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Design and synthesis of graphene/SnO₂/polyacrylamide nanocomposites as anode material for lithium-ion batteries

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Tin dioxide (SnO₂) is a promising anode material for lithium-ion batteries owing to its large theoretical capacity (1494 mA h g⁻¹). However, its practical application is hindered by these problems: the low conductivity, which restricts rate performance of the electrode, and the drastic volume change (400%). In this study, we designed a novel polyacrylamide/SnO₂ nanocrystals/graphene gel (PAAm@SnO₂NC@GG) structure, in which SnO₂ nanocrystals anchored in three-dimensional graphene gel network and the polyacrylamide layers could effectively prevent the agglomeration of SnO₂ nanocrystals, presenting excellent cyclability and rate performance. A capacity retention of over 90% after 300 cycles of 376 mA h g⁻¹ was achieved at a current density of 5 A g⁻¹. In addition, a stable capacity of about 989 mA h g⁻¹ at lower current density of 0.2 A g⁻¹ was achieved.

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1 Introduction

There has been great interest in electric vehicles as a result of shortage of fossil fuels and air pollution. The development of electric vehicles requires higher performance of lithium ion batteries (LIBs).¹ Nevertheless, graphite does not satisfy this need owing to its low capacity (372 mA h g⁻¹).^{2–6} Tin dioxide (SnO₂) with an excellent theoretical capacity of 1494 mA h g⁻¹ is a promising material of LIBs anode and has drawn a lot of attention.^{7–14} However, the application of SnO₂ in the industries is restricted by two severe problems. First, the poor conductivity (resistivity 93 Ω m) leads to its low electronic conduction rate and poor multiplier performance. Second, the pulverization problem caused by drastic volume change (~300%) during lithium insertion and extraction directly leads to severe capacity fading in the cycling.^{15–26}

To solve these two problems, much effort has been made, resulting in many significant achievements. Anchoring the SnO₂ nanoparticles into carbon-based materials such as graphene or carbon nanotubes (CNTs) is a good way to enhance the conductivity and stability of SnO₂-based anode materials.^{19,24,27–37} In our former study, we successfully anchored

SnO₂ nanocrystals on 3D graphene to tackle the problems caused by the large volume change of SnO₂ during lithiation and delithiation processes.³⁸ Assisted by this structure, high specific capacity and good cycling performance were achieved. However, when the current density is over 5 A g⁻¹ or higher, a significant fall in the capacity occurs during the long cycling test. After careful characterization of the unpacked cell, severe agglomeration of the nanoparticles was found, which directly led to the decay of the capacity at a high current density. Thus, the design of the nanostructure composed of SnO₂ and graphene is not strong enough to alleviate the agglomeration of nanoparticles at a high rate.

There were some reports that the composite structure of polymer and graphene would improve the performance of SnO₂-based anode.^{39–41} Moreover, recently, Li *et al.* reported a cost-effective approach to prepare elastic graphene aerogels with the aid of polyacrylamide (PAAm) during the gelation of a graphene oxide aqueous solution.⁴² The graphene aerogels modified by PAAm exhibited significantly improved mechanical strength. Inspiringly, we propose an *in situ*-polymerization method to modify the composite anode with PAAm. In this study, the SnO₂ nanoparticles anchored on the surface of graphene were further covered by a layer of polymer to restrict the agglomeration of the SnO₂ nanoparticles. Assisted by this polymer layer, the SnO₂ nanoparticles were anchored on the surface of graphene, resulting in a considerable increase in the current density. A capacity retention of over 90% after 300 cycles was achieved at a current density of 5 A g⁻¹. Moreover, a stable capacity of about 989 mA h g⁻¹ at lower current density of 0.2 A g⁻¹ was achieved, indicating that the excellent electrochemical performance was still maintained at low current

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density. This structure directly allayed the agglomeration of SnO_2 nanoparticles during cycling, particularly at higher current density.

