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Mechanisms of the polyol reduction of copper(II) salts depending on the anion type and diol chain length†

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The mechanisms of the polyol reduction of copper(II) compounds were investigated by various systematic experiments. The course of the reduction in ethylene glycol strongly depends on the anion present in solution. Elemental copper can only be obtained in high yields starting from $\text{Cu}(\text{NO}_3)_2$, $\text{Cu}(\text{OAc})_2$, or $\text{Cu}(\text{OH})_2$; but not from CuCl_2 or CuSO_4 . Intermediate compounds were observed, namely Cu_2O , and the alkoxide compounds $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ and $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$. $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$ was characterised by single-crystal X-ray diffraction. It is a 2D coordination polymer with bridging bidentate acetate and bidentate deprotonated EG ligands leading to vertex- and edge-linked CuO_5 polyhedra. Alkoxide compounds were also detected for higher glycol ethers. With growing glycol ether chain length the stability of alkoxide intermediates decreases, enhancing the stability range of Cu_2O . Because the majority of copper species is precipitated as copper(II) alkoxide compounds or Cu_2O before the reduction to elemental copper occurs, only few nuclei are formed in solution leading to microparticles instead of nanoparticles. Much smaller but strongly agglomerated particles were observed under alkaline reaction conditions.

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Introduction

Polyalcohols (polyols) can act as mild reducing agents for many metal cations, yielding metallic micro- or nanoparticles from easily available metal salts. This so-called polyol process has been known for over three decades and was successfully applied to produce a variety of mono- and multimetallic materials.^{1–3} In these reactions, the polyol normally acts as both solvent and reducing agent. Moreover, as a surface-active substance, it can also attach to the formed particles and affect the product's particle size and morphology.⁴ Polyols such as ethylene glycol (EG) have good chelating properties for metal cations, which compensate the lower polarity compared to water, and therefore the solvation behaviour of metal salts in water and polyols is quite similar.⁴ Applying polyols as high-boiling reaction media allows synthesis temperatures up to 320 °C facilitating the formation of crystalline nanoparticles without post-treatment.⁴

The redox potential of a polyol decreases with rising temperature.^{5,6} While noble metal ions are reduced at comparatively low temperatures, even at room temperature in the case of Pd^{2+} ,⁷ the reduction of less noble metals cations such as Co^{2+} , Ni^{2+} , or Cu^{2+} requires higher temperatures. For a given temperature the redox potential increases with increasing chain length of the polyol.⁸ At higher temperature the chelating properties of polyols have to be considered when planning a reduction reaction. Coordination compounds with polyol ligands that might represent intermediate species in the polyol process have been crystallised for several of the less-noble metals. During the reduction of $\text{Co}(\text{OH})_2$ or $\text{Ni}(\text{OH})_2$ in EG a solid phase precipitates before the reduction to the respective metals occurs.⁹ This solid is in equilibrium with solvated M^{2+} ions and is regarded as a metal reservoir governing the nucleation and growth of the final particles. Systematic reduction experiments of cobalt(II) salts in EG revealed the importance of the type of anion and the deprotonation of the polyol for the progress of the reduction reaction.^{10,11}

Like cobalt and nickel, copper particles can be obtained in polyols at elevated temperature. Copper and copper oxide nanomaterials, as cheap alternatives to precious metals, have found wide attention with respect to their potential application in catalysis, photovoltaics, and optics in the last years.¹² As copper(I) is fairly stable, its emergence in the course of the reaction is not unexpected: Cu_2O precipitates, e.g., as an intermediate when CuO^{13} , or CuSO_4 and 3 eq. of NaOH^{14} are

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heated in EG. Increasing the temperature leads to the formation of the final Cu particles. A solid Cu-EG compound similar to the Co-EG and Ni-EG complexes was described in the literature.¹⁵ By heating copper acetate in diethylene glycol (DEG) a Cu-DEG compound was obtained.¹⁶ However, these compounds were not structurally characterised.

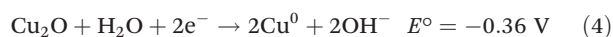
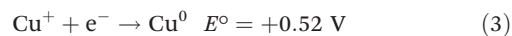
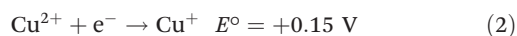
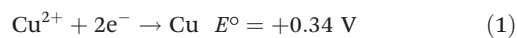
The knowledge about copper(II) species and their coordination modes in polyols is fragmentary. It was shown that the Jahn-Teller distorted tris-chelated complex $[\text{Cu}(\text{EG})_3]^{2+}$ is formed in a solution of $\text{Cu}(\text{CF}_3\text{SO}_3)_2$ in EG.¹⁷ Chloride anions have a higher coordinating affinity towards copper(II) than triflate anions, which is consistent with DFT calculations indicating the formation of the complex $[\text{CuCl}_2(\text{EG})_2]$ instead of $[\text{Cu}(\text{EG})_3]^{2+}$.¹⁸ The coordination ability of other anions has not been considered yet.

To establish a general understanding of the coordination and reduction behaviour of copper(II) in polyols, we carried out a series of systematic experiments. Several copper(II) precursor salts were reacted in EG, DEG, triethylene glycol (TrEG), and tetraethylene glycol (TEG) at different temperatures. Solid reaction intermediates were isolated and structurally characterised.

Results and discussion

Reduction of copper(II) in ethylene glycol

The following reduction potentials, which are only valid for aqueous solutions, have to be considered for the reduction of Cu^{2+} to Cu^0 .¹⁹ Reduction potentials of noble metal cations were found to follow the same order in EG and water.^{5,6} It is reasonable to assume that the order is also the same for the potentials considered here.



As reported before,²⁰ the reduction of Cu^{2+} ions to Cu^0 by polyols proceeds not in one step according to (1), but in two. Initially Cu^{2+} is reduced to Cu^+ (2). Although Cu^+ ions should be unstable and readily react to Cu^0 according to the more positive standard potential (3), insoluble Cu_2O immediately precipitates from the reaction solution. Cu_2O has a more negative reduction potential (4) and gets reduced to Cu^0 at a higher temperature.

