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UV-mediated hydrophosphinylation of unactivated alkenes with phosphinates under batch and flow conditions†

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A UV-mediated hydrophosphinylation of unactivated alkenes with *H*-phosphinates and hypophosphorous acid under radical free conditions is presented. The reaction affords selectively a large number of structurally diverse organophosphorous compounds in moderate to good yields under mild reaction conditions in the presence of an organic sensitizer as catalyst irradiated by UV-A LEDs. Furthermore, the high yielding hydrophosphinylation in continuous flow is disclosed.

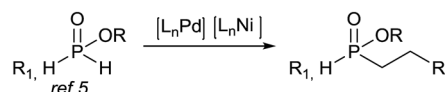
Introduction

Organophosphorus compounds have attracted much attention due to their wide range of applications in materials, catalysis, natural bioactive products and pharmaceuticals.¹ In particular, phosphinates (P(O)(OR)R¹R², R¹/R² = hydrogen and/or carbon) have been the target of numerous synthetic efforts as they are versatile precursors to organophosphorus compounds and represent a sustainable alternative to the use of phosphorous trichloride.² Indeed, phosphinates can be considered as the synthetic equivalents of phosphorous trichloride with several advantages in terms of stability and toxicity. In this regard, the development of efficient and selective methodologies for the preparation and functionalization of phosphinates (P(O)(OR)H₂) and *H*-phosphinates (P(O)(OR)R¹H) has been extensively studied, mostly by the group of Montchamp.^{2,3} Hydrophosphinylation of unactivated alkenes is one of these routes that provide a direct access to functionalized organophosphorous compounds.⁴ This atom economical process has been reported to proceed *via* transition metal catalysis,⁵ or radical activation (Scheme 1a and b).⁶ Despite the significant progress in this chemistry, these transformations suffer from some drawbacks like the relatively harsh conditions, narrow substrates scopes and costly catalytic system thus limiting their chemical and pharmaceutical applications. Recently, Lakhdar and co-workers reported a hydrophosphinylation of unactivated alkenes with ethyl and butyl phosphinates under photocatalytic conditions (Scheme 1c),⁷ even though diphenyl iodonium triflate which acted as sacrificial oxidant is not a heavy metal, the use of catalytic quantity is

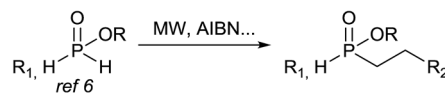
always preferred in term of sustainability and environmental benignity. Alternatively, pioneering work from the group of Dondoni⁸ demonstrated that P(O)-centered radical can be generated from *H*-phosphonates under UV-A irradiation, in the presence of 2,2-dimethoxy-2-phenylacetophenone (DMPA) as photoinitiator, and added to alkene functionalized carbohydrates. However, the reaction requires the use of a large excess of phosphonates (100 eq.). Mathé and coworkers⁹ then extended this free radical hydrophosphonylation to activated and unactivated alkenes, similar reaction with hypophosphorous acid and *H*-phosphinates derivatives¹⁰ have not been studied previously.

In this context, we envisaged that continuous-flow systems in combination with a sensitizer irradiated by UV-A LEDs ($\lambda = 365$

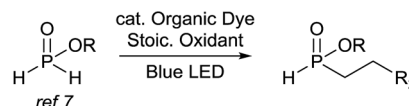
a) Transition metal catalysis



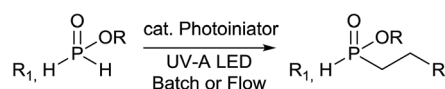
b) Radical activation



c) Visible light photocatalysis



This work



Scheme 1 Hydrophosphinylation of phosphinates with unactivated alkenes.

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nm) would result in a significant enhancement of the reaction. Over the last decade, interest has grown for flow chemistry based on microfluidic technology in particular towards large-scale application due to its significant improvement over traditional batch reactors concerning reduced consumption of chemicals, solvents and time together with enhanced yields, selectivity and control over reaction conditions. An easy scale up is also one of the characteristic of photo-reactions which are conducted in micro-reactors. In batch reactors, light penetration through the reaction media is limited which restrains the efficiency of photochemical processes. This drawback can be overcome with continuous microflow reactors since their small optical lengths improve sample irradiation and also enhance heat and mass transfer.¹¹ Therefore, in view of the growing demand to develop efficient, mild and sustainable methods to access phosphinates and as a continuation of our interest in photocatalyzed processes,¹² we herein report an original UV-mediated hydrophosphinylation of unactivated alkenes with hypophosphorous acid and less reactive *H*-phosphinates derivatives under batch and flow conditions.

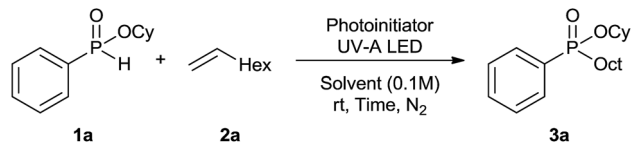
Results and discussion

Hydrophosphinylation of unactivated alkenes with *H*-phosphinates under batch conditions

Initially, the hydrophosphinylation of relatively challenging *H*-phosphinate (**1a**) with octene (**2a**) in equimolar amounts in the presence of a photoinitiator under UV-A irradiation was selected as a model to investigate the reaction. The use of one equivalent of 4,4'-dimethoxybenzophenone (4,4'-DMBP) was found to be effective furnishing **3a** after stirring in acetonitrile under an ambient inert atmosphere for 3 h (Table 1, entry 1, 42%). An investigation of solvents showed that the reaction efficiency was further enhanced when performed in DMSO (Table 1, entry 4, 69%) or acetic acid^{6b,13} (entry 5, 70%) as previously observed for radical process with *H*-phosphinates, whereas moderate yields were obtained with ethyl acetate (entry 2, 46%) or DMF (entry 3, 52%). Importantly a very low catalyst loading could promote the reaction yield up to 84% (entries 6 and 7). Moreover, the loading could be further decreased but a prolonged reaction time was required (Table 1, entry 8, 86%). Interestingly, the photoinitiator DMPA (2,2-dimethoxy-2-phenylacetophenone) was not efficient in our case since the phosphorylated product was formed in a moderate yield (Table 1, entry 15, 32%). Furthermore, none of the common photoinitiators (thioxantone, 4-methoxyacetophenone, benzophenone) exhibited a better catalytic effect (Table 1, entries 12–14). Control experiments revealed that the photoinitiator, inert conditions and light were essential for this reaction (Table 1, entries 9–11). It is noteworthy that using a 4,4'-DMBP derivative functionalized with an ionic liquid moiety¹⁴ (4,4'-(2-(1-methylimidazolium)ethoxy)benzophenone dibromide, DIBP) as photosensitizer led to the same conversions as previously observed furnishing an excellent way to simplify the purification procedure by simple washing of the reaction media thus improving the synthetic efficiency by avoiding chromatography (Table 1, entry 16).

