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Electrooxidation of Pd–Cu NP loaded porous carbon derived from a Cu-MOF†

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The development of new non-platinum catalysts for alcohol electrooxidation is of utmost importance. In this work, a bimetallic Pd–Cu loaded porous carbon material was first synthesized from a Cu-based metal–organic framework (MOF). The Cu loaded porous carbon was pre-synthesized through calcinating the Cu-based MOF under a N₂ atmosphere. After loading Pd onto the precursor and heating, Pd–Cu loaded porous carbon (Pd–Cu/C) was obtained for alcohol electrooxidation. Electrooxidation experiments revealed that this Pd–Cu bimetal loaded porous carbon assisted steady state electrolysis for alcohol oxidation in alkaline media. Moreover, different alcohols were electrooxidated using the present electrocatalyst for the purposes of discussing the oxidation mechanism. This electrooxidation study of Pd–Cu/C derived from a MOF demonstrates a good understanding of the electrooxidation of different alcohols, and provides useful guidance for developing new electrocatalyst materials for energy conversion and electronic devices.

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There is an immediate need to develop direct alcohol fuel cells (DAFCs), which have been proven to be a fine source of energy, which could probably replace fossil fuels use to fulfil global energy demand. As one of the most significant electrocatalytic procedures, the electrooxidation of alcohols is an important process in DAFCs, and has gathered much attention and is attractive, due to high power density output and low pollutant emissions. Generally, Pt based materials are the most common electrocatalysts for alcohol electrooxidation reactions. However, the high cost and limited supply of Pt severely restricts its commercial application. Therefore, the development of new efficient and inexpensive non-platinum alternative materials to Pt-based catalysts is of utmost importance.

Metal–organic frameworks (MOF) are assembled from metal ions linked by organic ligands, and are used in catalysis, guest molecule storage/separation, fluorescence, sensors and other devices.^{1–3} Due to their highly ordered porous structures and large surface areas, MOFs can also be used as templates/precursors for preparing porous carbon materials through thermal treatment.^{4–8} Several MOF derived carbon materials with good electrical conductivity are reported to show effective electrocatalytic performance,⁹ such as in the oxygen evolution reaction (OER),¹⁰ hydrogen evolution reaction (HER),¹¹ and oxygen reduction reaction (ORR).¹² Recently, a zeolitic

imidazolate framework ZIF-8 was calcinated in order to prepare porous carbon with both micro- and meso-pores to support Pd electrocatalysts for methanol electrooxidation.¹³ Unfortunately, this could not efficiently limit the use of the noble metal, which gives challenges to scientists for further exploration.

It has been demonstrated that the alloying of noble metals with transition metals has been used for the enhancement of catalytic activity and reduction of cost,¹⁴ because of the low cost and relatively high abundance of transition metals. The use of alloyed metal, Pd–M (where M is Cu, Co or Ni), binary electrocatalysts has been reported for effectively improving the catalytic properties.^{15–17} For this purpose, introducing a non-noble metal into a noble metal bimetallic system will fulfil the demand for a new catalyst and become an area of interest nowadays. The second metal (such as Cu) will behave as a donor, while Pd has an empty d orbital to accept electrons, and it is assumed that its electronic properties will be more similar to Pt.¹⁸ Therefore, Pd–Cu bimetal loaded porous carbon derived from a Cu-MOF can provide a good electrocatalyst for alcohol oxidation.

