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# Enhanced crystallinity of $\text{CH}_3\text{NH}_3\text{PbI}_3$ by the pre-coordination of $\text{PbI}_2$ -DMSO powders for highly reproducible and efficient planar heterojunction perovskite solar cells†

Jiyong Lee<sup>a</sup> and Seunghyun Baik <sup>\*b</sup>

Solution processable  $\text{CH}_3\text{NH}_3\text{PbI}_3$  has received considerable attention for highly-efficient perovskite solar cells. However, the different solubility of  $\text{PbI}_2$  and  $\text{CH}_3\text{NH}_3\text{I}$  is problematic, initiating active solvent engineering research using dimethyl sulfoxide (DMSO). Here we investigated the pre-coordination of  $\text{PbI}_2$ -DMSO powders for planar heterojunction perovskite solar cells fabricated by a low-temperature process ( $\leq 100$  °C). Pre-coordination was carried out by simple mechanical mixing using a mortar and pestle. The composition of  $\text{PbI}_2$ -DMSO<sub>x</sub> ( $x = 0, 1, \text{ or } 2$ ) in the powder mixture was investigated by gradually increasing mechanical mixing time, and a dominant composition of  $\text{PbI}_2$ -DMSO<sub>1</sub> was obtained after a 10 min mixing process. The pre-coordinated  $\text{PbI}_2$ -DMSO powders were then blended with  $\text{CH}_3\text{NH}_3\text{I}$  in DMF to make the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film by toluene-assisted spin-coating and heat treatment. Compared with the one-step blending of  $\text{CH}_3\text{NH}_3\text{I}$ ,  $\text{PbI}_2$ , and DMSO in DMF, the pre-coordination method resulted in better dissolution of  $\text{PbI}_2$ , larger grain size, and pinhole-free morphology. Consequently, absorption, fluorescence, carrier lifetime, and charge extraction were enhanced. The average open-circuit voltage (1.046 V), short-circuit current (22.9 mA cm<sup>-2</sup>), fill factor (73.5%), and power conversion efficiency (17.6%) were increased by 2–12% with decreased standard deviations (13–50%), compared with the one-step blending method. The best efficiency was 18.2%. The simple mechanical pre-coordination of  $\text{PbI}_2$ -DMSO powders was very effective in enhancing the crystallinity of  $\text{CH}_3\text{NH}_3\text{PbI}_3$  and photovoltaic performance.

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## 1. Introduction

Photovoltaic technology that directly converts solar light into electricity has received considerable attention to substitute fossil fuels.<sup>1–3</sup> However, the per-watt cost of silicon-based first generation photovoltaic devices is still greater than that of fossil-fuel-based electricity.<sup>1–3</sup> The low-cost titanium dioxide ( $\text{TiO}_2$ )-based photovoltaic devices such as dye-sensitized solar cells have been actively investigated to address this hurdle.<sup>1–4</sup> Recently, the low-cost and solution processable organometal halide perovskite absorbers, in combination with the double-layered  $\text{TiO}_2$  structure, have achieved over 20% power conversion efficiency (PCE).<sup>5–7</sup> The  $\text{TiO}_2$  double layers involved a mesoporous or nano-structured  $\text{TiO}_2$  layer, which was constructed at high temperatures ( $>550$  °C) or by complicated self-assembly

methods, on top of a blocking compact  $\text{TiO}_2$  layer to enhance interfacial coupling with perovskite.<sup>5–10</sup> On the other hand, low-temperature processable planar heterojunction perovskite solar cells are of great interest due to the simple single-layered  $\text{TiO}_2$  structure and cost-competitive solar power applications.<sup>2,11–14</sup>

It is challenging to construct highly crystalline perovskite on the  $\text{TiO}_2$  layer with high reproducibility by the solution process.<sup>11–25</sup> The irregular solvent evaporation rate, different solubility of  $\text{PbI}_2$  and  $\text{CH}_3\text{NH}_3\text{I}$ , and different crystal growth rate of components in the precursor solution lead to large morphological deviations of perovskite film and photovoltaic performance variations.<sup>12–25</sup> The additive-aided dissolution of  $\text{PbI}_2$  and/or control of solvent evaporation rate have been investigated to enhance reproducibility and PCE.<sup>20–25</sup> The solvent engineering using  $\text{CH}_3\text{NH}_3\text{I}$ ,  $\text{PbI}_2$ , and dimethyl sulfoxide (DMSO) in *N,N*-dimethylformamide (DMF) or gamma-butyrolactone was actively investigated.<sup>2,21,22,26,27</sup> The adduct of  $\text{CH}_3\text{NH}_3\text{I}$ - $\text{PbI}_2$ -DMSO was formed due to the interaction between Lewis base (lone pair electrons on oxygen in DMSO and  $\Gamma^-$  in  $\text{CH}_3\text{NH}_3\text{I}$ ) and Lewis acid ( $\text{PbI}_2$ ) enhancing the solubility of  $\text{PbI}_2$ .<sup>22</sup> This approach, together with the DMF evaporation control using nonpolar diethyl ether, enhanced both the

<sup>a</sup>Department of Energy Science, Sungkyunkwan University, Suwon 16419, Republic of Korea

<sup>b</sup>School of Mechanical Engineering, Sungkyunkwan University, Suwon 16419, Republic of Korea. E-mail: [sbaik@me.skku.ac.kr](mailto:sbaik@me.skku.ac.kr)

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reproducibility and PCE (average: 18.3%, best: 19.7%).<sup>22</sup> Pham *et al.* added lead chloride (PbCl<sub>2</sub>) with the Lewis acid–base adducts for efficient perovskite solar cells (average: 18.1%, best: 19%).<sup>27</sup> The PbCl<sub>2</sub> additive decreased defects and enhanced crystallinity and PCE.<sup>27</sup>

The water content in the lead halide was also systematically examined to improve solubility of PbI<sub>2</sub> and PCE of CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub>-based solar cells.<sup>28</sup> On the other hand, tin-based perovskites have been actively investigated due to the lower toxicity.<sup>29</sup> Solvent-coordinated tin halide complexes were investigated as impurity-free precursors for tin halide perovskite-based solar cells.<sup>29</sup>

Here we investigated the pre-coordination of PbI<sub>2</sub>–DMSO powders for the planar heterojunction CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub>-based solar cells. The pre-coordination was carried out by the simple mechanical mixing of DMSO and PbI<sub>2</sub> particles (1 : 1 mol%) using a mortar and pestle. The composition of PbI<sub>2</sub>–DMSO<sub>x</sub> ( $x = 0, 1, \text{ or } 2$ ) in the powder mixture was investigated by gradually increasing mechanical mixing time, and the dominant composition of PbI<sub>2</sub>–DMSO<sub>1</sub> was obtained after 10 min mixing process. The PbI<sub>2</sub>–DMSO powders were then mixed with CH<sub>3</sub>NH<sub>3</sub>I in DMF to make CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> film by spin-application and heat treatment (100 °C). The planar TiO<sub>2</sub> layer (50 nm) was formed on a fluorine-doped tin oxide (FTO) glass substrate using the low-temperature ( $\leq 100$  °C) TiCl<sub>4</sub> bath deposition method.<sup>2,13</sup> For comparison, the CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> film was also synthesized by one-step blending of CH<sub>3</sub>NH<sub>3</sub>I, PbI<sub>2</sub>, and DMSO in DMF following a previously published protocol.<sup>22</sup> The pre-coordination method resulted in better dissolution of PbI<sub>2</sub> which was confirmed by X-ray diffraction (XRD) analysis. A larger grain size and smaller grain boundary defects, such as pin holes, were achieved leading to an increase in absorption and photoluminescence. Accordingly, the charge carrier life time was increased and better charge extraction was achieved which was confirmed by time-resolved photoluminescence (TRPL) analysis. Overall, the average open circuit voltage ( $V_{oc}$ , 1.046 V), short-circuit current ( $J_{sc}$ , 22.9 mA cm<sup>-2</sup>), fill factor (FF, 73.5%), and PCE (17.6%) were increased by 2, 3, 6, and 12%, respectively, compared with the one-step blending method. The standard deviations of  $V_{oc}$  ( $\pm 0.009$  V),  $J_{sc}$  ( $\pm 0.4$  mA cm<sup>-2</sup>), FF ( $\pm 1.2\%$ ), and PCE ( $\pm 0.37\%$ ) were decreased by 36, 13, 54, and 50%, compared with the one-step blending method, demonstrating improved reproducibility. The best FF and PCE were 75.5% and 18.2%. This indicated that the simple mechanical pre-coordination of PbI<sub>2</sub>–DMSO powders was very efficient in enhancing the dissolution of PbI<sub>2</sub>, crystallinity of CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub>, and photovoltaic performance of planar heterojunction perovskite solar cells prepared by the low-temperature process ( $\leq 100$  °C).

