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Mechanisms of partial hydrogen sorption reversibility in a 3NaBH₄/ScF₃ composite

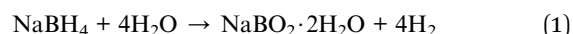
Ning Zhao,^a Jianxin Zou,^{id} *^{ab} Xiaoqin Zeng^{ab} and Wenjiang Ding^{ab}

A new hydrogen storage composite containing NaBH₄ and a 3d transition metal fluoride, 3NaBH₄/ScF₃, was synthesized *via* ball milling. The composite shows no reaction during milling and its dehydrogenating process can be divided into three steps upon heating: (i) partial substitution of H⁺ by F[−] in NaBH₄ to form NaBH_xF_{4−x} at the early stage, releasing about 0.19 wt% of hydrogen; (ii) formations of Na₃ScF₆, NaBF₄ and ScB₂ through the reaction between NaBH₄ and ScF₃, with 2.52 wt% of hydrogen release and a dehydrogenating activation energy of 162.67 kJ mol^{−1} H₂; (iii) further reaction of residual NaBH₄ and Na₃ScF₆ to form NaF, B and ScB₂, with a dehydrogenating activation energy of 169.37 kJ mol^{−1} H₂. The total hydrogen release of the composite reaches 5.54 wt% at 530 °C. The complete dehydrogenated composite cannot be rehydrogenated while the products after the second dehydrogenating step can be hydrogenated with an absorption activation energy of 44.58 kJ mol^{−1} H₂. These results demonstrate that by adding 3d transition metal fluorides into NaBH₄, a partial reversibility in NaBH₄ can be achieved.

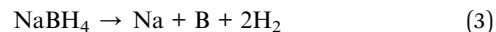
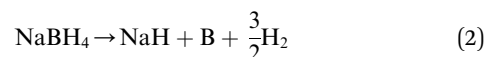
1 Introduction

Hydrogen is one of the most promising alternative and attractive clean energy sources that can substitute fossil fuels, with sufficient energy density and environment-friendliness.^{1,2} Nevertheless, almost a century since the concept of “hydrogen economy” was introduced by Jules Verne,^{3,4} it is still challenging to find reliable, flexible and cost-efficient hydrogen media for on-board, stationary and portable applications.^{5,6} In the past few decades, great attention has been paid to both hydrogen production technologies and a variety of hydrogen storage methods,^{7–10} including the use of different compounds,^{11,12} especially complex hydrides, of which borohydrides are typical representative.^{13–15} These solid-state hydrogen storage materials offer some advantages over high pressure gaseous storage and low temperature liquid storage, such as high capacity, high safety, and low cost. In particular, alkali metal borohydrides are regarded as possible hydrogen carriers, ascribed to their contribute to their high gravimetric and volumetric hydrogen density, together with good stability.¹⁶ However, current research results showed that few technologies regarding the use of metal borohydrides as hydrogen storage carriers are able to fulfill the requirements established by the US Department of Energy.¹⁷

Early investigations into borohydrides have shown that, compared to other borohydrides such as LiBH₄ and KBH₄, NaBH₄ is stable under alkaline conditions, and undergoes hydrolysis through the following reaction:¹⁸



However, the use of NaBH₄ as a hydrogen generator through hydrolysis faces several issues related to the catalyst durability, and/or poisoning, as well as storage vessels.^{19–21} In contrast, thermal decomposition of NaBH₄ has emerged as potential alternative method for hydrogen storage. J. Urgnani *et al.* investigated the thermal decomposition behaviors of NaBH₄, and proposed that NaBH₄ would decompose in two steps according to the following reactions:²²



For the sake of improving the thermodynamic and kinetic properties of the thermal decomposition of NaBH₄, many strategies have been taken, such as adding catalyst, building thermodynamic destabilization system, nano-engineering and chemical modifications.^{13,23–25} For instance, NaBH₄/MgH₂ system could fulfil its decomposition before melting due to the formation of MgB₂ in the system.²⁶ Czujiko *et al.* suggested that pure Mg could lower down the desorption temperature of NaBH₄ through catalytic effect.²⁷ In addition, the decomposition of NaBH₄ may occur at lower temperatures with some reversibility through the addition of Ni-based catalysts,²⁸ which

