

Cite this: *RSC Adv.*, 2018, 8, 8699

Effects of polymers on the properties of hydrogels constructed using sodium deoxycholate and amino acid†

Yi Guo, Ruijin Wang, Yazhuo Shang * and Honglai Liu*

The gelation behavior and properties of sodium deoxycholate (NaDC) and L-aspartic acid (Asp) in aqueous solution were investigated in detail at 25 °C. The linear polymer poly(2-(2-methoxyethoxy)ethyl methacrylate-co-oligo-(ethylene glycol) methacrylate) (P(MEO₂MA₉₀-co-OEGMA₁₀)) and star-shaped polymer poly(2-(dimethylamino)ethyl methacrylate-*b*-2-(2-methoxyethoxy)ethyl methacrylate) (CDPDPM) were introduced in NaDC/Asp hydrogels for exploring the effects of polymers on the properties of NaDC/Asp hydrogels and the mechanism underlying gelation processes by polymers was proposed. The hydrogels were characterized by phase behavior observation, polarized optical microscopy (POM), cryogenic scanning electron microscopy (cryo-SEM), X-ray powder diffraction (XRD), Fourier transform infrared (FT-IR) spectroscopy and rheological measurements. Moreover, the adsorption performances of hydrogels with and without polymers to methylene blue (MB) were studied using a UV-vis spectrometer. The results indicated that the transition from sol to gel state was observed with an increase in the Asp concentration in the system. Both linear and star-shaped polymers can participate in the formation of a gel network structure, so that the density of network structure and the mechanical strength of hydrogels increased. Furthermore, it was found that the viscoelasticity of the CDPDPM-containing hydrogel was much higher than that of the P(MEO₂MA₉₀-co-OEGMA₁₀)-containing hydrogel under the same condition, indicating that CDPDPM performed better in strengthening the network structure of the hydrogels than P(MEO₂MA₉₀-co-OEGMA₁₀) due to the special structure that provided more binding sites for hydrogen bonding and stronger hydrophobicity that inhibited the swelling and dissolution of hydrogels. On coming in contact with the MB solution, the CDPDPM-containing hydrogel can adsorb MB and maintain the hydrogel state for recycling. On the contrary, the NaDC/Asp hydrogel dissolved and P(MEO₂MA₉₀-co-OEGMA₁₀)-containing hydrogel collapsed in the MB solution. The properties of the hydrogels are expected to be tailored by introducing polymers with different properties, including the charge numbers, the number of available binding sites, and hydrophobic properties.

Received 7th January 2018
Accepted 9th February 2018

DOI: 10.1039/c8ra00171e

rsc.li/rsc-advances

1. Introduction

Supramolecular hydrogels, as an important class of soft matter, have received considerable attention in recent years because of their remarkable potentials in fields of biomaterials,¹ drug delivery,^{2,3} sensing,^{4,5} tissue engineering,⁶ separation science,⁷ *etc.* Generally, the supramolecular hydrogels are derived from the self-assembly of low-molecular-weight gelator *via* various noncovalent interactions, such as hydrophobic interactions, hydrogen bonding, van der Waals forces, and electrostatic

interactions.^{8–10} The gelators are usually based on small organic molecules.¹¹ Among these gelators, the biocompatible and biodegradable species are particularly popular for constructing hydrogels, such as nucleobases,¹² bile salts,¹³ carbohydrates¹⁴ and peptides,¹⁵ as they can be safely used in biological applications.¹⁶

Bile acids and their salts belong to not only important biological surfactants, but also small-molecule biological gelators that can act as solubilizers of cholesterol, bilirubin, and various fat-soluble vitamins in the intestine of vertebrates, including humans.^{17,18} Unlike the typical surfactant that carries a polar head and a nonpolar tail, the bile salt is well known as a facial amphiphile and possesses a rigid steroid backbone comprising hydroxyl groups and a carboxyl groups on the polar face and methyl groups on the nonpolar face. Due to the special structure, bile salts can self-assemble in aqueous solutions to form special nanostructures including nanofibers and nanohelices

Key Laboratory for Advanced Materials, School of Chemistry & Molecular Engineering, East China University of Science and Technology, Shanghai 200237, China. E-mail: shangyazhuo@ecust.edu.cn; hlliu@ecust.edu.cn; Tel: +86-21-64252767; +86-21-64252921

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c8ra00171e

that have the potential to further cross-link to form gels.^{19,20} Sodium deoxycholate (NaDC), as one of the most studied bile salt hydrogelators, is sensitive to pH and yields hydrogels at slightly alkaline pH.²¹ However, the viscoelastic and mechanical properties of NaDC hydrogels are not strong enough. Thus, some additives are needed to construct composite hydrogels and improve their function. Numerous studies have been performed on the effects of salts, pH, metal ions, and amino acids on the properties of NaDC hydrogels.^{20,22,23} The existing results explicitly elucidate how these additives essentially affect the properties of NaDC hydrogels. Jover *et al.* prepared NaDC hydrogels in a phosphate buffer solution (PBS) and reported that the viscoelasticity of the hydrogel from NaDC in PBS increased drastically with an increase in NaCl concentration. They found that decreasing the pH or increasing the ionic strength of aqueous solutions of NaDC led to the protonation of the carboxylate group, and consequently strengthened the hydrogen bonding and weakened the electrostatic repulsion between the polar faces of the NaDC molecules. The intermolecular hydrogen bonds between carboxylic acid groups may result in gelation.²⁴ Zhang *et al.* prepared supramolecular hydrogels in mixtures of NaDC and amino acids (glycine (Gly), alanine (Ala), lysine (Lys) and arginine (Arg)) in different pH buffer solutions. The results have shown that the introduction of Ala and Gly strengthens the NaDC hydrogels due to its suitable size and hydrophilicity, which endows them the probability of creating more linkages between the NaDC molecules, while the addition of Lys and Arg could break the hydrogen bonds and weaken the formation of gels.²⁵ Wang *et al.* designed a type of luminescent hydrogels through supramolecular self-assembly of biological surfactants (NaDC) and lanthanide salt (europium nitrate (Eu(NO₃)₃)) and reported that the mechanical strength of the hydrogels increased with an increase in the concentration of NaDC and Eu(NO₃)₃.²⁶ Clearly, hydrogen bonds are the foundation of forming hydrogels and the properties of hydrogels are primarily determined by the amount of hydrogen bonds formed in studied systems.

