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Transparent superhydrophobic hollow films (TSHFs) with superior thermal stability and moisture resistance†

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Highly transparent superhydrophobic hollow films (contact angle, 165.7° and sliding angle, 2.1°) have been prepared by using candle soot as a template, followed by methyltrimethoxysilane (MTMS) chemical vapor deposition (CVD) and calcination at 450 °C, without the addition of dibutyl tin laurate (DBTL). The influence of deposition time on the superhydrophobicity and transparency of the as-prepared films was discussed herein to get the optimum performance film. The film shows high transparency (transmittance close to 90%) when it was coated on a glass substrate. What is more, it shows good thermal stability and superior moisture resistance as well, retaining its superhydrophobicity after calcining up to 500 °C or even exposure to ambient conditions for 30 days. ²⁹Si NMR, XPS spectra and thermal analysis verify the existence of methyl groups linked with silicon matrix on film (low surface energy), while AFM and TEM analysis confirm the sub-100 nm roughness and hollow structure of the optimum film.

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1. Introduction

Wettability is an important property for solid surfaces in many industrial processes as well as in our daily life. Superhydrophobic surfaces, with contact angle (CA) greater than 150° and sliding angle (SA) lower than 5°, exhibit excellent water repellent properties and have attracted significant attention both in scientific research and industrial exploitation due to their potential applications in impermeable textiles, self-cleaning coatings, lab-on-chip devices, microfluidic devices and outdoor antennas or windows.^{1–7} Hierarchical micro/nanosized roughness and low surface energy are the common causes of the observed superhydrophobicity of surfaces, which have been proved by some researchers during their studies in biological surfaces, such as the lotus leaf, the gecko, the namib desert beetle and water strider.^{8–13} The effect of hierarchical roughness and surface energy on the wettability of artificial and natural superhydrophobic surfaces has been demonstrated as well by Li *et al.* in thermodynamic approaches.^{14–16}

In the past few years, the fabrication of superhydrophobic surfaces has achieved a major progress.^{17–19} Superhydrophobic surfaces have been obtained from various routes, including etching and lithography, sol-gel processing, template-based method, physical/chemical vapor deposition (PVD/CVD), electrochemical deposition, self-assembly routes, electrospinning, phase

separation means, and others.^{20–30} Yet, some problems still exist in the approaches described above, especially the optical transparency, which is crucially important for the application of superhydrophobic surfaces in solar cell panel and window treatment fields. However, few researchers paid their attention to improve the transparency of superhydrophobic surfaces.

It is a challenge for us to prepare superhydrophobic surfaces with good transparency. Superhydrophobic surfaces with high optical transparency have potential applications in the solar cell panel and the treatment of windows. To attain superhydrophobic surfaces, the surface roughness is necessary, but surface roughness also induces light scattering when the light falls on the surfaces. Light scattering is responsible for the blurring of superhydrophobic surfaces. Thus, to get transparency, light scattering should be lowered. In order to minimize the light scattering, surface roughness must be reduced below the wavelength of the incident light. Sub-100 nm roughness, which can lower the scattering intensity of incident light and give consideration to both superhydrophobicity and transparency, has been proved to be an effective route for producing superhydrophobic surfaces with high transparency in visible light. A few groups have made progress in designing sub-100 nm roughness and preparing transparent superhydrophobic surfaces in the recent past, such as the assembly and co-assembly of sub-100 nm nanoparticles by a simple dip-coating method, depositing fluorosilane on the sub-100 nm roughness to lower the surface energy, combining chemical vapor deposition (CVD) and template-based method.^{31–40} But most of above cases are limited by their strict preparation conditions, multi-step processing procedures, costly materials or poor ability to scale-up in practice. And the traditional silicon-

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containing film had to use dibutyl tin laurate (DBTL), which is very harmful.

Thermal stability is also an important property for the film, while most of the TSHFs change their properties when the treating temperature is too high. Zhang *et al.* reported multiwalled carbon nanotube@poly(dimethylsiloxane) (MWCNTs@PDMS) can only keep the superhydrophobic under 360 °C.³⁸ And Lin *et al.* reported CTMS can only keep the superhydrophobic under 400 °C.⁴¹ The low thermal stability limits the application of TSHFs.