^aNational Engineering Research Center of Light Alloy Net Forming, State Key Laboratory of Metal Matrix Composites, Shanghai Jiao Tong University, Shanghai, 200240, P. R. China. E-mail: zoujx@sjtu.edu.cn; Fax: +86-21-34203730; Tel: +86-21-54742381

^bShanghai Engineering Research Center of Mg Materials and Applications, School of Materials Science and Engineering, Shanghai Jiao Tong University, Shanghai, 200240, P. R. China



facilitates hydrogen release and improves the reversibility to some extent.²⁹ Moreover, the chemical reaction that regenerates borohydrides from metal-borides occurs much easier over the regeneration from boron since less energy is required for breaking the chemical bond between B-M (M means metal) relative to the B-B's.³⁰

In previous works, many research works regarding hydrogen storage composite systems have been carried out, and some of them have explored how rare earth element (RE) addition can effect the thermodynamics and kinetics properties of metal borohydrides based systems, *i.e.*, NaBH₄-YF₃,³¹ NaBH₄-ScCl₃,³² LiBH₄-YCl₃,³³ *etc.* In particular, hydrogen sorption reversibility was achieved in 3NaBH₄/LnF₃ systems with good thermodynamic and kinetic properties.³⁴

Considering that scandium lies in the III B column of the periodic table as the lanthanide elements, we attempt to prepare a new hydrogen storage system, 3NaBH₄/ScF₃, through ball milling method. Our previous investigations on NaBH₄-MF₃ (M = metal) systems proved that the molar ratio of 3 : 1 was the best one, *i.e.* 3NaBH₄/LnF₃ (Ln = La, Ce, Nd, Gd, Yb),³⁴ 3NaBH₄/YF₃,³¹ 3NaBH₄/PrF₃ and 3NaBH₄/HoF₃.^{35,36} If the molar ratio of NaBH₄ to MF₃ is higher than 3 : 1, some NaBH₄ will be left after the desorption due to the incomplete reaction. While if the ratio is less than 3 : 1, MF₃ is excessive, and the overall hydrogen sorption capacity is reduced since MF₃ can neither release nor absorb hydrogen. Consequently, the ratio of NaBH₄ to ScF₃ is set as 3 : 1 in the present work. We conducted a detailed study of hydrogen sorption behaviors of the 3NaBH₄/ScF₃ system, and proposed mechanisms of hydrogen sorption in this composite, depending on experimental analyses.

2 Experimental

2.1 Sample preparation

NaBH₄ (Aladdin Reagent Database Inc., 96%) and ScF₃ (Aladdin Reagent Database Inc., 98%) were used as starting materials without further purification and mixed in the molar ratio of 3 : 1 in a planetary ball miller whose type is QM-1SP2. The stainless steel vessel with 100 ml volume was used to load 0.3928 g of NaBH₄ and 0.3530 g of ScF₃ powders together with 25 stainless steel balls (diameter of 5 mm, average weight of 0.8950 g each). The ball to powder weight ratio is approximately 30 : 1. Ball milling is conducted at a rotation speed of 400 rpm for 180 min. The prearrangement, manipulation and storage of specimen were carried out in an Ar-filled Lab 2000 glove box (Etelux inert gas system Co., Ltd.) in which neither moisture nor oxygen concentration beyonds 10 ppm.

2.2 Characterization

The analyses of phase composition for the ball milled, dehydrogenated and rehydrogenated samples were accomplished using an X-ray diffraction apparatus (D/max 2550VL/PC), equipped with a Cu-K α radiation source. For XRD tests, composite powder at different states were kept into specific sample holders with arched glass on both sides and airtight PVC tape on the top, to isolate them from air. Meanwhile, the

broad peak at around $2\theta = 15^\circ$ in the patterns is caused by the tape. Using a Spectrum Nicolet iS5 produced by Thermo Fisher Scientific Inc., Fourier transform infrared spectroscopy (FTIR) tests were performed on samples with different states in an Ar filled glove box. In principle of volumetric methods, we carried out temperature-programmed-desorption (TPD) measurements on 0.2 g of the 3NaBH₄/ScF₃ composite sample from room temperature to about 530 °C under an initial vacuum condition, with a heating rate R_H of 3 °C min⁻¹. Evaluations of hydrogen absorption performance were implemented for 10 h at various temperatures under around 3.2 MPa H₂ pressure.