Polymer is a chemical compound composed essentially of a large number of repeating structural units and has been widely used in numerous fields such as medicine, food, and electronics.^{27–29} The adjustability of structure, chemical composition as well as polymerization degree of polymers endows it potential to be used in the gel forming system. On the one hand, the number of hydrogen bonds formed in the system can be altered by introducing special functional groups or changing the number of certain repeating units in the polymer. On the other hand, the mode of hydrogen bond arrangement can be tailored by adjusting the molecular configuration of polymers. Polymers should have significant effects on the properties of hydrogels. The organic fillers, such as graphene oxide (GO), are used in the NaDC hydrogel systems. Wang *et al.* found that GO could effectively disperse in NaDC/PBS hydrogels, which enhanced the mechanical strength of the gels due to GO that could provide more binding sites of hydrogen bonds.³⁰ Clearly, the properties of hydrogels are expected to be optimized by introducing the corresponding polymers. However, the intrinsic function and the mechanism of action for polymers in gelation

are less involved and the understanding of the role of polymers in gelation and the corresponding mechanism of action is neither comprehensive nor systematic. Hence, further studies are still needed.

L-Aspartic acid (Asp) as a common amino acid can be used as an ammoniac detoxicating agent, hepar function accelerator, and fatigue refresher in pharmaceutical agents; also, it is widely utilized for mineral supplementation in the form of its potassium, magnesium or calcium salts.^{31,32} In this study, we combined the biological surfactant NaDC and amino acid Asp together to construct a gel system and investigated the corresponding gelation behavior in detail. The linear polymer poly(2-(2-methoxyethoxy)ethyl methacrylate-*co*-oligo(ethylene glycol) methacrylate) (P(MEO₂MA₉₀-*co*-OEGMA₁₀)) and star-shaped polymer poly(2-(dimethylamino)ethyl methacrylate-*b*-2-(2-methoxyethoxy)ethyl methacrylate) (CDPDPM) were introduced into the NaDC/Asp hydrogels for exploring the effects of polymers on the properties of NaDC/Asp hydrogels systematically and the mechanism of polymer in gelation processes was proposed. The properties of the hydrogels were characterized by a variety of techniques, including polarized optical microscopy (POM), cryogenic scanning electron microscopy (cryo-SEM), X-ray powder diffraction (XRD), Fourier transform infrared (FT-IR) spectroscopy, and rheological measurements. Furthermore, the adsorption performances of hydrogels with and without polymers to methylene blue (MB) were studied using a UV-vis spectrometer. Our results may provide useful information not only for the adjustment of hydrogel properties on purpose, but also for the application of hydrogel in practice.

2. Materials and methods

2.1 Materials

Sodium deoxycholate (NaDC) and L-aspartic acid (Asp) were purchased from J&K Scientific with 98% purity and used without further purification. Poly(2-(2-methoxyethoxy)ethyl methacrylate-*co*-oligo(ethylene glycol) methacrylate) (P(MEO₂MA₉₀-*co*-OEGMA₁₀), $M_n = 2.20 \times 10^4$ g mol⁻¹, PDI = 1.63) and poly(2-(dimethylamino)ethyl methacrylate-*b*-2-(2-methoxyethoxy)ethyl methacrylate) (CD-*g*-P(DMAEMA)₁₁₄-*b*-P(MEO₂MA)₁₁₅, CDPDPM, $M_n = 4.23 \times 10^4$ g mol⁻¹, PDI = 1.37) were synthesized *via* atom transfer radical polymerization (ATRP) according to our previous study (Fig. 1).^{33,34} Ultrapure water (with a resistivity of 18.2 MΩ cm) was used throughout the experiments.

2.2 Sample preparation

The 1 g L⁻¹ solutions of the polymers were prepared first. Later, the stock solutions of NaDC and Asp were prepared by dissolving appropriate amounts of them in ultrapure water or polymer solutions. The solutions were mildly stirred at room temperature until all solids were dissolved. A series of sample solutions with desired concentrations of each component were prepared by mixing different amounts of stock solutions to a final volume of 5 mL. All samples were kept in a thermostatic incubator at 25.0 ± 0.1 °C for at least 4 weeks until the



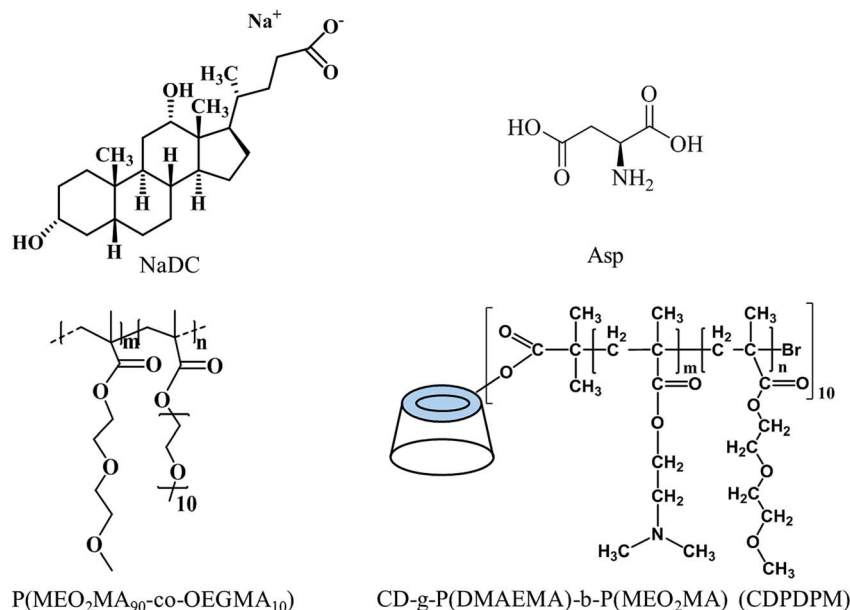


Fig. 1 Chemical structures of NaDC, Asp, P(MEO₂MA₉₀-co-OEGMA₁₀) and CDPDPM.

equilibrium state was attained and then, the phase behavior was inspected.

2.3 Identification and observation of birefringence

Photographs of sample birefringence were recorded using a Nikon Eclipse 50i pol polarizing optical microscope (POM) equipped with the Instec HCS302 heating/cooling stage.

