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Switching the optical and electrical properties of carbon nanotube hybrid films using a photoresponsive dispersant as a dopant†

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The light-induced switching of the optical and electrical properties of single-walled carbon nanotubes (SWCNTs) was demonstrated following hybridisation with an azobenzene-based photoresponsive dispersant in solid film. The resultant SWCNT/photoresponsive dispersant hybrid film showed switching of absorbance in the near-infrared region by the irradiation of UV and visible lights, which was caused by the change of the doping ability of the dispersant to SWCNT. In addition, the switching of the electrical properties of the SWCNT hybrid film was also observed.

Introduction

Carbon nanomaterials like single-walled carbon nanotubes (SWCNTs) and graphene are promising candidates in numerous applications such as electronics,^{1–4} optical devices,^{5,6} and nanocomposites^{7–10} because of their extraordinary properties. Recently, stimuli-responsive carbon nanomaterial hybrids have attracted considerable interest because they adopt the exceptional features of carbon nanomaterials with readily available stimuli. Among the various stimuli accessible, light is characteristic because of its high temporospatial resolution, its eco-friendliness, and the wide selection of possible wavelengths. Hybrids of carbon nanomaterials and photoresponsive compounds have great potential applications as sensors,¹¹ memory devices,¹² electrical devices,^{13–16} and storage for solar thermal energy.^{17–19} In fact, various types of photoactive carbon nanomaterials have been reported, with light being used most often to induce switching of the electrical conductivity of nanocarbon devices. For example, azobenzene-immobilised SWCNTs and graphene have been shown to exhibit switching of their electrical conductivity by UV and visible light irradiation, which may prove useful in memory and sensor devices.^{14,20} In another example, photoresponsive field-effect transistor devices were demonstrated using spiropyran-immobilised graphene.²¹

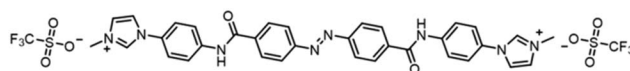
In our previous work, we reported a photoresponsive dispersant that can disperse carbon nanomaterials into water and organic solvent.^{22–25} Moreover, upon irradiation by UV and visible light, the dispersibility of these carbon nanomaterials

can be switched between dispersive and precipitative states based on the photoisomerisation of the dispersant.²³ In this study, the bulky tetra-ammonium cationic moiety of the photoresponsive dispersant was replaced with an imidazole moiety, giving rise to a greatly enhanced planarity of the entire molecule that is much higher than that of previously reported dispersants. The present work demonstrates novel photoactive carbon nanomaterials using SWCNTs and an azobenzene-based dispersant with a planer structure (**AB**, Scheme 1). The SWCNT/**AB** hybrid was obtained by a simple wet-coating method from a uniform dispersion of SWCNT/**AB**. Interestingly, the SWCNT/**AB** hybrid film shows not only electrical conductivity switching but also switching of its near-infrared (NIR) absorbance with irradiation of UV and visible light. Notably, these unique switching properties may lead to innovative applications in optoelectronic devices which can input UV and output both NIR and electric signals (*e.g.* sensor, memory, optical detector, and optical filter).

Experimental

Materials and instruments

The reagents and solvents were purchased from Aldrich, Wako Pure Chemical Industries Ltd., and Tokyo Chemical Industry. Single-walled carbon nanotubes (SWCNT), which were synthesised from a high-pressure carbon monoxide method (HiPco), were purchased from NanoIntegris (*d* = 0.8–1.2 nm, length = 100–1000 nm).



Scheme 1 Chemical structure of the azobenzene-derived cation with the trifluoromethanesulfonyl anion (**AB**).

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UV-Vis-NIR spectra were measured with a V-670 spectrometer (Jasco) using quartz cells (pass length = 0.5 mm). UV light was irradiated by a UV-LED (EXECURE H-1VH4-V1, HOYA Candeo Optonics) and Super-high pressure mercury lamp (USHIO, E016 with optical cut filter). Atomic Force Microscope (AFM) images (tapping mode) were recorded using an Innova (Veeco Instruments, Inc.) with a Si probe (RTESP-300, $f_0 = 300$ kHz, Bruker). Sheet resistance was measured with a Loresta-AX resistivity meter (MCP-T370, Mitsubishi Chemical Analytech). Raman spectroscopy was measured by an NRS-1000 spectrometer (Jasco) with a 532 nm laser source (2.5 mW). Elemental analysis was recorded using vario MICRO cube (Elementar).

