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Effect of a pH-controlled co-precipitation process on rhodamine B adsorption of MnFe_2O_4 nanoparticles

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We investigated the effect of a pH-controlled co-precipitation process on the adsorption behavior of manganese ferrite (MnFe_2O_4) nanoparticles as well as their structural and magnetic properties. The pH of prepared MnFe_2O_4 nanoparticles is typically an important factor affecting the adsorption capacity of an adsorbent. In this study, MnFe_2O_4 nanoparticles were prepared using a co-precipitation method at four different pH values of 9.0, 9.5, 10.0, and 10.5. The adsorption behaviors on rhodamine B (RhB) by MnFe_2O_4 nanoparticles prepared at different pH values were investigated. It was found that, *via* a pH-controlled process, MnFe_2O_4 nanoparticles prepared at pH 10.5 showed the highest RhB removal efficiency. The results indicated that the large pore size and surface charge of MnFe_2O_4 nanoparticles improved the adsorption capacities for RhB. Kinetic data were fitted to a pseudo-second order kinetic model and revealed that equilibrium was reached within 60 min. The isotherm data showed that the Langmuir maximum adsorption capacity of the MnFe_2O_4 nanoparticles prepared at pH 10.5 for RhB was 9.30 mg g^{-1} .

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1. Introduction

Many textile industries use processes, such as dyeing silk, wool, jute, leather and cotton that are major sources of severe worldwide water pollution problems.¹ The release of many dyes in the water is very harmful because these strong colored components create dangerous byproducts through oxidation, and inhibit benthic photosynthesis.^{2,3} The toxic and carcinogenic degraded products of dyes from industrial effluents released into clean water affect the central nervous system in humans and the environment.^{4,5} Rhodamine B (RhB), a cationic dye, is a major source of environmental pollution because of its stability and resistance to biodegradation.⁶ It can be toxic to human beings and animals, and causes the irritation of skin, eyes and respiratory tract.^{7,8} Thus, the development of efficient methods for RhB elimination is urgently needed. Several techniques such as advanced

oxidation processes,^{9,10} direct coagulation/flocculation-sand filtration or ozonation,^{11–14} membrane separation,¹⁵ coagulation and flocculation,¹⁶ as well as adsorption,^{17–19} have been used to remove dyes from wastewater. Adsorption processes are considered to be an effective, promising and economical alternative, which can be used to remove dyes in water.^{20–22} Recently, magnetic nanoparticles have been used to remove dyes from water. Among ferrites utilized for dye removal, manganese ferrite (MnFe_2O_4) is the most prominent ferrite because of its highest magnetization, high catalytic activity, and good durability. Moreover, MnFe_2O_4 provides an advantage for repeating magnetic separation with good stability.^{23–29} It has been shown that MnFe_2O_4 nanoparticles can be prepared *via* several methods such as hydrothermal,^{30,31} sol-gel,^{29,32} and co-precipitation.^{33–35} Among these methods, co-precipitation has been widely utilized for the preparation of magnetic ferrite. When compared to other methods, the products obtained from a co-precipitation method are relatively uniform, and can be prepared easily and scaled up for industrial purposes. From previous reports, the physical properties of MnFe_2O_4 nanoparticles were strongly dependent on the pH preparation value.^{36–38} Therefore, controlling the pH value within the co-precipitation process is important for the synthesis of MnFe_2O_4 nanoparticles with the desired morphology and size for efficient adsorption. To the best of our knowledge, the characteristics and adsorption properties of MnFe_2O_4 nanoparticles prepared at various pH values for RhB removal have not been investigated.

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In this study, the MnFe_2O_4 nanoparticles prepared at different pH values of 9.0, 9.5, 10.0, and 10.5 were synthesized by a co-precipitation method. The effect of controlling the pH during the co-precipitation process was studied to determine the crystal structure, morphology, magnetic properties of the MnFe_2O_4 nanoparticles. The adsorption capacities, adsorption kinetics and adsorption isotherms of the obtained MnFe_2O_4 nanoparticles were evaluated based on the efficiency of the RhB removal.

2. Experimental

2.1 Preparation of MnFe_2O_4 nanoparticles

MnFe_2O_4 nanoparticles were prepared using a co-precipitation method. Firstly, 25 mM of manganese nitrate hydrate ($\text{Mn}(\text{NO}_3)_2 \cdot x\text{H}_2\text{O}$, Sigma-Aldrich, USA, 98%) and 50 mM of iron nitrate nonahydrate ($\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$, Sigma-Aldrich, USA, 98%) were dissolved in deionized water. These solutions were mixed together using magnetic stirring at a constant speed. Then, the mixed solution was heated to 70 °C for 15 min, and then added dropwise 0.2 M of sodium hydroxide (NaOH, Sigma-Aldrich, 97%). The pH of each solution was adjusted to form different values of 9.0, 9.5, 10.0, and 10.5. Correspondingly, the MnFe_2O_4 samples produced during the pH-controlled preparation process were denoted as MnFe_2O_4 -pH 9.0, MnFe_2O_4 -pH 9.5, MnFe_2O_4 -pH 10.0, and MnFe_2O_4 -pH 10.5. The mixture solutions were then stirred magnetically for 24 h at room temperature. The obtained precipitates were washed with deionized water by centrifugation several times and were dried at 80 °C for 24 h. A flow diagram of MnFe_2O_4 nanoparticles preparation at the pH values of 9.0, 9.5, 10.0, and 10.5 by co-precipitation method are shown in Fig. 1.

2.2 Characterization

The crystal phases of all samples were characterized by X-ray diffraction (XRD, JEOL JDX-3530) with Cu K_α irradiation ($\lambda =$

1.5418 nm). The morphology and particle size were determined by scanning electron microscopy (SEM, JEOL JSM-6335F) and transmission electron microscopy (TEM, JEOL JEM-2010). The Brunauer–Emmett–Teller (BET) specific surface area and pore size of the prepared samples were analyzed by nitrogen adsorption–desorption isotherms obtained using an autosorb 1 MP (Quantachrome). The zeta potential determined by zetasizer nano instrument (ZS Malvern) was used to evaluate the isoelectric point (IEP) of the MnFe_2O_4 nanoparticles. A vibrating sample magnetometer (VSM, Lakeshore VSM 7400) was employed to measure the magnetic properties of the obtained samples. The Fourier transform infrared (FTIR) spectra of the samples were recorded in the 4000–400 cm^{-1} wavenumber region using KBr disks as standard on a Thermo Scientific model Nicolet 6700 FT-IR spectrophotometer. Surface composition and oxidation state were carried out using X-ray photoelectron spectroscopy (XPS, Kratos Axis ULTRA^{PLD}) with a monochromatic Al K_α source (1486.6 eV). All XPS spectra for the samples were referenced by setting the adventitious carbon C (1s) peak to 284.8 eV.

