The modular synthesis of rare earth-transition metal heterobimetallic complexes utilizing a redox-active ligand†

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We report a robust and modular synthetic route to heterometallic rare earth-transition metal complexes. We have used the redox-active bridging ligand 1,10-phentanoiline-5,6-dione (pd), which has selective N,N' or O,O' binding sites as the template for this synthetic route. The coordination complexes [Ln(hfac)3(N,N'-pd)] (Ln = Y [1], Gd [2]; hfac = hexafluoroacetylacetonate) were synthesised in high yield. These complexes have been fully characterised using a range of spectroscopic techniques. Solid state molecular structures of 1 and 2 have been determined by X-ray crystallography and display different pd binding modes in coordinating and non-coordinating solvents. Complexes 1 and 2 are unusually highly coloured in coordinating solvents, for example the vis-NIR spectrum of 1 in acetonitrile displays an electronic transition centred at 587 nm with an extinction coefficient consistent with significant charge transfer. The reaction between 1 and 2 and VCP2 or VCP2′ (CP = tetramethylcyclopentadienyl) resulted in the isolation of the heterobimetallic complexes, [Ln(hfac)3(N,N'-O,O'-pd)VCP2] (Ln = Y [3], Gd [4]) or [Ln(hfac)3(N,N'-O,O'-pd)VCP2′] (Ln = Y [5], Gd [6]). The solid state molecular structures of 3, 5 and 6 have been determined by X-ray crystallography. The spectroscopic data on 3–6 are consistent with oxidation of V(II) to V(IV) and reduction of pd to pd2− in the heterobimetallic complexes. The spin–Hamiltonian parameters from low temperature X-band EPR spectroscopy of 3 and 5 describe a 5A1 ground state, with a V(IV) centre. DFT calculations on 3 are in good agreement with experimental data and confirm the SOMO as the dν,v orbital localised on vanadium.

Introduction

The cooperativity between metal centres in multimetallic complexes, clusters and polymeric species gives rise to very remarkable chemistry in a diverse range of fields.1 The potential applications of complexes containing both transition metals and lanthanides are significant. In particular there is research interest in the design of rare earth-transition metal single-molecule magnets (SMM) and catalysts.2,3 The chemistry of multimetallic complexes containing f-elements is significantly less well developed than that of transition metals. This is attributable to the fact that the directed synthesis of f-d complexes is non-trivial. The inspiration for our approach was the Tb analogue of the [{(SiMe3)2N}Ln2N2K] molecule magnets (SMM) and catalysts.2,3

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could be used as a bipyridine equivalent to bind a second metal.14 This paper also reported the selective N,N′-pd binding of PdCl₂ and the first crystal structures of transition metal pd complexes (Chart 1). Eisenberg combined pd and bipyridine (bipy) to form the homobimetallic Pt complexes, [PtCl₂(N,N′-O,O′-pd)] (Chart 1). These were then used as precursors to synthesise [Pt(L)(N,N′-O,O′-pd)]Pt(di-tert-butyl-bipy)] (where L = dithiolate, dicatecholate or pd) and probe electronic structure.15

Directly relevant to this work, is the first report of the coordination and redox chemistry of group 4 and 5 metals with pd, by Pampaloni and Calderazzo.16 These groups subsequently published a number of studies demonstrating that pd will selectively bind N,N′,O,Os with a reducing organometallic precursor and can form homo- and heterometallic transition metal complexes.17–19 Prior to this work there are just two reports of N,N′-pd binding of Ln2O3 starting materials, their stability in H2O and facile dehydration by comparison to commercially available LnCl3·(H2O)6. The hfac (hfac = 1,1,1-5,5,5-hexafluoroacetylacetonate) ligand was chosen because [Ln(hfac)3(glyme)] are well-defined molecular species.22,23 The hfac ancillary ligand environment has good solubility and stability, with 19F NMR as an additional probe of solution behaviour. We note that while you can access [Ln(hfac)3(glyme)] directly from Ln2O3, the synthesis requires the addition of chelating polyethers to prevent aggregation.24

However, using a salt elimination reaction enables this synthetic route to be modular and applicable to other ligand environments. The coordination complexes [Ln(hfac)3(N,N′-pd)] were synthesised in high yield. Ln(OTf)3 (Ln = Y, Gd), pd and 3 equivalents of K(hfac) were combined dry, and THF added at room temperature with stirring, giving a colour change from yellow to deep green. The reaction was stirred at room temperature for 16 h, and THF removed in vacuo. Extraction into toluene, filtration to exclude K(OTf) and removal of toluene in vacuo gave the neutral complexes [Ln(hfac)3(N,N′-pd)] as pastel green solids in excellent yields (Ln = Y: 82% [1], Gd: 79% [2]) (Scheme 1). Elemental analyses of 1 and 2 crystallised from toluene at −35 °C support the [Ln(hfac)3(N,N′-pd)] formulation, with one toluene of crystallisation associated. 1 and 2 are soluble in both coordinating and non-coordinating solvents, to give green and yellow solutions, respectively.

