

Cite this: *Dalton Trans.*, 2018, **47**, 2406

Synthesis and properties of MoCl₄ complexes with thio- and seleno-ethers and their use for chemical vapour deposition of MoSe₂ and MoS₂ films†

Yao-Pang Chang, Andrew L. Hector,  William Levason,  Gillian Reid * and Joshua Whittam

Treatment of *trans*-[MoCl₄(MeCN)₂] with L (L = Me₂S, Me₂Se, THT, $\frac{1}{2}$ MeSCH₂CH₂SMe) in CH₂Cl₂ solution, or reaction of MoCl₅ with excess L' (L' = ⁿBu₂S, ⁿBu₂Se, $\frac{1}{2}$ MeSCH₂CH₂SMe, $\frac{1}{2}$ PrSCH₂CH₂SⁱPr, $\frac{1}{2}$ MeSCH₂CH₂CH₂SMe, $\frac{1}{2}$ MeSeCH₂CH₂CH₂SeMe) in MeCN, produces the Mo(IV) complexes, [MoCl₄(L)₂] and [MoCl₄(L')₂], respectively, in good yield. The new complexes have been characterised by IR and UV-vis spectroscopy, elemental analysis and magnetic measurements, whilst crystal structure analyses of *trans*-[MoCl₄(Me₂S)₂], *cis*-[MoCl₄{RS(CH₂)₂SR}] (R = Me, ⁱPr) and *cis*-[MoCl₄{MeS(CH₂)₃SMe}] confirmed their identities and distorted octahedral geometries. The potential of [MoCl₄(ⁿBu₂E)₂] (E = S, Se) as the first examples of molybdenum halide derived single source CVD precursors for the growth of MoE₂ thin films was first probed by TGA, which showed multi-step decomposition processes, with the masses of the final residues consistent with MoSe₂ (E = Se) and MoCl₄ (E = S), respectively. Low pressure CVD experiments conducted at 750 °C (E = S) and 525 °C (E = Se) gave silvery and golden yellow films, respectively. Grazing incidence and in plane XRD data confirmed these to be 2H-MoE₂, whilst the texture of the MoSe₂ was investigated using pole figure measurements. SEM and EDX data, optical and Raman data on the films are also reported.

Received 18th November 2017,
Accepted 15th December 2017

DOI: 10.1039/c7dt04352j

rsc.li/dalton

Introduction

The coordination chemistry of molybdenum extends over nine formal oxidation states (−II to +VI) and has been intensively studied over the last 50 years.^{1–3} Over this period, in addition to exploring the basic chemistry, much effort has been devoted to studying the binding and reduction of dinitrogen by molybdenum complexes,^{2,4} and in modelling the active sites of molybdoenzymes.^{4,5} In the last 10 years the study of early d-block dichalcogenides, including MoS₂ and MoSe₂, has become the focus of much effort. Layered early transition metal dichalcogenides, ME₂ (M = Nb, Ta, V, W, *etc.*; E = S, Se or Te), are inorganic analogues of graphene, and their properties can be tuned by varying the metal and the chalcogen. Production of the materials as thin films maximises the anisotropy of their magnetic or electronic properties and thus methods to deposit 2D layers are of particular interest currently.^{6–8} Applications of these materials in optoelectronics, spintronics, sensors, electrocatalysts and magnetic materials

are being actively developed.^{8,9} The most common methods of forming MoS₂ or MoSe₂ thin films have been exfoliation from bulk samples, or heating MoO₃ (or sometimes molybdenum metal) with S or Se in a hydrogen/inert gas atmosphere, to form the desired MoE₂ (E = S or Se) 2D films on an appropriate substrate.^{10,11–13} Thermal decomposition of [NH₄]₂[MoS₄] in a hydrogen atmosphere has also recently been reported.¹³

Chalcogenoether complexes of the early d-block metal halides can function as single source CVD precursors for some ME₂ (for example, M = Ti, V or Nb; E = S or Se).^{14,15} Chemical vapour deposition from molybdenum halide complexes has rarely been used to form molybdenum dichalcogenides, although MoSe₂ films were produced by dual source atmospheric pressure chemical vapour deposition (CVD) using MoCl₅ and either Et₂Se or ⁿBu₂Se.¹⁶ The tetrathiolate complex, [Mo(SⁱBu)₄], has also been used to deposit MoS₂ films by CVD.¹⁷

Although there are a few early reports of MoCl₄ complexes with mono-, di- and poly-thioethers and thiamacrocycles,^{1,2,18} the only recent work is the detailed characterisation, including the X-ray crystal structures, of *trans*-[MoCl₄(Et₂S)₂]¹⁹ and *trans*-[MoCl₄(THT)₂] (THT = tetrahydrothiophene).²⁰

Here we report the preparation, spectroscopic and structural characterisation of a series of new complexes of molybdenum(IV) chloride with thio- and seleno-ether ligands and the evaluation of a subset of these as single source low pressure

Chemistry, University of Southampton, Southampton SO17 1BJ, UK.

E-mail: G.Reid@soton.ac.uk

† Electronic supplementary information (ESI) available. CCDC 1557208–1557213. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7dt04352j



CVD reagents for the deposition of the corresponding MoE₂ thin films.

Experimental

Syntheses were performed by using standard Schlenk and glove-box techniques under a dry N₂ atmosphere. MoCl₅ was obtained from Sigma-Aldrich and stored in a glovebox. Tetrahydrothiophene, Me₂S and ⁿBu₂S were purchased from Sigma-Aldrich and Me₂Se from Strem. Solvents were dried by distillation from CaH₂ (CH₂Cl₂, MeCN), Na (Et₂O) or Na/benzophenone ketyl (*n*-hexane). Dichalcogenoethers MeE(CH₂)_nEMe (E = S, Se; *n* = 2, 3), ⁱPrS(CH₂)₂SⁱPr and ⁿBu₂Se were prepared *via* literature methods.^{21,22}

Infrared spectra were recorded using a PerkinElmer Spectrum 100 spectrometer in the range 4000–200 cm⁻¹, and samples prepared as Nujol mulls between CsI plates. UV/visible spectra were recorded as powdered solids, using the diffuse reflectance attachment of a PerkinElmer 750S spectrometer. Microanalyses on new compounds were undertaken by London Metropolitan University. Thermogravimetric analysis (TGA) used a Netzsch TG209 F1 Libra analyser under a flow of argon at 65 mL min⁻¹, contained within a dry, nitrogen purged glovebox. The temperature was increased at a rate of 10 °C min⁻¹.

Single crystal X-ray experimental

Data collections used a Rigaku AFC12 goniometer equipped with an enhanced sensitivity (HG) Saturn724+ detector mounted at the window of an FR-E+ SuperBright molybdenum ($\lambda = 0.71073$) rotating anode generator with VHF Varimax optics (70 micron focus) with the crystal held at 100 K (N₂ cryostream). Structure solution and refinement were performed using SHELX(S/L)97 and SHELX-2014/7 and were straightforward.^{23–25} H atoms were added and refined using a riding model. The X-ray data collection details are in Table S1 (ESI[†]). CCDC reference numbers in CIF format are [MoCl₄(NCMe)₂] 1557208; [MoCl₄(Me₂S)₂] 557209; [MoCl₄(MeSCH₂CH₂SMe)] 557210; [MoCl₄(ⁱPrSCH₂CH₂SⁱPr)] 1557211; [MoCl₄(MeSCH₂CH₂CH₂SMe)] 1557212; [MoCl₅(Me₂S)][Me₂SCH₂SMe] 1557213.†

Complex synthesis

trans-[MoCl₄(MeCN)₂]. Was made by a modified literature method.²⁶ MoCl₅ (270 mg, 1.0 mmol) was dissolved in MeCN (20 mL) with stirring for 30 min to give a dark brown solution. The solution was evaporated *in vacuo*. The resulting dark brown solid was washed by *n*-hexane (5 mL), the washings decanted off, and the solid dried *in vacuo*. Yield: 249 mg, 78%. Required for C₄H₆Cl₄MoN₂ (319.56): C, 15.02; H, 1.89; N, 8.76. Found: C, 14.84; H, 1.81; N, 8.61%. IR (Nujol/cm⁻¹): 2314, 2283 (MeCN), 335 (Mo–Cl).

