

Cite this: *Dalton Trans.*, 2018, **47**, 10692

The modular synthesis of rare earth-transition metal heterobimetallic complexes utilizing a redox-active ligand†

James R. Hickson,^{a,b} Samuel J. Horsewill,^a Christopher Bamforth,^a Jake McGuire,^{id}^a Claire Wilson,^{id}^a Stephen Sproules^{id}^a and Joy H. Farnaby^{id}^{*a}

We report a robust and modular synthetic route to heterometallic rare earth-transition metal complexes. We have used the redox-active bridging ligand 1,10-phenanthroline-5,6-dione (pd), which has selective *N,N'* or *O,O'* binding sites as the template for this synthetic route. The coordination complexes [Ln(hfac)₃(*N,N'*-pd)] (Ln = Y [**1**], Gd [**2**]; hfac = hexafluoroacetylacetonate) were synthesised in high yield. These complexes have been fully characterised using a range of spectroscopic techniques. Solid state molecular structures of **1** and **2** have been determined by X-ray crystallography and display different pd binding modes in coordinating and non-coordinating solvents. Complexes **1** and **2** are unusually highly coloured in coordinating solvents, for example the vis-NIR spectrum of **1** in acetonitrile displays an electronic transition centred at 587 nm with an extinction coefficient consistent with significant charge transfer. The reaction between **1** and **2** and VCp₂ or VCp[†]₂ (Cp[†] = tetramethylcyclopentadienyl) resulted in the isolation of the heterobimetallic complexes, [Ln(hfac)₃(*N,N'*-*O,O'*-pd)VCp₂] (Ln = Y [**3**], Gd [**4**]) or [Ln(hfac)₃(*N,N'*-*O,O'*-pd)VCp[†]₂] (Ln = Y [**5**], Gd [**6**]). The solid state molecular structures of **3**, **5** and **6** have been determined by X-ray crystallography. The spectroscopic data on **3–6** are consistent with oxidation of V(II) to V(IV) and reduction of pd to pd²⁻ in the heterobimetallic complexes. The spin-Hamiltonian parameters from low temperature X-band EPR spectroscopy of **3** and **5** describe a ²A₁ ground state, with a V(IV) centre. DFT calculations on **3** are in good agreement with experimental data and confirm the SOMO as the d_{x²-y²} orbital localised on vanadium.

Received 31st March 2018,
Accepted 2nd June 2018

DOI: 10.1039/c8dt01262h

rsc.li/dalton

Introduction

The cooperativity between metal centres in multimetallic complexes, clusters and polymeric species gives rise to very remarkable chemistry in a diverse range of fields.¹ The potential applications of complexes containing both transition metals and lanthanides are significant. In particular there is research interest in the design of rare earth-transition metal single-molecule magnets (SMM) and catalysts.^{2,3}

The chemistry of multimetallic complexes containing f-elements is significantly less well developed than that of transition metals. This is attributable to the fact that the directed synthesis of f-d complexes is non-trivial. The inspiration for our approach was the Tb analogue of the [(SiMe₃)₂N]Ln₂N₂K system. This complex was remarkable both for the (N₂)³⁺⁻

anion and the SMM behaviour.⁴ It generated significant subsequent interest in multimetallic lanthanide radical bridged complexes.⁵ The very recent reports of magnetic hysteresis at 60 K in a dysprosocenium cation are a testament to the power of the organometallic molecular design approach.^{6,7}

Our aim was to identify and use a redox-active bridging ligand with multiple and selective binding sites as the template for a general and modular synthetic route to heterometallic d-f or f-f' complexes. The ligand used in this work 1,10-phenanthroline-5,6-dione (pd) has been predominantly studied for its physical organic, biological and bio-inorganic chemistry, particularly as a precursor to DNA intercalation agents and photocatalysts.^{8,9} Metal complexes of pd have been targeted for applications as diverse as electrocatalysis, sensitised Ln emission and thermochromism.^{10–12}

There is transition metal literature precedent for using pd as a bridging ligand to synthesise multi-metallic complexes. The first transition metal complexes of pd were synthesised by Balch by the addition of Pt(PPh₃)₄ to pd to form [(*O,O'*-pd)Pt(PPh₃)₂].¹³ The electrochemical and electrocatalytic properties of these complexes were subsequently investigated by Abruna.¹⁰ Pierpont demonstrated that [(*O,O'*-pd)Pt(PPh₃)₂]

^aSchool of Chemistry, WestCHEM, University of Glasgow, Glasgow, UK.

E-mail: Joy.Farnaby@glasgow.ac.uk

^bDepartment of Chemistry, Imperial College London, South Kensington, London, UK

† Electronic supplementary information (ESI) available. CCDC 1832429–1832434 and 1841191. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8dt01262h



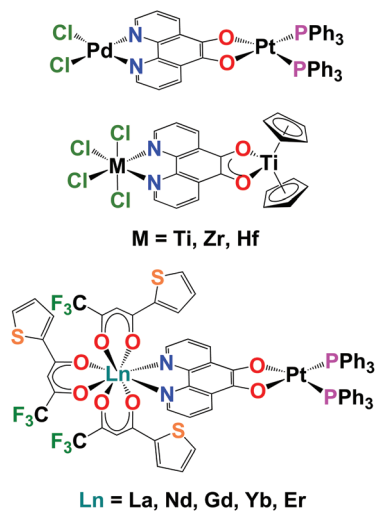


Chart 1 Bimetallic complexes of pd.

could be used as a bipyridine equivalent to bind a second metal.¹⁴ This paper also reported the selective N,N' -pd binding of PdCl_2 and the first crystal structures of transition metal pd complexes (Chart 1). Eisenberg combined pd and bipyridine (bipy) to form the homobimetallic Pt complexes, $[\text{PtCl}_2(N,N'-O,O'-\text{pd})\text{Pt}(\text{di-}t\text{-tert-butyl-bipy})]$. These were then used as precursors to synthesise $[\text{Pt}(\text{L})(N,N'-O,O'-\text{pd})\text{Pt}(\text{di-}t\text{-tert-butyl-bipy})]$ (where L = dithiolate, dicatocateholate or pd) and probe electronic structure.¹⁵

Directly relevant to this work, is the first report of the coordination and redox chemistry of group 4 and 5 metals with pd, by Pampaloni and Calderazzo.¹⁶ These groups subsequently published a number of studies demonstrating that pd will selectively bind N,N' with a coordination chemistry precursor or O,O' with a reducing organometallic precursor and can form homo- and heterometallic transition metal complexes.^{17–19} Prior to this work there are just two reports of structurally characterised molecular lanthanide complexes of pd, both by Faulkner and Ward, which describe the sensitised NIR emission of $[\text{Ln}(\text{tta})_3(N,N'-O,O'-\text{pd})\text{Pt}(\text{PPh}_3)_2]$ (Ln = La, Nd, Gd, Er, Yb; tta = 1-thenoyl-5,5,5-trifluoroacetylacetonate).^{12,20} We note that there are reports of the biological activity of combinations of Ln and pd.

Here we report a robust and modular synthesis and first structural characterisation of $[\text{Ln}(\text{hfac})_3(N,N'-\text{pd})]$ (Ln = Y, Gd)

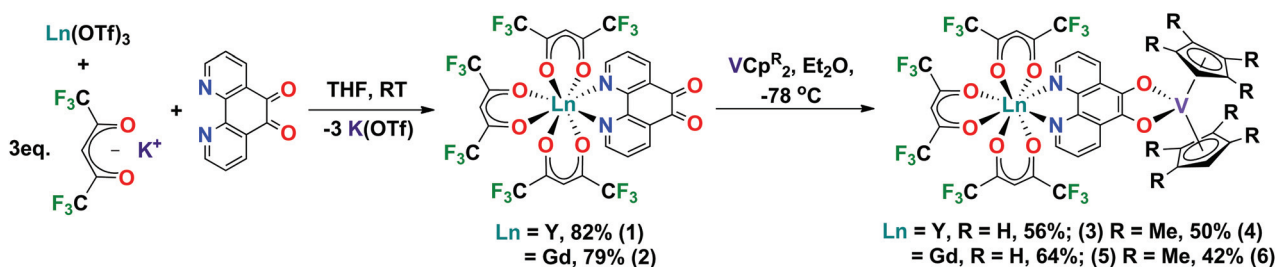
complexes. We also report the reduction of these coordination chemistry complexes using vanadocenes to form hetero-bimetallic rare earth-transition metal complexes. The $[\text{Ln}(\text{hfac})_3(N,N'-O,O'-\text{pd})\text{VCp}^{\text{R}}_2]$ complexes have been fully characterised by a range of spectroscopic techniques. These complexes are spectroscopically rich and represent the first building-blocks in our studies of rare earth-transition metal complexes.