2 Experimental section

2.1 Synthesis of SnO_2 nanoparticles

First, 0.8 g $\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$ (Sinopharm Chemical Reagent Co., Ltd.) was dissolved in 40 mL deionized (DI) water and then transferred to a 50 mL poly(tetrafluoroethylene) (Teflon)-lined stainless steel autoclave to be heated at 120 °C for 28 h to obtain a white precipitate, which was further dispersed in 40 mL DI water with a concentration of about 15 mg mL^{-1} .

2.2 Synthesis of $\text{PAAm@SnO}_2\text{@GG}$

GO powder (synthesized by Hummer's method⁴³) was added into SnO_2 dispersions and sonicated for several minutes to obtain a concentration of about 5 mg mL^{-1} . Then, 37.5 mg acrylamide (AAM), 1.65 mg methylene-bis-acrylamide (MBAA, the crosslinker), 1.2 mg potassium peroxydisulfate (KPS, the initiator), and 30 mg ascorbic acid (the chemical reducing agent) were added into the $\text{SnO}_2\text{@GG}$ dispersions. The mixed dispersions were placed in an oil bath (90 °C) for 10 min and cooled to 70 °C for 12 h in order to partially reduce GO. Subsequently, the dispersions turned into a weak gel of the reduced GO and polymerized AAM (PAAm). The as-prepared $\text{PAAm@SnO}_2\text{@GG}$ were taken out with a tweezer and blotted with a filter paper to remove surface adsorbed water and subsequently freeze-dried.

2.3 Electrode fabrication

The working electrode material consists of 80 wt% $\text{PAAm@SnO}_2\text{@GG}$, 10 wt% super P, and 10 wt% poly(vinylidene fluoride) (PVDF), which was pasted on a rough Cu

foil. The loading mass of active material was about 0.9 mg cm^{-2} . The electrolyte used for all the tests was 1 M LiPF_6 in ethylene carbonate/diethyl carbonate/dimethyl carbonate (1 : 1 : 1 v/v/v, Shenzhen Kejing Star Technology Company). The electrochemical properties were evaluated by galvanostatic cycling of 2032 coin-type half cells with the $\text{PAAm@SnO}_2\text{@GG}$ as the working electrode and lithium foil as the counter/reference electrode. The charge and discharge measurements of the batteries were carried out on a LAND CT2001A system in the fixed voltage window between 0.05 V and 3 V vs. Li^+/Li at room temperature. Electrochemical impedance spectral measurements were determined on a PARSTAT 2273 advanced electrochemical system over the frequency range from 100 kHz to 100 mHz. Cyclic voltammetry was also performed on a PARSTAT 2273 advanced electrochemical system at a scan rate of 0.1 mV s^{-1} .

3 Results and discussion

As shown in Fig. 1a, we successfully synthesized $\text{PAAm@SnO}_2\text{-NC@GG}$ by thermochemical reduction and *in situ* polymerization. To further understand the structure of $\text{PAAm@SnO}_2\text{-NC@GG}$, different techniques such as X-ray diffraction (XRD), thermogravimetric analysis (TGA), Fourier transform infrared spectroscopy (IR), scanning electron microscopy (SEM), and transmission electron microscopy (TEM) were performed.

Fig. 2a shows the XRD patterns of $\text{PAAm@SnO}_2\text{-NC@GG}$, in which all the peaks are identical. It exhibits prominent peaks at 26°, 33° and 52°, which can be well indexed to the (110), (101) and (211) planes of tetragonal rutile SnO_2 , respectively (JCPDS #41-1445). In order to quantify the amount of SnO_2 in the hybrid, TGA was carried out in the air atmosphere. The samples were heated from 25 °C to 800 °C. Fig. 2b shows the TGA curve for the powders of $\text{PAAm@SnO}_2\text{-NC@GG}$. Clearly, SnO_2 remains

