Synthesis and properties of a POM-based trinuclear copper(II) triazole framework†

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A novel POM-based trinuclear copper(II) triazole framework, namely, \( \left[ \text{H}_2\text{Cu}_6(\text{trz})_6(\mu_3-\text{OH})_3\text{Mo}_5\text{O}_{18}\right]\cdot3\text{H}_2\text{O} \) (1) was isolated using a hydrothermal method, which displays a 3D network constructed from trinuclear copper(II) units and triazole ligands with \( \text{Mo}_5\text{O}_{18} \) anions as templates. 1 has been identified by single crystal X-ray diffraction, elemental analysis, thermogravimetric analysis, powder X-ray diffraction and FT-IR. Magnetic studies indicate that antiferromagnetic interactions exist in 1. In addition, 1 exhibits good Lewis acid catalytic activity for the synthesis of cyclohexanone ethylene ketal with 95% conversion. The HOMO–LUMO gap \( (E_g) \) of 1 is 2.34 eV calculated using the Kubelka–Munk equation \((Fh\nu)^0.5 \), indicating that its forbidden bandwidth belongs to the semiconductor category. Visible-light photodegradation of RhB catalyzed by 1 was investigated, which shows high activity with an above 98% degradation rate.

Introduction

Polyoxometalates (POMs), as a class of nanosized metal-oxo clusters composed of group VI (Mo and W) and group V (V, Nb and Ta) elements in their highest oxidation state, have received increasing attention owing to their aesthetically appealing structures and wide applications such as in the fields of magnetism,¹ and catalysis.² Due to their strong acidity and fast multielectron transformation activity, POMs can be employed as acid, redox and bifunctional (acid and redox) catalysts in many synthetically useful organic reactions. Moreover, they can act as excellent photocatalysts in the oxidation of organic compounds owing to their ability as electron and oxygen relays.³ Trinuclear triangular copper(n) (TTC) clusters with the formula \( \{\text{Cu}_3(\mu_1\text{-OH})(\text{L})_3\} \) (L = pyrazole, imidazole, triazole, terazole and the derivatives of them) possess a planar structure composed of three copper centers held together by a central μ₁-O atom.⁴ In the TTC cluster, each Cu(n) center is coordinately unsaturated with open coordination sites, which is a potential active site in Lewis acid and oxidation catalytic reactions.⁵ Designing MOFs based on TTC units have attracted great attention of researchers owing to their promising applications in catalysis.⁶ For examples, Zhang et al.⁷ synthesized a active-site-accessible porous MOF using TTC as SBU, which shows high activity for the CO catalytic oxidation, giving a conversion rate of 100%. Mondal et al.⁸ reported a series of MOFs based on TTC units, which show high activity in the photocatalytic degradation of toxic dye molecules. Furthermore, due to their interesting magnetic properties, TTC clusters have attracted attention in the construction of magnetic high-nuclear clusters and MOFs materials. For an example, Ouellette et al.⁹ reported a microporous metal–organic framework constructed with magnetic \( \{\text{Cu}_3(\mu_2\text{-OH})(\text{trz})_3\} \) building blocks.⁷

Inspired by the properties of POMs and TTC clusters, we explored to encapsulate POMs into the host matrix of MOFs built by TTC units. Firstly, due to the promising catalytic performances of both POMs and TTC units, the obtained POMs-based TTC frameworks may show enhanced catalytic activity by the combination and synergies of the functions of POMs and TTC units. Secondly, the inlay of POMs into the host matrix of MOFs built by TTC units would effectively enhance the stability of both POMs and TTC frameworks, which is an essential prerequisite for practical applications. Finally, the combination of magnetic TCC units and POMs, interesting magnetic properties may be evoked. Although some POM-based MOFs have been synthesized, to our knowledge, the POM-based MOFs combined POM and TCC units have never been reported. Herein, a POM-based trinuclear–copper(n) triazole framework, \( \{\text{H}_2\text{Cu}_6(\text{trz})_6(\mu_3\text{-OH})_3\text{Mo}_5\text{O}_{18}\}\cdot3\text{H}_2\text{O} \) (1) \((\text{trz} = 1,2,4-	ext{triazole})\), has been successfully synthesized under hydrothermal condition. The catalytic performance of 1 was assessed by choosing the synthesis of cyclohexanone ethylene ketal catalyzed by Lewis acid and visible-light photodegradation of RhB as testing reactions. In addition, the magnetic property of 1 has been investigated.
Experimental