2.4 Cryogenic scanning electron microscopy (cryo-SEM)

The microstructures of hydrosols and hydrogels were observed using cryo-SEM. The sample (5 μ L) was loaded on the cryo-specimen holder and then transferred into a liquid nitrogen bath until liquid nitrogen ceased boiling. Then, the liquid nitrogen bath was transferred into the vacuum space for vacuuming immediately. After the liquid nitrogen ceased boiling, the specimen holder was transferred into the sample preparation cryo-cavity for sublimation (from -180 $^{\circ}$ C to -90 $^{\circ}$ C) for 4 min and then cooled to -120 $^{\circ}$ C for coating. The exposed surface topography was coated with a conducting deposit of gold within 1 min. The coated sample was then transferred at -120 $^{\circ}$ C into the lens of a field-emission scanning electron microscope (Hitachi S-4800, Japan) equipped with a cryo-stage (Gatan Alto-2500, UK) and observed under an accelerating voltage of 5 kV at a working distance of 19 mm.

2.5 Rheological measurements

The rheological measurements were carried out on the TA ARES rheometer equipped with a conical-cylinder geometry (cup diameter: 30 mm, bob diameter: 27 mm, bob length: 49 mm, gap: 1 mm). In oscillatory measurements, a dynamic strain sweep experiment at a fixed angular frequency of 6.28 rad s^{-1} was carried out prior to the subsequent frequency sweep in

order to ensure that the selected strain was in a linear viscoelastic region. The viscoelastic properties of the samples were determined by oscillatory measurements in the angular frequency range of 0.1 – 100 rad s^{-1} . The samples were measured at 25.0 ± 0.1 $^{\circ}$ C with a cyclic water bath. The samples remained in the conical-cylinder geometry for 10 min before measurement.

2.6 X-ray powder diffraction (XRD) measurements

The XRD patterns of the freeze-dried samples were recorded on the D/MAX-2550 VB/PC rotating anode X-ray powder diffractometer (Rigaku Corporation, Japan) with Cu K α radiation ($\lambda = 0.15418$ nm) and a graphite monochromator. The samples were examined at room temperature over 1 – 10° in the 2θ mode (1° min^{-1}).

2.7 Infrared spectroscopy

Fourier transform infrared (FT-IR) spectra were recorded using the Thermo Fisher Nicolet 6700 FT-IR spectrometer. The samples were prepared by KBr disc or film technique and the scan range was 4000 – 400 cm^{-1} .

2.8 Adsorption study of methylene blue

For the typical time-dependent dye removal measurement, 1 mL hydrogels (100 mM NaDC/30 mM Asp in ultrapure water or polymer solutions) were first prepared and then, 3 mL methylene blue solution (0.025 mM) was added undisturbedly. The residual dye concentration after adsorption with hydrogel was measured using a UV-vis spectrophotometer (UV-2450, Shimadzu, Japan) at 25.0 ± 0.1 $^{\circ}$ C. All measurements were performed in a 1 cm-path length quartz cuvette in the wavelength range of 400 – 800 nm.



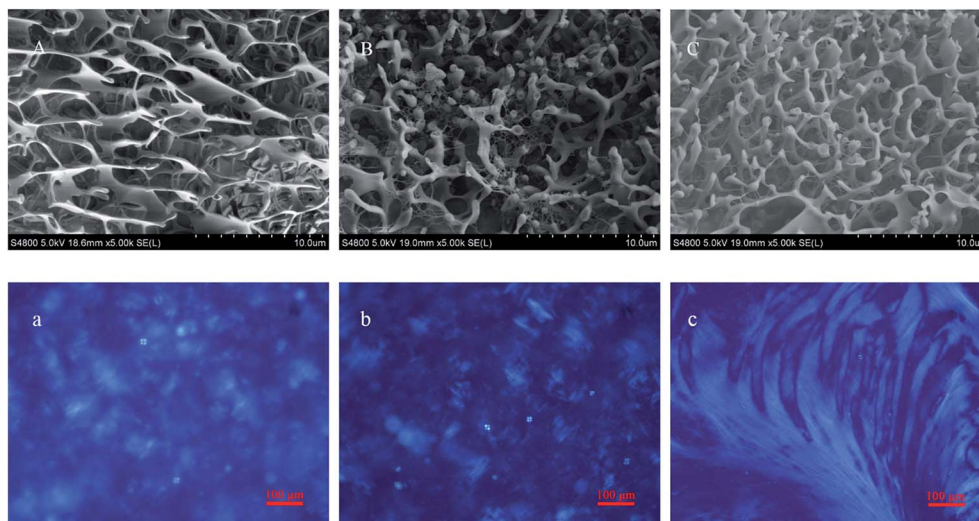


Fig. 3 SEM (A–C) and POM (a–c) images of (A and a) 100 mM NaDC/20 mM Asp, (B and b) 100 mM NaDC/30 mM Asp, and (C and c) 100 mM NaDC/40 mM Asp.

measurements indicate that G' and G'' increase with an increase in Asp concentration along with a change in appearance from hydrosols to hydrogels, indicating the enhancement in the strength of network structures. The η^* value of both hydrosols and hydrogels decreases with an increase in angular frequency, indicating shear-thinning behavior (Fig. 4(c)). Moreover, the value of η^* increases with an increase in Asp concentration, which implies that the addition of Asp favors the enhancement of the mechanical property of NaDC/Asp hydrogels.

From the above experimental results, it can be inferred that for NaDC/Asp system, the addition of Asp can improve the

mechanical properties of the hydrogels, which is in accordance with the results obtained from cryo-SEM. However, the ability of Asp to enhance the viscoelastic properties of the hydrogels is limited. The mechanical strength of network structures is still low, which limits the application of the NaDC/Asp hydrogels.

3.2 Effects of different polymers on the properties of NaDC/Asp hydrogels

Due to the poor mechanical properties of NaDC/Asp hydrogels, the polymer was introduced into the system. The additional