The reflux temperature which can be reached for a given reaction mixture strongly depends on the water content. Polyols are hygroscopic and therefore contain a small amount of water. In this study the water content was kept low by using freshly supplied polyols. Water was also introduced by the hydrated copper salts that were used as starting materials. When 5 vol% of water were added to reaction mixtures of EG and suitable copper(II) compounds (e.g. acetate and hydrox-

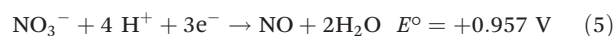
ide), the boiling point dropped to ca. 160 °C and no elemental copper was obtained under reflux conditions. However, microwave-assisted heating of mixtures in a closed vessel to 210 °C led to the formation of elemental copper. Hence, a small quantity of water does not change the general reduction behaviour but limits the boiling temperature. The water content has to be kept low to enable a complete reduction to elemental copper under ordinary reflux conditions.

The reduction of copper(II) precursor salts in EG was monitored at different temperatures. It was observed, that the anion of the copper salt has a major influence on the reduction reaction. The reduction of Cu^{2+} to Cu^0 does not proceed when starting from **copper(II) chloride**. Reacting CuCl_2 in EG for 2 h at 210 °C yielded no metallic copper but a clear green solution. This can be explained by the coordination affinity of chloride ligands towards the Cu^{2+} ion, which leads to the formation of solution species like $[\text{CuCl}_2(\text{EG})_2]$.¹⁸ This comparatively stable complex has a lower reduction potential than $[\text{Cu}(\text{EG})_3]^{2+}$, which prevents the reduction of Cu^{2+} to metallic copper at the given temperature. A decreasing reduction potential with increasing chloride concentration is known from copper chloride complexes in aqueous solution.²¹ Although elemental copper could not be produced, CuCl was obtained by reacting CuCl_2 in DEG at 240 °C and TrEG at 280 °C, respectively. Powder X-ray diffraction (PXRD) showed no other crystalline phases. The light green solution turns dark greenish-brown upon heating, and above 220 °C it abruptly turns into a white suspension of CuCl .

Copper(II) sulfate can be transformed to elemental copper by refluxing in EG at 210 °C. Copper was obtained only in low yields (max. 40%), and a small amount of Cu_2O was observed as a byproduct even after refluxing for 60 min (Fig. S1†). Conducting the reaction at 240 °C in DEG did not significantly change the product's purity or yield.

Obviously, CuCl_2 and CuSO_4 are not well suited for the precipitation of metallic copper in boiling EG without further modifying the process. Syntheses starting from these materials described in the literature always required the addition of 3–4 eq. of hydroxide.^{14,18}

The light blue solution of **copper(II) nitrate** in EG does not change up to 160 °C. At 180 °C the formation of nitrous gases is observed, and after 20 min a brown suspension occurs. The precipitate isolated by centrifugation mainly consists of Cu_2O (Fig. 1). The dispersion is rather stable and difficult to separate completely by centrifugation. This agrees with the previously described colloidal stability of oxide nanoparticles in polyols up to 20 wt%.²² A small reflection in the powder patterns can be assigned to copper(II) oxalate $\text{Cu}(\text{C}_2\text{O}_4) \cdot x\text{H}_2\text{O}$ ($x = 0-1$, PDF number 21-297). IR spectra confirm the presence of carboxyl groups (Fig. S2†). The formation of the byproduct is enabled by the oxidation of EG to oxalate under simultaneous reduction of nitrate to nitrous gases according to (5).



Under reflux conditions at 210 °C a dark red suspension of pure elemental copper is formed after some minutes. No Cu_2O



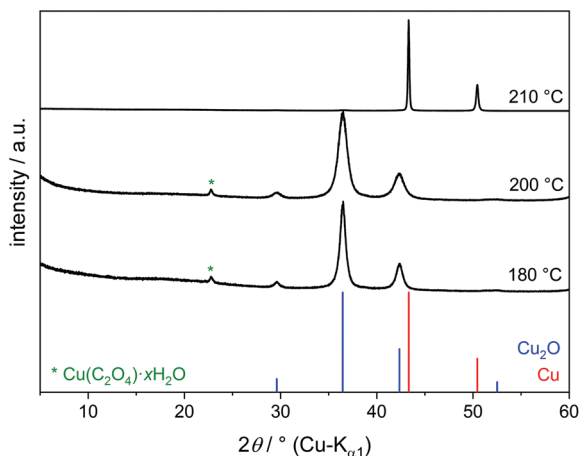


Fig. 1 PXRD patterns of reaction products of $\text{Cu}(\text{NO}_3)_2 \cdot 2.5\text{H}_2\text{O}$ in EG at various temperatures.

was detected by PXRD (Fig. 1). However, the surface of copper particles is easily oxidised by air if no precautions are taken to prevent this.¹² Such a surface layer of copper oxide is likely to be present for all the products in this study since the reactions were carried out under air and without surfactants. The amorphous surface layer is probably only a few nanometres thick. Rietveld refinements of some of the finally obtained copper powders, mixed with a known amount of silicon standard material, showed no notable amorphous content in the products.

Copper(II) hydroxide does not dissolve in EG even upon prolonged heating but transforms into an EG-containing solid. The PXRD patterns of the products obtained at 120–180 °C show primarily one intense broad reflection at $2\theta = 10.8^\circ$ (Fig. 2). A Cu–EG compound with a similar powder pattern was obtained from copper acetate in EG and described as a precursor for the production of nanostructured CuO. However, this compound was not further characterised.¹⁵

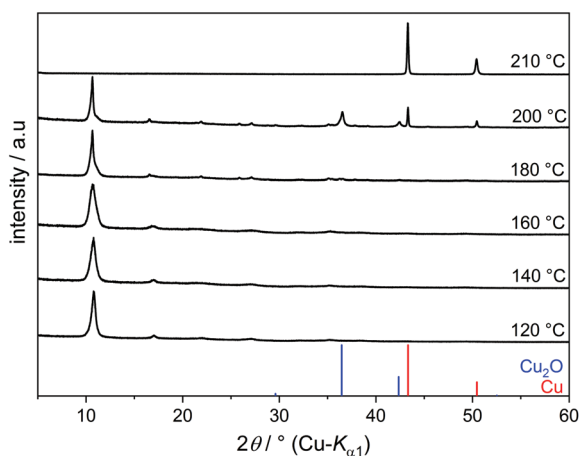


Fig. 2 PXRD patterns of reaction products of $\text{Cu}(\text{OH})_2$ in EG at various temperatures.