Materials and methods

(TBA)$_2$[Mo$_6$O$_{19}$] (TBA = tetrabutylammonium) was prepared according to the literature procedures. Other chemicals were obtained commercially and used without further purification. Crystal morphology was characterized by using a Quanta 250 FEG scanning electron microscope. Fourier transform infrared spectroscopies (FT-IR) were recorded in the range 400–4000 cm$^{-1}$ on an Alpha Centaurt FT-IR spectrophotometer using KBr pellets. Single crystal XRD analysis of 1 was conducted on a Bruker Smart Apex CCD diffractometer with Mo K{$\alpha$} monochromated radiation ($\lambda = 0.71073$ Å) at room temperature. Powder X-ray diffraction (PXRD) measurements were performed on a Rigaku D/MAX-3 instrument with Cu K{$\alpha$} ($\lambda = 1.5418$ Å) radiation in the angular range 20 3–50° at 293 K. Thermogravimetric analysis (TGA) was performed on a Perkin-Elmer TGA7 instrument, with a heating rate of 10 °C min$^{-1}$ under a nitrogen atmosphere. The magnetic susceptibility data were obtained on a SQUID magnetometer (Quantum Design, MPMS-5) in the temperature range 2–300 K with applied field of 1000 Oe. GC analyses were performed on an Agilent 6820 instrument with a flame ionization detector.

Procedure for catalytic condensation of cyclohexanone and glycol

The ketalization reaction was performed in a three-necked round-bottom flask with a given mass of 1. Then, reasonable volume ratio of cyclohexanone and ethylene glycol, and cyclohexane were added with magnetic stirring, the device was equipped with thermometer refluxing condenser tube and water segregator in atmospheric pressure in 3 hours. The solution was heated to boil and refluxed until no water flowed off. The resultant solution was distilled and purified. Samples should be taken every half an hour, and the sample should be diluted with acetone, the yield of cyclohexanone ethylene ketal can be measured by GC analysis with sample diluted by acetone.

Procedure for photocatalysis of RhB

The photodegradation of RhB is evaluated under Xe light irradiation from high pressure Xe lamp through a typical process: 10 mg of the compound 1 was mixed together with 100 mL of 10 mg L$^{-1}$ RhB solution in a beaker by ultrasonic dispersion for 30 min until reached the surface adsorption equilibrium on the particles of compound 1. Then, the mixture was stirred continuously under visible light irradiation. At 0, 20, 40, 60, 90, 120, 150, 180 and 200 min, 3 mL of the sample was taken out from the beaker respectively, followed by centrifugation to remove 1 and a clear solution was obtained for UV-vis analysis. The UV-vis diffuse reflectance (DR) spectrum of 1 was recorded between 380 and 550 nm. The photodecomposition rate was calculated by the formula: $D$ (%) = ($A_0 - A_t$)$/A_0 \times 100$%; $D$ is the degradation rate; $A_0$ is the initial absorption; $A_t$ is the concentration after a period of reaction time.

Synthesis of 1

A mixture of Na$_2$MoO$_4$·2H$_2$O (1.0 mmol, 0.242 g), (TBA)$_2$[Mo$_6$O$_{19}$] (0.5 mmol, 0.638 g), CuCl$_2$·2H$_2$O (0.341 g, 2.0 mmol), trz (2.0 mmol, 0.138 g) was added to distilled water (10 mL) with stirring. Then the solution was stirred for 0.5 h and the pH was adjusted to 5.2 with 1 M HCl solution. The resulting suspension was transferred to a Teflon-lined stainless-steel autoclave and sealed. After being heated at 170 °C for five days, it was cooled to room temperature at a rate of 10 °C h$^{-1}$. Yellow block crystals of 1 were separated, washed with water and air-dried. Although (TBA)$_2$[Mo$_6$O$_{19}$] does not appear in the system, but he is essential in the reaction, because we tried not to add it in the system, has not been a product, it is due to (TBA)$_2$[Mo$_6$O$_{19}$] was decomposed into MoO$_6$ polyhedron in the reaction system, then, MoO$_6$ polyhedron reactant with the MoO$_4$$^{2-}$ get the target product 1. Yield: 0.19 g (30% based on Mo). Anal. calcd (%) for 1: C 8.76, H 1.23, N 15.32, Cu 23.20, Mo 29.14; found: C 8.80, H 1.22, N 15.33, Cu 23.17, Mo 29.16. IR (solid KBr pellet): 3268(s), 2932(s), 1559(w), 1373(w), 1221(w), 1045.99(w), 936(s), 823(s), 714(s), 526(s), 542(m), 409 (m) cm$^{-1}$.