Results and discussion

Synthesis of $[\text{Ln}(\text{hfac})_3(N,N'-\text{pd})]$

Metal precursor choice is particularly important in f-element chemistry. In this work, we chose to use the $\text{Ln}(\text{OTf})_3$ (Ln = Y, Gd, $\text{OTf} = \text{CF}_3\text{SO}_3$) more commonly applied in organic synthesis as Lewis acid catalysts.²¹ This was due to the ease of preparation from inexpensive Ln_2O_3 starting materials, their stability in H_2O and facile dehydration by comparison to commercially available $\text{LnCl}_3 \cdot (\text{H}_2\text{O})_x$. The hfac (hfac = 1,1,1-5,5,5-hexafluoroacetylacetonate) ligand was chosen because $[\text{Ln}(\text{hfac})_3(\text{glyme})]$ are well-defined molecular species.^{22,23} The hfac ancillary ligand environment has good solubility and stability, with ^{19}F NMR as an additional probe of solution behaviour. We note that while you can access $[\text{Ln}(\text{hfac})_3(\text{glyme})]$ directly from Ln_2O_3 , the synthesis requires the addition of chelating polyethers to prevent aggregation.²⁴

However, using a salt elimination reaction enables this synthetic route to be modular and applicable to other ligand environments. The coordination complexes $[\text{Ln}(\text{hfac})_3(N,N'-\text{pd})]$ were synthesised in high yield. $\text{Ln}(\text{OTf})_3$ (Ln = Y, Gd), pd and 3 equivalents of $\text{K}(\text{hfac})$ were combined dry, and THF added at room temperature with stirring, giving a colour change from yellow to deep green. The reaction was stirred at room temperature for 16 h, and THF removed *in vacuo*. Extraction into toluene, filtration to exclude $\text{K}(\text{OTf})$ and removal of toluene *in vacuo* gave the neutral complexes $[\text{Ln}(\text{hfac})_3(N,N'-\text{pd})]$ as pastel green solids in excellent yields (Ln = Y: 82% [1], Gd: 79% [2]) (Scheme 1). Elemental analyses of 1 and 2 crystallised from toluene at -35°C support the $[\text{Ln}(\text{hfac})_3(N,N'-\text{pd})]$ formulation, with one toluene of crystallisation associated. 1 and 2 are soluble in both coordinating and non-coordinating solvents, to give green and yellow solutions, respectively.



Scheme 1 The modular synthesis of rare earth-transition metal complexes.



Spectroscopy of $[\text{Ln}(\text{hfac})_3(\text{N},\text{N}'\text{-pd})]$

The room temperature ^1H NMR spectrum of **1** in d_6 -benzene is consistent with complexation to form the 1 : 1 adduct. The pd resonances in **1** at $\delta = 6.37$, 7.53, and 9.16 ppm are shifted downfield relative to the pd free ligand and the aromatic splitting pattern (dd) is no longer observed for the protons neighbouring the N,N' -pocket. The expected ratio of 2 : 2 : 2 : 3 is observed between the pd resonances and the hfac singlet at $\delta = 6.23$ ppm. The ^{19}F NMR spectrum displays a singlet at $\delta = -76.9$ ppm. The X-band EPR spectrum of **2** at room temperature (see ESI†) is common for Gd(III) and simulation gave $g_{\text{iso}} = 1.989$, close to that of the free electron, which is typical for Gd(III).²⁵ The $^8\text{S}_0$ ground state of Gd(III) has no angular orbital momentum, and so unlike its f-block counterparts, Gd(III) gives an EPR response at room temperature. The Evans Method was used to determine the magnetic moment of **2** in d_8 -thf at room temperature. The corrected magnetic moments are in the range of 7.56–7.71 μ_{B} for **2**, consistent with Gd(III).

IR spectroscopy cannot definitively assign the coordination mode of the pd ligand but does provide a useful measure of its carbon–oxygen bond order. The reported ranges of ν_{CO} are: neutral pd = 1678 cm^{-1} , radical anion $\text{pd}^{\cdot-}$ 1500–1400 cm^{-1} and dianion pd^{2-} 1400–1300 cm^{-1} .^{14,16} The IR spectra of **1** and **2** have several absorptions in the ν_{CO} region, complex **1** at 1699, 1668 and 1647 cm^{-1} and **2** at 1697, 1661, 1647 cm^{-1} . The ν_{CO} at 1647 is assigned to the hfac ligand based on the IR of closely related complexes, $[\text{Ln}(\text{hfac})_3(\text{glyme})]$, which have very strong absorption between 1650 and 1660 cm^{-1} .²³ The ν_{CO} at 1699 and 1697 cm^{-1} are assigned as $\text{N},\text{N}'\text{-pd}$. This is consistent with complexes formulated as, $[\text{PdCl}_2(\text{N},\text{N}'\text{-pd})]$,¹³ $[\text{Fe}(\text{N},\text{N}'\text{-pd})_3][\text{PF}_6]_2$,¹⁰ and $[\text{MCl}_4(\text{N},\text{N}'\text{-pd})]$ ($\text{M} = \text{Ti}, \text{Zr}, \text{Hf}$)¹⁶ which have strong absorptions in the range 1680–1705 cm^{-1} . We propose additionally that the ν_{CO} at 1668 and 1661 cm^{-1} may result from the short contacts seen between the Ln metal centre and one pd-O atom of the neighbouring molecule (Fig. 3 and ESI†).

The deep green colour of **1** and **2** in coordinating solvents is an unexpected property as neither $4d^0$ Y(III) nor $4f^7$ Gd(III) often give rise to strongly coloured complexes. It is of note that related Ln complexes with 1,10-phenanthroline (phen) or neutral 2,2-bipyridyl (bipy) ligands are not strongly coloured.^{24,26,27} The electronic spectra of **1** and **2** in acetonitrile display peaks centred at 587 nm ($\epsilon = 2390 \text{ M}^{-1} \text{ cm}^{-1}$) and 597 ($\epsilon = 129 \text{ M}^{-1} \text{ cm}^{-1}$), respectively (see ESI† for spectra). We assign this as ligand to metal charge transfer.

Crystallography of $[\text{Ln}(\text{hfac})_3(\text{N},\text{N}'\text{-pd})]$

Crystals of **1** suitable for X-ray diffraction were grown from either a saturated Et_2O solution at -35 °C over 7 days (**1a**) or a saturated toluene solution at -15 °C overnight (**1b**). Similarly crystals of **2** suitable for X-Ray diffraction were grown from either a saturated Et_2O solution at -35 °C overnight (**2a**) or a saturated toluene solution at -15 °C overnight (**2b**). The solid state molecular structures and selected parameters are shown in Fig. 1–3 or in the ESI.†

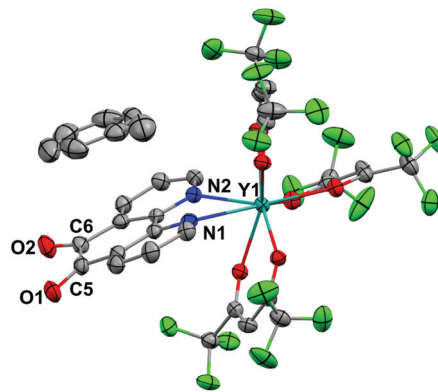


Fig. 1 Molecular structure of **1b**. Hydrogen atoms omitted for clarity. Thermal ellipsoids drawn at 50% probability. Selected distances (Å): Y1–N1 2.501(3), Y1–N2 2.552(3), C5–C6 1.524(6), C6–O2 1.216(5), C5–O1 1.213(5).

