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Introduction

Perovskite materials have attracted considerable interest due to a wide range of technologically important properties displayed by materials with this structure-type, including superconductivity, ionic conductivity, colossal magnetoresistance, ferroelectric properties, and the ability to catalyse a range of reactions. In addition to this rich wealth of properties, perovskites also display a wealth of interesting and, at time, unexpected structural diversity. In particular, research on high $T_{\rm c}$ cuprate superconductors showed the ability of the perovskite structure to accommodate carbonate and other oxyanions (borate, nitrate, sulfate, phosphate).¹⁻¹² In these situations, the C, B, N, P, S of the oxyanion group resides on the perovskite B cation site, while the oxide ions of this group fill 3 (C, B, N)-4 (P, S) of the available 6 oxide ion positions around this site, albeit displaced so as to achieve the required geometry for the oxyanion. The incorporation of oxyanions into other perovskite transition metal containing systems has also subsequently been reported, e.g. $Sr(Co/Fe/Mn)O_{3-\delta}$, $La_{1-x}Sr_xCo_{0.8}Fe_{0.2}O_{3-\delta}$,

Carbonate: an alternative dopant to stabilize new perovskite phases; synthesis and structure of Ba₃Yb₂O₅CO₃ and related isostructural phases Ba₃Ln₂O₅CO₃ (Ln = Y, Dy, Ho, Er, Tm and Lu)⁺

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In this paper we report the synthesis of the new layered perovskite oxide carbonate, $Ba_3Yb_2O_5CO_3$. This phase is formed when $3BaCO_3 : 1Yb_2O_3$ mixtures are heated in air at temperatures ≤ 1000 °C, while above this temperature the carbonate is lost and the simple oxide phase $Ba_3Yb_4O_9$ is observed. The structure of $Ba_3Yb_2O_5CO_3$ was determined from neutron diffraction studies and consists of a tripled perovskite with double Yb–O layers separated by carbonate layers, the first example of a material with such a structure. Further studies showed that analogous $Ba_3Ln_2O_5CO_3$ phases could be formed for other rare earths (Ln = Y, Dy, Ho, Er, Tm and Lu). The results highlight the ability of the perovskite structure to accommodate carbonate groups, and emphasise the need to consider their potential presence particularly for perovskite systems prepared in lower temperature synthesis routes.

 $Ba_{1-x}Sr_xCo_{0.8}Fe_{0.2}O_{3-\delta}$, CaMnO₃ and La_{1-x}Sr_xMnO₃-type materials.¹³⁻³⁰ Recently the incorporation of sulfate and phosphate have also been reported in $Ba_2(In/Sc)_2O_5$ leading to new cubic perovskites with high oxide ion conductivity/proton conductivity.³¹⁻³³

In terms of these doping studies, it has been shown that oxyanions such as sulfate, borate, and phosphate exhibit higher thermal stability in the structure than carbonate. Therefore, work on incorporating carbonate has commonly employed reaction in evacuated quartz tubes to achieve reaction while preventing carbonate loss.¹⁵⁻¹⁹ As such, other researchers may be viewing such carbonate containing systems as "exotic" compounds, which would not be formed under conventional synthesis conditions. This is a particularly worrying omission, since there is a growing trend in research to move to lower temperature synthesis (e.g. sol gel) routes where carbonate may be present. Indeed evidence for the importance of carbonate in materials is illustrated by recent work on other structure types; for example Li et al. have shown that hexagonal "YAlO₃" is an oxide carbonate,³⁴ with a true composition of Y₃Al₃O₈CO₃, while orthorhombic Ba₂TiO₄ has also been recently shown to contain carbonate.35

While prior carbonate doped perovskite work has focused on systems containing transition metals, such as Cu, here we extend such studies to investigate whether a perovskite phase could be formed for non-transition metal systems, with synthesis in air rather than evacuated quartz tubes. In particular, given that a cubic perovskite, $Ba_2Yb_{1.5}P_{0.5}O_{5.5}$ has been

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recently shown to be stabilised by phosphate incorporation, we therefore investigated the possibility that a perovskite containing carbonate could be formed in the BaCO₃: Yb₂O₃ phase diagram at lower temperatures ($\leq 1000 \text{ °C}$) to allow carbonate incorporation. This led to the identification of a new phase with BaCO₃: Yb₂O₃ ratio of 3:1, and in this paper we report the synthesis and structural characterisation of this new carbonate containing perovskite system. In addition, in order to investigate the versatility for carbonate incorporation in perovskites, we also illustrate the formation of isostructural phases for a range of other rare earths.

