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Enhancement of CO₂ capture performance of aqueous MEA by mixing with [NH₂e-mim][BF₄]

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Alcohol amine solutions have a high absorption capacity and rate for CO₂ capture, however, there are some shortcomings such as high energy-consumption and low stability. To enhance CO₂ capture performance of aqueous MEA, a functional ionic liquid ([NH₂e-mim][BF₄]) was introduced based on the advantages for CO₂ capture. Absorbents were prepared with the molar concentration ratio of [NH₂e-mim][BF₄] to the 30 vol% aqueous MEA of 0 : 10, 1 : 9, 2 : 8, 3 : 7, 4 : 6 and 6 : 4. The density and the viscosity of the investigated absorbents were measured and the effects of the molar fraction of [NH₂e-mim][BF₄] (n_1) and temperature on CO₂ absorption performance were investigated. CO₂ desorption performance of the solvent at different temperatures was discussed. The stability performance of the absorbent with n_1 of 2 : 8 (I/M_{2:8}) was examined by five consecutive cyclic tests. The results showed that for pure CO₂, the I/M_{2:8} displayed the highest absorption performance at 303 K under 1 bar: a comparable CO₂ absorption capacity of the 30 vol% aqueous MEA and a higher CO₂ absorption rate at the later absorption stage. Moreover, with the increase of temperature, CO₂ absorption capacity and rate decreased, while CO₂ desorption efficiency and rate increased. 393 K was chosen as the optimum desorption temperature with the desorption efficiency of 99.31%. The introducing of IL contributed to CO₂ desorption performance of the absorbents significantly. The properties (CO₂ absorption capacity, mass loss, density and viscosity) of the I/M_{2:8} during the cycles suggested that the IL-MEA mixture had an excellent stability performance.

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1. Introduction

Global warming, caused by excessive emission of carbon dioxide (CO₂), has become one of the world's major environmental issues.¹⁻⁴ The reduction of CO₂ emissions by the capture of CO₂ from flue gases is considered as an effective method to mitigate the greenhouse effect.⁵⁻⁷ Currently, the leading technology involves chemical absorption with aqueous amine solutions (typically 30 vol% amine by volume).^{8,9} However, the commercially available aqueous amine solutions, represented by monoethanolamine (MEA), present many disadvantages including high regenerative energy and degradation in the presence of oxygen.¹⁰⁻¹² Moreover, the volatilization of amines causes environmental pollution and corrosion, as well as raises the cost of operation and amortized installation.^{13,14}

In recent years, considerable research efforts have been made to study the capture performance of the solvents that could overcome the aforementioned disadvantages. Ionic liquids (ILs), which are salts with a melting point below 100 °C and very low volatility, have great promise in the near future

considering their high CO₂ capture performance and reutilization.¹⁵⁻¹⁹ Among these, functionalized ILs, which could simultaneously improve absorption rate and selectivity of CO₂ capture through the reversible reactions between reactive group of the ILs and CO₂, have been intensively investigated in the past several decades.^{1,20-22} Bates *et al.* synthesized the IL (1-(1-aminopropyl)-3-butylimidazole fluoroborate, [NH₂p-bim][BF₄]) with amine moieties as the functional groups.²³ The IL shows a high adsorption capacity of 0.5 mol CO₂ per mol IL. In addition, some other studies also found that the amine-functionalized ILs have high CO₂ capture performance.²⁴⁻²⁷ However, the high viscosity of this amine-functionalized ILs influenced the mass transfer between the liquids and the gas seriously, resulting in a low CO₂ absorption capacity.²⁸⁻³²

In this paper, a new kind of solvent was developed by mixing an amine-functionalized IL (1-(1-aminoethyl)-3-methylimidazole fluoroborate [NH₂e-mim][BF₄]), with 30 vol% aqueous MEA solutions to investigate whether there was a synergetic effect on CO₂ capture performance. The effects of the molar concentration ratio of [NH₂e-mim][BF₄] to MEA in the mixture and the temperature on CO₂ absorption performance of the solvents were explored. Further, the desorption performance of the solvents at different temperature were discussed. Moreover, the cyclic stability of the absorbents were evaluated by five consecutive CO₂ absorption-desorption tests.