trans-[MoCl₄(THT)₂]. MoCl₅ (250 mg, 0.915 mmol) was dissolved in CH₂Cl₂ (10 mL). THT (0.26 mL, 2.97 mmol) was then added with stirring for 30 min during which the solution

changed to orange. The solution was taken to dryness *in vacuo*, to produce a dark orange solid. Yield: 296 mg, 78%. Required for C₈H₁₆Cl₄MoS₂ (414.0): C, 23.20; H, 3.89. Found: C, 23.28; H, 3.73%. IR (Nujol/cm⁻¹): 338 (Mo–Cl). UV/vis (diffuse reflectance)/cm⁻¹: 22 700, 19 400, 16 600 (sh).

trans-[MoCl₄(Me₂S)₂]. MoCl₅ (200 mg, 0.75 mmol) was dissolved in CH₂Cl₂ (15 mL). Anhydrous Me₂S (1 mL) was then added with stirring for 30 min and gave a red-orange solution. The solution was filtered and evaporated *in vacuo*. *n*-Hexane (10 mL) was added to wash the solid and then removed with a syringe. The dark orange solid was dried *in vacuo*. Yield: 206 mg, 76%. Required for C₄H₁₂Cl₄MoS₂ (360.02): C, 13.27; H, 3.34. Found: C, 13.36; H, 3.48%. IR (Nujol/cm⁻¹): 332 (Mo–Cl). UV/vis (diffuse reflectance)/cm⁻¹: 22 000, 19 400, 16 800 (sh). μ_{eff} : 2.22 B.M (298 K).

Orange crystals of *trans*-[MoCl₄(Me₂S)₂] grew by allowing a CH₂Cl₂ solution to evaporate slowly in a N₂ atmosphere. A few orange-red crystals were also obtained by recrystallisation of crude [MoCl₄(Me₂S)₂] from CH₂Cl₂; these were found to be [Me₂SCH₂SMe][MoCl₅(Me₂S)] from an X-ray structure determination.

Alternative method: MoCl₅ (150 mg, 0.55 mmol) was dissolved in MeCN (10 mL) with stirring to give a dark brown solution. The solution was stirred for another 10 min and taken to dryness to produce a dark brown solid. The solid was suspended in CH₂Cl₂ (10 mL) before adding Me₂S (3 mL). The suspended solid dissolved immediately after the Me₂S was added in to give a red-brown solution. The solution was stirred for another 10 min and then pumped to dryness, leaving a dark orange solid. The product was spectroscopically identical with the product made from MoCl₅ in CH₂Cl₂.

trans-[MoCl₄(ⁿBu₂S)₂]. MoCl₅ (205 mg, 0.75 mmol) was dissolved in CH₂Cl₂ (10 mL). A solution of ⁿBu₂S (274 mg, 2.5 mmol) in CH₂Cl₂ (*ca.* 3 mL) was then added with stirring for 30 min and gave a red-orange solution. The solution was filtered and evaporated to dryness *in vacuo*. The residue was washed with *n*-hexane (10 mL), and the dark orange oil dried *in vacuo*. Yield: 250 mg, 62%. Required for C₁₆H₃₆Cl₄MoS₂ (526.31): C, 36.51; H, 6.13. Found: C, 36.15; H, 6.72%. IR (Nujol/cm⁻¹): 334, 305sh (Mo–Cl). UV/vis (diffuse reflectance)/cm⁻¹: 41 700, 20 900, 19 100, 18 500 (sh).

Alternative method: MoCl₅ (136 mg, 0.5 mmol) was dissolved in MeCN (10 mL) with stirring for 30 min to give a dark brown solution. The solution taken to dryness *in vacuo* and left a dark brown solid. CH₂Cl₂ (15 mL) was added, followed by a solution of ⁿBu₂S (0.4 mL, 2.0 mmol) in CH₂Cl₂ (2 mL). The dark solution changed to red-orange immediately. The solution was then stirred for 15 min then evaporated *in vacuo* to afford a dark orange oil. The product was spectroscopically identical to that made from MoCl₅ in CH₂Cl₂.

trans-[MoCl₄(Me₂Se)₂]. MoCl₅ (205 mg, 0.75 mmol) was suspended in CH₂Cl₂ (15 mL) and Me₂Se (196 mg, 1.96 mmol) was added with stirring for 30 min and produced an orange-red solution. The solution was filtered and taken to dryness *in vacuo*. The solid was washed with *n*-hexane (10 mL), the washings removed with a syringe, and the residual dark orange



powder dried *in vacuo*. Yield: 283 mg, 77%. Required for $C_4H_{12}Cl_4MoSe_2$ (455.81): C, 10.54; H, 2.65. Found: C, 10.69; H, 2.43%. IR (Nujol/ cm^{-1}): 306 (Mo–Cl). μ_{eff} : 2.59 B.M. UV/vis (diffuse reflectance)/ cm^{-1} : 37 000, 33 300, 26 800, 21 800, 19 000.

Alternative method: $[MoCl_4(MeCN)_2]$ (100 mg, 0.31 mmol) was dissolved in 10 mL CH_2Cl_2 with stirring. A solution of Me_2Se (73 mg, 0.65 mmol) and CH_2Cl_2 (5 mL) was then added. The solution was stirred for 1 h. resulting in a red-orange solution. Evaporation *in vacuo* left a dark orange powder. Yield: 99 mg, 70%. Required for $C_4H_{12}Cl_4MoSe_2$ (455.81): C, 10.54; H, 2.65. Found: C, 10.61; H, 2.54%. The product was spectroscopically identical to that made from $MoCl_5$.

trans-[$MoCl_4(^nBu_2Se)_2$]. $MoCl_5$ (136 mg, 0.5 mmol) was dissolved in MeCN (15 mL) with stirring for 30 min to give a dark brown solution. The solution was taken to dryness *in vacuo* and left a dark brown solid. CH_2Cl_2 (25 mL) was added, followed by a solution of nBu_2Se (193 mg, 1.0 mmol) in CH_2Cl_2 (2 mL). The dark solution changed to red-orange immediately. The solution was then stirred for 15 min then evaporated *in vacuo* to afford a dark red-orange oil. Yield: 292 mg, 94%. Required for $C_{16}H_{36}Cl_4MoSe_2$ (624.13): C, 30.79; H, 5.81. Found: C, 30.23; H, 6.00%. IR (Nujol/ cm^{-1}): 342s, 305 (sh) (Mo–Cl). UV/vis (diffuse reflectance)/ cm^{-1} : 40 000, 32 000, 28 000, 22 900, 21 000.

[$MoCl_4(MeSCH_2CH_2SMe)$]. $MoCl_5$ (205 mg, 0.75 mmol) was dissolved in CH_2Cl_2 (10 mL). A solution of $MeSCH_2CH_2SMe$ (92 mg, 0.75 mmol) in CH_2Cl_2 (*ca.* 3 mL) was added and the mixture stirred overnight. A deep orange solution and a brown solid formed. More CH_2Cl_2 (10 mL) was added and the solution was filtered. The filtrate was evaporated to dryness *in vacuo*. *n*-Hexane (10 mL) was added to wash the solid and subsequently removed by a syringe. The solid was dried *in vacuo* to give a dark orange solid. Yield: 157 mg, 58%. IR (Nujol/ cm^{-1}): 356, 319, 293 (Mo–Cl). UV/vis (diffuse reflectance)/ cm^{-1} : 41 500, 27 300, 21 000, 19 600. μ_{eff} : 2.66 B.M. Orange crystals grew on allowing a CH_2Cl_2 solution to evaporate in a N_2 atmosphere.

Alternative method: $MoCl_5$ (136 mg, 0.5 mmol) was dissolved in MeCN (10 mL) with stirring for 10 min to give a dark brown solution. The solution taken to dryness *in vacuo*, CH_2Cl_2 (10 mL) was added, followed by a solution of $MeSCH_2CH_2SMe$ (180 mg, 1.5 mmol) in CH_2Cl_2 (2 mL). The dark solution changed to light orange-green immediately and deposited a brown solid, which was filtered off. The brown solid was dried *in vacuo* and washed with Et_2O (*ca.* 5 mL), then extracted with CH_2Cl_2 (*ca.* 5 mL). The resulting solution was again taken to dryness *in vacuo*. Yield: 60 mg, 33%. Required for $C_4H_{10}Cl_4MoS_2 \cdot 0.1Et_2O$ (367.4): C, 14.38; H, 3.02. Found: C, 14.38; H, 3.02% (sample contained *ca.* 10% Et_2O from wash solvent; this was identified *via* 1H NMR spectroscopy).