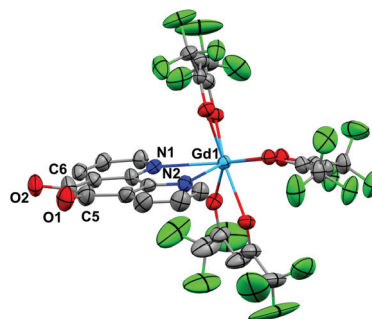


Fig. 2 Molecular structure of **2a**. Hydrogen atoms and Et_2O of crystallisation omitted for clarity. Thermal ellipsoids drawn at 50% probability. Selected distances (Å): Gd1–N1 2.585(8), Gd1–N2 2.555(9), C5–C6 1.54(6), C6–O2 1.23(2), C5–O1 1.24(2).

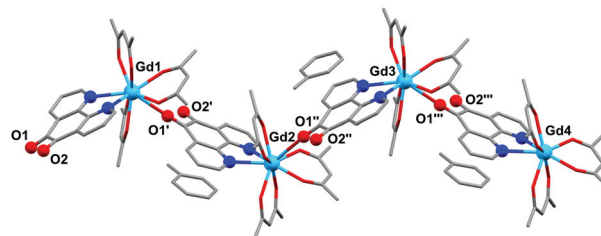


Fig. 3 Molecular structure of **2b**, showing the polymeric structure. All atoms depicted as wireframe or ball and stick.

All structures adopt a distorted square antiprismatic geometry, and confirm $\text{N},\text{N}'\text{-pd}$ coordination. As expected the pd C–O distances in **1a–2b** are unchanged from the free ligand (1.213(5) and 1.216(5) Å).²⁸ The Ln–O(hfac) distances to the hfac ligands are consistent between both structures, and are in agreement with the crystal structures of other examples of $\text{Ln}(\text{hfac})_3$ complexes in the range of 2.299(2)–2.343(3) Å.^{23,29} The Ln–N(pd) distances in structures **1a–2b** of 2.501(3)–2.585(8) Å are directly comparable to those reported in $[\text{Ln}(\text{tta})_3]$



(*N,N'*-*O,O'*-pd)Pt(PPh₃)₂] (Ln = Er, Gd; tta = 1-thenoyl-5,5,5-trifluoroacetylacetonate) of 2.499(7)–2.539(7) Å,¹² as the ionic radius of Er(III) is an appropriate size match for Y(III).³⁰ Complex **1** is the first crystallographically characterised example of an yttrium pd complex. The Y–N(pd) distances are also comparable to those seen in [Y(pyrMe₂-acac)₃(phen)] (pyrMe₂-acac = 1,5-(1,3-dimethyl-1*H*-pyrazol-4-yl)acetylacetonate) of 2.562(4)–2.579(4) Å.²⁴

The structure of **2b** from toluene is a coordination polymer as a result of *N,N'*-*O*-pd binding to the next Gd in the chain (Fig. 3). The only other structure in the CCDC that displays this binding mode is the Ag(I) coordination polymer [Ag(ClO₄)(*N,N'*-*O*-pd)]_n.³¹ The Gd1–O2 distance is 2.830(2) Å, and the coordination polymer exhibits a zig-zag structure with an angle of 77.56° between Gd–pd planes. Analysis of short contacts in **1b** reveals a Y1–O2 interaction within the sum of the van der Waals radii at 3.154 Å, giving a pseudo-coordination polymer structure with an angle of 83.98° between Y–pd planes (see ESI†). As a result of the *N,N'*-*O*-pd binding, both **2b** and **1b** feature significant distortion within the pd ring. The three rings of pd are no longer coplanar and there is an out-of-plane O2-ring distortion and associated out-of-plane displacement of O2.

Reduction chemistry of **1** and **2**

The pd ligand can be sequentially reduced by two electrons, first to the semiquinone radical anion pd^{•−} and then to the fully reduced and diamagnetic catecholate pd^{2−}; the reduction potentials are −0.85 and −1.71 V vs. Fc⁺/Fc, respectively.^{10,32} Almost all published metal complexes of the reduced pd ligand contain pd^{2−} and the few examples of the radical anion of pd^{•−} were generated *in situ*, usually under electrochemical conditions and characterised by electronic spectroscopies such as EPR and UV-vis-NIR. This is in spite of the ease of reduction of pd to pd^{•−} relative to the structurally related 1,10-phenanthroline¹⁰ and 2,2'-bipyridine.³³

Complex **1** is unstable under standard electrochemical conditions. Therefore, we used the reduction potentials of the free pd ligand to rationalise the choice of organometallic reducing agent.³⁴ The first reversible oxidation process of V(II) to V(III) in VCp₂ to VCp₂⁺ is a good match at −0.93 V vs. Fc⁺/Fc.³⁵ We note that VCp₂ has been previously shown to reduce pd, with the product formulated as [(*O,O'*-pd)VCp₂] and in our hands this reaction formed intractable dark green solids.¹⁶ Additionally, we used a substituted vanadocene, VCp^t₂ (Cp^t = tetramethylcyclopentadienyl) to test whether the added steric bulk would promote the formation of an ion pair. The substituted Cp^t ligand environment is also a good model for extending this synthetic route to incorporate f-element organometallics.³⁶

Synthesis of [Ln(hfac)₃(*N,N'*-*O,O'*-pd)VCp^R₂]

A Schlenk flask charged with **1** or **2** and VCp₂ was cooled to −78 °C and Et₂O was added with stirring. The colour of the resultant solution was dark green-black. The reaction mixture was warmed to room temperature and Et₂O was removed *in vacuo*. Recrystallisation of the solids from a mixture of

toluene and hexane at −15 °C gave [Ln(hfac)₃(*N,N'*-*O,O'*-pd)VCp₂] as black microcrystalline solids (Ln = Y: 56% [**3**], Ln = Gd: 64% [**4**]). The analogous reaction between **1** or **2** and VCp^t₂ (t = Me₄), resulted in the isolation of dark green-black microcrystalline solids of [Ln(hfac)₃(*N,N'*-*O,O'*-pd)VCp^t₂] (Ln = Y: 50%, [**5**] Ln = Gd: 42% [**6**]) (Scheme 1). The composition and purity of complexes **4**–**6** was confirmed by elemental analysis.

Spectroscopy of [Ln(hfac)₃(*N,N'*-*O,O'*-pd)VCp^R₂]

Complexes **3**–**6** are paramagnetic and NMR silent. The Evans Method was used to determine the magnetic moment of **3**, **5** and **6**, in *d*₈-thf at room temperature. The corrected magnetic moments are in the range of 1.71–1.74 μ_B for **3**, 1.72–1.75 μ_B for **5** and 7.99–8.41 μ_B for **6**. The data for **3** and **5** are consistent with a single unpaired electron *i.e.* 3d¹ V(IV), resulting from the oxidative addition of V(II) to V(IV) and the concomitant reduction of pd by two electrons to pd^{2−}. The data for **6** are consistent with the value predicted by the spin only formula for Gd(III) and V(IV). This is chemical behaviour similar to the oxidative addition of VCp₂ to acetylene to form the metallocyclopropane.³⁷

The IR spectra of complexes **3**–**6**, are consistent with pd reduction. The ν_{CO} at 1699 and 1697 cm^{−1} assigned to neutral *N,N'*-pd and at 1668 and 1661 cm^{−1} assigned to *N,N'*-*O*-pd are no longer visible. Instead, complexes **3**–**6**, display ν_{CO} in the range 1370–1381 cm^{−1}, consistent with the reported examples of pd^{2−} complexes.^{14,16}

An overlay of the electronic spectra of **3** and **4** collected in MeCN are contrasted with the electronic spectra of **5** and **6** collected in THF and presented in Fig. 4. The d–d transitions at

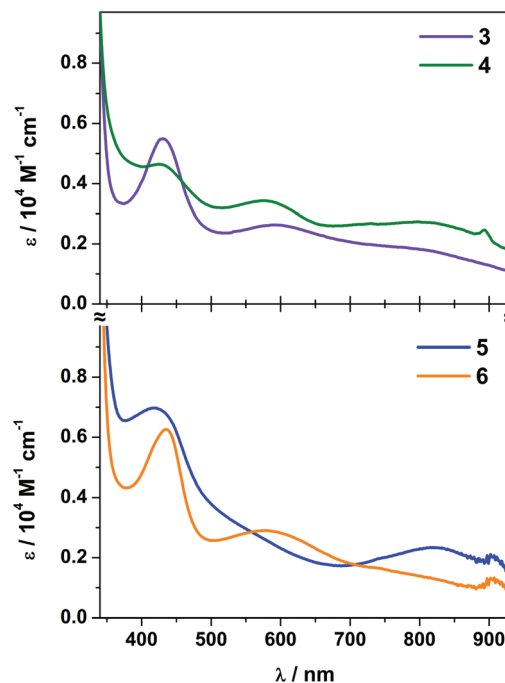


Fig. 4 Overlay of the electronic spectra of **3** and **4** in MeCN (top) and **5** and **6** in THF (bottom) recorded at room temperature.