## 2. Experimental

### 2.1. Transparent conductive substrate-electron transport layer preparation

All chemicals were of analytical grade and used as received without further purification. The chemical bath deposition method was used to construct a planar TiO<sub>2</sub> layer on a fluorine-

doped SnO<sub>2</sub> (FTO)-coated glass substrate (Pilkington, TEC 8). The cleaned FTO glass was immersed in 0.2 M titanium tetrachloride (TiCl<sub>4</sub>) aqueous solution at 60 °C for 65 min, washed with de-ionized water, and dried on a hot plate (100 °C, 1 h).<sup>1,13</sup> The entire TiO<sub>2</sub> deposition process was carried out at low temperature ( $\leq 100$  °C).

### 2.2. CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> precursor solution preparation

The CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> film was fabricated on top of the planar TiO<sub>2</sub> layer by two different methods; one-step blending and DMSO pre-coordination (DPC) methods. For the DPC method, PbI<sub>2</sub> powders (1.22 g, Alpha Aesar) were mechanically mixed with DMSO (207 mg, Aldrich) using a mortar and pestle (1 : 1 mol%). The mechanical mixing time was varied (3, 6, or 10 min). The CH<sub>3</sub>NH<sub>3</sub>I was prepared by stirring methylamine (27.8 mL, 40% in methanol, TCI) with hydroiodic acid (30 mL, 57 wt% in water, Aldrich) in a round-bottom flask (0 °C, 2 h).<sup>2</sup> The CH<sub>3</sub>NH<sub>3</sub>I precipitate was then recovered by evaporation (50 °C for 5 h), washing with diethyl ether (30 min, 3 times), and drying under vacuum (60 °C for 24 h). Finally, the pre-coordinated PbI<sub>2</sub>–DMSO powders were stir-mixed (12 h) with CH<sub>3</sub>NH<sub>3</sub>I (0.42 g) in DMF (2 mL, Aldrich) to prepare the CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> precursor solution. For the one-step blending method, the CH<sub>3</sub>NH<sub>3</sub>I powder (0.42 g), PbI<sub>2</sub> powder (1.22 g), and DMSO (207 mg) were simultaneously added in DMF (2 mL) (1 : 1 : 1 mol%) and stirred at room temperature for 12 h to obtain the CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> precursor solution following a previously published protocol.<sup>22</sup>

### 2.3. CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> film fabrication

The same procedure was used to fabricate the CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> film using the precursor solution prepared either by the one-step blending method or DPC method. The CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> precursor solution was coated on the planar TiO<sub>2</sub> layer the by two-step spinning process (1000 rpm for 3 s and 4500 rpm for 8 s) in a N<sub>2</sub> filled glove box. The evaporation rate of DMF was controlled by toluene dripping (0.3 mL) in the final stage ( $\sim 2$  s) of spinning process.<sup>21</sup> The final CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub> film was obtained by heat treatment (100 °C) at ambient air condition for 20 min.

### 2.4. Hole transport layer-electrode fabrication

The spiro-OMeTAD powder (72.3 mg, Solaronix) was dissolved in chlorobenzene (1.5 mL, Aldrich) with 4-*tert*-butylpyridine (28.8  $\mu$ L, Aldrich) and Li-TFSI solution (17.5  $\mu$ L, 520 mg in 1 mL acetonitrile) for the hole transporting layer precursor solution.<sup>2</sup> The spiro-OMeTAD layer was spin-deposited (4000 rpm, 40 s) on the CH<sub>3</sub>NH<sub>3</sub>PbI<sub>3</sub>/TiO<sub>2</sub>/FTO/glass followed by drying in vacuum for 12 h. Finally, the gold electrode ( $\sim 100$  nm) was deposited on the spiro-OMeTAD layer by thermal evaporation. The active solar cell area was maintained constant (0.15 cm<sup>2</sup>) using a photo mask.

### 2.5. Characterization

The pre-coordinated PbI<sub>2</sub>–DMSO powders were investigated by Fourier transform infrared (FT-IR) spectroscopy (Bruker, IFS-66/S), XRD (Rigaku, D/MAX-2500/PC, Cu K <sub>$\alpha$</sub>  radiation,



$\lambda = 1.5418 \text{ \AA}$ ), differential scanning calorimetry (DSC, SEIKO, 7020, heating rate =  $2 \text{ }^\circ\text{C min}^{-1}$ ), and thermogravimetric analysis (TGA, SEIKO, Exstar 6000). The microstructure was investigated by scanning electron microscopy (SEM, JEOL, JSM-7600F) and XRD. The UV-Vis spectroscopy (Agilent, RSA-HP-8453) was performed on the planar specimen ( $\text{CH}_3\text{NH}_3\text{PbI}_3/\text{TiO}_2/\text{FTO}/\text{glass}$ ). Photoluminescence and TRPL were measured using a confocal microscope system (NTEGRA SPECTRA, NT-MDT) with an objective lens ( $100\times$ , NA 0.7) and a grating (150 lines per mm, 400 nm blaze). The excitation source was a pulsed laser (405 nm, 20 MHz,  $1 \mu\text{W cm}^{-2}$ ). Photoluminescence data were collected using a thermoelectrically-cooled CCD detector. TRPL data were collected with a high-speed photomultiplier tube detector (Photonic Solutions, PMC-100, integration time = 10 s) for a time-correlated single-photon-counting system. The current density–voltage ( $J$ - $V$ ) characteristics were measured using a digital source meter (Keithley 2400) under the AM 1.5 one sun illumination (Oriel, Sol3ATM,  $100 \text{ mW cm}^{-2}$ ). The voltage scan rate was  $0.3 \text{ V s}^{-1}$ . The external quantum efficiency (EQE) was also measured (Oriel, IQE 200). The electrochemical impedance spectroscopy (EIS) was performed using a potentiostat/galvanostat (Biologic, SP-240), and the data were analyzed using a Z-fit software (EC-Lab).

### 3. Results and discussion

A schematic of the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film synthesis process is shown in Fig. 1A. In contrast to the one-step blending method where  $\text{PbI}_2$ ,  $\text{CH}_3\text{NH}_3\text{I}$ , DMSO, and DMF are initially mixed at one

time,<sup>21,22,27</sup>  $\text{PbI}_2$  powders (2.65 mM) and DMSO (2.65 mM) were mechanically pre-mixed using a mortar and pestle in the DPC method. The mechanical mixing time was gradually increased (3, 6, and 10 min) to systematically investigate  $\text{PbI}_2$ -DMSO<sub>*x*</sub> ( $x = 0, 1, \text{ or } 2$ ) adduct formation in the powder. The loss of DMSO was negligible during the mortar mixing process due to the low saturation vapor pressure (0.76 kPa at  $60 \text{ }^\circ\text{C}$ ) and high boiling point ( $189 \text{ }^\circ\text{C}$ ).<sup>24</sup> The thermogravimetric analysis of the pre-coordinated  $\text{PbI}_2$ -DMSO<sub>*x*</sub> powders indicated a 14.4% weight decrease at high temperatures ( $>140 \text{ }^\circ\text{C}$ ) which was consistent with the initial weight ratio between DMSO and  $\text{PbI}_2$  powders (ESI Fig. S1†).

