


 Cite this: *RSC Adv.*, 2018, 8, 14833

Received 22nd March 2018

Accepted 27th March 2018

DOI: 10.1039/c8ra02526f

rsc.li/rsc-advances

The influence of substituent field and resonance effects on the ease of N-heterocyclic carbene formation from imidazolium rings†

 Wojciech P. Oziminski ^{*ab} and Christopher A. Ramsden ^{*c}

Using a set of twelve selected substituents, the influence of substituent properties on the ease of deprotonation of imidazolium cations and mesoionic imidazolium-4-olates measured by the CREF index has been investigated. Significant correlations between CREF values and the Swain and Lupton field (*F*) and resonance (*R*) substituent constants have been found. In all cases the field effect has the greatest influence but resonance effects are also significant.

Introduction

We have demonstrated that the deprotonation energy (ZPE corrected) of N-heterocyclic carbene (NHC) precursors, formulated as an index, provides a convenient overview of the relative ease of formation of a wide range of normal (nNHC), abnormal (aNHC) and remote (rNHC) neutral and anionic NHCs.^{1,2} This index, which we describe as the CREF (Carbene Relative Energy of Formation) index, is easily calculated using modern DFT methods. The results direct attention to properties of novel or unexplored molecules of potential interest, including their potential σ -bonding strength as ligands.

In addition, we have observed, in accord with the Hammond postulate,³ that there is good correlation between CREF indexes and experimental rate constants for deprotonation or H–D exchange for a range of heterocycles for which kinetic data is available.

Our previous studies have been largely restricted to the influence of ring heteroatoms (N, O and S) on NHC formation. We now report a study of the influence of ring substituents.

Results and discussion

To explore the influence of ring substituents on NHC formation, we selected a set of eleven common substituents that represents a broad spectrum of sigma- and pi-electron-donor and electron-acceptor characteristics. The unsubstituted molecules were also

included for reference. These substituents are shown in Table 1 together with their Swain and Lupton field (*F*) and resonance (*R*) constants.⁴ Each parameter shows good variation (*F* 0.0 to +0.65 and *R* –0.73 to +0.42), and there is low correlation between the two (*r* 0.56).

Initially we studied the influence of 4-substituents on the formation of the NHCs **2** by deprotonation of the imidazolium cations **1**. The CREF values for these transformations are shown in Table 2 and multiple regression analysis shows a statistically significant relationship (eqn (1), Table 2A) between CREF values and the Swain and Lupton constants. As might be expected, since deprotonation primarily involves heterolytic cleavage of a σ bond, the larger effect is the field effect, but the resonance effect is also significant. Neither parameter alone (*F* nor *R*) gives a significant correlation. Using eqn (1) the CREF values of most common organic substituents can be estimated using published *F* and *R* values.

The observed substituent effects depend on the relative effects of substituents on reactant and product. Electron-

Table 1 Substituent set and Swain and Lupton substituent constants (*F* and *R*)⁴

Substituent	<i>F</i>	<i>R</i>
NHMe	0.03	–0.73
OMe	0.29	–0.56
F	0.45	–0.39
Cl	0.42	–0.19
Me	0.01	–0.18
cycHex	0.03	–0.18
OCF ₃	0.39	–0.04
H	0.00	0.00
NO ₂	0.65	0.13
CN	0.51	0.15
CF ₃	0.38	0.16
NO	0.49	0.42

^aNational Medicines Institute, 30/34 Chelmska Street, 00-725 Warsaw, Poland

^bFaculty of Pharmacy, Medical University of Warsaw, 1 Banacha Street, 02-097 Warsaw, Poland

^cSchool of Chemical and Physical Sciences, Lennard-Jones Laboratories, Keele University, Staffordshire ST5 5BG, UK. E-mail: c.a.ramsden@keele.ac.uk; wojozim@gmail.com