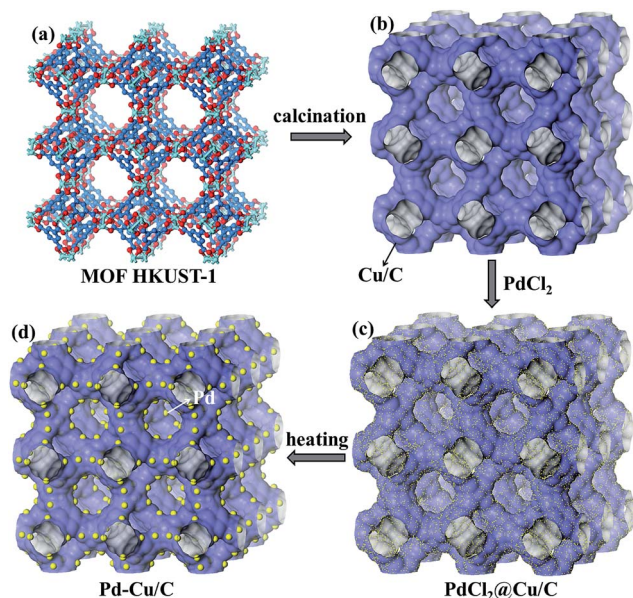
In this work, as shown in Scheme 1, we have synthesized Cu loaded porous carbon through the direct carbonization of a porous Cu based MOF. After loading Pd onto the precursor and heating, Pd–Cu bimetal loaded porous carbon (Pd–Cu/C) was obtained. Chronoamperometric studies revealed that this Pd–Cu bimetal loaded porous carbon assisted steady state electrolysis for alcohol oxidation in alkaline media. In addition, alcohols with different numbers of carbon atoms (such as ethanol, 1-propanol and 2-propanol) were also investigated for electrooxidation. It is important to note that the effectiveness of this bimetallic NP loaded carbon means that it can serve as

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Scheme 1 The preparation procedure for Pd-Cu/C derived from HKUST-1 and PdCl₂: (a) MOF HKUST-1; (b) Cu/C calcinated from HKUST-1; (c) PdCl₂ loaded on Cu/C; and (d) the Pd-Cu/C material.

a catalyst for the electrooxidation of low-molecular weight alcohols, which probably can be used as energy sources in portable electronic devices.

Here, a 3-D MOF, HKUST-1 (also called Cu₃(BTC)₂, BTC = 1,3,5-benzenetricarboxylate), was chosen as the precursor for preparing the Cu/C material, due to the structure having high porosity and it being a rich Cu source. The as-synthesized HKUST-1 was calcinated at 700 °C for 5 h under a N₂ atmosphere, and the Cu/C material was obtained. In addition, the guest species PdCl₂ was loaded onto the calcinated HKUST-1 through immersing the pre-calcinated HKUST-1 into a PdCl₂ ethanolic solution (1 mM) for 2 h (Scheme 1). The PdCl₂ loaded Cu/C (PdCl₂@Cu/C) was heated at 300 °C for 1 h under a N₂ atmosphere. Finally, an alloy of Pd and Cu loaded porous carbon material (Pd-Cu/C) was obtained and characterized through powder XRD, BET and XPS analyses.

The PXRD data (Fig. 1a) from as-synthesized HKUST-1 powder and the bimetallic Pd-Cu NP loaded carbon porous material derived from HKUST-1 show that the samples contain bimetallic palladium and copper mostly. The XRD peak appearing at 43.3° corresponds to the (fcc) (111) facet plane of Cu. Due to Pd being dispersed homogeneously at a low concentration through the sample, the XRD pattern could not display the obvious peak from Pd. However, inductively coupled plasma emission spectroscopy (ICP) data (Table S1†) from the sample showed 0.76% Pd and 36.68% Cu, indicating the existence of Pd and Cu. The porosity of Pd-Cu/C was demonstrated through BET data, which shows N₂ adsorption of ~150 cm³ g⁻¹. The Pd XPS spectrum showed two definite peaks at 335.5 and 341 eV, respectively assigned to 3d_{5/2} and 3d_{3/2} and matching well with Pd⁰. XPS peaks at 932.4 and 952.1 eV indicate the valence states of Cu ions in the Cu 2p_{3/2} and Cu 2p_{1/2} orbitals in the Pd-Cu/C