Fig. 1B shows optical images of as-received  $\text{PbI}_2$  and  $\text{PbI}_2$ -DMSO<sub>*x*</sub> adduct powders. The adduct powder exhibited lighter yellowish color with increasing mechanical mixing time. Fig. 1C compares FT-IR spectra of the powders. The sulfoxide oxygen in DMSO (electron-rich Lewis base) donates an electron pair to  $\text{PbI}_2$  (electron-deficient Lewis acid) in the  $\text{PbI}_2$ -DMSO adduct formation *via* Pb–O coordinate covalent bonding.<sup>22,26,30</sup> The Pb–O bond formation was confirmed by the red shift of S=O stretching vibration.<sup>22,26,30</sup> The  $\nu(\text{S=O})$  peaks at  $1020 \text{ cm}^{-1}$  and  $978 \text{ cm}^{-1}$  correspond to  $\text{PbI}_2$ -DMSO<sub>1</sub> and the peak at  $989 \text{ cm}^{-1}$  indicates  $\text{PbI}_2$ -DMSO<sub>2</sub>.<sup>30</sup> The  $\text{PbI}_2$ -DMSO<sub>2</sub> peak decreased and stoichiometric  $\text{PbI}_2$ -DMSO<sub>1</sub> peaks increased as the mechanical mixing time increased. This was consistent with XRD characterization (Fig. 1D). The  $\text{PbI}_2$ -DMSO<sub>2</sub> peaks decreased as the mechanical mixing time increased. The  $\text{PbI}_2$ -DMSO<sub>1</sub> peaks became dominant after 10 min mixing. DSC analysis also revealed the dominant  $\text{PbI}_2$ -DMSO<sub>1</sub> formation as the

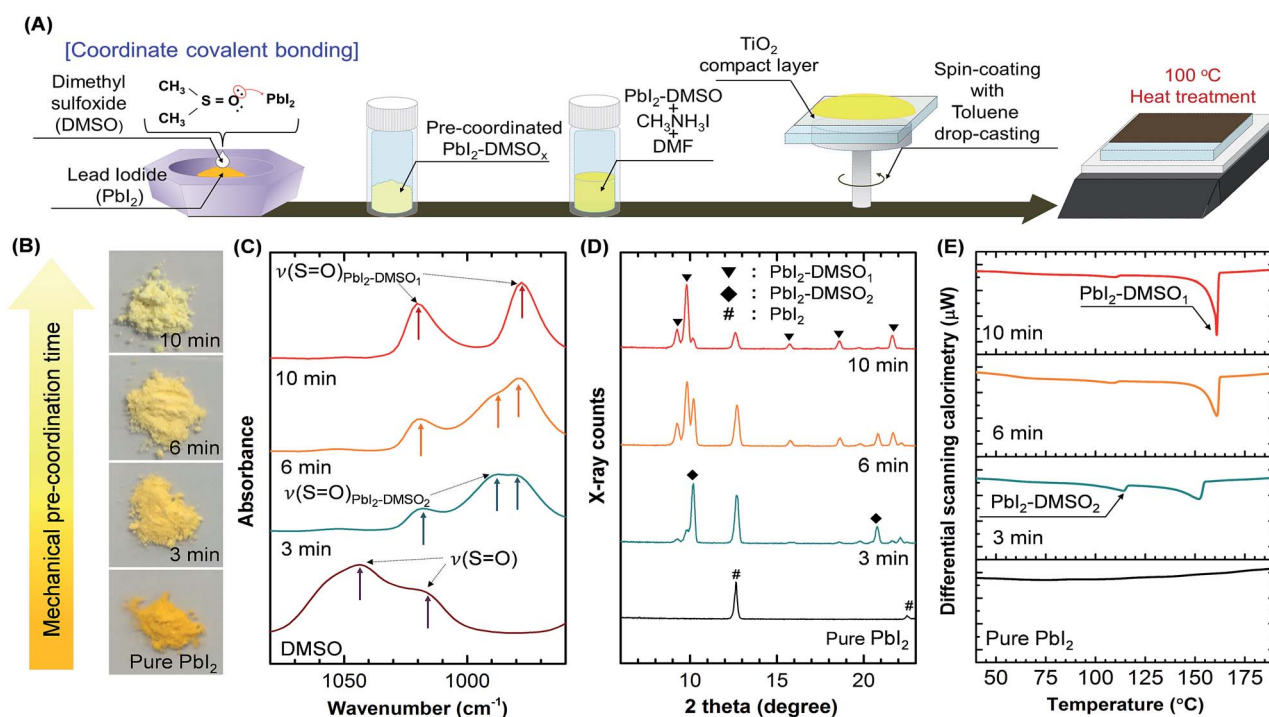


Fig. 1 The  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film synthesis on the planar  $\text{TiO}_2$  by the DMSO pre-coordination (DPC) method. (A) Schematic of the synthesis process. (B–E) The  $\text{PbI}_2$ -DMSO<sub>*x*</sub> ( $x = 0, 1, \text{ or } 2$ ) adduct formation in the powders was investigated as a function of the mechanical mixing time by optical images (B), FT-IR (C), XRD (D), and DSC (E) analyses.



mechanical mixing time increased (Fig. 1E). The endothermic peak at 114 °C indicates  $\text{PbI}_2\text{-DMSO}_2$ , and the peak at a higher temperature corresponds to  $\text{PbI}_2\text{-DMSO}_1$ .<sup>31</sup> The different analytical methods consistently revealed the dominant  $\text{PbI}_2\text{-DMSO}_1$  adduct formation as the mechanical mixing time increased up to 10 min in the DPC method. This left only a small amount of uncoordinated  $\text{PbI}_2$  and  $\text{PbI}_2\text{-DMSO}_2$  in the powder mixture. The pre-coordinated  $\text{PbI}_2\text{-DMSO}$  powders were then mixed with  $\text{CH}_3\text{NH}_3\text{I}$  powders in DMF by stirring. The precursor solution was then coated on a planar  $\text{TiO}_2/\text{FTO}/\text{glass}$  substrate by the two-step spinning process. The evaporation rate of DMF was controlled by toluene dripping in the final stage ( $\sim 2$  s) of spinning process.<sup>21</sup> The residual solvent was quickly removed by the toluene dripping resulting in spontaneous formation of the  $\text{CH}_3\text{NH}_3\text{I-PbI}_2\text{-DMSO}$  intermediate phase.<sup>21,22</sup> Finally, the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film was cured at 100 °C in ambient air environment (20 min).<sup>12</sup>

It is interesting to note that the  $\text{PbI}_2\text{-DMSO}$  coordination was maintained even after dissolving the pre-coordinated  $\text{PbI}_2\text{-DMSO}_x$  powders (DPC-10 min) in DMF. As shown in ESI Fig. S2,† there was negligible difference in the TGA spectra of DPC-10 min powders before and after dissolving in DMF. The fully dissolved DPC-10 min powders in DMF were retrieved by precipitation using toluene and drying in vacuum for the TGA measurement. In contrast,  $\text{PbI}_2$  powders dissolved in DMF and retrieved by precipitation ( $\text{PbI}_2\text{-DMF}$ ) shows a decrease in weight concentration at lower temperatures. This clearly indicated the maintenance of  $\text{PbI}_2\text{-DMSO}$  coordination in DMF solvent.