The generality of the method was investigated, with the scope of the reaction being explored with respect to the *H*-

Table 1 Optimization of the reaction conditions^a



Entry	Photoinitiator	Equivalent	Solvent	Time (h)	Yield ^b (%)
1	4,4'-DMBP ^c	1	MeCN	3	42
2	4,4'-DMBP	1	EtOAc	3	46
3	4,4'-DMBP	1	DMF	3	52
4	4,4'-DMBP	1	DMSO	3	69
5	4,4'-DMBP	1	AcOH	3	70
6	4,4'-DMBP	0.5	DMSO	3	72
7	4,4'-DMBP	0.1	DMSO	5	84
8	4,4'-DMBP	0.05	DMSO	17	86
9	—	—	DMSO	5	10
10 ^d	4,4'-DMBP	0.1	DMSO	5	—
11 ^e	4,4'-DMBP	0.1	DMSO	5	5
12	Thioxantone	0.1	DMSO	5	76
13	4-MAP ^f	0.1	DMSO	5	19
14	Benzophenone	0.1	DMSO	5	77
15	DMPA ^g	0.1	DMSO	5	32
16	DIBP ^h	0.1	DMSO	16	79 (76) ⁱ

^a Reaction condition: **1a** (0.3 mmol, 1 equiv.), **2a** (0.3 mmol, 1 equiv.), photoinitiator (*x* equiv.), solvent ([0.1 M]), under nitrogen and UV-A LED irradiation ($\lambda = 365 \pm 15$ nm, 230 mW cm⁻²) at room temperature. ^b Derived from ³¹P crude NMR spectra on integration of all formed species. ^c 4,4'-DMBP = 4,4'-dimethoxybenzophenone. ^d Without irradiation. ^e Under ambient atmosphere. ^f 4-MAP = 4-methoxyacetophenone. ^g DMPA = 2,2-dimethoxy-2-phenylacetophenone. ^h DIBP = 4,4'-(2-(1-methylimidazolium)ethoxy)benzophenone dibromide. ⁱ The isolated yield is shown in parentheses.

phosphinate component, and the results are summarized in Table 2. Various P(O)–H compounds were employed to react with 1-octene **2a** giving the corresponding **3a–g** with isolated yields ranging from 30 to 76% (Table 2, entries 1–7). As expected, in all of the cases the anti Markovnikov product was observed. The reaction is not limited to *H*-phosphinate esters as *H*-phosphinate acid **1b** could also deliver the desired product with a good yield (entry 2, 64%) in contrast to *H*-phosphinate acid **1c** whose corresponding phosphinate acid product **3c** was difficult to obtain in high purity (entry 3, 30%). However, *H*-phosphinate including unactivated alkyl and benzyl derivatives reacted successfully regardless of the ester chosen (Table 2, entries 4–6). When using (hydroxymethyl)-*H*-phosphinate ester, which was reported to have several application,¹⁵ the branched product **3g** was obtained with a significant yield of 75% (Table 2, entry 7). To further evaluate the substrate scope, a series of alkenes were tested with *H*-phosphinate **1a** (Table 2, entries 8–12). The reaction with allyl benzene gave the corresponding product **3h** with 48% of yield (Table 2, entry 8). Interestingly, **1a** reacted with hindered alkene giving **3i** with a modest yield of 27% (Table 2, entry 9). Note that halogens like Br and alcohol groups were well tolerated which indicates a potential for further functionalization (Table 2, entries 10–12).



Table 2 Hydrophosphinylation of alkenes with *H*-phosphinates^a 1

Entry	<i>H</i> -Phosphinate 1	Alkene 2	Product 3	Yield (%)
1		1-Octene 2a		76
2		1-Octene 2a		64
3		1-Octene 2a		30 ^b
4		1-Octene 2a		73
5		1-Octene 2a		71
6 ^c		1-Octene 2a		51
7		1-Octene 2a		75
8				48
9 ^c				27
10				38
11 ^c				56
12 ^c				69

^a Reaction condition: all reactions unless specified were carried out using **1** (0.3 mmol, 1 equiv.), **2** (0.3 mmol, 1 eq.), DIBP (0.1 equiv.) in DMSO ([0.1 M]) under nitrogen and UV-A LED irradiation ($\lambda = 365 \pm 15$ nm, 230 mW cm⁻²) at room temperature for 16 h, isolated yield.

^b Derived from ³¹P crude NMR spectra on integration of all formed species. ^c (0.5 equiv.) of DIBP.

Although the detailed reaction mechanism is unclear, based on literature precedents^{64,66} all these results could be well explained by a radical chain mechanism as depicted in Scheme 2. Upon irradiation, the excited state of the photoinitiator¹⁷ abstracts a proton from *H*-phosphinates to form the phosphoryl radical¹⁸ which adds to the terminal carbon of the alkenes. The carbon centered radical can then abstract a proton from *H*-phosphinate **1** to regenerate the phosphoryl radical.

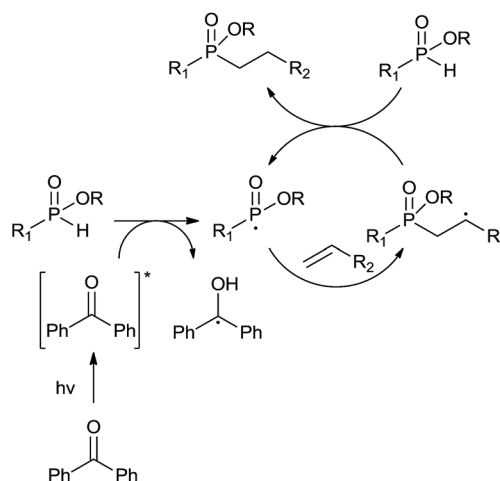
Hydrophosphinylation of unactivated alkenes with *H*-phosphinates under flow conditions

With the successful results for the hydrophosphinylation under batch conditions, the reaction was then performed under continuous flow to further improve its efficiency in a shorter amount of time. The continuous flow hydrophosphinylation was performed using a continuous flow microfluidic system composed of a high pressure syringe pump delivering the homogeneous reaction mixture at specific flow rates to a commercially available microreactor (Mikroglas Dwell Device® microreactor from Invenios Europe, Langen, Germany) irradiated by HP UV-A lamp. This microreactor is made up of Foturan® glass that is transparent up to 300 nm allowing to work at a wide range of wavelengths (UV-A & visible).

We were delighted to find that *H*-phosphinate derivatives (**3a**, **b**, **k**) could be obtained with 4,4'-DMBP as photoinitiator within a shorter reaction time than in batch (30 min vs. 5 h), indicating the importance of the short light path length provided by the microreactor, with slightly higher isolated yield, highlighting the potentials of this process (Scheme 3). This results offer the possibility to conduct the reaction on large scale without erosion of the yields.

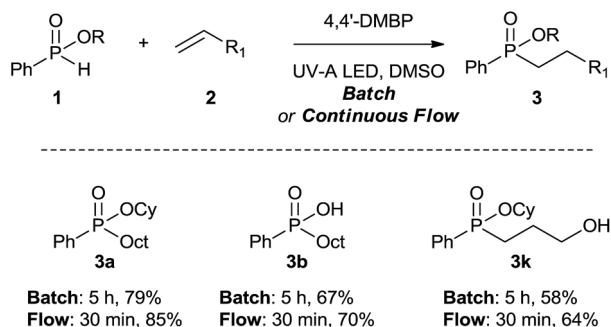
Hydrophosphinylation of unactivated alkenes with hydrophosphorous acid under batch conditions

To further demonstrate the utility of our protocol, we sought to test its potential for the hydrophosphinylation of various



Scheme 2 Proposed reaction mechanism for the photoinduced hydrophosphinylation of unactivated alkenes with *H*-phosphinates.