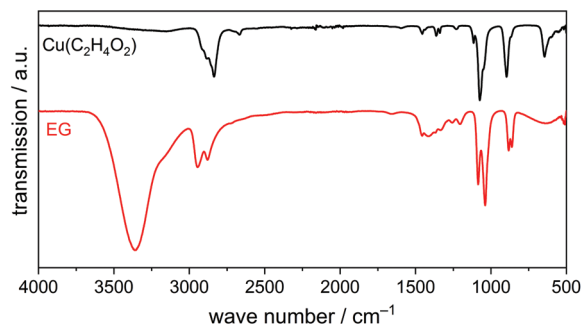
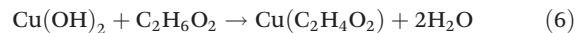


Fig. 3 IR spectra of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ and liquid EG.

The IR spectrum of this precipitate (Fig. 3) shows some absorption bands similar to those of liquid EG: bands due to CH stretching ($2950\text{--}2850\text{ cm}^{-1}$), C–O stretching ($1100\text{--}1000\text{ cm}^{-1}$), C–C stretching and CH_2 deformation (*ca.* 900 cm^{-1}) are clearly visible in the spectrum. The absence of the broad OH stretching band at $3600\text{--}3000\text{ cm}^{-1}$ and C–OH deformation bands at $1450\text{--}1300\text{ cm}^{-1}$, on the other hand, indicates fully deprotonated EG molecules. The CH stretching vibrations are split as compared to liquid EG most probably because of a chelating coordination of the EG molecules. In combination with results from elemental analysis (measured: C 18.1 wt%, H 2.9 wt%, calculated: C 19.4 wt%, H 3.3 wt%) and the Cu content determined from the CuO residual after heating in air (determined: Cu 52.3 wt%, calculated: Cu 51.4 wt%), we propose that the solid alkoxide yielded in this study is (ethylene glycolato)-copper(II), $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$.

Hence, copper hydroxide reacts with EG molecules to form the alkoxide upon the release of water:



The powder pattern of the produced alkoxide is very similar to those commonly found for layered hydroxide salts.²³ In the crystal structure of $\text{Cu}(\text{OH})_2$, strongly Jahn–Teller distorted CuO_6 octahedra are edge-linked to corrugated layers.²⁴ These layers are stacked along the *b* axis with a distance of 530 pm (Fig. 4a). Corrugated layers of edge-linked CuO_6 octahedra are also found in copper(II) layered hydroxide salts, *e.g.* $\text{Cu}_2(\text{OH})_3(\text{OAc})$.²⁵ Assuming a similar structure for $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$, the *d*-spacing of the most intense reflection corresponds to an interlayer spacing of 820 pm, probably due

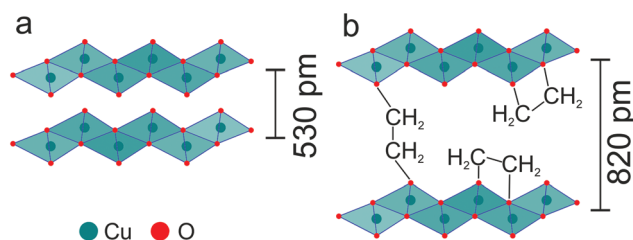


Fig. 4 Crystal structure of $\text{Cu}(\text{OH})_2$ (a) and a proposed structure model of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ (b).



to deprotonated EG ligands occupying the interlayer space (Fig. 4b). The EG ligands might show an intralayer–chelating or interlayer–bridging coordination mode or a mixture of both.

Similar to $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$, the solid alkoxides $\text{Co}(\text{C}_2\text{H}_4\text{O}_2)^{26}$ and $\text{Ni}(\text{C}_2\text{H}_4\text{O}_2)^{27}$ have been obtained by heating $\text{Co}(\text{OAc})_2$ and $\text{Ni}(\text{OH})_2$ in EG, respectively. Both compounds were described as layered brucite-like structures with turbostratic disorder. The interlayer spacing of 830 pm is close to the value for $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ determined in this study. The reported PXRD patterns and IR spectra show some similarities to those of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$.

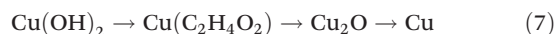
According to DTA-TG measurements in air (Fig. S3†), $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ is mainly reduced to Cu_2O in the first thermal decomposition step at about 200 °C. The oxidation of Cu_2O to the final product CuO leads to a mass increase between 220 °C and 400 °C. $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ is paramagnetic. The temperature dependence of the inverse susceptibility $\chi^{-1}(T)$ (Fig. S4†) can be modelled above 100 K by a Curie–Weiss law with the Curie constant $C_{\text{mol}} = 0.350 \text{ cm}^3 \text{ K mol}^{-1}$ and the Curie–Weiss temperature $\theta = -61 \text{ K}$, indicating antiferromagnetic coupling. The effective magnetic moment per Cu atom of $\mu_{\text{eff}} = 1.67 \mu_{\text{B}}$ derived from C_{mol} is in good agreement with the spin-only value of 1.73 expected for Cu^{2+} ions in a Jahn–Teller distorted octahedral environment. The $\chi^{-1}(T)$ curve shows a steady decrease upon lowering the temperature to 2 K which might be an indication of magnetic frustration.

When reacting $\text{Cu}(\text{OH})_2$ in EG at 120–180 °C, the main fraction (more than 80%) of Cu^{2+} precipitates as $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ and only a minor fraction of Cu^{2+} species remains dissolved. At 200 °C, Cu^{2+} is reduced to Cu_2O . The Bragg reflections of Cu_2O are broad, indicating small domain sizes. Elemental copper is obtained under reflux conditions at 210 °C.

SEM images (Fig. 5) show drastic changes in the size and morphology of the solids collected from the reaction mixtures. The starting material $\text{Cu}(\text{OH})_2$ consists of prolate nanoparticles. The appearance of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ changes between 120 °C and 200 °C from highly agglomerated almost isotropic

nanoparticles (Fig. 5b) to faceted crystals up to 3 μm length (Fig. 5c). The cubic Cu_2O particles with edge lengths up to 100 nm are well-stabilised in the reaction medium and rather difficult to precipitate completely by centrifugation. Due to the low Cu^{2+} concentration in solution prior to nucleation, the number of formed nuclei is limited and therefore copper sub-micron-sized particles (Fig. 5d) rather than nanoparticles are formed. These particles are poorly stabilised in EG and tend to agglomerate rapidly as the reaction time is increased.

Combining the PXRD and IR results the reaction scheme (7) of $\text{Cu}(\text{OH})_2$ with EG can be deduced.