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2. Experimental

Materials

CO₂ (purity ≥ 99.99%) was purchased from Wuhan Minghui gas Co., Ltd., China. Ethanolamine (MEA, AR) and sodium borate (NaBF₄, CP) were purchased from Sinopharm Chemical Reagent Co., Ltd., China. 2-Bromine ethylamine hydrobromide (C₂H₇Br₂N, 98%) and 1-methylimidazole (C₄H₆N₂, 99%) were provided by Aladdin Reagent Co., Ltd., China.

Characterization and measurement

Preparation of the investigated solvents. [NH₂e-mim][BF₄] was synthesized and the purity was measured according to our previous work.^{22,33,34}

An appropriate amount of material was taken in a closed flask according to the molar concentration ratio of the [NH₂e-mim][BF₄] to the 30 vol% aqueous MEA of 0 : 10, 1 : 9, 2 : 8, 3 : 7, 4 : 6 and 6 : 4. The corresponding molar fraction of [NH₂e-mim][BF₄] (*n*₁) was 0, 0.1, 0.2, 0.3, 0.4, 0.6. It was mixed evenly by ultrasonic vibration.

Measurement of physical properties. The density was measured using a density meter (DM45 Delta Range, Mettler Toledo of Switzerland) that operated *via* electromagnetically induced oscillation of a glass U-form tube, with automatic compensation for variations in atmospheric pressure. The accuracy of the density meter measurements was ±0.00005 g cm⁻³ for all operating conditions.

The viscosity was measured using a viscometer (DV-II+ Pro, Brookfield of USA). The “ULA” spindle and jacketed sample cell were used for these relatively low viscosity absorbents. The accuracy of viscometer was ±1% of the reading for torque measurement with a repeatability of ±0.2% of the reading. The temperature of the jacketed sample chamber was controlled *via* a circulating bath (TC-602P, Brookfield of USA) with a temperature stability of ±0.01 K.

The mass of the solution was measured by a precision electronic balance (AR2140, Mettler Toledo of Switzerland). The accuracy of the mass measurements was ±0.00001 g.

Determination of CO₂ absorption/desorption performance. CO₂ absorption and desorption performance of the solvents were measured by a homemade apparatus as shown as Fig. 1.

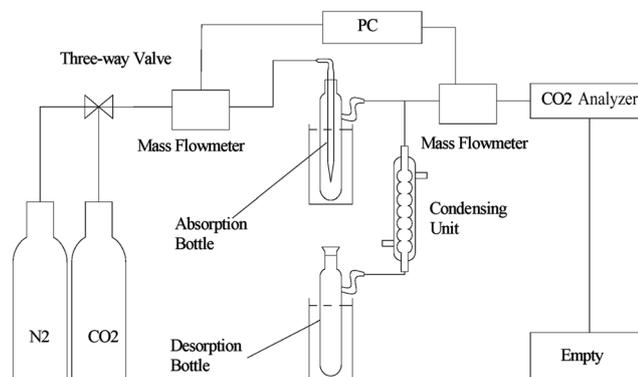


Fig. 1 Experimental apparatus for CO₂ absorption/desorption.

The error between the measured CO₂ solubility and the theoretical value is less than 5%.

For CO₂ absorption experiment, CO₂ intake speed was controlled at approximately 60 mL min⁻¹ under 1 bar. CO₂ intake and outlet flow was determined by mass flow meters (50L type, SIERRA Flow Measurement and Control Technology Company of USA) with measurement error range of ±1%, which was recorded by a computer every 30 s. CO₂ intake flow was recorded as *V*_{*i*} and *V*'_{*i*} for CO₂ outlet flow at time “*i*”. The volume of CO₂ solubility in the investigated solvents (Δ*V*_{*i*}, mL) could be obtained by eqn (1). CO₂ absorption volume at time “*t*” (*Q*_{*a*}, mL) and the molar fraction of CO₂ (*X*_{CO₂}, unit: mol CO₂ per mol mixture, expressed as mol per mol) in the solvents was calculated from eqn (2) and (3).

$$\Delta V_i = V_i - V'_i \quad (1)$$

$$Q = \sum_{i=0}^t \Delta V_i \quad (2)$$

$$X_{\text{CO}_2} = \frac{Q_a}{n_M + n_1} \quad (3)$$

where *n*_M and *n*₁ was the molar amount of MEA and [NH₂e-mim][BF₄] in the investigated absorbents.