740 and 640 nm usually present in pseudo-tetrahedral vanadocenes^{38,39} are not directly observed as they are hidden underneath the charge transfer transitions. The transitions are assigned with the aid of computational analysis (*vide infra*).

Crystallography of [Ln(hfac)₃(N,N'-O,O'-pd)VCp^R₂]

Crystals of **3**, **5** and **6** suitable for X-ray diffraction were grown by diffusion of hexane vapour into a toluene or Et₂O solution at -35 °C or by cooling a toluene/hexane mixture to -35 °C over several days. The structures and selected structural parameters are shown in Fig. 5 and 6 or in the ESI† Complexes **3**, **5** and **6** feature yttrium or gadolinium in a distorted square antiprismatic geometry and N,N'-pd bound analogous to **1a** and **2a**. The VCp^R₂ units are O,O'-pd bound, resulting in a pseudo-tetrahedral geometry typical for V(IV) metallocenes.^{40,41} The metal to ancillary ligand bonding metrics are normal. The five membered ring formed between the V and the O,O'-pocket of the pd

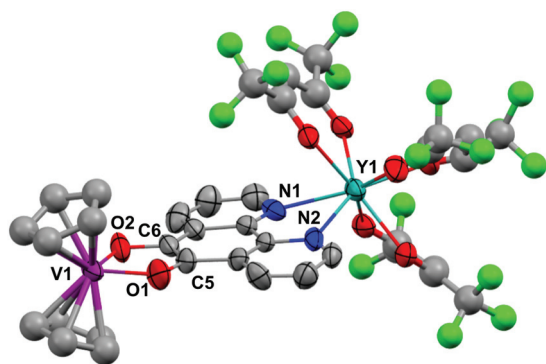


Fig. 5 Molecular structure of **3**. Hydrogen atoms and solvent of crystallisation omitted for clarity. Thermal ellipsoids drawn at 50% probability for pd, V and Y. Other ancillary ligand atoms depicted as ball and stick due to CF₃ and Cp disorder (see cif). Selected distances (Å): Y1–N1 2.474(17), Y1–N2 2.482(17), C5–C6 1.396(15), C6–O2 1.24(2), C5–O1 1.40(2), V1–O1 1.972(16), V1–O2 1.971(16).

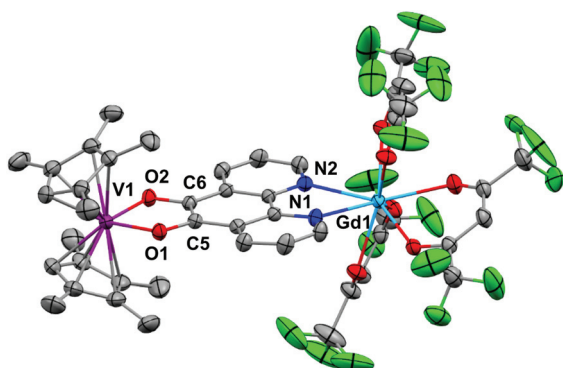


Fig. 6 Molecular structure of **6**. Hydrogen atoms and solvent of crystallisation omitted for clarity. Thermal ellipsoids drawn at 50% probability. Selected distances (Å) for **5**: Y1–N1 2.48(1), Y1–N2 2.49(1), C5–C6 1.38(2), C6–O2 1.35(1), C5–O1 1.34(2), V1–O1 1.975(8), V1–O2 1.977(9) and **6**: Gd1–N1 2.532(2), Gd1–N2 2.527(2), C5–C6 1.387(3), C6–O2 1.333(3), C5–O1 1.333(3), V1–O1 1.9869(18), V1–O2 1.9862(18).

is a common structural motif. The V1–O1 and V1–O2 distances in **3**, **5** and **6** are consistent with those observed in related complexes [Na₂(Et₂O)₂][V{dippCO₂}₃], [VCp₂(acac)][OTf] and V(IV) catecholates such as [NH₄Et₃][V(cat)₃] and [H₂PrDA][VO(cat)₂] (cat = 1,2-dihydroxybenzene, PrDA = 1,3-diaminopropane).^{42–45}

The metrics that are important in assessing the reduction of the pd ligand are C5–C6, C5–O1 and C6–O2. The C5–C6 bond distances in **3** of 1.40(3) Å, in **5** of 1.38(2) Å and in **6** of 1.387(3) Å are significantly shortened from that of **1b** (1.524(6) Å) consistent with double bond formation. The C5–O1 and C6–O2 distances in **3** of 1.40(2) and 1.24(2) Å, in **5** of 1.34(2) and 1.35(2) Å and in **6** of 1.333(3) Å are significantly elongated from **1b** (1.213(5) and 1.216(5) Å). The average C–O distances in **3**, **5** and **6** are consistent with reduction to single C–O bonds. We note the disparity in the C–O distances in **3** but do not think it is chemically significant. These metrics in **3**, **5** and **6** are consistent with pd²⁻ and are essentially identical to those found in [Ln(tta)₃(N,N'-O,O'-pd)Pt(PPh₃)₂] (Ln = Gd, Er).¹²

EPR of [Ln(hfac)₃(N,N'-O,O'-pd)VCp^R₂]

The room temperature X-band EPR spectra of **3** and **5** shown in Fig. 7 are typical for V(IV) with an 8-line pattern from coupling of the electron spin ($S = 1/2$) to the $I = 7/2$ nuclear spin of the ⁵¹V (99.8% abundance). The spectral simulation of **3** was given $g_{\text{iso}} = 1.9799$ and $A_{\text{iso}} = 65.8 \times 10^{-4} \text{ cm}^{-1}$, which are synonymous with non-oxo vanadyl species.^{40–42,44} In contrast, the spectrum of **5** exhibits the two overlapping 8-line signals; simulations yielded an equal mix of one entity with $g_{\text{iso}} = 1.9819$, $A_{\text{iso}} = 68.4 \times 10^{-4} \text{ cm}^{-1}$, and the other $g_{\text{iso}} = 1.9779$, $A_{\text{iso}} = 64.9 \times 10^{-4} \text{ cm}^{-1}$ (Table 1).

The corresponding frozen solution spectra recorded in toluene at 130 K clearly depicts a signal indicative of a single chemical species of **3** and **5** (Fig. 8). The simulation parameters are nearly identical, with $g = (1.9657, 1.9776, 1.9978)$ and $A = (106.6, 78.1, 11.0) \times 10^{-4} \text{ cm}^{-1}$ for **3**, and $g = (1.9654,$

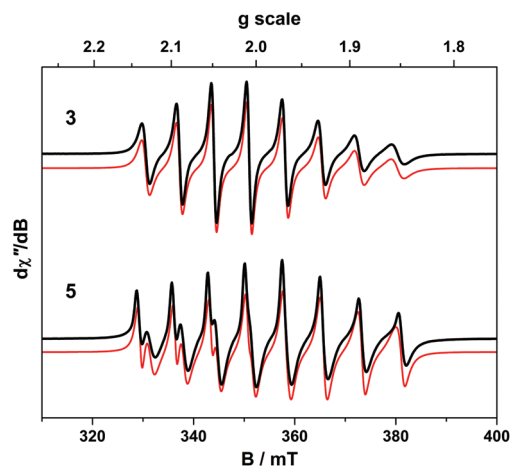


Fig. 7 X-band EPR spectra of **3** and **5** recorded in toluene solution at 293 K (experimental conditions: frequency, 9.8634 GHz; power, 0.63 mW; modulation, 0.2 mT). Experimental data are represented by the black line; simulation is depicted by the red trace.