Synthesis of photoresponsive dispersant AB

4-(1H-Imidazol-1-yl)aniline (1.5 g, 9.4 mmol) and triethylamine (2.0 g, 19.8 mmol) dissolved dichloromethane (dehydrated, 20 mL) was added to a solution of 4,4'-azodibenzoyl dichloride (1.0 g, 3.3 mmol) in dichloromethane (dehydrated, 30 mL). The mixture was stirred for 12 h at 25 °C. The precipitates was filtered, washed with dichloromethane, and dried *in vacuo*, to give **1** (1.7 g, 93%) as a yellow powder. δ_{H} (400 MHz; pyridine- d_5 ; Me $_4$ Si): 8.11–8.15 (8H, m, Ph), 8.21 (4H, m, Ph), 8.32 (2H, s, Ph), 8.45 (4H, m, Ph), 8.56 (4H, m, Ph), 11.4 (2H, s, NH). The obtained **1** (1.7 g, 3.1 mmol) and methyl trifluoromethanesulfonate (2.0 g, 12 mmol) in *N,N*-dimethylformamide (200 mL) was stirred for 48 h at 25 °C. The solution was added to acetone. Obtained precipitates were filtered and washed with acetone to give **AB** as a yellow powder (2.4 g, 89%). ν_{max} (KBr)/ cm^{-1} 3159 (ν_{CH} , Ar), 3113 (ν_{CH} , Ar), 1667 ($\nu_{\text{C=O}}$, amide), 1518 (σ_{NH}). λ_{max} (propylene carbonate)/nm 340 ($\epsilon/\text{dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$, 43 100). δ_{H} (400 MHz; DMSO- d_6 ; Me $_4$ Si): 3.95 (6H, s, NMe), 7.78 (4H, m, Ph), 7.94 (4H, m, Ph), 8.04–8.11 (8H, m, Ph), 8.22–8.27 (4H, m, Ph), 9.70 (2H, s, Ph), 10.7 (2H, s, NH). Found: C, 48.8; H, 3.4; N, 12.7. Calc. for C $_{36}$ H $_{30}$ F $_6$ N $_8$ O $_8$ S $_2$: C, 49.1; H, 3.4; N, 12.7%.

Preparation of SWCNT/AB and SWCNT/Sodium deoxycholate dispersions

SWCNT (1.0 mg) and **AB** (2.0 mg) were added to propylene carbonate (3 mL) and sonicated (80 W, 35 kHz) for 60 min at 25 °C using an ultrasonic bath sonicator (SHARP UT-105) followed by centrifugation (28500g, 180 min). The supernatant (80 vol%) was collected as a uniform dispersion of SWCNT/**AB**.

SWCNT (1.0 mg) and sodium deoxycholate (SDOC, 2.0 mg) were added to D $_2$ O (3 mL) and sonicated (80 W, 35 kHz) for 60 min at room temperature using an ultrasonic bath sonicator (SHARP UT-105) followed by centrifugation (28500g, 180 min). The supernatant (80 vol%) was collected as a uniform dispersion of SWCNT/SDOC.

Results and discussion

The hybrid film of SWCNT with **AB** was prepared by the following procedure. First, an azobenzene-derived cationic molecule with the trifluoromethanesulfonyl anion (**AB**, Scheme 1), a photoresponsive dispersant, was synthesised in two steps

with an overall yield of 85%. The synthesized **AB** exhibited good dispersibility of SWCNT in polar organic solvent as previously reported dispersants,^{26–28} and photoisomerisation between the *trans* to *cis* states by irradiation of UV or visible lights in propylene carbonate (PC) solution (Fig. S1†). Second, a dispersion of SWCNT/**AB** in PC was prepared by sonication and subsequent centrifugation (see the Experimental section for further details). The dispersion was then casted on a glass substrate and dried in air for 24 h, followed by drying *in vacuo* for 12 h. After the fabrication processes, the SWCNT/**AB** hybrid was obtained as a uniform thin film.