During CO₂ desorption, temperature was regulated to range from 383 K to 398 K. CO₂ liberation volume was measured by a mass flowmeter and recorded in a computer. CO₂ desorption capacity (*Q*_{*d*}, mL) was also calculated by eqn (2). And CO₂ desorption efficiency (*η*, %) was defined as the percentage of CO₂ desorption capacity to corresponding CO₂ absorption capacity, which could be calculated by eqn (4).

$$\eta = \frac{Q_d}{Q_a} \times 100\% \quad (4)$$

The investigated solvents were subjected to the steps previously mentioned to carry out CO₂ absorption–desorption cycle experiments.

3. Results and discussion

Physical properties of the investigated solvents

Density and viscosity of the investigated solvents at 303 K under 1 bar were shown in Fig. 2.

The density value of the mixed absorbents was between the density value of 30 vol% solution (1007.2 kg m⁻³) and the [NH₂e-mim][BF₄] (1472.2 kg m⁻³) which were measured at the same condition. And the viscosity of the mixed absorbents also showed the same law. The viscosity of [NH₂e-mim][BF₄] was 3589.2 mPa s determined by our group,^{22,34} while the viscosity of the absorbents was below 10 mPa s with *n*₁ of 0–0.6, which showed that the addition of 30 vol% solution reduced the viscosity of absorbents significantly. According to Fig. 2, both the density and the viscosity of the investigated absorbents increased with the increase of *n*₁. The density was basically linear related to *n*₁ and the slope of viscosity kept increasing with the increase of *n*₁ when *n*₁ exceed 0.2.



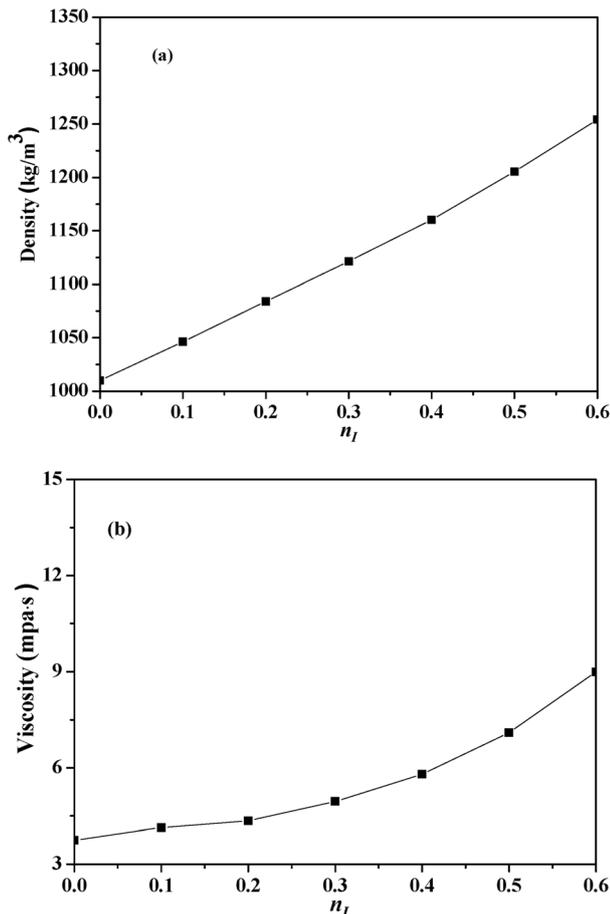


Fig. 2 Physical properties of the absorbents with different n_1 at 303 K: (a) density; (b) viscosity.

