr}_{-0.5\text{V}}$ (0.14 V, Fig. 3B, first row) and $\text{CoCr}_{1\text{V}}$ (0.6 V, Fig. 3B, third row). The lower barrier towards $\text{Ru}(\text{NH}_3)_6^{3+/2+}$ redox reaction shown by $\text{CoCr}_{-0.5\text{V}}$ agrees well with the thinner oxide layer observed at cathodic potentials by XPS, while on $\text{CoCr}_{1\text{V}}$ the inhomogeneous texture of the transpassive oxide might explain the low resistivity of the surface.⁴⁷

After PAP binding (black lines in Fig. 3B), no big changes are visible on the CVs measured on CoCr_{OCP} and $\text{CoCr}_{1\text{V}}$ (Fig. 3B). This result confirms that PAP binding was inhibited on the oxidized surfaces of CoCr_{OCP} and $\text{CoCr}_{1\text{V}}$. On the contrary, the PAP film grown on $\text{CoCr}_{-0.5\text{V}}$ strongly blocks the redox reactions of $\text{Ru}(\text{NH}_3)_6^{3+/2+}$, as shown by the marked increase in ΔE_p and the significant attenuation of the cathodic and anodic currents (Fig. 3B). The PAP layer formed on $\text{CoCr}_{-0.5\text{V}}$ is inhomogeneous (Fig. S5†) and a few nanometer thick, causing an increase in RMS roughness from 1.7 nm before coating to 6 nm after coating (Fig. S6†).

We further assessed PAP formation on the substrates by performing XPS on the PAP coated samples (Fig. 4). As for the uncoated CoCr substrates, the survey scans showed the presence of Cr, Co, O, C and N on all samples (Fig. S4†).

After PAP binding, we observed a significant increase in C accompanied with a decrease of Co and Cr. The O content remained almost constant and N increased slightly after electrografting on $\text{CoCr}_{-0.5\text{V}}$ and CoCr_{OCP} (Fig. 4A). This result indicates the presence of a carbon based organic coating at all CoCr surfaces (regardless of the nature of the oxide layer). The coating might be attributed to PAP film formation, but also to adventitious carbon-based contamination, since electrochemistry did not show evidence of electrografting taking place on CoCr_{OCP} and $\text{CoCr}_{1\text{V}}$ (Fig. 3).

To further assess the nature of this organic coating, we analysed the high resolution C_{1s} spectra of the three CoCr substrates (Fig. 4B). The C–C and C–O components increased significantly on all surfaces and are higher on the $\text{CoCr}_{-0.5\text{V}}$ than on others; no significant differences among surfaces are detected for the C–O component, and less C=O is found on $\text{CoCr}_{-0.5\text{V}}$ than on the other surfaces (Fig. 4C). However, again, these components may originate from both PAP layer formation