The UV-Vis-NIR spectra of the SWCNT/**AB** hybrid dispersion in PC and the hybrid thin film were measured, respectively. As shown in Fig. 1a, the SWCNT/**AB** hybrid dispersion showed sharp absorption bands at 1000–1600 nm. These absorption bands are derived from the semiconducting SWCNT (S_{11}).^{29,30} After UV irradiation (365 nm, 25 mW cm^{-2} , 2 min), the absorption spectrum showed a decrease of the absorption band around 340 nm (from 1.81 to 1.46), which corresponds to the π – π^* transition of *trans*-azobenzene, indicating the *trans*–*cis* photoisomerisation of **AB**. Additionally, a slight increase of the absorption bands of S_{11} was observed (Fig. 1a inset). By comparison, the SWCNT/**AB** hybrid film did not exhibit clear S_{11} absorption bands (Fig. 1b). Instead, by the UV light irradiation, the hybrid film showed a minimal decrease of the azobenzene absorption band (from 2.29 to 2.25 at 340 nm) with a drastic increase of the S_{11} bands (Fig. 1b inset). To reveal the effect of the photoresponsive dispersant, a non-photoactive SWCNT film was prepared using SDOC. Fig. S2† shows the absorption spectra of the SWCNT/SDOC dispersion and its solid film. As seen, no spectral changes were evident for the SWCNT/SDOC hybrids, even for long UV irradiation time (10 min). In addition, the present drastic difference in UV spectra between solid film and solution has never been observed in our previous studies on other photoresponsive dispersants. Therefore, we consider that the high molecular planarity of **AB** induces a strong interaction with the SWCNT, especially in the solid state, in which solvation and dynamic exchange of the dispersant are suppressed.

Based on these results, it appears that the spectral changes are caused by doping and de-doping effects of the

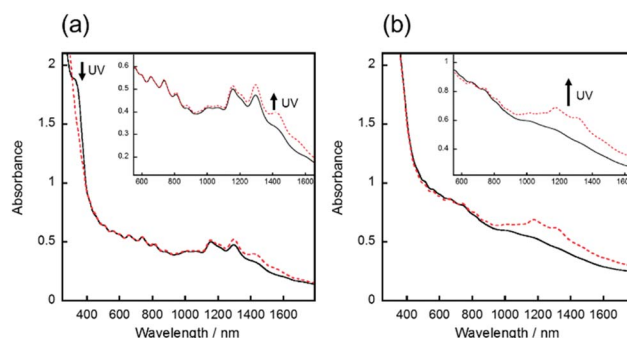


Fig. 1 (a) UV-Vis-NIR spectra of the SWCNT/**AB** hybrids in (a) propylene carbonate solution, and (b) solid film. The black solid and red dashes lines indicate the absorption spectra before and after UV irradiation (365 nm, 20 mW cm^{-2} , and 2 min), respectively.



photoisomerised **AB**. Indeed, it has been known that the absorption bands of S_{11} are bleached by doping with acids (e.g. H_2SO_4 and HNO_3).^{31–33} In our case, the absorption spectrum of the SWCNT/**AB** film indicates the bleached S_{11} bands compared with that of the SWCNT/SDOC film and the SWCNT/**AB** film with UV irradiation. Collectively, these results showed that the as-prepared SWCNT film was doped by **AB**, and the appearance of the S_{11} bands following UV irradiation should correspond the de-doping of photoisomerised **AB**. The photochemical change of the sheet resistance of the SWCNT/**AB** film supports this scheme.³³ After UV irradiation, the sheet resistance of the SWCNT/**AB** film increased 3 times (from 3.1×10^3 to $9.5 \times 10^3 \Omega \text{ sq}^{-1}$), suggesting the de-doping of SWCNT by the structural change of **AB** on the hybrid (Table S1†). These results indicate that the SWCNT/**AB** hybrid can output two signals, as evidenced by changes in NIR absorption and conductivity, following the input of one signal (UV light). Notably, almost all previously reported photoactive carbon nanomaterials comprise one-to-one relations between input and output.^{34,35}