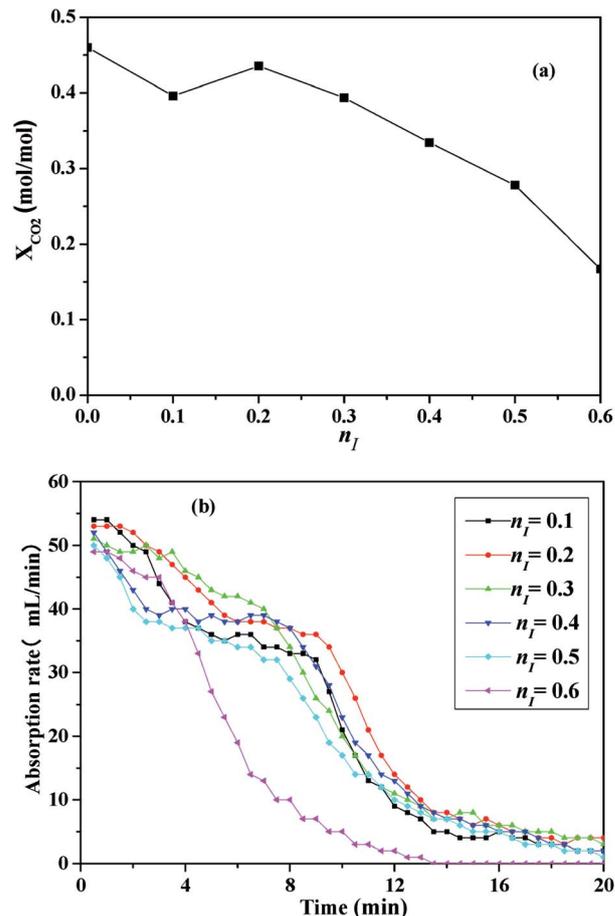


Fig. 3 CO_2 absorption performance of the absorbents with different n_1 at 303 K: (a) the absorption capacity; (b) the absorption rate.

Effects of the n_1 on the absorption performance

Fig. 3 illustrated the effects of n_1 on CO_2 absorption performance at 303 K under 1 bar.

As shown in Fig. 3(a), with the increase of n_1 , CO_2 absorption capacity increased initially, and reached a maximum of 0.4554 mol per mol (contrasting with a value of 0.4601 mol CO_2 per mol MEA measured at the same condition) when n_1 was 0.2, then decreased, which suggested there was an optimum n_1 in the absorbent for CO_2 absorption. In Fig. 3(b), as time went on, the absorption rate of all absorbents decreased rapidly within 12 minutes, and then decreased gently. With the increasing of n_1 , CO_2 absorption rate increased firstly and then decreased, which suggested that an optimal n_1 also existed.

CO_2 absorption performance was influenced by the viscosity of the absorbents, the imidazole ring content and the amine group content in the absorbents.^{35,36} CO_2 absorption capacity was defined as the amount of CO_2 absorption per mol amine in this work, thus the variation of CO_2 absorption performance was mainly caused by the viscosity and the imidazole ring content of the absorbents. On the one hand, $[\text{NH}_2\text{e-mim}][\text{BF}_4]$ concentration increased as n_1 increased and consequently increased the viscosity of the absorbents just as Fig. 2(b) shown, which militated against the contact of CO_2 and the absorbents.

On the other hand, the increase of n_1 should cause the increase of the imidazole ring content in the absorbent, which benefits CO_2 absorption.³⁶ The comprehensive effects of the factors caused that CO_2 absorption performance increased and then decreased with the increase of n_1 .

As illustrated in Fig. 3, the absorbent prepared with n_1 of 0.2 (I/M_{2:8}) showed the best absorption performance. Thus, a comparison between the absorption rate of the I/M_{2:8} and the MEA solution at 303 K under 1 bar had been made in Fig. 4. At the early stage (0–10 minutes), the aqueous MEA solution showed higher absorption rate than the mixture of $[\text{NH}_2\text{e-mim}][\text{BF}_4]$ and MEA. After ten minutes, the mixture absorbent showed a higher absorption rate, which suggested that an moderate amount of $[\text{NH}_2\text{e-mim}][\text{BF}_4]$ would contribute to increasing CO_2 absorption rate at the later stage mainly.

Considering that the mixed absorbent with the I/M_{2:8} showed the best absorption performance, the mixed absorbent mentioned below was prepared with this ratio.