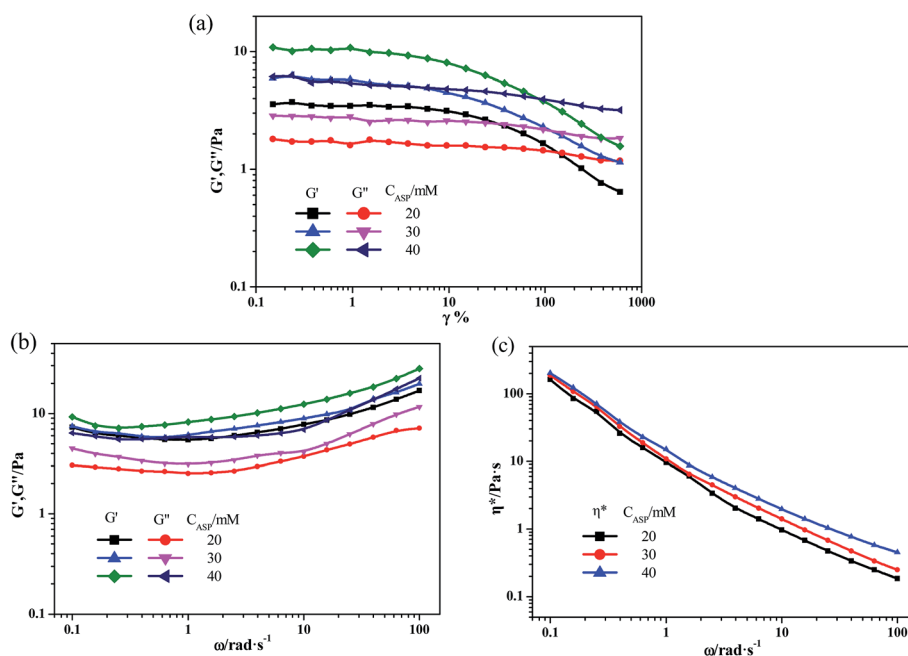


Fig. 4 Rheological properties of NaDC/Asp hydrogels. (a) The elastic modulus G' and the viscous modulus G'' as a function of strain γ at $25 \pm 0.1^\circ\text{C}$ under a fixed angular frequency $\omega = 6.28 \text{ rad s}^{-1}$. Dynamic rheology ((b) G' and G'' and (c) complex viscosity η^*) as a function of ω at $25 \pm 0.1^\circ\text{C}$ under a constant γ of 5%. The contents of NaDC and Asp are: $C_{\text{NaDC}} = 100 \text{ mM}$, and $C_{\text{Asp}} = 20, 30, 40 \text{ mM}$, respectively.



polymer may participate in the network of hydrogels and the oxygen-containing functional groups of polymer may provide numerous binding sites for hydrogen bonding, which could benefit for the mechanical strength of the NaDC/Asp hydrogels. To improve the mechanical strength of the NaDC/Asp hydrogels, polymers with different structures including linear polymer P(MEO₂MA_{90-co}-OEGMA₁₀) and star-shaped polymer CDPDPM (as shown in Fig. 1) were introduced into the NaDC/Asp system.

3.2.1 Microstructures of NaDC/Asp/polymer hydrogels. The microstructures of NaDC/Asp/polymer hydrogels were observed through cryo-SEM and POM. The appearances of hydrogels formed in NaDC/Asp systems with and without polymers show no visible difference. From Fig. 5(A–C), cryo-SEM images show that all hydrogels display a compact structure composed of irregular networks, suggesting that the polymers can incorporate into the network structure without changing the structure of the hydrogels. Moreover, with the addition of polymer, the density of network structures increases remarkably, indicating that the mechanical strength of hydrogels increases. Correspondingly, POM images display that Maltese crosses disappear in the polymer-containing hydrogels. Instead, the polymer-containing hydrogels show the fan-like textures of hexagonal structures and the birefringence property of the hydrogels becomes stronger (Fig. 5(a–c)). It can be noted that the addition of polymer benefits the enhancement of both mechanical strength and birefringence property of the NaDC/Asp hydrogels. Furthermore, the star-shaped polymer CDPDPM performs better in optimizing the properties of the hydrogels than the linear polymer P(MEO₂MA_{90-co}-OEGMA₁₀) because of the denser network structure and stronger birefringence property of the hydrogels.

3.2.2 Rheological properties of NaDC/Asp/polymer hydrogels. The mechanical properties of the hydrogels formed in NaDC/Asp/polymer systems were also investigated by dynamic rheological technique. Similar to the NaDC/Asp system, all the

hydrogels predominantly exhibit elastic properties, for which G' exceeds G'' over the investigated angular frequency range shown in Fig. 6(b). Moreover, both hydrogels show a shear-thinning behavior; correspondingly, the η^* decreases with an increase in angular frequency (Fig. 6(c)). However, the introduction of polymer improves the viscoelastic properties of hydrogels significantly. In stress sweep measurements (Fig. 6(a)), the γ^* value is about 20% for polymer-containing hydrogels, which is larger than that observed in the NaDC/Asp system with γ^* value of approximately 10%. Furthermore, the addition of polymer increases the G' by different degrees. For the system containing linear polymer P(MEO₂MA_{90-co}-OEGMA₁₀), the G' is 10 Pa, which is slightly higher than that of NaDC/Asp hydrogel (5 Pa). However, for CDPDPM-containing hydrogel, the G' value increases significantly and reaches 300 Pa, which is much higher than that of NaDC/Asp hydrogel. Furthermore, in oscillatory rheological measurements (Fig. 6(b)), both G' and G'' increase after the addition of the polymers. For P(MEO₂MA_{90-co}-OEGMA₁₀)-containing hydrogel, G' and G'' are about 11 Pa and 5 Pa, respectively, which are slightly larger than the values for the hydrogel induced by NaDC and Asp with G' and G'' around 6 Pa and 4 Pa, respectively. However, after addition of the star-shaped polymer CDPDPM, G' and G'' significantly improve to 75 Pa and 154 Pa, respectively. Moreover, the value of η^* increases drastically after the addition of polymer, particularly for the CDPDPM-containing hydrogel (Fig. 6(c)). These results indicate that the mechanical strength of the hydrogels can be improved by introducing a polymer. Moreover, the viscoelasticity of the CDPDPM-containing hydrogel is much higher than that of the P(MEO₂MA_{90-co}-OEGMA₁₀)-containing hydrogel under the same condition, indicating that the star-shaped polymer CDPDPM performs better in strengthening the network structure of the hydrogels than the linear polymer P(MEO₂MA_{90-co}-OEGMA₁₀).

