Nano metal oxides as efficient catalysts for selective synthesis of 1-methoxy-2-propanol from methanol and propylene oxide†

Jiawei Zhang, Qinghai Cai, Jingxiang Zhao and Shuying Zang*

Nano metal oxides such as Fe₂O₃, Fe₃O₄, CuO, NiO, ZnO and SnO₂ were prepared and characterized using XRD, SEM and TEM analysis. These as-prepared metal oxide materials were used as catalysts for the etherification of methanol with propylene oxide (PO). The results showed that α-Fe₂O₃ exhibited outstanding catalytic performance with 97.7% conversion and 83.0% selectivity to MP-2 at 160 °C for 8 h. Furthermore, the relationship between the catalytic activity or selectivity and surface basicity or energy gap was investigated. This catalyst could be easily recovered and reused due to its heterogeneous catalytic nature.

Introduction

Glycol ethers are considered to be important chemicals that combine the best solvency features of alcohols and ethers for synthesizing organic compounds. They are widely used as industrial solvents for coating materials, printing ink, dyeing leather etc. Due to the low toxicity of propylene glycol ether, it is expected to be a safe substitute for toxic ethylene glycol ether. Although there are several methods for synthesizing propylene glycol ether, the propylene oxide route is the most convenient and industrially feasible in terms of atom-economy and energy-efficiency. In this method, the epoxide ring of propylene oxide may open at the C−O bond and react with alcohols to form propylene glycol ether, catalyzed by acid or base catalysts. These catalysts include earlier homogeneous acids or bases, such as NaOH, alcoholic sodium, H₂SO₄ and BF₃ etc. and later solid acids or bases, e.g. acidic zeolites, Mg/Al hydrotalcite, Zn−Mg-Al oxides, MgO and amine-modified porous silica etc. Very recently, brucite-layer materials, Al₂O₃/MgO, basic and acidic ionic liquids have been developed as highly efficient catalysts for the selective synthesis of propylene glycol ether. Because there is much research interest in developing environmentally benign catalysis, we wish to report the preparation of nano metal oxides using a simple method, and highlight their catalytic performance for the synthesis of 1-methoxymethane-2-propanol (MP-2) via an atom-economic reaction. The simply prepared Fe₂O₃ catalyst, which is inexpensive and environmentally benign, demonstrated excellent catalytic performance for the synthesis of MP-2. The results are disclosed herein.

Experimental

Preparation of the catalysts

Iron oxide was prepared using the following procedure. 30 mL of FeCl₃ aqueous solution (1 mol L⁻¹), 2.5 g of urea and 2 mL of polyethylene glycol (PEG-400) as a dispersant were added to a three-necked flask fitted with a reflux condenser and an electric heater. Ammonia solution (28%) was added dropwise to the mixed solution to maintain a pH of 3.7 to afford Fe(OH)₃. Thereafter, the reactor was heated to 80 °C and aged at this temperature for 8 h. The mixed solution was then ultrasonically treated at room temperature for 20 min, followed by filtration to remove the filtrate. The obtained solid was washed with deionized water and anhydrous ethanol. Thereafter, it was dried at 100 °C for 12 h, and then calcined at 500 °C for 2 h in a muffle furnace to afford nano-iron oxide. Other oxides, such as CuO, ZnO, NiO and SnO₂, were prepared with the same method to that of iron oxide, using CuSO₄, ZnSO₄, NiCl and SnCl₄ aqueous solutions as precursors. However, the pH values for precipitation of Cu(OH)₂, Zn(OH)₂, Ni(OH)₂ and Sn(OH)₄ were adjusted to 4.7, 6.0, 7.2 and 2.5, respectively. Fe₂O₃ was prepared according to the literature.

Catalyst characterization

XRD of the sample was performed on a Bruker-D8 Advance X-ray diffractometer with Cu Kα radiation (40 kV and 36 mA). TEM measurements were taken on a JEM-2100 electron microscope (JEOL Japan) with an acceleration voltage of 200 kV. The
morphology of the catalysts was observed using a Hitachi S-4800 scanning electron microscope (SEM).

Catalytic test and analytic procedure

1.0 mL (14.4 mmol) of propylene oxide, 21 mL (0.52 mol) of methanol and 0.1 g of iron oxide catalyst were added, in turn, into a stainless steel 100 mL autoclave with a magnetic stirrer in an electric heater. The reactor was heated to 160 °C with fast stirring and the reaction was carried out at this temperature for 8 h. When the reaction was complete, the mixture was filtered to remove the catalyst and the filtrate was analyzed using GC (Agilent 7820) with a FID detector equipped with HP-5, a 30 m × 0.32 mm × 0.25 μm column and GC-MS (Agilent GC 7890A-5975C).