The crystallinity of perovskite films was investigated by XRD analysis (Fig. 2A), and corresponding SEM images are shown in

ESI Fig. S2.† The  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film prepared without DMSO additive exhibited a residual  $\text{PbI}_2$  (001) diffraction peak at 12.6°. The  $\text{PbI}_2$  with low solubility and  $\text{CH}_3\text{NH}_3\text{I}$  with high solubility resulted in the rapid crystal growth of nanorod-like  $\text{PbI}_2$  with the porous  $\text{CH}_3\text{NH}_3\text{PbI}_3$  morphology during the fast evaporation of solvent (*i.e.*, DMF) as shown in ESI Fig. S2A.†<sup>20,22,25</sup> The residual  $\text{PbI}_2$  peak was reduced in the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film prepared by the one-step blending method. The film crystallinity was significantly improved although a number of pin-holes were observed (ESI Fig. S2B†). The  $\text{CH}_3\text{NH}_3\text{I-PbI}_2\text{-DMSO}$  complex enhanced solubility of  $\text{PbI}_2$  and conversion to  $\text{CH}_3\text{NH}_3\text{PbI}_3$  with more uniform film morphology.<sup>21–27</sup> Interestingly, the unconverted residual  $\text{PbI}_2$  diffraction peak became negligible in the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film as the mechanical mixing time increased in the DPC method. The number of pin-holes also decreased with the increasing mixing time (ESI Fig. S2C–F†). The (110) peaks of perovskite films were magnified and fitted with Gaussian function (Fig. 2B).<sup>11</sup> The full width at half maximum (FWHM) was 0.255° for the one step blending method. The FWHM decreased from 0.253° to 0.230° as the mechanical mixing time increased from 3 to 10 min in the DPC method demonstrating improved crystallinity.<sup>11</sup> The XRD analysis indicated that the dominant  $\text{PbI}_2\text{-DMSO}_1$  composition in the mechanically mixed powder (Fig. 1) resulted in highly crystalline perovskite film without leaving unconverted  $\text{PbI}_2$ . There was negligible difference in the thickness ( $\sim 500$  nm) of  $\text{CH}_3\text{NH}_3\text{PbI}_3$  films prepared by the one-step blending method and the DPC method with different mechanical mixing times (see cross-sectional SEM images in ESI Fig. S3†). The enlarged grain size of the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film synthesized by the DPC

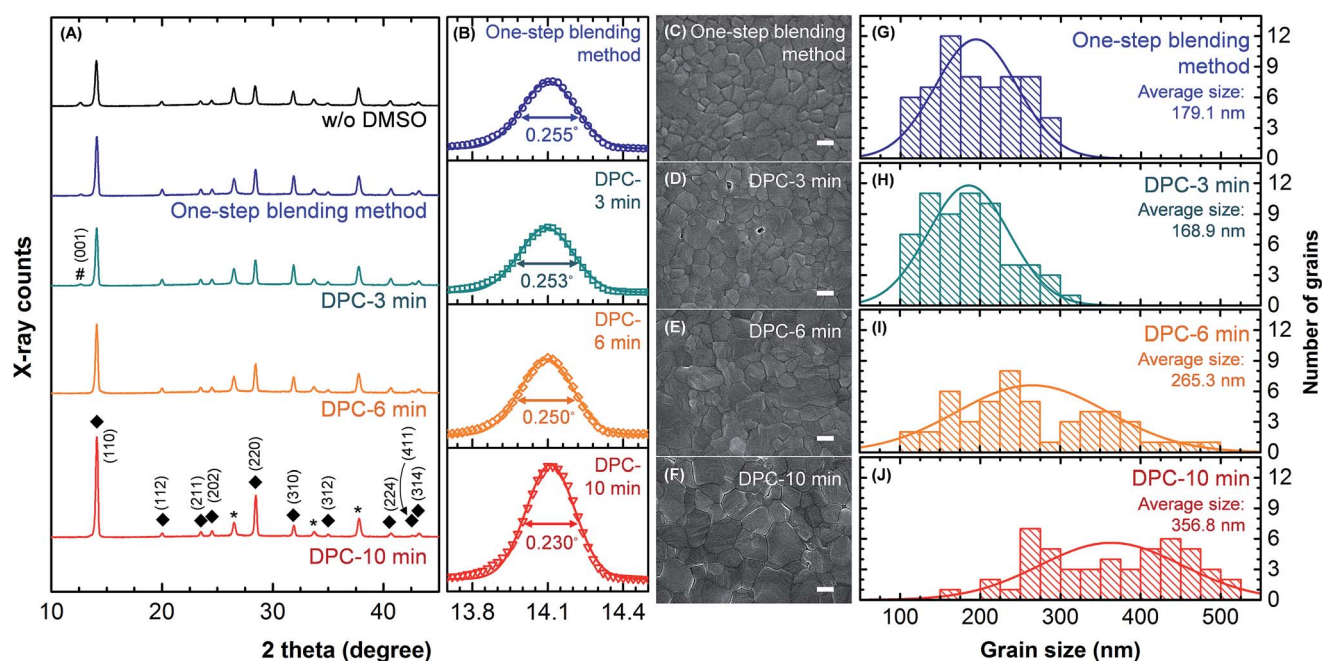


Fig. 2 The crystallinity of  $\text{CH}_3\text{NH}_3\text{PbI}_3$  films prepared by the one-step blending method and DPC method (3, 6, and 10 min mechanical mixing). (A) XRD analysis. The film prepared without DMSO was also compared. The  $\text{PbI}_2$ ,  $\text{CH}_3\text{NH}_3\text{PbI}_3$ , and FTO peaks are denoted by #,  $\blacklozenge$ , and \*, respectively. (B) The (110) peak of  $\text{CH}_3\text{NH}_3\text{PbI}_3$  was magnified and fitted with Gaussian distribution function. (C–F) SEM images of the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  films. The scale bar is 200 nm. (G–J) Histogram of the grain size distribution.



method could be clearly observed from SEM images (Fig. 2C–F). The corresponding grain size distribution is shown in Fig. 2G–J. Each grain area was measured from SEM images, and the grain size was defined as a diameter of a circle with an identical area of the grain. The one-step blending method produced the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  grain size distribution from 109 to 288 nm with an average of 179.1 nm. The grain size increased as the mechanical mixing time increased in the DPC method. The average grain size was 356.8 nm with a size distribution from 173 to 506 nm after the 10 min mixing process. Larger grains, reduced grain boundaries, and pin-hole free morphology were achieved by the 10 min pre-coordinated  $\text{PbI}_2$ -DMSO powders. Perovskite films prepared by the 10 min mechanical mixing process were used for the further DPC method analysis unless otherwise stated.

Fig. 3A compares UV-Vis absorption spectra of the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  films synthesized without DMSO, by the one-step blending method, and by the DPC method. The different synthesis method did not affect the film thickness as shown in ESI Fig. S3.† The absorption onset centered at  $\sim 770$  nm was observed similar to other  $\text{CH}_3\text{NH}_3\text{PbI}_3$  films.<sup>32</sup> The film synthesized without DMSO provided the minimum absorbance. This could be due to the fast growth of nanorod-like  $\text{PbI}_2$  with surrounding large pin-holes reducing the coverage of  $\text{CH}_3\text{NH}_3\text{PbI}_3$  (ESI Fig. S2†). In contrast, the film synthesized by the

DPC method showed the highest absorbance. The pinhole free morphology was observed for the film synthesized by the DPC method (ESI Fig. S2†) increasing the absorbance of the film. The enlarged perovskite grains also reduced the grain boundaries and defective electronic trap sites dominantly located at grain boundaries.<sup>6,31</sup> The elimination of non-radiative recombination levels within the band gap, caused by removal of the defective sites at grain boundaries, enhanced photoluminescence intensity (DPC-10 min) as shown in Fig. 3B. The quenching of photoluminescence by the defective sites at grain boundaries was also reported in literature.<sup>6,32,34</sup>

TRPL spectroscopy was employed to investigate the carrier recombination life time which is related with the charge diffusion length.<sup>32,35</sup> The TRPL spectra of the perovskite/glass configuration is shown in Fig. 3C.<sup>6</sup> The TRPL data were fitted using a bi-exponential decay function to calculate the average life time.<sup>6</sup>

$$y = a_1 \times e^{-\frac{x}{\tau_1}} + a_2 \times e^{-\frac{x}{\tau_2}} + y_0 \quad (1)$$

where  $a_1$  and  $a_2$  are weight constants,  $\tau_1$  and  $\tau_2$  are time decay constants, and  $y_0$  is an offset. The average carrier lifetime can be calculated by eqn (2)–(4).