3.2.3 Adsorption performances to MB. Toxic dye-based industrial effluents are hazardous to both ecological balance

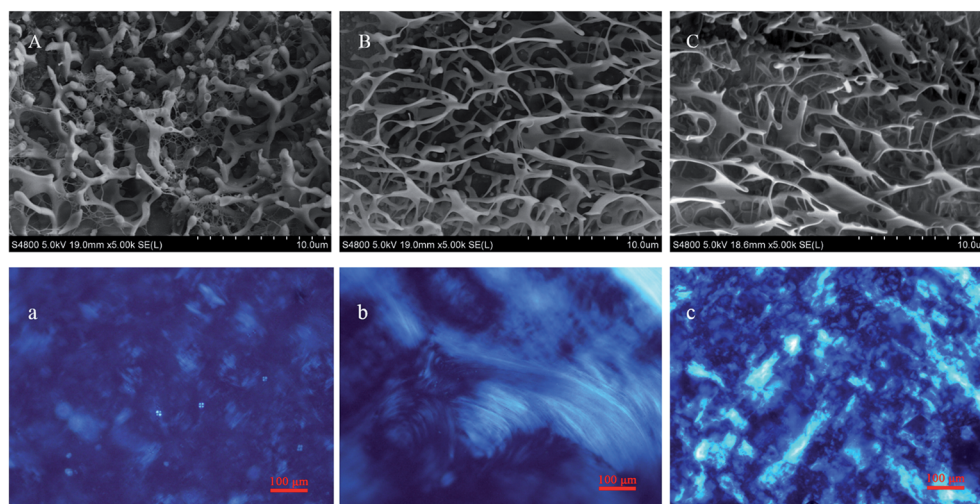


Fig. 5 SEM (A–C) and POM (a–c) images of 100 mM NaDC/30 mM Asp (A and a) without polymer, (B and b) with 1 g L^{−1} P(MEO₂MA_{90-co}-OEGMA₁₀), and (C and c) with 1 g L^{−1} CDPDPM.



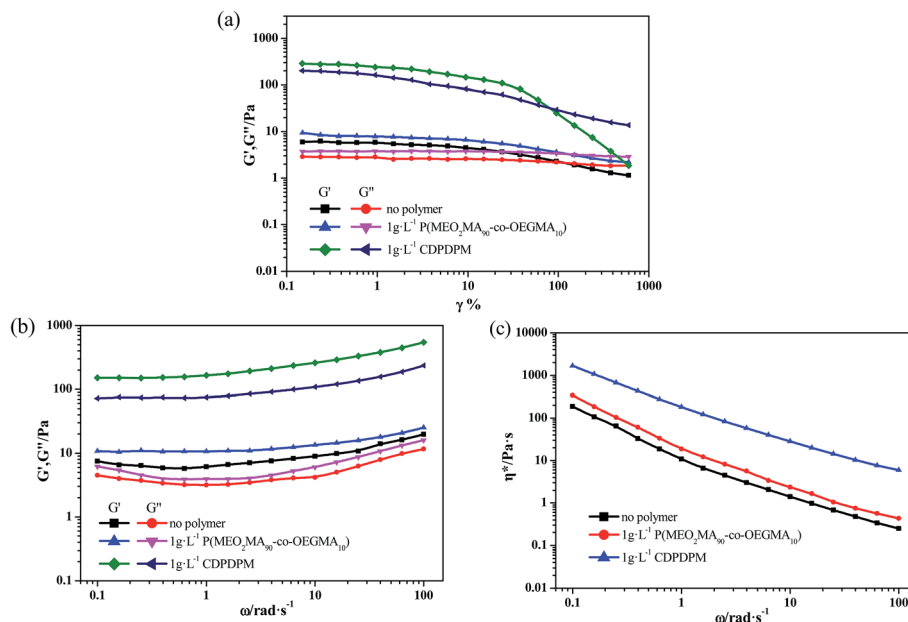


Fig. 6 Rheological properties of NaDC/Asp hydrogels with and without polymers. (a) The elastic modulus G' and the viscous modulus G'' as a function of strain γ at $25 \pm 0.1^\circ\text{C}$ under a fixed angular frequency $\omega = 6.28 \text{ rad s}^{-1}$. Dynamic rheology ((b) G' and G'' and (c) complex viscosity η^*) as a function of ω at $25 \pm 0.1^\circ\text{C}$ under a constant γ of 5%. The contents of NaDC and Asp are: $C_{\text{NaDC}} = 100 \text{ mM}$ and $C_{\text{Asp}} = 30 \text{ mM}$. The concentration of the polymer is 1 g L^{-1} .

and environment. Seeking convenient and efficient adsorbents for decontamination of waste waters is the goal that the researchers pursue. MB, an aromatic heterocyclic compound, is clinically used as the antidote of cyanide and nitrite poisoning as well as a photosensitizer to inactivate the virus in a single bag of fresh plasma. Therefore, MB is inevitably present in the medical sewage and the effective treatment of this type of medical sewage is significant to human health and environment.³⁹ In this study, MB was selected as the object to explore the adsorption properties of NaDC/Asp/polymer hydrogels.

The gels fabricated by 100 mM NaDC/30 mM Asp mixed system with and without polymer were used for evaluating the dye removal behavior. Fig. 7 provides the pictorial diagrams of dye adsorption for different gel systems. It can be observed that the hydrogel formed in the NaDC/Asp system shows no adsorption ability to MB, but tends to capture water molecules of MB solution. Therefore, the hydrogel swells and dissolves within 48 h (Fig. 7(a)). Though the introduction of P(MEO₂MA₉₀-co-OEGMA₁₀) increases the mechanical strength of the hydrogel slightly due to certain hydrophobicity, the hydrogel still swells after 4 h and collapses finally (Fig. 7(b)). However, for NaDC/Asp/CDPDPM hydrogel, the MB is efficiently entrapped inside the gel, gradually accompanied by the blue solution becomes clearer and clearer (Fig. 7(c)). Correspondingly, the variation in MB concentration in the aqueous solution was monitored by UV-vis spectra once it came in contact with the hydrogel (Fig. 7(d)). It is clearly shown that the absorbance decreases sharply after the MB aqueous solution comes in contact with the hydrogel and reaches stability after about 300 min. Moreover, it is worth noting that NaDC/Asp/CDPDPM hydrogel does not dissolve and almost does not swell after adsorption of MB,

which makes the recycling of hydrogel possible. This may be due to the special structure that provides more binding sites for hydrogen bonding, resulting in a significant increase in the mechanical strength. Moreover, the stronger hydrophobic CDPDPM can incorporate into the hydrophobic microdomains to inhibit the entry of water, thus avoiding the swelling and dissolution of the hydrogels. It can be concluded that due to the efficient adsorption of the toxic dye molecules and recyclable property, the CDPDPM-containing hydrogel can be used as an environmentally friendly water-purifying agent.