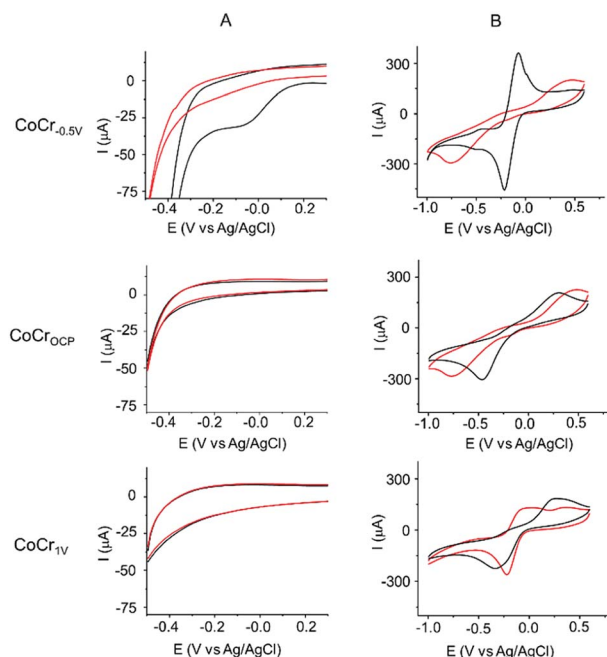


Fig. 3 CVs measured on $\text{CoCr}_{-0.5\text{V}}$ (first row), CoCr_{OCP} (second row) and $\text{CoCr}_{1\text{V}}$ (third row) during and after PAP grafting. (A) CVs performed in de-aerated 0.5 M HCl containing 2 mM of PPD and 1 equivalent of NaNO_2 . The black and red lines indicate the first and fifth cycles, respectively. (B) CVs performed in aqueous solution of 5 mM $\text{Ru}(\text{NH}_3)_6\text{Cl}_3/0.1$ M KBr before (black) and after (red) electrochemical binding of PAP. Scan rate was 50 mV s^{-1} .



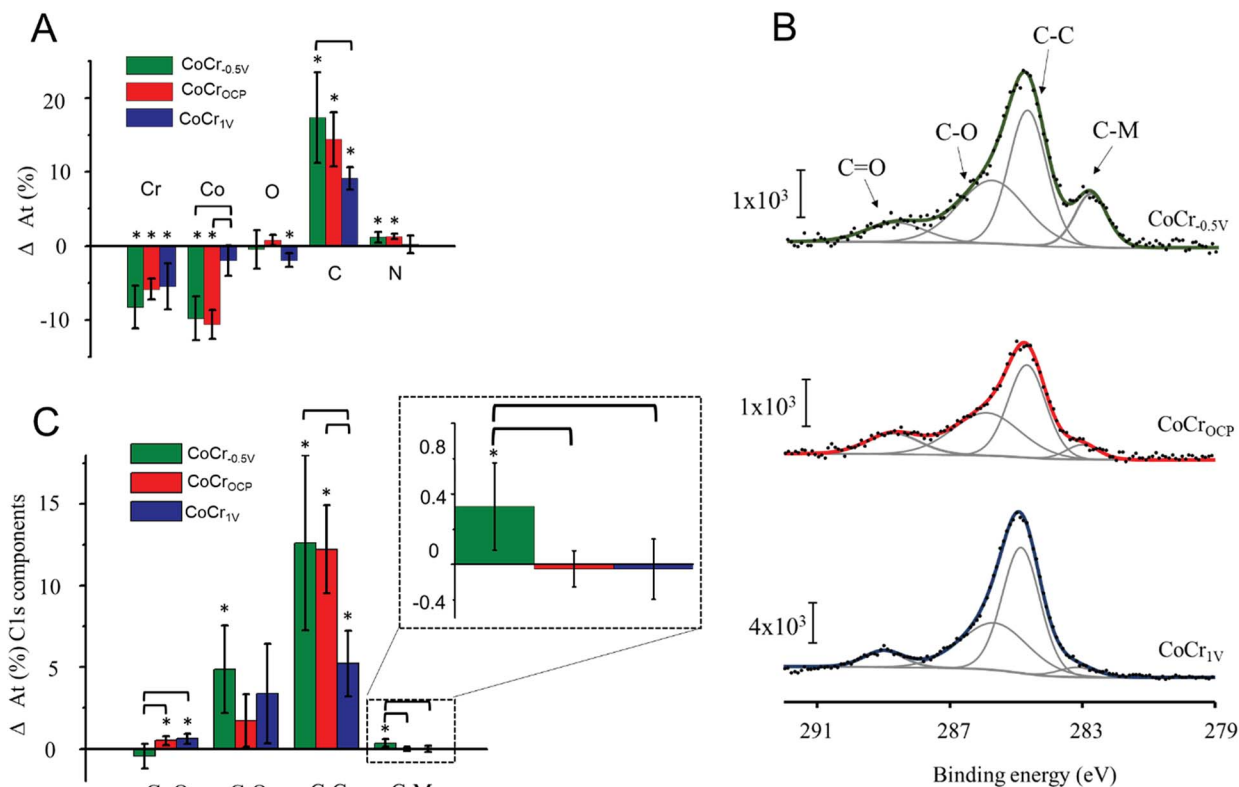


Fig. 4 PAP electrografting on $\text{CoCr}_{-0.5\text{V}}$, CoCr_{OCP} , and $\text{CoCr}_{1\text{V}}$ analyzed by XPS. (A) Differences in elemental composition measured by XPS survey scans before and after electrografting. (B) Representative high resolution $\text{C}_{1\text{s}}$ XPS spectra recorded after PAP electrografting. (C) Differences in relative intensities of the components used to deconvolute $\text{C}_{1\text{s}}$ spectra measured on samples before and after PAP electrografting. Asterisk signs indicate a significant change within the same sample group following PAP treatment ($p < 0.05$). Brackets indicate statistically significant differences between two groups ($p < 0.05$).