As mentioned above, there are few reports in literature focusing on the development of simple method for the preparation of highly transparent superhydrophobic hollow films by using easily obtained raw materials. Herein, we synthesized highly transparent superhydrophobic hollow films by using candle soot as template, followed by methyltrimethoxysilane (MTMS) CVD and calcination at 450 °C. The influence of deposition time on the superhydrophobicity and transparency of as-prepared films was discussed hereafter to get the optimum performance film. The optimum performance film has high water contact angle (165.7°), low sliding angle (2.1°) and high transmittance (closes to 90%) when it was coated on glass substrate without using harmful DBTL. The selected film possesses an excellent thermal stability and moisture resistance, which retains its superhydrophobicity after calcining up to 500 °C or even exposing to ambient condition for 30 days.

2. Experimental

A general schematic of the preparation of transparent superhydrophobic hollow films is given in Fig. 1. The route and experiments for synthesis are described in detail in ESI† (online). Various technologies including ²⁹Si CP MAS NMR, X-ray photoelectron spectroscopy (XPS), thermal analysis, atomic force microscopy (AFM) and transmission electron microscopy (TEM) were conducted to confirm and characterize the structure, composition and morphology of the optimum film. The detailed characterizations are described in detail in ESI† (online).

3. Results and discussion

3.1. FT-IR spectra of TSHFs

The transparent superhydrophobic hollow films were scratched from substrates for further FT-IR analysis. Herein, we only measured TSHF samples from CVD time higher than 6 h. This is due to that the amount of sample is too small to get for analysis when the CVD time is lower than 6 h. The FT-IR spectra of TSHF samples from different CVD time (8, 10, 12, 16, 20 and 24 h) are displayed in Fig. S1 (ESI†). All samples show strong absorption bands between 1047 cm⁻¹ and 1137 cm⁻¹ and a peak located at 439 cm⁻¹, which are related to the stretching and bending modes of Si–O–Si groups, respectively. Besides, peaks at 1274 cm⁻¹ and 773 cm⁻¹ can be observed on the spectra of TSHFs as well, indicating the existence of –Si–CH₃ groups on samples. The peaks around 2978 cm⁻¹ are due to stretching and bending vibrations of –CH₃ groups, while the weak flat peak at 3475 cm⁻¹ is related to the slight remaining –OH groups on TSHF samples. The FT-IR spectra verify the existence of –Si–CH₃ groups, which also indicate minute amount of hydroxyl groups on TSHFs sample.

3.2. Water CA, SA and transparency measurement of TSHFs

The transparent superhydrophobic hollow films (TSHFs) were prepared by using candle soot as template, followed by MTMS CVD and calcination at 450 °C. The influence of different CVD time on the water CA, SA and transparency of the obtained films is displayed in Fig. 2, and the inset photographs are the water drop profiles of corresponding samples. As shown in Fig. 2(a), the water contact angle of films increases by adding the CVD time, reaching 165.2° when the CVD time is 10 h, after that, the water contact angle of films increases slightly by further raising the CVD time of MTMS. Meanwhile, the sliding angle of films also responds to the CVD time changes, which decreases from 5.6° for TSHF-8 to the minimum value of 2.1° for TSHF-24. Note that too low CVD time can not achieve superhydrophobic films, such as 2, 4 and 6 h samples. It might be due to unformed shell and insufficient methyl groups on the shell are obtained for low