3.3 Gelation mechanism

For further studying the gelation mechanism of hydrogels formed in the NaDC/Asp/polymer system, FT-IR spectra were recorded as shown in Fig. 8. For all the samples, the wide peak at 3440 cm^{-1} is well known for being attributed to the symmetric and antisymmetric O-H stretching modes of NaDC. It can also be noted that the asymmetric and symmetric methylene stretching bands of NaDC are located at 2947 and 2858 cm^{-1} , respectively. For pure NaDC (Fig. 8(A)) curve (a), there is no peak observed at around 1703 cm^{-1} (attributed to the symmetric stretching vibration of $-\text{COOH}$). For the hydrogels formed by 100 mM NaDC with varied amounts of Asp (Fig. 8(A)) curves (b–d), the peak appearing at 1583 cm^{-1} is assigned to the antisymmetric stretching vibration of $-\text{COO}^-$, indicating the existence of the deprotonated form $-\text{COO}^-$ in NaDC molecules.^{40,41} The bands at 1703 , 1583 , and 1406 cm^{-1} indicate that the carboxyl and carboxylate species coexist as aggregates in solutions,²¹ proving the existence of $\text{O-H}\cdots\text{O}$ hydrogen bonding between NaDC and deoxycholic acid or Asp molecules. For NaDC/Asp/polymer hydrogels (Fig. 8(B)) curves



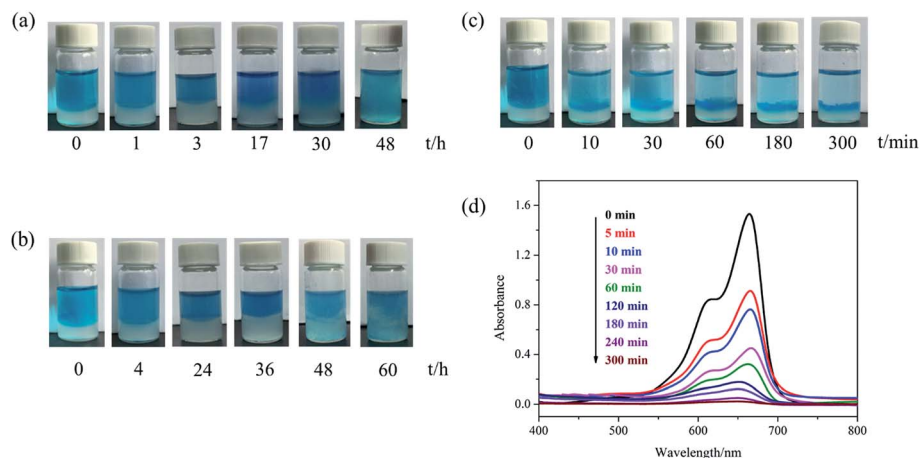


Fig. 7 The pictorial diagrams of dye adsorption for 100 mM NaDC/30 mM Asp hydrogel (a) without polymer, (b) with 1 g L^{-1} P(MEO₂MA₉₀-co-OEGMA₁₀), and (c) with 1 g L^{-1} CDPDPM. (d) UV-vis spectra of MB solution with time for the NaDC/Asp/CDPDPM hydrogel ($C_{\text{NaDC}} = 100 \text{ mM}$, $C_{\text{Asp}} = 30 \text{ mM}$ and $C_{\text{CDPDPM}} = 1 \text{ g L}^{-1}$).

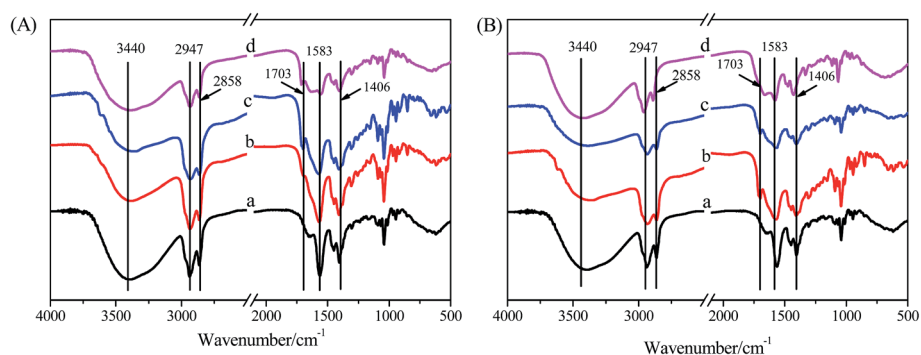


Fig. 8 (A) FT-IR spectra of NaDC (a) and hydrogels of 100 mM NaDC with varied amounts of Asp: (b) 20, (c) 30, and (d) 40 mM. (B) FT-IR spectra of NaDC (a) and hydrogels of 100 mM NaDC/30 mM Asp (b) without polymer, (c) with 1 g L^{-1} P(MEO₂MA₉₀-co-OEGMA₁₀), and (d) with 1 g L^{-1} CDPDPM.

(c and d), all bands of the FT-IR spectra are similar to that of NaDC/Asp hydrogels under the same concentration (Fig. 8(B)) curve (b), indicating that the addition of polymer does not change the gelation route of NaDC/Asp hydrogels though it incorporates in the network structure of the hydrogels.

In order to get more detailed information about the studied hydrogels, XRD spectra were recorded to detect the mechanism of formation of the hydrogels.^{23,26} According to Bragg's law, the spacing distance (d) can be calculated from the reflection peaks, which are considered to be related to some cyclic units induced by the arrangements of DC[−] species. Fig. 9 shows the XRD spectra of hydrogels formed in the NaDC/Asp system with and without polymers. For hydrogel formed in 100 mM NaDC/30 mM Asp system, a strong peak centered at $2\theta = 2.7$ appears, from which the calculated d is 3.27 nm, which is slightly larger than twice that of the deoxycholate backbone length ($1.5 \text{ nm} \times 2$).¹⁹ Rich *et al.* have similarly reported that the diameter of deoxycholate complexes increases once other molecules such as amino acids or peptides appear in the solution during the formation of the complexes.²⁰ The arrangement patterns of molecules in NaDC/Asp/polymer hydrogels were also