2.3 Adsorption experiments

The adsorption properties of MnFe_2O_4 nanoparticles prepared at different pH values for RhB removal were examined using a batch adsorption experiment. Initially, at room temperature, 100 mg of the adsorbent was added to 100 mL of RhB reaction solution with a concentration of 2×10^{-5} M. During the adsorption process, samples were taken at 5 min intervals for 60 min. The MnFe_2O_4 nanoparticles were separated from the RhB solution by centrifugation at 3000 rpm for 5 min. The solute concentration was determined by a UV-vis spectrophotometer (Thermo Scientific Evolution 201) at $\lambda_{\text{max}} = 550$ nm. The adsorption capacity and removal efficiency were calculated using the following equations:^{39,40}

$$q_e = \frac{(C_0 - C_e)V}{W} \quad (1)$$

$$E = \frac{(C_0 - C_e)}{C_0} \times 100\% \quad (2)$$

where q_e is the equilibrium adsorption capability (mg g^{-1}), C_0 and C_e are the initial and equilibrium concentrations of RhB (mg L^{-1}), W is the mass (g) of the magnetic adsorbent, and V is the volume of the solution (mL).

3. Results and discussion

3.1 Crystal structure

The XRD pattern of the MnFe_2O_4 powders prepared at different pH values are shown in Fig. 2. The considered diffraction peaks with 2θ values for MnFe_2O_4 -pH 10.5 were 18.08°, 29.74°, 35.02°, 36.66°, 42.57°, 52.82°, 56.26°, and 61.74° corresponding to the crystal planes (111), (220), (311), (222), (400), (422), (511), and (440), respectively. According to the JCPDS standard (card no. 88–1965), the crystal structure of MnFe_2O_4 powders can be indexed as spinel structure belonging to the face center cubic (FCC) system. No diffraction peak was attributed to manganese

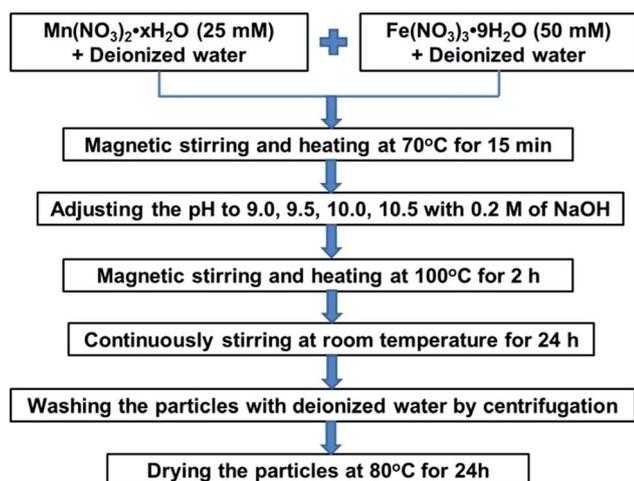


Fig. 1 Flow diagram of the MnFe_2O_4 nanoparticles preparation by co-precipitation at the pH values of 9.0, 9.5, 10.0, and 10.5.



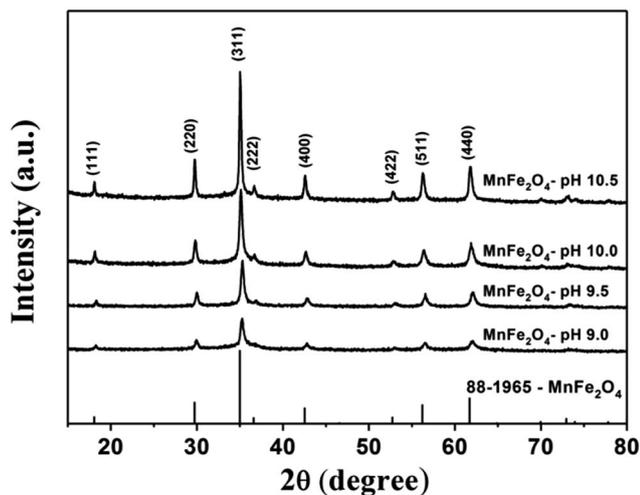


Fig. 2 XRD patterns of MnFe_2O_4 nanoparticles prepared by preparation by co-precipitation at the pH values of 9.0, 9.5, 10.0, and 10.5.

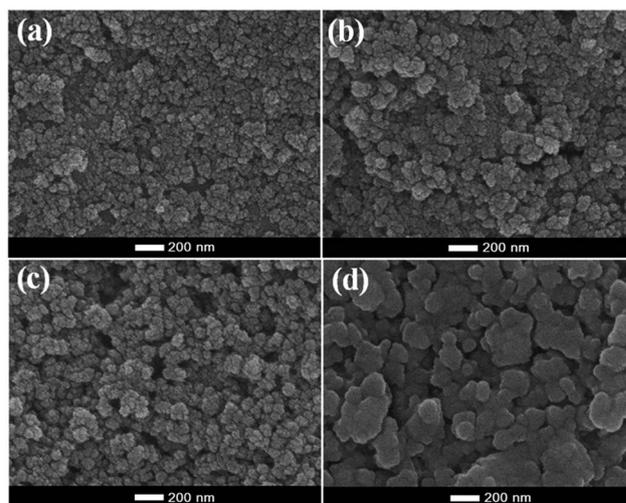


Fig. 3 SEM images of (a) MnFe_2O_4 -pH 9.0, (b) MnFe_2O_4 -pH 9.5, (c) MnFe_2O_4 -pH 10.0 and (d) MnFe_2O_4 -pH 10.5.

oxide and/or iron oxide materials. This suggests that the well-crystallized MnFe_2O_4 phase with a spinel structure has formed during the co-precipitation process.

As seen from Fig. 2, the XRD diffraction pattern of MnFe_2O_4 -pH 9.0 shows the lowest intensity due to its small crystallite size with low crystallinity. On the other hand, the intensity of the (311) peak increased with increasing pH values from 9.5 to 10.5, indicating high crystallinity and large crystallite sizes for particles obtained under these preparation conditions. These results imply that the crystallite size and crystallinity of MnFe_2O_4 nanoparticles can be adjusted with the pH value used during the co-preparation process. The crystallite size of all samples was calculated using the Debye–Scherrer formula.⁴¹ The calculated crystallite size of MnFe_2O_4 at the pH values of 9.0, 9.5, 10.0, and 10.5 were 29.5, 41.2, 54.2, and 103.7 nm, respectively (Table 1).