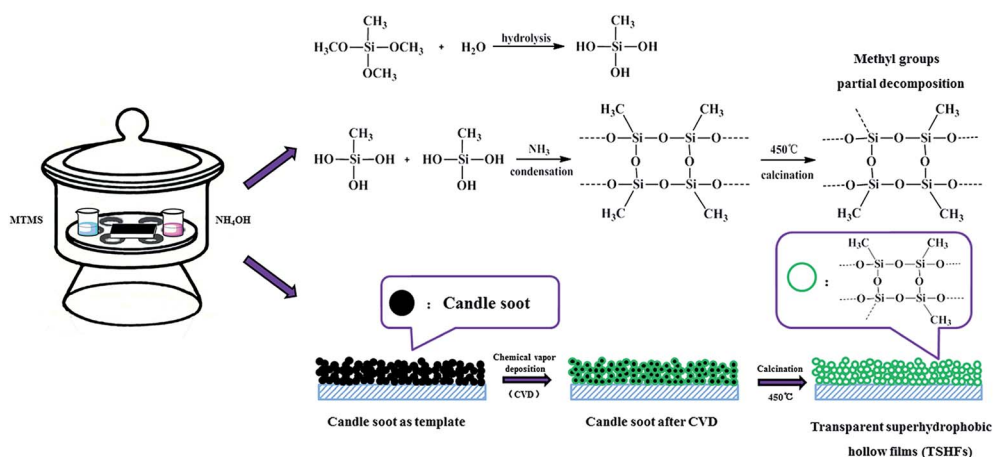


Fig. 1 A general schematic of the preparation of transparent superhydrophobic hollow films.



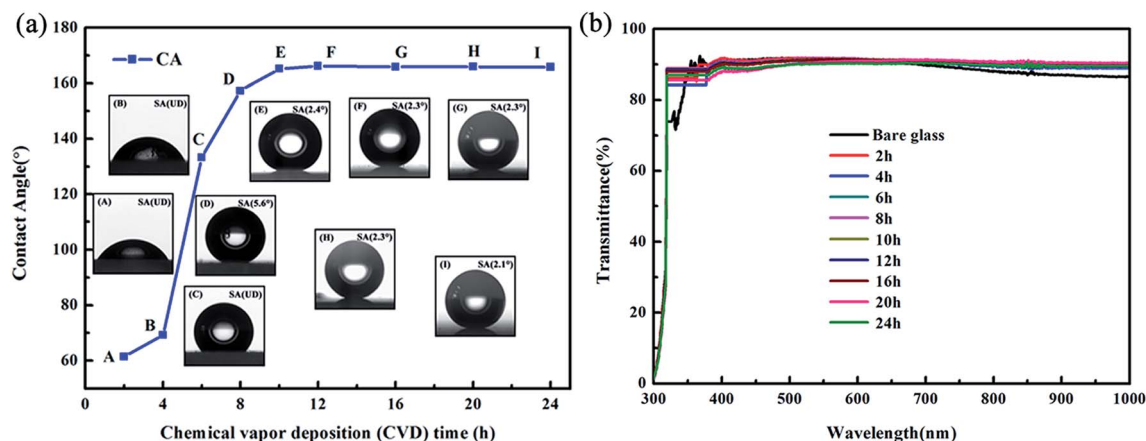


Fig. 2 The influence of different CVD time on the water contact angle (CA), sliding angel (SA) and transparency of films (a), note that the word "UD" in the images of the water drop profiles means undeterminable; the transmittance of bare glass and corresponding TSHFs coated glasses (b).

CVD time, which are not strong enough to render the films superhydrophobicity.

Fig. 2(b) presents the transmittance of bare glass and corresponding TSHFs coated glasses. It was found that the transmittance of TSHFs coated glasses is almost unchanged by adding the CVD time. All coated glasses maintain their optical transparency in the visible wavelength (transmittance closes to 90%) when compared with bare glass. Even in high visible wavelength (>750 nm), the transmittance of coated glasses is somewhat higher than bare glass. It is ascribed to that sub-100 nm roughness from hollow films reduces the reflection of incident light and increases the transmittance. Similar observation has also been reported in the former publications.^{41,45}

To aid the clearness, the transmittance of bare glass and TSHFs coated glasses in three visible wavelengths is summarized in Table S1 (ESI†). Clearly, TSHF-24 possesses the highest water contact angle (165.7°), lowest sliding angle (2.1°) and similar transparency with other film samples. Taking the water contact angle (CA), sliding angle (SA) and transparency of films into account together, TSHF-24 is chosen hereafter to investigate the structure, elemental composition, thermal stability, morphology and moisture resistance of as-prepared films.