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c8ra02526f



Table 2 CREF index values and correlations for formation of normal NHCs

$$\text{CREF} = 0.416(\pm 0.002) - 0.038(\pm 0.004)F - 0.021(\pm 0.003)R \quad (1),$$

$$n = 12, r = 0.98, s = 0.003, F = 134.0, p < 0.0001$$

$$\text{CREF} = 0.418(\pm 0.004) - 0.069(\pm 0.009)F - 0.028(\pm 0.006)R \quad (2),$$

$$n = 12, r = 0.97, s = 0.006, F = 69.1, p < 0.0001$$

	Subst. R	Precursor 1 E + ZPE ^a	Product 2 E + ZPE ^a	CREF		Subst. R	Precursor 3 E + ZPE ^a	Product 4 E + ZPE ^a	CREF
a	NHMe	−399.8209	−399.3944	0.427	a	NHMe	−494.4616	−494.0342	0.427
b	OMe	−419.6950	−419.2753	0.420	b	OMe	−534.2104	−533.7884	0.422
c	F	−404.4263	−404.0218	0.405	c	F	−503.6798	−503.2840	0.396
d	Cl	−764.7876	−764.3815	0.406	d	Cl	−1224.4070	−1224.0071	0.400
e	Me	−344.4758	−344.0567	0.419	e	Me	−383.7819	−383.3579	0.424
f	cycHex	−539.7656	−539.3424	0.423	f	cycHex	−774.3580	−773.9273	0.431
g	OCF ₃	−717.5314	−717.1286	0.403	g	OCF ₃	−1129.8927	−1129.4988	0.394
h	H	−305.1679	−304.7548	0.413	h	H	−305.1679	−304.7548	0.413
i	NO ₂	−509.6969	−509.3096	0.387	i	NO ₂	−714.2134	−713.8449	0.369
j	CN	−397.4147	−397.0229	0.392	j	CN	−489.6589	−489.2844	0.374
k	CF ₃	−642.2949	−641.8977	0.397	k	CF ₃	−979.4124	−979.0276	0.385
l	NO	−434.4790	−434.0888	0.390	l	NO	−563.7831	−563.4090	0.374

^a Hartrees.

^a Hartrees.

withdrawing substituents reduce density around neighbouring atomic nuclei, exposing greater nuclear charge, and thus lower σ -orbital energies in both reactant and product.

Since deprotonation effectively involves the transfer of one electron from a hydrogen 1s environment to a carbon σ environment, the lowering of σ energy levels might be expected to not only polarise the C–H bond in the precursor but also stabilise the product, which gains an extra sp^2 σ electron. This is consistent with electron-withdrawing substituents, *e.g.*, NO₂, CN, CF₃, lowering CREF values. Other effects may contribute; substituents that are π donors, *e.g.*, NHMe, OMe, may stabilise the cationic precursors **1** relative to the products **2** contributing to higher CREF values. The origins of the electronic effects of substituents may be complex but eqn (1) conveniently summarises the net effect and there are both σ and π contributions.

Fig. 1 illustrates the effect of three structural changes (cations **5**–**7**) on the CREF index of the imidazolium cation **1h** (CREF 0.413). A 4-trifluoromethyl substituent lowers the value by 0.016 units, which is comparable to the effect of 4-aza substitution. Cations **5** and **6** both have acidity comparable to the 1,3-thiazolium cation **7** (Fig. 1). The effect of a 4-nitro substituent (**1i**, Table 2) is even larger (Δ CREF 0.026 units). We have previously estimated that a reduction of the index by 0.01 units leads to a reduction of aqueous pK_a of 2.4.² On this basis, the estimated pK_a of 16.8 for the 4-nitroimidazolium ring is comparable to that of the 1,3-oxazolium cation (pK_a 16.9).⁵ The associated weakening of the C(2)–H bond in 4-

nitroimidazolium rings is consistent with their formation of hydrogen bonds with chloride ions.⁶

The effect of a second substituent is cumulative. An equally significant correlation (eqn (2), Table 2B) was obtained for deprotonation of the disubstituted derivatives **3** giving the NHCs **4**. Relative to eqn (1), the coefficients for *F* and *R* in eqn (2) are increased by 1.82 and 1.40, respectively. A number of 4,5-disubstituted imidazolium salts with electron-withdrawing substituents have been investigated as potential energetic ionic liquids.⁷ Based on the calculated CREF values, the dicyano derivative **3j** (CREF 0.374) can be predicted to have $pK_a \sim 13.6$, and to have C–H acidity comparable to 1,4-dialkyltetrazolium salts (CREF 0.372).^{8–10}