3.2 Morphology

The morphology and particle size of MnFe_2O_4 nanoparticles with varying pH values (pH 9.0, 9.5, 10.0, and 10.5) were investigated by SEM and TEM images, and are shown in Fig. 3 and 4 as well as Table 1. Fig. 3 displays the SEM images of MnFe_2O_4 prepared at different pH values of 9.0, 9.5, 10.0, and 10.5. The images indicated that synthesized manganese ferrites were composed of agglomerated small particles. The morphology of MnFe_2O_4 nanoparticles appears to be not-so-

hard-grained clusters. These nanoparticles consist of the spherical particles with average diameters of 69, 98, 149, and 247 nm for the preparation at pH values of 9.0, 9.5, 10.0, and 10.5, respectively. Furthermore, TEM images can reveal internal structures, and provide more accurate measurement of particle sizes and morphology. The morphology of MnFe_2O_4 nanoparticles from TEM images displayed spherical-like shapes. The average particle sizes of MnFe_2O_4 nanoparticles were 4.4, 4.5, 5.5, and 6.4 nm at the pH values 9.0, 9.5, 10.0, and 10.5, respectively (Fig. 4e–h). Fig. 4a–d show that the agglomerated particles diameter of MnFe_2O_4 increased as the pH value increased. The increase in the pH of the mixture solution resulted in more agglomeration of particles as the average particle size increased.^{42,43} The agglomeration of MnFe_2O_4 nanoparticles affects the BET-specific surface area (SSA) (these results are shown in the Section 3.3 BET-specific surface area and pore size). The corresponding HRTEM images in Fig. 4i–l show lattice spacings of the MnFe_2O_4 nanoparticles prepared at pH 9.0 to 10.5. The d spacing of approximately 0.25 nm corresponds to the (311) plane of MnFe_2O_4 .

3.3 BET-specific surface area and pore size

Fig. 5 shows the N_2 adsorption–desorption isotherms area and the pore size distribution of the as-prepared MnFe_2O_4 nanoparticles. The N_2 adsorption–desorption isotherm of MnFe_2O_4 nanoparticles was matched to a type IV hysteresis loop

Table 1 Physical properties of MnFe_2O_4 nanoparticles prepared by a co-precipitation at different pH values

MnFe_2O_4	XRD-crystallite size (nm)	TEM-particle size (nm)	SSA ($\text{m}^2 \text{g}^{-1}$)	Average pore size diameter (nm)	M_s (emu g^{-1})
pH 9.0	29.5	4.4	182.82	2.934	24.58
pH 9.5	41.2	4.5	139.25	3.371	30.60
pH 10.0	54.2	5.5	157.59	3.849	36.72
pH 10.5	103.7	6.4	53.99	5.094	41.48



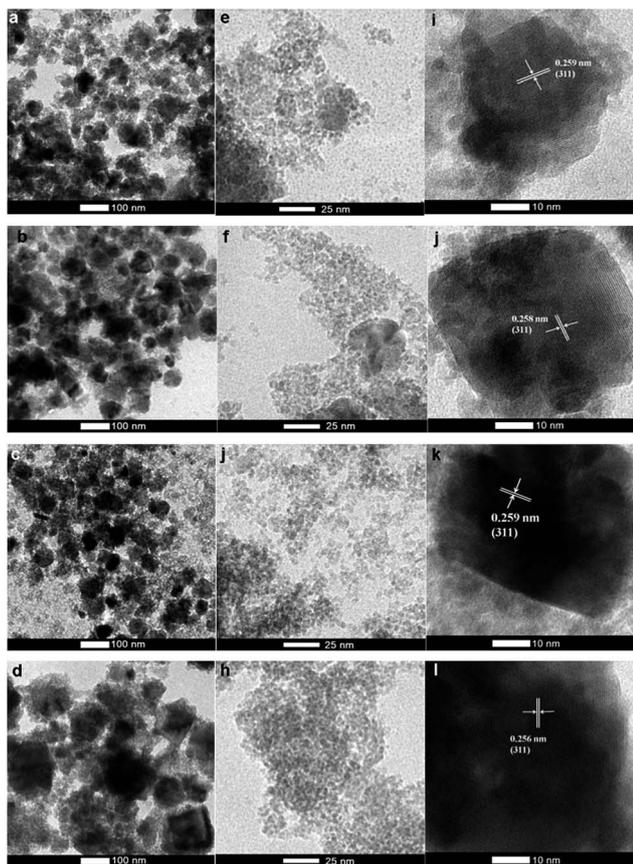


Fig. 4 TEM images and lattice fringes of MnFe_2O_4 nanoparticles prepared at the pH values of pH 9.0 (a, e and i), pH 9.5 (b, f and j), pH 10.0 (c, g and k), and pH 10.5 (d, h and l).

(according to the IUPAC classification) indicating a mesoporous structure. The BET-specific surface areas of the MnFe_2O_4 nanoparticles prepared at different pH values corresponded to the average pore size diameter as listed in Table 1. The MnFe_2O_4 -pH 9.0 shows a specific surface area of $182.82 \text{ m}^2 \text{ g}^{-1}$

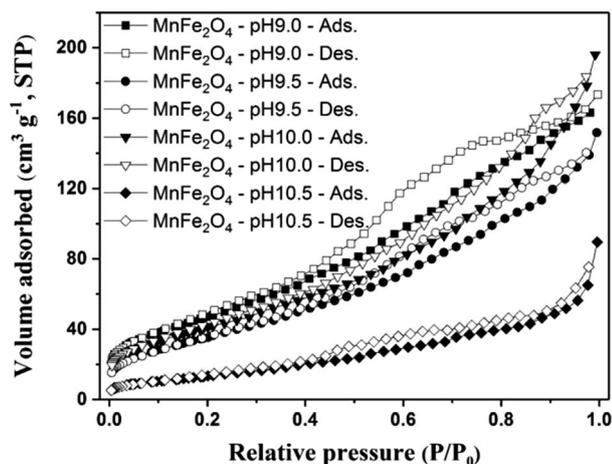


Fig. 5 N_2 adsorption-desorption isotherms for MnFe_2O_4 nanoparticles prepared at pH 9.0, 9.5, 10.0, and 10.5.

and increasing the pH to 10.5 shows a surface area notably diminished to $53.99 \text{ m}^2 \text{ g}^{-1}$. This is possibly because of the progressive aggregation of small crystallites into larger particles. The average pore size diameter was observed to increase as the pH increased. Furthermore, the most significant increase to a pore size of 5.094 nm was noted at the highest pH of 10.5. This result suggests that the larger pore size is beneficial to an improved ability of RhB to adsorb to the surface of MnFe_2O_4 nanoparticles. In addition to pore size of the adsorbent, other factors such as initial dye concentration and surface charge of the adsorbent influence the adsorption capacity of adsorbent materials.^{43–46}