$$A_1 (\%) = \frac{a_1}{a_1 + a_2} \times 100 \quad (2)$$

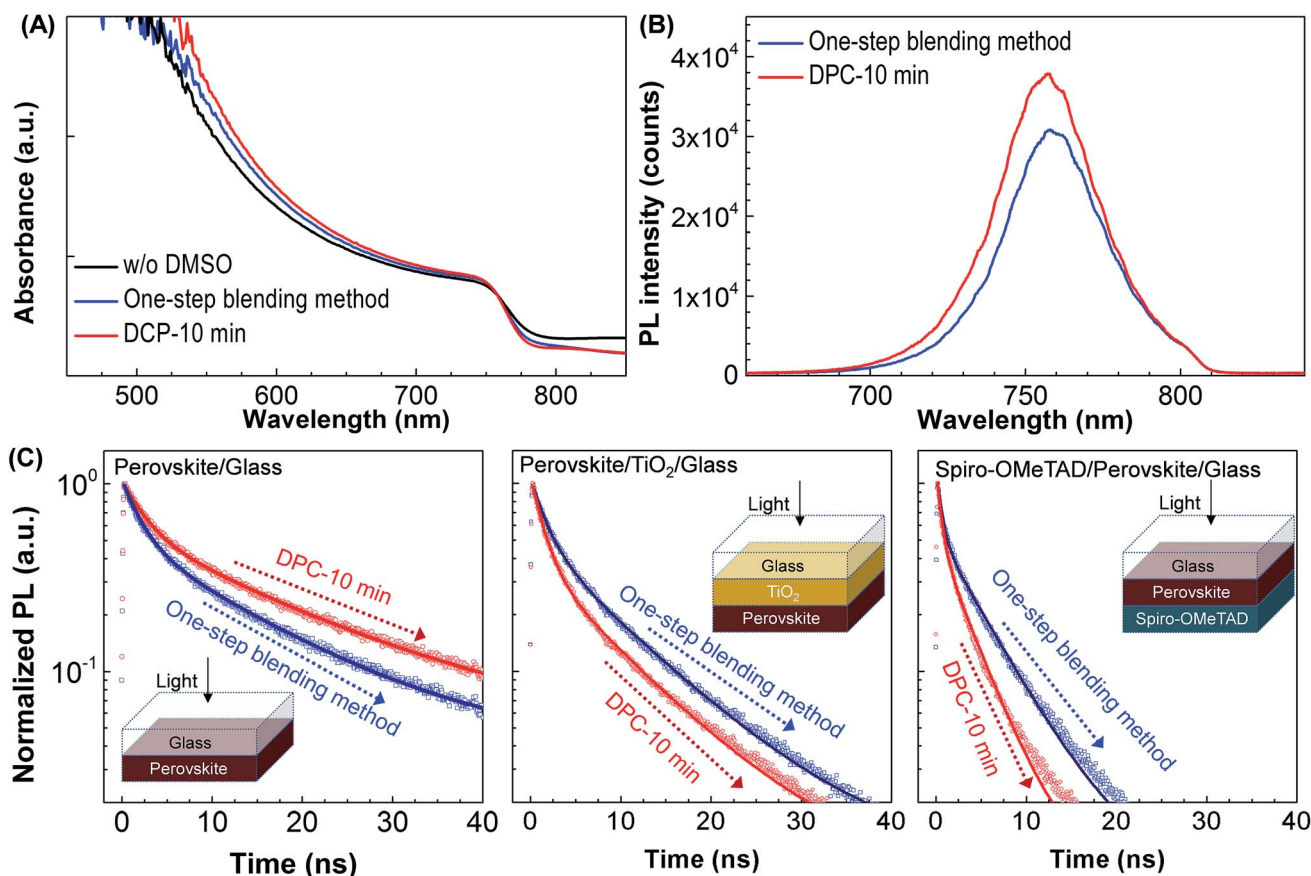


Fig. 3 Optical characteristics of the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  films prepared by the one-step blending and DPC methods. (A, B) UV-Vis absorption and photoluminescence spectra (perovskite/ $\text{TiO}_2$ /FTO/glass). (C) TRPL characteristics of perovskite/glass, perovskite/ $\text{TiO}_2$ /glass, and spiro-OMeTAD/perovskite/glass. The specimen was excited from the glass side by a pulsed laser (20 MHz, 407 nm,  $1 \mu\text{W cm}^{-2}$ ).



$$A_2 (\%) = \frac{a_2}{a_1 + a_2} \times 100 \quad (3)$$

$$\text{Average lifetime (ns)} = \frac{A_1 \tau_1^2 + A_2 \tau_2^2}{A_1 \tau_1 + A_2 \tau_2} \quad (4)$$

The detailed fitting parameters are provided in ESI Table S1.† The fast decay time constant ( $\tau_1$ ) is closely related with the non-radiative recombination by defective charge trap sites at grain boundaries, and the slow decay time constant ( $\tau_2$ ) is related with the intrinsic radiative recombination in bulk perovskite crystal.<sup>6</sup> The  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film synthesized by the DPC method exhibited a longer  $\tau_1$  (2.51 ns) implying reduced non-radiative recombination at grain boundaries and a longer  $\tau_2$  (17.08 ns) due to the improved bulk perovskite crystallinity.<sup>6</sup> Overall, the average carrier lifetime (15.34 ns) was 29% longer than that (11.93 ns) of the perovskite film synthesized by the one-step blending method. The perovskite film with an electron transporting layer (perovskite/ $\text{TiO}_2$ /glass) or a hole transporting layer (spiro-OMeTAD/perovskite/glass) showed much faster quenching behavior (Fig. 3C). Compared with the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film synthesized by the one-step blending method, all the time constants ( $\tau_1$ ,  $\tau_2$ , and average life time) were smaller for the perovskite film synthesized by the DPC method either with the electron or hole transporting layer (ESI Table S1†). This indicated a more efficient charge extraction due to the stronger

interfacial coupling with the electron and hole transporting layer.<sup>30</sup>

The increased carrier life time in the perovskite and efficient charge carrier extraction with the transporting layers resulted in enhanced photovoltaic performance as shown in Fig. 4. Fig. 4A compares the best  $J$ - $V$  curves of perovskite solar cells synthesized by the one-step blending and DPC methods. The  $J$ - $V$  data were measured at reverse scan direction. All the photovoltaic parameters ( $J_{\text{sc}} = 22.9 \text{ mA cm}^{-2}$ ,  $V_{\text{oc}} = 1.053 \text{ V}$ , FF = 75.5%, and PCE = 18.2%) were improved by the DPC method compared with those ( $J_{\text{sc}} = 22.3 \text{ mA cm}^{-2}$ ,  $V_{\text{oc}} = 1.037 \text{ V}$ , FF = 72.7%, and PCE = 16.9%) obtained by the one-step blending method. Fig. 4B shows the EQE spectrum of the best performing solar cells. The DPC method possessed a greater integrated  $J_{\text{sc}}$ , compared with the one-step blending method, which is consistent with the  $J$ - $V$  data in Fig. 4A. The hysteresis between the forward and reverse scan was decreased by the DPC method (ESI Fig. S6†). The decrease in charge trap sites, including pinholes, typically located at grain boundaries contributed to the decrease in hysteresis.<sup>33,36,37</sup> The average values and standard deviations of  $V_{\text{oc}}$ ,  $J_{\text{sc}}$ , FF, and PCE from 30 devices are compared in Fig. 4C. The DPC method showed a clear enhancement in photovoltaic performance ( $V_{\text{oc}} = 1.046 \pm 0.009 \text{ V}$ ,  $J_{\text{sc}} = 22.9 \pm 0.40 \text{ mA cm}^{-2}$ , FF =  $73.5 \pm 1.2\%$ , and PCE =  $17.6 \pm 0.37\%$ ), compared with the one-step blending method ( $V_{\text{oc}} = 1.027 \pm 0.014 \text{ V}$ ,  $J_{\text{sc}} = 22.30 \pm 0.46 \text{ mA cm}^{-2}$ , FF =  $69.1 \pm$