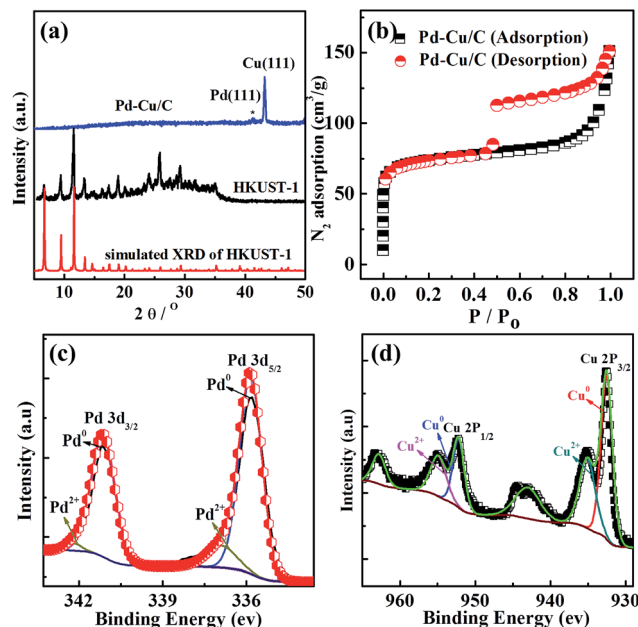


Fig. 1 (a) XRD data from HKUST-1 and Pd-Cu/C; (b) N₂ sorption isotherms for Pd-Cu/C; and XPS data from (c) Pd and (d) Cu in a sample of Pd-Cu/C.

material. Cu²⁺ is present in the porous carbon material, with respective peaks at 933.7 eV and 934.4 eV from CuO and Cu(OH)₂, with a prominent satellite observed in the 938–946 eV range. A few Pd²⁺ ions also exist in the sample due the easy oxidation of the surface. The Raman spectrum of Pd-Cu/C (Fig. S4†) shows typical graphitic carbon. The results of the characterization studies clearly reveal that the nanoparticles have a Pd and Cu bimetallic nature.

SEM images with EDS (Fig. 2a and b) results show that the sample contained much more copper than palladium, which clearly suggests that the presence of copper in the sample would probably be the reason for the expected electrooxidation of alcohols. It could be possible to replace the use of high-cost Pd or Pt based catalysts. The morphology of the Pd-Cu NPs was further characterized *via* TEM imaging and TEM element mapping (Fig. 2c, d and S5†), demonstrating that the nano-sized NPs were dispersed homogeneously. The HR-TEM image in Fig. 2c gives insight into the bimetallic nature of the synthesized nanoparticles, with two noticeable lattice fringes (0.225 nm for Pd(111) and 0.202 nm for Cu(111)). The mean size of the Pd-Cu NPs was 7.38 nm, as shown in Fig. 2d. The homogenous distribution, with well-defined bimetallic Pd-Cu based carbon material, was good for the electrooxidation of alcohols. The electrochemical active surface area (ECSA) for Pd-Cu/C was high compared with commercial Pd/C, which suggested that the synthesized Pd-Cu/C has ample available surface area, mainly because of synergistic effects from the Cu-MOF based carbon material and the morphology of the electrocatalyst.

In Fig. 3, CV profiles for commercial Pd/C and the presented Pd-Cu/C show two distinct peaks (forward (iF) and backward