3.4 Magnetic properties

The magnetic properties of MnFe_2O_4 nanoparticles were investigated by using a vibrating sample magnetometer (VSM) at room temperature under magnetic fields up to $\pm 10 \text{ kOe}$ as shown in Fig. 6a. All MnFe_2O_4 samples exhibited soft magnetic behavior with a narrow hysteresis loop. The values of saturation magnetization (M_s) of MnFe_2O_4 nanoparticles prepared at pH 9.0, 9.5, 10.0, and 10.5 are shown in Table 1. The saturation magnetization of the MnFe_2O_4 nanoparticles increased as the increased pH value of the prepared MnFe_2O_4 . MnFe_2O_4 -pH 10.5 showed the highest M_s value (41.48 emu g^{-1}) due to the large magnetic dipole moments from the large particle size and the high crystallinity of the particles. The results also demonstrate a faster magnetic attraction by external magnetic field and probably a more effective removal of the MnFe_2O_4 particles from the large volume of water. In Fig. 6b, MnFe_2O_4 -pH 10.5 was separated in solution after applying a magnetic field. It can be

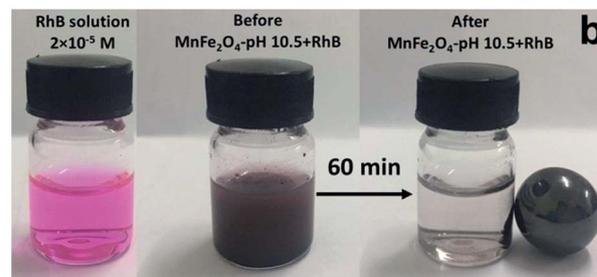
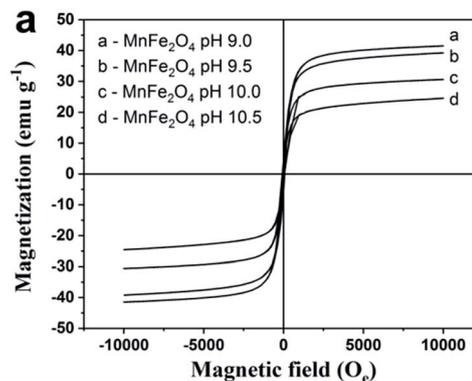


Fig. 6 (a) M–H curves of MnFe_2O_4 nanoparticles prepared at different pH values and (b) photograph of MnFe_2O_4 nanoparticles separation in RhB solution by using a magnet.



seen that the magnetic powders of MnFe_2O_4 -pH 10.5 immediately moved towards the magnet and the RhB solution became clear.

3.5 Adsorption kinetics

In Fig. 7, the maximum removal percentage of RhB (96.92%) was observed for MnFe_2O_4 -pH 10.5. The adsorption capacity of MnFe_2O_4 nanoparticles over RhB could be partially explained by the surface charges characterized by zeta potential. The point of zero charge (pzc) values of MnFe_2O_4 nanoparticles prepared at pH 9.0, 9.5, 10.0, and 10.5 were 5.58, 6.13, 6.69, and 5.18, respectively. The pH value of the suspension of the MnFe_2O_4 nanoparticles and RhB solution was measured to be approximately 8.8 which explains the negative charges on the surface of the MnFe_2O_4 . For this reason, the cationic RhB molecules would preferentially adsorb onto the negatively charged MnFe_2O_4 nanoparticles through electrostatic interaction.⁴⁷

In order to further analyze the adsorption kinetics for removing RhB by MnFe_2O_4 nanoparticles, two kinetic models, the pseudo-first-order and pseudo-second-order, models, were applied. The adsorption capacities of MnFe_2O_4 nanoparticles for RhB adsorption are shown in Fig. 8a. The pseudo-first order equation is generally represented according to the following equation^{48,49} (eqn (3)):

$$\frac{dq_t}{dt} = k_1(q_e - q_t) \quad (3)$$

where k_1 and q_t are the rate constant of pseudo-first order kinetics (min^{-1}) and adsorption capacity at time t (min), respectively. After integration by applying the conditions, $q_t = 0$ at $t = 0$ and $q_t = q_t$ at $t = t$, the eqn (3) becomes

$$\log(q_e - q_t) = \log(q_e) - \left(\frac{k_1}{2.303}\right)t \quad (4)$$

The k_1 , calculated q_e ($q_{e, \text{cal}}$), and correlation coefficients (R^2) of MnFe_2O_4 with varying pH values were determined as shown in Table 2. The plot of $\log(q_e - q_t)$ versus time after fitting the

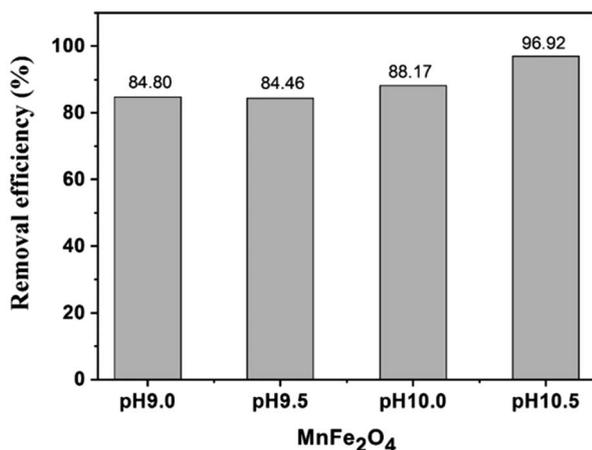


Fig. 7 Removal efficiency of RhB by MnFe_2O_4 nanoparticles prepared by co-precipitation at different pH values.

pseudo-first order model is presented in Fig. 8b. The $q_{e, \text{cal}}$ values of all samples were calculated from the intercept of the plot of $\log(q_e)$ versus t . From the obtained data, R^2 was relatively low, indicating that the pseudo-first order model was not well suited to predict the kinetic rate constant. Therefore, the pseudo-second order model was applied to determine the