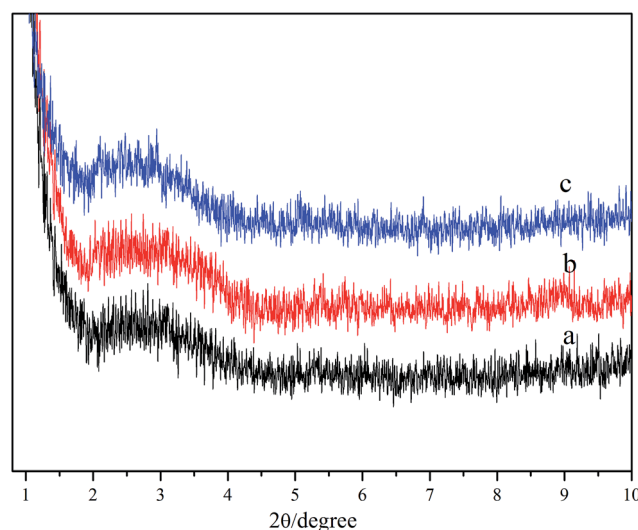
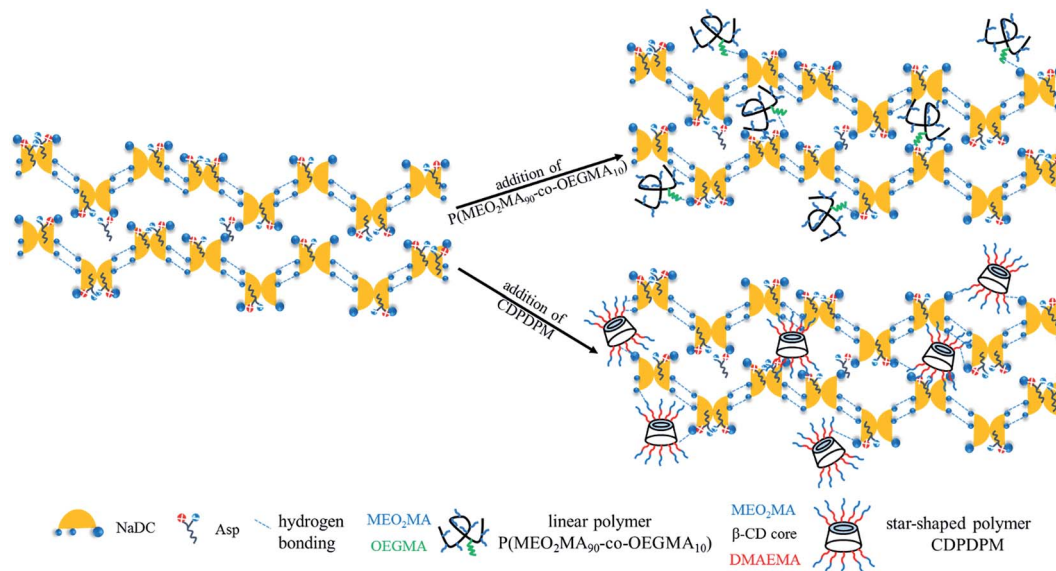


Fig. 9 XRD patterns of 100 mM NaDC/30 mM Asp xerogels (a) without polymer, (b) with 1 g L^{-1} P(MEO₂MA₉₀-co-OEGMA₁₀), and (c) with 1 g L^{-1} CDPDPM.





Scheme 1 Schematic illustration of the gelation mechanism for NaDC/Asp/polymer hydrogel.

investigated by XRD. As shown in Fig. 9, as compared to the hydrogel formed in the system without polymers, the d values of polymer-containing hydrogels show no visible change (3.29 nm for P(MEO₂MA₉₀-*co*-OEGMA₁₀)-containing hydrogel and 3.32 nm for CDPDPM-containing hydrogel), indicating that the additional polymer either linear or star-shaped cannot change the style of the cycled units induced by the arrangement of DC[−] species.

Based on the above results, the possible gelation mechanism of the NaDC/Asp/polymer hydrogels is proposed (Scheme 1). For NaDC/Asp hydrogel, the deoxycholate anions and amino acids may adopt lamellar arrangement, for which the hydrogen-bonded deoxycholate pairs are stacked together at their hydrophilic edges, allowing an opening angle between the planes of the two ring systems⁴² as shown in Scheme 1. In addition, two connected deoxycholate species form the continuous hydrophilic cavities and such interior cores act as the pocket, in which water molecules are held through interfacial tension.⁴¹ In fact, the process of the formation of hydrogels is complicated. The formation of this structure is driven by not only hydrogen bonding, but also hydrophobic interactions, electrostatic interactions, steric effects, and van der Waals forces. After the addition, the polymers can incorporate into the network of the hydrogels and correspondingly hydrogen bonds between the oxygen-containing functional groups of the polymer and the COO[−] or OH of NaDC molecules will be formed. In this case, the polymer provides numerous binding sites, so that NaDC can easily connect with the polymer, leading to a drastic enhancement in the mechanical strength of the gels. Clearly, the special structure enables the CDPDPM to provide more binding sites for hydrogen bonding and the stronger hydrophobicity endows CDPDPM the ability to inhibit the entry of water, thus avoiding the swelling and dissolution of hydrogels. Therefore, the star-shaped polymer CDPDPM has more advantages in strengthening the network structure of the hydrogels than the linear polymer.

4. Conclusion

The gelation behavior and properties of NaDC with Asp in aqueous solution were investigated in detail. In addition, the linear polymer P(MEO₂MA₉₀-*co*-OEGMA₁₀) and the star-shaped polymer CDPDPM were introduced into the NaDC/Asp hydrogels for exploring the effects of polymers on the properties of NaDC/Asp hydrogels systematically. The mechanism of gelation processes by polymers was also proposed. The results have indicated that the polymer can participate in the formation of a gel-network structure and provide numerous binding sites, so that NaDC can easily connect with the polymer, leading to an enhancement in the mechanical strength of the gels. Due to the special structure that provides more binding sites for hydrogen bonding and stronger hydrophobicity that inhibits the swelling and dissolution of hydrogels, the star-shaped polymer performs better in strengthening the network structure of the hydrogels than the linear polymer. Moreover, the introduction of the star-shaped polymer CDPDPM imparts additional functions to the composite hydrogels. The NaDC/Asp/CDPDPM hydrogel can adsorb MB efficiently without losing the gel state, thus making the recycling of hydrogel possible, while the NaDC/Asp hydrogel dissolves totally and NaDC/Asp/P(MEO₂MA₉₀-*co*-OEGMA₁₀) hydrogel collapses in the MB aqueous solution. This indicates that the NaDC/Asp/CDPDPM hydrogel can be used as a novel and environmentally friendly material to recover MB from medical sewage. Moreover, the NaDC/Asp/CDPDPM hydrogel has potential to be used in drug delivery systems and bioengineering.

Conflicts of interest

The authors declare no competing financial interest.



Acknowledgements

This work is supported by the National Natural Science Foundation of China (Project No. 21476072) and the Fundamental Research Funds for the Central Universities.