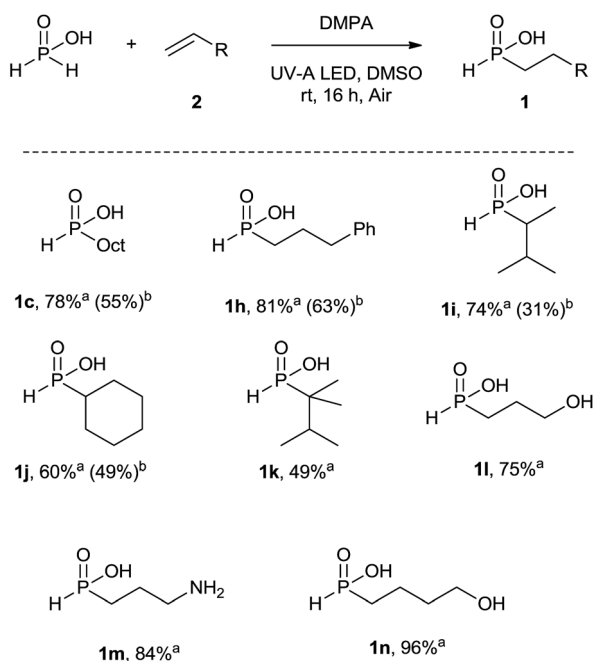




Scheme 3 Hydrophosphinylation of alkenes with *H*-phosphinates **1** under batch and flow conditions. (a) Reactions conditions in batch: **1** (0.3 mmol, 1 equiv.), **2** (0.3 mmol, 1 eq.), 4,4'-DMPB (0.1 equiv.) in DMSO ([0.1 M]) under nitrogen and UV-A LED irradiation ($\lambda = 365 \pm 15$ nm, 230 mW cm^{-2}) at room temperature, 5 h, isolated yield. (b) Reaction conditions in continuous flow: **1** (0.3 mmol, 1 equiv.), **2** (0.3 mmol, 1 equiv.), 4,4'-DMPB (0.1 equiv.) in DMSO ([0.1 M]) under nitrogen and UV-A LED irradiation ($\lambda = 365 \pm 15$ nm, 230 mW cm^{-2}) with Dwell device manufactured by Mikroglass Chemtech Mainz, Germany with a rectangular shape of dimensions 115 mm \times 2 mm \times 0.5 mm as photo-microreactor, at room temperature for 30 min of residence time, isolated yield.

alkenes with hypophosphorous acid to form the previous precursor *H*-phosphinate acid (Scheme 4).

A slight modification of the previous batch conditions (see Table 1, ESI†) allowed us to obtain the corresponding *H*-phosphinates **1** with ^{31}P NMR yield ranging from 49 to 96%. A



Scheme 4 Hydrophosphinylation of alkenes with hypophosphorous acid. All reactions (unless otherwise specified) were carried out using H_3PO_2 (0.3 mmol, 2 equiv.), **2** (0.15 mmol, 1 equiv.), DMPA (0.2 equiv.) in DMSO ([0.1 M]) under UV-A LED irradiation ($\lambda = 365 \pm 15$ nm, 230 mW cm^{-2}) at room temperature for 16 h. ^aDerived from ^{31}P crude NMR spectra on integration of all formed species. ^bIsolated yield.

2 : 1 molar ratio of hypophosphorous acid–alkenes was necessary to obtain a high conversion without the formation of the disubstituted by-product. However despite the observed high conversions, isolated yields were modest due to difficult purifications. In all of the cases, the reaction was chemo-selective where only mono substituted phosphinates were observed. As for *H*-phosphinate, the reaction proceeded well with terminal (**1c–h**) and cyclic alkenes (**1j**). Hindered alkenes delivered the desired compounds like **1i** with a moderate yield of 31%, which is still higher than that obtained with *H*-phosphinate ester. *H*-Phosphinate acid **1k** however was difficult to obtain in high purity. Similar to the first reaction assessed, the reaction tolerates functional groups such as amines or alcohols (**1l–n**) but the obtained *H*-phosphinate acids were not isolated. Although the scope of this reaction seems to be limited, *H*-phosphinate acids could be easily esterified¹⁹ *in situ* and therefore the resulting procedure offers significant synthetic advantages.

Experimental section

General information

All reagents were purchased from commercial suppliers (Strem Chemicals Inc., Sigma-Aldrich or Alfa Aesar) and were used without further purification. Thin-layer chromatography (TLC) was performed on Silica gel 60 F254 plates (Merck) and visualized under UV (254 nm) or by staining with potassium permanganate or phosphomolybdic acid. Column chromatography was performed with 63–200 mesh silica gel. NMR spectra were recorded on a Bruker AVANCE 300 spectrometer at 300 MHz (75 MHz). Chemical shifts are reported in parts per million relative to solvent signal and coupling constants are reported in hertz (Hz). High-resolution mass spectra (HRMS) were performed on a Thermo LTQ Orbitrap mass spectrometer using nanoESI ionization. *H*-Phosphinates derivatives **1** and 4,4'-(2-(1-methylimidazolium)ethoxy)benzophenone dibromide (DIBP)¹⁴ were prepared according to the reported procedure. The illumination was performed by UV-A LEDs (365 nm, irradiance = 230 mW cm^{-2}) Omnicure® AC475 model from Lumen Dynamics (Excelitas Technologies, Waltham, MA, USA). Note that the irradiance was measured at the surface of the reactor using a radiometer.

General procedure A for the hydrophosphinylation with *H*-phosphinate **1** derivatives under batch conditions

A solution of a selected *H*-phosphinate **1** (0.3 mmol, 1 equiv.), an alkene **2** (0.3 mmol, 1 equiv.) and DIPB (0.03–0.15 mmol, 10–50 mol%) in degassed DMSO (1.5 mL) was irradiated under 365 nm and N_2 for 16 h. Ethyl acetate (30 mL) was added and the reaction mixture was washed with saturated solution of NaHCO_3 (2×15 mL) and brine (2×15 mL). The organic layer was dried over MgSO_4 , filtered and concentrated under reduce pressure. The residue was then purified by flash chromatography on silica gel if necessary to afford the corresponding organophosphorous compound.



General procedure B for the hydrophosphinylation with hypophosphorous acid

To a 50% aqueous solution of hypophosphorous acid (79.2 mg, 0.6 mmol, 2 equiv.) in DMSO (3 mL) was added an alkene **2** (0.3 mmol, 1 equiv.), 4,4'-DMPA (15.5 mg, 0.06 mmol, 20 mol%). The reaction mixture was irradiated under 365 nm and ambient atmosphere for 16 h. Ethyl acetate (30 mL) was added and the organic layer was washed with brine (3 × 15 mL). To the organic layer a 0.03 M solution of NaHCO₃ (30 mL, 1 mmol) was added. The aqueous layer was washed with ethyl acetate (2 × 15 mL), acidified with 1 M HCl (2 mL, 2 mmol), saturated with NaCl and extracted with ethyl acetate (30 mL). The last ethyl acetate layer was dried over Na₂SO₄, filtered and concentrated under reduce pressure. The residue was then purified by flash chromatography on silica gel using 9 : 1 : 0.5 DCM–MeOH–AcOH as the eluant to afford the desired pure *H*-phosphinate acid **1**.