By adding 2 eq. of NaOH to a suspension of $\text{Cu}(\text{OH})_2$ in EG, a clear, dark blue solution is formed. Cu^{2+} might exist in this solution in the form of square-planar $[\text{Cu}(\text{OH})_4]^{2-}$ or $[\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)_2]^{2-}$ complexes. The $[\text{Cu}(\text{OH})_4]^{2-}$ species exists in alkaline aqueous solutions.²⁸ $[\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)_2]^{2-}$ ions, which might be more stable in the EG solution because of the chelate effect, are known from crystals formed in alkaline mixtures of Cu^{2+} and group 1 or 2 metal ions in EG.^{29–32}

A copper alkoxide intermediate precipitates from the dark blue solution by heating to 120 °C. IR spectra and TG curves of this precipitate are very similar to those of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ obtained without NaOH, but the PXRD pattern (Fig. 6) is slightly different. Instead of one dominating reflection at $2\theta = 10.8^\circ$ one pronounced reflection at 10.4° and a less intense one at 11.2° are found, corresponding to interlayer distances of 850 pm and 790 pm, respectively. This suggests a structural variability of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$, which might arise from different coordination modes of the EG ligands influencing the interlayer spacing. The yield of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ (ca. 40%) is much lower than with pure $\text{Cu}(\text{OH})_2$ as the starting material due to a higher solubility of copper(II) species in the alkaline solution.

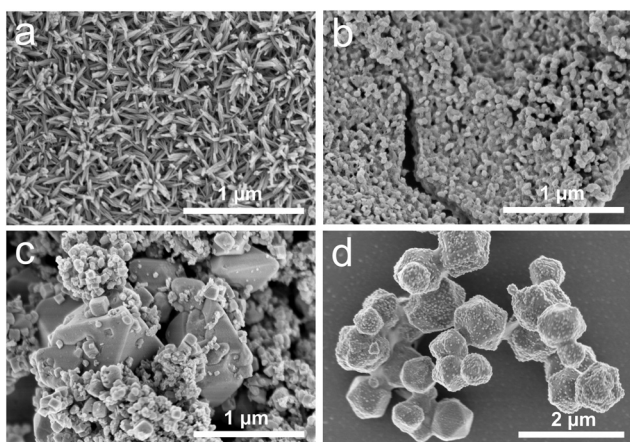


Fig. 5 SEM pictures of starting material $\text{Cu}(\text{OH})_2$ (a), $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$ (b), Cu_2O (c), and copper particles (d).

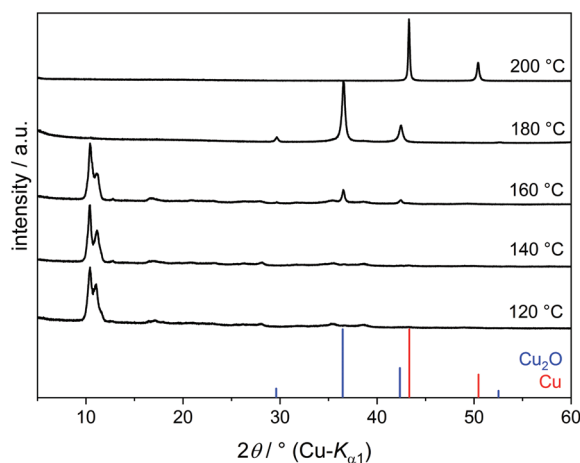


Fig. 6 PXRD patterns of reaction products of $\text{Cu}(\text{OH})_2$ and 2 eq. of NaOH in EG at various temperatures.



The reduction of $[\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)_2]^{2-}$ to elemental copper proceeds according to the reaction scheme (7) shown above, but at a significantly lower temperature compared to pure $\text{Cu}(\text{OH})_2$. According to PXRD results, Cu_2O begins to form at 160 °C and quantitatively transforms to Cu at 200 °C. This could be due to the solubility of the precursor as well as the increased pH value. Hydroxide ions are assumed to initiate the reaction by deprotonation of the α -carbon atom of the $(\text{C}_2\text{H}_4\text{O}_2)^{2-}$ ligand, thus leaving two electrons for the reduction of the metal cation.¹⁸

The majority of the finally obtained, severely agglomerated Cu particles (Fig. 7) have a significantly smaller size compared to those obtained from pure $\text{Cu}(\text{OH})_2$ due to a higher concentration of Cu^{2+} in the reaction solution prior to nucleation.

Copper(II) acetate completely dissolves in EG at *ca.* 80 °C yielding a bright turquoise solution. Further heating to 120 °C leads to the precipitation of a dark green crystalline compound (Fig. 8). The IR spectrum of this compound (Fig. 9) shows CH stretching ($2950\text{--}2850\text{ cm}^{-1}$), C–O stretching ($1100\text{--}1000\text{ cm}^{-1}$), C–C stretching and CH_2 deformation (*ca.* 900 cm^{-1}), but no OH stretching band, indicating fully deprotonated EG groups. Bands in the range $1600\text{--}1300\text{ cm}^{-1}$ are due to COO stretching and CH_3 deformation because of the presence of acetate ions.

The structure determination of the obtained product by single-crystal X-ray diffraction revealed the coordination polymer $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$. It crystallises in the triclinic space group $P\bar{1}$ and shows a 2D network structure generated by bridging bidentate acetate and deprotonated EG ligands (Fig. 10). The three crystallographically independent Cu atoms have approximately square-pyramidal coordination environments of O atoms. Four O atoms are located in distances of 192–200 pm forming a distorted square-planar array. The apical O atoms are more remote with distances between 224.6 pm and 251.4 pm. Valence-bond sums³³ of 2.11 (Cu1) and 2.04 (Cu2, Cu3) unambiguously indicate copper(II). The CuO_5 pyramids are linked *via* their edges and vertices forming

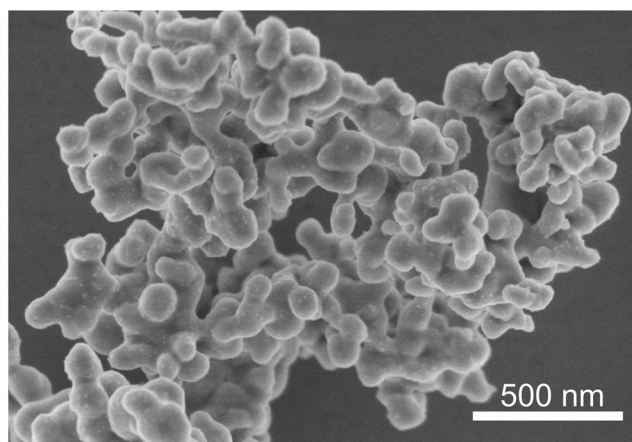


Fig. 7 SEM picture of copper particles obtained from $\text{Cu}(\text{OH})_2$ and 2 eq. of NaOH in EG at 200 °C.