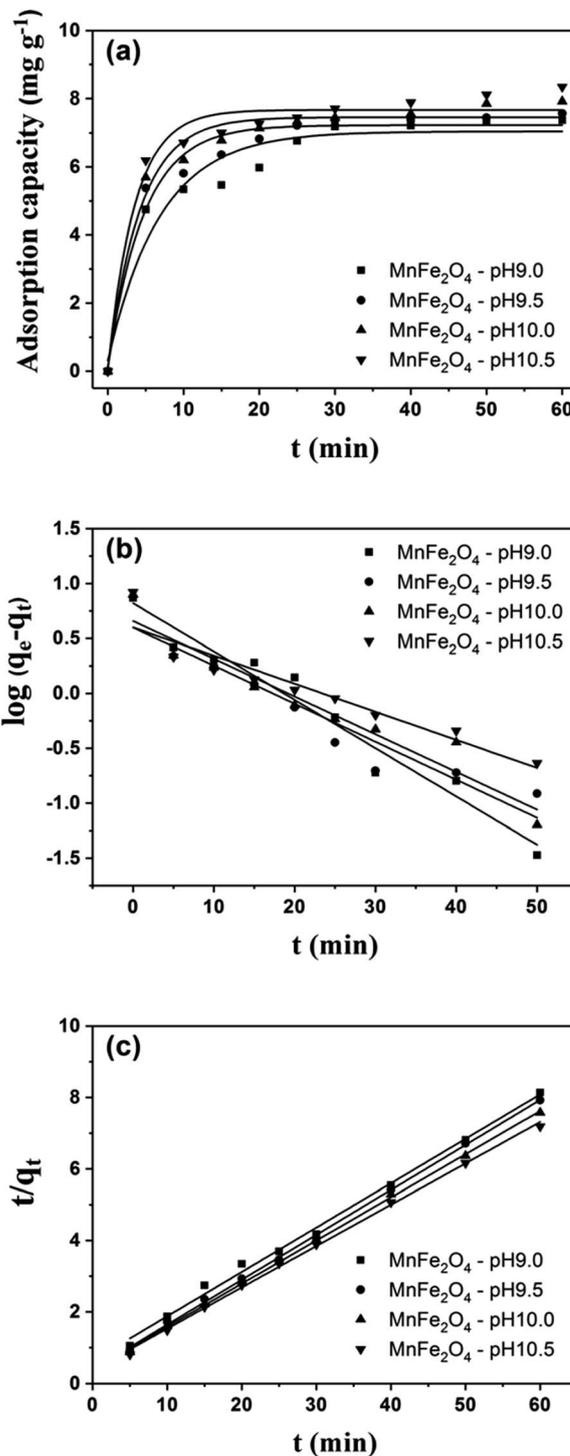


Fig. 8 Adsorption kinetics of (a) MnFe_2O_4 , the plots of (b) pseudo-first order, and (c) pseudo-second order kinetic models for RhB adsorption onto MnFe_2O_4 nanoparticles with varying pH preparation values.



Table 2 Kinetic parameters, q_e , $q_{e, cal}$, k_1 , k_2 , and correlation of coefficient R^2 for the adsorption of RhB onto $MnFe_2O_4$ nanoparticles prepared at varying pH values

Samples	q_e ($mg\ g^{-1}$)	Pseudo-first order			Pseudo-second order		
		$q_{e, cal}$ ($mg\ g^{-1}$)	k_1 (min^{-1})	R^2	$q_{e, cal}$ ($mg\ g^{-1}$)	k_2 ($g\ mg^{-1}\ min^{-1}$)	R^2
$MnFe_2O_4$ -pH 9.0	7.37	6.61	0.101	0.9587	7.76	0.283	0.9916
$MnFe_2O_4$ -pH 9.5	7.57	3.97	0.079	0.9143	7.78	0.469	0.9975
$MnFe_2O_4$ -pH 10.0	7.92	4.55	0.078	0.9295	8.10	0.460	0.9976
$MnFe_2O_4$ -pH 10.5	8.35	3.97	0.059	0.8975	8.45	0.443	0.9967

suitable kinetic model which is expressed as the following equation:

$$\frac{dq_t}{dt} = k_2(q_e - q_t) \quad (5)$$

where k_2 is the pseudo-second order rate constant ($g\ mg^{-1}\ min^{-1}$). After integrating eqn (5), the linear fitting is expressed as the following equation:

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \quad (6)$$

Fig. 8c shows pseudo-second-order kinetic models for RhB adsorption onto $MnFe_2O_4$ nanoparticles with varying pH preparation values. The adsorption kinetics and parameters of k_2 and $q_{e, cal}$ from the intercept and slope of the plot of t/q_t versus t are shown in Table 2. The R^2 values show that the RhB removal kinetics can be approximated using a pseudo-second order model. Thus, the pseudo-second order model is more appropriate for explaining this adsorption process.

3.6 Adsorption isotherms

Two adsorption isotherm models, Langmuir and Freundlich equations, were used to further describe the adsorption equilibrium. The Langmuir equation describes monolayer adsorption with limited homogeneous sites. The Langmuir model is given by the following equation:^{50,51}

$$q_e = \frac{q_m K_L C_e}{(1 + K_L C_e)} \quad (7)$$

where C_e , K_L , q_e , q_m , are the concentration of adsorbate at equilibrium ($mg\ L^{-1}$), Langmuir constant ($L\ mg^{-1}$), the amount of solute adsorbed at the equilibrium of RhB ($mg\ g^{-1}$), and the maximum adsorption capacity of RhB ($mg\ g^{-1}$), respectively. Eqn (7) can be expressed in a linear form as displayed in the following equation:

$$\frac{C_e}{q_e} = \frac{1}{K_L q_m} + \frac{C_e}{q_m} \quad (8)$$

Fig. 9a shows linear forms of the isotherms for representative RhB adsorption by $MnFe_2O_4$ nanoparticles with varying preparation of pH values. The Langmuir isotherm for RhB adsorption on $MnFe_2O_4$ nanoparticles was examined from a linear plot of C_e/q_e versus C_e as presented in Fig. 9b.

The Freundlich isotherm is one of the applicable models to describe heterogeneous surface adsorption and multilayer adsorption under various non-ideal conditions. The Freundlich

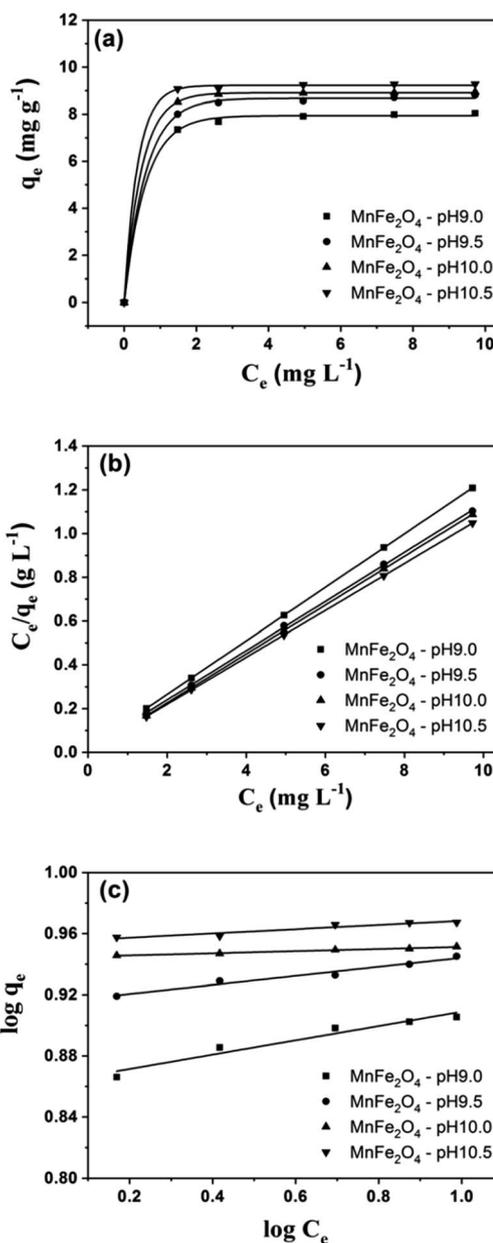


Fig. 9 (a) Adsorption isotherms, (b) a typical linear Langmuir, and (c) Freundlich isotherms of RhB removal using $MnFe_2O_4$ nanoparticles with varying pH preparation values.