Effects of the temperature on the absorption performance

Fig. 5 showed the effects of the temperature on CO_2 absorption capacity and rate of the 30 vol% MEA solution and the I/M_{2:8}. According to Fig. 5(a), the CO_2 absorption capacity decreased as



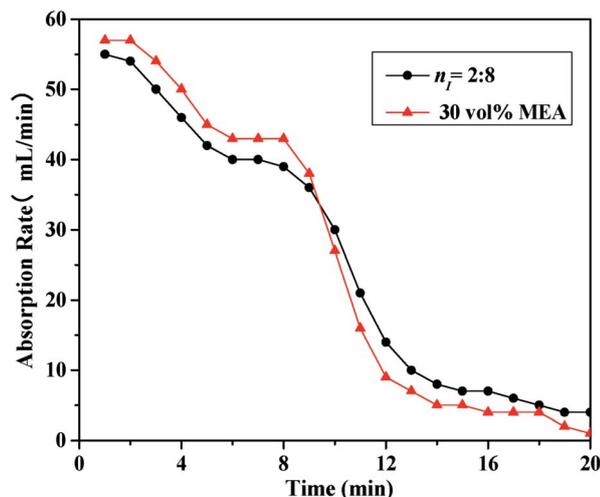


Fig. 4 Comparison of CO₂ absorption rate between the absorbent with n_1 of 0.2 and the 30 vol% MEA solution at 303 K.

the temperature increased. The comparison between the CO₂ absorption capacity of the mixed absorbent and the MEA solution illustrated that CO₂ absorption capacity of the latter was comparable to the former. From Fig. 5(b), the CO₂ absorption

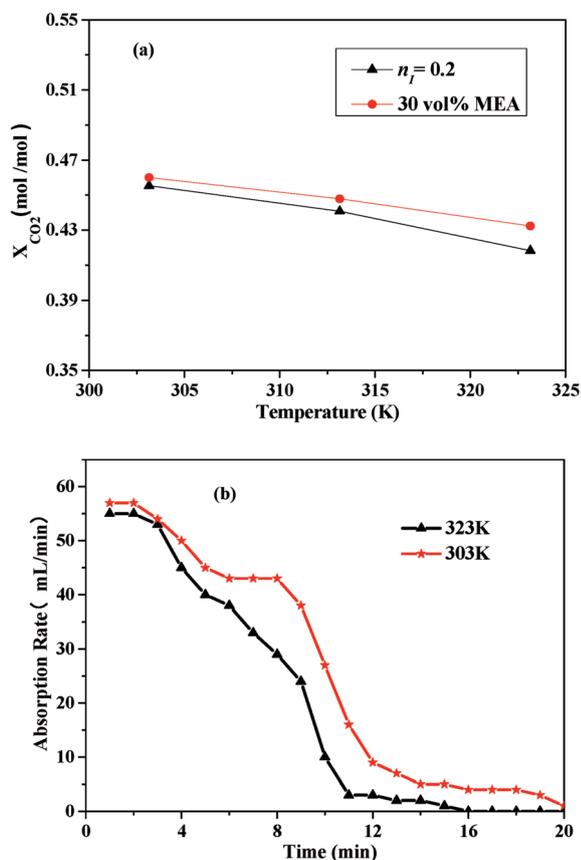


Fig. 5 CO₂ absorption performance of the I/M_{2.8} absorbent: (a) CO₂ absorption capacity vs. temperature; (b) CO₂ absorption rate vs. time at 303 K and 323 K.

rate decreased with the temperature ranged from 303 K to 323 K. Absorption equilibrium could be reached a little bit earlier at a higher temperature of 323 K, which might due to a lower CO₂ absorption capacity.

Therefore, both CO₂ absorption capacity and CO₂ absorption rate of the absorbents decreased with increasing of temperature. Similar result could be found in the previous studies.^{37–39}

Effects of the temperature on the desorption performance

Fig. 6 presented the desorption performance of the I/M_{2.8} at different temperature and the desorption performance of 30 vol% MEA at 383 K under 1 bar. As shown in Fig. 6(a), all the desorption efficiency increased significantly at the beginning (0–40 minutes) and then increased gently until desorption equilibrated. With the increasing of temperature, the desorption efficiency increased. The desorption efficiency of the I/M_{2.8} was much higher than that of the aqueous MEA at the same temperature and it exceeded 50% in 30 minutes. By contrast, the MEA solution needed an hour to reach the desorption efficiency of 50%. The desorption rate decreased intensively first and then decreased slightly until desorption balanced, as shown in Fig. 6(b). With increasing temperature, the desorption rate increased, which followed the same trend as the desorption efficiency. As it could be seen in Fig. 6, the desorption efficiency and the desorption rate at 393 K were close to the values at 398