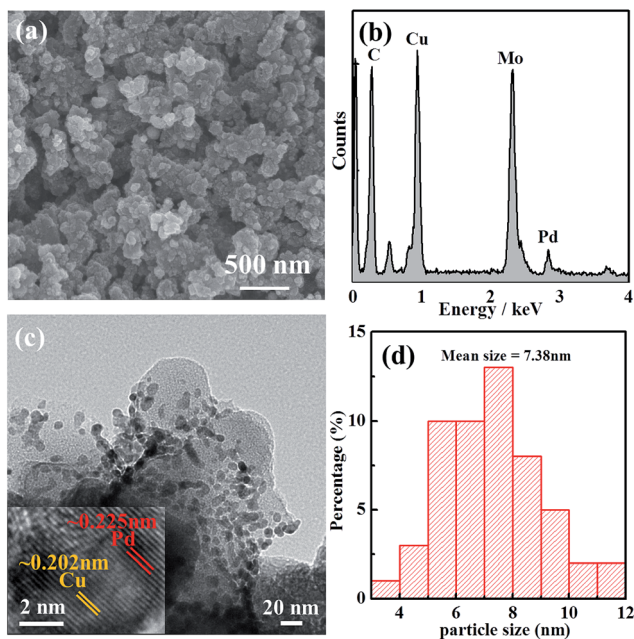


Fig. 2 (a) SEM image of and (b) EDS data from Pd-Cu/C; (c) a TEM image of Pd-Cu NPs in the hybrid carbon material; and (d) the size distribution of the Pd-Cu NPs.

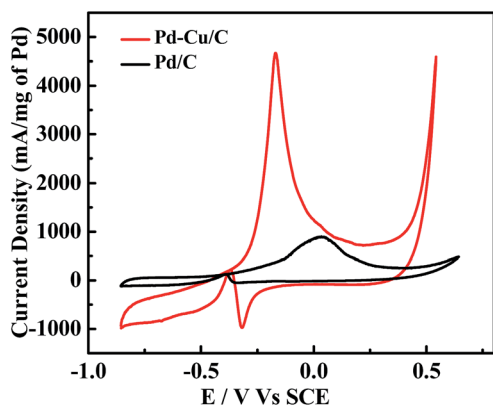


Fig. 3 CV curves from Pd/C and Pd-Cu/C electrocatalysts during CH₃OH (1 M) oxidation in 1 M KOH solution, at a scan rate of 50 mV s⁻¹, at room temperature.

(iB) peaks) during the oxidation of methanol-containing 1 M KOH solution. The peak at -0.37 V indicates the oxidation of aforementioned carbonaceous species, such as Pd-CO_{ads}, along with newly formed alcohol adsorbates, following the removal of surface intermediates at lower potentials.¹⁹ For the forward peak potential, a shift in the *i*F value is observed, mainly because of Cu existing with Pd in the material. This results in the oxidation of poisonous species, such as Pd adsorbed CO, at higher potentials,²⁰ leading to such high activity. ATR-IR (Fig. S7†) and GC analyses (Fig. S8†) show the methanol oxidation reaction (MOR) pathway during the formation of the final CO₂ product. The catalytic activity of Pd-Cu/C is found to be ~ 13 times higher than commercial Pd/C for methanol oxidation, demonstrating that the presence of Cu with Pd in Pd-

Cu based catalysts increases CO oxidation because of a strong binding ability. Cu binds to CO more strongly than Pd, as a result of electronic structure differences,^{21,22} thus preventing the electrode from undergoing CO poisoning, a major issue for Pd-based catalysts during the methanol oxidation reaction (MOR). The mechanism of methanol oxidation is shown in the ESI (eqn (8)–(10)).† The high *i*F value for Pd-Cu/C can be ascribed to the fast formation of reactive intermediates, such as Pd-CH₂OH, Pd-COOH, Pd-H, Pd-(CHO)_{ads}, and Pd-(COOH)_{ads}.^{23–25} The removal of these intermediates is necessary for a high current density. Furthermore, formaldehyde (HCHO), formic acid (HCOOH) and CO₂ would be the final products in the MOR.^{26,27} Pd-Cu/C has good catalytic activity for the MOR, leading to further investigation into the electrooxidation of different alcohols, such as ethanol, 1-propanol, and 2-propanol (Fig. 4 and Table S2†).