Table 3 Extracted values of q_m , K_L , K_F , and n from RhB adsorption using MnFe_2O_4 nanoparticles by curve fitting based on Langmuir and Freundlich isotherms

Samples	q_e (mg g^{-1})	Langmuir equation			Freundlich equation		
		q_m (mg g^{-1})	K_L (L mg^{-1})	R^2	K_F (L mg^{-1})	$1/n$	R^2
MnFe_2O_4 -pH 9.0	8.04	8.09	10.75	0.9998	7.28	0.0469	0.9052
MnFe_2O_4 -pH 9.5	8.81	8.83	13.33	0.9998	8.22	0.0295	0.9688
MnFe_2O_4 -pH 10.0	8.94	8.94	22.82	0.9999	8.80	0.0170	0.9869
MnFe_2O_4 -pH 10.5	9.28	9.30	34.68	0.9999	9.01	0.0139	0.9139

isotherm can be shown in the simplest model^{50,51} as the following:

$$q_e = K_F C_e^{\frac{1}{n}} \quad (9)$$

where K_F and $1/n$ are the adsorption capacity at unit concentration (L mg^{-1}) and adsorption intensity, respectively. The $1/n$ indicates the degree of non-linearity between solution concentration and adsorption, $1/n = 0$ (irreversible), $0 < 1/n < 1$ (favorable), and $1/n > 1$ (unfavorable). The linear form of the Freundlich isotherm can be expressed as the following equation:

$$\log q_e = \log K_F + \frac{1}{n} \log C_e \quad (10)$$

The values of K_F , n , and R^2 are shown in Table 3. The fitting result of the Freundlich model using eqn (10) for RhB adsorption on MnFe_2O_4 nanoparticles is shown in Fig. 9c. From the fitting results of the Langmuir and Freundlich models, the experiment isotherm fits quite well with the Langmuir model based on an examination of the values of each model's constant and correlation coefficient (R^2). This implies that the RhB adsorption onto MnFe_2O_4 is monolayer adsorption and the adsorption sites of MnFe_2O_4 nanoparticles are homogeneous. Moreover, the practical value of the equilibrium adsorption ability is much closer to the theoretical value obtained from the Langmuir model.

The adsorption capacity of the as-prepared MnFe_2O_4 nanoparticles was compared with other reports as shown in Table 4.^{52–57} It was found that the adsorption capacity of adsorbents

depends on several factors such as size, morphology, and specific surface area. Moreover, the extrinsic properties of adsorption capacity such as adsorbent loading and initial dye concentration significantly affected to the removal efficiency. No report has been found for MnFe_2O_4 reacting with RhB from previous work (Table 5).^{58–63} It was found that MnFe_2O_4 nanoparticles can adsorb RhB from water with adsorption capacity of 9.30 mg g^{-1} . This indicates that the prepared MnFe_2O_4 nanoparticles can be considered as a promising adsorbent for the removal of rhodamine B from aqueous solution.

3.7 Fourier transform infrared analysis

The FTIR spectroscopy was performed to investigate the possible interactions between MnFe_2O_4 nanoparticles and RhB as shown in Fig. 10. The MnFe_2O_4 -pH 10.5 exhibits an adsorption peak at 1624 cm^{-1} , corresponding to O–H group due to the adsorbed water or moisture in samples. The broad band at around 3392 cm^{-1} is attributed to the symmetric vibration of –OH groups of H_2O adsorbed by the sample. Meanwhile, the two bands at 576 and 1384 cm^{-1} can be attributed to the intrinsic vibrations of octahedral coordinated metal ions in the spinel structure and OH bending vibration, respectively. Typically, the metal oxide vibrations occur below 1000 cm^{-1} .^{64,65} For spectrum of RhB, the weak intensity band at 2983 cm^{-1} is associated with the aromatic C–H vibrations, which also appear at 1130 cm^{-1} (plane bending) and at 682 cm^{-1} (wagging vibrations). Moreover, the vibration bands in 2928 – 2869 cm^{-1} is inferred to the presence of RhB. In addition, the band at 1696 cm^{-1} corresponds to the C=O stretching vibration, while the band at 1341 cm^{-1} is associated with aromatic skeletal C–C

Table 4 Comparison of the studied dyes with MnFe_2O_4 and MnFe_2O_4 -based composites as adsorbents based on the removal capacities

Adsorbent	Properties			Adsorption properties				
	Size (nm)	SSA ($\text{m}^2 \text{ g}^{-1}$)	M_s (emu g^{-1})	Dye	Initial dye concentration (mg L^{-1})	Adsorbent loading (g L^{-1})	q_e (mg g^{-1})	Reference
MnFe_2O_4	80–100	—	78.4	Congo red	100	0.3	92.4	52
MnFe_2O_4	200–290	155.7	18.8	Congo red	150	16	25.78	53
MnFe_2O_4	2470	68.6	42.8	Azo dye acid red B	100	4	53.8	54
MnFe_2O_4 /bentonite	300–800	130	—	AR138	100	0.1 g/30 ml	—	55
MnFe_2O_4 /PW	21.1–25.6	—	1.53	Congo red	50–100	0.05–0.1 g	86.96	56
MnFe_2O_4 @resin	<5 (TEM)	—	6.2	Congo red	10	0.2	1000	57
MnFe_2O_4	5	53.99	41.48	RhB	7.5	0.1	9.30	In this study



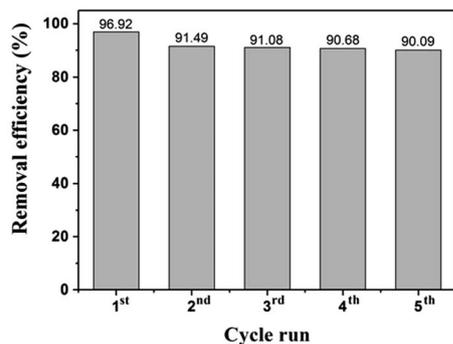


Fig. 12 Removal efficiencies of RhB during adsorption cycles for MnFe_2O_4 -pH 10.5.

MnFe_2O_4 particles.^{78,79} The significant binding energy shifts observed from XPS study as well as the results from UV-vis and zeta potential analyses suggest there is strong electrostatic and chemical interactions between MnFe_2O_4 and RhB. These results confirm the adsorption of RhB on the MnFe_2O_4 surface.

3.9 Reusability of MnFe_2O_4 nanoparticles

Reusability is one of the prerequisites to verify the practicality of a direct application of adsorbents for dye removal. The stability of the MnFe_2O_4 -pH 10.5 was evaluated by repeating experiments on the RhB adsorption as shown in Fig. 12.