**Results and discussion**

**Synthesis of [Ln(hfac)3(N,N′-pd)]**

Metal precursor choice is particularly important in f-element chemistry. In this work, we chose to use the Ln(OTf)3 (Ln = Y, Gd, OTf = CF3SO3) more commonly applied in organic synthesis as Lewis acid catalysts.21 This was due to the ease of preparation from inexpensive Ln2O3 starting materials, their stability in H2O and facile dehydration by comparison to commercially available LnCl3·(H2O)6. The hfac (hfac = 1,1,1-5,5,5-hexafluoroacetylacetonate) ligand was chosen because [Ln(hfac)3(glyme)] are well-defined molecular species.22,23 The hfac ancillary ligand environment has good solubility and stability, with 19F NMR as an additional probe of solution behaviour. We note that while you can access [Ln(hfac)3(glyme)] directly from Ln2O3, the synthesis requires the addition of chelating polyethers to prevent aggregation.24

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**Scheme 1** The modular synthesis of rare earth-transition metal complexes.
Spectroscopy of [Ln(hfac)]₃(N,N′-pd)]

The room temperature ¹H NMR spectrum of 1 in d₆-benzene is consistent with complexation to form the 1 : 1 adduct. The pd resonances in 1 at δ = 6.37, 7.53, and 9.16 ppm are shifted downfield relative to the pd free ligand and the aromatic splitting pattern (dd) is no longer observed for the protons neighbouring the N,N′-pd pocket. The expected ratio of 2 : 2 : 2 : 3 is observed between the pd resonances and the hfac singlet at δ = 6.23 ppm. The 19F NMR spectrum displays a singlet at δ = −76.9 ppm. The X-band EPR spectrum of 2 at room temperature (see ESI†) is common for Gd(III) and simulation gave giso = 1.989, close to that of the free electron, which is typical for Gd(III).23 The 8S0 ground state of Gd(III) has no angular orbital momentum, and so unlike its f-block counterparts, Gd(III) gives an EPR response at room temperature. The Evans Method was used to determine the magnetic moment of 2 in d₅-thf at room temperature. The corrected magnetic moments are in the range of 7.56–7.71µB for 2, consistent with Gd(III).

IR spectroscopy cannot definitively assign the coordination mode of the pd ligand but does provide a useful measure of its carbon–oxygen bond order. The reported ranges of νCO are: neutral pd = 1678 cm⁻¹, radical anion pd⁻ = 1500–1400 cm⁻¹ and diamion pd²⁻ = 1400–1300 cm⁻¹.14,16 The IR spectra of 1 and 2 have several absorptions in the νCO region, complex 1 at 1699, 1668 and 1647 cm⁻¹ and 2 at 1697, 1661, 1647 cm⁻¹. The νCO at 1647 is assigned to the hfac ligand based on the IR of closely related complexes, [Ln(hfac)₃(glyme)], which have very strong absorption between 1650 and 1660 cm⁻¹.23 The νCO at 1699 and 1697 cm⁻¹ are assigned as N,N′-pd. This is consistent with complexes formulated as, [PdCl₂(N,N′-pd)]13 [Fe(N,N′-pd)[PF₆]₂]10 and [MCl₄(N,N′-pd)] (M = Ti, Zr, Hf)16 which have strong absorptions in the range 1680–1705 cm⁻¹. We propose additionally that the νCO at 1668 and 1661 cm⁻¹ may result from the short contacts seen between the Ln metal centre and one pd-O atom of the neighbouring molecule (Fig. 3 and ESI†).

The deep green colour of 1 and 2 in coordinating solvents is an unexpected property as neither 4d⁰ Y(III) nor 4f⁷ Gd(III) strongly absorptions in the range 1680–1699 and 1697 cm⁻¹.13 

Crystallography of [Ln(hfac)]₃(N,N′-pd)]

Crystals of 1 suitable for X-ray diffraction were grown from either a saturated Et₂O solution at −35 °C over 7 days (1a) or a saturated toluene solution at −15 °C overnight (1b). Similarly crystals of 2 suitable for X-Ray diffraction were grown from either a saturated Et₂O solution at −35 °C overnight (2a) or a saturated toluene solution at −15 °C overnight (2b). The solid state molecular structures and selected parameters are shown in Fig. 1–3 or in the ESI†.