Hydrophosphinylation of cyclohexyl phenyl-*H*-phosphinate **1a** under flow condition

A solution of a cyclohexyl phenyl-*H*-phosphinate²⁰ **1a** (67.2 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and 10 mol% of 4,4'-dimethoxybenzophenone (7.2 mg) in degazed DMSO (1.5 mL) was pumped through the Mikroglas Dwell Device reactor ($V_{int} = 1.15$ mL) at 38.33 $\mu\text{L min}^{-1}$, residence time of 30 min and irradiated by UV-A LEDs. Ethyl acetate (30 mL) was added to the collected reaction mixture and the reaction mixture was washed with saturated solution of NaHCO₃ (2 × 15 mL) and brine (2 × 15 mL). To the organic layer was dried over MgSO₄, filtered and the solvent was removed under reduce pressure to afford **3a** (85.6 mg, 0.25 mmol, 85%) as a colorless oil.

Cyclohexyl octyl(phenyl)phosphinate (3a). Following the general procedure A, using cyclohexyl phenyl-*H*-phosphinate **1a** (67.2 mg, 0.3 mmol, 1 equiv.), 1-octene (33.6 mg, 0.3 mmol, 1 equiv.) and 10 mol% of catalyst (17.8 mg) affords **3a** (76.6 mg, 0.23 mmol, 76%) as a colorless oil. ³¹P NMR (121 MHz, CDCl₃): $\delta = 43.5$ (s); ¹H NMR (300 MHz, CDCl₃): $\delta = 7.81$ – 7.72 (m, 2H), 7.55 – 7.39 (m, 3H), 4.30 – 4.16 (m, 1H), 2.06 – 2.11 (m, 24H), 0.83 (t, $J = 6.8$ Hz, 3H); ¹³C NMR (75 MHz, CDCl₃): $\delta = 132.3$ (d, $J_{PC} = 122$ Hz), 132.0 (d, $J_{PCCC} = 2.7$ Hz), 131.7 (d, $J_{PCCC} = 9.7$ Hz, 2C), 128.5 (d, $J_{PCC} = 12.2$ Hz, 2C), 74.3 (d, $J_{POC} = 6.8$ Hz), 34.4 (d, $J_{POCC} = 2.9$ Hz), 33.8 (d, $J_{POCC} = 4.3$ Hz), 31.9 , 31.0 (d, $J_{PCC} = 16.0$ Hz), 30.4 (d, $J_{PC} = 101$ Hz), 29.1 (2C), 25.3 , 23.8 , 23.8 , 22.7 , 21.8 (d, $J_{PCCC} = 3.7$ Hz), 14.2 . HRMS (ESI) m/z calcd for C₂₀H₃₄O₂P ([M + H]⁺) 337.2290, found 337.2289.

Octyl phenylphosphinic acid (3b). A solution of phenyl phosphinic acid **1b** (47.4 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and DIPB (6.7 mg, 0.03 mmol, 10 mol%) in degazed DMSO (3 mL) under N₂ was irradiated under 365 nm for 15 h. Ethyl acetate (30 mL) was added and the reaction mixture was washed with brine (3 × 15 mL). To the organic layer was added a 0.03 M solution of sodium bicarbonate (30 mL, 1 mmol). The aqueous layer was washed with ethyl acetate (2 × 15 mL), acidified with 1 M HCl (2 mL, 2 mmol) and extracted with ethyl acetate (30 mL). The last ethyl acetate layer was dried over Na₂SO₄, filtered and the solvent was

removed under reduce pressure to afford the desired pure compound **3b** (51.8 mg, 0.19 mmol, 64%). ³¹P NMR (121 MHz, CDCl₃): $\delta = 45.8$ (s); ¹H NMR (300 MHz, CDCl₃): $\delta = 11.65$ (br s, 1H), 7.72 – 7.58 (m, 2H), 7.45 – 7.24 (m, 3H), 1.83 – 1.64 (m, 2H), 1.46 – 1.26 (m, 2H), 1.24 – 1.03 (m, 10H), 0.78 (t, $J = 6.7$ Hz, 3H); ¹³C NMR (75 MHz, CDCl₃): $\delta = 132.5$ (d, $J_{PC} = 127$ Hz), 131.9 , 131.2 (d, $J_{PCCC} = 9.4$ Hz, 2C), 128.4 (d, $J_{PCC} = 11.4$ Hz, 2C), 31.9 , 30.8 (d, $J_{PCC} = 16.3$ Hz), 30.6 (d, $J_{PC} = 98.7$ Hz), 29.2 (2C), 22.7 , 21.8 (d, $J_{PCCC} = 2.6$ Hz), 14.2 ; HRMS (ESI) m/z calcd for C₁₄H₂₄O₂P ([M + H]⁺) 255.1508, found 255.1507.

Octyl octylphosphinic acid (3c). A solution of Octyl phosphinic acid^{5b} **1c** (53.5 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and DIPB (6.7 mg, 0.03 mmol, 10 mol%) in degazed DMSO (3 mL) was irradiated under 365 nm and argon for 16 h. Ethyl acetate (30 mL) was added and the reaction mixture was washed with brine (3 × 15 mL). To the organic layer was added a 0.03 M solution of sodium bicarbonate (30 mL, 1 mmol). The aqueous layer was washed with ethyl acetate (2 × 15 mL), acidified with 1 M HCl (2 mL, 2 mmol) and extracted with ethyl acetate (30 mL). The last ethyl acetate layer was dried over Na₂SO₄, filtered and the solvent was removed under reduce pressure to afford compound **3c** with impurities (51.8 mg, 0.19 mmol, derived from ³¹P NMR spectra on integration of all formed species). All attempts to purify the compound were unfruitful. ³¹P NMR (121 MHz, CDCl₃): $\delta = 56.5$ (s).

Cyclohexyl dioctylphosphinate (3d). Following the general procedure A, using cyclohexyl octyl-*H*-phosphinate²⁰ **1d** (78 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and 10 mol% of catalyst (17.8 mg) affords **3d** (81.2 mg, 0.22 mmol, 73%) as a colorless oil. ³¹P NMR (121 MHz, CDCl₃): $\delta = 56.5$ (s); ¹H NMR (300 MHz, CDCl₃): $\delta = 4.41$ – 4.27 (m, 1H), 1.92 – 1.81 (m, 2H), 1.74 – 1.41 (m, 12H), 1.40 – 1.17 (m, 24H), 0.85 (t, $J = 6.7$ Hz, 6H); ¹³C NMR (75 MHz, CDCl₃): $\delta = 73.2$ (d, $J_{POC} = 6.8$ Hz), 34.4 (d, $J_{POCC} = 4.3$ Hz, 2C), 31.9 (2C), 31.0 (d, $J_{PCC} = 15.1$ Hz, 2C), 29.2 (2C), 29.2 (2C), 28.9 (d, $J_{PC} = 89.9$ Hz, 2C), 25.4 , 23.9 (2C), 22.7 (2C), 22.1 (d, $J_{PCCC} = 3.9$ Hz, 2C), 14.2 (2C); HRMS (ESI) m/z calcd for C₂₂H₄₆O₂P ([M + H]⁺) 373.3229, found 373.3227.