Experimental

High purity BaCO₃, Yb₂O₃ were used to prepare Ba₃Yb₂O₅CO₃. The powders were intimately ground in a 3 : 1 BaCO₃, Yb₂O₃ ratio and heated initially to 900 °C for 12 h in air. They were then reground and reheated to 1000 °C in air for a further 24 h with regrinding every 12 hours. After the initial synthesis in air, we also investigated the formation of this compound under dry N₂ in order to allow higher temperature heat treatment by reducing CO₂ loss through reaction with moisture in the air. In this case, it was found that Ba₃Yb₂O₅CO₃ could be prepared in a shorter timescale (12 h) by employing a higher temperature (1050 °C). Note heat treatment in air at 1050 °C leads to large Ba₃Yb₄O₉ impurities due to carbonate loss. These synthesis routes were then used to make a wider range of isostructural Ba₃Ln₂O₅CO₃ systems (Ln = Y, Dy, Ho, Er, Tm and Lu).

Phase identification and initial structure determination was carried out by Rietveld profile refinement using powder X-ray diffraction data (XRD) collected on a Panalytical Empyrean diffractometer (Cu K α radiation) or a Bruker D2 diffractometer (Co K α radiation).

Raman data for Ba₃Yb₂O₅CO₃ were collected on a Renishaw inVia Raman microscope using a 532 nm laser in order to confirm the presence of carbonate in the material.

For the detailed structure determination on Ba₃Yb₂O₅CO₃, time of flight powder neutron diffraction (NPD) data were recorded on the HRPD diffractometer at the ISIS pulsed spallation source (Rutherford Appleton Laboratory, UK). Structure refinements using the NPD data were performed using the Rietveld method with the General Structure Analysis System GSAS-II suite of programs.³⁶

Results and discussion

Phase identification and structure determination of $Ba_3Yb_2O_5CO_3$

Initially a range of $BaCO_3 : Yb_2O_3$ mixtures with different ratios were investigated, and it was found that a new phase was observed with an optimum $BaCO_3 : Yb_2O_3$ ratio of 3 : 1. The formation of this phase was shown to be very sensitive to the synthesis temperature (Fig. 1). At temperatures up to 900 °C, no



Fig. 1 X-ray diffraction patterns for a 3:1 mixture of $BaCO_3: Yb_2O_3$ heated at different temperatures, showing the formation of a new layered perovskite phase, $Ba_3Yb_2O_5CO_3$ between 950–1000 °C, with decomposition of this phase to give $Ba_3Yb_4O_9$ at higher temperatures.

reaction was observed, and the XRD pattern showed simply the presence of the starting materials, $BaCO_3$ and Yb_2O_3 . Heating to 950 °C led to the appearance of the new perovskite phase, with the purity improving by heating at 1000 °C. At temperatures above 1000 °C, this compound was shown to decompose, and the formation of the simple oxide $Ba_3Yb_4O_9$ was observed, which suggested that the initial perovskite phase contained carbonate.

Indexing the pattern for this new phase gave a tetragonal perovskite-related cell which was tripled along the c direction. From the unit cell parameters obtained, a structure was then surmised based on the assumption that oxygen vacancies around the Yb sites would be less favourable. This predicted structure consisted of double Yb–O layers separated by carbonate layers.

This structure was then used as the starting point for the Rietveld refinement of the neutron diffraction data. The refinement was performed using the GSAS-II programme and constraints were initially placed upon the carbonate group within the system which included O–C–O angles constrained to 120° and C–O bond lengths constrained to 1.28 Å, in line with those expected for carbonate. In line with expectations that oxygen vacancies would be less favourable around the Yb site, the carbonate group was shown to be orientated with two of its oxygens directed towards the Yb sites, with the remaining oxygen equatorial to the carbonate layer. This gives an essentially square pyramidal Yb with a sixth longer bond to the carbonate oxygen.