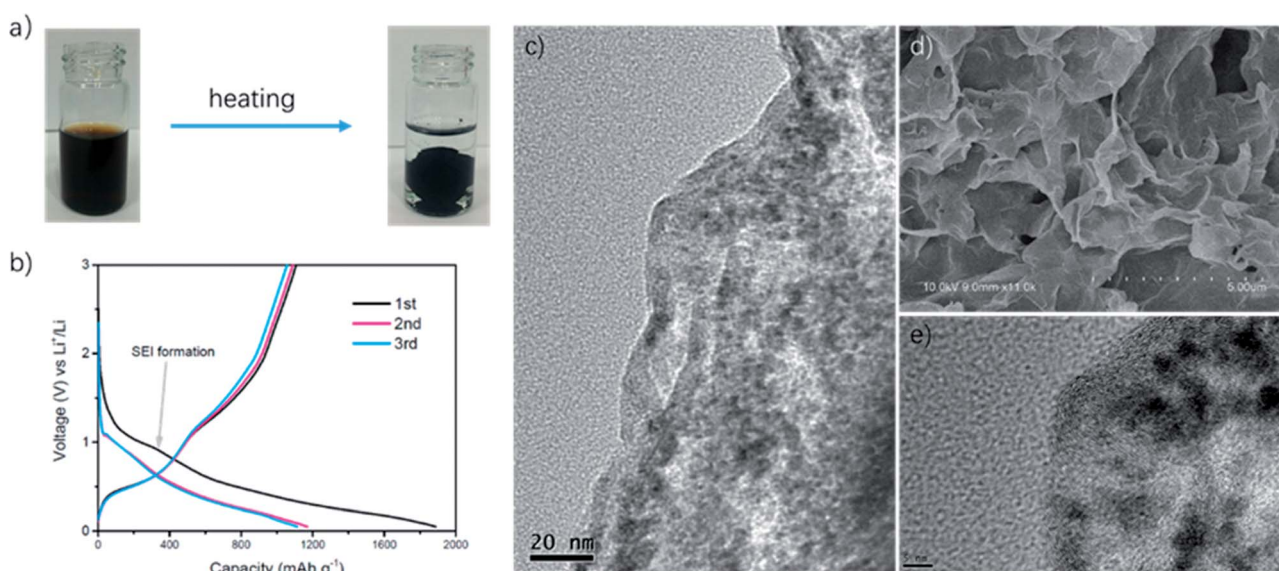


Fig. 1 (a) Synthesis of $\text{PAAm@SnO}_2\text{-NC@GG}$ (b) galvanostatic charge/discharge profiles plotted for the 1st, 2nd and 3rd cycles at a current density of 0.2 A g^{-1} (c) TEM image of $\text{PAAm@SnO}_2\text{-NC@GG}$ (d) SEM image of $\text{PAAm@SnO}_2\text{-NC@GG}$ (e) the magnification of (c).



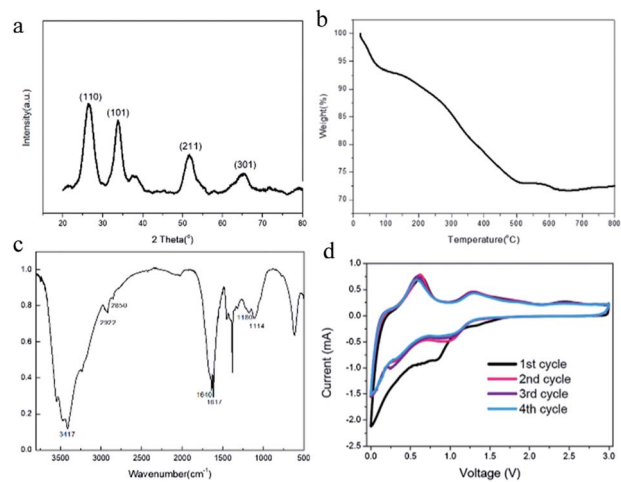


Fig. 2 (a) XRD patterns of PAAm@SnO₂NC@GG (b) TGA curve of PAAm@SnO₂NC@GG (c) IR spectrum of PAAm@SnO₂NC@GG (d) CV measurement of PAAm@SnO₂NC@GG based cell at a scan rate of 0.1 mV s⁻¹.