Dehydrogenation behaviors of composites were determined by synchronous thermal analyzer (Differential Scanning Calorimetry/Thermal Gravimetry, DSC/TG), in a Netzsch, STA 449 F3 Jupiter equipment, in which $R_H = 3$ °C min⁻¹, 5 °C min⁻¹, 7 °C min⁻¹ and 10 °C min⁻¹, respectively, starting from room temperature to about 500 °C, under the protection of 0.1 MPa Ar flow. To compare the properties of hydrogen storage performances of different composites, data of weight percent considering hydrogen release and uptake during the tests was assessed based on the samples' initial weight value.

3 Results and discussions

3.1 Dehydrogenation of the 3NaBH₄/ScF₃ composite

The effect of the ScF₃ addition on the hydrogen desorption behaviors of NaBH₄ was examined by DSC measurements at different R_H , namely 3 °C min⁻¹, 5 °C min⁻¹, 7 °C min⁻¹ and 10 °C min⁻¹, as well as TG measurement at the R_H of 10 °C min⁻¹, as Fig. 1 showed.

Two major endothermic peaks upon heating appeared on the DSC curves, indicating that two main desorption steps took place during dehydrogenation. According to the starting point of dehydriding in DSC profiles, the first major dehydrogenation begins at 356 °C with $R_H = 3$ °C min⁻¹ heating rate condition. In contrast, pure NaBH₄ shows a dehydrogenation temperature of 517 °C under the same condition.³¹ In addition, a subsequent broad endothermic peak was also recorded. TG curve obtained at the heating rate of 10 °C min⁻¹ shows that the total mass loss reaches 4.10 wt% at 500 °C. According to the DSC/TG profiles, the dehydrogenation enthalpies of the first and second major desorption reactions are determined to be 27.43 ± 5 kJ mol⁻¹ H₂ and 29.54 ± 2 kJ mol⁻¹ H₂, respectively.

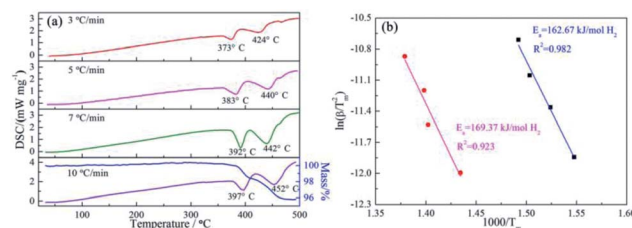


Fig. 1 DSC curves of 3NaBH₄/ScF₃ composite samples under heating rates of 3 °C min⁻¹, 5 °C min⁻¹, 7 °C min⁻¹ and 10 °C min⁻¹ and TG profile at the heating rate of 10 °C min⁻¹ (a) and the corresponding Kissinger plots for the two major endothermic desorption steps (b).



The apparent dehydrogenation activation energy (E_a) of the $3\text{NaBH}_4/\text{ScF}_3$ sample could be determined using the Kissinger method,³⁷ as described below:

$$\frac{d\left(\ln\frac{\beta}{T_m^2}\right)}{d\left(\frac{1}{T_m}\right)} = -\frac{E_a}{R} \quad (4)$$

where heating rate (β), peak temperature (T_m), and gas constant (R) show a specific relationship. Table 1 gives the peak temperatures in DSC curves at various R_H obtained from Fig. 1a. The fitting plot displays that $\ln(\beta/T_m^2)$ and $1/T_m$ have good linearity, as shown in Fig. 1b. According to eqn (4), the E_a value is calculated to be $162.67 \text{ kJ mol}^{-1} \text{ H}_2$ and $169.37 \text{ kJ mol}^{-1} \text{ H}_2$ for the first major desorption step and second major desorption step, respectively.

To obtain further information of dehydrogenation process of the target system, TPD measurement was performed on the ball-milled composite with a constant R_H of 3°C min^{-1} from ambient temperature to 530°C and the results are shown in Fig. 2. The results exhibit that $3\text{NaBH}_4/\text{ScF}_3$ composite has an appropriate onset dehydrogenation temperature, which is actually lower than 200°C in vacuum. Meanwhile, the figure shows that the desorption behavior may be subdivided into three consecutive processes, with a small amount of hydrogen ($\sim 0.19 \text{ wt\%}$) released at temperature lower than 310°C in the initial process, and the latter two steps range from 310°C to 420°C , and above 420°C , releasing about 2.52 wt\% and 2.83 wt\% of hydrogen, respectively. In comparison to the DSC measurements, the second and third desorption steps shown in TPD profile should correspond to the two major endothermic peaks on DSC curves.