Table 1 Spin-Hamiltonian g - and A -values^a derived from spectral simulation of complexes

Complex	Spin centre	g_{iso}	A_{iso}
3	V ($S = 1/2$)	1.9799	65.8
5	V ^b ($S = 1/2$)	1.9819	68.4
	V ^c ($S = 1/2$)	1.9790	64.9
2	Gd ($S = 7/2$)	1.989	—
4	Gd ($S = 7/2$)	1.991	—
4	V ^b ($S = 1/2$)	1.9779	72.0
	V ^c ($S = 1/2$)	1.9950	72.0

^a In 10^{-4} cm^{-1} . ^b Major conformer. ^c Minor conformer.

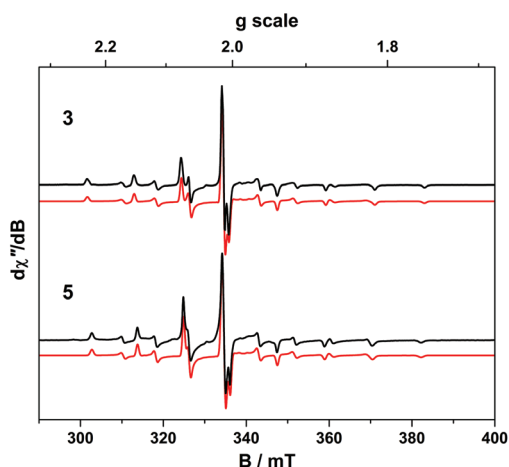


Fig. 8 X-band EPR spectra of **3** and **5** recorded in toluene solution at 130 K (experimental conditions: frequency, 9.4244 GHz; power, 0.2 mW; modulation, 0.1 mT). Experimental data are represented by the black line; simulation is depicted by the red trace.

1.9794, 1.9978) and $A = (104.1, 78.0, 9.5) \times 10^{-4} \text{ cm}^{-1}$ for **5**. The existence of one signal in the frozen matrix suggests two conformations of **5** are present in fluid solution with subtle differences in their spin-Hamiltonian parameters. The spectral profile for **3** and **5** describes a 2A_1 ground state ($d_{x^2-y^2}$ orbital in C_{2v} symmetry) with archetypal $g_x \approx g_y < g_z < g_e$ and $A_x \approx A_y > A_z$ splitting patterns.³⁸ The narrow lines neatly show a slight rhombicity in the spectrum with a notable inequivalence of g_x and g_y , and A_x and A_y in this ligand field. This is in contrast to three-fold symmetric tris(chelate)vanadium complexes, which have the same 2A_1 ground state but overall three-fold symmetry, such that $g_x = g_y$, and $A_x = A_y$.^{38,41,42}

The X-band EPR spectrum of **4** is shown in Fig. 9 and displays a superposition of the single broad signal for Gd(III) entwined with the 8-line pattern of V(IV). The spectrum was simulated with three subspectra, two of which arise from different conformations of the complex which are distinguished by the g -values of the V(IV) centre because of the narrow line-width as seen for **5**. In **4** the major (10.5%) and minor (1.8%) conformer sum to a seventh of proportion of the intensity as expected for a heterobimetallic species with one Gd(III) $S = 7/2$

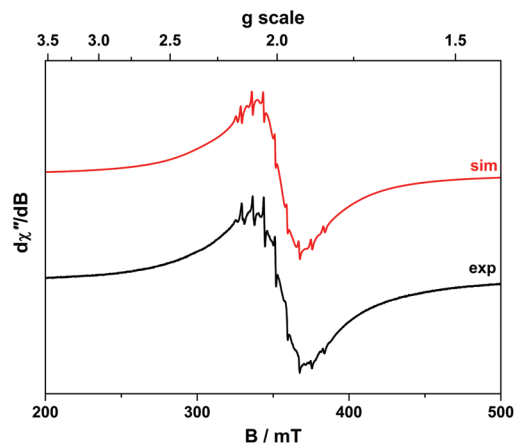


Fig. 9 X-band EPR spectrum of **4** recorded in toluene solution at 293 K (experimental conditions: frequency, 9.8717 GHz; power, 0.63 mW; modulation, 0.3 mT). Experimental data are represented by the black line; simulation is depicted by the red trace.

ion and one V(IV) $S = 1/2$ ion. In addition we observe no manifestation of exchange coupling between the Gd(III) and V(IV) centres such that $J \approx 0$ (*i.e.* two uncoupled spins).

DFT calculations on **3**

Density functional theoretical (DFT) calculations were carried out on **3** ($S = 1/2$). The optimised structure is in excellent agreement with the crystallographic structures of both **3** and **5** (Table S2†). The computed C–C and average C–O bond distances of 1.391 and 1.336 Å, respectively, are representative of a bridging pd ligand in its dianionic (catecholate) form. The highest occupied molecule orbital (HOMO) is singly-occupied (SOMO) with 91% vanadium d character that depicts the $d_{x^2-y^2}$ orbital which is a_1 in C_{2v} symmetry as shown by EPR (*vide supra*). The Mulliken spin population analysis shows +1.35 spins (α -spin) located on the V(IV) centre with matching β -spin on the first coordination sphere atoms (Fig. 10). This is the result of bond polarisation inherent to the relatively Lewis acidic V(IV) ion.⁴¹

Time-dependent (TD) DFT calculations were carried out on **3** in order to assign the electronic spectra of this series of heterobimetallic compounds. Each complex (**3–6**) displays two

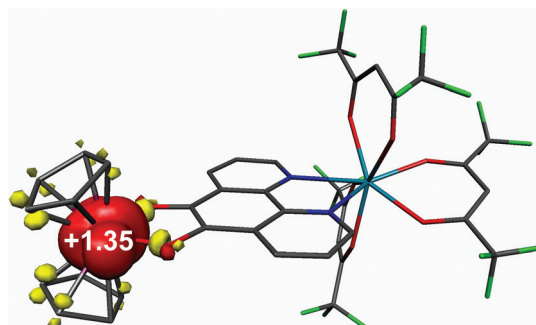


Fig. 10 Mulliken spin population analysis for **3**.



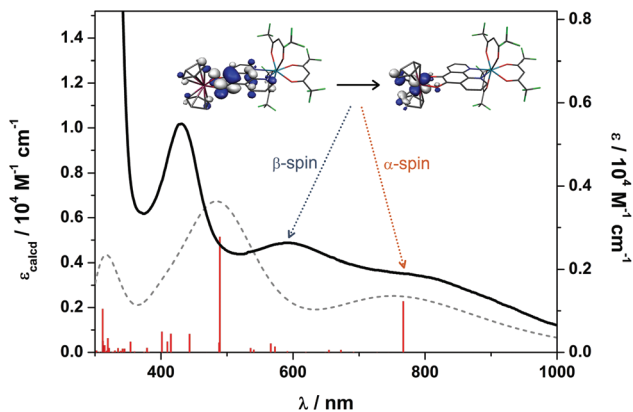


Fig. 11 Overlay of the experimental (solid line) and calculated (dashed line) electronic absorption spectra for **3**. The vertical bars show the individual calculated transitions. Inset orbitals depict the HOMO–1 → LUMO transition that constitutes the two lowest energy bands designated the α -spin and β -spin excitations, respectively.

low-energy bands at *ca.* 800 and 600 nm (Fig. 4). These signature spectral features were well-reproduced in the TD-DFT analysis (Fig. 11). Both bands are ligand-to-metal charge transfer (LMCT) transitions between the same donor and acceptor orbitals. These are the b_2 symmetric HOMO–1, which is localised on the O,O' end of the pd bridging ligand, and the LUMO, which is the vanadium d_{xy} orbital (Fig. 11). The computed spectrum reveals the spin-up (α -spin) excitation at 767 nm which perfectly matches the experimental peak at 800 nm in **3**. The corresponding spin-down (β -spin) excitation appears at 489 nm which is *ca.* 100 nm blue shifted from the experimental peak. This aforementioned polarisation inherent to Lewis acidic V(IV) sees the α -spin manifold is exchange stabilised relative to the β -spin manifold by virtue of the unpaired electron in the $d_{x^2-y^2}$ orbital. This is the same phenomenon that results in >1 spins located at the V(IV) centre in the spin population analysis (Fig. 10). The experimental polarisation is 0.54 eV is overestimated by the calculations at 0.92 eV. This is typical for DFT as energies of virtual (unoccupied) orbitals are less reliable especially given the TD-DFT computed spectrum is derived from energy differences between donor and acceptor orbitals based on the ground state calculation. This neglects additional factors that can (de) stabilise excited states, and therein fail to accurately reproduce electronic spectra of paramagnetic compounds. Nevertheless, this approximation is sufficient to assign the observed spectral features. The prominent band at 410 nm is tentatively assigned as a $\pi \rightarrow \pi^*$ transition from the pd^{2-} ligand.