thermally stable, the content of which in PAAm@SnO₂NC@GG composite is about 70%. Fig. 2c shows the FT-IR spectrum of PAAm@SnO₂NC@GG, which shows the presence of various groups. In the figure, we can find the peak at 3417 cm⁻¹ attributed to -NH₂, the peaks at 2922 cm⁻¹ and 2850 cm⁻¹ assigned to the vibrations of the anti-symmetric and symmetric -CH₂- stretching mode, respectively, and the peaks at 1640 cm⁻¹ and 1617 cm⁻¹ attributed to the carbonyl group and the bending vibration of -N-H- on the acylamide, respectively. Furthermore, the remaining bands at 1180 cm⁻¹ and 1114 cm⁻¹ can be assigned to the stretching vibrations of the C-N bond. The IR spectrum indicates the existence of PAAm in the entire structure.

Fig. 1d shows the SEM images of PAAm@SnO₂NC@GG, which exhibits three-dimensional porous layer structure of densely stacked graphene sheets. Fig. 1c and e show the TEM images of PAAm@SnO₂NC@GG. In Fig. 1c, we can find SnO₂ nanocrystals well-loaded on the graphene aerogel. Fig. 1e shows that the diameter of SnO₂ nanocrystals is around 5 nm. All SEM and TEM images demonstrate that SnO₂ nanocrystals are well-dispersed in the 3D graphene gel.

Electrochemical measurements of PAAm@SnO₂NC@GG were carried out at room temperature in two-electrode 2031 coin-type half-cells. All the specific capacities and current densities are reported based on the weight of PAAm@SnO₂NC@GG hybrid. Cyclic voltammetry (CV) measurements were performed on half cells at a scan rate of 0.1 mV s⁻¹ versus Li/Li⁺ (Fig. 2d). As shown in Fig. 2d, the cathodic peaks at 0.83 V and close to 0 V in the first cycle could be assigned to the formation of Li₂O and Li_xSn alloy (Sn + xLi⁺ + xe⁻ = Li_xSn (0 ≤ x ≤ 4.4)), respectively. In addition, the broad peak between 0.5 V and 0.75 V can be ascribed to the decomposition of the electrolyte on the surface of the 3-D graphene in the PAAm@SnO₂NC@GG, leading to the formation of solid electrolyte interphase (SEI). During the charging process, the strong peak at 0.63 V and the