All structures adopt a distorted square antiprismatic geometry, and confirm N,N′-pd coordination. As expected the pd C–O distances in 1a–2b are unchanged from the free ligand (1.213(5) and 1.216(5) Å).28 The Ln–O(hfac) distances to the hfac ligands are consistent between both structures, and are in agreement with the crystal structures of other examples of Ln[hfac]₃ complexes in the range of 2.299(2)–2.343(3) Å.25,29 The Ln–N(pd) distances in structures 1a–2b of 2.501(3)–2.585 (8) Å are directly comparable to those reported in [Ln(tta)]₃.
(N,N-O,O'-pd)Pt(PPh3)2] (Ln = Er, Gd; tta = 1-thenoyl-5,5,5-
trifluoroacetylacetone) of 2.499(7)–2.539(7) Å, as the ionic
radius of Er(III) is an appropriate size match for Y(III).10
Complex 1 is the first crystallographically characterised
example of an yttrium pd complex. The Y-N(pd) distances
are also comparable to those seen in [Y(pyrrMe2-acac)3(phen)]
(pyrrMe2-acac = 1,3-dimethyl-1H-pyrazol-4-yl)acetyl-
acetone) of 2.562(4)–2.579(4) Å.24

The structure of 2b from toluene is a coordination polymer
as a result of N,N-O-pd binding to the next Gd in the chain
(Fig. 3). The only other structure in the CCDC that displays
this binding mode is the Ag(i) coordination polymer [Ag(ClO4)
(N,N-O-pd)]3.31 The Gd1–O2 distance is 2.830(2) Å, and the
coordination polymer exhibits a zig-zag structure with an angle
of 77.56° between Gd-pd planes. Analysis of short contacts in
1b reveals a Y1–O2 interaction within the sum of the van der Waals radii at 3.154 Å, giving a pseudo-coordination polymer
structure with an angle of 83.98° between Y–pd planes
(see ESI†). As a result of the N,N-O-pd binding, both 2b and
1b feature significant distortion within the pd ring.
The three rings of pd are no longer coplanar and there is an
out-of-plane O2-ring distortion and associated out-of-plane
displacement of O2.

Reduction chemistry of 1 and 2

The pd ligand can be sequentially reduced by two electrons,
first to the semiquinone radical anion pd− and then to the
fully reduced and diamagnetic catecholate pd2−; the reduction
potentials are −0.85 and −1.71 V vs. Fe+/Fe3+, respectively.10,32
Almost all published metal complexes of the reduced pd
ligand contain pd2− and the few examples of the radical anion
of pd− were generated in situ, usually under electrochemical
conditions and characterised by electronic spectroscopies such
as EPR and UV-vis-NIR. This is in spite of the ease of reduction
of pd to pd− relative to the structurally related 1,10-phenan-
throline10 and 2,2′-bipyridine.33

Complex 1 is unstable under standard electrochemical
conditions. Therefore, we used the reduction potentials of the free
pd ligand to rationalise the choice of organometallic reducing
agent.34 The first reversible oxidation process of V(II) to V(III) in
VCP3 to VCP2+ is a good match at −0.93 V vs. Fe+/Fe3+.35 We note
that VCP2 has been previously shown to reduce pd, with the
product formulated as [(O,O'-pd)VCP2] and in our hands this
reaction formed intractable dark green solids.16 Additionally,
we used a substituted vanadocene, VCP5 (Cp5 = tetramethyl-
cyclopentadienyl) to test whether the added steric bulk would
promote the formation of an ion pair. The substituted Cp5
ligand environment is also a good model for extending this
synthetic route to incorporate f-element organometallics.36

Synthesis of [Ln(hfac)3(N,N-O,O'-pd)VCP2]

A Schlenk flask charged with 1 or 2 and VCP2 was cooled to
−78 °C and Et2O was added with stirring. The colour of the
resultant solution was dark green-black. The reaction mixture
was warmed to room temperature and Et2O was removed
in vacuo. Recrystallisation of the solids from a mixture of
toluene and hexane at −15 °C gave [Ln(hfac)3(N,N-O,O'-pd)
VCP2] as black microcrystalline solids (Ln = Y: 56% [3], Ln =
Gd: 64% [4]). The analogous reaction between 1 or 2 and VCP3
(t = Me4), resulted in the isolation of dark green-black micro-
crystalline solids of [Ln(hfac)3(N,N-O,O'-pd)VCP3] (Ln = Y:
50%, [5] Ln = Gd: 42% [6] (Scheme 1). The composition and
purity of complexes 4–6 was confirmed by elemental analysis.

Spectroscopy of [Ln(hfac)3(N,N-O,O'-pd)VCP3]

Complexes 3–6 are paramagnetic and NMR silent. The Evans
Method was used to determine the magnetic moment of 3, 5
and 6, in d8-thf at room temperature. The corrected magnetic
moments are in the range of 1.71–1.74 μB for 3, 1.72–1.75 μB for
5 and 7.99–8.41 μB for 6. The data for 3 and 5 are consistent
with a single unpaired electron i.e. 3d1 V(nv), resulting from the
oxidative addition of V(nv) to V(nv) and the concomitant
reduction of pd by two electrons to pd2−. The data for 6 are
consistent with the value predicted by the spin only formula
for Gd(III) and V(III). This is chemical behaviour similar to the
oxidative addition of VCP2 to acetylene to form the
metallocyclopropane.37

The IR spectra of complexes 3–6, are consistent with pd
reduction. The νCO at 1699 and 1697 cm−1 assigned to neutral
N,N'-pd and at 1668 and 1661 cm−1 assigned to N,N'-O-pd are
no longer visible. Instead, complexes 3–6, display νCO in the
range 1370–1381 cm−1, consistent with the reported examples of
d2− complexes.14,16

An overlay of the electronic spectra of 3 and 4 collected in
MeCN are contrasted with the electronic spectra of 5 and 6
collected in THF and presented in Fig. 4. The d-d transitions at
740 and 640 nm usually present in pseudo-tetrahedral vanado
cenes\textsuperscript{18,39} are not directly observed as they are hidden
underneath the charge transfer transitions. The transitions are
assigned with the aid of computational analysis (vide infra).