Butyl dioctylphosphinate (3e). Following the general procedure A, using butyl octyl-*H*-phosphinate^{5d} **1e** (70.3 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and 10 mol% of catalyst (17.8 mg) affords **3e** (74 mg, 0.21 mmol, 71%) as a colorless oil. ³¹P NMR (121 MHz, CDCl₃): $\delta = 57.6$ (s); ¹H NMR (300 MHz, CDCl₃): $\delta = 3.92$ (q, $J = 6.7$ Hz, 2H), 1.71 – 1.19 (m, 32H), 0.89 (t, $J = 7.4$ Hz, 3H), 0.84 (t, $J = 6.7$ Hz, 6H); ¹³C NMR (75 MHz, CDCl₃): $\delta = 63.7$ (d, $J_{POC} = 6.9$ Hz), 32.9 (d, $J_{POCC} = 5.8$ Hz), 31.9 (2C), 31.0 (d, $J_{PCC} = 15.0$ Hz, 2C), 29.2 (2C), 29.1 (2C), 28.0 (d, $J_{PC} = 89.6$ Hz, 2C), 22.7 (2C), 22.0 (d, $J_{PCCC} = 3.9$ Hz, 2C), 18.9 , 14.1 (2C), 13.7 . HRMS (ESI) m/z calcd for C₂₀H₄₄O₂P ([M + H]⁺) 347.3073, found 347.3073.

Butyl benzyl(octyl)phosphinate (3f). Following the general procedure A, using butyl benzyl-*H*-phosphinate^{4a} **1f** (63.6 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and 50 mol% of catalyst (89 mg) affords **3f** (49.6 mg, 0.15 mmol, 51%) as a colorless oil. ³¹P NMR (121 MHz, CDCl₃): $\delta = 52.9$ (s); ¹H NMR (300 MHz, CDCl₃): $\delta = 7.28$ – 7.16 (m, 5H), 3.96 – 3.76 (m,



2H), 3.06 (d, $J_{\text{PCH}} = 16.7$ Hz, 2H), 1.60–1.39 (m, 6H), 1.35–1.13 (m, 12H), 0.84 (t, $J = 7.3$ Hz, 3H), 0.80 (t, $J = 6.7$ Hz, 3H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 132.2$ (d, $J_{\text{PCC}} = 7.2$ Hz), 129.8 (d, $J_{\text{PCCC}} = 5.6$ Hz, 2C), 128.8 (d, $J_{\text{PCCCC}} = 2.7$ Hz, 2C), 126.9 (d, $J_{\text{PCCCCC}} = 3.2$ Hz), 64.3 (d, $J_{\text{POC}} = 6.9$ Hz), 36.5 (d, $J_{\text{PC}} = 84.0$ Hz), 32.9 (d, $J_{\text{POCC}} = 5.8$ Hz), 31.9, 30.9 (d, $J_{\text{PCC}} = 15.2$ Hz), 29.2, 29.1, 27.6 (d, $J_{\text{PC}} = 92.5$ Hz), 22.8, 21.8 (d, $J_{\text{PCCC}} = 4.3$ Hz), 18.9, 14.2, 13.4. HRMS (ESI) m/z calcd for $\text{C}_{19}\text{H}_{34}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 325.2290, found 325.2289.

Cyclohexyl hydroxymethyl(octyl)phosphinate (3g). To a solution of hypophosphorous acid (6.3 g, 48 mmol, 1 equiv.) was added paraformaldehyde (1.575 g, 52.5 mmol, 1.1 equiv.) and the reaction mixture was stirred under N_2 for 24 h. Then, cyclohexanol (10 mL, 96 mmol, 2 equiv.) and toluene (85 mL) were added and the solution was refluxed for 16 h under Ar using a Dean-Stark apparatus. During the reaction, the formation of a gel occurred. The reaction mixture was cooled to room temperature and transferred into another round bottom flask to remove the gel and concentrated under vacuum to afford the desired hydroxymethyl-*H*-phosphinate **1g** (9.4 g, 57% purity, 30.1 mmol, 63%) which was used without any further purification for the next step.²¹ A solution of compound **1g** (93.7 mg, 0.3 mmol, 1 equiv.), 1-octene **2a** (50.4 mg, 0.45 mmol, 1.5 equiv.) and DIPB (17.5 mg, 0.03 mmol, 10 mol%) in degazed DMSO (3 mL) under N_2 was irradiated under 365 nm for 16 h. Ethyl acetate (30 mL) was added and the reaction mixture was washed with saturated solution of NaHCO_3 (2×15 mL) and brine (2×15 mL). The organic layer was dried over MgSO_4 , filtered and the solvent was removed under reduce pressure. The residue was purified on a column of silica gel to afford the desired pure desired *H*-phosphinate **3g** (65.2 mg, 0.23 mmol, 75%). ^{31}P NMR (121 MHz, CDCl_3): $\delta = 52.9$ (s); ^1H NMR (300 MHz, CDCl_3): $\delta = 4.82$ (s br, 1H), 4.43–4.28 (m, 1H), 3.87–3.70 (m, 2H), 1.92–1.14 (m, 24H), 0.84 (d, $J = 6.6$ Hz, 3H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 74.4$ (d, $J_{\text{POC}} = 7.2$ Hz), 59.3 (d, $J_{\text{PC}} = 104.4$ Hz), 34.3 (d, $J_{\text{POCC}} = 3.2$ Hz), 34.2 (d, $J_{\text{POCC}} = 3.3$ Hz), 31.9, 31.0 (d, $J_{\text{PCC}} = 14.9$ Hz), 29.2 (2C), 26.5 (d, $J_{\text{PC}} = 89.6$ Hz), 25.2, 23.8 (2C), 22.7, 21.5 (d, $J_{\text{PCCC}} = 4.2$ Hz), 14.2; HRMS (ESI) m/z calcd for $\text{C}_{15}\text{H}_{32}\text{O}_3\text{P}$ ($[\text{M} + \text{H}]^+$) 291.2083, found 291.2080.

Cyclohexyl phenyl(3-phenylpropyl)phosphinate (3h). Following the general procedure A, using cyclohexyl phenyl-*H*-phosphinate **1a** (67.2 mg, 0.3 mmol, 1 equiv.), allylbenzene **2b** (35.4 mg, 0.3 mmol, 1 equiv.) and 50 mol% of catalyst (89 mg) affords **3h** (48.3 mg, 0.14 mmol, 48%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 43.0$ (s); ^1H NMR (300 MHz, CDCl_3): $\delta = 7.74$ –7.64 (m, 2H), 7.49–7.33 (m, 3H), 7.21–7.00 (m, 5H), 4.24–4.09 (m, 1H), 2.66–2.49 (m, 2H), 1.98–1.04 (m, 14H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 141.2$, 132.1 (d, $J_{\text{PC}} = 123$ Hz), 132.0 (d, $J_{\text{PCCCC}} = 2.7$ Hz), 131.6 (d, $J_{\text{PCCC}} = 9.8$ Hz, 2C), 128.5 (2C), 128.5 (d, $J_{\text{PCC}} = 12.2$ Hz, 2C), 128.4 (2C), 126.0, 74.4 (d, $J_{\text{POC}} = 6.7$ Hz), 36.6 (d, $J_{\text{PCC}} = 16.0$ Hz), 34.4 (d, $J_{\text{POCC}} = 2.9$ Hz), 33.7 (d, $J_{\text{POCC}} = 4.4$ Hz), 29.8 (d, $J_{\text{PC}} = 101$ Hz), 25.2, 23.7, 23.7, 23.5 (d, $J_{\text{PCCC}} = 3.2$ Hz); HRMS (ESI) m/z calcd for $\text{C}_{21}\text{H}_{28}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 342.1749, found 426.0564.