The current densities for different alcohol oxidation processes are summarized in Fig. 4a and b. The normalized *i*F (calculated using Pd mass) for the MOR (~ 4643 mA mg⁻¹) was higher than for three other alcohols, *i.e.*, it was ~ 139 , ~ 94 and ~ 26.5 mA mg⁻¹ for ethanol, 1-propanol and 2-propanol, respectively. The *i*F/*i*B ratio for methanol is ~ 12 times higher than that for ethanol, ~ 13 times that for 1-propanol and ~ 4 times that for 2-propanol. The reactivity order for Pd-Cu/C is methanol > ethanol > 1-propanol > 2-propanol. 2-Propanol electrooxidation showed a lower current density on a Pd-Cu/C electrode in alkaline medium, although *i*F/*i*B is ~ 5.4 . The negative shift in the onset of the ethanol oxidation reaction (EOR) suggested that a high copper content with very low amount of Pd was suitable for EOR kinetics using a Pd-Cu/C catalyst. The ethoxy (CH₃CO)_{ads} was strongly adsorbed, and blocked hydrogen absorption/adsorption. The current intensity of *i*F increased due to the formation of fresh Pd-OH, through stripping carbonaceous residue from the electrode (eqn (11)–(14)†). In addition, the increased current at high potentials sharply reached the largest value then started to decline, because a PdO layer formed on the electrode, blocking the further adsorption of reactive species.²⁸ ATR-IR spectra (Fig. S9†) show the presence of CO₂ and CO_{ads}, whereas bands appear at 1670 cm⁻¹ and 1390 cm⁻¹ because of the formation of

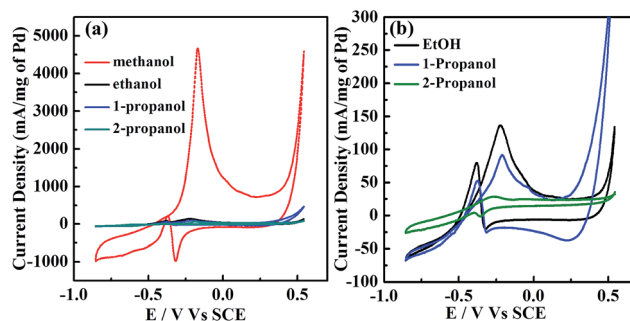


Fig. 4 (a) CV curves from: the Pd-Cu/C electrocatalyst for C₁–C₃ aliphatic alcohol (1 M) oxidation in KOH (1 M) solution; and (b) Pd-Cu/C in 1 M EtOH, 1-propanol and 2-propanol at a scan rate of 50 mV s⁻¹ at room temperature.



acetic acid.²⁹ The Pd–Cu/C electrocatalyst has the potential to oxidize the intermediate to the final product, CO₂, during EOR to some extent; Cu promotes oxidation through increasing the production of OH_{ads}[−]/H₂O to eliminate the intermediate CH₃–CO_{ads} simultaneously on Pd.³⁰ The stability of Pd–Cu/C in all four alcohols (methanol, ethanol, 1-propanol and 2-propanol) was studied using chronoamperometry at a potential of −0.25 V, as shown in Fig. S2.† The slow current decay showed that the stability of the Pd–Cu/C electrocatalyst in methanol is best. Comparing the results, the current for methanol oxidation was higher than that for the other three alcohols. However, the oxidation currents from ethanol and 1-propanol were larger than that from 2-propanol. This suggested that Pd–Cu/C is less stable and shows lower anti-poisoning ability during 2-propanol oxidation in an alkaline medium.

During this oxidation, 1-propanol oxidizes to propanal first, and its further oxidation results in the formation of a stable product, propanoic acid (Scheme S1†). 1-Propanol is converted, with its carboxylate as the major product, as verified using ATR-IR (Fig. S10†). 2-Propanol forms acetone as an intermediate product, leading to the poisoning of the electrode.³¹ ATR-IR spectra (Fig. S11†) of Pd–Cu/C also confirm that the electrocatalyst follows a dual pathway through acetone and propene intermediates to oxidize to CO₂ finally (Scheme S2†).^{32–34} However, acetone formation is kinetically favored.³⁵ The results show that the location of the –OH group in the alcohol influences the electrooxidation reaction kinetics. In contrast methanol oxidation using Pd–Cu/C has much higher catalytic activity than ethanol, 1-propanol and 2-propanol oxidation, which makes it a good candidate for direct methanol fuel cells.