After each run, the MnFe_2O_4 nanoparticles were collected using a magnetic field and washed by ultrasonic cleaning with deionized water following with absolute ethanol. Then, the separated MnFe_2O_4 nanoparticles were re-dispersed in fresh RhB solution with the same concentration (2×10^{-5} M), and next the new adsorption cycle began again. After 5 cycles, the removal efficiency decreased by a small fraction from 96.92% to 90.09%.

4. Conclusions

The magnetically separable MnFe_2O_4 nanoparticles were successfully prepared by a co-precipitation method with varying pH values. It was found that the physical, magnetic, and adsorption properties of MnFe_2O_4 nanoparticles were changed by controlling the pH value during MnFe_2O_4 nanoparticles preparation. The MnFe_2O_4 -pH 10.5 sample exhibited the highest adsorption capacity towards RhB solution. The highest adsorption of RhB on the MnFe_2O_4 -pH 10.5 sample depends on two factors that are large pore size and negative charge surface. Moreover, the reusability of MnFe_2O_4 -pH 10.5 particles can be achieved by an external magnetic field separation with magnetization of 41.48 emu g^{-1} . The experimental results correlated reasonably well with the Langmuir model. The adsorption kinetics of RhB by MnFe_2O_4 followed the pseudo-second order model. This study concludes that the facile synthesized magnetic MnFe_2O_4 nanoparticles prove to be effective as a potential adsorbent for the removal of RhB from aqueous solutions and a promising adsorbent for further utilizations in dye removal from colored textile wastewater.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 S. Zhu, T. Xu, H. Fu, J. Zhao and Y. Zhu, *Environ. Sci. Technol.*, 2007, **41**, 6234.
- 2 P. Saritha, C. Aparna, V. Himabindu and Y. Anjaneyulu, *J. Hazard. Mater.*, 2007, **149**, 609.
- 3 S. C. Jung, *Water Sci. Technol.*, 2011, **63**(7), 1491.
- 4 T. Yao, S. Guo, C. Zeng, C. Wang and L. Zhang, *J. Hazard. Mater.*, 2015, **292**, 90.
- 5 Q. Li, Q. Yue, Y. Su, B. Gao and J. Li, *J. Hazard. Mater.*, 2009, **165**, 1170.
- 6 E. Baldev, D. Mubarakali, A. Ilavarasi, D. Pandiaraj, K. A. Sheik Syed Ishack and N. Thajuddin, *Colloids Surf., B*, 2013, **105**, 207.
- 7 I. K. Konstantinou and T. A. Albanis, *Appl. Catal., B*, 2004, **49**, 1.
- 8 H. Lee, S. H. Park, Y.-K. Park, B. H. Kim, S.-J. Kim and S.-C. Jung, *Chem. Cent. J.*, 2013, **7**(156), 1.
- 9 C. Galindo, P. Jacques and A. Kalt, *Chemosphere*, 2001, **45**, 997.
- 10 W. S. Kuo and P. H. Ho, *Chemosphere*, 2001, **45**, 77.
- 11 A. Y. Zahrim and N. Hilal, *Water Resources and Industry*, 2014, **3**, 23.
- 12 S. A. Avlonitis, I. Poulis, D. Sotiriou, M. Pappas and K. Moutesidis, *Desalination*, 2008, **221**, 259.
- 13 Q. Y. Yue, B. Y. Gao, Y. Wang, H. Zhang, X. Sun, S. G. Wang and R. R. Gu, *J. Hazard. Mater.*, 2008, **152**, 221.
- 14 S. Wang, *Dyes Pigm.*, 2008, **76**, 714.
- 15 M. Sundrarajan, G. Vishnu and K. Joseph, *Dyes Pigm.*, 2007, **75**, 273.
- 16 J. Liu, S. Ma and L. Zang, *Appl. Surf. Sci.*, 2013, **265**, 393.
- 17 H. Tajizadegan, M. Jafari, M. Rashidzadeh and A. Saffar-Teluri, *Appl. Surf. Sci.*, 2013, **276**, 317.
- 18 W. Deligeer, Y. W. Gao and S. Asuha, *Appl. Surf. Sci.*, 2011, **257**, 3524.
- 19 M. Anbia, S. Asl Hariri and S. N. Ashrafzadeh, *Appl. Surf. Sci.*, 2010, **256**, 3228.
- 20 J. Hu, I. M. C. Lo and G. Chen, *Langmuir*, 2005, **21**, 11173.
- 21 H. F. Liang and Z. C. Wang, *Mater. Chem. Phys.*, 2010, **124**, 964.