Cyclohexyl (3-methylbutan-2-yl)(phenyl)phosphinate (3i). Following the general procedure A, using cyclohexyl phenyl-*H*-phosphinate **1a** (67.2 mg, 0.3 mmol, 1 equiv.), 2-methylbut-2-

ene **2c** (22.1 mg, 0.3 mmol, 1 equiv.) and 50 mol% of catalyst (89 mg) affords **3i** (24 mg, 0.08 mmol, 27%) as a colorless oil (mixture 1 : 1 of the both diastereoisomers). ^{31}P NMR (121 MHz, CDCl_3): $\delta = 42.7$ (s), 46.3 (s); ^1H NMR (300 MHz, CDCl_3): $\delta = 7.76$ –7.66 (m, 2H), 7.49–7.34 (m, 3H), 4.26–4.10 (m, 1H), 2.31–1.48 (m, 7H), 1.42–1.03 (8H), 0.96–0.79 (m, 6H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 132.3$ (d, $J_{\text{PCC}} = 5.1$ Hz, $2 \times 0.5\text{C}$), 132.2 (d, $J_{\text{PCC}} = 5.2$ Hz, $2 \times 0.5\text{C}$), 131.9 (d, $J_{\text{PCCCC}} = 2.2$ Hz, 0.5C), 131.9 (d, $J_{\text{PCCCC}} = 2.2$ Hz, 0.5C), 128.5 (d, $J_{\text{PCC}} = 3.2$ Hz, $2 \times 0.5\text{C}$), 128.3 (d, $J_{\text{PCC}} = 3.1$ Hz, $2 \times 0.5\text{C}$), 74.2 (d, $J_{\text{POC}} = 7.2$ Hz, 0.5C), 74.0 (d, $J_{\text{POC}} = 7.1$ Hz, 0.5C), 39.6 (d, $J_{\text{PC}} = 100$ Hz $2 \times 0.5\text{C}$), 34.4 (d, $J_{\text{POCC}} = 2.4$ Hz, 0.5C), 34.4 (d, $J_{\text{POCC}} = 2.3$ Hz, 0.5C), 33.8 (d, $J_{\text{POCC}} = 1.9$ Hz, 0.5C), 33.8 (d, $J_{\text{POCC}} = 1.9$ Hz, 0.5C), 26.8 (d, $J_{\text{PCCC}} = 2.8$ Hz, 0.5C), 26.1 (0.5C), 25.4 ($2 \times 0.5\text{C}$), 23.8 ($2 \times 0.5\text{C}$), 23.7 ($2 \times 0.5\text{C}$), 22.6 (d, $J_{\text{PCC}} = 6.1$ Hz, 0.5C), 22.4 (d, $J_{\text{PCC}} = 8.7$ Hz, 0.5C), 18.3 (d, $J_{\text{PCC}} = 3.2$ Hz, 0.5C), 17.7 (d, $J_{\text{PCCC}} = 2.2$ Hz, 0.5C), 7.7 (0.5C), 6.9 (d, $J_{\text{PCC}} = 4.6$ Hz, 0.5C) (1 signal missing); HRMS (ESI) m/z calcd for $\text{C}_{17}\text{H}_{28}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 295.1821, found 295.1819.

Cyclohexyl (5-bromopentyl)(phenyl)phosphinate (3j). Following the general procedure A, using cyclohexyl phenyl-*H*-phosphinate **1a** (67.2 mg, 0.3 mmol, 1 equiv.), 5-bromo-1-pentene **2d** (46.6 mg, 0.3 mmol, 1 equiv.) and 10 mol% of catalyst (17.8 mg) affords **3j** (43 mg, 0.12 mmol, 38%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 42.7$ (s); ^1H NMR (300 MHz, CDCl_3): $\delta = 7.82$ –7.72 (m, 2H), 7.56–7.39 (m, 3H), 4.31–4.18 (m, 1H), 3.34 (t, $J = 6.8$ Hz, 2H), 2.06–1.11 (m, 18H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 132.2$ (d, $J_{\text{PC}} = 123$ Hz), 132.1 (d, $J_{\text{PCCCC}} = 2.7$ Hz), 131.7 (d, $J_{\text{PCC}} = 9.7$ Hz, 2C), 128.6 (d, $J_{\text{PCC}} = 12.3$ Hz, 2C), 74.5 (d, $J_{\text{POC}} = 6.7$ Hz), 34.5 (d, $J_{\text{POCC}} = 2.9$ Hz), 33.8 (d, $J_{\text{POCC}} = 4.4$ Hz), 33.5, 32.4, 30.3 (d, $J_{\text{PC}} = 101$ Hz), 29.3 (d, $J_{\text{PCC}} = 15.7$ Hz), 25.3, 23.8, 23.8, 21.2 (d, $J_{\text{PCCC}} = 3.6$ Hz); HRMS (ESI) m/z calcd for $\text{C}_{17}\text{H}_{27}\text{BrO}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 372.0854, found.

Cyclohexyl phenyl(3-hydroxypropyl)phosphinate (3k). Following the general procedure A, using cyclohexyl phenyl-*H*-phosphinate **1a** (67.2 mg, 0.3 mmol, 1 equiv.), allyl alcohol **2e** (17.4 mg, 0.3 mmol, 1 equiv.) and 50 mol% of catalyst (89 mg) affords **3k** (47.0 mg, 0.168 mmol, 56%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 44.9$ (s); ^1H NMR (300 MHz, CDCl_3): $\delta = 7.84$ –7.72 (m, 2H), 7.58–7.40 (m, 3H), 4.32–4.18 (m, 1H), 3.63–3.70 (m, 2H), 2.12–1.12 (m, 15H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 132.3$ (d, $J_{\text{PCCCC}} = 2.6$ Hz), 131.7 (d, $J_{\text{PCCC}} = 9.9$ Hz, 2C), 131.6 (d, $J_{\text{PC}} = 123.9$ Hz), 128.7 (d, $J_{\text{PCC}} = 12.4$ Hz, 2C), 74.9 (d, $J_{\text{POC}} = 6.7$ Hz), 62.8 (d, $J_{\text{PCC}} = 11.0$ Hz), 34.4 (d, $J_{\text{POCC}} = 2.7$ Hz), 33.8 (d, $J_{\text{POCC}} = 4.3$ Hz), 27.9 (d, $J_{\text{PC}} = 101.3$ Hz), 25.4 (d, $J_{\text{PCCC}} = 3.8$ Hz), 25.2, 23.7, 14.3; HRMS (ESI) m/z calcd for $\text{C}_{15}\text{H}_{24}\text{O}_3\text{P}$ ($[\text{M} + \text{H}]^+$) 283.1457, found 283.1455.