The entire hydrogen release obtained from the experimental process up to 530°C is 5.54 wt\% , which is over 95% of the theoretical hydrogen content. Previous study shows that during dehydrogenation process of pure NaBH_4 under the same condition, only 0.68 wt\% weight loss is observed when heated up to 482°C .¹⁹ Thus, the hydrogen desorption properties of NaBH_4 were significantly promoted by the addition of ScF_3 . However, from Fig. 1, there is no endothermic peak present from ambient to 300°C in DSC curves. C. Bonatto Minella reported a related phenomenon and the difference between DSC analyses and volumetric measurements was ascribed to dissimilar experimental conditions.³⁸

XRD analyses were performed on samples treated under a series of controlled conditions in order to have a better understanding of mechanisms of de/rehydrogenation in the

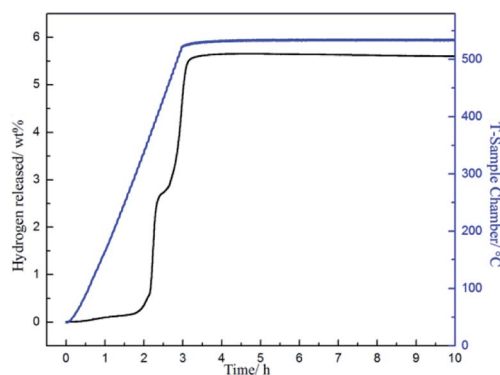


Fig. 2 TPD profile of the $3\text{NaBH}_4/\text{ScF}_3$ sample at a heating rate of 3°C min^{-1} .

$3\text{NaBH}_4/\text{ScF}_3$ composite, as shown in Fig. 3. The results show that no new product can be found in the sample after ball milling, except NaBH_4 (JCPDS no. 09-0386) and ScF_3 (JCPDS no. 46-1243), which means that only physical mixing takes place during milling process. Before dehydrogenation at various temperatures of 300°C , 420°C and 530°C , the loaded sample container was first placed under a 4.5 MPa H_2 pressure, then followed by a quick temperature rising and heat preservation. At last, the sample chamber was evacuated and powders were treated at the expected temperature for 3 h. After dehydrogenation at 300°C under vacuum, it can be clearly seen in Fig. 4 that diffraction peaks from NaBH_4 in the $3\text{NaBH}_4/\text{ScF}_3$ composite shifted slightly from high angle side to lower angle side, demonstrating a lattice expansion of NaBH_4 after the first dehydrogenation.

On the basis of the XRD patterns and using the RIR analysis, the lattice parameters of NaBH_4 in ball milled sample is calculated to be: $a = 0.6165 \text{ nm}$, $b = 0.6184 \text{ nm}$, $c = 0.6153 \text{ nm}$ and $\alpha = \beta = \gamma = 90^\circ$. After the first step dehydrogenation at 300°C , the lattice constants of NaBH_4 changed into $a = 0.6185 \text{ nm}$, $b = 0.6225 \text{ nm}$, $c = 0.6166 \text{ nm}$ and $\alpha = \beta = \gamma = 90^\circ$. Such a lattice expansion of NaBH_4 might be attributed to the fact that H^- was partially substituted by F^- in the unit cell. Similar phenomenon was also observed in some previous

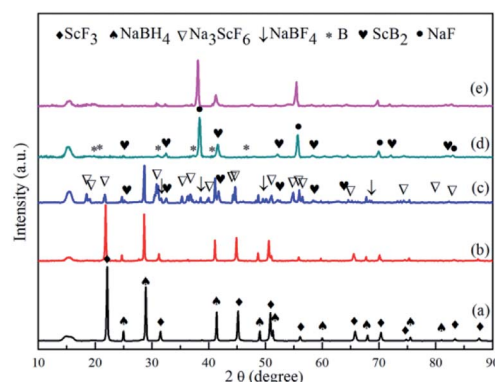


Fig. 3 XRD patterns of the $3\text{NaBH}_4/\text{ScF}_3$ composite after ball milling (a), dehydrogenated at 300°C (b), dehydrogenated at 420°C (c), fully dehydrogenated at 530°C (d) and rehydrogenated at 420°C (e).