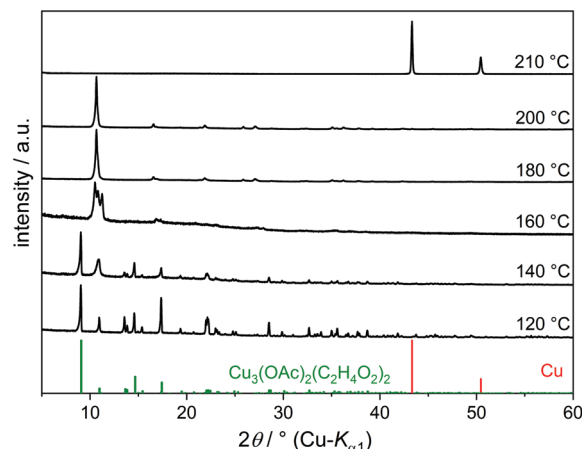


Fig. 8 PXRD patterns of reaction products of $\text{Cu}(\text{OAc})_2\cdot\text{H}_2\text{O}$ in EG obtained at various temperatures. The reference pattern for $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$ was simulated from the single-crystal structure data.

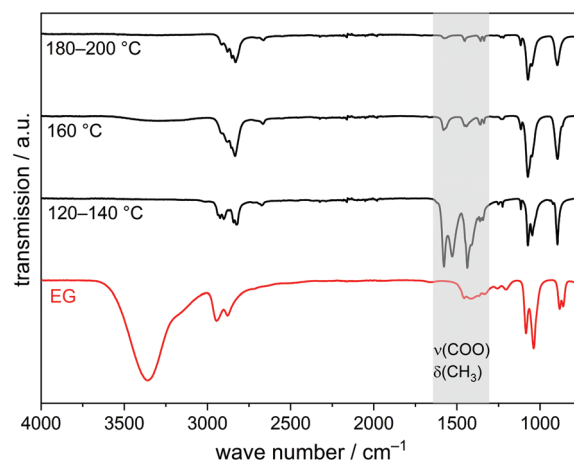


Fig. 9 IR spectra of reaction products from $\text{Cu}(\text{OAc})_2\cdot\text{H}_2\text{O}$ in EG obtained at different temperatures.

bands along [100] that consist of alternating Cu_6O_6 and Cu_4O_4 rings. The longest Cu–O distance of 251.4 pm (at Cu2) links the chains to a strongly corrugated layer parallel (001). The layers are stacked along the *c* axis and not linked *via* hydrogen bonds.

The mixed coordination of metal cations to solely double-deprotonated EG and acetate ligands, which occurs in $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$, was also reported for the compound $\text{Sb}(\text{OAc})(\text{C}_2\text{H}_4\text{O}_2)$.³⁴ Mono-deprotonated EG molecules were found in $\text{Zn}_2(\text{OAc})_3(\text{C}_2\text{H}_5\text{O}_2)$.³⁵ The occurrence of deprotonated EG is rather unusual, especially regarding the fact that no base was added to the reaction mixture. Acetic acid ($\text{p}K_a = 4.75$) is a much stronger acid than EG ($\text{p}K_a = 15.1$),¹⁹ and therefore the deprotonation of EG by acetate anions is an unlikely process. However, the stability of the product compound or the evaporation of acetic acid from the reaction solution may drive the reaction. The yield of the precipitated alkoxide is *ca.* 50%.



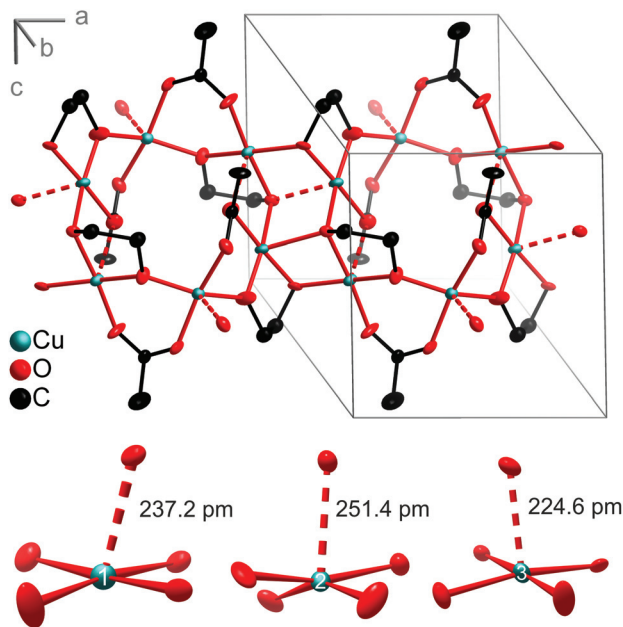
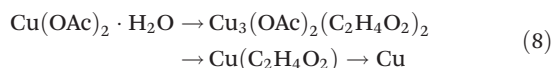


Fig. 10 Crystal structure of $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$ and coordination environment of the three independent Cu atoms. The longest Cu–O distances are represented by dashed lines. Hydrogen atoms are omitted for clarity. Ellipsoids represent a probability density of 90%.

DTA-TG measurements of dry $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$ in air (Fig. S5†) showed that the compound is stable up to 200 °C and decomposes in an exothermic reaction to Cu_2O and then to the final product CuO . $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$ shows paramagnetic behaviour. The χ^{-1} plot (Fig. S6†) can be modelled above 200 K by a Curie–Weiss fit. The obtained Curie constant $C_{\text{mol}} = 0.412 \text{ cm}^3 \text{ K mol}^{-1}$ corresponds to an effective magnetic moment $\mu_{\text{eff}} = 1.81 \mu_{\text{B}}$, which is slightly higher than the spin-only value of $1.73 \mu_{\text{B}}$ expected for Cu^{2+} ions.

The acetate–alkoxide coordination polymer is not stable at reaction temperatures of 160 °C and higher. Products isolated after 30 min at 180 °C and 200 °C show PXRD patterns (Fig. 8) and IR spectra (Fig. 9) characteristic for the already described compound $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$. The solid alkoxide is formed in high yields and only a small amount of Cu^{2+} remains in solution. Elemental copper is obtained under reflux conditions at 210 °C. Cu_2O could not be detected after 30 min because of its small stability range in this particular reaction system, which is accompanied by the high stability of $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$.

By combining the PXRD and IR results, reaction scheme (8) can be deduced for the reduction of copper acetate in EG.