Similar statistically significant correlations between CREF index values and the Swain and Lupton constants were found for the formation the substituted aNHCs **9** (Table 3A) and **11** (Table 3B). For the deprotonation **8** \rightarrow **9** the contributions of *F*

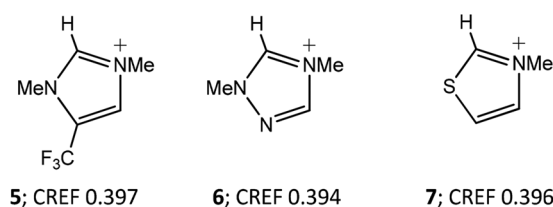
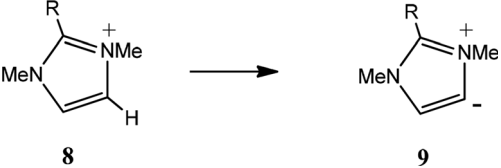
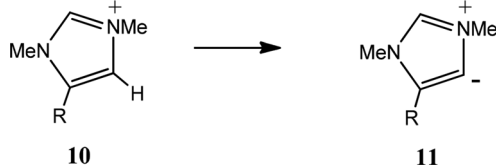


Fig. 1 Three structural changes to the imidazolium cation that result in similar CREF values.

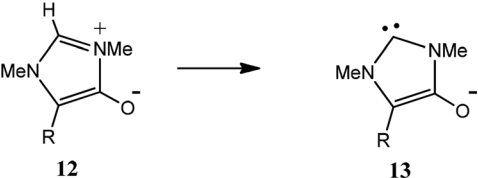
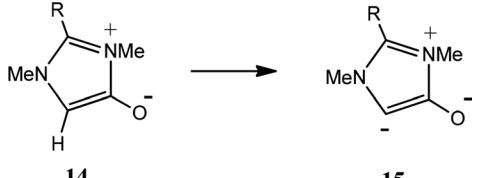


Table 3 CREF index values and correlations for formation of abnormal NHCs

A					B				
									
CREF = 0.445(±0.002) – 0.042(±0.006) <i>F</i> – 0.022(±0.004) <i>R</i> (3), <i>n</i> = 12, <i>r</i> = 0.97, <i>s</i> = 0.004, <i>F</i> = 72.8, <i>p</i> < 0.0001					CREF = 0.445(±0.002) – 0.048(±0.005) <i>F</i> – 0.016(±0.003) <i>R</i> (4), <i>n</i> = 12, <i>r</i> = 0.98, <i>s</i> = 0.003, <i>F</i> = 106.0, <i>p</i> < 0.0001				
Subst. R	Precursor 8 E + ZPE ^a	Product 9 E + ZPE ^a	CREF		Subst. R	Precursor 10 E + ZPE ^a	Product 11 E + ZPE ^a	CREF	
a	NHMe	–399.8211	–399.3675	0.454	a	NHMe	–399.8209	–399.3696	0.451
b	OMe	–419.6930	–419.2447	0.448	b	OMe	–419.6950	–419.2521	0.443
c	F	–404.4274	–403.9925	0.435	c	F	–404.4263	–403.9974	0.429
d	Cl	–764.7879	–764.3522	0.436	d	Cl	–764.7876	–764.3553	0.432
e	Me	–344.4771	–344.0280	0.449	e	Me	–344.4758	–344.0276	0.448
f	cycHex	–539.7674	–539.3132	0.454	f	cycHex	–539.7656	–539.3137	0.452
g	OCF ₃	–717.5305	–717.0970	0.434	g	OCF ₃	–717.5314	–717.1054	0.426
h	H	–305.1679	–304.7258	0.442	h	H	–305.1679	–304.7258	0.442
i	NO ₂	–509.6897	–509.2772	0.413	i	NO ₂	–509.6969	–509.2861	0.411
j	CN	–397.4117	–396.9924	0.419	j	CN	–397.4147	–396.9984	0.416
k	CF ₃	–642.2870	–641.8610	0.426	k	CF ₃	–642.2949	–641.8727	0.422
l	NO	–434.4773	–434.0640	0.413	l	NO	–434.4790	–434.0628	0.416

^a Hartrees.