Table 1 The DSC peak temperatures of the $3\text{NaBH}_4/\text{ScF}_3$ at different heating rates under 0.1 MPa argon atmosphere

Sample	Heating rate/ $^\circ\text{C min}^{-1}$	Temperature of peaks/ $^\circ\text{C}$	
$3\text{NaBH}_4/\text{ScF}_3$	3	373	424
	5	383	440
	7	392	442
	10	397	452



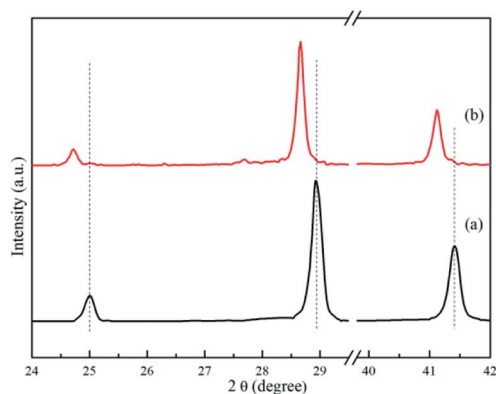
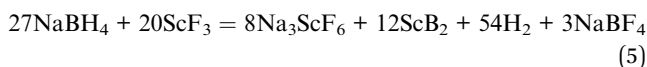
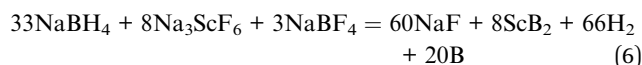


Fig. 4 XRD patterns of 3NaBH₄/ScF₃ composite after ball milling (a) and after dehydrogenation at 300 °C (b) in the range of low diffraction angle.

works,^{29,39–43} for which the lattice expansion was attributed to the formation of an intermediate compound NaBH_xF_{4–x}. In the present work, the formation of NaBH_xF_{4–x} occurred at the first desorption step between NaBH₄ and ScF₃, together with releasing small amount of hydrogen, as also seen in the 3NaBH₄/NdF₃, 3NaBH₄/PrF₃, 3NaBH₄/HoF₃ systems,^{19,35,36} and was regarded as an energy favorable process in theory.⁴⁴ Upon heating at temperatures higher than 310 °C, along with the reduction of NaBH₄ and disappearance of ScF₃ (Fig. 3c), Na₃ScF₆ appeared in the system, indicating that a major reaction between NaBH₄ and ScF₃ occurred. It has been also indicated by Radovan Cerny *et al.* that in the case of NaBH₄/ScCl₃ system, Na₃ScCl₆ and NaSc(BH₄)₄ formed as a result of the reaction between NaBH₄ and ScCl₃.³² Chong *et al.* reported that the reaction occurred at 250 °C between NaBH₄ and HoF₃ could produce NaHo(BH₄)₄ and NaHo₂F₇ phases.³⁶ In the 3NaBH₄/ScF₃ composite, NaSc(BH₄)₄ might form upon heating and then decomposed to ScB₂ and H₂. Based on the XRD analysis, the second dehydrogenation step before 420 °C can be described as:



Such a reaction has a theoretical hydrogen release of 2.53 wt%, close to the value measured from TPD for the second step dehydrogenation. At around 500 °C, the remaining NaBH₄ reacts with Na₃ScF₆ and NaBF₄ to produce NaF, ScB₂ and B, as shown in the indexed XRD pattern of Fig. 3d. Thus, the third dehydrogenation step can be described as follows:



with a theoretical hydrogen desorption value of 3.08 wt%. This value is also close to what is observed in the TPD analysis for the third dehydrogenation step. According to Garroni *et al.*,⁴⁵ Na₂[B₁₂H₁₂] is usually a byproduct during desorption of NaBH₄ based composites, which forms in an intermediate step and is still present at the end of reaction. Na₂[B₁₂H₁₂] is found to be

a stable byproduct and cannot be re-hydrogenated to NaBH₄, thus is regarded as an unfavorable product for the reversibility.^{46,47} However, the Na₂[B₁₂H₁₂] phase was not found in the XRD pattern of the partial or the complete dehydrogenated 3NaBH₄/ScF₃ samples, which means that very small amount or even no such a byproduct was generated during the decomposition of the 3NaBH₄/ScF₃ composite.