[$MoCl_4(^iPrSCH_2CH_2S^iPr)$]. $MoCl_5$ (136 mg, 0.5 mmol) was dissolved in CH_3CN (10 mL) forming dark brown solution. The solution was stirred for 10 min and then evaporated *in vacuo*, giving dark brown solid. The solid was dissolved in CH_2Cl_2 (15 mL) and a solution of $^iPrSCH_2CH_2S^iPr$ (270 mg, 1.5 mmol) in CH_2Cl_2 (2 mL) was added. The colour changed to orange

and solid formed immediately. The solution was stirred for 30 min and then evaporated *in vacuo*. The solid was washed with Et_2O (5 mL), the Et_2O removed *via* a syringe, and the solid was dried *in vacuo*. Orange powder. Yield: 160 mg, 77%. Required for $C_8H_{18}Cl_4MoS_2$ (416.11): C, 23.09; H, 4.36. Found: C, 23.90; H, 4.56%. IR (Nujol/ cm^{-1}): 371sh, 350, 309 (Mo–Cl). UV/vis (diffuse reflectance)/ cm^{-1} : 41 300, 26 800, 25 000 (sh), 21 000, 18 700 (sh). μ_{eff} : 2.18 B.M. Orange red crystals grew by allowing a CH_2Cl_2 solution to evaporate slowly under N_2 .

[$MoCl_4(MeSCH_2CH_2CH_2SMe)$]. $MoCl_5$ (205 mg, 0.75 mmol) was suspended in CH_2Cl_2 (20 mL) and a solution of $MeSCH_2CH_2CH_2SMe$ (109 mg, 0.8 mmol) in CH_2Cl_2 (*ca.* 3 mL) was added, and stirred overnight to give an orange solution and some orange solid. The solution was removed by a syringe, and the solid washed with *n*-hexane (10 mL). The solid was subsequently dried *in vacuo*. Yield: 83 mg, 30%. Required for $C_5H_{12}Cl_4MoS_2$ (374.03): C, 16.06; H, 3.23. Found: C, 16.22; H, 3.35%. IR (Nujol/ cm^{-1}): 362, 342, 327 (Mo–Cl). UV/vis (diffuse reflectance)/ cm^{-1} : 40 500, 22 070, 19 500. μ_{eff} : 2.21 B.M. Orange crystals grew by allowing a CH_2Cl_2 solution to evaporate under N_2 .

[$MoCl_4(MeSeCH_2CH_2CH_2SeMe)$]. Was made as described for $[MoCl_4(^iPrSCH_2CH_2S^iPr)]$ and obtained as a light orange powder. Yield: 59%. Required for $C_5H_{12}Cl_4MoSe_2$ (467.82): C, 12.84; H, 2.59. Found: C, 13.01; H, 2.57%. IR (Nujol/ cm^{-1}): 332, 308 (Mo–Cl). UV/vis (diffuse reflectance)/ cm^{-1} : 41 100, 26 800, 21 000, 18 700.

LPCVD of MoS_2 films using $[MoCl_4(^nBu_2S)_2]$. The precursor (190 mg) dissolved in CH_2Cl_2 (1 mL) was loaded into the closed end of a silica tube in an N_2 purged glove box, and the solvent removed. Silica substrates (*ca.* $1 \times 8 \times 20$ mm³) were then loaded in the tube and placed end-to-end. The tube was set in a furnace so that the substrates were in the heated zone and the precursor was *ca.* 2 cm away from the start of the heated zone. The tube was evacuated to 0.1 mm Hg, and the furnace was heated to 750 °C and left for 10 minutes to allow the temperature to stabilise. The tube was gradually moved into the hot zone until evaporation of the precursor began to occur. The position was then maintained until the all the precursor had evaporated. A yellow/orange film was observed at the opposite end of the tube (S). After *ca.* 30 min the tube was cooled to room temperature and the tiles were unloaded at ambient temperature in a dinitrogen atmosphere. Silver films were observed on the first two substrates, which corresponds to 735 °C (determined by temperature profiling). The silver films are identical by grazing incidence and in-plane XRD and SEM/EDX analysis.

$MoSe_2$ films using $[MoCl_4(^nBu_2Se)_2]$. Using the same deposition procedure and with *ca.* 50 mg of precursor complex. Depositions were undertaken with the furnace temperature set at 400, 450, 500 and 500 °C. A red (Se) film was observed at the far end of the tube and all experiments produced golden/brown films on the substrates. Temperature profiling for the best $MoSe_2$ films confirmed that the actual deposition temperature was 535 °C. Using larger precursor quantities (*ca.* 200 mg) led to thicker films being formed.



Film characterisation

X-Ray diffraction (XRD) patterns were collected in grazing incidence mode ($\theta_1 = 1^\circ$) or in-plane mode ($\theta_1 = 0.5^\circ$, $2\theta_z$ scan with the detector scanning in the film plane) using a Rigaku SmartLab diffractometer (Cu-K α , $\lambda = 1.5418 \text{ \AA}$) with parallel X-ray beam and a DTex Ultra 250 1D detector. Phase matching, lattice parameter calculations (MoS $_2$) and Le Bail fitting (MoSe $_2$) used the PDXL2 software package²⁷ and diffraction patterns from ICSD.²⁸ Scanning electron microscopy (SEM) was performed on samples at an accelerating voltage of 10 to 15 kV using a JEOL JSM6500 or a Philips XL30 ESEM. Film thicknesses were measured by fracturing the substrate and gold sputtering the edge in order to control charging of the insulating silica surface. Energy dispersive X-ray (EDX) analysis data were obtained with an Oxford INCA x-act X-ray detector (JSM6500) or Thermofisher Ultradrly NSS 3 (XL30). Raman spectra were collected using a Renishaw InVia Raman Microscope with a 100 mW He-Ne 785 nm Laser. Optical spectra were obtained in transmission mode using a PerkinElmer 750S UV-visible spectrometer.

Results and discussion

Two routes were used to prepare the MoCl $_4$ complexes (Scheme 1), either substitution of the MeCN ligands in *trans*-[MoCl $_4$ (MeCN) $_2$] with the chalcogenoether ligand in CH $_2$ Cl $_2$ solution or reaction of MoCl $_5$ with excess ligand in CH $_2$ Cl $_2$.

[MoCl $_4$ (RCN) $_2$] (R = Me, Et, ⁿPr) are well known starting materials^{29–31} and an X-ray crystal structure determination of the complex with R = Me showed it to be the *trans* isomer (see Fig. S1, ESI †). The crystals are isomorphous with *trans*-[WCl $_4$ (MeCN) $_2$].³¹ In the preparations using MoCl $_5$, the chalcogen ligand reduces the molybdenum centre, being halogenated in the process, and careful washing is needed to remove the by-products. A few crystals of a minor by-product from the synthesis of [MoCl $_4$ (Me $_2$ S) $_2$] from MoCl $_5$ were identified by a

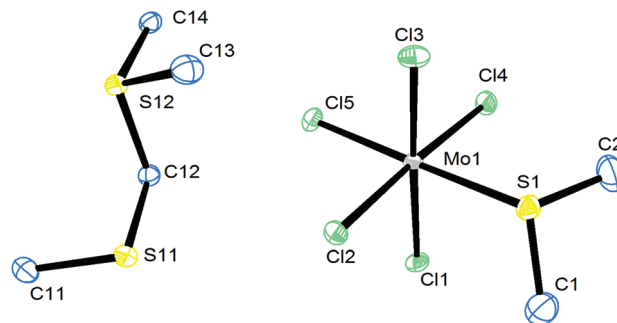
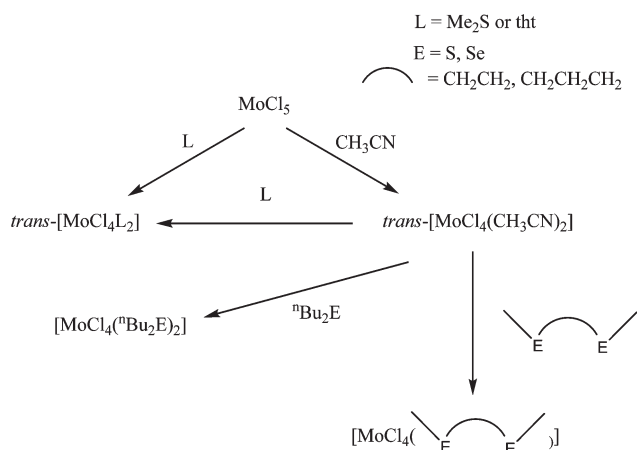