- 22 S. Zhang, H. Niu, Y. Cai, X. Zhao and Y. Shi, *Chem. Eng. J.*, 2010, **158**, 599.
- 23 X. Bao, Z. Qiang, W. Ling and J. Chang, *Sep. Purif. Technol.*, 2013, **117**, 104.
- 24 J. G. Parsons, M. L. Lopez, J. R. Peralta-Videa and J. L. Gardea-Torresdey, *Microchem. J.*, 2009, **91**, 100.
- 25 J. Dui, G. Zhu and S. Zhou, *ACS Appl. Mater. Interfaces*, 2013, **5**, 10081.
- 26 I. Ibrahim, I. O. Ali, T. M. Salama, A. A. Bahgat and M. M. Mohamed, *Appl. Catal., B*, 2016, **181**, 389.
- 27 D. Chen, Y. Zhang and Z. Kang, *Chem. Eng. J.*, 2013, **215–216**, 235.
- 28 W. Wang, Z. Ding, M. Cai, H. Jian, Z. Zeng, F. Li and J. P. Liu, *Appl. Surf. Sci.*, 2015, **346**, 348.
- 29 Y. Ren, N. Li, J. Feng, T. Luan, Q. Wen, Z. Li and M. Zhang, *J. Colloid Interface Sci.*, 2012, **367**, 415.
- 30 X. Hou, J. Feng, X. Xu and M. Zhang, *J. Alloys Compd.*, 2010, **491**, 258.
- 31 M. P. Reddy, A. M. A. Mohamed, M. V. Ramana, X. B. Zhou and Q. Huang, *J. Magn. Magn. Mater.*, 2015, **395**, 185.
- 32 X. Hou, J. Feng, Y. Ren, Z. Fan and M. Zhang, *Colloids Surf., A*, 2010, **363**, 1.
- 33 M. Augustin and T. Balu, *Mater. Today*, 2015, **2**, 923.
- 34 C. Pereira, A. M. Pereira, C. Fernandes, M. Rocha, R. Mendes, M. P. Fernández-García, A. Guedes, P. B. Tavares, J.-M. Grenèche, J. P. Araújo and C. Freire, *Chem. Mater.*, 2012, **24**, 1496.
- 35 M. J. Akhta and M. Younas, *Solid State Sci.*, 2012, **14**, 1536.
- 36 N. S. Gajbhiye and G. Balaji, *Thermochim. Acta*, 2002, **385**, 143–151.
- 37 J. Wang, T. Deng, Y. Lin, C. Yang and W. Zhan, *J. Alloys Compd.*, 2008, **450**, 532.
- 38 M. Vadivel, R. R. Babu, K. Sethuraman, K. Ramamurthi and M. Arivanandhan, *J. Magn. Magn. Mater.*, 2014, **362**, 122.
- 39 Y. Xiao, H. Liang, W. Chen and Z. Wang, *Appl. Surf. Sci.*, 2013, **285**, 498.
- 40 X. Fang, J. Xiao, S. Yang, H. He and C. Sun, *Appl. Catal., B*, 2015, **162**, 544.
- 41 T. Huang, S. Mao, J. Yu, Z. Wen, G. Lu and J. Chen, *RSC Adv.*, 2013, **3**, 16657.
- 42 A. Molea, V. Popescu, N. A. Rowson and A. M. Dinescu, *Powder Technol.*, 2014, **253**, 22.
- 43 J. Jaramillo, B. A. Garzón and L. TiradoMejía, *J. Phys.: Conf. Ser.*, 2016, **687**, 12.
- 44 Q. Hu, Y. Liu, X. Gu and Y. Zhao, *Chemosphere*, 2017, **181**, 328.
- 45 S. Zhang, H. Niu, Y. Cai, X. Zhao and Y. Shi, *Chem. Eng. J.*, 2010, **158**, 599.
- 46 M. Iram, C. Guo, Y. Guan, A. Ishfaq and H. Liu, *J. Hazard. Mater.*, 2010, **181**, 1039.
- 47 S. Wang, X. Zhang, L. Pan, F.-M. Zhao, J.-J. Zou, T. Zhang and L. Wang, *Appl. Catal., B*, 2015, **64**, 234.
- 48 L. Wang, J. Li, Y. Wang, L. Zhao and Q. Jiang, *Chem. Eng. J.*, 2012, **181–182**, 72.
- 49 Y.-S. Ho, *J. Hazard. Mater.*, 2006, **136**, 681.
- 50 Y. C. Wong, Y. S. Szeto, W. H. Cheung and G. McKay, *Langmuir*, 2003, **19**, 7888.
- 51 G. A. Saygili, *J. Mol. Liq.*, 2015, **211**, 515.
- 52 L. Wang, J. Li, Y. Wang, L. Zhao and Q. Jiang, *Chem. Eng. J.*, 2012, **181–182**, 72.
- 53 L. Yang, Y. Zhang, X. Liu, X. Jiang, Z. Zhang, T. Zhang and L. Zhang, *Chem. Eng. J.*, 2014, **246**, 88.
- 54 R. Wu and J. Qu, *J. Chem. Technol. Biotechnol.*, 2005, **80**, 20.
- 55 S. Hashemian, *Afr. J. Biotechnol.*, 2010, **9(50)**, 8667.
- 56 G. A. Saygili, *J. Mol. Liq.*, 2015, **211**, 515.
- 57 M. H. Beyki, H. Alijani and Y. Fazli, *J. Mol. Liq.*, 2016, **216**, 6.
- 58 T. A. Khan, I. Ali, V. Vati Singh and S. Sharma, *J. Environ. Prot.*, 2009, **3**, 11.
- 59 W. Konicki, D. Siber and U. Narkiewicz, *Pol. J. Chem. Technol.*, 2017, **19(4)**, 65.
- 60 J. Feng, Y. Hou, Y. Wang, X. Wang, H. Chen and L. Li, *Sci. China: Chem.*, 2016, **59**, 1033.
- 61 T. Soltani and M. H. Entezari, *Chem. Eng. J.*, 2013, **223**, 145.
- 62 W. Yin, S. Hao and H. Cao, *RSC Adv.*, 2017, **7**, 4062.
- 63 M. Mehrali, E. Sadeghinezhad, A. R. Akhiani, S. T. Latibari, H. S. C. Metselaar, A. S. Kherbeet and M. Mehrali, *Powder Technol.*, 2017, **308**, 149.
- 64 S. Kanagesan, S. B. A. Aziz, M. Hashim, I. Ismail, S. Tamilselvan, N. B. B. M. Alitheen, M. K. Swamy and B. P. C. Rao, *Molecules*, 2016, **21**, 312.
- 65 P. Iranmanesh, S. Saeednia, M. Mehran and S. Rashidi Dafeh, *J. Magn. Magn. Mater.*, 2017, **425**, 31.
- 66 D. Zhao, C. Chen, Y. Wang, W. Ma, J. Zhao, T. Rajh and L. Aang, *Environ. Sci. Technol.*, 2008, **42**, 308.
- 67 Q. He, J. Shi, X. Cui, J. Zhao, Y. Chen and J. Zhou, *J. Mater. Chem.*, 2009, **19**, 3395.
- 68 R. M. Dukali, I. M. Radović, D. B. Stojanović, D. M. Šević, V. J. Radojević, D. M. Jocić and R. R. Aleksić, *J. Serb. Chem. Soc.*, 2014, **79(7)**, 867.
- 69 J. Fu, Q. Xin, X. Wu, Z. Chen, Y. Yan, S. Liu, M. Wang and Q. Xu, *J. Colloid Interface Sci.*, 2016, **461**, 292.
- 70 Q. Tang, L. Jiang, J. Liu, S. Wang and G. Sun, *ACS Catal.*, 2014, **4**, 457.
- 71 J. Chen, W. Wen, L. Kong, S. Tian, F. Ding and Y. Xiong, *Ind. Eng. Chem. Res.*, 2014, **56**, 6297.
- 72 Y. Yao, Y. Cai, F. Lu, F. Wei, X. Wang and S. Wang, *J. Hazard. Mater.*, 2014, **270**, 61.
- 73 Z. Zhang, Y. Wang, Q. Tan, Z. Zhong and F. Su, *J. Colloid Interface Sci.*, 2013, **398**, 185.
- 74 J. Zhang, T. Yao, C. Guan, N. Zhang, H. Zhang, X. Zhang and J. Wu, *J. Colloid Interface Sci.*, 2017, **505**, 130.
- 75 L. Chen, F. He, N. Zhao and R. Guo, *Appl. Surf. Sci.*, 2017, **420**, 669.
- 76 F.-T. You, G.-W. Yua, Y. Wang, Z.-J. Xing, X.-J. Liu and J. Li, *Appl. Surf. Sci.*, 2017, **413**, 387.
- 77 Y. J. Oh, J. J. Yoo, Y. I. Kim, J. K. Yoon, H. N. Yoon, J.-H. Kim and S. B. Park, *Electrochim. Acta*, 2014, **116**, 118.
- 78 T. Serizawa, H. Takeshita and M. Akashi, *Langmuir*, 1998, **14**, 4088.
- 79 S. Banerjee and M. C. Chattopadhyaya, *Arabian J. Chem.*, 2017, **10**, S1629.