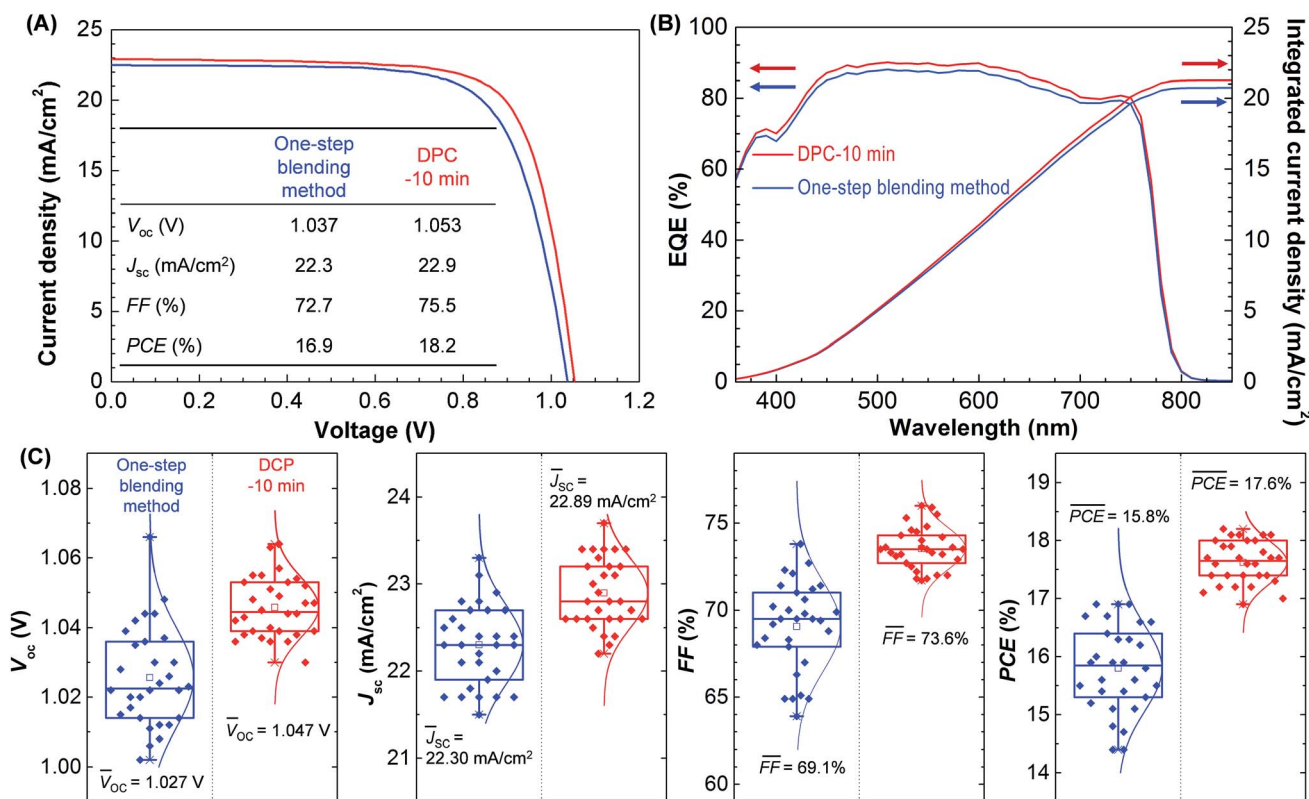


Fig. 4 Current density–voltage ( $J$ - $V$ ) characteristics of perovskite solar cells synthesized by the one-step blending and DPC methods. (A, B) The best  $J$ - $V$  curve and external quantum efficiency (EQE) spectrum with integrated  $J_{\text{sc}}$  obtained by each method. (C) The average values and standard deviations of  $V_{\text{oc}}$ ,  $J_{\text{sc}}$ , FF, and PCE from 30 devices.



2.6%, and PCE = 15.8 ± 0.74%). The average  $V_{oc}$ ,  $J_{sc}$ , FF, and PCE values were increased by 2, 3, 6, and 12%, respectively. The standard deviations of  $V_{oc}$ ,  $J_{sc}$ , FF, and PCE were decreased by 36, 13, 54, and 50% demonstrating enhanced reproducibility. The reduced grain boundary defects including pin holes enabled reproducible photovoltaic performance of perovskite solar cells.<sup>21–25</sup>

The nominal composition of  $\text{CH}_3\text{NH}_3\text{I}$ ,  $\text{PbI}_2$ , DMSO, and DMF was same in the precursor solutions for the one-step blending method and DPC method. However, the dominant  $\text{PbI}_2$ -DMSO<sub>1</sub> formation, after the mechanical mixing process in the DPC method, left only a small amount of uncoordinated  $\text{PbI}_2$  and  $\text{PbI}_2$ -DMSO<sub>2</sub> in the powder mixture (Fig. 1). This resulted in highly crystalline  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film with negligible residual  $\text{PbI}_2$ , negligible pin-holes, and large grains (Fig. 2). Consequently, the photovoltaic performance was enhanced compared with the one-step method (Fig. 4). Among the solar cells prepared by the DPC method, the photovoltaic performance was systematically enhanced from 15.3 to 18.2% as the mechanical mixing time increased from 3 to 10 min (ESI Fig. S6†). The average PCE also increased with a decrease in standard deviation. A longer mechanical mixing time (15 min) in the DPC method did not further improve the photovoltaic performance. It is interesting to note that an increased stirring time from 12 to 36 hours in the one-step blending method did not further increase the power conversion efficiency (ESI Fig. S7†). This indicated the importance of forming  $\text{PbI}_2$ -DMSO<sub>1</sub> in the powder mixture by the DPC method for the enhancement of photovoltaic performance.

As shown in Fig. 5, a decrease in series resistance ( $R_s$ ) and an increase in shunt resistance ( $R_{sh}$ ) were observed with smaller deviations for the solar cells fabricated by the DPC method.  $R_s$  and  $R_{sh}$  were calculated using the lumped equivalent circuit model.<sup>2,38,39</sup>

$$I = I_{ph} - I_0 \left\{ \exp \left[ \frac{q(V + IR_s)}{nkT} \right] - 1 \right\} - \frac{V + IR_s}{R_{sh}} \quad (5)$$

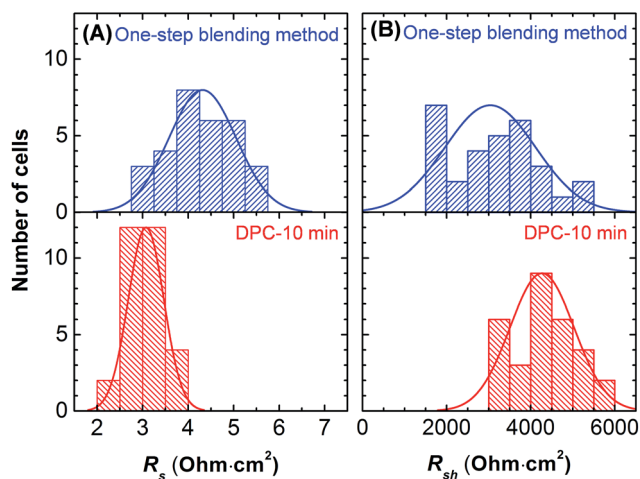


Fig. 5 The series (A) and shunt resistance (B) of perovskite solar cells prepared by the one-step blending and DPC methods. Thirty devices were fabricated for each type.

where  $I_{ph}$  is the photo-generated current delivered by constant current source,  $I_0$  is the saturation current of the equivalent diode,  $q$  is the elementary charge,  $k$  is the Boltzmann's constant,  $T$  is the cell temperature in kelvin,  $n$  is the diode ideality factor,  $I$  is the measured current, and  $V$  is the measured cell voltage.  $R_s$  and  $R_{sh}$  can be approximated by differentiating eqn (5) in case of negligible  $I_0$  and a significantly larger  $R_{sh}$  compared with  $R_s$ .<sup>38,39</sup>  $R_s$  was calculated by assuming  $n = 1$  (ref. 40) and  $T = 300$  K.

$$R_s = - \left( \frac{dV}{dI} \right)_{V=V_{oc}} - \frac{nkT}{q} \times \frac{1}{I_{sc}}, \quad R_{sh} = - \left( \frac{dV}{dI} \right)_{I=I_{sc}} \quad (6)$$

A more efficient charge carrier extraction with the transporting layers (Fig. 3C) resulted in a decrease in  $R_s$ . The increase in  $R_{sh}$  was attributed to the reduced grain boundary with pinhole free morphology (Fig. 2 and ESI S2†).<sup>37</sup> The dense  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film increased absorbance (Fig. 3a) and reduced