Fig. 1 The structure of one of the two crystallographically independent cations and molybdenum anions in [Me $_2$ SCH $_2$ SMe][MoCl $_5$ (Me $_2$ S)] showing the atom numbering scheme and with ellipsoids drawn at the 50% probability level. The other cation and anion are very similar. Hydrogen atoms are omitted for clarity. Selected bond lengths (\AA) and angles ($^\circ$): Mo1–Cl4 = 2.3251(17), Mo1–Cl1 = 2.3509(17), Mo1–Cl3 = 2.3633(18), Mo1–Cl5 = 2.3847(18), Mo1–Cl2 = 2.3885(17), Mo1–S1 = 2.5538(19), (*cis*) Cl–Mo–Cl = 88.10(6)–92.87(7), Cl–Mo1–S1 = 83.14(6)–91.74(6).

crystal structure determination as the sulfonium salt, [Me $_2$ SCH $_2$ SMe][MoCl $_5$ (Me $_2$ S)] (Fig. 1). This cation has also been identified as a product of reaction of MoCl $_5$ with DMSO,³² in the latter case it was obtained with an [MoOCl $_4$] $^-$ anion.³²

The isolated, coordinatively saturated MoCl $_4$ L $_2$ complexes are mildly hydrolytically sensitive. The crystal structures were determined for *trans*-[MoCl $_4$ (Me $_2$ S)], *cis*-[MoCl $_4$ {RS(CH $_2$) $_2$ SR}] (R = Me, ⁱPr) and *cis*-[MoCl $_4$ {MeS(CH $_2$) $_3$ SMe}] and show the expected six-coordinate geometries (Fig. 2–4), with the chelate bite of the dithioethers responsible for the deviations from regular octahedral geometry. The bond lengths are similar to those in related Mo(IV) complexes.^{19,20,33}

The complexes are paramagnetic with μ_{eff} in the range 2.0–2.7 B.M. The values are lower than the predicted spin-only value for a d 2 ion (2.82 B.M.) due to the significant spin-orbit coupling, as expected for 4d elements,³⁴ and are similar to



Scheme 1 The preparative routes to the MoCl $_4$ chalcogenoether complexes.

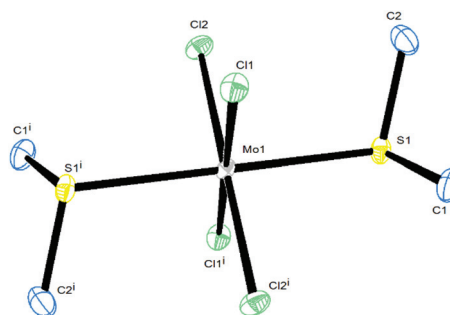


Fig. 2 The structure of the centrosymmetric [MoCl $_4$ (Me $_2$ S) $_2$] showing the atom numbering scheme and with ellipsoids drawn at the 50% probability level. Hydrogen atoms are omitted for clarity. Symmetry operation: $i = -x, -y, -z + 1$. Selected bond lengths (\AA) and angles ($^\circ$): Mo1–Cl1 = 2.3457(5), Mo1–Cl2 = 2.3323(5), Mo1–S1 = 2.5297(6); Cl1–Mo1–Cl2 = 89.39(2), Cl1 i –Mo1–Cl2 i = 90.61(2), Cl1–Mo1–S1 i = 88.65(2), Cl1 i –Mo1–S1 i = 91.35(2), Cl2–Mo1–S1 = 90.51(2), Cl2 i –Mo1–S1 = 89.49(2).



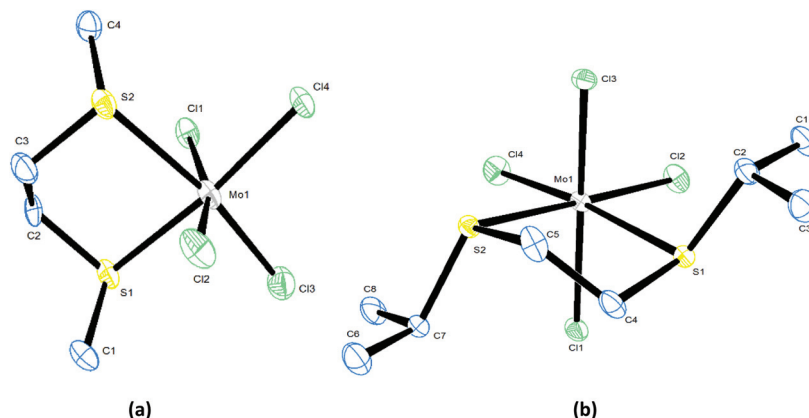


Fig. 3 (a) The structure of $[\text{MoCl}_4(\text{MeSCH}_2\text{CH}_2\text{SMe})]$ showing the atom numbering scheme and with ellipsoids drawn at the 50% probability level. Hydrogen atoms are omitted for clarity. Selected bond lengths (\AA) and angles ($^\circ$): $\text{Mo1}-\text{Cl3} = 2.247(3)$, $\text{Mo1}-\text{Cl1} = 2.305(3)$, $\text{Mo1}-\text{Cl2} = 2.305(3)$, $\text{Mo1}-\text{Cl4} = 2.339(3)$, $\text{Mo1}-\text{S1} = 2.519(3)$, $\text{Mo1}-\text{S2} = 2.591(3)$, $\text{Cl3}-\text{Mo1}-\text{Cl1} = 97.61(12)$, $\text{Cl3}-\text{Mo1}-\text{Cl2} = 98.42(13)$, $\text{Cl3}-\text{Mo1}-\text{Cl4} = 96.83(11)$, $\text{Cl1}-\text{Mo1}-\text{Cl4} = 92.47(10)$, $\text{Cl2}-\text{Mo1}-\text{Cl4} = 92.95(10)$, $\text{Cl3}-\text{Mo1}-\text{S1} = 87.75(11)$, $\text{Cl1}-\text{Mo1}-\text{S1} = 82.18(10)$, $\text{Cl2}-\text{Mo1}-\text{S1} = 91.07(10)$, $\text{Cl1}-\text{Mo1}-\text{S2} = 85.34(10)$, $\text{Cl2}-\text{Mo1}-\text{S2} = 77.66(11)$, $\text{Cl4}-\text{Mo1}-\text{S2} = 92.13(10)$, $\text{S1}-\text{Mo1}-\text{S2} = 83.62(9)$. (b) The structure of $[\text{MoCl}_4(\text{PrSCH}_2\text{CH}_2\text{SPr})]$ showing the atom numbering scheme and with ellipsoids drawn at the 50% probability level. Hydrogen atoms are omitted for clarity. Selected bond lengths (\AA) and angles ($^\circ$): $\text{Mo1}-\text{Cl1} = 2.3203(7)$, $\text{Mo1}-\text{Cl2} = 2.3295(7)$, $\text{Mo1}-\text{Cl3} = 2.3135(7)$, $\text{Mo1}-\text{Cl4} = 2.3016(7)$, $\text{Mo1}-\text{S1} = 2.5872(7)$, $\text{Mo1}-\text{S2} = 2.5730(7)$, $\text{Cl2}-\text{Mo1}-\text{Cl4} = 93.37(3)$, $\text{Cl4}-\text{Mo1}-\text{Cl3} = 97.30(3)$, $\text{Cl4}-\text{Mo1}-\text{Cl1} = 96.60(3)$, $\text{Cl3}-\text{Mo1}-\text{Cl2} = 96.57(3)$, $\text{Cl1}-\text{Mo1}-\text{Cl2} = 96.14(3)$, $\text{Cl4}-\text{Mo1}-\text{S2} = 92.81(3)$, $\text{Cl3}-\text{Mo1}-\text{S2} = 79.16(2)$, $\text{Cl1}-\text{Mo1}-\text{S2} = 86.63(2)$, $\text{Cl3}-\text{Mo1}-\text{S1} = 87.01(2)$, $\text{Cl1}-\text{Mo1}-\text{S1} = 78.40(2)$, $\text{Cl2}-\text{Mo1}-\text{S1} = 89.61(2)$, $\text{S2}-\text{Mo1}-\text{S1} = 84.52(2)$.