Reduction of copper(II) in higher glycol ethers

Besides EG, the reduction of copper(II) hydroxide and acetate was also monitored in the longer-chain diols DEG, TrEG, and TEG. The resulting solid phases starting from copper hydroxide and acetate are summarised in Tables 1 and 2, respectively.

Table 1 Stability ranges (in °C) of obtained solid compounds starting from $\text{Cu}(\text{OH})_2$ after a reaction time of 30 min

Polyol	$\text{Cu}(\text{OH})_2$	Alkoxide	Cu_2O	Cu
EG		120–200	200	≥ 210
DEG		120–200	180–200	≥ 220
TrEG	120	120–180	180–200	≥ 220
TEG	120–160		140–220	≥ 200

Table 2 Stability ranges (in °C) of obtained solid compounds starting from $\text{Cu}(\text{OAc})_2 \cdot \text{H}_2\text{O}$ after a reaction time of 30 min

Polyol	Acetate–alkoxide	Alkoxide	Cu_2O	Cu
EG	120–140	160–200		≥ 210
DEG	120	140–200	200–220	≥ 220
TrEG			140–240	≥ 240
TEG			120–240	≥ 240

On reacting $\text{Cu}(\text{OH})_2$ with DEG, an alkoxide compound is obtained between 120 °C and 200 °C. PXRD patterns (Fig. 11) show an intense reflection at $2\theta = 11.0^\circ$ and weaker reflections at 16.2° ($d = 550 \text{ pm}$) and 32.7° ($d = 270 \text{ pm}$). IR spectra indicate the absence of acetate in this compound as they only show bands associated with the DEG molecules. No OH stretching band can be observed, although in some samples there is a broad signal around 3000 cm^{-1} arising from adsorbed water or DEG molecules. The formula deduced for this alkoxide compound is $\text{Cu}(\text{C}_4\text{H}_8\text{O}_3)$ which is in agreement with results from elemental analysis (measured: C 28.7 wt%, H 4.8 wt%, calculated: C 28.7 wt%, H 4.8 wt%) and the Cu content calculated from the CuO residual after heating in air (determined: 36.8 wt%, calculated: 37.9 wt%). The compound described here had already been synthesised and two structural models, containing either mono-deprotonated or a mixture of double-deprotonated and neutral DEG molecules, had been proposed.¹⁶ However, no or only very few OH

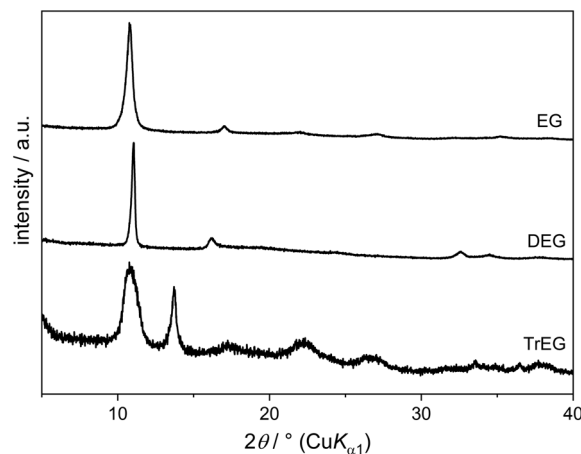


Fig. 11 PXRD patterns of alkoxides synthesised from the reaction of $\text{Cu}(\text{OH})_2$ with EG, DEG, and TrEG.



groups might be present in the structure according to our results.

The starting compound $\text{Cu}(\text{OH})_2$ can still be found in the solid products obtained with TrEG and TEG up to 140 °C and 160 °C, respectively. Similar to EG and DEG, an alkoxide compound is formed with TrEG. The PXRD pattern (Fig. 11) shows intense broad reflections at $2\theta = 10.8^\circ$ ($d = 820$ pm) and 13.7° ($d = 645$ pm).

In contrast to the previously discussed polyols TEG does not form an alkoxide compound with Cu^{2+} . Only $\text{Cu}(\text{OH})_2$, Cu_2O and copper can be detected in the obtained solid by PXRD. IR spectra show weak CH and C–O stretching bands, possibly due to TEG molecules adsorbed to the particle surface.

With copper acetate as precursor material (Table 2), only with EG and DEG alkoxides are formed. Similar to the reaction sequence in EG, an acetate–alkoxide compound is formed when copper acetate is heated in DEG to 120 °C. Both DEG and carboxylate vibration bands are visible in the IR spectrum of the precipitate. The powder diffractogram shows one intense and sharp reflection at $2\theta = 9.4^\circ$ ($d = 940$ pm). At 140 °C desert-rose shaped particles of $\text{Cu}(\text{C}_4\text{H}_8\text{O}_3)$ are obtained (Fig. S11†). $\text{Cu}(\text{C}_4\text{H}_8\text{O}_3)$ transforms to Cu_2O and finally to copper. No alkoxide compounds could be detected for TrEG and TEG. However, the reduction to Cu^+ starts at a much lower temperature and the stability range of Cu_2O is significantly enhanced. There is a small amount of a byproduct at 180 °C for TrEG and at 160 °C for TEG (Fig. S8†). IR spectra of the samples show intense carbonyl bands at 1650 cm^{-1} (Fig. S14†). Elemental copper is produced at higher temperatures when compared to copper acetate as the starting material. This is due to the stronger alkalinity of $\text{Cu}(\text{OH})_2$ which leads to deprotonation of EG molecules.

Coordination affinity of anions to copper(II) in ethylene glycol

Experiments including more than one anion in the solution showed some trends from which a general sequence of affinity towards Cu^{2+} could be deduced. Addition of NaOH to mixtures of CuCl_2 , CuSO_4 , $\text{Cu}(\text{NO}_3)_2$, or $\text{Cu}(\text{OAc})_2$ in EG first yields a light blue precipitate (2 eq. of NaOH), which subsequently dissolves resulting in a dark blue solution (4 eq. of NaOH). The reaction behaviour of these solutions is very similar to that of $\text{Cu}(\text{OH})_2$ or $[\text{Cu}(\text{OH})_4]^{2-}/[\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)_2]^{2-}$. With 4 eq. of NaOH, the complete reduction of the starting material to elemental copper can be achieved at 180 °C. Seemingly, anions other than hydroxide do not affect the reduction. However, with sulfate anions Na_2SO_4 was formed as byproduct during the reaction (Fig. S7†). Na_2SO_4 is only sparingly soluble in EG and has to be removed from the solid product by multiple washing cycles.