and adventitious contaminants that are hardly avoidable during XPS analysis. The C–M component instead can only originate from the covalent binding between PAP and CoCr. Due to the complex alloy surface composition, we were unable to determine the exact nature of the carbide bond. However, we hypothesize that PAP forms a bond with Cr, since this element is predominantly present on the surface of $\text{CoCr}_{-0.5\text{V}}$ (Fig. 2).

The C–M component was previously reported to be a direct proof for the existence of a covalent bond between the aryl group and different substrates (*i.e.* iron⁴⁸ and stainless steel⁴⁹).

After electrografting, this component increased significantly on $\text{CoCr}_{-0.5\text{V}}$ compared to CoCr_{OCP} and $\text{CoCr}_{1\text{V}}$. This result indicates that a covalent bond forms only between PAP and the surface of $\text{CoCr}_{-0.5\text{V}}$, which is covered with a thin and inhomogeneous oxide layer. The high resolution $\text{N}_{1\text{s}}$ spectra of $\text{CoCr}_{-0.5\text{V}}$ show a significant increase in the amine peak at 400 eV and decrease of the metal nitride peak at 394 eV after grafting compared to the pristine surface (Fig. S7†). These changes further confirm the presence of a PAP layer on this sample.

This result correlates well with previous studies on aryl diazonium grafting on oxide-free metals such as Au¹² and Cu.²¹ From the XPS and electrochemistry data we can conclude that the electrografting occurs on CoCr when its oxide is reduced.

Efficacy of PAP layer in promoting binding of PMMA to CoCr

We tested the ability of the electrografted PAP layer to improve the adhesion of PMMA to CoCr by simply adding PMMA to the PAP-treated CoCr surfaces and measuring the tensile bond strength at the PMMA–CoCr interface (Fig. 5A). Our hypothesis was that the branched structure of the PAP layer could be exploited to promote PMMA binding through polymeric chain entanglement and hydrogen bonding.

While there were no significant differences between the bond strength measured between PMMA and bare CoCr (0.9 ± 0.6 MPa) vs. PMMA and PAP-treated CoCr_{OCP} (1.4 ± 1.0 MPa), the tensile bond strength measured on PAP-treated $\text{CoCr}_{-0.5\text{V}}$ was almost five times higher (5.1 ± 1.2 MPa) (Fig. 5B). This result confirms that the presence of a PAP layer on $\text{CoCr}_{-0.5\text{V}}$ significantly increases the bond strength between PMMA and CoCr, likely due to chain entanglement and hydrogen bonds that form between the amine groups of PAP and the carbonyls of PMMA. The lack of significant differences between untreated CoCr and PAP-treated CoCr_{OCP} confirms the failure of PAP coating in the presence of an oxide layer on CoCr.

Several adhesives are reported in the literature to bind PMMA to metals, namely those based on silane and phosphate molecules.^{37–39} However, due to differences in test methodology and the fact that the reported coupling agents required