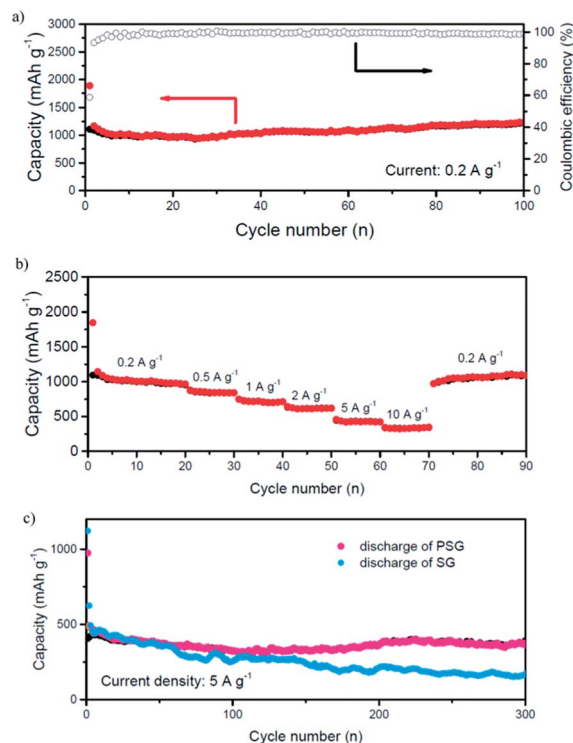


Fig. 3 (a) Electrochemical cycling performance of PAAm@SnO₂NC@GG electrode at a current density 0.2 A g⁻¹. (b) Capacity of the PAAm@SnO₂NC@GG electrode cycled at various current densities: 0.2 A g⁻¹, 0.5 A g⁻¹, 1 A g⁻¹, 2 A g⁻¹, 5 A g⁻¹ and 10 A g⁻¹ based on the total mass of the hybrid. (c) Electrochemical cycling performance of PAAm@SnO₂NC@GG (PSG) and SnO₂NC@GG (SG) electrode at a current density 5 A g⁻¹.

broad peak at 1.25 V correspond to the phase transition from Li_xSn alloy and Li₂O to SnO₂, respectively. The CV curves of the second, third and fourth cycle almost overlap, indicating an excellent cyclability of the PAAm@SnO₂NC@GG electrode.

Deep charge/discharge galvanostatic cycling from 3 to 0.05 V was performed to evaluate the electrochemical cycling performance of PAAm@SnO₂@GG composite electrodes as shown in Fig. 3a. The results indicated that the PAAm@SnO₂@GG composite exhibited excellent cycling stability. Normally, the capacity drop of SnO₂-based anode is due to the pulverization of the anode owing to the drastic volume change of SnO₂. However, through our composite anode, this problem was resolved due to the three aspects below:

(a) The 3-D graphene framework provides a tight buffer space for the expansion of SnO₂, which can avoid the breach by volume change.

(b) The SnO₂ nanocrystals are wrapped in the 3-D graphene, so most of the SEI is formed on the surface of graphene sheets and only a small amount of SEI would be unstable due to the huge volume change of SnO₂. This can reduce the capacity fade by the reconstruction of SEI.

(c) The 3-D graphene framework provides conductive pathways for the lithium ions and electrons.

Furthermore, the rate performance of PAAm@SnO₂NC@GG composite anode is also excellent. As shown in Fig. 3b, while the



current densities increase from 0.2 to 0.5, 1, 2, 5 and 10 A g⁻¹, the anode exhibits good capacity retention, varying from around 1000 to 750, 700, 600, 490, and 300 mA h g⁻¹, respectively. When the current density returns to 0.2 A g⁻¹, the charge capacity reverts to around 1000 mA h g⁻¹. The outstanding rate performance should benefit from the conductive 3-D graphene framework and the short diffusion path for the ions and electrons provided by SnO₂ nanocrystals constrained in it.

To reveal the influences of PAAM in the composite, we cycled different cells of SnO₂NC@GG without polymer and PAAM@SnO₂NC@GG at a high current density (5 A g⁻¹) for 300 times. As shown in Fig. 3c, we could easily find that the two electrodes had similar capacity at the beginning of the cycles. However, as the cycling process progressed, capacity fading was found in the SnO₂NC@GG electrode. The difference between the results obtained for the two electrodes indicates that the polymer plays an important role in the cycling performance.

In order to further understand how PAAM influences the cycling process, we removed the two batteries after 300 cycles at a current density of 5 A g⁻¹. The TEM images of the two removed electrodes are shown in Fig. 4. Fig. 4a and b show the TEM images of PAAM@SnO₂NC@GG electrode before after cycling, respectively. Fig. 4c and d show TEM images of SnO₂NC@GG electrode before and after cycling, respectively. The SnO₂ nanocrystals agglomerated in the SnO₂NC@GG electrode after 300 cycles (Fig. 4d). In contrast, no agglomeration was found in the PAAM@SnO₂NC@GG electrode after 300 cycles (Fig. 4b). Apparently, the capacity loss is because of the agglomeration of SnO₂ nanocrystals. To explain the influence of polymer in the composite anode, we put forward a mechanism as shown in Fig. 5. When the SnO₂NC@GG electrode was at a low current density, no agglomeration occurs. However, when the current density increases to 5 A g⁻¹, the SnO₂ nanocrystals agglomerate severely, resulting in a capacity loss. In the PAAM@SnO₂NC@GG electrode, the polymer layer can prevent the nanocrystals from agglomeration. Therefore, excellent stability was