There was no evidence for superlattice reflections/lowering of symmetry indicative of ordering of the orientation of the carbonate groups. In the final Rietveld refinement of the $Ba_3Yb_2O_5CO_3$ structure, the constraints on the O–C–O angles and C–O distances were removed (the only constraint included was that the sum of the occupancies of these O sites equal 3 in line with requirements for a carbonate group). This refinement led to a good fit to the data with the structural parameters and bond distances given in Tables 1 and 2, and the observed,

Table 1 Structural parameters for Ba₃Yb₂O₅CO₃

Atom	x	у	z	Site occupancy	Wyckoff position	$U_{\rm iso}$
C1 Ba1 Ba2 Yb1 O1 O2 O3 O4	$\begin{array}{c} 0.5 \\ 0.0 \\ 0.0 \\ 0.5 \\ 0.0 \\ 0.5 \\ 0.5 \\ 0.397(2) \end{array}$	$\begin{array}{c} 0.435(2) \\ 0.0 \\ 0.0 \\ 0.5 \\ 0.5 \\ 0.5 \\ 0.148(3) \\ 0.5 \end{array}$	$\begin{array}{c} 0.5\\ 0.0\\ 0.3358(6)\\ 0.1773(2)\\ 0.1995(3)\\ 0.0\\ 0.5\\ 0.4019(9) \end{array}$	$\begin{array}{c} 0.250\\ 1.000\\ 1.000\\ 1.000\\ 1.000\\ 0.210(14)\\ 0.270(7) \end{array}$	40 1a 2g 2h 4i 1c 40 8t	0.009(1) a 0.018(1) a a a a
Atom	U_{11}	U_{22}	U_{33}	U_{12}	U_{13}	U_{23}
Ba1 Ba2 O1 O2 O3 O4	$\begin{array}{c} 0.034(1)\\ 0.024(1)\\ 0.008(1)\\ 0.088(1)\\ 0.020(5)\\ 0.043(2) \end{array}$	0.034(1 0.024(1 0.026(1 0.088(1 0.020(5 0.052(2	$ \begin{array}{c} 1) & 0.015(1) \\ 0.030(2) \\ 1) & 0.071(2) \\ 1) & 0.004(2) \\ 5) & 0.008(3) \\ 2) & 0.052(2) \end{array} $	3) 0.000 2) 0.000 1) 0.000 2) 0.000 5) 0.000 2) 0.000	$\begin{array}{c} 0.000\\ 0.000\\ 0.000\\ 0.000\\ 0.000\\ -0.013(1)\end{array}$	0.000 0.000 0.000 0.000 0.000 0.000

 $R_{\rm wp}$ = 2.93% GOF = 1.71 *P*4/*mmm a* = 4.3258(2) Å, *c* = 11.9036(5) Å. ^{*a*} Anisotropic atomic displacement parameters.

Table 2 Selected interatomic distances for Ba₃Yb₂O₅CO₃

Bond	Bond length/Å	Bond	Bond length/Å
C1-O3 C1-O4 Ba1-O1 Ba1-O2 Ba2-O1	$\begin{array}{c} 1.242(16) (\times 1) \\ 1.376(10) (\times 2) \\ 3.212(2) (\times 8) \\ 3.059(0) (\times 4) \\ 2.704(5) (\times 4) \end{array}$	Ba2-O3 Ba2-O4 Yb1-O1 Yb1-O2 Yb1-O4	$\begin{array}{c} 2.985(5) (\times 2)^{a} \\ 2.871(7) (\times 4)^{a} \\ 2.179(1) (\times 4) \\ 2.110(3) (\times 1) \\ 2.711(11) (\times 1) \end{array}$

^a Dependent on orientation of carbonate gp.