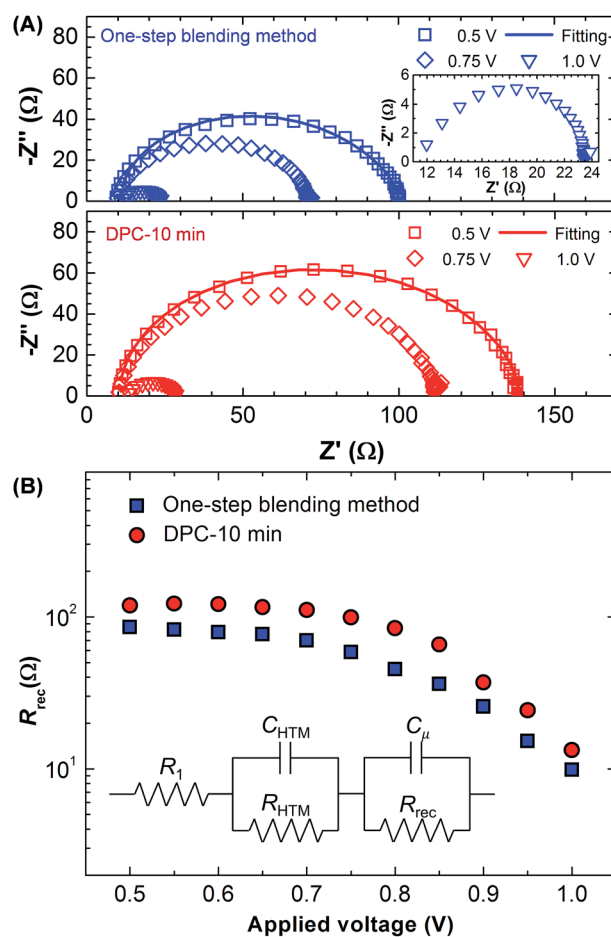


Fig. 6 Electrochemical impedance spectroscopy of perovskite solar cells prepared by the one-step blending and DPC methods. (A) Nyquist plots measured at the applied bias of 500, 750, and 1000 mV (100 mW  $\text{cm}^{-2}$  illumination). The low impedance region is magnified in the inset. (B) The recombination resistance is shown as a function of the applied bias. The equivalent circuit is provided in the inset.  $R_1$  is the series resistance of cell,  $C_\mu$  is the chemical capacitance, and  $R_{rec}$  is the recombination resistance.  $C_{HTM}$  and  $R_{HTM}$  are the capacitance and resistance of hole transport materials.



**Table 1** The key parameters used for fitting the EIS data in the equivalent circuit model. The applied voltage was 500 mV

Method	$R_1$ ( $\Omega$ )	$R_{\text{HTM}}$ ( $\Omega$ )	$C_{\text{HTM}}$ ( $\mu\text{F}$ )	$R_{\text{rec}}$ ( $\Omega$ )	$C_{\mu}$ ( $\mu\text{F}$ )
One-step	10.85	9.79	1.31	79.1	0.03
DPC-10 min	10.48	8.72	1.57	119.1	0.03

recombination sites. Overall, this lead to an impressive increase in FF and PCE.

The EIS analysis was performed to measure the recombination resistance ( $R_{\text{rec}}$ ).<sup>27,41–43</sup> Nyquist plots for the DPC and one-step blending methods were compared at the applied bias of 500, 750, and 1000 mV as shown in Fig. 6A. The frequency range was 1–10<sup>6</sup> Hz. The arc feature consists of two semicircles. The high frequency semicircle corresponds to the resistance of hole transport materials ( $R_{\text{HTM}}$ ), and the dominant low frequency feature is related with  $R_{\text{rec}}$ .<sup>27,42,43</sup> The arc decreased as the applied bias increased. The DPC method exhibited a larger arc than the one-step blending method.  $R_{\text{rec}}$  was calculated using the equivalent circuit model shown in the inset of Fig. 6A. The same equivalent circuit model has been widely employed to describe the EIS data of perovskite solar cells prepared by the solution process.<sup>27,41–43</sup> There was an excellent fitting with the experimental data, and key fitting parameters are provided in Table 1. Compared with the one-step blending method, the DPC method exhibited enhanced  $R_{\text{rec}}$  at the entire bias range indicating suppressed charge recombination (Fig. 6B). This contributed to the higher  $J_{\text{sc}}$  and  $V_{\text{oc}}$  in the DPC method.<sup>44</sup>  $R_{\text{rec}}$  started to decrease over 0.7 V, and the slope was steeper for the DPC method. This indicated a lower diode ideality factor for the DPC method enhancing FF.<sup>43,45</sup> Overall, the EIS analysis was consistent with the  $J$ – $V$  characteristics (Fig. 4).

The statistical photovoltaic performance analysis demonstrated a clear advantage of the pre-coordinated  $\text{PbI}_2$ –DMSO powders prepared by the simple mechanical mixing method. A further optimization of the  $\text{TiO}_2$  structures, solvents, materials, and detailed processing parameters may enhance the photovoltaic performance of solar cells fabricated by the DPC method in the future.

## 4. Conclusions

The pre-coordinated  $\text{PbI}_2$ –DMSO powders enhanced the photovoltaic performance of planar heterojunction perovskite solar cells fabricated by the low temperature process ( $\leq 100$  °C). The pre-coordination was carried out by the simple mechanical mixing of DMSO and  $\text{PbI}_2$  particles (1 : 1 mol%) using a mortar and pestle. The dominant composition of  $\text{PbI}_2$ –DMSO<sub>1</sub> was obtained after the 10 min mixing process. The  $\text{PbI}_2$ –DMSO powders were then mixed with  $\text{CH}_3\text{NH}_3\text{I}$  in DMF to make the  $\text{CH}_3\text{NH}_3\text{PbI}_3$  film by the toluene-assisted spin-coating and heat treatment. This method resulted in better dissolution of  $\text{PbI}_2$ , larger grain size, and pinhole free morphology. The TRPL analysis revealed an increased charge carrier life time and better charge extraction. Overall, the average  $V_{\text{oc}}$  (1.046 V),  $J_{\text{sc}}$  (22.9 mA  $\text{cm}^{-2}$ ), FF (73.5%), and PCE (17.6%) were enhanced by 2, 3, 6,

and 12%, respectively, compared with the one-step mixing method. The standard deviations were decreased indicating improved reproducibility. The simple mechanical pre-coordination of  $\text{PbI}_2$ –DMSO powders effectively enhanced crystallinity of  $\text{CH}_3\text{NH}_3\text{PbI}_3$  leading to the enhanced photovoltaic performance with the best PCE of 18.2%.

## Conflicts of interest

There are no conflicts to declare.

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## Notes and references

- 1 Q. Zheng, H. Kang, J. Yun, J. Lee, J. H. Park and S. Baik, *ACS Nano*, 2011, **5**, 5088–5093.
- 2 J. Lee, M. M. Menampambath, J. Y. Hwang and S. Baik, *ChemSusChem*, 2015, **8**, 2358–2362.
- 3 J. H. Heo, S. H. Im, J. H. Noh, T. N. Mandal, C.-S. Lim, J. A. Chang, Y. H. Lee, H. Kim, A. Sarkar, M. K. Nazeeruddin, M. Grätzel and S. I. Seok, *Nat. Photonics*, 2013, **7**, 486–491.
- 4 B. O'Regan and M. Grätzel, *Nature*, 1991, **353**, 737–740.
- 5 W. S. Yang, J. H. Noh, N. J. Jeon, Y. C. Kim, S. Ryu, J. Seo and S. I. Seok, *Science*, 2015, **348**, 1234–1237.
- 6 D. Y. Son, J. W. Lee, Y. J. Choi, I. H. Jang, S. Lee, P. J. Yoo, H. Shin, N. Ahn, M. Choi, D. Kim and N. G. Park, *Nat. Energy*, 2016, **1**, 1–8.
- 7 M. Saliba, T. Matsu, J. Y. Seo, K. Domanski, J. P. Correa-Baena, M. K. Nazeeruddin, S. M. Zekeeruddin, W. Tress, A. Abate, A. Hagfeldt and M. Grätzel, *Energy Environ. Sci.*, 2016, **9**, 1989–1997.
- 8 J. W. Lee, D. J. Seol, A. N. Cho and N. G. Park, *Adv. Mater.*, 2014, **26**, 4991–4998.
- 9 D. Zhong, B. Cai, X. Wang, Z. Yang, Y. Xing, S. Miao, W. H. Zhang and C. Li, *Nano Energy*, 2015, **11**, 409–418.
- 10 M. S. Seo, I. Jeong, J. S. Park, J. Lee, I. K. Han, W. I. Lee, H. J. Son, B. H. Sohn and M. J. Ko, *Nanoscale*, 2016, **8**, 11472–11479.
- 11 Q. Chen, H. Zhou, Y. Fang, A. Z. Stieg, T. B. Song, H. H. Wang, X. Xu, Y. Liu, S. Lu, J. You, P. Sun, J. McKay, M. S. Goorsky and Y. Yang, *Nat. Commun.*, 2015, **6**, 7269.
- 12 H. Zhou, Q. Chen, G. Li, S. Luo, T. Song, H. S. Duan, Z. Hong, J. You, Y. Liu and Y. Yang, *Science*, 2014, **345**, 542–546.
- 13 A. Yella, L. P. Heinger, P. Gao, M. K. Nazeeruddin and M. Grätzel, *Nano Lett.*, 2014, **14**, 2591–2596.
- 14 H. Kim, K. G. Lim and T. W. Lee, *Energy Environ. Sci.*, 2016, **9**, 12–30.