Crystallography of [Ln(hfac)]\textsubscript{3}(N,N’-O,O’-pd)VCp\textsubscript{R}\textsuperscript{2+}

Crystals of 3, 5 and 6 suitable for X-ray diffraction were grown
by diffusion of hexane vapour into a toluene or Et\textsubscript{2}O solution
at \textdegree35 \textdegreeC or by cooling a toluene/hexane mixture to \textdegree35 \textdegreeC
over several days. The structures and selected structural para-
metric values are shown in Fig. 5 and 6 or in the ESL.\textsuperscript{1} Complexes 3, 5
and 6 feature yttrium or gadolinium in a distorted square anti-

tiprismatic geometry and N,N’-pd bound analogous to 1a and 2a.
The VCP\textsuperscript{R}\textsubscript{2+} units are O,O’-pd bound, resulting in a pseudo-tetra-
hedral geometry typical for V(IV) metallocenes.\textsuperscript{40,41} The metal to
ancillary ligand bonding metrics are normal. The five mem-
bered ring formed between the V and the O,O’-pocket of the pd
is a common structural motif. The V1–O1 and V1–O2 distances in
3, 5 and 6 are consistent with those observed in related com-
plexes [Na\textsubscript{2}(Et\textsubscript{2}O)\textsubscript{6}][V(dippCO)]\textsubscript{3}], [VCP\textsubscript{2}(acac)][OTf] and V(IV)
catecholates such as [NH\textsubscript{2}Et\textsubscript{3}]\textsubscript{3}V(cat)\textsubscript{3} and [H\textsubscript{2}PrDA]VO(cat)\textsubscript{3}
(cat = 1,2-dihydroxybenzene, PrDA = 1,3-diaminopropane).\textsuperscript{42–45}

The metrics that are important in assessing the reduction of
the pd ligand are C5–C6, C5–O1 and C6–O2. The C5–C6
bond distances in 3 of 1.40(3) \texttildeÅ, in 5 of 1.38(2) \texttildeÅ and in 6 of
1.387(3) \texttildeÅ are significantly shortened from that of 1b (1.524(6) \texttildeÅ)
consistent with double bond formation. The C5–O1 and
C6–O2 distances in 3, 5 and 6 are consistent with reduction to single C–O bonds.
We note the disparity in the C–O distances in 3 but do not think
it is chemically significant. These metrics in 3, 5 and 6 are
consistent with pd\textsuperscript{2+} and are essentially identical to those found in
[Ln(tta)\textsubscript{3}(N,N’-O,O’-pd)Pt(PPh\textsubscript{3})\textsubscript{2}] (Ln = Gd, Er).\textsuperscript{12}

EPR of [Ln(hfac)]\textsubscript{3}(N,N’-O,O’-pd)VCp\textsubscript{R}\textsuperscript{2+}

The room temperature X-band EPR spectra of 3 and 5 shown
in Fig. 7 are typical for V(IV) with an 8-line pattern from
coupling of the electron spin (S = 1/2) to the I = 7/2 nuclear
spin of the \textsuperscript{51}V (99.8% abundance). The spectral simulation of 3
was gave \(g_{iso} = 1.9799\) and \(A_{iso} = 65.8 \times 10^{-4}\) cm\textsuperscript{–1}, which are
synonymous with non-oxo vanadyl species.\textsuperscript{10,42,44} In contrast,
the spectrum of 5 exhibits the two overlapping 8-line signals;
simulations yielded an equal mix of one entity with \(g_{iso} = 1.9819\),
\(A_{iso} = 68.4 \times 10^{-4}\) cm\textsuperscript{–1} and the other \(g_{iso} = 1.9779\)
\(A_{iso} = 64.9 \times 10^{-4}\) cm\textsuperscript{–1} (Table 1).

The corresponding frozen solution spectra recorded in
toluene at 130 K clearly depicts a signal indicative of a single
chemical species of 3 and 5 (Fig. 8). The simulation para-

eters are nearly identical, with \(g = (1.9637, 1.9776, 1.9978)\)
and \(A = (106.6, 78.1, 11.0) \times 10^{-4}\) cm\textsuperscript{–1} for 3, and \(g = 1.9654,\)

Fig. 5 Molecular structure of 3. Hydrogen atoms and solvent of crystal-
lisation omitted for clarity. Thermal ellipsoids drawn at 50% probability
for pd, V and Y. Other ancillary ligand atoms depicted as ball and stick
due to CF\textsubscript{3} and Cp disorder (see cif). Selected distances (\texttildeÅ): Y1–N1
2.474(17), Y1–N2 2.482(17), C5–C6 1.396(15), C6–O2 1.24(2), C5–O1
1.40(2), V1–O1 1.972(16), V1–O2 1.971(16).