Cyclohexyl (4-hydroxybutyl) (phenyl) phosphinate (3l). Following the general procedure A, using cyclohexyl phenyl-*H*-phosphinate **1a** (67.2 mg, 0.3 mmol, 1 equiv.), 3-buten-1-ol **2f** (21.6 mg, 0.3 mmol, 1 equiv.) and 50 mol% of catalyst (89 mg) affords **3l** (61 mg, 0.207 mmol, 69%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 43.4$ (s); ^1H NMR (300 MHz, CDCl_3): $\delta = 7.81$ –7.72 (m, 2H), 7.58–7.41 (m, 3H), 4.30–4.16 (m, 1H), 3.65–3.57 (m, 2H), 2.06–1.13 (m, 17H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 132.2$ (d, $J_{\text{PCCCC}} = 2.6$ Hz), 132.0 (d, $J_{\text{PC}} = 123.0$ Hz), 131.7 (d, $J_{\text{PCCC}} = 9.8$ Hz, 2C), 128.6 (d, $J_{\text{PCC}} = 12.3$ Hz, 2C), 74.6 (d, $J_{\text{POC}} =$



6.8 Hz), 62.0, 34.5 (d, $J_{\text{POCC}} = 2.8$ Hz), 33.8 (d, $J_{\text{POCC}} = 4.4$ Hz), 33.5 (d, $J_{\text{PCC}} = 14$ Hz), 29.8 (d, $J_{\text{PC}} = 101$ Hz), 25.3, 23.8, 18.2 (d, $J_{\text{PCC}} = 3.6$ Hz) (1 signal missing); HRMS (ESI) m/z calcd for $\text{C}_{16}\text{H}_{26}\text{O}_3\text{P}$ ($[\text{M} + \text{H}]^+$) 297.1614, found 297.1612.

Octyl phosphinic acid (1c). Following the general procedure B, using 50% aqueous solution of hypophosphorous acid (79.2 mg, 0.6 mmol, 2 equiv.), 1-octene **2a** (33.6 mg, 0.3 mmol, 1 equiv.) and 20 mol% of catalyst (15.5 mg, 0.06 mmol) affords **1b** (29 mg, 0.165 mmol, 55%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 37.7$ (d, $J_{\text{PH}} = 544$ Hz); ^1H NMR (300 MHz, CDCl_3): $\delta = 12.53$ (br s, 1H), 7.11 (d, $J_{\text{PH}} = 544$ Hz, 1H), 1.85–1.71 (m, 2H), 1.70–1.53 (m, 2H), 1.49–1.24 (m, 10H), 0.91 (t, $J = 6.7$ Hz, 3H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 31.9$, 30.5 (d, $J_{\text{PCC}} = 16$ Hz), 29.4 (d, $J_{\text{PC}} = 94$ Hz), 29.2, 29.1, 22.7, 20.5 (d, $J_{\text{PCC}} = 2.7$ Hz), 14.2; HRMS (ESI) m/z calcd for $\text{C}_8\text{H}_{20}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 179.1195, found 179.1190.

3-Phenylpropyl phosphinic acid (1h). Following the general procedure B, using 50% aqueous solution of hypophosphorous acid (79.2 mg, 0.6 mmol, 2 equiv.), allylbenzene **2b** (35.4 mg, 0.3 mmol, 1 equiv.) and 20 mol% of catalyst (15.5 mg, 0.06 mmol) affords **1h** (34.7 mg, 0.189 mmol, 63%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 34.2$ (d, $J_{\text{PH}} = 541$ Hz); ^1H NMR (300 MHz, CDCl_3): $\delta = 7.32$ – 7.24 (m, 2H), 7.23–7.10 (m, 3H), 7.07 (d, $J_{\text{PH}} = 541$ Hz, 1H), 2.76–2.63 (m, 2H), 2.02–1.62 (m, 4H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 140.9$, 128.6 (4C), 126.4, 36.4 (d, $J_{\text{PCC}} = 16$ Hz), 28.8 (d, $J_{\text{PC}} = 94$ Hz), 22.4 (d, $J_{\text{PCC}} = 2.9$ Hz); HRMS (ESI) m/z calcd for $\text{C}_9\text{H}_{14}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 185.0725, found 185.0720.

1,2-Dimethylpropyl phosphinic acid (1i). Following the general procedure B, using 50% aqueous solution of hypophosphorous acid (79.2 mg, 0.6 mmol, 2 equiv.), 2-methylbut-2-ene **2c** (21 mg, 0.3 mmol, 1 equiv.) and 20 mol% of catalyst (15.5 mg, 0.06 mmol) affords **1i** (12.6 mg, 0.09 mmol, 31%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 44.0$ (d, $J_{\text{PH}} = 534$ Hz); ^1H NMR (300 MHz, CDCl_3): $\delta = 8.90$ (br s, 1H), 7.01 (d, $J_{\text{PH}} = 534$ Hz, 1H), 2.22–2.04 (m, 1H), 1.72–1.54 (m, 1H), 1.18–0.95 (m, 9H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 39.5$ (d, $J_{\text{PC}} = 94$ Hz), 27.1, 21.6 (d, $J_{\text{PCC}} = 12$ Hz), 19.0 (d, $J_{\text{PCC}} = 5.9$ Hz), 7.3; HRMS (ESI) m/z calcd for $\text{C}_5\text{H}_{14}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 137.0725, found 137.0720.

Cyclohexyl phosphinic acid (1j). Following the general procedure B, using 50% aqueous solution of hypophosphorous acid (79.2 mg, 0.6 mmol, 2 equiv.), cyclohexene (21 mg, 0.3 mmol, 1 equiv.) and 20 mol% of catalyst (15.5 mg, 0.06 mmol) affords **1k** (21.7 mg, 0.147 mmol, 49%) as a colorless oil. ^{31}P NMR (121 MHz, CDCl_3): $\delta = 42.8$ (d, $J_{\text{PH}} = 537$ Hz); ^1H NMR (300 MHz, CDCl_3): $\delta = 6.82$ (d, $J_{\text{PH}} = 537$ Hz, 1H), 2.01–1.54 (m, 6H), 1.39–1.14 (m, 5H); ^{13}C NMR (75 MHz, CDCl_3): $\delta = 37.6$ (d, $J_{\text{PC}} = 95$ Hz), 26.0 (2C), 25.8, 24.0 (2C); HRMS (ESI) m/z calcd for $\text{C}_6\text{H}_{14}\text{O}_2\text{P}$ ($[\text{M} + \text{H}]^+$) 149.0725, found 149.0720.