calculated and difference neutron profiles shown in Fig. 2 (the data indicated a small amount (2.5 wt%) of Yb_2O_3 impurity, which was included in the final refinement). Given the presence of small Yb_2O_3 impurity, attempts were also made to refine the Yb site occupancy. This gave a value just above 1,

Raman studies of the Ba₃Yb₂O₅CO₃ material (Fig. 4) confirmed the presence of carbonate bands, with the most significant band (ν_1 symmetric stretch) seen at 1054.51 cm⁻¹. In comparison, BaCO₃ was also analysed with the carbonate ν_1 symmetric stretch observed at a higher energy of 1059.53 cm⁻¹. This difference confirms that the CO₃²⁻ peak was due to the Ba₃Yb₂O₅CO₃ material and not just small amounts of BaCO₃ reagent.

The final structural model is shown in Fig. 3, which shows that the structure, consists of double Yb–O layers separated by carbonate layers, representing a new example of an ordered



Fig. 3 (a) Refined structure of $Ba_3Yb_2O_5CO_3$ showing the split carbonate group positions attributed to local differences in carbonate orientation (b) structure illustrating a single carbonate orientation to illustrate clearer the bonding around the Yb and carbonate.



Fig. 4 Raman data for Ba₃Yb₂O₅CO₃.

perovskite stabilised through carbonate incorporation. The anisotropic atomic displacement parameters for O1 and O2 suggest local tilting of the YbO₆ octahedra to coincide with the adjacent carbonate orientation.

Synthesis of other Ba₃Ln₂O₅CO₃ (Ln = rare earth) phases

Following the successful synthesis and structure determination of $Ba_3Yb_2O_5CO_3$, the possible synthesis of analogous phases with different rare earths was investigated. Similar synthesis in air led to the successful formation of $Ba_3Ln_2O_5CO_3$ for Ln = Lu, Tm. For other rare earths (Ln = Y, Er, Ho and Dy) air synthesis led to samples with large $Ba_3Ln_4O_9$ or $BaLn_2O_4$ impurities. For these systems, better quality samples could be prepared by utilising the dry N₂ synthesis approach to limit loss of CO₂, and hence maintain the presence of carbonate in the sample. Attempts to prepare these phases for larger rare earths, *e.g.* Sm, Nd, resulted in no presence of the perovskite oxide carbonate phase. The structure obtained for Ba₃Yb₂O₅CO₃ was used for preliminary structure refinement to obtain cell parameters for these Ba₃Ln₂O₅CO₃ systems. Each refinement showed a good fit to the data. However, through the refinements it became apparent that each still contained small amounts (~3-4 wt%) of their respective Ln₂O₃ starting materials.



Fig. 5 (a) Variation in cell volume with Ln^{3+} ionic radius. (b) Variation in *a* cell length with Ln^{3+} ionic radius. (c) Variation in *c* cell length with Ln^{3+} ionic radius showing no systematic variation.

The refined unit cell parameters are given in Table 3, with observed, calculated and difference profile fits shown in ESI.[†] As expected, there is a linear increase in unit cell volume with increasing rare earth size (Table 3, Fig. 5a). However, the variation in the individual cell parameters follow a less systematic trend. While the *a* parameter shows a similar general increase with rare earth size (Fig. 5b), *c* follows a non-systematic variation (Fig. 5c). More detailed neutron diffraction structure determination would be required to explain this, which may be related to changes in local carbonate orientations. For example, it is possible that there might be some carbonate groups rotated 90° so that two oxygens are in the equatorial positions.

Conclusions

In this work, we illustrate the formation of perovskite oxide carbonate phases in the BaCO₃ : Ln₂O₃ phase diagram for a ratio of 3 : 1. These Ba₃Ln₂O₅CO₃ phases are shown to be new layered perovskites, with double Ln–O layers separated by carbonate layers, the first example of a material with such a structure. The results highlight the need to consider the potential incorporation of carbonate in perovskite materials prepared at temperatures ≤ 1000 °C. Thus it raises questions whether many perovskite systems prepared at low temperature (*e.g. via* sol–gel routes) may contain carbonate, affecting properties rather than simply previously considered morphological effects.

Conflicts of interest

There are no conflicts to declare.

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