Reversible changes of absorbance were observed following the irradiation of visible light or thermal treatment at 100 °C. Following the visible light irradiation (436 nm, 10 mW cm^{-2}) to the UV-irradiated SWCNT/**AB** film, the gradual decrease of the S_{11} absorption bands was observed (Fig. 2a). After 20 min of irradiation, the absorption spectra were recovered to that before UV irradiation. This spectral change demonstrated that the doping of the SWCNT by **AB** occurred again due to the reverse photoreaction of **AB** from the *cis* state to *trans* one.¹⁴ Over five cycles of switching between doping and de-doping of **AB** by alternating UV and visible irradiation, the absorbance change was basically consistent (Fig. 2b). Additionally, the switching of the sheet resistance of the SWCNT/**AB** film was also observed following the UV and visible light irradiations (Fig. S3†). When the UV-irradiated SWCNT/**AB** film was held at 100 °C, a decrease of the S_{11} absorption bands was evident (Fig. 3a). It has been known that azobenzene can interchange between the *cis* and *trans* states not only by visible light but also by thermal treatment.²³ Clearly, the presence of **AB** in the hybrid suggests a thermal reverse reaction, and the SWCNT/**AB** film indeed exhibited a decrease in the S_{11} absorption bands at high temperature. In contrast, the similar decrease of absorbance was not observed when the UV-

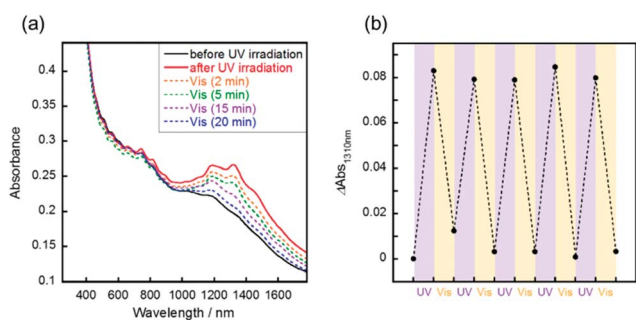


Fig. 2 (a) UV-Vis-NIR spectral change of the SWCNT/**AB** hybrid film with UV (365 nm, 20 mW cm^{-2} , 2 min) and visible light (436 nm, 10 mW cm^{-2} , 20 min) irradiation. (b) Absorbance change ratio at 1310 nm (ΔAbs_{1310}) of the SWCNT/**AB** hybrid film with UV and visible light irradiation.

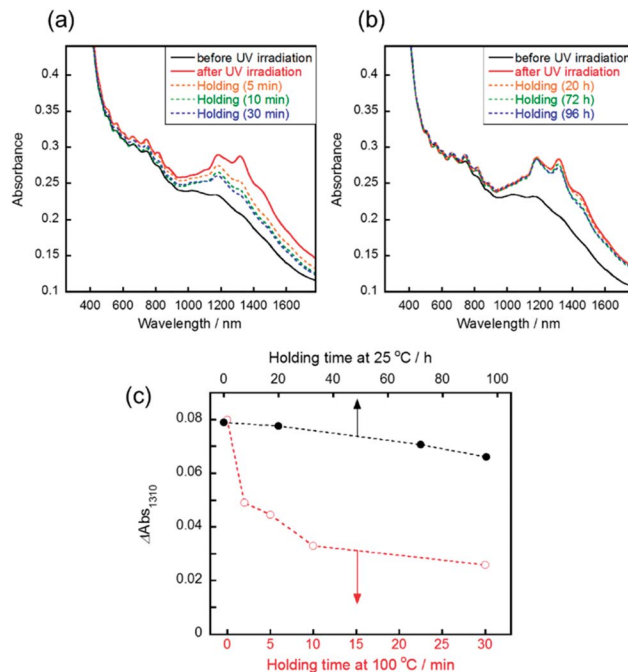


Fig. 3 (a) UV-Vis-NIR spectral change of the SWCNT/**AB** hybrid film with holding at (a) 100 °C and (b) 25 °C. (c) Time-dependent change of ΔAbs_{1310} at 25 °C (black solid circle) and (b) 25 °C (red open circle).

irradiated SWCNT/**AB** film was held at relatively low temperature of 25 °C (Fig. 3b and c). Held at 25 °C for 96 h, the SWCNT/**AB** film showed only a slight decrease of absorbance (the change of absorbance at 1310 nm was 0.012) due to the suppression of the thermal *cis* to *trans* back reaction.