Conclusions

In conclusion, we developed an efficient UV-mediated hydrophosphinylation of unactivated alkenes with *H*-phosphinate under free radical conditions giving a straightforward access to a wide range of phosphorous compounds providing a metal free

alternative for the preparation and functionalization of P(O)–H bonds. This reaction can be carried out using an ionic liquid soluble photosensitizer while maintaining good yield. Furthermore, the reaction with hypophosphorous acid is disclosed for the first time. Finally, a continuous flow hydrophosphinylation was also reported which allowed faster reaction time with high yields, highlighting the potential of our methodology for rapid scale up and in-line synthesis. Taking into account the simplicity of our reaction conditions we believe this procedure will be appealing for further chemical and pharmaceutical applications.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- (a) L. D. Quin, *A Guide to Organophosphorous Chemistry*, Wiley Interscience, New York, 2000; (b) *New Aspects in Phosphorous Chemistry*, ed. J.-P. Majoral, Springer, Berlin, 2002–2005, vol. 1–5; (c) D. E. C. Corbridge, in *Phosphorus: Chemistry, Biochemistry and Technology*; CRC Press: London, edn 6, 2013.
- J.-L. Montchamp, *Acc. Chem. Res.*, 2014, **47**, 77.
- For selected reviews see: (a) O. Berger and J.-L. Montchamp, *Beilstein J. Org. Chem.*, 2014, **10**, 732; (b) N. Z. Kiss and G. Keglevich, *Curr. Org. Chem.*, 2014, **18**, 2673; (c) J.-L. Montchamp, *Top. Curr. Chem.*, 2015, **361**, 217.
- (a) L. Coudray and J.-L. Montchamp, *Eur. J. Org. Chem.*, 2008, 3601; (b) X.-Q. Pan, J.-P. Zou, W.-B. Yi and W. Zhang, *Tetrahedron*, 2015, **71**, 7481.
- (a) C. Petit, F. Fécourt and J.-L. Montchamp, *Adv. Synth. Catal.*, 2011, **353**, 1883; (b) S. Deprele and J.-L. Montchamp, *J. Am. Chem. Soc.*, 2002, **124**, 9386; (c) P. Ribiere, K. Bravo-Altamirano, M. I. Antczak, J. D. Hawkins and J.-L. Montchamp, *J. Org. Chem.*, 2005, **70**, 4064; (d) S. Ortial, H. C. Fisher and J.-L. Montchamp, *J. Org. Chem.*, 2013, **78**, 6599.
- (a) E. E. Nifant'ev, R. K. Magdeeva and N. P. Shchepet'eva, *J. Gen. Chem.*, 1980, **50**, 1416; (b) S. Deprele and J.-L. Montchamp, *J. Org. Chem.*, 2001, **66**, 6745; (c) O. Dubert, A. Gautier, E. Condamine and S. R. Piettre, *Org. Lett.*, 2002, **4**, 359; (d) A. Gautier, G. Garipova, C. Salcedo, S. Balieu and S. R. Piettre, *Angew. Chem., Int. Ed.*, 2004, **43**, 5963; (e) M. I. Antczak and J.-L. Montchamp, *Synthesis*, 2006, 3080; (f) Y. Feng and J. K. Coward, *J. Med. Chem.*, 2006, **49**, 770; (g) T. Hirai and L.-B. Han, *Org. Lett.*, 2007, **9**, 53; (h) H. C. Fisher, O. Berger, F. Gelat and



- J.-L. Montchamp, *Adv. Synth. Catal.*, 2014, **356**, 1199; (i) P. Troupa, G. Katsiouleri and S. Vassiliou, *Synlett*, 2015, **26**, 2714; (j) Z. Li, F. Fan, Z. Zhang, Y. Xiao, D. Liu and Z.-Q. Liu, *RSC Adv.*, 2015, **5**, 27853.
- 7 G. Fausti, F. Morlet-Savary, J. Lalevee, A.-C. Gaumont and S. Lakhdar, *Chem.-Eur. J.*, 2017, **23**, 2144.
- 8 A. Dondoni, S. Staderini and A. Marra, *Eur. J. Org. Chem.*, 2013, 5370.
- 9 P.-Y. Geant, B. S. Mohamed, C. Perigaud, S. Peyrottes, J.-P. Uttaro and C. Mathe, *New J. Chem.*, 2016, **40**, 5318.
- 10 During the preparation of the manuscript a similar approach was developed: P.-Y. Geant, J.-P. Uttaro, S. Peyrottes and C. Mathé, *Eur. J. Org. Chem.*, 2017, 2850.
- 11 For recent reviews on flow photochemistry, see: (a) Z. J. Garlets, J. D. Nguyen and C. R. J. Stephenson, *Isr. J. Chem.*, 2014, **54**, 351; (b) Y. H. Su, N. J. W. Straathof, V. Hessel and T. Noel, *Chem.-Eur. J.*, 2014, **20**, 10562; (c) D. Cambie, C. Bottecchia, N. J. W. Straathof, V. Hessel and T. Noel, *Chem. Rev.*, 2016, **116**, 1023.
- 12 (a) M. Penhoat, T. Vanbésien, A. Cocud, A. Addad, H. Vezin and C. Rolando, *New J. Chem.*, 2016, **40**, 9460; (b) N. El Achi, M. Penhoat, Y. Bakkour, C. Rolando and L. Chausset-Boissarie, *Eur. J. Org. Chem.*, 2016, 4284; (c) N. El Achi, Y. Bakkour, L. Chausset-Boissarie, M. Penhoat and C. Rolando, *RSC Adv.*, 2017, **7**, 29815.
- 13 (a) F. Effenberger and H. Kottmann, *Tetrahedron*, 1985, **41**, 4171; (b) O. Berger and J.-L. Montchamp, *Chem.-Eur. J.*, 2014, **20**, 12385.
- 14 S. C. Hubbard and P. B. Jones, *Tetrahedron*, 2005, **61**, 7425.
- 15 O. Berger, L. Gavara and J.-L. Montchamp, *Org. Lett.*, 2012, **14**, 3404.
- 16 (a) O. Tayama, A. Nakano, T. Iwahama, S. Sakaguchi and I. Yasutaka, *J. Org. Chem.*, 2004, **69**, 5494; (b) S.-I. Kawaguchi, A. Nomoto, M. Sonoda and A. Ogawa, *Tetrahedron Lett.*, 2009, **50**, 624; (c) W.-J. Yoo and S. Kobayashi, *Green Chem.*, 2013, **15**, 1844.
- 17 (a) A. Beckett and G. Porter, *Faraday Trans.*, 1963, **59**, 2038; (b) G. Porter and P. Suppan, *Faraday Trans.*, 1965, **61**, 1664; (c) J. Lalevee, F. Morlet-Savary, M. A. Tehfe, B. Graff and J. P. Fouassier, *Macromolecules*, 2012, **45**, 5032; (d) M. Marazzi, S. Mai, D. Roca-Sanjuán, M. I. G. Delcey, R. Lindh, L. González and A. Monari, *J. Phys. Chem. Lett.*, 2016, **7**, 622–626; (e) G. R. Dormán, H. Nakamura, A. Pulsipher and G. D. Prestwich, *Chem. Rev.*, 2016, **116**, 15284.
- 18 C. S. Demmer, N. Krogsgaard-Larsen and L. Bunch, *Chem. Rev.*, 2011, **11**, 7981.
- 19 S. Deprèle and J.-L. Montchamp, *J. Organomet. Chem.*, 2002, **643**, 154.
- 20 O. Berger, C. Petit, E. L. Deal and J.-L. Montchamp, *Adv. Synth. Catal.*, 2013, **355**, 1361.
- 21 All attempts of purification by aqueous work-up or chromatography failed or conducted in a very low yield. Purity was determined by ³¹P NMR on a sample by mixing it with phenylphosphinic acid as a standard.