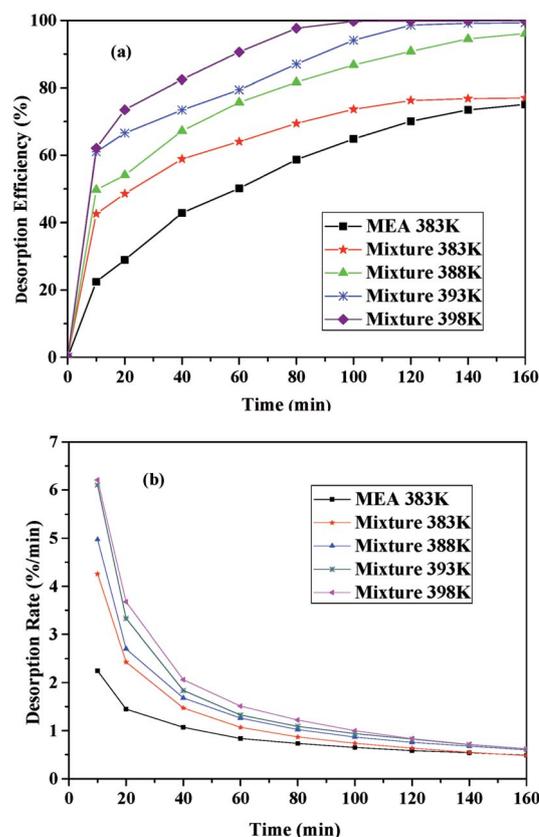


Fig. 6 CO₂ desorption performance the I/M_{2.8} absorbent at different temperature: (a) the desorption efficiency vs. time; (b) the desorption rate vs. time.



K. Considering the energy-consumption, 393 K was chosen as the optimum desorption temperature of the mixture with CO₂ desorption efficiency of 99.31%, which was lower than that of the 30 vol% MEA solution (398 K).

According to Fig. 6, CO₂ desorption efficiency and rate of the I/M_{2:8} were higher than that of the aqueous MEA at the same conditions, which suggested that the mixture absorbents had a lower energy consumption for regenerations compared with the aqueous alcohol. The energy-consumption of CO₂ desorption involved three parts: the energy for the increase of the mixture temperature, which was the product of the mole quantity, molar heat capacity and the differential of the temperature; the energy for the escape of CO₂, which was equal to the absorption heat of CO₂; and the energy for the vaporization of the absorbent.¹ The regeneration temperature of the functional ILs was from 263 K to 343 K,⁴⁰ which was lower than that of the alcohol amine. Thus, the ILs needed less energy for the increase of the temperature. The absorption heat depended on the heat of the reaction between the absorbents and CO₂, which were nearly the same for these two absorbents. Therefore, the energy for the escape of CO₂ from the two absorbents showed little change. The regeneration temperature of the MEA solution was higher than ILs, accordingly, the vaporization of water and MEA needed more energy. Consequently, CO₂ desorption of ILs system required less energy than the MEA system, which suggested that the ILs was an economical solution for CO₂ capture.

Stability performance of the absorbent

To examine the stability performance of the absorbents, properties including CO₂ adsorption capacity, mass loss, density and viscosity of the MEA solution and the I/M_{2:8} were circularly tested for 5 times. The absorption temperature was 303 K, and the time of duration was 30 min. The desorption temperature of MEA solution was 398 K, while the desorption temperature of I/M_{2:8} was 393 K, and the time of duration was 120 min. Fig. 7 showed the properties of the absorbents in each cycle.

As shown in Fig. 7(a), the values of CO₂ adsorption capacity of the MEA solution and the I/M_{2:8} decreased gradually in five absorption-desorption cycles. And the decrease percentage of CO₂ adsorption capacity of the I/M_{2:8} were slightly less than those of MEA. However, the mass weight loss of the two absorbents in the 5 cycles showed different trends. The weight loss percentage of the I/M_{2:8} was 3% relative to the first circle after one cycle, and then the weight was a small loss during the next four cycles. By contrast, the mass of MEA system decreased more obviously, and the percent of the decline values increased as the number of the circles went up. The difference of the two absorbents might be caused by the water and MEA which was more volatile at a higher temperature of 393 K. To verify the reason of the phenomenon, the density and the viscosity of the absorbents were measured respectively in each circle. The results were shown in Fig. 7(b).