Addition of 2 eq. of NaOAc to solutions of CuCl_2 , CuSO_4 , and $\text{Cu}(\text{NO}_3)_2$ produces bright turquoise solutions which show the same reaction behaviour as those of pure $\text{Cu}(\text{OAc})_2$. Remarkably, CuCl_2 can be reduced to elemental copper at 210 °C (Fig. S7†). Because of its chelating ability, acetate seems to be a stronger ligand for Cu^{2+} than chloride.

The addition of 2 eq. of NaCl prevents the reduction of $\text{Cu}(\text{NO}_3)_2$ or CuSO_4 to elemental copper. Clear solutions without any solid products were obtained after 30 min at 210 °C. Chloride acts as a Lewis base and directly binds to Cu^{2+} , probably forming the solution species $[\text{CuCl}_2(\text{EG})_2]$, which leads to the same reaction behaviour as for CuCl_2 .

Based on these results, we propose a sequence of decreasing affinity towards Cu^{2+} ions in EG solutions: hydroxide > acetate > chloride > nitrate/sulfate.

Conclusions

Systematic experiments were carried out to study the anion-dependent course of the reduction of copper(II) salts in polyols. Elemental copper can only be obtained in high yields starting from $\text{Cu}(\text{NO}_3)_2$, $\text{Cu}(\text{OAc})_2$, or $\text{Cu}(\text{OH})_2$; but not with CuCl_2 or CuSO_4 . With EG, the intermediate compounds $\text{Cu}_3(\text{OAc})_2(\text{C}_2\text{H}_4\text{O}_2)_2$, $\text{Cu}(\text{C}_2\text{H}_4\text{O}_2)$, and Cu_2O were observed during the reaction. Experiments with higher glycol ethers show that increasing the polyol chain length decreases the stability of alkoxides and therefore enhances the stability range of Cu_2O . In the case of TEG, no alkoxide compound but phase-pure Cu_2O was detected in a very wide temperature range. The formation of Cu_2O or alkoxide compounds reduces the Cu^{2+} concentration in the reaction solution and therefore leads to a small number of nuclei resulting in micro- rather than nano-sized copper particles. Under alkaline reaction conditions, e.g. by addition of NaOH, the lowest reduction temperature and the smallest particle sizes are achieved due to a higher Cu^{2+} concentration prior to nucleation.

This study shows, that the anion-dependent reduction behaviour of copper in polyols is similar to the less noble metals cobalt and nickel, although the occurrence of the stable oxidation state +I is a peculiarity of the copper system. Less noble metals should be treated under alkaline conditions to ensure a complete reduction at the lowest possible temperature.

Experimental

Chemicals

Freshly supplied polyols were used as received: ethylene glycol (Fluka, 99.5%), diethylene glycol (Aldrich, 99%), triethylene glycol (Alfa Aesar, 99%), and tetraethylene glycol (Acros, 99.5%). Hydrated copper(II) salts were used for all syntheses: $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ (Aldrich, 99%), $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ (Aldrich, 99%), $\text{Cu}(\text{NO}_3)_2 \cdot 2.5\text{H}_2\text{O}$ (ABCR, 99.5%), $\text{Cu}(\text{OAc})_2 \cdot \text{H}_2\text{O}$ (Aldrich, 99%). $\text{Cu}(\text{OH})_2$ was prepared by precipitation of an aqueous solution of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ with 2 eq. of NaOH. The light blue powder was filtrated, washed several times with water and ethanol, and dried under vacuum at ambient temperature overnight. For some experiments, solutions of NaOH (Grüssing, 99%), NaOAc $\cdot 3\text{H}_2\text{O}$ (Grüssing, 99.5%), or NaCl (VWR Prolabo, 99.9%) in EG with a concentration of 1 mol l^{-1} were used.



Synthesis

For each experiment, the respective copper(II) compound (0.5 mmol) was mixed with 5 ml of the polyol in a flask equipped with a magnetic stirring bar and a reflux condenser. In some experiments a specified amount of NaOH, NaOAc·3H₂O, or NaCl solution was added, maintaining the total volume of 5 ml. Heating was performed with a CEM Discover-SP microwave device under open vessel conditions. The reaction temperature was controlled with the in-built infrared sensor. The mixture was heated to the desired temperature in 3 min, annealed for 30 min, and rapidly cooled down to 60 °C with compressed air. The solid products were collected by centrifugation, repeatedly washed with ethanol, and finally dried under vacuum at ambient temperature.

Powder characterisation

Powder X-ray diffraction (PXRD) patterns were collected with a transmission-geometry diffractometer Stadi P (Stoe) equipped with a Mythen 1K detector (Dectris) using Ge(111)-monochromated Cu-K α radiation ($\lambda = 154.056$ pm). The powder samples were placed between two layers of adhesive tape. Crystallographic data from the Inorganic Crystal Structure Database was used to calculate reference powder patterns for Cu (CSD 627114), Cu₂O (CSD 180846) and Cu(OH)₂ (CSD 68459). Fourier-transform infrared (FT-IR) spectra were measured with a Vertex 80 (Bruker) spectrometer in attenuated total reflectance mode in the wavelength range 4000–500 cm⁻¹. Scanning electron microscopy (SEM) was performed using a SU8020 electron microscope (Hitachi) equipped with a triple-detector system for secondary and low-energy backscattered electrons with an acceleration voltage of 3 kV. DTA-TG curves were recorded with a STA 409 thermal analyser (Netzsch). The samples were heated under air in an alumina crucible with a rate of 5 K min⁻¹ up to 600 °C. Elemental analysis was carried out with a Vario MICRO cube CHNS analyser (Elementar). Magnetisation measurements were conducted with a Helium-cooled vibrating sample magnetometer (Cryogenic) in DC mode in the range 2–300 K.