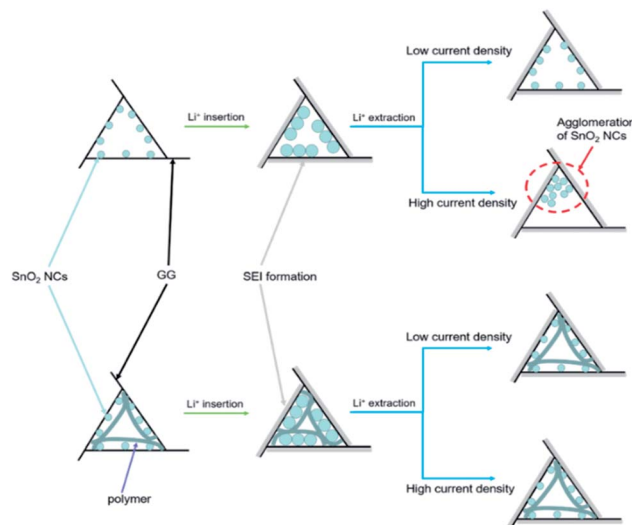


Fig. 5 Insertion and extraction of lithium ion for SnO₂NC@GG and PAAM@SnO₂NC@GG hybrid.

obtained at high current density with the help of PAAM. Furthermore, the PAAM will not influence the rate capability.

4 Conclusions

In conclusion, we developed a facile method to synthesize PAAM@SnO₂NC@GG composite electrode by thermochemical reduction and *in situ* polymerization. In this composite structure, the SnO₂ nanocrystals were enclosed in the 3D graphene aerogel framework and polymer layers, which provide empty space for accommodation of the drastic volume change of SnO₂. Additionally, the PAAM layer improved the performance at a high current density on the basis of excellent performance at lower current density by avoiding the agglomeration. The PAAM@SnO₂NC@GG composite electrode delivered an outstanding capacity retention of over 90% after 300 cycles was achieved at a current density of 5 A g⁻¹ and a stable capacity of about 989 mA h g⁻¹ at lower current density of 0.2 A g⁻¹. The novel graphene/SnO₂/polyacrylamide ternary nanocomposites provide us a new way to improve high rate capability for lithium-ion batteries.

Conflicts of interest

There are no conflicts to declare.

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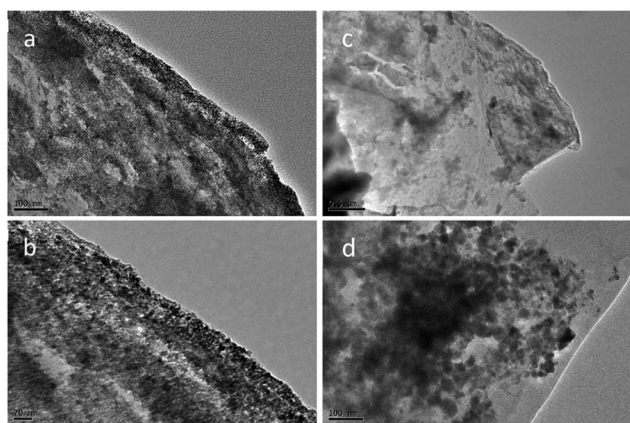


Fig. 4 (a) TEM images of PAAM@SnO₂NC@GG electrode before cycles. (b) TEM images of PAAM@SnO₂NC@GG electrode after 300 cycles at a current density of 5 A g⁻¹. (c) TEM images of SnO₂NC@GG electrode before cycles. (d) TEM images of SnO₂NC@GG electrode after 300 cycles at a current density of 5 A g⁻¹.



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