3.3. ²⁹Si CP MAS NMR, XPS spectra and thermal analysis of TSHF-24

²⁹Si CP (cross-polarization) MAS NMR is employed herein to get the detailed information of the binding structures among MTMS, as shown in Fig. S2 (ESI†). Strangely, two weak Q species' peaks can be detected on the ²⁹Si NMR spectrum of TSHF-24, indicating the presence of siloxane (Q⁴, Si(OSi)₄), free silanols (Q³, Si(OSi)₃OH) groups on the sample, which is in conflict with the former publication.⁴⁴ It is assigned to that the methyl groups partially decompose and form Q species when calcined at 450 °C in air to remove the candle soot. Bidental (T²) and tridental (T³) anchoring structures, detected at −56.7 (T²) and −65.9 (T³) ppm, are adopted in the reaction among MTMS. Additionally, the peak strength of tridental (T³) anchoring structure is higher than that of bidental (T²) structure,

indicating the high reactivity of MTMS. The ²⁹Si NMR spectrum of TSHF-24 also confirms the presences of methyl groups (T² and T³) on TSHF sample, which is in conformity with the FT-IR analysis above. The chemical shift δ_{Si} and the anchoring structures among MTMS are summarized in Table S2 (ESI†) to aid clearness.

XPS spectra are displayed as well in Fig. S3† to get the surface elemental composition of as-prepared film. The photoelectron peaks of O 1s, C 1s and Si 2p are located at 531.7 eV, 283.7 eV and 102.9 eV, on the spectrum of TSHF-24 sample. The spectra of O 1s and C 1s show only one peak at 531.7 and 283.7 eV, assigning to −O*−Si− and −Si−C*H₃ groups, respectively.^{42–46} The spectrum of Si 2p can be divided into two peaks. A peak at 102.7 eV is related to −O−Si*−CH₃, while the peak at 104.1 eV is the corresponding peak of Si*O₂.^{42–46} The XPS spectra in Fig. 5 indicate the existence of methyl groups and silica matrix (Si*O₂) on TSHF-24 sample, which is in good accordance with the result of ²⁹Si NMR.

The thermal analysis of TSHF-24 is also shown in Fig. 3 to get the thermal stability and grafting density of methyl groups on

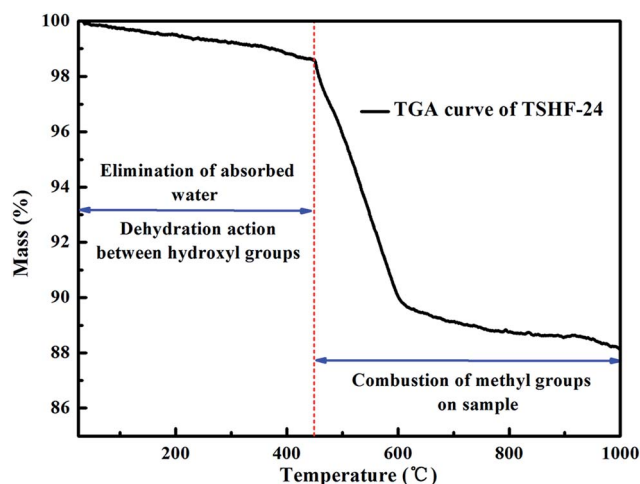


Fig. 3 The thermal analysis of TSHF-24 sample.



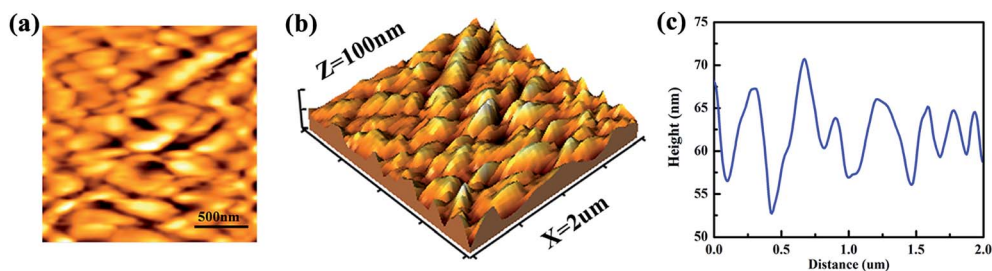


Fig. 4 2D AFM image (a), 3D AFM image (b) and height profile curve (c) of TSHF-24 sample.