Single-crystal X-ray crystallography

A suitable single crystal of Cu₃(OAc)₂(C₂H₄O₂)₂ was mounted on the tip of a glass fibre and cooled to 100 K under flowing nitrogen gas. Intensity data were collected with a four-circle diffractometer Kappa Apex2 (Bruker) equipped with a CCD-detector using graphite-monochromated Mo-K α radiation ($\lambda = 71.073$ pm). Data were corrected for Lorentz and polarisation factors,³⁶ and multi-scan absorption correction³⁷ was applied. The structure was solved using SHELXT intrinsic phasing³⁸ and refined with full-matrix least-squares based on F_o^2 with SHELXL³⁹ using the OLEX2⁴⁰ program package. Refinement included anisotropic displacement parameters for all non-hydrogen atoms. Hydrogen atoms were refined with riding coordinates and displacement parameters. Graphics were developed with Diamond.⁴¹ The crystal structure data have

been deposited at the Cambridge Crystallographic Data Centre, deposit number 1857785.†

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- M. Figlarz, F. Fiévet and J.-P. Lagier, Process for the Reduction of Metallic Compounds by Polyols, and Metallic Powders Obtained by This Process, *US* 4539041A, 1985.
- F. Fiévet, J. P. Lagier, B. Blin, B. Beaudoin and M. Figlarz, *Solid State Ionics*, 1989, **32–33**, 198–205.
- F. Fiévet, J.-P. Lagier and M. Figlarz, *MRS Bull.*, 1989, **14**, 29–34.
- H. Dong, Y.-C. Chen and C. Feldmann, *Green Chem.*, 2015, **17**, 4107–4132.
- F. Bonet, C. Guéry, D. Guyomard, R. Herrera Urbina, K. Tekaia-Elhsissen and J.-M. Tarascon, *Int. J. Inorg. Mater.*, 1999, **1**, 47–51.
- F. Bonet, C. Guéry, D. Guyomard, R. Herrera Urbina, K. Tekaia-Elhsissen and J.-M. Tarascon, *Solid State Ionics*, 1999, **126**, 337–348.
- M. Heise, J.-H. Chang, R. Schönemann, T. Herrmannsdörfer, J. Wosnitza and M. Ruck, *Chem. Mater.*, 2014, **26**, 5640–5646.
- A. J. Biacchi and R. E. Schaak, *ACS Nano*, 2011, **5**, 8089–8099.
- G. Viau, F. Fiévet-Vincent and F. Fiévet, *Solid State Ionics*, 1996, **84**, 259–270.
- T. Matsumoto, K. Takahashi, K. Kitagishi, K. Shinoda, J. L. C. Huaman, J.-Y. Piquemal and B. Jeyadevan, *New J. Chem.*, 2015, **39**, 5008–5018.
- K. Takahashi, S. Yokoyama, T. Matsumoto, J. L. C. Huaman, H. Kaneko, J.-Y. Piquemal, H. Miyamura and J. Balachandran, *New J. Chem.*, 2016, **40**, 8632–8642.
- P. Lignier, R. Bellabarba and R. P. Tooze, *Chem. Soc. Rev.*, 2012, **41**, 1708–1720.
- F. Fiévet, F. Fiévet-Vincent, J.-P. Lagier, B. Dumont and M. Figlarz, *J. Mater. Chem.*, 1993, **3**, 627–632.
- J. Sun, Y. Jing, Y. Jia, M. Tillard and C. Belin, *Mater. Lett.*, 2005, **59**, 3933–3936.
- A. Cao, J. D. Monnell, C. Matranga, J. Wu, L. Cao and D. Gao, *J. Phys. Chem. C*, 2007, **111**, 18624–18628.
- A. Anžlovar, Z. C. Orel and M. Žigon, *J. Nanosci. Nanotechnol.*, 2008, **8**, 3516–3525.



- 17 S. E. Okan and P. S. Salmon, *Mol. Phys.*, 1995, **85**, 981–998.
- 18 K. J. Carroll, J. U. Reveles, M. D. Shultz, S. N. Khanna and E. E. Carpenter, *J. Phys. Chem. C*, 2011, **115**, 2656–2664.
- 19 *CRC Handbook of Chemistry and Physics*, ed. W. M. Haynes, CRC Press, Boca Raton, Florida, USA, 2016.
- 20 Z. C. Orel, E. Matijević and D. V. Goia, *J. Mater. Res.*, 2003, **18**, 1017–1022.
- 21 Y. Meng and A. J. Bard, *Anal. Chem.*, 2015, **87**, 3498–3504.
- 22 C. Feldmann and H.-O. Jungk, *Angew. Chem., Int. Ed.*, 2001, **40**, 359–362.
- 23 G. G. C. Arizaga, K. G. Satyanarayana and F. Wypych, *Solid State Ionics*, 2007, **178**, 1143–1162.
- 24 H. R. Oswald, A. Reller, H. W. Schmalle and E. Dubler, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 1990, **46**, 2279–2284.
- 25 S. Švarcová, M. Klementová, P. Bezdička, W. Lasocha, M. Dušek and D. Hradil, *Cryst. Res. Technol.*, 2011, **46**, 1051–1057.
- 26 N. Chakroune, G. Viau, S. Ammar, N. Jouini, P. Gredin, M. J. Vaulay and F. Fiévet, *New J. Chem.*, 2005, **29**, 355–361.
- 27 K. Tekaia-Elhsissen, A. Delahaye-Vidal, G. Nowogrocki and M. Figlarz, *C. R. Acad. Sci., Ser. II*, 1989, **309**, 469–472.
- 28 Y.-Y. H. Chao and D. R. Kearns, *J. Phys. Chem.*, 1977, **81**, 666–668.
- 29 N. Habermann, G. Jung, M. Klaassen and P. Klüfers, *Chem. Ber.*, 1992, **125**, 809–814.
- 30 C. P. Love, C. C. Torardi and C. J. Page, *Inorg. Chem.*, 1992, **31**, 1784–1788.
- 31 A. R. Pico, C. S. Houk, T. J. R. Weakley and C. J. Page, *Inorg. Chim. Acta*, 1997, **258**, 155–160.
- 32 J. H. Rivers, K. J. Carroll, R. A. Jones and E. E. Carpenter, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2010, **66**, m83–m85.
- 33 O. C. Gagné and F. C. Hawthorne, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2015, **71**, 562–578.
- 34 S. M. Biroš, B. M. Bridgewater, A. Villeges-Estrada, J. M. Tanski and G. Parkin, *Inorg. Chem.*, 2002, **41**, 4051–4057.
- 35 L. Poul, N. Jouini and F. Fiévet, *Z. Kristallogr.*, 1998, **213**, 416–418.
- 36 *SAINT+: Area-Detector Integration Software, Version 8.34A*, Bruker AXS Inc., Madison, Wisconsin, USA, 2013.
- 37 *SADABS: Area-Detector Absorption Correction, Version 2014/4*, Bruker AXS Inc., Madison, Wisconsin, USA, 2014.
- 38 G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Adv.*, 2015, **71**, 3–8.
- 39 G. M. Sheldrick, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2015, **71**, 3–8.
- 40 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.
- 41 K. Brandenburg, *Diamond 4: Crystal and Molecular Structure Visualization*, Crystal Impact GbR, Bonn, Germany, 2017.