In summary, we have first synthesized a bimetallic Pd–Cu NP loaded porous carbon material from a Cu-based MOF for alcohol electrooxidation. The Cu loaded porous carbon was pre-synthesized by calcinating the Cu-based MOF HKUST-1 under a N₂ atmosphere. Afterwards, Pd–Cu NP loaded porous carbon was obtained for alcohol electrooxidation. Electrooxidation experiments revealed that Pd–Cu/C was suitable for steady state electrolysis for alcohol oxidation in alkaline media. In addition, different alcohols were electrooxidated using the present electrocatalyst to discuss the oxidation mechanism. This electrooxidation study of Pd–Cu/C derived from a MOF offers good understanding into the electrooxidation of different alcohols and it could provide useful guidance for the development of new electrocatalyst materials.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 C.-J. Zhong, J. Luo, P. N. Njoki, D. Mott, B. Wanjala, R. Loukrakpam, S. Lim, L. Wang, B. Fang and Z. Xu, *Energy Environ. Sci.*, 2008, **1**, 454.
- 2 M. Jahan, Q. Bao and K. P. Loh, *J. Am. Chem. Soc.*, 2012, **134**, 6707.
- 3 (a) J. Lee, O. K. Farha, J. Roberts, K. A. Scheidt, S. T. Nguyen and J. T. Hupp, *Chem. Soc. Rev.*, 2009, **38**, 1450; (b) F. Luo, C. S. Yan, L. L. Dang, R. Krishna, W. Zhou, H. Wu, X. L. Dong, Y. Han, T. L. Hu, M. O'Keeffe, L. L. Wang, M. B. Luo, R. B. Lin and B. L. Chen, *J. Am. Chem. Soc.*, 2016, **138**, 5678.
- 4 (a) L. Shang, H. J. Yu, X. Huang, T. Bian, R. Shi, Y. F. Zhao, G. I. N. Waterhouse, L. Z. Wu, C. H. Tung and T. R. Zhang, *Adv. Mater.*, 2016, **28**, 1668; (b) A. Aijaz, J. Masa, C. Rosler, W. Xia, P. Weide, A. J. R. Botz, R. A. Fischer, W. Schuhmann and M. Muhler, *Angew. Chem., Int. Ed.*, 2016, **55**, 4087; (c) T. Wang, L. Shi, J. Tang, V. Malgras, S. Asahina, G. G. Liu, H. B. Zhang, X. G. Meng, K. Chang, J. P. He, O. Terasaki, Y. Yamauchi and J. H. Ye, *Nanoscale*, 2016, **8**, 6712; (d) Z. G. Gu, D. X. Zhang, W. Q. Fu, Z. H. Fu, M. I. Vohra, L. Zhang, C. Woll and J. Zhang, *Inorg. Chem.*, 2017, **56**, 3526.
- 5 B. Liu, H. Shioyama, T. Akita and Q. Xu, *J. Am. Chem. Soc.*, 2008, **130**, 5390.
- 6 J. Tang and Y. Yamauchi, *Nat. Chem.*, 2016, **8**, 638.
- 7 S. Dang, Q.-L. Zhu and Q. Xu, *Nat. Rev. Mater.*, 2017, **3**, 17075.
- 8 M. E. Davis, *Nature*, 2002, **417**, 813.
- 9 H. Wang, Q.-L. Zhu, R. Zou and Q. Xu, *Chem*, 2017, **2**, 52.
- 10 (a) Y. J. Han, J. F. Zhai, L. L. Zhang and S. J. Dong, *Nanoscale*, 2016, **8**, 1033; (b) M. H. Zhuang, X. W. Ou, Y. B. Dou, L. L. Zhang, Q. C. Zhang, R. Z. Wu, Y. Ding, M. H. Shao and Z. T. Luo, *Nano Lett.*, 2016, **16**, 4691.
- 11 (a) M. Xu, L. Han, Y. J. Han, Y. Yu, J. F. Zhai and S. J. Dong, *J. Mater. Chem. A*, 2015, **3**, 21471; (b) X. Y. Gan, R. J. Zheng, T. L. Liu, J. Meng, R. P. Chen, X. Sun and X. Sun, *Chem.–Eur. J.*, 2017, **23**, 7264.
- 12 W. Zhang, Z. Y. Wu, H. L. Jiang and S. H. Yu, *J. Am. Chem. Soc.*, 2014, **136**, 14385.
- 13 J. Li, Q. L. Zhu and Q. Xu, *Chem. Commun.*, 2015, **51**, 10827.
- 14 (a) S. T. Hunt, M. Milina, A. C. Alba-Rubio, C. H. Hendon, J. A. Dumesic and Y. Roman-Leshkov, *Science*, 2016, **352**, 974; (b) C. Y. Song, D. M. Zhang, B. Wang, Z. Cai, P. Yan, Y. Sun, K. Ye, D. X. Cao, K. Cheng and G. L. Wang, *Nano Res.*, 2016, **9**, 3322.
- 15 Z. Yan, M. Zhang, J. Xie, J. Zhu and P. K. Shen, *Appl. Catal., B*, 2015, **165**, 636.
- 16 B. Y. Xia, H. B. Wu, X. Wang and X. W. Lou, *J. Am. Chem. Soc.*, 2012, **134**, 13934.
- 17 J. Bagchi and S. K. Bhattacharya, *Transition Met. Chem.*, 2007, **32**, 47.
- 18 S. V. Myers, A. I. Frenkel and R. M. Crooks, *Chem. Mater.*, 2009, **21**, 4824.
- 19 P. S. Roy, J. Bagchi and S. K. Bhattacharya, *Catal. Sci. Technol.*, 2012, **2**, 2302.