Conclusions

We have demonstrated the synthesis and first structural characterisation of $[\text{Ln}(\text{hfac})_3(\text{N},\text{N}'\text{-pd})]$ for both Y and Gd complexes. The synthetic route works in exactly the same way for both diamagnetic Y and paramagnetic Gd. Considering the

size of Gd and Y, this synthetic route would be expected to be directly applicable to all the Ln elements from Sm–Lu. As the route is modular, a range of ancillary ligand environments would also be expected to be synthetically accessible. We have also shown the reduction of these coordination chemistry complexes using vanadocenes to form heterobimetallic rare earth-transition metal complexes. The spectroscopic data for $[\text{Ln}(\text{hfac})_3(\text{N},\text{N}'\text{-O},\text{O}'\text{-pd})\text{VCp}^{\text{R}}_2]$ are consistent with oxidation of V(II) to V(IV) and reduction of pd to pd^{2-} in these complexes. This redox chemistry is driven by the O,O' -pd binding of the VCp^{R}_2 unit. Work is ongoing to isolate rare-earth transition metal complexes containing the radical anion of pd bridging the metal centres and to synthesise heterobimetallic f–f' complexes using 4f and 5f elements.

Experimental

General details

All air-sensitive manipulations were carried out in an MBraun glovebox (O_2 and $\text{H}_2\text{O} < 1$ ppm) or by using standard Schlenk techniques under N_2 . All glassware was dried at 130 °C overnight prior to use. Filter cannulas were prepared using Whatman 25 mm glass microfiber filters and were pre-dried at 130 °C overnight. Dry solvents were obtained using an Innovative Technology Inc. Pure Solv 400-5-MD solvent purification system (activated alumina columns). Solvents were sparged with N_2 and stored in ampoules over activated molecular sieves under N_2 . Deuterated benzene and THF were dried by refluxing over K. Deuterated acetonitrile was dried by refluxing over CaH_2 . Dry deuterated solvents were degassed by three freeze–thaw cycles, vacuum distilled, and kept in ampoules in the glovebox under N_2 . The following starting materials were prepared according to literature procedures: VCp_2 , VCp^{t}_2 .⁴⁶ Hexafluoroacetylacetonone (Hhfac) was purchased from Aldrich and degassed by three freeze–thaw cycles before use. Potassium hydride in mineral oil was purchased from Aldrich, the mineral oil removed by washing with anhydrous hexanes, and stored under N_2 . Ln_2O_3 (Ln = Y, Gd), trifluoromethanesulfonic acid and 1,10-phenanthroline were purchased from Aldrich and used without further purification.

Physical methods

^1H NMR data were recorded on an AVIII 400 MHz instrument and were referenced internally to the appropriate residual protio-solvent and are reported relative to tetramethylsilane ($\delta = 0$ ppm). ^{19}F and $^{19}\text{F}\{^1\text{H}\}$ NMR spectra were recorded on a Bruker AVIII 400 MHz spectrometer and were referenced to CFCl_3 ($\delta = 0$ ppm). All spectra were recorded at a constant temperature of 300 K. Coupling constants (J) are reported in hertz (Hz). Standard abbreviations indicating multiplicity were used as follows: m = multiplet, d = doublet, t = triplet, s = singlet. UV/vis/NIR spectra were collected using a Shimadzu UV-3600 UV/vis/NIR spectrometer using anhydrous solvents, which were filtered through Celite® prior to use. ATR-IR spectra were collected using either a Shimadzu IRAffinity-1S or



a Shimadzu FTIR 8400S spectrometer. Abbreviations indicating strength of bands were used as follows: s = strong, m = medium, w = weak.

Elemental analysis was performed at the University of Strathclyde, and a combustible (tungstic oxide) was used to ensure complete combustion of all complexes.

Crystallographic data were collected for single crystals of $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ (**1a/1b**), $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2]$ (**3**) and $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2^{\text{t}}]$ (**6**) on a Bruker D8 VENTURE diffractometer at 100 K using Mo-K α radiation ($\lambda = 0.71073 \text{ \AA}$). Single-crystal structure data for $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]\cdot\text{Et}_2\text{O}$ (**2a**), $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]\cdot\text{Tol}$ (**2b**) and $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2^{\text{t}}]$ (**5**) were collected by the EPSRC UK National Crystallography Service using a Rigaku AFC12 goniometer, mounted at the window of a FR-E+ SuperBright molybdenum rotating anode generator, and equipped with a (HG) Saturn724+ detector. CCDC numbers 1832429–1832434 and 1841191 \dagger contain the crystallographic information for this paper.

X-band EPR spectra were collected on a Bruker ELEXSYS E500 spectrometer and simulations were performed using Bruker's Xsophe software package.⁴⁷

Synthesis of $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ (**1**) \ddagger

THF (15 mL) was added to a Schlenk flask charged with a dry mixture of $\text{Y}(\text{OTf})_3$ (1.16 g, 2.17 mmol), pd (0.456 g, 2.17 mmol) and $\text{K}(\text{hfac})$ (1.60 g, 6.51 mmol) at room temperature with stirring. The yellow suspension was stirred at room temperature for 1 h, whereupon the solution became deep green. The green solution was stirred for 16 h, THF was removed *in vacuo* and the green solids were extracted into toluene ($3 \times 5 \text{ mL}$), giving a yellow solution. Removal of toluene *in vacuo* gave green solids, which were washed with hexane ($2 \times 5 \text{ mL}$) to give $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ as a pale green solid (1.70 g, 1.84 mmol, 85% yield). Single crystals suitable for X-ray diffraction were grown from both a saturated Et_2O solution at $-35 \text{ }^\circ\text{C}$ over 1 week ($[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]\cdot\text{Et}_2\text{O}$, **1a**) and a saturated toluene solution at $-15 \text{ }^\circ\text{C}$ overnight ($[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]\cdot\text{Tol}$, **1b**). ^1H NMR (C_6D_6): δ 6.23 (3H, s, hfac-CH) 6.37 (2H, dd, $J_{\text{HH}} = 7.8, 5.5 \text{ Hz}$, pd $\text{C}^{3,8}\text{-H}$) 7.53 (2H, d, $J_{\text{HH}} = 8.7 \text{ Hz}$, pd $\text{C}^{4,7}\text{-H}$) 9.16 (2H, d, $J_{\text{HH}} = 5.2 \text{ Hz}$, pd $\text{C}^{2,9}\text{-H}$) ppm. $^{19}\text{F}\{^1\text{H}\}$ NMR (C_6D_6): δ -76.93 (s, hfac- CF_3) ppm. Anal. calcd for $\text{C}_{27}\text{H}_9\text{N}_2\text{O}_8\text{F}_{18}\text{Y}\cdot\text{C}_7\text{H}_8$: C, 35.24%; H, 0.99%; N, 3.04%. Found: C, 35.06%; H, 0.96%; N, 3.14%. IR(ATR) 1699 (w, ν_{CO}) 1668 (w) 1647 (s) 1610 (w) 1579 (w) 1495 (m) 1426 (m) 1251 (s) 1197 (s) 1136 (s) 794 (s) 651 (s) 582 (m) cm^{-1} . λ_{max} (MeCN)/nm ($\epsilon/\text{M}^{-1} \text{ cm}^{-1}$): 590 (2300).