Results and discussion

Characterization of the catalysts

Fig. 1 presents the XRD pattern of the prepared iron oxide sample recorded in the 2θ range of 10–90°. It can be seen from the figure that typical diffraction peaks at 2θ values of 24.1, 33.2, 35.6, 41.0, 49.5, 54.2, 57.7, 62.4 and 64.0° correspond to the diffractions of the (012), (104), (110), (024), (116), (018), (214) and (300) planes, respectively. The intense peaks of the XRD pattern indicate that the Fe₂O₃ product was well crystallized and no impurity diffraction peaks were observed, implying that almost single phase Fe₂O₃ was successfully prepared. All diffraction peaks were perfectly indexed, which was in agreement with the data of α-Fe₂O₃ (hematite, JCPDS 33-0664). Moreover, the average crystallite size of the product was estimated to be 31.7 nm using the Debye–Scherrer equation. The XRD patterns of the other metal oxides, such as CuO, NiO, ZnO, SnO₂ and Fe₃O₄, are shown in Fig. SI1–5.† SEM and TEM. Fig. 2 shows the scanning electron microscopy (SEM) image of the Fe₂O₃ sample. As seen from the figure, iron oxide particles of about 30 nm in size were well dispersed and their sizes were in agreement with the value calculated using the Debye–Scherrer equation. In addition, the transmission electron microscopy (TEM) image of the Fe₂O₃ sample also showed the same result as the SEM of Fe₂O₃. Nano-particles of about 30 nm in size are shown in Fig. 3. The SEM and/or TEM images of other metal oxides, such as CuO, NiO, ZnO, SnO₂ and Fe₃O₄ (TEM), are shown in Fig. SI6–10.† The morphologies of these oxides can also be seen in Fig. SI6–10:† particles of about 20 nm for Fe₃O₄, particles of 200–300 nm for CuO, particles of about 100 nm for ZnO, and aggregations composed of ~30 nm sized particles for SnO₂ and NiO.

Catalytic activity. The catalytic activity of α-Fe₂O₃ and the other metal oxides such as Fe₃O₄, CuO, NiO, ZnO and SnO₂ for the etherification of methanol with propylene oxide (PO) was assessed. The results are summarized in Table 1. The data show that the α-Fe₂O₃ catalyst possessed very high activity for etherification with a conversion of 97.7% and 83.0% selectivity to 1-methoxy-2-propanol (MP-2) at 160 °C for 8 h, corresponding to a TON (turnover number, converted moles per mole catalyst) of 11.5 mol mol⁻¹Fe⁻¹ (entry 1). Also, the CuO catalyst displayed the same activity as α-Fe₂O₃, with a conversion of 97.9% (entry 2), but the selectivity to MP-2 was relatively low with a MP-2/MP-1 (2-methoxy-1-propanol) ratio of only 1.17. Although Fe₃O₄ and NiO had a higher MP-2/MP-1 ratio, their catalytic activities for the etherification were lower (entry 3 and 4), exhibiting 26.5 and 56.3% conversions with TONs of 3.0 mol mol⁻¹Fe⁻¹ and 6.1 mol

Fig. 1 XRD pattern of Fe₂O₃.

Fig. 2 SEM image of iron oxide sample.

Fig. 3 TEM image of the iron oxide sample.
Catalytic activity of some metal oxides

<table>
<thead>
<tr>
<th>Entry</th>
<th>Oxides</th>
<th>Conv. (%)</th>
<th>Sel. (%)</th>
<th>TON</th>
<th>Rat.</th>
<th>$E_g$</th>
<th>Bas.</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>Fe$_3$O$_4$</td>
<td>26.5</td>
<td>80.4</td>
<td>3.0</td>
<td>4.10</td>
<td>—</td>
<td>—</td>
</tr>
<tr>
<td>2</td>
<td>NiO</td>
<td>56.3</td>
<td>80.5</td>
<td>6.1</td>
<td>4.13</td>
<td>3.80</td>
<td>0.997</td>
</tr>
<tr>
<td>3</td>
<td>ZnO</td>
<td>63.9</td>
<td>59.3</td>
<td>7.7</td>
<td>1.46</td>
<td>3.40</td>
<td>1.001</td>
</tr>
<tr>
<td>4</td>
<td>SnO$_2$</td>
<td>59.7</td>
<td>59.0</td>
<td>13.0</td>
<td>1.44</td>
<td>3.80</td>
<td>0.927</td>
</tr>
</tbody>
</table>

Reaction conditions: reaction temperature = 110 °C, catalyst = 0.1 g, molar ratio of methanol to PO = 36:1; Conv. = conversion, Sel. = selectivity to MP-2, TON = turn over number (mol mol$_{-1}$M$^{-1}$), rat. = ratio of MP-2/MP-1, $E_g$ = energy gap, Bas. = basicity.

The thermodynamic equilibrium of the reaction was almost achieved at that time. Furthermore, the selectivity to MP-2 increased as the time was prolonged from 3 to 5 h, while on the contrary, the selectivity to MP-1 was decreased over this range. The selectivity to MP-2, as well as MP-1, was almost constant during the reaction time from 5 to 9 h.