Fig. 6 Molecular structure of 6. Hydrogen atoms and solvent of crystal-
lisation omitted for clarity. Thermal ellipsoids drawn at 50% probability.
Selected distances (\texttildeÅ) for 5: Y1–N1 2.481(1), Y1–N2 2.49(1), C5–C6
1.38(2), C6–O2 1.35(4), C5–O1 1.34(2), V1–O1 1.975(8), V1–O2 1.977(9)
and 6: Gd1–N1 2.532(2), Gd1–N2 2.527(2), C5–C6 1.387(3), C6–O2
1.333(3), C5–O1 1.333(3), V1–O1 1.9869(18), V1–O2 1.9862(18).

Fig. 7 X-band EPR spectra of 3 and 5 recorded in toluene solution at
293 K (experimental conditions: frequency, 9.8634 GHz; power,
0.63 mW; modulation, 0.2 mT). Experimental data are represented by
the black line; simulation is depicted by the red trace.
metry, such that profile for differences in their spin-Hamiltonian parameters. The spectral formations of existence of one signal in the frozen matrix suggests two con-
entwined with the 8-line pattern of V(IV). The spectrum was have the same \(2A_1\) ground state but overall three-fold sym-
the result of bond polarisation inherent to the relatively Lewis acidic V(IV) ion.41

**DFT calculations on 3**

Density functional theoretical (DFT) calculations were carried out on 3 \((S = 1/2)\). The optimised structure is in excellent agreement with the crystallographic structures of both 3 and 5 (Table S2†). The computed \(C-C\) and average \(C-O\) bond distances of 1.391 and 1.336 \(\text{Å}\), respectively, are representative of a bridging pd ligand in its dianionic (catecholate) form. The highest occupied molecule orbital (HOMO) is singly-occupied (SOMO) with 91% vanadium d character that depicts the d \(2\) orbital which is a 1 in \(C_{2v}\) symmetry as shown by EPR (vide supra). The Mulliken spin population analysis shows +1.35 spins (\(\alpha\)-spin) located on the V(IV) centre with matching \(\beta\)-spin on the first coordination sphere atoms (Fig. 10). This is the result of bond polarisation inherent to the relatively Lewis acidic V(IV) ion.41

Time-dependent (TD) DFT calculations were carried out on 3 in order to assign the electronic spectra of this series of heterobimetallic compounds. Each complex (3–6) displays two ion and one V(IV) \(S = 1/2\) ion. In addition we observe no manifestation of exchange coupling between the Gd(u) and V(IV) centres such that \(J \approx 0\) (i.e. two uncoupled spins).

**Fig. 10** Mulliken spin population analysis for 3.

**Table 1** Spin–Hamiltonian \(g\)- and \(A\)-values\(^a\) derived from spectral simulation of complexes

<table>
<thead>
<tr>
<th>Complex</th>
<th>Spin centre</th>
<th>(g_{iso})</th>
<th>(A_{iso})</th>
</tr>
</thead>
<tbody>
<tr>
<td>3</td>
<td>(V (S = 1/2))</td>
<td>1.9799</td>
<td>65.8</td>
</tr>
<tr>
<td>5</td>
<td>(V^4 (S = 1/2))</td>
<td>1.9819</td>
<td>68.4</td>
</tr>
<tr>
<td></td>
<td>(V (S = 1/2))</td>
<td>1.9790</td>
<td>64.9</td>
</tr>
<tr>
<td>2</td>
<td>(\text{Gd (S = 7/2)})</td>
<td>1.989</td>
<td>—</td>
</tr>
<tr>
<td>4</td>
<td>(\text{Gd (S = 7/2)})</td>
<td>1.991</td>
<td>—</td>
</tr>
<tr>
<td></td>
<td>(V^5 (S = 1/2))</td>
<td>1.9779</td>
<td>72.0</td>
</tr>
<tr>
<td></td>
<td>(V (S = 1/2))</td>
<td>1.9950</td>
<td>72.0</td>
</tr>
</tbody>
</table>

\(^a\) In \(10^{-4}\) cm\(^{-1}\). \(^b\) Major conformer. \(^c\) Minor conformer.

**Fig. 8** X-band EPR spectra of 3 and 5 recorded in toluene solution at 130 K (experimental conditions: frequency, 9.4244 GHz; power, 0.2 mW; modulation, 0.1 mT). Experimental data are represented by the black line; simulation is depicted by the red trace.