To clarify the doping effect of **AB** on SWCNT in the SWCNT/**AB** solid film, Raman spectroscopy was also performed. Upon UV irradiation, the G band of the SWCNT in the SWCNT/**AB** film showed a blue-shift (from 1590 cm^{-1} to 1594 cm^{-1}) with respect to that of the unirradiated film (Fig. 4a). Following visible light

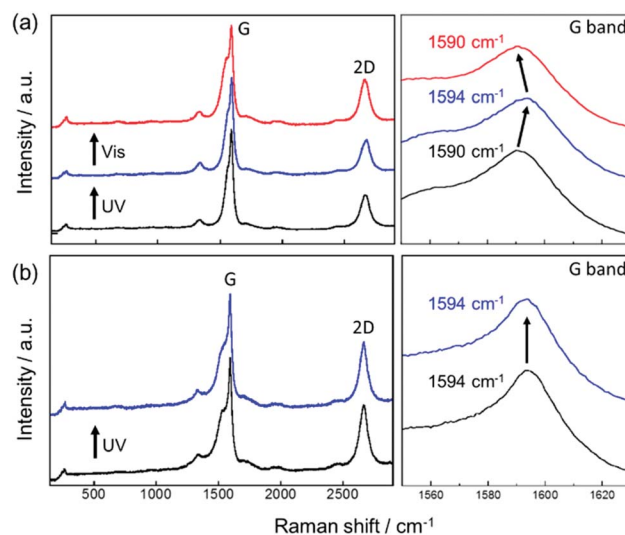


Fig. 4 Raman spectra of (a) SWCNT/**AB** and (b) SWCNT/SDOC films on glass substrate as prepared (black line), with UV irradiation (blue line), and with Vis irradiation (red line). Excitation wavelength is 532 nm.



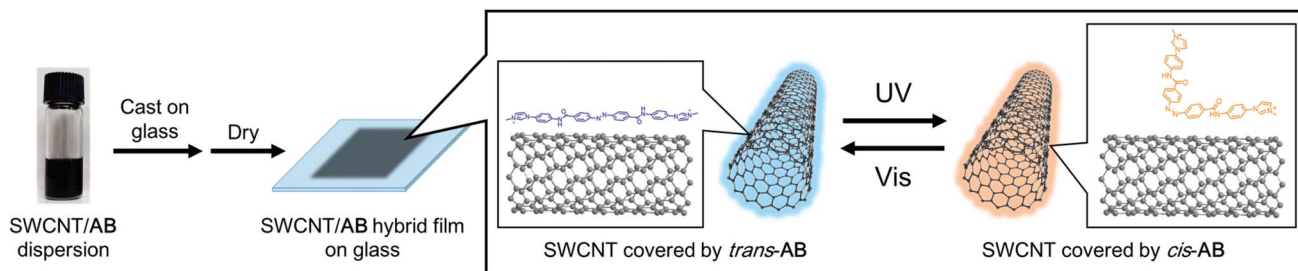


Fig. 5 Schematic representation of the light-induced switching of AB (only the cationic part is displayed) on the SWCNT surface in solid film.

irradiation, the G band was red-shifted to the original position (from 1594 cm^{-1} to 1590 cm^{-1}). As a control, the Raman spectrum of the SWCNT/SDOC film was also recorded. The G band of the SWCNT in this film was always observed at 1590 cm^{-1} with and without UV irradiation (Fig. 4b). On the basis of these studies, the observed shifts in the G band with UV and visible light for SWCNT/AB can be attributed to the photoisomerisation of the attached AB.¹⁴ Upon UV irradiation, AB converts from the *trans* form to the *cis* one leading doping and de-doping, respectively. These conclusions are supported by the I_{2D}/I_G ratio, as doping effects are apparent from a decrease of this ratio.³⁶ Table S1† summarises the I_{2D}/I_G ratio in several conditions. The I_{2D}/I_G of SWCNT/AB is 0.32, which is reduced relative to that of SWCNT/SDOC (0.53), indicating that the SWCNT was doped by AB. Upon UV irradiation, the I_{2D}/I_G increased to 0.36, indicating de-doping by the photo-isomerised AB.

The surface morphology of the SWCNT/AB film on glass substrate before and after UV irradiation was also examined using AFM (Fig. S4†). No clear differences were observed before or after UV irradiation, indicating that the switching behaviour occurred only as a molecular-scale photoreaction and not as a macroscopic-scale phenomenon.