- 15 M. Liu, M. B. Johnston and H. J. Snaith, *Nature*, 2013, **501**, 395–398.
- 16 Q. Chen, H. Zhou, Z. Hong, S. Luo, H. S. Duan, H. H. Wang, Y. Liu, G. Li and Y. Yang, *J. Am. Chem. Soc.*, 2014, **136**, 622–625.
- 17 S. Bai, N. Sakai, W. Zhang, Z. Wang, J. T. W. Wang, F. Gao and H. J. Snaith, *Chem. Mater.*, 2017, **29**, 462–473.
- 18 G. E. Eperon, V. M. Burlakov, P. Docampo, A. Goriely and H. J. Snaith, *Adv. Funct. Mater.*, 2014, **24**, 151–157.
- 19 B. Conings, L. Baeten, C. D. Dobbelaere, J. D'Hean, J. Manca and H. G. Boyer, *Adv. Mater.*, 2014, **26**, 2041–2046.
- 20 Y. J. Jeon, S. Lee, R. Kang, J. E. Kim, J. S. Yeo, S. H. Lee, S. S. Kim, J. M. Yun and D. Y. Kim, *Sci. Rep.*, 2014, **4**, 6953.
- 21 N. J. Jeon, J. H. Noh, Y. C. Kim, W. S. Yang, S. Ryu and S. I. Seok, *Nat. Mater.*, 2014, **13**, 897–903.
- 22 N. Ahn, D. Y. Son, I. H. Jang, S. M. Kang, M. Choi and N. G. Park, *J. Am. Chem. Soc.*, 2015, **137**, 8695–8699.
- 23 W. Li, J. Fan, J. Li, Y. Mai and L. Wang, *J. Am. Chem. Soc.*, 2015, **137**, 10399–10405.
- 24 Y. Wu, A. Islam, X. Yang, C. Qin, J. Liu, K. Zhang, W. Peng and L. Han, *Energy Environ. Sci.*, 2014, **7**, 2934–2938.
- 25 X. Cao, C. Li, Y. Li, F. Fang, X. Cui, Y. Yao and J. Wei, *Nanoscale*, 2016, **8**, 19804–19810.
- 26 X. Li, D. Bi, C. Yi, J. D. Décoppet, J. Luo, S. M. Zakeeruddin, A. Hagfeldt and M. Grätzel, *Science*, 2016, **353**, 58–62.
- 27 N. D. Pham, V. T. Tiong, P. Chen, L. Wang, G. J. Wilson, J. Bell and H. Wang, *J. Mater. Chem. A*, 2017, **5**, 5195–5203.
- 28 A. Wakamiya, M. Endo, T. Sasamori, N. Tokitoh, Y. Ogomi, S. Hayase and Y. Murata, *Chem. Lett.*, 2014, **43**, 711–713.
- 29 M. Ozaki, Y. Katsuki, J. Liu, T. Handa, R. Nishikubo, S. Yakumaru, Y. Hashikawa, Y. Murata, T. Saito, Y. Shimakawa, Y. Kanemitsu, A. Saeki and A. Wakamiya, *ACS Omega*, 2017, **2**, 7016–7021.
- 30 I. Wharf, T. Gramstad, R. Makhija and M. Onyszczuk, *Can. J. Chem.*, 1976, **54**, 3430–3438.
- 31 J. Muhammad, J. H. Heo, C. E. Song, K. Kong, W. S. Shin, J. Lee, S. H. Im and S. Moon, *Nano Energy*, 2016, **27**, 330–339.
- 32 C. Roldán-Carmona, P. Gratia, I. Zimmermann, G. Grancini, P. Gao, M. Grätzel and M. K. Nazeeruddin, *Energy Environ. Sci.*, 2015, **8**, 3550–3556.
- 33 J. Xu, A. Buin, A. H. Ip, W. Li, O. Voznyy, R. Comin, M. Yuan, S. Jeon, Z. Ning, J. J. McDowell, P. Kanjanaboos, J. P. Sun, X. Lan, L. N. Quan, D. H. Kim, L. G. Hill, P. Maksymovych and E. H. Sargent, *Nat. Commun.*, 2015, **6**, 7081.
- 34 F. Wang, A. Shimazaki, F. Yang, K. Kanahashi, K. Matsuki, Y. Miyauchi, T. Takenobu, A. Wakamiya, Y. Murata and K. Matsuda, *J. Phys. Chem. C*, 2017, **121**, 1562–1568.
- 35 S. D. Stranks, G. E. Eperon, G. Grancini, C. Menelaou, M. J. P. Alcocer, T. Leijtens, L. M. Herz, A. Petrozza and H. J. Snaith, *Science*, 2013, **342**, 341–344.
- 36 W. Nie, H. Tsai, R. Asadpour, J. C. Blancon, A. J. Neukirch, G. Gupta, J. J. Crochet, M. Chhowalla, S. Tretiak, M. A. Alam, H. L. Wang and A. D. Mohite, *Science*, 2015, **347**, 522–525.
- 37 W. Qiu, T. Merckx, M. Jaysankar, C. M. Huerta, L. Rakocevic, W. Zhang, U. W. Paetzold, R. Gehlhaar, L. Froyen, J. Poortmans, D. Chenyns, H. J. Snaith and P. Heremans, *Energy Environ. Sci.*, 2016, **9**, 484–489.
- 38 J. Lee, H. Kang, J. Y. Hwang, S. W. Kim and S. Baik, *Carbon*, 2014, **79**, 337–345.
- 39 N. Koide, A. Islam, Y. Chiba and L. Han, *J. Photochem. Photobiol., A*, 2006, **182**, 296–305.
- 40 J. H. Im, I. H. Jang, N. Pellet, M. Grätzel and N. G. Park, *Nat. Nanotechnol.*, 2014, **9**, 927–932.
- 41 C. Chen, Y. Zhai, F. Li, F. Tan, G. Yue, W. Zhang and M. Wang, *J. Power Sources*, 2017, **341**, 396–403.
- 42 H. Kim, J. Lee, N. Yantara, P. P. Boix, S. A. Kulkarni, S. Mhaisalkar, M. Grätzel and N. G. Park, *Nano Lett.*, 2013, **13**, 2412–2417.
- 43 J. A. Christians, R. C. M. Fung and P. V. Kamat, *J. Am. Chem. Soc.*, 2014, **136**, 758–764.
- 44 F. Fabregat-Santiago, G. Garcia-Belmonte, I. Mora-Seró and J. Bisquert, *Phys. Chem. Chem. Phys.*, 2011, **13**, 9083–9118.
- 45 Y. Yang, K. Ri, A. Mei, L. Liu, M. Hu, T. Liu, X. Li and H. Han, *J. Mater. Chem. A*, 2015, **3**, 9103–9107.