**Fig. 9** X-band EPR spectrum of 4 recorded in toluene solution at 293 K (experimental conditions: frequency, 9.8717 GHz; power, 0.63 mW; modulation, 0.3 mT). Experimental data are represented by the black line; simulation is depicted by the red trace.
low-energy bands at ca. 800 and 600 nm (Fig. 4). These signature spectral features were well-reproduced in the TD-DFT analysis (Fig. 11). Both bands are ligand-to-metal charge transfer (LMCT) transitions between the same donor and acceptor orbitals. These are the $b_2$ symmetric HOMO–1, which is localised on the $O,O'$ end of the pd bridging ligand, and the LUMO, which is the vanadium d$_{xy}$ orbital (Fig. 11). The computed spectrum reveals the spin-up (α-spin) excitation at 767 nm which perfectly matches the experimental peak at 800 nm in 3. The corresponding spin-down (β-spin) excitation appears at 489 nm which is ca. 100 nm blue shifted from the experimental peak. This aforementioned polarisation inherent to Lewis acidic V(iv) sees the α-spin manifold is exchange stabilised relative to the β-spin manifold by virtue of the unpaired electron in the d$_{x^2−y^2}$ orbital. This is the same phenomenon that results in >1 spins located at the V(iv) centre in the spin population analysis (Fig. 10). The experimental polarisation is 0.54 eV is overestimated by the calculations at 0.92 eV. This is typical for DFT as energies of virtual (unoccupied) orbitals are less reliable especially given the TD-DFT computed spectrum is derived from energy differences between donor and acceptor orbitals based on the ground state calculation. This neglects additional factors that can (de) stabilise excited states, and therein fail to accurately reproduce electronic spectra of paramagnetic compounds. Nevertheless, this approximation is sufficient to assign the observed spectral features. The prominent band at 410 nm is tentatively assigned as a $\pi \rightarrow \pi^*$ transition from the pd$^{2-}$ ligand.

Conclusions

We have demonstrated the synthesis and first structural characterisation of [Ln(hfac)$_3$(N,N$'$-O$_2$O′-pd)] for both Y and Gd complexes. The synthetic route works in exactly the same way for both diamagnetic Y and paramagnetic Gd. Considering the size of Gd and Y, this synthetic route would be expected to be directly applicable to all the Ln elements from Sm–Lu. As the route is modular, a range of ancillary ligand environments would also be expected to be synthetically accessible. We have also shown the reduction of these coordination chemistry complexes using vanadocenes to form heterobimetallic rare earth-transition metal complexes. The spectroscopic data for [Ln(hfac)$_3$(N,N$'$-O$_2$O′-pd)] are consistent with oxidation of V(ii) to V(iv) and reduction of pd to pd$^{2-}$ in these complexes. This redox chemistry is driven by the $O,O'$-pd binding of the VCp$^2_2$ unit. Work is ongoing to isolate rare-earth transition metal complexes containing the radical anion of pd bridging the metal centres and to synthesise heterobimetallic f–f complexes using 4f and 5f elements.

Experimental

General details

All air-sensitive manipulations were carried out in an MBraun glovebox (O$_2$ and H$_2$O < 1 ppm) or by using standard Schlenk techniques under N$_2$. All glassware was dried at 130 °C overnight prior to use. Filter cannulas were prepared using Whatman 25 mm glass microfiber filters and were pre-dried at 130 °C overnight. Dry solvents were obtained using an Innovative Technology Inc. Pure Solv 400-5-MD solvent purification system (activated alumina columns). Solvents were sparged with N$_2$ and stored in ampoules over activated molecular sieves under N$_2$. Deuterated benzene and THF were dried by refluxing over K. Deuterated acetonitrile was dried by refluxing over CaH$_2$. Dry deuterated solvents were degassed by three freeze–thaw cycles, vacuum distilled, and kept in ampoules in the glovebox under N$_2$. The following starting materials were prepared according to literature procedures: V Cp$_2$, V Cp$_2$$^2$, Hexafluoroacetylacetonate (Hhfac) was purchased from Aldrich and used without further purification.

Physical methods

$^1$H NMR data were recorded on an AVIII 400 MHz instrument and were referenced internally to the appropriate residual protio-solvent and are reported relative to tetramethylsilane (δ = 0 ppm).$^{19}$F and $^{19}$F($^1$H) NMR spectra were recorded on a Bruker AVIII 400 MHz spectrometer and were referenced to CFC$_3$ (δ = 0 ppm). All spectra were recorded at a constant temperature of 300 K. Coupling constants (J) are reported in hertz (Hz). Standard abbreviations indicating multiplicity were used as follows: m = multiplet, d = doublet, t = triplet, s = singlet. UV/vis/NIR spectra were collected using a Shimadzu UV-3600 UV/vis/NIR spectrometer using anhydrous solvents, which were filtered through Celite® prior to use. ATR-IR spectra were collected using either a Shimadzu IRAfinity-1S or
a Shimadzu FTIR 8400S spectrometer. Abbreviations indicating strength of bands were used as follows: s = strong, m = medium, w = weak.

Elemental analysis was performed at the University of Strathclyde, and a combustible (tungstic oxide) was used to ensure complete combustion of all complexes.