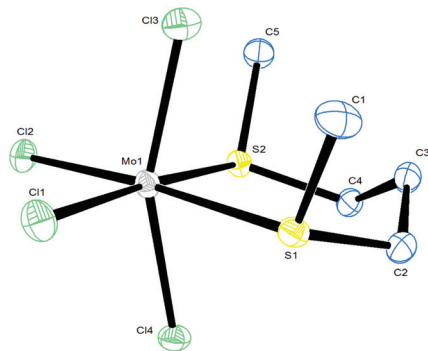


Fig. 4 The structure of $[\text{MoCl}_4(\text{MeSCH}_2\text{CH}_2\text{CH}_2\text{SMe})]$ showing the atom numbering scheme and with ellipsoids drawn at the 50% probability level. Hydrogen atoms are omitted for clarity. Selected bond lengths (\AA) and angles ($^\circ$): $\text{Mo1}-\text{Cl3} = 2.2686(6)$, $\text{Mo1}-\text{Cl1} = 2.3143(6)$, $\text{Mo1}-\text{Cl4} = 2.3380(6)$, $\text{Mo1}-\text{Cl2} = 2.3486(6)$, $\text{Mo1}-\text{S1} = 2.5282(6)$, $\text{Mo1}-\text{S2} = 2.5390(6)$, $\text{Cl3}-\text{Mo1}-\text{Cl1} = 98.40(2)$, $\text{Cl1}-\text{Mo1}-\text{Cl4} = 97.61(2)$, $\text{Cl3}-\text{Mo1}-\text{Cl2} = 96.29(2)$, $\text{Cl1}-\text{Mo1}-\text{Cl2} = 92.82(2)$, $\text{Cl4}-\text{Mo1}-\text{Cl2} = 97.58(2)$, $\text{Cl3}-\text{Mo1}-\text{S1} = 87.41(2)$, $\text{Cl1}-\text{Mo1}-\text{S1} = 83.61(2)$, $\text{Cl4}-\text{Mo1}-\text{S1} = 79.79(2)$, $\text{Cl3}-\text{Mo1}-\text{S2} = 86.13(2)$, $\text{Cl4}-\text{Mo1}-\text{S2} = 78.20(2)$, $\text{Cl2}-\text{Mo1}-\text{S2} = 85.94(2)$, $\text{S1}-\text{Mo1}-\text{S2} = 97.370(19)$.

those observed in related complexes.^{18–20,29,33} The solid state UV/visible spectra exhibit two overlapping bands in the range $18\,000\text{--}22\,000\text{ cm}^{-1}$, which can be assigned in approximate O_h symmetry as the d–d transitions, ${}^3T_{1g} \rightarrow {}^3T_{1g}(\text{P})$ and ${}^3T_{1g} \rightarrow {}^3T_{2g}$,³⁵ with ill-defined charge transfer transitions evident at $>30\,000\text{ cm}^{-1}$.

CVD of molybdenum dichalcogenide films

The thermal decomposition of the $[\text{MoCl}_4(\text{Bu}_2\text{E})_2]$ complexes at atmospheric pressure was first probed using thermo-

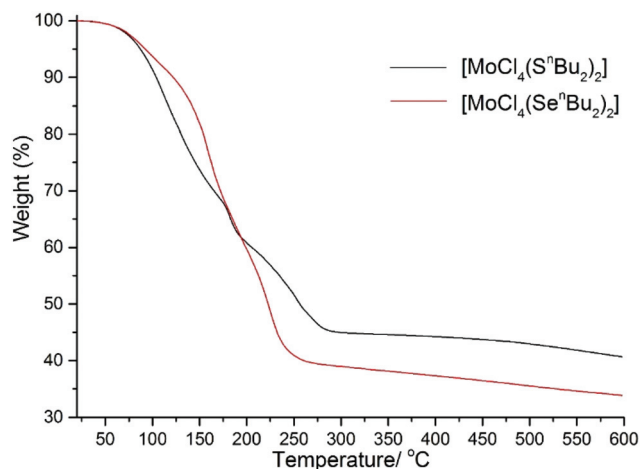


Fig. 5 TGA of $[\text{MoCl}_4(\text{Bu}_2\text{E})_2]$ ($\text{E} = \text{S}$ or Se) obtained under flowing Ar (65 mL min^{-1}) at ambient pressure with a heating rate of $10\text{ }^\circ\text{C min}^{-1}$.

gravimetric analysis (TGA). $[\text{MoCl}_4(\text{Bu}_2\text{Se})_2]$ underwent three mass loss steps at $75\text{--}125\text{ }^\circ\text{C}$, $180\text{--}200\text{ }^\circ\text{C}$ and $200\text{--}280\text{ }^\circ\text{C}$, leaving a residual mass of *ca.* 39.4%, which gradually reduced further as the temperature was raised to $600\text{ }^\circ\text{C}$ (Fig. 5). This suggests a complex decomposition pathway with MoSe_2 (40.7%) or MoCl_4 (38%) as the final residue. $[\text{MoCl}_4(\text{Bu}_2\text{S})_2]$ also shows three mass loss steps at $25\text{--}160\text{ }^\circ\text{C}$, $170\text{--}200\text{ }^\circ\text{C}$ and $210\text{--}300\text{ }^\circ\text{C}$, leaving a residual mass of *ca.* 44.9%, which undergoes a further small mass loss (*ca.* 5%) on heating to $600\text{ }^\circ\text{C}$ (Fig. 5). The residual mass is higher than that expected for MoS_2 (30%), but similar to that expected for MoCl_4 (44.7%); some sublimation of MoCl_4 at high temperature would be expected; we note that $[\text{NbCl}_4(\text{Bu}_2\text{S})_2]$ left NbCl_4 upon heating



in vacuo.³⁶ It is important to note that the TGA experiments are performed at 1 atmosphere pressure under flowing argon, and hence very different from LPCVD conditions. However, TGA does provide a guide to the volatility of the precursors.

CVD test experiments focussed on those complexes bearing ⁿBu substituents since, unlike Me substituents, these tend to be promote elimination of the organic by-products, most likely *via* β -hydride elimination reactions. CVD onto silica tiles using *ca.* 30–70 mg of the $[\text{MoCl}_4(\text{}^n\text{Bu}_2\text{Se})_2]$ precursor resulted in deposition of a reflective golden film on tiles positioned in the hottest part of the furnace at 400–550 °C. Grazing incidence XRD (Fig. 6) showed these films to be the 2H form of MoSe_2 ($P6_3/mmc$). Lattice parameters determined from a Le Bail fit to the pattern are $a = 3.2666(10)$ and $c = 13.170(16)$ Å ($R_{\text{wp}} = 1.8\%$, $R_p = 1.3\%$; see ESI Fig. S2† for fit). The lattice parameters listed in ICSD for this phase are in the ranges $a = 3.288\text{--}3.290$ and $c = 12.90\text{--}12.94$ Å.²⁸ The films are air and moisture stable (over many months), but easily scratched with a metal spatula. With deposition temperatures higher than 550 °C, MoO_2 thin films were recovered, presumably resulting from deposition of

molybdenum metal that oxidised post-deposition. Parkin and co-workers previously deposited films containing a mixture of the 2H- and 3R-type MoSe_2 films using dual source atmospheric pressure CVD,¹⁶ but we are not aware of any other reports of single phase MoSe_2 films from CVD from molybdenum chloride derived precursors.