Schematic images of the light-induced doping and de-doping of AB in the solid film are shown in Fig. 5. It is known that the energy level of SWCNTs are affected by the *trans*-to-*cis* transformation of nearly-located azobenzene.^{14,37} As in our previous work in a liquid condition,²³ we observed macroscopic changes of the SWCNT/AB hybrid dispersion from a uniform dispersive state to a precipitative one caused by the detachment of the photo-isomerised AB from the SWCNT surface into a solvent following UV light irradiation (Fig. S5†). Thus, absorption bands of the SWCNT were not observed in the solution state. By comparison, molecular movement is limited in the present solid films, so the SWCNT and AB can interact strongly and the photo-isomerisation cycle of AB is believed to occur on the SWCNT surface. That is the reason why the SWCNT/AB film showed drastic and reversible changes of optical and electrical properties.

It was found that the absorbance of the S_{11} bands could be controlled not only by irradiation time but also by the intensity of the UV light. We irradiated UV light to the SWCNT/AB film at different intensities and time (Fig. S6a and S6b†). Following irradiation with low intensity UV light (2 mW cm^{-2}), the rate of increase and the change of absorbance were lower than that

following irradiation with high intensity light (20 mW cm^{-2}) (Fig. S6c†). When we convert the irradiation time to irradiation energy (mJ cm^{-2}), a similar rate increase of absorbance is observed under different irradiation intensities, indicating that the increase of the rate of absorbance is strongly dependent on the total irradiated energy (Fig. S6d†).

Conclusions

In conclusion, we demonstrated the light-induced switching of both the optical and electrical properties of SWCNTs film by using the photoresponsive dispersant AB. The hybridisation of SWCNT and AB selectively reduces the absorbance of the semiconducting SWCNTs (S_{11}) because of the doping effect of AB. UV irradiation causes the photoisomerisation of AB, thereby reducing its doping ability and thus increasing of the absorbance of S_{11} . These absorbance changes were controllable by changing the intensity of the UV light and the irradiation time, and they were switchable by visible light irradiation and thermal treatment. Further, the hybrid film was fabricated using a simple wet-coating method, which is advantage in terms of cost, controllability of film properties, and can fabricate on various substrates. Overall, this SWCNT hybrid shows great potential for optoelectronic devices, sensors, and optical filters.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- 1 J. Du, S. Pei, L. Ma and H. M. Cheng, *Adv. Mater.*, 2014, **26**, 1958–1991.
- 2 J. W. Jo, J. W. Jung, J. U. Lee and W. H. Jo, *ACS Nano*, 2010, **4**, 5382–5388.
- 3 C. Cao, J. B. Andrews, A. Kumar and A. D. Franklin, *ACS Nano*, 2016, **10**, 5221–5229.