References

- 1 A. R. Hirst, B. Escuder, J. F. Miravet and D. K. Smith, *Angew. Chem., Int. Ed.*, 2008, **47**, 8002–8018.
- 2 J. Nanda and A. Banerjee, *Soft Matter*, 2012, **8**, 3380–3386.
- 3 S. Banerjee, R. K. Das and U. Maitra, *J. Mater. Chem.*, 2009, **19**, 6649–6687.
- 4 I. Yoshimura, Y. Miyahara, N. Kasagi, H. Yamane, A. Ojida and I. Hamachi, *J. Am. Chem. Soc.*, 2004, **126**, 12204–12205.
- 5 A. Wada, S.-i. Tamaru, M. Ikeda and I. Hamachi, *J. Am. Chem. Soc.*, 2009, **131**, 5321–5330.
- 6 G. A. Silva, C. Czeisler, K. L. Niece, E. Beniash, D. A. Harrington, J. A. Kessler and S. I. Stupp, *Science*, 2004, **303**, 1352–1355.
- 7 S. Yamamichi, Y. Jinno, N. Haraya, T. Oyoshi, H. Tomitori, K. Kashiwagi and M. Yamanaka, *Chem. Commun.*, 2011, **47**, 10344–10346.
- 8 G. C. Yu, X. Z. Yan, C. Y. Han and F. H. Huang, *Chem. Soc. Rev.*, 2013, **42**, 6697–6722.
- 9 M. Zhang, S. Strandman, K. C. Waldron and X. X. Zhu, *J. Mater. Chem. B*, 2016, **4**, 7506–7520.
- 10 J. W. Steed, *Chem. Soc. Rev.*, 2010, **39**, 3686–3699.
- 11 F. G. Nataliya, S. L. Anatolii, E. B. Vladimir and T. Aslan Yu, *Russ. Chem. Rev.*, 2017, **86**, 269.
- 12 K. Araki and I. Yoshikawa, in *Low Molecular Mass Gelator*, Springer Berlin Heidelberg, Berlin, Heidelberg, 2005, pp. 133–165, DOI: 10.1007/b107173.
- 13 U. Maitra, S. Mukhopadhyay, A. Sarkar, P. Rao and S. Indi, *Angew. Chem.*, 2001, **113**, 2341–2343.
- 14 N. Yan, G. He, H. Zhang, L. Ding and Y. Fang, *Langmuir*, 2010, **26**, 5909–5917.
- 15 A. Dasgupta, J. H. Mondal and D. Das, *RSC Adv.*, 2013, **3**, 9117–9149.
- 16 J. Moorthy, G. A. Mensing, D. Kim, S. Mohanty, D. T. Eddington, W. H. Tepp, E. A. Johnson and D. J. Beebe, *Electrophoresis*, 2004, **25**, 1705–1713.
- 17 S. Mukhopadhyay and U. Maitra, *Curr. Sci.*, 2004, **87**, 1666–1683.
- 18 J. W. Zhang and X. X. Zhu, *Sci. China, Ser. B: Chem.*, 2009, **52**, 849–861.
- 19 Y. Qiao, Y. Y. Lin, Y. J. Wang, Z. Y. Yang, J. Liu, J. Zhou, Y. Yan and J. B. Huang, *Nano Lett.*, 2009, **9**, 4500–4504.
- 20 D. M. Blow and A. Rich, *J. Am. Chem. Soc.*, 1960, **82**, 3566–3571.
- 21 G. H. Li, Y. Y. Hu, J. F. Sui, A. X. Song and J. C. Hao, *Langmuir*, 2016, **32**, 1502–1509.
- 22 Y. Qiao, Y. Y. Lin, S. F. Zhang and J. B. Huang, *Chem.–Eur. J.*, 2011, **17**, 5180–5187.
- 23 X. F. Sun, X. Xin, N. Tang, L. W. Guo, L. Wang and G. Y. Xu, *J. Phys. Chem. B*, 2014, **118**, 824–832.
- 24 A. Jover, F. Meijide and E. Rodríguez Núñez, *Langmuir*, 2002, **18**, 987–991.
- 25 Y. J. Zhang, X. Xin, J. L. Shen, W. Y. Tang, Y. J. Ren and L. Wang, *RSC Adv.*, 2014, **4**, 62262–62271.
- 26 Y. T. Wang, X. Xin, W. Z. Li, C. Y. Jia, L. Wang, J. L. Shen and G. Y. Xu, *J. Colloid Interface Sci.*, 2014, **431**, 82–89.
- 27 S. Kumar and P. Maiti, *RSC Adv.*, 2016, **6**, 67449–67480.
- 28 K. Liu, X. Jiang and P. Hunziker, *Nanoscale*, 2016, **8**, 16091–16156.
- 29 A. Duro-Castano, J. Movellan and M. J. Vicent, *Biomater. Sci.*, 2015, **3**, 1321.
- 30 L. Wang, X. Xin, M. Z. Yang, X. Ma, J. L. Shen, Z. H. Song and S. L. Yuan, *Colloids Surf., A*, 2015, **483**, 112–120.
- 31 D. D. Alford, A. P. Wheeler and C. A. Pettigrew, *J. Environ. Polym. Degrad.*, 1994, **2**, 225–236.
- 32 D. Voet, J. G. Voet and C. W. Pratt, *Fundamentals of biochemistry: life at the molecular level*, 2016.
- 33 Y. Guo, H. J. Liu, J. Q. Chen, Y. Z. Shang and H. L. Liu, *Acta Phys.-Chim. Sin.*, 2015, **31**, 1914–1923.
- 34 Y. Guo, M. Li, X. R. Li, Y. Z. Shang and H. L. Liu, *React. Funct. Polym.*, 2017, **116**, 77–86.
- 35 H. G. Li and J. C. Hao, *J. Phys. Chem. B*, 2008, **112**, 10497–10508.
- 36 L. Xu, G. Y. Xu, T. Liu, Y. J. Chen and H. J. Gong, *Carbohydr. Polym.*, 2013, **92**, 516–522.
- 37 G. Montalvo, M. Valiente and E. Rodenas, *Langmuir*, 1996, **12**, 5202–5208.
- 38 R. G. Weiss, *J. Am. Chem. Soc.*, 2014, **136**, 7519–7530.
- 39 K. Y. Foo and B. H. Hameed, *Chem. Eng. J.*, 2012, **203**, 81–87.
- 40 Y. Sun, Z. L. Yang, L. Zhang, N. F. Zhou, S. F. Weng and J. G. Wu, *J. Mol. Struct.*, 2003, **655**, 321–330.
- 41 L. Schefer, A. Sánchez-Ferrer, J. Adamcik and R. Mezzenga, *Langmuir*, 2012, **28**, 5999–6005.
- 42 L. B. Pártay, M. Segá and P. Jedlovský, *Langmuir*, 2007, **23**, 12322–12328.