Results and discussion

Synthesis and crystal structure

The reaction of Na$_2$MoO$_4$·2H$_2$O, (TBA)$_2$[Mo$_6$O$_{19}$], CuCl$_2$·2H$_2$O, trz and water under hydrothermal conditions generated yellow block crystals of 1. Notably, the formation of 1 was highly sensitive to the pH value of the reaction system, which was controlled in the range of 5.2 to 5.5, and 1 could not be obtained outside this pH range. Although (TBA)$_2$[Mo$_6$O$_{19}$] does not appear in the composition of 1, it is essential for the synthesis of 1. We cannot obtain 1 in the absence of (TBA)$_2$[Mo$_6$O$_{19}$]. It may be due to (TBA)$_2$[Mo$_6$O$_{19}$] decomposed into MoO$_6$ polyhedra in the reaction system, then MoO$_6$ polyhedra react with MoO$_4$$^{2-}$ yielding the target product 1. The SEM image showed that 1 was block crystal and its surface was slippy (Fig. S2, ESI†). The PXRD pattern of 1 and the simulated PXRD pattern are shown in Fig. S3, ESI.† The diffraction peaks on the patterns correspond well in position, confirming that the product is in a pure phase. 1 displayed a trinuclear copper triazole framework with [MoO$_3$]$_{3-}$ anions as templates. There are three Cu$^{2+}$ ions, six trz ligands, one [Mo$_5$O$_{18}$]$^{6-}$ anion and one coordination water molecule in the basic structural unit of 1. Three Cu$^{2+}$ centres are connected by a μ$_{3}$-O and each pair of adjacent Cu$^{2+}$ was connected by a triazole ligand to form a three core triangular Cu$_3$(μ$_2$-OH)[trz]$_3$ unit (Fig. 1). In the trinuclear copper unit, three Cu$^{2+}$ centres display distinct coordination environments: Cu$_1$ was coordinated by the μ$_1$-O atom, O atom from [Mo$_5$O$_{18}$]$^{6-}$ anions and three N from trz; Cu$_2$, was connected with the μ$_1$-O atom, three N atoms from trz ligands and one Ow atom, Cu$_3$ was coordinated by the μ$_1$-O atom, one O atom from a [Mo$_5$O$_{18}$]$^{6-}$ anion and three N atoms from trz. The average Cu–O and Cu–N bond distances are 2.151 Å and 1.994 Å, respectively. [Mo$_5$O$_{18}$]$^{6-}$ anion is constructed by three MoO$_6$ octahedra and two MoO$_3$ polyhedra by corner-sharing connections with Mo–O bond lengths in the range of 1.707–2.606 Å. In the structure of 1, each
Chemical stability

1 was air-stable because no efflorescence of the crystal was observed. After being steeped in common organic solvents (such as dimethylformamide, dichloromethane, methanol, acetone, hexanone and glycol) for 5 days at room temperature, there were no changes in its PXRD patterns demonstrated that 1 was stable in these organic solvents (Fig. 2a). Furthermore, 1 was stable in acid and basic aqueous solution with pH range from 2 to 12 (Fig. 2b). In order to compare the chemical stability of 1, K₃HMo₅O₁₇ and Cu₃(trz)₃(OH)₃(H₂O)₄ were measured the PXRD patterns of K₃HMo₅O₁₇ and Cu₃(trz)₃(OH)₃(H₂O)₄ under the same condition (Fig. S9, ESI†). The result showed that K₃HMo₅O₁₇ was not stable in glycol and aqueous solution with pH ≤ 10. Cu₃(trz)₃(OH)₃(H₂O)₄ was stable in all of these organic solvents, but it was not stable in aqueous solution with pH ≥ 10. The above results show that the chemical stability of 1 is better than K₃HMo₅O₁₇ and Cu₃(trz)₃(OH)₃(H₂O)₄, which is attributed to the structure of 1, the introduction of POM into TTC MOFs making 1 more stable.