sample. The slight weight loss up to 450 °C is due to the elimination of absorbed water and dehydration action between hydroxyl groups, the weight loss between 450 and 1000 °C is ascribed to the combustion of methyl groups on TSHF-24 sample. The whole weight loss of TSHF-24 is 11.61%, and the weight loss assigning to methyl groups is 10.18%. The grafting density of methyl groups on TSHF-24 can be calculated according to the following equation:

$$\text{Methyl (mol g}^{-1}\text{)} = \frac{\text{wt}_{\text{methyl}}}{\text{MW}_{\text{methyl}}}$$

$\text{MW}_{\text{methyl}}$ stands for molecular weight of methyl groups, 15 g mol^{-1} , while $\text{wt}_{\text{methyl}}$ represents the weight loss related to methyl groups. Thus, the grafting density of methyl groups on film can be obtained, $6.79 \times 10^{-3} \text{ mol g}^{-1}$.

3.4. AFM analysis of TSHF-24

A low surface energy and a sub-100 nm roughness have been considered to be account for the observed superhydrophobicity and transparency of TSHF-24, and the methyl groups linked with silica matrix (verified in Section 3.3) on the surface of film endow the TSHF-24 sample a low surface energy. AFM analysis

of TSHF-24, as shown in Fig. 4, is conducted to verify the hierarchical structure and sub-100 nm roughness of film.

It can be observed from 2D AFM image that the surface of substrate has been coated homogeneously by TSHF-24 sample (Fig. 4(a)). In addition, a large quantity of sharp bumps stands on the surface of TSHF-24 sample (Fig. 4(b)), indicating hierarchical roughness of film. The height profile curve verifies the result of 3D AFM image, from which a frequently fluctuant height profile can be seen for TSHF-24 sample. The RMS roughness value of TSHF-24 can be calculated from the software, about 9.7 nm, indicating the sub-100 nm roughness of film. So it could be deduced from the AFM images that the surface of TSHF-24 sample is hierarchical structure and sub-100 nm roughness.

3.5. The TEM analysis and the optical photographs of water droplets experiment

Fig. 5 presents the TEM images of TSHF-24 samples before and after calcination, the inset images are the closer look of corresponding samples. The optical photographs of water droplets on the treated glass are also displayed in Fig. 5. Clearly, candle soot exhibits a particle chain structure, which is made up of 42

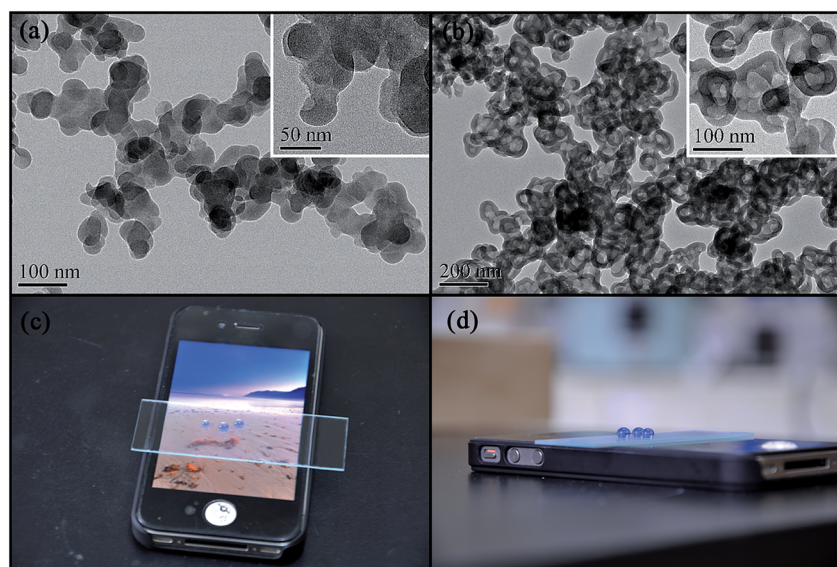


Fig. 5 TEM images of TSHF-24 samples before (a) and after (b) calcination (the inset images are the closer look of corresponding samples). The optical photographs of water droplets on the treated glass (c and d).