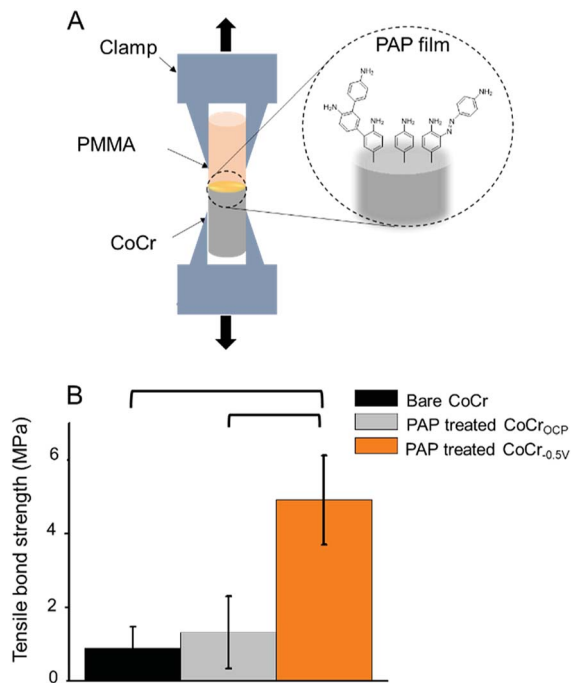


Fig. 5 (A) Schematic of the mechanical test designed to measure the interfacial bond strength between different CoCr samples and PMMA. (B) Tensile bond strength between PMMA and bare CoCr (black), PAP-treated CoCr_{OCP} (grey) and PAP-treated CoCr_{-0.5V} (orange). Brackets indicate statistically significant differences between groups ($p < 0.05$).

some type of surface pretreatment such as sandblasting,^{44,50} silica coating,⁵¹ and alloy priming,^{52,53} a straightforward comparison between these adhesives and diazonium chemistry is not possible. Instead, the difference between the control and treated groups could be used as an indication of the efficiency of diazonium treatment: in fact, the five fold improvement in tensile bond strength obtained with diazonium is comparable⁵⁴ than most of the values reported in literature for silane and phosphonates based adhesives, which range between 1.5 and 5 fold.

4. Conclusions

We show that the efficacy of electrografting PAP on CoCr is closely related to the composition of the CoCr surface. The presence of a passivating oxide layer on CoCr at OCP or when electrochemically treated at anodic potentials hinders PAP electrografting. In contrast, reducing the oxide thickness by treating CoCr at cathodic potentials favors PAP electrografting through the formation of metal carbide bonds. The presence of a PAP coating on CoCr_{-0.5V} results in a 5-fold increase in bond strength between PMMA and CoCr compared to uncoated CoCr samples. Thus, aryl diazonium coating offers a strong and stable alternative to the commonly used silane and phosphonate adhesives: in addition to the comparable or higher increase in strength measured in this work, the bonds formed by diazonium are much more stable in aqueous environments and large pH intervals thanks to the remarkable resistance of C–M bond to hydrolysis.^{46,55}

Conflicts of interest

There are no conflicts of interest to declare.