- 4 Y. Yang, L. Ding, J. Han, Z. Zhang and L.-M. Peng, *ACS Nano*, 2017, **11**, 4124–4132.
- 5 X. He, W. Gao, L. Xie, Q. Zhang, S. Lei, B. Li, J. M. Robinson, S. K. Doorn, R. Vajtai, P. M. Ajayan and W. W. Adams, *Nat. Nanotechnol.*, 2016, **11**, 633–640.
- 6 X. He, N. Fujimura, J. M. Lloyd, K. J. Erickson, A. A. Talin, Q. Zhang, W. Gao, Q. Jiang, Y. Kawano, R. H. Hauge, F. Léonard and J. Kono, *Nano Lett.*, 2014, **14**, 3953–3958.
- 7 T. Fukumaru, T. Fujigaya and N. Nakashima, *Macromolecules*, 2013, **46**, 4034–4040.
- 8 G. W. Huang, H. M. Xiao and S. Y. Fu, *J. Mater. Chem. C*, 2014, **2**, 2758–2764.
- 9 C. Subramaniam, A. Sekiguchi, T. Yamada, D. N. Futaba and K. Hata, *Nanoscale*, 2015, **8**, 3888–3894.
- 10 D. Benetti, K. T. Dembele, J. Benavides, H. Zhao, S. Cloutier, I. Concina, A. Vomiero and F. Rosei, *J. Mater. Chem. C*, 2016, **4**, 3555–3562.
- 11 Q. Wang, S.-F. Lee, C. Tang, J. C.-M. Yu, K. C.-F. Leung and K.-W. Wong, *RSC Adv.*, 2014, **4**, 18193–18197.
- 12 L. Zhou, Q. Liu, X. Lv, L. Gao, S. Fang and H. Yu, *J. Mater. Chem. C*, 2016, **4**, 9993–9997.
- 13 K. R. Devi, C. L. Lekshmi, K. Joseph and S. Mahesh, *ACS Appl. Mater. Interfaces*, 2016, **9**, 1167–1172.
- 14 S. Li, Y. Feng, P. Long, C. Qin and W. Feng, *J. Mater. Chem. C*, 2017, **5**, 5068–5075.
- 15 Y. Wang, H. Mi, Q. Zheng, Z. Ma and S. Gong, *ACS Appl. Mater. Interfaces*, 2015, **7**, 2641–2647.
- 16 W. Feng, C. Qin, Y. Shen, Y. Li, W. Luo, H. An and Y. Feng, *Sci. Rep.*, 2015, **4**, 3777.
- 17 J. Zhao and J. Ma, *J. Phys. Chem. C*, 2016, **120**, 25131–25141.
- 18 A. M. Kolpak and J. C. Grossman, *Nano Lett.*, 2011, **11**, 3156–3162.
- 19 A. M. Kolpak and J. C. Grossman, *J. Chem. Phys.*, 2013, **138**, 34303.
- 20 M. Kim, N. S. Safron, C. Huang, M. S. Arnold and P. Gopalan, *Nano Lett.*, 2012, **12**, 182–187.
- 21 A. R. Jang, E. K. Jeon, D. Kang, G. Kim, B. S. Kim, D. J. Kang and H. S. Shin, *ACS Nano*, 2012, **6**, 9207–9213.
- 22 Y. Matsuzawa, Y. Takada, T. Kodaira, H. Kihara, H. Kataura and M. Yoshida, *J. Phys. Chem. C*, 2014, **118**, 5013–5019.
- 23 H. Jintoku, Y. Matsuzawa, H. Kihara and M. Yoshida, *Chem. Lett.*, 2016, **45**, 1307–1309.
- 24 H. Jintoku, T. Sato, T. Nakazumi, Y. Matsuzawa, H. Kihara and M. Yoshida, *ACS Appl. Mater. Interfaces*, 2017, **9**, 30805–30811.
- 25 Y. Matsuzawa, H. Kato, H. Ohyama, D. Nishide, H. Kataura and M. Yoshida, *Adv. Mater.*, 2011, **23**, 3922–3925.
- 26 O. V. Kharissova, B. I. Kharisov and E. G. de Casas Ortiz, *RSC Adv.*, 2013, **3**, 24812.
- 27 T. Fujigaya and N. Nakashima, *Sci. Technol. Adv. Mater.*, 2015, **16**, 24802.
- 28 K. Mulla, S. Liang, H. Shaik, E. a. Younes, A. Adronov and Y. Zhao, *Chem. Commun.*, 2015, **51**, 149–152.
- 29 M. J. O'Connell, S. M. Bachilo, C. B. Huffman, K. L. Rialon, P. J. Boul and W. H. Noon, *Science*, 2002, **297**, 593–597.
- 30 Y. Kato, A. Inoue, Y. Niidome and N. Nakashima, *Sci. Rep.*, 2012, **2**, 733.
- 31 Y. Nonoguchi, K. Ohashi, R. Kanazawa, K. Ashiba, K. Hata, T. Nakagawa, C. Adachi, T. Tanase and T. Kawai, *Sci. Rep.*, 2013, **3**, 3344.
- 32 H. Wang, P. Wei, Y. Li, J. Han, H. R. Lee, B. D. Naab, N. Liu, C. Wang, E. Adijanto, B. C.-K. Tee, S. Morishita, Q. Li, Y. Gao, Y. Cui and Z. Bao, *Proc. Natl. Acad. Sci.*, 2014, **111**, 4776–4781.
- 33 J. L. Blackburn, T. M. Barnes, M. C. Beard, Y. Kim, C. Robert, T. J. McDonald, B. To, T. Coutts and M. J. Heben, *Renewable Energy*, 2008, **2**, 1266–1274.
- 34 W. Feng, W. Luo and Y. Feng, *Nanoscale*, 2012, **4**, 6118–6134.
- 35 X. Zhang, L. Hou and P. Samorì, *Nat. Commun.*, 2016, **7**, 11118.
- 36 A. C. Ferrari and D. M. Basko, *Nat. Nanotechnol.*, 2013, **8**, 235–246.
- 37 J. M. Simmons, I. In, V. E. Campbell, T. J. Mark, F. Léonard, P. Gopalan and M. A. Eriksson, *Phys. Rev. Lett.*, 2007, **98**, 86802.