- 20 P. S. Roy and S. K. Bhattacharya, *Catal. Sci. Technol.*, 2013, **3**, 1314.
- 21 Y. Debauge, M. Abon, J. Bertolini, J. Massardier and A. Rochefort, *Appl. Surf. Sci.*, 1995, **90**, 15.
- 22 F. Delbecq, *Surf. Sci.*, 1997, **389**, L1131.
- 23 X. Xia, T. Iwasita, F. Ge and W. Vielstich, *Electrochim. Acta*, 1996, **41**, 711.
- 24 J. S. Spendelow and A. Wieckowski, *Phys. Chem. Chem. Phys.*, 2007, **9**, 2654.
- 25 Y. X. Chen, A. Miki, S. Ye, H. Sakai and M. Osawa, *J. Am. Chem. Soc.*, 2003, **125**, 3680.
- 26 W. F. Lin, J. T. Wang and R. Savinell, *J. Am. Chem. Soc.*, 1997, **119**, 1917.
- 27 E. Batista, G. Malpass, A. Motheo and T. Iwasita, *Electrochem. Commun.*, 2003, **5**, 843.
- 28 A. Tripković, K. D. Popović and J. Lović, *Electrochim. Acta*, 2001, **46**, 3163.
- 29 T. Iwasita and E. Pastor, *Electrochim. Acta*, 1994, **39**, 531.
- 30 Q. Dong, Y. Zhao, X. Han, Y. Wang, M. Liu and Y. Li, *Int. J. Hydrogen Energy*, 2014, **39**, 14669.
- 31 S.-G. Sun and Y. Lin, *Electrochim. Acta*, 1998, **44**, 1153.
- 32 J. Otomo, X. Li, T. Kobayashi, C.-j. Wen, H. Nagamoto and H. Takahashi, *J. Electroanal. Chem.*, 2004, **573**, 99.
- 33 J. Ye, J. Liu, C. Xu, S. P. Jiang and Y. Tong, *Electrochem. Commun.*, 2007, **9**, 2760.
- 34 Y. Cheng, Y. Liu, D. Cao, G. Wang and Y. Gao, *J. Power Sources*, 2011, **196**, 3124.
- 35 I. d. A. Rodrigues, J. P. I. De Souza, E. Pastor and F. C. Nart, *Langmuir*, 1997, **13**, 6829.