UV-vis spectroscopy

The UV-vis diffuse reflectance absorption spectrum of compound 1, K₃HMo₅O₁₇ and Cu₃(trz)₃(OH)₃(H₂O)₄ were tested as presented in Fig. 3, S10 and S11, ESI†. The result showed that their Eₙ were 2.34 eV, 2.87 eV and 1.67 eV, respectively. We take 1 as model to show the relationship of UV-vis diffuse reflectance absorption spectrum and Eₙ. In the ultraviolet range, the lowest energy band at 265 nm can be assigned to the charge transfer transition from O → Mo, the higher energy band at 780 nm arises from the d–d electronic transition of the Cu centers. The optical band gaps (Eₙ) of 1 have been calculated using Tauc’s relation:14 (αhv) = A(hv – Eₙ)²; where α, r, h, n, Eₙ and A stand for absorption coefficient, frequency, Planks’ constant, transition coefficient, band gap and band tailing parameter, respectively. Among them, A is a constant and α is the optical absorption coefficient given by the Kubelka–Munk function: F(R∞) = (1 – R∞)/(2R∞), R∞ = (Iₒ/Iₐ)diff is the diffuse reflectivity. The index n is a number which characterizes the transition process involved. The intercept of the linear plots (αhv)¹⁰ with the E-axis
gives a direct transition ($n = 1/2$) at 2.34 eV (Fig. 3). The results demonstrate the semiconductor nature for 1.

**Magnetic property**

The temperature dependence of magnetic susceptibility of 1 was investigated at 2–300 K in an applied magnetic field of 1000 Oe. The $\chi_M$ versus $T$ plot demonstrated the $\chi_M$ value slowly increases from 0.0071 emu mol$^{-1}$ at 300 K to 0.045 emu mol$^{-1}$ at 2 K. The $\chi_M T$ value of 1 is 2.13 emu K mol$^{-1}$ at 300 K (Fig. 4), which is larger than the theoretical value of 1.125 emu K mol$^{-1}$ expected from three isolated Cu$^{2+}$ ($S = 0.5, g = 2.0$). As the temperature is lowered, the $\chi_M T$ value decreases slowly up to 50 K and then sharply falls to a minimum value of 0.94 emu K mol$^{-1}$ at 2 K. The temperature dependences of $\chi_M$ and $\chi_M T$ versus $T$ plots for 1 suggest the presence of antiferromagnetic coupling interactions between Cu$^{2+}$ centers. The $\chi_M^{-1}$ versus $T$ plot follows the Curie–Weiss law with $C = 3.56$ emu K mol$^{-1}$ and $\theta = -29.97$ K in the range of 48–300 K (Fig. S12, ESI†), which further demonstrated the presence of antiferromagnetic interactions between Cu$^{2+}$ ions in 1. The result was correspondent with the previously report$^6$ about TTC compound, antiferromagnetic interactions existing in the TTC core.

**Catalysis for condensation reaction of cyclohexanone and glycol**

In many organic reactions, carbonyl protection reaction is one of the most important intermediate processes. Cyclohexanone reacts with glycol to form cyclohexanone ethylene ketal is a typical carbonyl protection reaction, in which acid catalyst plays a significant role.$^{14}$ Herein, we chose the synthesis of cyclohexanone ethylene ketal as a model reaction to research the acid catalytic activity of 1. With single factor level analysis, we obtained the optimum reaction conditions of yields of ketals: reaction time were 3.0 hours, cyclohexanone/glycol molar ratio is 1 : 1.5, the amount of acid catalyst 1 was 0.1 g, the most suitable volume of water-carrying agent cyclohexane was 16 mL, reaction temperature was allowed to reflux at 100–105 °C (Fig. 5). Under the optimum experimental conditions, the yield of cyclohexanone ketal was 95%. The ketal product was characterized by FT-IR (Fig. S13, ESI†) and PXRD (Fig. S14, ESI†). After 1 h of reaction, the catalyst was removed through hot filtration and the filtrate was further heated for 2 h at 100 °C. No further growth of the conversion (Fig. 5e) indicated.
that 1 was a heterogeneous catalyst in the catalytic reaction and was not dissolved in the reaction system. 1 was recovered by centrifugation after 3 h of reaction. The catalyst could be reused for at least 5 times, without significant loss of catalytic activity (Fig. 5f). Additionally, as exhibited in Table 1, compound 1 showed better catalytic activity than the parent materials K₃Mo₅O₁₇ and Cu₃(trz)₃(OH)₃(H₂O)₄, in the presence of which the yield of cyclohexanone ethylene ketal was 95%, 44% and 36%, respectively (Table 1). The reaction mechanism of the condensation of cyclohexanone and glycol catalyzed by 1 was Lewis acid active sites, which come from unsaturated coordination sites of Cu centers attack on the carbonyl carbon forming C⁺ ion, and then C⁺ ion accepts the attack of nucleophile glycol forming ketals (Fig. S15†).