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Notes and references

- 1 M. Delamar, R. Hitmi, J. Pinson and J. M. Saveant, *J. Am. Chem. Soc.*, 1992, **114**, 5883–5884.
- 2 A. Chaussé, M. M. Chehimi, N. Karsi, J. Pinson, F. Podvorica and C. Vautrin-UI, *Chem. Mater.*, 2002, **14**, 392–400.
- 3 C. Cao, Y. Zhang, C. Jiang, M. Qi and G. Liu, *ACS Appl. Mater. Interfaces*, 2017, **9**, 5031–5049.
- 4 S. Mahouche-Chergui, S. Gam-Derouich, C. Mangeney and M. M. Chehimi, *Chem. Soc. Rev.*, 2011, **40**, 4143–4166.
- 5 H. Mahjoubi, J. M. Kinsella, M. Murshed and M. Cerruti, *ACS Appl. Mater. Interfaces*, 2014, **6**, 9975–9987.
- 6 O. Alageel, M. N. Abdallah, Z. Y. Luo, J. Del-Rio-Highsmith, M. Cerruti and F. Tamimi, *Dent. Mater.*, 2015, **31**, 105–114.
- 7 X. T. Le, N. D. Doan, T. Dequivre, P. Viel and S. Palacin, *ACS Appl. Mater. Interfaces*, 2014, **6**, 9085–9092.
- 8 S. Maldonado, T. J. Smith, R. D. Williams, S. Morin, E. Barton and K. J. Stevenson, *Langmuir*, 2006, **22**, 2884–2891.
- 9 P. Allongue, M. Delamar, B. Desbat, O. Fagebaume, R. Hitmi, J. Pinson and J.-M. Savéant, *J. Am. Chem. Soc.*, 1997, **119**, 201–207.
- 10 J. Pinson and F. Podvorica, *Chem. Soc. Rev.*, 2005, **34**, 429–439.
- 11 T. Itoh and R. L. McCreery, *J. Am. Chem. Soc.*, 2002, **124**, 10894–10902.
- 12 L. Laurentius, S. R. Stoyanov, S. Gusarov, A. Kovalenko, R. Du, G. P. Lopinski and M. T. McDermott, *ACS Nano*, 2011, **5**, 4219–4227.
- 13 J. M. Hicks, Z. Y. Wong, D. J. Scurr, N. Silman, S. K. Jackson, P. M. Mendes, J. W. Aylott and F. J. Rawson, *Langmuir*, 2017, **33**, 4924–4933.
- 14 V. Rebutini, E. Fazio, S. Santangelo, F. Neri, G. Caputo, C. Martin, T. Brousse, F. Favier and N. Pinna, *Chemical Modification of Graphene Oxide through Diazonium Chemistry and Its Influence on the Structure-Property Relationships of Graphene Oxide-Iron Oxide Nanocomposites*, *Chem.–Eur. J.*, 2015, **21**, 12465–12474.
- 15 M. G. Paulik, P. A. Brooksby, A. D. Abell and A. J. Downard, *J. Phys. Chem. C*, 2007, **111**, 7808–7815.
- 16 A. Jacques, S. Devillers, J. Delhalle and Z. Mekhalif, *Electrochim. Acta*, 2013, **109**, 781–789.