Synthesis of $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ (**2**)

THF (15 mL) was added to a Schlenk flask charged with a dry mixture of $\text{Gd}(\text{OTf})_3$ (0.759 g, 1.26 mmol), $\text{K}(\text{hfac})$ (0.997 g, 3.77 mmol) and 1,10-phenanthroline-5,6-dione (0.263 g,

1.25 mmol). The yellow suspension was stirred for 20 h, turning green after *ca.* 1 h. THF was removed *in vacuo* and $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ was extracted into toluene ($3 \times 10 \text{ mL}$) to give a yellow solution. Toluene was removed *in vacuo* to give a green solid, which was washed with hexanes to give $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ as a pale green solid (0.992 g, 1.00 mmol, 79% yield). Single crystals suitable for X-ray diffraction were grown from both saturated Et_2O ($[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]\cdot\text{Et}_2\text{O}$, **2a**) and toluene ($[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]\cdot\text{Tol}$, **2b**) solutions cooled to $-35 \text{ }^\circ\text{C}$ overnight. Anal. calcd for $\text{C}_{27}\text{H}_9\text{N}_2\text{O}_8\text{F}_{18}\text{Gd}\cdot\text{C}_7\text{H}_8$: C, 37.79%; H, 1.59%; N, 2.59%. Found: C, 37.58%; H, 1.47%; N, 2.54%. IR (ATR): 1697 (w, ν_{CO}) 1661 (w) 1647 (s) 1495 (m) 1252 (m) 1202 (m) 1142 (s) 1098 (w) 802 (s) 732 (m) 660 (s). λ_{max} (MeCN)/nm ($\epsilon/\text{M}^{-1} \text{ cm}^{-1}$): 597 (129). μ_{eff} (298 K, $d_8\text{-THF}$) 7.56–7.71 BM.

Synthesis of $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2]$ (**3**)

Et_2O (10 mL) was added to a $-78 \text{ }^\circ\text{C}$ cooled Schlenk flask charged with a dry mixture of **1** (0.50 g, 0.54 mmol) and VCp_2 (0.10 g, 0.54 mmol), immediately giving a dark green solution. The reaction was allowed to warm to room temperature and stirred for 2 h. Et_2O was removed *in vacuo* and the green/black product recrystallised from toluene/hexane at $-15 \text{ }^\circ\text{C}$ to give $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2]$ as a black microcrystalline solid (0.335 g, 0.30 mmol, 56% yield). Crystals suitable for X-ray diffraction were grown from a toluene/hexane mixture at $-35 \text{ }^\circ\text{C}$ overnight. Anal. calcd For $\text{C}_{37}\text{H}_{19}\text{F}_{18}\text{N}_2\text{O}_8\text{YV}$: C, 40.35%; H, 1.74%; N, 2.54%. Found: C, 39.88%; H, 1.81%; N, 2.56%. IR(ATR): 1651 (s) 1504 (m) 1371 (m, ν_{CO}) 1250 (s) 1197 (s) 1134 (s) 794 (s) 651 (s) 582 (m) cm^{-1} . λ_{max} (MeCN)/nm ($\epsilon/\text{M}^{-1} \text{ cm}^{-1}$): 431 (5400), 600 (2630), 788 (1840). μ_{eff} (298 K, $d_8\text{-THF}$): 1.71–1.74 BM.

Synthesis of $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2]$ (**4**)

THF (15 mL) was added to a dry mixture of $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ (500 mg, 0.505 mmol) and $\text{V}(\text{Cp})_2$ (91.3 mg, 0.504 mmol) and the dark suspension was stirred for 24 h. THF was removed *in vacuo*. $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{V}(\text{Cp})_2]$ was extracted into toluene and cooled to $-15 \text{ }^\circ\text{C}$, whereupon it precipitated as a black powder (378 mg, 0.322 mmol, 64% yield). Anal. calcd for $\text{C}_{37}\text{H}_{19}\text{F}_{18}\text{N}_2\text{O}_8\text{VGd}$: C, 37.99%; H, 1.64%; N, 2.39%. Found: C, 37.11%; H, 1.70%; N, 2.38%. IR(ATR): 1651 (m) 1497 (m) 1381 (w, ν_{CO}) 1258 (m) 1204 (w) 1134 (s) 1018 (s) 795 (s) 656 (m). λ_{max} (MeCN)/nm ($\epsilon/\text{M}^{-1} \text{ cm}^{-1}$): 425 (4650) 576 (3430) 797 (2740).

Synthesis of $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2^{\text{t}}]$ (**5**)

Et_2O (30 mL) was added to a $-78 \text{ }^\circ\text{C}$ cooled Schlenk flask charged with $\text{VCp}_2^{\text{tet}}$ (250 mg, 0.83 mmol) and $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ (770 mg, 0.83 mmol), giving a dark green suspension. The reaction was allowed to warm to room temperature, and stirred for 16 hours. The dark suspension was filtered into a Schlenk and dried *in vacuo* to give deep-green solids, which were recrystallised from a saturated solution of toluene and hexane at $-15 \text{ }^\circ\text{C}$ overnight. The micro-crystals were isolated, washed with hexanes ($3 \times 15 \text{ mL}$), and dried *in vacuo* to give $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-O},\text{O}^{\prime}\text{-pd})\text{VCp}_2^{\text{t}}]$ as a dark green powder (520 mg,

\ddagger We note that a complex formulated as $[\text{Y}(\text{hfac})_3(\text{N},\text{N}^{\prime}\text{-pd})]$ was reported in the following paper, *Chem. Papers*, 2016, 70, 43. However, the spectroscopic data reported therein is not consistent with our full characterisation of compound **1**.



10.43 mmol, 50% yield). Crystals suitable for X-ray diffraction were grown by diffusion of hexane vapour into a toluene solution at $-35\text{ }^{\circ}\text{C}$ over several days. Anal. calcd for $\text{C}_{45}\text{H}_{35}\text{F}_{18}\text{N}_2\text{O}_8\text{VY}$: C, 44.54%; H, 2.91%; N, 2.31%. Found: C, 44.49%; H, 3.78%; N, 2.49%. IR(ATR): 1651 (s), 1506 (m), 1373 (m, ν_{CO}), 1250 (s), 794 (m), 727 (w), 659 (m). λ_{max} (THF)/nm ($\epsilon/\text{M}^{-1}\text{ cm}^{-1}$): 418 (6970), 553 (2920), 816 (2340). μ_{eff} (298 K, $d_{\text{g}}\text{-THF}$) 1.72–1.75 BM.

Synthesis of $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}'\text{-O},\text{O}'\text{-pd})\text{VCp}^t_2]$ (6)

Et_2O (30 mL) was added to a $-78\text{ }^{\circ}\text{C}$ cooled Schlenk flask charged with $\text{VCp}^{\text{tet}}_2$ (38 mg, 0.13 mmol) and $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}'\text{-pd})]$ (130 mg, 0.13 mmol), giving a dark green suspension. The reaction was allowed to warm to room temperature, and stirred for 16 hours. The dark suspension was filtered into a Schlenk and dried *in vacuo* to give deep-green solids, which were recrystallised from a saturated solution of toluene and hexane at $-15\text{ }^{\circ}\text{C}$ overnight. The micro-crystals were isolated, washed with hexanes ($3 \times 15\text{ mL}$), and dried *in vacuo* to give $[\text{Gd}(\text{hfac})_3(\text{N},\text{N}'\text{-O},\text{O}'\text{-pd})\text{VCp}^t_2]$ as a dark green powder (70 mg, 5.40 mmol, 42% yield). Crystals suitable for X-ray diffraction were grown by diffusion of hexane vapour into an Et_2O solution at $-35\text{ }^{\circ}\text{C}$ over several days. Anal. calcd for $\text{C}_{45}\text{H}_{35}\text{F}_{18}\text{N}_2\text{O}_8\text{VGd}$: C, 42.16%; H, 2.75%; N, 2.19%. Found: C, 42.45%; H, 3.18%; N, 2.04%. IR(ATR): 1650 (s), 1500 (m), 1370 (w, ν_{CO}), 1250 (s), 800 (m), 733 (w), 662 (m). λ_{max} (THF)/nm ($\epsilon/\text{M}^{-1}\text{ cm}^{-1}$): 435 (6260), 792 (1420). μ_{eff} (298 K, $d_{\text{g}}\text{-THF}$) 7.99–8.41 BM.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We acknowledge the EPSRC for PhD studentships to JRH (EP/L504786/1) and SJH (EP/M508056/1), Imperial College London (JRF to JHF) and the University of Glasgow for funding. We thank Professor Nick Long for his mentorship while at ICL and for useful discussions. We also acknowledge the EPSRC UK National Crystallography Service at the University of Southampton.