**Photocatalysis property**

POMs-based coordination polymer was introduced transition metal and organic ligand as functional groups, which could enrich the potential applications of POMs, such as catalysis and conductivity. Some POMs-based coordination polymers have been proven to be an excellent kind of green and cheap photocatalyst and can be used to photodecomposition of organic dyes.¹³ Liu et al.¹³a synthesized a POMs-based coordination polymer [WO₄Ni(en)₂(H₂O)₃][Ni(en)₃][P₂W₁₈O₆₂]·[Ni(en)₃]·CO₃·H₂O (en = ethanediamine), which achieved 97% degradation rate for the photocatalytic degradation of methylene blue (MB) under visible light irradiation. Luo et al.¹³b reported a POMs-based coordination polymer {Cu₃(L)₃[CrMo₅O₁₇H]·O₁₅}(H₂O)₃·2H₂O (L = N₂N'-bis(3-pyridinecarboxamide)-1,3-propane), which exhibited photocatalytic activity on the photodecomposition of RhB under visible light with 97.6% degradation rate. Herein, we investigated the photodecomposition activity of 1 for RhB. We chose visible light as the light source, because the calculated HOMO-LUMO gap (Eₒ) of compound 1 was 2.34 eV, which corresponds to visible light absorption. The UV-vis diffuse reflectance spectrum of 1 was presented in Fig. 3 and S11, ESI†.

The photodegradation rate of RhB by 1 was 98% in 200 min. After reuse five times, the result did not change evidently (Fig. 6a and c). For a comparative purpose, the photodecomposition rate in the presence of (TBA)₃[Mo₅O₁₇H], Cu₃(trz)₃(OH)₃(H₂O)₄ and blank were only 37%, 13% and 8% respectively (Fig. 6b), which were much lower than that by 1.

We proposed that the photodegradation of RhB by 1 was realized by the synergy effect of Cu₃(trz)₃(OH)₃²⁺ and Mo₅O₁₇⁻ units of 1. The maximum absorbance wavelength of UV-vis diffuse reflectance spectrum of (TBA)₃[Mo₅O₁₇H] and Cu₃(trz)₃(OH)₃(H₂O)₄ were 280 nm and 520 nm, respectively (Fig. S9 and b, ESI†). Therefore, Cu₃(trz)₃(OH)₃²⁺ acts as sensitizer absorbing visible light to produce photo-introduced electron. The photo-introduced electron is transferred to the lower LUMO orbit of Mo₅O₁₇⁻ and then combined with O₂⁻ radical, which acts as oxidizing agent to effectively attack and oxidize RhB molecule. The proposed mechanism was showed in Fig. 7.

### Table 1 Catalytic performances of 1 compared with the reference catalysts for the synthesis of cyclohexanone ethylene ketal

<table>
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<tr>
<th>Entry</th>
<th>Catalyst</th>
<th>Time/h</th>
<th>Yield/%</th>
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<tr>
<td>1</td>
<td>K₃HMnO₅₇</td>
<td>3</td>
<td>44</td>
</tr>
<tr>
<td>2</td>
<td>Cu₃(trz)₃(OH)₃(H₂O)₄</td>
<td>3</td>
<td>36</td>
</tr>
<tr>
<td>3</td>
<td>Compound 1</td>
<td>3</td>
<td>95</td>
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*Reaction conditions: the molar ratio of cyclohexanone to glycol was 1:1.5; water-carrying agent: 16 mL of cyclohexane; reaction temperature: 100–105 °C; the molar ratio of the catalyst was 0.06 mmol.
Conclusions

In summary, a new 3D coordination polymer built by triangle trinuclear copper units and triazole with [Mo$_5$O$_{18}$]$^{6-}$ anions as templates were successfully synthesized. Its catalytic activities in the synthesis of cyclohexanone ethylene ketal and visible light photodegradation of RhB were explored. It exhibits good Lewis acid catalytic activity for the synthesis of cyclohexanone ethylene ketal with 95% conversion after 3 h. Its forbidden bandwidth belongs to the semiconductor category, and it can catalyze the photodecomposition of RhB under visible-light with above 98% degradation rate, demonstrating it is an effective catalyst for the removal of organic pollutant under visible light. This work demonstrates that it is a feasible way to achieve POMs-based coordination polymer with excellent catalytic performance by combining catalytic active building blocks. To some degree, this may offer an access to obtain a series of POMs-based catalysts bearing attractive catalytic activities. This continuous research is currently going on in our group.

Conflicts of interest

There are no conflicts to declare.

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Notes and references