3.2 Rehydrogenation in the 3NaBH₄/ScF₃ composite

Fig. 3e shows the XRD pattern of the complete dehydrogenated 3NaBH₄/ScF₃ composite (530 °C for 3 h) that is maintained under a pressure of 3.2 MPa H₂ at 420 °C for 10 h. The pattern shows no change compared to that of the complete dehydrogenated composite, which means that the latter one has no hydrogen absorption ability.

As the complete dehydrogenated 3NaBH₄/ScF₃ composite shows no reversibility, the hydrogen absorption after the second dehydrogenation step was attempted to study the possible reversibility. The profiles of hydrogen absorption are given in Fig. 5a, which are obtained under the condition of 380 °C, 400 °C and 420 °C at 3.2 MPa H₂ pressure for the sample that has gone through dehydrogenation at 420 °C for 3 h. These curves clearly show the reversible hydrogen absorption of the partial dehydrogenated 3NaBH₄/ScF₃ composite: under the pressure of 3.2 MPa H₂, a hydrogenation capacity of 1.59 wt% can be achieved at 420 °C for 8 h, while it can absorb 1.28 wt% at 400 °C and 1.19 wt% at 380 °C, respectively. By contrast, under 3.5 MPa hydrogen pressure, pure NaBH₄ shows no hydrogen absorption at 400 °C.³⁶

The hydrogenation activation energy (*E*_{ab}) is generally utilized to discriminate kinetics of absorption, by analyzing the entire energy barriers of hydrogen absorption process. Based on the Johanson–Mehl–Avrami (JMA) model, the following equation can be used to evaluate absorption kinetics:⁴⁸

$$\ln[-\ln(1 - \alpha_A)] = \eta \ln k + \eta \ln t \quad (7)$$

where $\alpha(t)$ is a function of time *t*, *k* is a parameter describing kinetic, η is the Avrami exponent which matches transformation mechanism. Then, the following Arrhenius equation is used to calculate *E*_{ab}:

$$k = A \exp\left(-\frac{E_{ab}}{RT}\right) \quad (8)$$

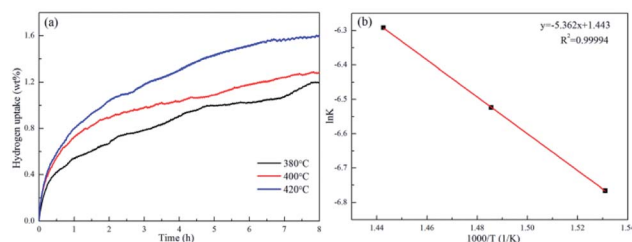


Fig. 5 Isothermal hydrogen absorption curves for the partially dehydrogenated 3NaBH₄/ScF₃ composite sample at temperatures of 380 °C, 400 °C and 420 °C (a) and the ln *k*–1000/*T* fitting plot (b).



where A represents temperature-independent constant, R represents universal gas constant, and T represents the absolute temperature. The scheme of $\ln k$ versus $1000/T$, which is shown in Fig. 5, displays a good linear relationship. Therefore, the E_{ab} value obtained from the slope is therefore estimated to be $44.58 \text{ kJ mol}^{-1} \text{ H}_2$ for the partially dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ composite.

To elucidate the mechanism of hydrogen absorption in the partially dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ composite, XRD analysis is carried out on the rehydrogenated sample and the result is shown in Fig. 6. The $3\text{NaBH}_4/\text{ScF}_3$ sample was dehydrided at 420°C for 3 h in vacuum and then rehydrogenated at 420°C for 10 h under the pressure of 3.2 MPa H_2 . In Fig. 6, the diffraction peaks from NaBF_4 and ScB_2 phases became weaker and even disappeared along with the increment in peak intensities of NaBH_4 and ScF_3 as compared to those in Fig. 3c. Therefore, the hydrogen absorption in the partial dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ composite follows exactly the reverse reaction path of the second step dehydrogenation. That is, the rehydrogenation consumes Na_3ScF_6 , NaBF_4 and ScB_2 , accompanied with the regeneration of NaBH_4 and ScF_3 . Compared to the completely dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ composite, the partially dehydrogenated composite contains Na_3ScF_6 and NaBF_4 phases, indicating that these two phases play the key role for the rehydrogenation.