As the reaction temperature was raised, the conversion rapidly increased to a maximum at 160 °C, and basically remained unchanged at 170 °C (Fig. 5). The change in the conversion with temperature was probably due to the control of the reaction kinetics, that is, high temperature can accelerate the reaction rate and shorten the time needed to reach equilibrium. Thus, increasing the reaction temperature is greatly advantageous to forming the product. Unexpectedly, the selectivity to MP-2 was independent of the reaction temperature in the range from 110 to 190 °C, retaining the value at 83.0% over the whole temperature range. This is rare in the catalytic synthesis of organic compounds.

Reactivity of other alcohols

Other alcohols such as ethanol, propanol, butanol, iso-butanol, amyl alcohol, iso-amyl alcohol and octanol were then applied as alternatives to methanol, and the etherification reactions of PO

Effect of reaction conditions

Using Fe$_3$O$_4$ as the catalyst, the effect of reaction time on the catalytic activity was investigated. As shown in Fig. 4, the conversion of PO increased as the reaction proceeded from 3 to 8 h at 110 °C. The conversion approached to 57.6% as the reaction was carried out for 8 h. Then, the conversion slightly increased as the reaction time was prolonged to 9 h, implying that the thermodynamic equilibrium of the reaction was almost achieved at that time. Furthermore, the selectivity to MP-2 increased as the time was prolonged from 3 to 5 h, while on the contrary, the selectivity to MP-1 was decreased over this range. The selectivity to MP-2, as well as MP-1, was almost constant during the reaction time from 5 to 9 h.

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with these alcohols were also examined in the presence of α-Fe₂O₃ as a catalyst under the same conditions (Table 2). As shown in the table, the reactivity was reduced as the length of the carbon chain in the alcohols increased (entry 1–8), so that octanol exhibited very low reactivity for this reaction with 31.6% conversion (entry 8). By comparison, the lower reactivity of iso-butanol and iso-amyl alcohol compared to that of butanol and amyl alcohol is related to the steric effect of secondary alcohols (entry 4–7). Furthermore, the selectivity to 1-alkoxy-2-alcohol was also lower than that of the methanol system, with 48.0 to 64.2% selectivity achieved.

### Table 2 Reactivity of other alcohols with PO

<table>
<thead>
<tr>
<th>Entry</th>
<th>Alcohols</th>
<th>Conv. (%)</th>
<th>Sel. to MP-2 (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>Methanol</td>
<td>97.7</td>
<td>82.6</td>
</tr>
<tr>
<td>2</td>
<td>Ethanol</td>
<td>69.8</td>
<td>50.9</td>
</tr>
<tr>
<td>3</td>
<td>Propanol</td>
<td>64.7</td>
<td>49.6</td>
</tr>
<tr>
<td>4</td>
<td>Butanol</td>
<td>63.9</td>
<td>64.2</td>
</tr>
<tr>
<td>5</td>
<td>Iso-butanol</td>
<td>50.2</td>
<td>52.1</td>
</tr>
<tr>
<td>6</td>
<td>Amyl alcohol</td>
<td>51.6</td>
<td>50.0</td>
</tr>
<tr>
<td>7</td>
<td>Iso-amyl alcohol</td>
<td>41.6</td>
<td>48.0</td>
</tr>
<tr>
<td>8</td>
<td>Octanol</td>
<td>31.6</td>
<td>50.0</td>
</tr>
</tbody>
</table>

The stability of the catalyst

In order to further evaluate the performance of the α-Fe₂O₃ catalyst, reuse tests of α-Fe₂O₃ were conducted. The catalyst was separated by filtration after the first test, and then reused for the next run under the same conditions. The results are depicted in Fig. 6.

As shown in the figure, the catalytic activity of the reused catalyst was almost unaffected even at the fourth run, exhibiting >95.0% conversion and retaining 83.0% selectivity to MP-2. The SEM image of the reused catalyst is shown in Fig. 7. Aggregation of the Fe₂O₃ nanoparticles was observed when they were reused four times, but their phase composition did not change, as shown in the XRD pattern of the used Fe₂O₃ catalyst (Fig. 8). These results imply that the catalyst can be efficiently recovered and recycled.

### Conclusions

Nano metal oxide materials were prepared and characterized using XRD, SEM and TEM analysis. These oxides were used for the etherification of methanol with propylene oxide (PO) to form 1-methoxy-2-propanol (MP-2). The as-prepared Fe₂O₃ material, recognized as α-Fe₂O₃, was found to be an effective catalyst for the etherification of methanol with propylene oxide (PO), with 97.7% conversion and 83.0% selectivity to MP-2 at 160 °C for 8 h, and achieving a TON of 11.5 mol molFe⁻¹. The high catalytic activity and selectivity to MP-2 of this catalyst, which was prepared using a simple method, can be attributed to its low energy gap and high basicity. This catalyst can be easily recovered and reused due to its heterogeneous catalytic nature. Thus, the simply prepared and environmentally benign α-Fe₂O₃ catalyst demonstrates great potential for application in industry.

### Conflicts of interest

There are no conflicts to declare.

### Acknowledgements

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Notes and references