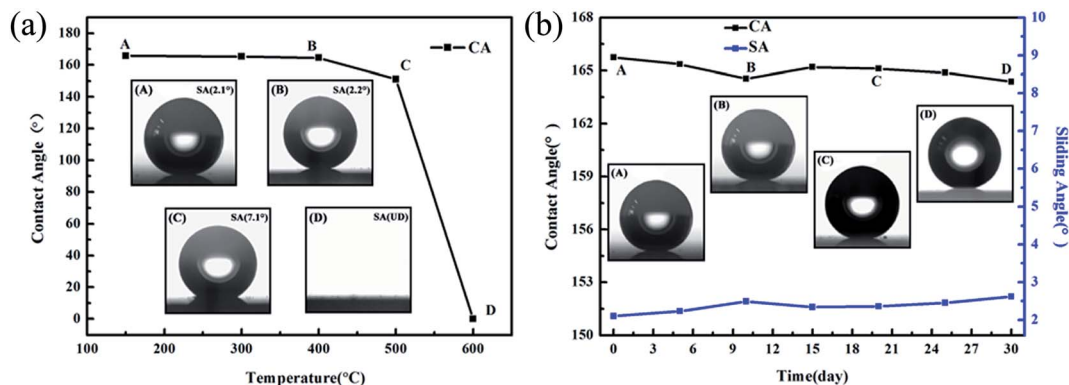


Fig. 6 The thermal stability (a) and moisture resistance (b) of TSHF-24 film.

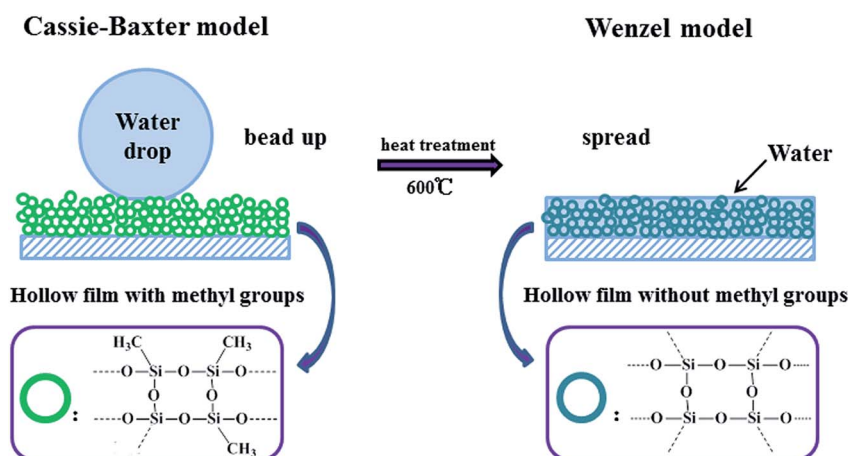


Fig. 7 The mechanism for the wettability change of TSHF-24 film after heat treatment at 600 °C.

± 5 nm spherical carbon balls. On these carbon balls, a distinct core-shell structure (Fig. 5(a)) can be observed on before treatment. After calcining in air at 450 °C, the template is removed, and a hollow structure film (Fig. 5(b)) can be seen subsequently. The thickness of shell can be measured (about 20 ± 5 nm) as well from Fig. 5. Additionally, the dyed water droplets are rolling up on TSHF-24 coated glass (Fig. 5(c) and (d)) and the images underneath the coated glass are still visible (Fig. 5(c)), indicating the superhydrophobicity and transparency of obtained hollow film. Fig. 5 demonstrates the hollow structure, transparency and superhydrophobicity of TSHF-24 film.