The results of FTIR analyses for the ball-milled $3\text{NaBH}_4/\text{ScF}_3$ sample, sample dehydrogenated at 420°C for 3 h and corresponding products rehydrogenated at 400°C for 3 h can be found in Fig. 7. In Fig. 7a, the FTIR spectrum of sample after ball milling has the signatures of B–H stretching band in the position of 2226 cm^{-1} , 2306 cm^{-1} and 2366 cm^{-1} , and B–H bending band peak at 1119 cm^{-1} , all of which are supposed to be originated from borohydride. These peaks are considered to be from NaBH_4 .¹⁹ However, it should be noted that, the height of those peaks, which represent the intensity of B–H bonds vibration from the $[\text{BH}_4]^-$ group, gradually become weaker as dehydrogenation reaction proceeds, indicating the decomposition of NaBH_4 , as seen at 1121 cm^{-1} , 2221 cm^{-1} , 2338 cm^{-1} and 2369 cm^{-1} in Fig. 7b. According to the work of D. Syamala,⁴⁴ peak located at 1065 cm^{-1} can be marked as $[\text{BF}_4]^-$ asymmetric stretching, indicating the formation of NaBF_4 after the second step

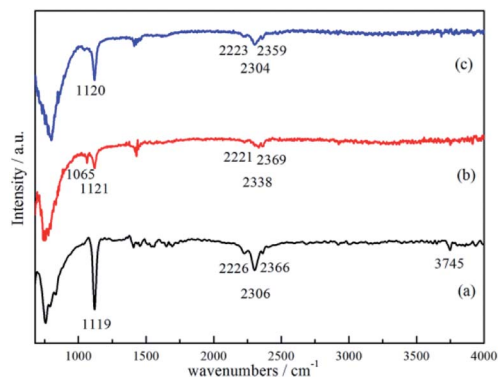


Fig. 7 FTIR spectra of the $3\text{NaBH}_4/\text{ScF}_3$ composite after ball-milling (a), half dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ sample (b), rehydrogenation of partially dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ sample (c).

dehydrogenation, which is in good agreement with the XRD results (Fig. 3c). In Fig. 7c, the signatures of $[\text{BH}_4]^-$ bending at 1120 cm^{-1} and $[\text{BH}_4]^-$ stretching at 2223 cm^{-1} , 2304 cm^{-1} and 2359 cm^{-1} were clearly revealed for the rehydrogenated sample and the intensity of these peaks increased, indicating the regeneration of NaBH_4 .⁴⁹ Meantime, the peak from $[\text{BF}_4]^-$ asymmetric stretching disappeared after rehydrogenation, showing the consumption of NaBF_4 along with rehydrogenation. Peaks at wave numbers between 1330 cm^{-1} and 1800 cm^{-1} (Fig. 7a) were subtracted from the unavoidable moisture absorption and atmospheric humidity absorbed by the sample during measurement, while peak located at around 3745 cm^{-1} was identified as stretching band vibration of O–H.⁵⁰

3.3 Mechanisms of hydrogenation in $3\text{NaBH}_4/\text{ScF}_3$ composite

It is shown in the present work that the hydrogen storage performance of NaBH_4 can be effectively improved by introducing ScF_3 as a reagent. In particular, rehydrogenation can be achieved in the partially dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ composite. Analyses revealed that both Sc^{3+} cation and F^- anion show irreplaceable importance during the re/dehydrogenation processes of the composite. Firstly, F anion can replace H anion in the initial process of dehydrogenation, from NaBH_4 to form $\text{NaBH}_x\text{F}_{4-x}$. Then, Sc cation loses electron to form ScB_2 , in which Sc cation has the calculated valence of $+4.08$,⁵¹ rather than served as a three-valent cation. This is accompanied with the formation of hydrogen gas. Meanwhile, during the second dehydrogenation step, a portion of F anions from ScF_3 incorporate into Na_3ScF_6 crystallites, which might serve as the nucleation center for the growth of other products.