As illustrated in Fig. 7(b), as the number of absorption-desorption cycles went up, the density of the absorbents increased gently and then became invariant after the third cycle. In the second cycle, the density of the I/M_{2:8} increased about

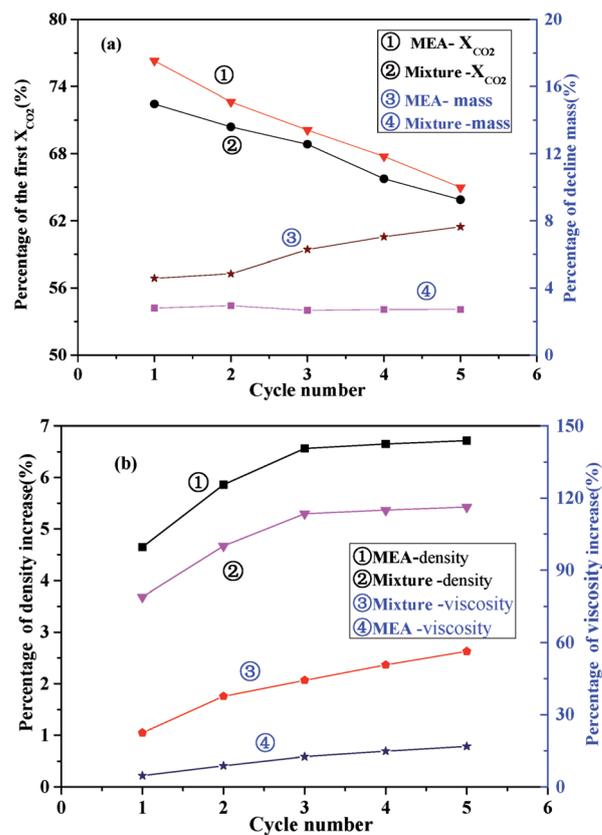


Fig. 7 Properties of the I/M_{2:8} absorbent during 5 times cycle: (a) CO₂ adsorption capacity as a percentage of the first time and percentage of quality decline compared with the first time; (b) percentage of density and viscosity increase compared with the first time.

3.5% relative to the first circle, while the increase amount of the MEA solution was about 4.5%. The viscosity of the absorbents also increased as the number of absorption-desorption cycles went up. The viscosity of the MEA solution increased in a range of 2–5% with the increase of the cyclic times. And for the I/M_{2:8}, it increased about 12% relative to the first circle, and the growth rate slowed down after the second circle with the increase amount of about 5%. The viscosity value of the mixture was in a range of 4–10 mPa s at 303 K, which was still a moderate viscosity for CO₂ adsorption. From the above results, it suggested that the mixture of ILs and MEA showed a better stability of density and a larger change of viscosity, which mitigated the loss of volatile MEA and water in the absorbents.

According to Fig. 7, it was concluded that the mixture of IL and MEA had a better stability in CO₂ adsorption capacity, mass loss and density and had a better thermodynamic stability than the MEA solution.

4. Conclusions

In this work, absorbents with excellent CO₂ capture performance were developed by mixing 30 vol% MEA solution with an ionic liquid [NH₂e-mim][BF₄]. The density and the viscosity of the absorbents increased with the increase of *n*₁. The addition of



MEA solution reduced the viscosity of absorbents to no more than 10 mPa s.

The absorbent exhibited a rather excellent CO₂ capture performance when n_1 was 0.2: (1) a moderate [NH₂e-mim][BF₄] could contribute to increasing CO₂ absorption rate at the later stage; (2) CO₂ absorption capacity of the I/M_{2.8} could comparable with the 30 vol% MEA solution. It contributed to advance the CO₂ absorption process to increase temperature, however, CO₂ absorption capacity and rate decreased.

Both CO₂ desorption efficiency and CO₂ desorption rate increased with increasing temperature. And CO₂ desorption efficiency and rate was higher than the 30 vol% solution at the same conditions. Considering the energy-consumption, 393 K was chosen as the optimum desorption temperature, which was 5 K lower than the MEA solution.

Compared with MEA solution, the absorbent of I/M_{2.8} showed a better cyclic stability. The reason was that the addition of [NH₂e-mim][BF₄] into the MEA could improve the thermodynamic stability of the absorbent.

Conflicts of interest

There are no conflicts to declare.

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