3.6. The thermal stability and moisture resistance of TSHF-24

The thermal stability of film, displayed in Fig. 6(a), was performed by calcining the coated glass at the temperature of 150, 300, 400, 500 and 600 °C under air atmosphere for 1 hour, respectively. The water contact angle of film decreases slowly from 165.7° to 164.5°, and the sliding angle of film declines slightly from 2.1° to 2.2° when raising the treating temperature of film up to 400 °C, which might be due to the dehydration action between hydroxyl groups and elimination of the combined water in film. The superhydrophobicity of film shows

apparent decrease at 500 °C, when the treating temperature is up to 600 °C, the wettability of film switches from superhydrophobic (>160°) to superhydrophilic (0°). The mechanism for the wettability change of film is proposed in Fig. 7. As known that Cassie-Baxter model and Wenzel model are common models for explaining the wettability of surfaces. In Wenzel's model, the rough surface could be wetted by driving off the trapped air by water, and the adhesion strength between water droplet and surface is large. However, in Cassie-Baxter model, the trapped air in rough surface could be hardly cleared by water, and the surface is difficult to be wetted. The adhesion strength between water droplet and surface is weak and even could be neglected in Cassie-Baxter model.

Cassie-Baxter model:⁴⁷

$$\cos \theta^c = f(\cos \theta_0 + 1)^{-1}$$

Wenzel model:⁴⁸

$$\cos \theta^w = r \cos \theta_0$$

where θ^c is contact angle from Cassie-Baxter model, θ^w is Wenzel contact angle and θ_0 is the equilibrium contact angle on



Table 1 Comparison with other superhydrophobic films in literatures

	Contact angle (°)	Sliding angel (°)	Transmittance (%)	Reference
TSHFs	165.7	2.1	90	This work
MWCNTs@PDMS	165	3	83	38
APTEOS	159.9	4.5	91	41
SPHSNs	160	1	89	49
(SSNs) ₁ /(SPHSNs) ₂	156	5.0	98.5	50
SCPs/DDs-SNPs	153	5	92.67	51
POTS	164	1	92.8	52

flat surface. Herein, r is the roughness factor, while f is the fraction of liquid–solid contact. Note that f decreases with the increase of the surface roughness. When $\theta_0 > 90^\circ$, θ^w and θ^c increase with the increase of surface roughness. When $\theta_0 < 90^\circ$, θ^w and θ^c decrease with the increase of surface roughness.

Clearly, the wettability of prepared film conforms to the model of Cassie–Baxter when the treating temperature is up to 500 °C. The methyl groups from MTMS molecules lower the surface energy of film, while hierarchical sub-100 nm roughness of film increases the trapped air fraction and decreases fraction of liquid–solid contact. As a result, the film exhibits superhydrophobicity and the water droplets are rolling up on film. When the film experiences calcination at a temperature of 600 °C, the methyl groups from MTMS are burned out, and the rough surface can be obtained. The film is in accord with Wenzel's model subsequently, on which the water could spread simply and shows superhydrophilicity. The results obtained from Fig. 6(a) are in good accordance with the thermal analysis above.

The moisture resistance of film, displayed in Fig. 6(b), is conducted by exposing the film to ambient condition for 30 days. Clearly, the prepared film possesses a good moisture resistance, which maintains its superhydrophobicity (higher than 165°) even after exposing to ambient condition for 30 days. What is more, it shows that the TSHFs film has very good performance after making comparison with other superhydrophobic films in literatures (Table 1). The TSHFs film has the biggest CA, and it also has very small SA and high transmittance at the same time.

4. Conclusions

In summary, the obtained optimum film exhibits superhydrophobicity (CA and SA, 165.7° and 2.1°), high transparency (transmittance closes to 90%), good thermal stability (up to 500 °C) and superior moisture resistance, without the usage of DBTL. And it is interesting that the film switches from superhydrophobic (>160°) to superhydrophilic (0°) when it is calcined at 600 °C. Further work will give insight into the relationship between experiment parameters (e.g. CVD time and candle soot thickness) and morphology (e.g. surface roughness and shell thickness) of as-prepared films.

Conflicts of interest

There are no conflicts to declare.

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