The grazing incidence diffraction patterns of the MoSe_2 films are dominated by the 002 reflection, suggesting $\langle 00l \rangle$ alignment of the crystallites. These layered materials typically grow as platelet-shaped crystallites with the *ab* crystallographic axes in the plane of the crystallites, so this would suggest platelets lying flat on the substrate. An in-plane XRD measurement (Fig. 6) suppressed the 002 reflection and enhanced the 100 and 110, supporting this conclusion. SEM images (Fig. 7) showed mainly the edges of platelets suggesting they were standing on edge rather than lying flat, but imaging the edge of a cross-section showed a dense, flat film with thickness of around 130 nm underlying these orthogonally oriented platelets. A pole figure measurement (Fig. 8) taken with $2\theta = 13.54^\circ$ (002 reflection) exhibits a single and sharp peak (FWHM $\sim 10^\circ$) with $\alpha = 90^\circ$, consistent with the vast majority of crystallites adopting the $\langle 00l \rangle$ orientation. The pole figure with $2\theta = 56.10^\circ$ (110 reflection), exhibits a ring with α close to zero (the edge of the peak is cut off as the measurement cannot be made right at the substrate plane), also consistent with the $\langle 00l \rangle$ preferred orientation. Unfortunately, the pole figure at the 103 reflection position was featureless. This may suggest stacking errors reducing its intensity.

Energy dispersive X-ray (EDX) analysis data showed significant Si and O from the substrate, in addition to peaks due to Mo and Se, indicating that the films are thin. There was no evidence for any residual Cl in the films. The Mo:Se ratio was quantified as 1:1.9, consistent within error with the formation of MoSe_2 .

The crystallite size in the MoSe_2 film produced at 550 °C was calculated from the grazing incidence XRD data using the Williamson–Hall method. The 6.6(9) nm size that resulted is much smaller than the largest dimensions of the platelets observed in Fig. 7 (*ca.* 100 nm), but the observed preferred orientation shows the bulk of the material to be contained within the dense underlying film. The diffraction crystallite size appears to be dominated by this part of the sample. Using a larger amount of precursor (*ca.* 200 mg) resulted in thicker

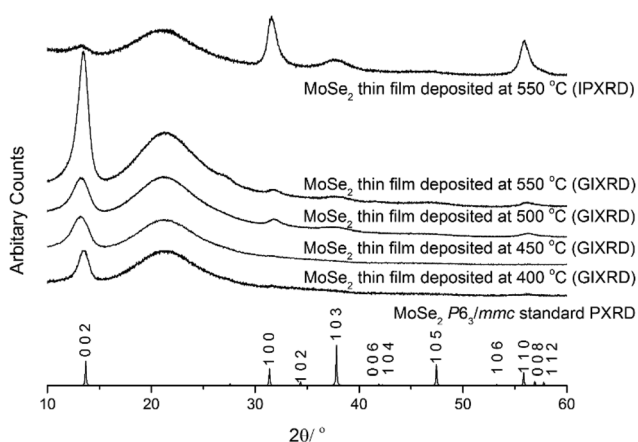


Fig. 6 Grazing incidence XRD (GIXRD, incidence angle = 1°) and in-plane XRD (IPXRD, incidence angle = 0.5°) from MoSe_2 thin films deposited by CVD from $[\text{MoCl}_4(\text{}^n\text{Bu}_2\text{Se})_2]$ at temperatures as shown. The broad feature at $2\theta \sim 22^\circ$ is from the SiO_2 substrate. The calculated pattern for bulk polycrystalline 2H- MoSe_2 ($P6_3/mmc$) with Miller indices is shown for comparison.²⁸

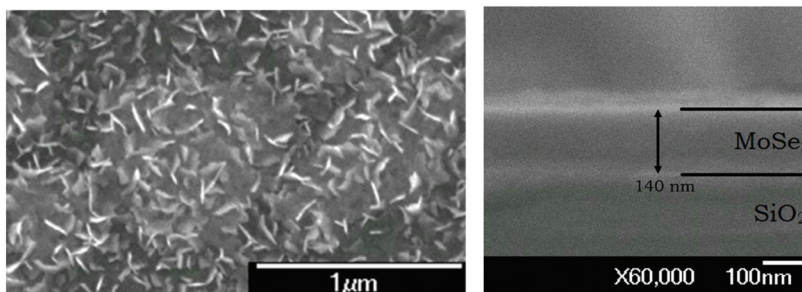


Fig. 7 SEM images of MoSe_2 films produced by CVD using ~ 70 mg $[\text{MoCl}_4(\text{}^n\text{Bu}_2\text{Se})_2]$ at 550 °C. Top view (left) and cross-section (right).



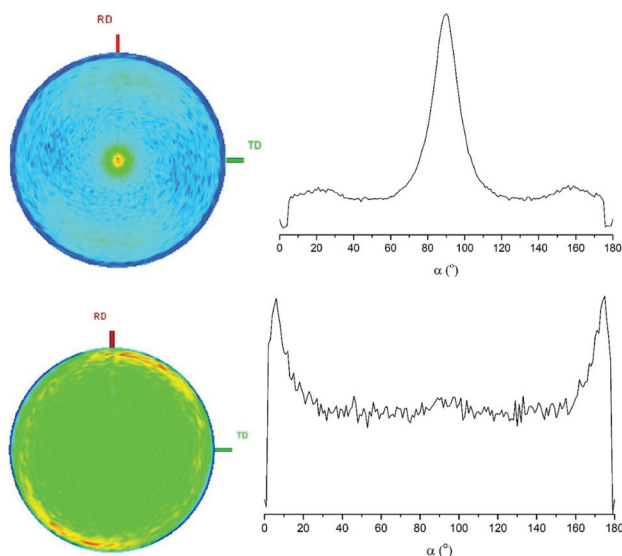


Fig. 8 002 (top) and 110 (bottom) pole figure measurements shown as 2D plots (left) and integrated plots with α (right) for MoSe₂ thin films produced by CVD using ~ 70 mg [MoCl₄(ⁿBu₂Se)₂] at 550 °C.

films, but the optical quality of the films became worse. SEM images (ESI, Fig. S3†) showed that some crystallites continued to grow preferentially and thus the film thickness was less consistent over the area of the film.

The Raman spectrum from the MoSe₂ film was collected using 785 nm excitation and shows three peaks at 140, 241 and 290 cm⁻¹ assigned to the E_{1g}, A_{1g} and E_{2g} vibration modes of 2H-MoSe₂, respectively (Fig. 9),^{16,37–40} whilst peaks at *ca.* 317, 455 and 595 cm⁻¹ can be attributed to contribution from acoustic phonons to the Raman scattering spectrum.⁴¹ The optical spectra (ESI, Fig. S4†) showed the expected excitonic transition features around 700 nm.⁴²

When using similar amounts of [MoCl₄(ⁿBu₂S)₂] (*ca.* 30–70 mg) to that used in most of the MoSe₂ depositions

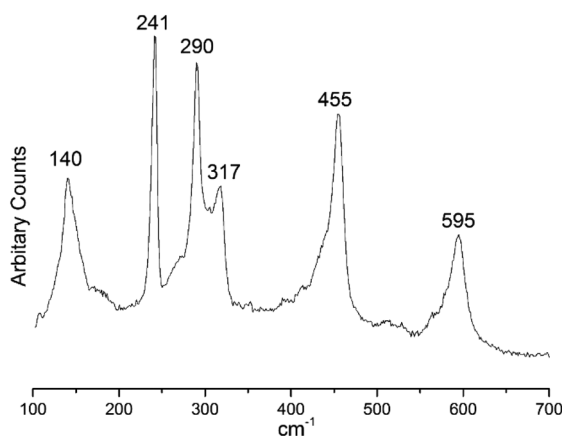


Fig. 9 Raman spectrum of the MoSe₂ film deposited by LPCVD from [MoCl₄(ⁿBu₂Se)₂] at 550 °C.