Crystallographic data were collected for single crystals of $[\text{Y(hfac)}_3(N,N'-pd)]$ (1a,1b), $[\text{Y(hfac)}_3(N,N'-O,O'-pd)]$ (2), and $[\text{Gd(hfac)}_3(N,N'-O,O'-pd)]$ (3) on a Bruker D8 VENTURE diffractometer at 100 K using Mo-Kα radiation ($\lambda = 0.71073$ Å). Single-crystal structure data for $[\text{Gd(hfac)}_3(N,N'-pd)]$·Et2O (2a), $[\text{Gd(hfac)}_3(N,N'-O,O'-pd)]$·Tol (2b) and $[\text{Y(hfac)}_3(N,N'-O,O'-pd)]$·Tol (3) were collected by the EPSRC UK National Crystallography Service using a Rigaku AFC12 goniometer, mounted at the window of a FR-E+ SuperBright molybdenum rotating anode generator, and equipped with a (HG) Saturn214+ detector. CCDC numbers 1832429-1832434 and 1841119† contain the crystallographic information for this paper.

X-band EPR spectra were collected on a Bruker ELEXSYS E500 spectrometer and simulations were performed using Bruker’s Xsophe software package.47

**Synthesis of $[\text{Y(hfac)}_3(N,N'-pd)]$ (1):**

THF (15 mL) was added to a Schlenk flask charged with a dry mixture of $[\text{Y(Ot}f)_3]$ (1.16 g, 2.17 mmol), pd (0.456 g, 2.17 mmol) and K(hfac) (1.60 g, 6.51 mmol) at room temperature with stirring. The yellow suspension was stirred at room temperature for 1 h, whereupon the solution became deep green. The green solution was stirred for 16 h, THF was removed in vacuo and the green solids were extracted into toluene (3 × 5 mL), giving a yellow solution. Removal of toluene in vacuo gave green solids, which were washed with hexane (2 × 5 mL) to give $[\text{Y(hfac)}_3(N,N'-pd)]$ as a pale green solid (1.70 g, 1.84 mmol, 85% yield). Single crystals suitable for X-ray diffraction were grown from both saturated Et2O ($[\text{Gd(hfac)}_3(N,N'-pd)]$·Et2O, 1a) and saturated toluene solution at −35 °C overnight ($[\text{Gd(hfac)}_3(N,N'-pd)]$·Tol, 1b). [H] NMR (C6D6): δ 6.23 (3H, s, hfac-CH) 6.37 (2H, dd, $J_{HH} = 7.8, 5.5$ Hz, pd C5-H) 7.53 (2H, d, $J_{HH} = 8.7$ Hz, pd C4,7-H) 9.16 (2H, d, $J_{HH} = 5.2$ Hz, pd C2,5-H) ppm. 19F [1H] NMR (C6D6): δ −76.93 (s, hfac-CF3) ppm. Anal. calcd for C35H27F18N2O10Y·C7H8: C, 35.24%; H, 0.99%; N, 2.59%. Found: C, 35.85%; H, 1.47%; N, 2.54%. IR (ATR): 1697 (w, νC=O) 1654 (s) 1518 (m) 1372 (m) 1197 (s) 1136 (s) 794 (s) 656 (m).

**Synthesis of $[\text{Gd(hfac)}_3(N,N'-O,O'-pd)]$ (2):**

THF (15 mL) was added to a Schlenk flask charged with a dry mixture of Gd(OtF)3 (0.759 g, 1.26 mmol), K(hfac) (0.997 g, 3.77 mmol) and 1,10-phenanthroline-5,6-dione (0.263 g, 1.25 mmol). The yellow suspension was stirred for 20 h, turning green after ca. 1 h. THF was removed in vacuo and $[\text{Gd(hfac)}_3(N,N'-pd)]$ was extracted into toluene (3 × 10 mL) to give a yellow solution. Toluene was removed in vacuo to give a green solid, which was washed with hexanes to give $[\text{Gd(hfac)}_3(N,N'-pd)]$ as a pale green solid (0.992 g, 1.00 mmol, 79% yield). Single crystals suitable for X-ray diffraction were grown from both saturated Et2O ($[\text{Gd(hfac)}_3(N,N'-pd)]$·Et2O, 2a) and toluene ($[\text{Gd(hfac)}_3(N,N'-pd)]$·Tol, 2b) solutions cooled to −35 °C overnight. Anal. calcd for C37H19F18N2O8VY: C, 40.35%; H, 1.74%; N, 2.59%. Found: C, 37.58%; H, 1.47%; N, 2.54%. IR (ATR): 1697 (w, νC=O) 1654 (s) 1518 (m) 1372 (m) 1197 (s) 1136 (s) 794 (s) 656 (m).