References

- J. F. Berry and C. M. Thomas, *Dalton Trans.*, 2017, **46**, 5472–5473.
- K. Liu, W. Shi and P. Cheng, *Coord. Chem. Rev.*, 2015, **289–290**, 74–122.
- K. Griffiths, A. C. Tsipis, P. Kumar, O. P. E. Townrow, A. Abdul-Sada, G. R. Akien, A. Baldansuren, A. C. Spivey and G. E. Kostakis, *Inorg. Chem.*, 2017, **56**, 9563–9573.
- J. D. Rinehart, M. Fang, W. J. Evans and J. R. Long, *J. Am. Chem. Soc.*, 2011, **133**, 14236–14239.
- S. Demir, L.-R. Jeon, J. R. Long and T. D. Harris, *Coord. Chem. Rev.*, 2015, **289–290**, 149–176.
- C. A. P. Goodwin, F. Ortu, D. Reta, N. F. Chilton and D. P. Mills, *Nature*, 2017, **548**, 439–442.
- F.-S. Guo, B. M. Day, Y.-C. Chen, M.-L. Tong, A. Mansikkamäki and R. A. Layfield, *Angew. Chem., Int. Ed.*, 2017, **56**, 11445–11449.
- A. E. Friedman, J.-C. Chambron, J.-P. Sauvage, N. J. Turro and J. K. Barton, *J. Am. Chem. Soc.*, 1990, **112**, 4960–4962.
- K. Peuntinger, T. D. Pilz, R. Staehle, M. Schaub, S. Kaufhold, L. Petermann, M. Wunderlin, H. Görls, F. W. Heinemann, J. Li, T. Drewello, J. G. Vos, D. M. Guldi and S. Rau, *Dalton Trans.*, 2014, **43**, 13683–13695.
- C. A. Goss and H. D. Abruna, *Inorg. Chem.*, 1985, **24**, 4263–4267.
- J. Yuasa, T. Suenobu and S. Fukuzumi, *J. Phys. Chem. A*, 2005, **109**, 9356–9362.
- N. M. Shavaleev, L. P. Moorcraft, S. J. A. Pope, Z. R. Bell, S. Faulkner and M. D. Ward, *Chem. – Eur. J.*, 2003, **9**, 5283–5291.
- A. Y. Girgis, Y. S. Sohn and A. L. Balch, *Inorg. Chem.*, 1975, **14**, 2327–2331.
- G. A. Fox, S. Bhattacharya and C. G. Pierpont, *Inorg. Chem.*, 1991, **30**, 2895–2899.
- W. Paw and R. Eisenberg, *Inorg. Chem.*, 1997, **36**, 2287–2293.
- F. Calderazzo, F. Marchetti, G. Pampaloni and V. Passarelli, *J. Chem. Soc., Dalton Trans.*, 1999, 4389–4396.
- E. K. Brechin, L. Calucci, U. Englert, L. Margheriti, G. Pampaloni, C. Pinzino and A. Prescimone, *Inorg. Chim. Acta*, 2008, **361**, 2375–2384.
- F. Calderazzo, G. Pampaloni and V. Passarelli, *Inorg. Chim. Acta*, 2002, **330**, 136–142.
- F. Calderazzo and G. Pampaloni, *J. Organomet. Chem.*, 1987, **330**, 47–59.
- N. M. Shavaleev, L. P. Moorcraft, S. J. A. Pope, Z. R. Bell, S. Faulkner and M. D. Ward, *Chem. Commun.*, 2003, 1134–1135.
- S. Kobayashi, M. Suguira, H. Kitagawa and W. W.-L. Lam, *Chem. Rev.*, 2002, **102**, 2227–2302.
- K. D. Pollard, H. A. Jenkins and R. J. Puddephatt, *Chem. Mater.*, 2000, **12**, 701–710.
- G. Malandrino, R. L. Nigro, I. L. Fragala and C. Benelli, *Eur. J. Inorg. Chem.*, 2004, 500–509.
- I. V. Taidakov, Y. A. Strelenko, R. S. Borisov, A. Z. Temerdashev, N. P. Datskevich and A. G. Vitukhnovskii, *Russ. J. Coord. Chem.*, 2015, **41**, 230–239.
- M. Hardiman, J. Pellisson, S. E. Barnes, P. E. Bisson and M. Peter, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1980, **22**, 2175–2194.
- D. Roitershtein, A. Domingos, L. C. J. Pereira, J. R. Ascenso and N. Marques, *Inorg. Chem.*, 2003, **42**, 7666–7673.
- V. S. Herzog and K. Gustav, *Z. Anorg. Allg. Chem.*, 1966, **346**, 150–161.
- J.-W. Dai, Z.-Y. Li and O. Sato, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2014, **70**, o573.



- 29 A. Y. Rogachev, A. V. Mironov, S. I. Troyanov, N. P. Kuzmina and A. V. Nemukhin, *J. Mol. Struct.*, 2006, **789**, 187–194.
- 30 R. D. Shannon, *Acta Crystallogr., Sect. A: Cryst. Phys., Diffraction, Theor. Gen. Crystallogr.*, 1976, **32**, 751–767.
- 31 M. Wen, E. Kejia, M. Munakata, Y. Suenaga, T. Kuroda-Sowa, M. Maekawa and S. G. Yan, *Mol. Cryst. Liq. Cryst.*, 2006, **457**, 203–213.
- 32 D. M. Murphy, K. McNamara, P. Richardson, V. Sanchez-Romaguera, R. E. P. Winpenny and L. J. Yellowlees, *Inorg. Chim. Acta*, 2011, **374**, 435–441.
- 33 F. Ortu, J. Liu, M. Burton, J. M. Fowler, A. Formanuk, N. F. Chilton, M-E. Boulon and D. P. Mills, *Inorg. Chem.*, 2017, **56**, 2496–2505.
- 34 N. G. Connelly and W. E. Geiger, *Chem. Rev.*, 1996, **96**, 877–910.
- 35 J. D. L. Holloway and W. E. Geiger Jr., *J. Am. Chem. Soc.*, 1979, **101**, 2038–2044.
- 36 W. J. Evans, *Organometallics*, 2016, **35**, 3088–3100.
- 37 J. L. Petersen and L. Griffith, *Inorg. Chem.*, 1980, **19**, 1852–1858.
- 38 J. L. Petersen and L. F. Dahl, *J. Am. Chem. Soc.*, 1975, **97**, 6422–6432.
- 39 D. P. Bakalik and R. G. Hayes, *Inorg. Chem.*, 1972, **11**, 1734–1738.
- 40 J. A. Belot, R. D. McCullough, A. L. Rheingold and G. P. A. Yap, *Organometallics*, 1996, **15**, 5062–5065.
- 41 B. Gleeson, J. Claffey, M. Hogan, H. Müller-Bunz, D. Wallis and M. Tacke, *J. Organomet. Chem.*, 2009, **694**, 1369–1374.
- 42 P. Ghosh, S. Ghosh, C. Navara, R. K. Narla, A. Benyumov and F. M. Uckun, *J. Inorg. Biochem.*, 2001, **84**, 241–253.
- 43 G. H. Spikes, S. Sproules, E. Bill, T. Weyhermu and K. Wieghardt, *Inorg. Chem.*, 2008, **47**, 10935–10944.
- 44 S. R. Cooper, Y. B. Koh and K. N. Raymond, *J. Am. Chem. Soc.*, 1982, **104**, 5092–5102.
- 45 C. Zixiang, Z. Linli, L. Xiaoming, Y. Hongtao and L. Bing, *Z. Anorg. Allg. Chem.*, 2012, **638**, 1523–1530.
- 46 L. E. Manzer, E. A. Mintz and T. J. Marks, *Inorg. Synth.*, 1989, **28**, 260–263.
- 47 G. R. Hanson, K. E. Gates, C. J. Noble, M. Griffin, A. Mitchell and S. Benson, *J. Inorg. Biochem.*, 2004, **98**, 903–916.