It has been established that the regeneration of NaBH_4 in the $\text{NaBH}_4\text{--MF}_x$ systems is associated with electronegativity (χ_p) of the metal cations.⁵² Previous works have shown that after adding transition metal fluorides into NaBH_4 based composites, when the Pauling's electronegativity of the transition metal lies in around between 1.23–1.54, hydrogen sorption reversibility has larger thermodynamic tendency to occur.⁵² The χ_p value of Sc^{3+} is 1.415, which lies in such specific range, thus the

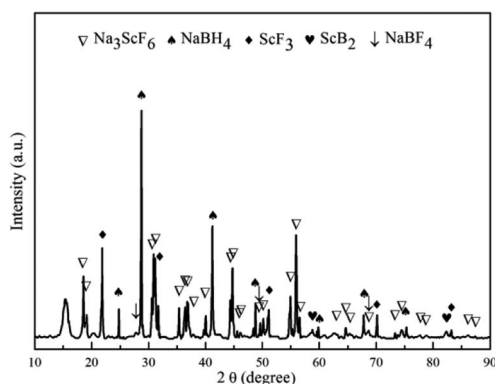


Fig. 6 XRD pattern of the partially dehydrogenated $3\text{NaBH}_4/\text{ScF}_3$ composite sample after rehydrogenation.



regeneration of NaBH_4 in the dehydrided $3\text{NaBH}_4/\text{ScF}_3$ system is favorable.⁵³

Using a database of density functional theory,^{54–56} the enthalpies of desorption reactions are calculated to be $41.01 \text{ kJ mol}^{-1} \text{ H}_2$ for the second dehydriding step, and $43.31 \text{ kJ mol}^{-1} \text{ H}_2$ for the final step. These values are comparably higher than the values obtained from DSC/TG analyses, but significant lower than that of pristine NaBH_4 ($108 \pm 3 \text{ kJ mol}^{-1}$ of H_2).⁵⁷ The differences between the calculated and measured enthalpies can be explained by fact that H^- was partially substituted by F^- in NaBH_4 , which is also observed in other NaBH_4 based systems containing fluorides.⁵⁸ However, the enthalpy for complete dehydrogenation is still fairly high, about $56.97 \text{ kJ mol}^{-1} \text{ H}_2$ calculated from DSC analyses. Consequently, the rehydrogenation of the complete dehydrided $3\text{NaBH}_4/\text{ScF}_3$ composite is difficult from the thermodynamic point of view.⁵²

During the rehydrogenation process, the experimental results show that only partial dehydrogenated products ($\text{NaBF}_4 + \text{ScB}_2 + \text{Na}_3\text{ScF}_6$) have reversibility, while the final dehydriding products ($\text{NaF} + \text{ScB}_2 + \text{B}$) cannot be rehydrogenated. Apart from the thermodynamic factors, this might also be understood from structural similarity between $[\text{BF}_4]^-$ in NaBF_4 and $[\text{BH}_4]^-$ in NaBH_4 , which may facilitate the regeneration of NaBH_4 through the exchange between H^- and F^- . The structural similarity was also observed between dehydrogenated and rehydrogenated products in the $3\text{NaBH}_4/\text{LnF}_3$ systems, which led to the improved rehydrogenation kinetics in these systems.²⁵ Researchers have also found that substitution reaction could be well understood from the hydride–fluoride isostructure, which has been proposed and confirmed in various hydrides–fluorides compounds having different stoichiometries.^{59–61}

4 Conclusions

In this study, the $3\text{NaBH}_4/\text{ScF}_3$ composite was prepared through mechanical milling. The behaviors and mechanisms of hydrogen de/absorption of the composite were explained by using TPD, DSC/TG, XRD and FTIR techniques. Following are the summarized results:

(1) TPD and DSC analyses confirmed that NaBH_{x-4-x} compound formed at the early dehydriding stage due to the partial substitution of H^- anion by F^- anion in NaBH_4 , releasing about 0.19 wt% of hydrogen. When temperature further increases, Na_3ScF_6 , NaBF_4 and ScB_2 formed through the reaction between NaBH_4 and ScF_3 with 2.52 wt% of hydrogen released. Finally, the reaction of residual NaBH_4 with Na_3ScF_6 produces NaF , B and ScB_2 , releasing about 2.83 wt% of hydrogen.

(2) The partially dehydrogenated products, Na_3ScF_6 , ScB_2 and NaBF_4 , can be rehydrogenated to generate NaBH_4 with an activation energy of $44.58 \text{ kJ mol}^{-1} \text{ H}_2$. In contrast, the fully dehydrogenated products, $\text{NaF} + \text{ScB}_2 + \text{B}$, cannot be hydrogennated. The hydrogen sorption reversibility of the partially dehydrogenated composite can be understood through thermodynamic point of view and the structural similarity between $[\text{BF}_4]^-$ in NaBF_4 and $[\text{BH}_4]^-$ in NaBH_4 .

Conflicts of interest

There are no conflicts to declare.

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