- 17 H. Ma, L. Lee, P. A. Brooksby, S. A. Brown, S. J. Fraser, K. C. Gordon, Y. R. Leroux, P. Hapiot and A. J. Downard, *J. Phys. Chem. C*, 2014, **118**, 5820–5826.
- 18 B. Cui, J.-Y. Gu, T. Chen, H.-J. Yan, D. Wang and L.-J. Wan, *Langmuir*, 2013, **29**, 2955–2960.
- 19 D. R. Jayasundara, R. J. Cullen and P. E. Colavita, *Chem. Mater.*, 2013, **25**, 1144–1152.
- 20 D. Belanger and J. Pinson, *Chem. Soc. Rev.*, 2011, **40**, 3995–4048.
- 21 B. L. Hurley and R. L. McCreery, *J. Electrochem. Soc.*, 2004, **151**, B252–B259.
- 22 Y. A. Atmane, L. Sicard, A. Lamouri, J. Pinson, M. I. Sicard, C. Masson, S. Nowak, P. Decorse, J.-Y. Piquemal and A. Galtayries, *J. Phys. Chem. C*, 2013, **117**, 26000–26006.
- 23 K. Bell, P. Brooksby, M. Polson and A. Downard, *Chem. Commun.*, 2014, **50**, 13687–13690.
- 24 M. Hinge, M. Ceccato, P. Kingshott, F. Besenbacher, S. U. Pedersen and K. Daasbjerg, *New J. Chem.*, 2009, **33**, 2405–2408.
- 25 K. Brymora, J. Fouineau, A. Eddarir, F. Chau, N. Yaacoub, J.-M. Grenèche, J. Pinson, S. Ammar and F. Calvayrac, *J. Nanopart. Res.*, 2015, **17**, 438.
- 26 F. Lamberti, S. Agnoli, L. Brigo, G. Granozzi, M. Giomo and N. Elvassore, *ACS Appl. Mater. Interfaces*, 2013, **5**, 12887–12894.
- 27 M. Hinge, E. S. Gonçalves, S. U. Pedersen and K. Daasbjerg, *Surf. Coat. Technol.*, 2010, **205**, 820–827.
- 28 S. Mercier, F. Iozzelli, M. P. Bacos and P. Josso, in *Materials Science Forum*, 2004, vol. 461–464, pp. 949–956.
- 29 Y. S. Wang, S. Narasimhan, J. M. Larson and S. K. Schaefer, *J. Mater. Eng. Perform.*, 1998, **7**, 53–65.
- 30 S. H. Teoh, *Engineering materials for biomedical applications*, World scientific, 2004.
- 31 L. Shi, D. O. Northwood and Z. Cao, *J. Mater. Sci.*, 1994, **29**, 1233–1238.
- 32 Y. S. Al Jabbari, *J. Adv. Prosthodont.*, 2014, **6**, 138–145.
- 33 B. Bergman, A. Hugoson and C. O. Olsson, *J. Oral Rehabil.*, 1995, **22**, 595–599.
- 34 B. E. Pjetursson, W. C. Tan, K. Tan, U. Brägger, M. Zwahlen and N. P. Lang, *Clin. Oral. Implants Res.*, 2008, **19**, 131–141.
- 35 D. G. Castner and B. D. Ratner, *Surf. Sci.*, 2002, **500**, 28–60.
- 36 S. Taksali, J. N. Grauer and A. R. Vaccaro, *Spine J.*, 2004, **4**, S231–S238.
- 37 A. Hodgson, S. Kurz, S. Virtanen, V. Fervel, C.-O. Olsson and S. Mischler, *Electrochim. Acta*, 2004, **49**, 2167–2178.
- 38 H. E. Placko, S. A. Brown and J. H. Payer, *J. Biomed. Mater. Res., Part A*, 1998, **39**, 292–299.
- 39 Y. Liao, E. Hoffman, M. Wimmer, A. Fischer, J. Jacobs and L. Marks, *Phys. Chem. Chem. Phys.*, 2013, **15**, 746–756.
- 40 S. Hiromoto, E. Onodera, A. Chiba, K. Asami and T. Hanawa, *Biomaterials*, 2005, **26**, 4912–4923.
- 41 A. Katerski, A. Mere, V. Kazlauskienė, J. Miskinis, A. Saar, L. Matisen, A. Kikas and M. Krunkas, *Thin Solid Films*, 2008, **516**, 7110–7115.
- 42 T. Kilner, R. Pilliar, G. Weatherly and C. Allibert, *J. Biomed. Mater. Res., Part A*, 1982, **16**, 63–79.
- 43 J. Wilson, J. Walton and G. Beamson, *J. Electron Spectrosc. Relat. Phenom.*, 2001, **121**, 183–201.
- 44 T. E. Jacobson, *Int. J. Prosthodont.*, 1989, **2**, 163–172.
- 45 H. Shimizu, K. S. Kurtz, Y. Tachii and Y. Takahashi, *J. Dent.*, 2006, **34**, 117–122.
- 46 M. Rosso, A. Arafat, K. Schroën, M. Giesbers, C. S. Roper, R. Maboudian and H. Zuilhof, *Langmuir*, 2008, **24**, 4007–4012.
- 47 M. Bojinov, G. Fabricius, T. Laitinen, T. Saario and G. Sundholm, *Electrochim. Acta*, 1998, **44**, 247–261.
- 48 K. Boukerma, M. M. Chehimi, J. Pinson and C. Blomfield, *Langmuir*, 2003, **19**, 6333–6335.
- 49 X. T. Le, G. Zeb, P. Jégou and T. Berthelot, *Electrochim. Acta*, 2012, **71**, 66–72.
- 50 L. Mudford, R. Curtis and J. Walter, *J. Dent.*, 1997, **25**, 415–421.
- 51 K. B. May, J. Fox, M. E. Razzoog and B. R. Lang, *J. Prosthet. Dent.*, 1995, **73**, 428–431.
- 52 C. Ohkubo, I. Watanabe, T. Hosoi and T. Okabe, *J. Prosthet. Dent.*, 2000, **83**, 50–57.
- 53 M. Bulbul and B. Kesim, *J. Prosthet. Dent.*, 2010, **103**, 303–308.
- 54 S.-S. Kim, M.-S. Vang, H.-S. Yang, S.-W. Park and H.-P. Lim, *J. Adv. Prosthodont.*, 2009, **1**, 41–46.
- 55 R. A. Franking, E. C. Landis and R. J. Hamers, *Langmuir*, 2009, **25**, 10676–10684.