**Synthesis of $[\text{Y(hfac)}_3(N,N'-O,O'-pd)]$ (3):**

Et2O (10 mL) was added to a −78 °C cooled Schlenk flask charged with a dry mixture of 1 (0.50 g, 0.54 mmol) and V(Cp)2 (0.10 g, 0.54 mmol), immediately giving a dark green solution. The reaction was allowed to warm to room temperature and stirred for 2 h. Et2O was removed in vacuo and the green/black product recrystallised from toluene/hexane at −15 °C to give $[\text{Y(hfac)}_3(N,N'-O,O'-pd)]$·VCp2 as a black microcrystalline solid (0.335 g, 0.30 mmol, 56% yield). Crystals suitable for X-ray diffraction were grown from a toluene/hexane mixture at −35 °C overnight. Anal. calcd for C39H29F18N2O10Y: C, 40.35%; H, 1.74%; N, 2.54%. Found: C, 39.88%; H, 1.81%; N, 2.56%. IR (ATR): 1651 (s) 1500 (m) 1371 (s) 1136 (s) 794 (s) 651 (s) 582 (m) cm−1. max (MeCN)/nm (ε/M−1 cm−1): 431 (5400), 600 (2630), 788 (1840).

**Synthesis of $[\text{Gd(hfac)}_3(N,N'-O,O'-pd)]$ (4):**

THF (15 mL) was added to a dry mixture of $[\text{Gd(hfac)}_3(N,N'-pd)]$ (500 mg, 0.505 mmol) and V(Cp)2 (91.3 mg, 0.504 mmol) and the dark suspension was stirred for 24 h. THF was removed in vacuo. $[\text{Gd(hfac)}_3(N,N'-O,O'-pd)]$·V(Cp)2 was extracted into toluene and cooled to −15 °C, whereupon it precipitated as a black powder (378 mg, 0.322 mmol, 64% yield). Anal. calcd for C45H23F18N2O10V: C, 40.94%; H, 1.64%; N, 3.29%. Found: C, 40.85%; H, 1.59%; N, 3.29%. IR (ATR): 1651 (m) 1504 (s) 1468 (w) 1070 (s) 1069 (m) 955 (s) 794 (s) 651 (m) 582 (m) cm−1. max (MeCN)/nm (ε/M−1 cm−1): 425 (6450), 576 (3430), 797 (2740).

**Synthesis of $[\text{Y(hfac)}_3(N,N'-O,O'-pd)]$ (5):**

Et2O (30 mL) was added to a −78 °C cooled Schlenk flask charged with V(Cp)2 (500 mg, 0.533 mmol) and $[\text{Y(hfac)}_3(N,N'-pd)]$ (770 mg, 0.83 mmol), giving a dark green suspension. The reaction was allowed to warm to room temperature and stirred for 16 hours. The dark suspension was filtered into a Schlenk and dried in vacuo to give deep-green solids, which were recrystallised from a saturated solution of toluene and hexane at −15 °C overnight. The micro-crystals were isolated, washed with hexanes (3 × 15 mL), and dried in vacuo to give $[\text{Y(hfac)}_3(N,N'-O,O'-pd)]$·V(Cp)2 as a dark green powder (520 mg, 70%, 56% yield).
10.43 mmol, 50% yield). Crystals suitable for X-ray diffraction were grown by diffusion of hexane vapour into a toluene solution at ~35 °C over several days. Anal. calcld for C14H18F3N2O2Vg: C, 44.54%; H, 2.91%; N, 2.31%. Found: C, 44.49%; H, 3.78%; N, 2.49%. IR(ATR): 1651 (s), 1506 (m), 1423 (m), 1370 (w), 1250 (s), 974 (m), 727 (w), 659 (m). λmax (THF)/nm (εM−1 cm−1): 418 (6970), 553 (2920), 816 (2340). μeff (298 K, d2e-THF) 1.72–1.75 BM.

Synthesis of \([\text{Gd(hfac)}_3(N,N′,O,O′-pd)]\) (6)

Et₂O (30 mL) was added to a ~78 °C cooled Schlenk flask charged with VCP₆H₂ (38 mg, 0.13 mmol) and \([\text{Gd(hfac)}_3(N,N′-pd)]\) (130 mg, 0.13 mmol), giving a dark green suspension. The reaction was allowed to warm to room temperature, and stirred for 16 hours. The dark suspension was filtered into a Schlenk and dried in vacuo to give deep-green solids, which were recrystalised from a saturated solution of toluene and hexane at ~15 °C overnight. The micro-crystals were isolated, washed with hexanes (3 × 15 mL), and dried in vacuo to give \([\text{Gd(hfac)}_3(N,N′-O,O′-pd)]\) as a dark green powder (70 mg, 5.40 mmol, 42% yield). Crystals suitable for X-ray diffraction were grown by diffusion of hexane vapour into an Et₂O solution at ~35 °C over several days. Anal. calcld for C₁₄H₁₈F₃N₂O₂Vg: C, 42.16%; H, 2.75%; N, 2.19%. Found: C, 42.45%; H, 3.18%; N, 2.04%. IR(ATR): 1650 (s), 1500 (m), 1370 (w), 1250 (s), 800 (m), 733 (w), 662 (m). μeff (298 K, d2e-THF) 7.99–8.41 BM.

Conflicts of interest

There are no conflicts to declare.

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