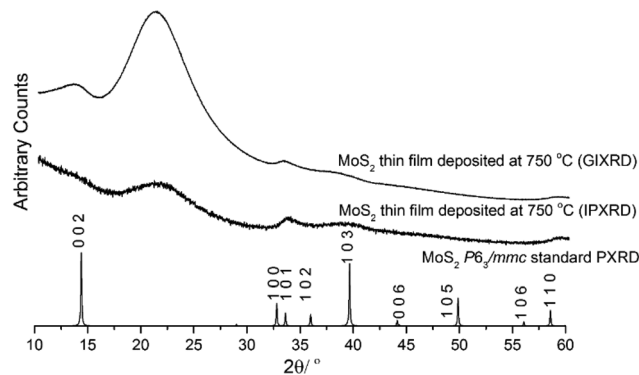


Fig. 10 Grazing incidence (incidence angle = 1°) and in-plane (incidence angle = 0.5°) XRD patterns from the MoSe₂ thin film deposited from [MoCl₄(ⁿBu₂S)₂] at 750 °C, and the simulated XRD pattern from bulk 2H-MoSe₂.²⁸ The broad feature at $2\theta \sim 22^\circ$ is from the SiO₂ substrate.

described above, CVD on SiO₂ substrates at 750 °C resulted in very thin yellow films. These were too thin to generate any X-ray diffraction in grazing incidence or in plane measurement, or to image effectively in the SEM, but did produce optical spectra consistent with MoSe₂⁴² (ESI, Fig. S5†). At lower temperatures no deposition occurred. However, by using larger amounts of [MoCl₄(ⁿBu₂S)₂] (200 mg) at 750 °C, silver films were obtained. Grazing incidence XRD (Fig. 10) showed similar features in these films to those seen for MoSe₂ in Fig. 7, but with much weaker intensities. The pattern was dominated by the interlayer spacing peak (the 002 reflection), which gives no information on the exact phase of MoSe₂ produced. However, the in-plane XRD pattern contained stronger 100, 103 and 110 reflections, confirming that, like MoSe₂, MoSe₂ adopted the 2H structure. The lattice parameters were refined as $a = 3.13(5)$ and $c = 13.7(8)$ Å, although there is significant uncertainty in this result due to the weak, broad peaks. Lattice parameters in ICSD are in the range $a = 3.14$ – 3.16 Å and $c = 12.29$ – 12.53 Å.²⁸ The suppression of the 002 reflection in the in-plane pattern suggests $\langle 00l \rangle$ preferred orientation like that observed in MoSe₂, but unfortunately, the scattering from these films was not strong enough for pole figure measurements. The Raman spectrum recorded from the MoSe₂ film collected using 785 nm excitation showed weak bands at *ca.* 373 and 406 cm⁻¹, assigned as the E_{2g} and A_{1g} vibrational modes in 2H-MoSe₂.^{37,40,41,43–45}

SEM images confirm that the orthogonally oriented crystallites on the MoSe₂ film are very small, but that the underlying dense layer has a similar thickness to that found for MoSe₂ (~ 150 nm, Fig. 11). The weak and broad diffraction data suggest the crystallite sizes may also be smaller in the dense layer. Due to a close overlap between the Mo L_α (2.293 keV) and S K_α (2.307 keV) fluorescence peaks, it was not possible to obtain the Mo : S ratio from the EDX data. However, these did confirm the presence of both elements and there is no evidence for any residual Cl in the films.



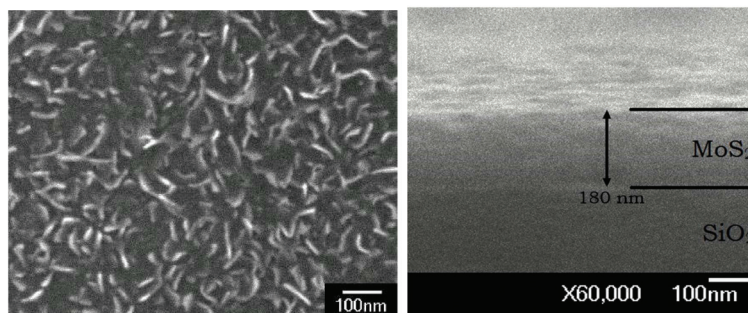


Fig. 11 SEM images of MoS₂ films produced by CVD using ~200 mg [MoCl₄(ⁿBu₂S)₂] at 750 °C.

Conclusions

The synthesis and characterisation of a new series of well-defined, distorted octahedral Mo(IV) chloride complexes with mono- and bi-dentate thioether and selenoether ligands has been demonstrated, with single crystal structures determined for several representative examples. TGA analysis of [MoCl₄(ⁿBu₂S)₂] and [MoCl₄(ⁿBu₂Se)₂] suggested they may be sufficiently volatile to be used as sources of MoS₂ and MoSe₂, respectively. This was borne out by LPCVD experiments which produced silver (2H-MoS₂) and golden yellow (2H-MoSe₂) thin films. The identities of the films, crystallite sizes and morphologies were determined *via* XRD, SEM/EDX, Raman and optical measurements. The successful identification of single source CVD precursors for their deposition is an important development given the very high level of interest in thin films of these materials currently and their range of potential applications.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We thank the EPSRC for support (through EP/K00509X/1, EP/K009877/1 and EP/M508147/1) and the University of Southampton for a VC Scholarship (to Y.-P. C).

References

- C. D. Garner and J. M. Charnock, in *Comprehensive Coordination Chemistry*, ed. G. Wilkinson, R. D. Gillard and J. A. McCleverty, Pergamon, Oxford, 1987, ch. 36.4; E. I. Stiefel, in *Comprehensive Coordination Chemistry*, ed. G. Wilkinson, R. D. Gillard and J. A. McCleverty, Pergamon, Oxford, 1987, ch. 36.5; G. J. Leigh and R. L. Richards, in *Comprehensive Coordination Chemistry*, ed. G. Wilkinson, R. D. Gillard and J. A. McCleverty, Pergamon, Oxford, 1987, ch. 36.2.
- C. G. Young, in *Comprehensive Coordination Chemistry II*, ed. J. A. McCleverty and T. J. Meyer, Elsevier, Oxford, 2004, Vol. 4, Section 4.7.
- E. I. Stiefel, *Prog. Inorg. Chem.*, 1977, **22**, 1.
- M. Hidai and Y. Mizobe, in *Metal Ions in Biological Systems*, ed. A. Sigel and H. Sigel, Marcel Dekker, NY, 2002, vol. 39, p. 121.
- Molybdenum and Tungsten; their role in biological processes*, Metal ions in biological systems, ed. A. Sigel and H. Sigel, Marcel Dekker, NY, 2002, vol. 39.
- M. Chhowalla, Z. Liu and H. Zhang, *Chem. Soc. Rev.*, 2015, **44**, 2584, (Eds. of themed issue on early transition metal dichalcogenides).
- M. Chhowalla, H. S. Shin, G. Eda, L.-J. Li, K. P. Loh and H. Zhang, *Nat. Chem.*, 2013, **5**, 263; J. Liu and X.-W. Liu, *Adv. Mater.*, 2012, **24**, 4097.
- J. R. Brent, N. Savjani and P. O'Brien, *Prog. Mater. Sci.*, 2017, **89**, 411.
- Q. Xiang, J. Yu and M. Jaroniec, *J. Am. Chem. Soc.*, 2012, **134**, 6575; K. Lee, R. Gatensby, N. McEvoy, T. Hallam and G. S. Duesberg, *Adv. Mater.*, 2013, **25**, 6699; Z. Yan, C. Jiang, T. R. Pope, C. F. Tsang, J. L. Stickney, P. Goli, J. Renteria, T. T. Salguero and A. A. Balandin, *J. Appl. Phys.*, 2013, **114**, 20430.
- J. S. Rhyee, J. Kwon, P. Dak, J. H. Kim, S. M. Kim, J. Park, Y. K. Hong, W. G. Song, I. Omkaram, M. A. Alam and S. Kim, *Adv. Mater.*, 2016, **28**, 2316; Y.-H. Chang, W. Zhang, Y. Zhu, Y. Han, J. Pu, J.-K. Chang, W.-T. Hsu, J.-K. Huang, C.-L. Hsu, M.-H. Chiu, T. Takenobu, H. Li, C.-I. Wu, W.-H. Chang, A. T. S. Wee and L.-J. Li, *ACS Nano*, 2014, **8**, 8582; X. Wang, Y. Gong, G. Shi, W. L. Chow, K. Keyshar, G. Ye, R. Vajtai, J. Lou, Z. Liu, E. Ringe, B. K. Tay and P. M. Ajayan, *ACS Nano*, 2014, **8**, 5125.
- X. Lu, M. I. Utama, J. Lin, X. Gong, J. Zhang, Y. Zhao, S. T. Pantelides, J. Wang, Z. Dong, Z. Liu, W. Zhou and Q. Xiong, *Nano Lett.*, 2014, **14**, 2419; A. S. Pawbake, M. S. Pawar, S. R. Jadhkar and D. J. Late, *Nanoscale*, 2016, **8**, 3008.
- J. Chen, X. Zhao, S. J. Tan, H. Xu, B. Wu, B. Liu, D. Fu, W. Fu, D. Geng, Y. Liu, W. Liu, W. Tang, L. Li, W. Zhou, T. C. Sum and K. P. Loh, *J. Am. Chem. Soc.*, 2017, **139**, 1073; Y. H. Lee, X. Q. Zhang, W. Zhang, M. T. Chang, C. T. Lin,