Table 4 CREF index values and correlations for formation of anionic NHCs

A					B				
									
CREF = 0.573(±0.003) – 0.040(±0.008) <i>F</i> – 0.023(±0.005) <i>R</i> (5), <i>n</i> = 12, <i>r</i> = 0.95, <i>s</i> = 0.005, <i>F</i> = 44.0, <i>p</i> < 0.0001					CREF = 0.615(±0.004) – 0.069(±0.011) <i>F</i> – 0.026(±0.008) <i>R</i> (6), <i>n</i> = 12, <i>r</i> = 0.95, <i>s</i> = 0.007, <i>F</i> = 46.3, <i>p</i> < 0.0001				
Subst. R	Precursor 12 E + ZPE ^a	Product 13 E + ZPE ^a	CREF		Subst. R	Precursor 14 E + ZPE ^a	Product 15 E + ZPE ^a	CREF	
a	NHMe	–474.6740	–474.0966	0.577	a	NHMe	–474.6720	–474.0542	0.618
b	OMe	–494.5473	–493.9711	0.576	b	OMe	–494.5498	–493.9350	0.615
c	F	–479.2917	–478.7231	0.569	c	F	–479.2966	–478.6949	0.602
d	Cl	–839.6561	–839.0907	0.565	d	Cl	–839.6556	–839.0623	0.593
e	Me	–419.3327	–418.7548	0.578	e	Me	–419.3331	–418.7112	0.622
f	cycHex	–614.6214	–614.0457	0.576	f	cycHex	–614.6186	–613.9984	0.620
g	OCF ₃	–792.4036	–791.8447	0.559	g	OCF ₃	–792.4013	–791.8160	0.585
h	H	–380.0312	–379.4554	0.576	h	H	–380.0312	–379.4121	0.619
i	NO ₂	–584.5940	–584.0559	0.538	i	NO ₂	–584.5863	–584.0241	0.562
j	CN	–472.3062	–471.7592	0.547	j	CN	–472.2981	–471.7166	0.582
k	CF ₃	–717.1766	–716.6242	0.552	k	CF ₃	–717.1658	–716.5785	0.587
l	NO	–509.3800	–508.8396	0.540	l	NO	–509.3748	–508.8135	0.561

^a Hartrees.

and R are similar to those for the formation of the nNHCs **2** (Table 2A). This is consistent with the fact that the relationships between substituent location and carbene centre are similar. For the deprotonation **10** \rightarrow **11** the contribution of F is greater and that of R is slightly smaller; in this case the substituents are adjacent to the carbene centre.

We have also investigated the effect of substituents on anionic NHC formation (Table 4). The CREF index values for anionic NHCs are higher, not least because the deprotonation process involves charge separation. Again, significant correlations are observed. For the formation of the nNHCs (**12** \rightarrow **13**) (Table 4A), the correlation between CREF values and the electronic parameters F and R (eqn (5)) is very similar to that for the deprotonations **1** \rightarrow **2** (Table 2A) and **8** \rightarrow **9** (Table 3A). For formation of the anionic aNHCs (**14** \rightarrow **15**) (Table 4B), the contribution of the field effect (F) (eqn (6)) is appreciably larger. The reason for this enhanced field effect is not clear, but it may be due to the association of position 4 with particularly high electron density.

Computational details

Calculations were performed using the Gaussian 09 program,¹¹ and the hybrid B3LYP functional,^{12,13} accompanied by the 6-311++G(d,p) basis set,¹⁴ was used. All geometry optimizations were followed by frequency calculations to ensure that the stationary points obtained were true minima on the potential energy surface and to calculate the zero-point vibrational corrections (ZPE) to energy. The CREF index was calculated as the (ZPE-corrected) energy difference between protonated precursor and the carbene (obtained *via* C–H bond heterolytic cleavage).

The multiple regression relationships and associated statistics based on the ordinary least squares method were determined using online software.¹⁵ The meaning of the statistical parameters are the following: n = sample size, r = multiple regression correlation coefficient, s = residual standard deviation, F = Fisher test value, p = p -value.