- K. D. Chang, Y. C. Yu, J. T. Wang, C. S. Chang, L. J. Li and T. W. Lin, *Adv. Mater.*, 2012, **24**, 2320.
- 13 K. K. Liu, W. Zhang, Y. H. Lee, Y. C. Lin, M. T. Chang, C. Y. Su, C. S. Chang, H. Li, Y. Shi, H. Zhang, C. S. Lai and L. J. Li, *Nano Lett.*, 2012, **12**, 1538; J. Pu, Y. Yomogida, K.-K. Liu, L.-J. Li, Y. Iwasa and T. Takenobu, *Nano Lett.*, 2012, **12**, 4013.
- 14 Y.-P. Chang, W. Levason and G. Reid, *Dalton Trans.*, 2016, **45**, 18393.
- 15 A. L. Hector, M. Jura, W. Levason, S. D. Reid and G. Reid, *New J. Chem.*, 2009, **33**, 64115; Y.-P. Chang, A. L. Hector, W. Levason and G. Reid, *Dalton Trans.*, 2017, **46**, 9824; S. L. Benjamin, C. H. de Groot, C. Gurnani, A. L. Hector, R. Huang, K. Ignatyev, W. Levason, S. J. Pearce, F. Thomas and G. Reid, *Chem. Mater.*, 2013, **25**, 4719; S. L. Benjamin, Y.-P. Chang, C. Gurnani, A. L. Hector, M. Huggon, W. Levason and G. Reid, *Dalton Trans.*, 2014, **43**, 16640; Y.-P. Chang, A. L. Hector, W. Levason and G. Reid, *Dalton Trans.*, 2017, **46**, 9824.
- 16 N. D. Boscher, C. J. Carmalt, R. G. Palgrave, J. J. Gil-Tomas and I. P. Parkin, *Chem. Vap. Deposition*, 2006, **12**, 692.
- 17 J. Cheon, J. E. Gozum and G. S. Girolami, *Chem. Mater.*, 1997, **9**, 1847.
- 18 A. D. Westland and U. Uzelac, *Can. J. Chem.*, 1970, **40**, 2871; D. Sevdic and L. Fekete, *Inorg. Chim. Acta*, 1982, **57**, 111; W. Levason, C. A. McAuliffe, F. P. McCullough, S. G. Murray and C. A. Rice, *Inorg. Chim. Acta*, 1977, **22**, 227.
- 19 P. Dierkes, G. Frenzen, S. Wocadlo, S. Berger, J. Pebler and K. Dehnicke, *Z. Naturforsch. Teil B*, 1995, **50**, 159.
- 20 L. Favero, F. Marchetti, G. Pampaloni and S. Zacchini, *Dalton Trans.*, 2014, **43**, 495.
- 21 F. R. Hartley, S. G. Murray, W. Levason, H. E. Soutter and C. A. McAuliffe, *Inorg. Chim. Acta*, 1979, **35**, 265.
- 22 D. J. Gulliver, E. G. Hope, W. Levason, S. G. Murray, D. M. Potter and G. L. Marshall, *J. Chem. Soc., Perkin Trans. 2*, 1984, 429.
- 23 G. M. Sheldrick, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2015, **71**, 3.
- 24 *CrysAlis PRO*, Agilent Technologies Ltd, Yarnton, Oxfordshire, England.
- 25 *CrystalClear-SM Expert 2.1 b29*, Rigaku Corporation, Tokyo, Japan, 2012.
- 26 J. R. Dilworth and R. L. Richards, *Inorg. Synth.*, 1990, **28**, 33.
- 27 S. Grazulis, D. Chateigner, R. T. Downs, A. F. Yokochi, M. Quiros, L. Lutterotti, E. Manakova, J. Butkus, P. Moeck and A. Le Bail, *J. Appl. Crystallogr.*, 2009, **42**, 726.
- 28 ICSD: Inorganic Crystal Structure Database (ICSD), Fachinformationszentrum Karlsruhe (FIZ), accessed via the EPSRC funded National Chemical Database Service hosted by the Royal Society of Chemistry.
- 29 F. Stoffelbach, D. Saurens and R. Poli, *Eur. J. Inorg. Chem.*, 2001, 2699.
- 30 A. V. Butcher and J. Chatt, *J. Chem. Soc. A*, 1970, 2652.
- 31 A. Manteghetti, C. Belin, M. Tillard-Charbonnel, J.-L. Pascal, E. Clot and F. Favier, *New J. Chem.*, 1999, **23**, 165.
- 32 M. Bortoluzzi, E. Ferretti, M. Hayatifar, F. Marchetti, G. Pampaloni and S. Zacchini, *Eur. J. Inorg. Chem.*, 2016, 3838.
- 33 B. Modec, J. V. Brenčić and L. Golic, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2000, **56**, 780.
- 34 B. N. Figgis and J. Lewis, *Prog. Inorg. Chem.*, 1964, **6**, 37.
- 35 A. B. P. Lever, *Inorganic electronic spectroscopy*, Elsevier, NY, 2nd edn, 1984.
- 36 Y.-P. Chang, W. Levason, M. E. Light and G. Reid, *Dalton Trans.*, 2016, **45**, 16262.
- 37 S. Sugai and T. Ueda, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1982, **26**, 6554.
- 38 A. Bachmatiuk, R. F. Abelin, H. T. Quang, B. Trzebicka, J. Eckert and M. H. Rummeli, *Nanotechnology*, 2014, **25**, 365603.
- 39 Y.-H. Chang, W. Zhang, Y. Zhu, Y. Han, J. Pu, J.-K. Chang, W.-T. Hsu, J.-K. Huang, C.-L. Hsu, M.-H. Chiu, T. Takenobu, H. Li, C.-I. Wu, W.-H. Chang, A. T. S. Wee and L.-J. Li, *ACS Nano*, 2014, **8**, 8582.
- 40 A. Roy, H. C. Movva, B. Satpati, K. Kim, R. Dey, A. Rai, T. Pramanik, S. Guchhait, E. Tutuc and S. K. Banerjee, *ACS Appl. Mater. Interfaces*, 2016, **8**, 7396.
- 41 P. Soubelet, A. E. Bruchhausen, A. Fainstein, K. Nogajewski and C. Faugeras, *Phys. Rev. B*, 2016, **93**, 155407.
- 42 N. Dong, Y. Li, Y. Feng, S. Zhang, X. Zhang, C. Chang, J. Fan, L. Zhang and J. Wang, *Sci. Rep.*, 2015, **5**, 14646.
- 43 K. K. Liu, W. Zhang, Y. H. Lee, Y. C. Lin, M. T. Chang, C. Y. Su, C. S. Chang, H. Li, Y. Shi, H. Zhang, C. S. Lai and L. J. Li, *Nano Lett.*, 2012, **12**, 1538.
- 44 Y. Shi, W. Zhou, A. Y. Lu, W. Fang, Y. H. Lee, A. L. Hsu, S. M. Kim, K. K. Kim, H. Y. Yang, L. J. Li, J. C. Idrobo and J. Kong, *Nano Lett.*, 2012, **12**, 2784.
- 45 W. Park, J. Baik, T. Y. Kim, K. Cho, W. K. Hong, H. J. Shin and T. Lee, *ACS Nano*, 2014, **8**, 4691.