Conclusions

Statistically significant correlations have been found between the CREF index, which is a measure of C–H acidity, and electronic substituent constants for imidazolium cations and mesoionic imidazolium-4-olates. Both σ and π contributions to the acidity are observed. Derivatives disubstituted with electron-withdrawing groups (*e.g.*, CF_3 , CN , NO_2) are predicted to have acidities comparable to 1,3-dialkyltetrazolium cations,⁹ but without the instability associated with rapid loss of nitrogen.¹⁶

The CREF index is a measure of the strength of C–H σ bonds in heteroaromatic rings, and, by analogy, the strength of C–C σ bonds (relevant to decarboxylation) and C-metal σ bonds, for example when NHCs act as metal ligands. In the latter case, there is no measure of C-metal π bonding (or steric effects). In the context of ligand binding, the effectiveness of metal complexes of the imidazole-ylidenes **4** as catalysts has been found to vary with ring substituents.¹⁷ This has been attributed

to possible variation in π bonding between ligand and metal. This may be the case, but the correlations described in Table 2 show that π effects also influence σ bond strength and thus catalytic efficiency.

In 1966 Olofson and Landesberg concluded, justifiably, that the use of LCAO-MO methods to predict H/D exchange in heterocyclic cations was futile.¹⁸ The DFT approach that we have explored previously^{1,2} and extended in this study seems more promising. Experimentally determined physical properties of heterocycles are limited and there are gaps in the literature.¹⁹ It is hoped that the DFT-based approach described above will be helpful in understanding variations in physical properties of heteroaromatic rings and augment experimental data.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Computational Grant G36-9 from the Interdisciplinary Centre for Mathematical and Computational Modelling at Warsaw University (ICM UW) is gratefully acknowledged.

Notes and references

- 1 C. A. Ramsden and W. P. Oziminski, *J. Org. Chem.*, 2016, **81**, 10295.
- 2 C. A. Ramsden and W. P. Oziminski, *J. Org. Chem.*, 2017, **82**, 12485.
- 3 G. S. Hammond, *J. Am. Chem. Soc.*, 1955, **77**, 334.
- 4 C. Hansch, A. Leo and R. W. Taft, *Chem. Rev.*, 1991, **91**, 165.
- 5 T. L. Amyes, S. T. Diver, J. P. Richard, F. M. Rivas and K. Toth, *J. Am. Chem. Soc.*, 2004, **126**, 4366.
- 6 H. Ihm, S. Yun, H. G. Kim, J. K. Kim and K. S. Kim, *Org. Lett.*, 2002, **4**, 2897.
- 7 A. R. Katritzky, H. Yang, D. Zhang, K. Kirichenko, M. Smiglak, J. D. Holbrey, W. M. Reichert and R. D. Rogers, *New J. Chem.*, 2006, **30**, 349.
- 8 R. A. Olofson, W. R. Thompson and J. S. Michelman, *J. Am. Chem. Soc.*, 1964, **86**, 1865.
- 9 A. C. Rochat and R. A. Olofson, *Tetrahedron Lett.*, 1969, **10**, 3377.
- 10 G. D. Frey, K. Öfele, H. G. Krist, E. Herdtweck and W. A. Herrmann, *Inorg. Chim. Acta*, 2006, **359**, 2622.
- 11 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo,



- J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc., Wallingford CT, 2009.
- 12 A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648.
- 13 P. J. Stephens, F. J. Devlin, C. F. Chabalowski and M. J. Frisch, *J. Phys. Chem.*, 1994, **98**, 11623.
- 14 R. Krishnan, J. S. Binkley, R. Seeger and J. A. Pople, *J. Chem. Phys.*, 1980, **72**, 650.
- 15 P. Wessa, *Multiple Regression (v1.0.48) in Free Statistics Software (v1.2.1)*, Office for Research Development and Education, 2017, https://www.wessa.net/rwasp_multipleregression.wasp.
- 16 D. M. Zimmerman and R. A. Olofson, *Tetrahedron Lett.*, 1970, **11**, 3453.
- 17 D. M. Khranov, E. L. Rosen, J. A. V. Er, P. D. Vu, V. M. Lynch and C. W. Bielawski, *Tetrahedron*, 2008, **64**, 6853.
- 18 R. A. Olofson and J. M. Landesberg, *J. Am. Chem. Soc.*, 1966, **88**, 4263.
- 19 A. R. Katritzky, C. A. Ramsden, J. A. Joule and V. V. Zhdankin, *Handbook of Heterocyclic Chemistry*, 3rd edn, Elsevier, Oxford, 2010.

