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The use of semiconductor photocatalysts to generate electrons with efficient reducing capability for organic photoreduction synthesis and the removal of harmful substances has become a hotspot in the field of green chemistry research. In this work, α -MnO₂ nanocubes and α -MnO₂@MnIn₂S₄ hybrid photocatalysts with a core-shell structure were synthesized successively by a two-step method. XRD and XPS verified the coexistence of the two substances (α -MnO₂ and MnIn₂S₄) in hybrid systems. According to the SEM and TEM characterization, it is clearly seen that MnIn₂S₄ nanosheets grow on α -MnO₂ nanocubes to form a hierarchical structure. Furthermore, HRTEM showed that the interface contact between α -MnO₂ and MnIn₂S₄ resulted in an atomically defined junction. The photocatalytic performance of the composite catalyst was evaluated by reducing 4-nitroaniline to 4-phenylenediamine and Cr(VI) to Cr(III), respectively. The results show that the catalytic activity of the composite material is effectively improved compared to that of the single components. The Z-scheme electron transport mechanism was proved by ultraviolet-visible diffuse reflectance spectroscopy, valence band XPS, energy band structure calculation and active species detection experiments. The constructed Z-scheme hierarchical α -MnO₂@MnIn₂S₄ system with an atomically defined junction can improve the redox performance of semiconductors for organic synthesis and environmental remediation.

Received 11th October 2020
Accepted 3rd December 2020

DOI: 10.1039/d0na00848f

rsc.li/nanoscale-advances

1 Introduction

Recently, with the development of industrialization, human society's demand for industrial chemicals has been increasing, and the issue of environmental pollution has also been paid more and more attention to.^{1–3} 4-Phenylenediamine (4-PDA) is a widely used industrial chemical raw material. It can be used in the preparation of azo dyes, as well as in the preparation of developers, rubber antioxidants, and copper-iron indicators.^{4–8} The industrial preparation of 4-PDA is mainly carried out by using 4-nitroaniline (4-NA) as a raw material and adopting industrial hydrogenation.^{9,10} This method is restricted by strict reaction conditions and complicated procedures. In recent years, some research studies involving photocatalytic and electrocatalytic methods to prepare 4-PDA have attracted more and more attention.^{11,12} Among these methods, highly selective photoreduction of 4-NA to 4-PDA in water systems has become a powerful technology. Moreover, dealing with highly toxic hexavalent chromium is a difficult task in the treatment of world environmental pollution.^{13,14} Based on trivalent

chromium's low toxicity and easy flocculation into Cr(OH)₃ in an alkaline environment, reducing Cr(VI) to trivalent chromium has become a key procedure for environmental remediation.^{15,16} Among the currently developed reduction methods, photocatalytic reduction has shown potential application prospects compared with high-cost chemical reduction methods¹⁷ and biological reduction methods¹⁸ with high environmental requirements. Therefore, visible-light-driven semiconductor photocatalysis has received extensive attention in these fields because of its low cost, environmental friendliness and unique reaction mechanisms and easy availability of materials.^{19–22} Importantly, photogenerated electrons generated by semiconductors with a deeper conduction band top have extremely strong reducibility and have attracted the attention of researchers.

Manganese-based materials (MnO₂, Mn₃O₄, MnFeO₄, MnSiO₃, and metal-Mn LDH) are widely used in photocatalytic environmental remediation, supercapacitors, lithium-sulfur batteries, and guided cancer treatments due to their low-cost, easy availability, environmental friendliness, high theoretical capacitance, and narrow band gap.^{23–27} The visible light absorption capability of manganese oxide among them is high due to its narrow band gap which makes it a promising candidate for photocatalysis. Gaigneaux *et al.* reported a plasma chemical reduction method to prepare three kinds of α -, γ - and δ -MnO₂. Tartrazin yellow was selected as the model pollutant

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0na00848f

for photocatalytic activity evaluation, and the results showed that α -MnO₂ has achieved the best overall photocatalytic performance.²⁸ On the other hand, the weaker photoreduction ability caused by the higher conduction band potential and the high recombination efficiency of photogenerated carriers caused by the narrow band gap limit the in-depth application of MnO₂ in the photocatalytic industry. Fortunately, the preparation of composite nanomaterials and the introduction of the concept of heterojunction catalysis greatly improve this defect. Leung *et al.*²⁹ *in situ* synthesized a TiO₂/MnO₂ heterojunction catalyst by a hydrothermal method and used the catalyst for vacuum ultraviolet (VUV) photocatalytic oxidation of toluene degradation. The good contact of the heterojunction structure of TiO₂/MnO₂ effectively inhibited the recombination of photo-excited electron-hole pairs and improved the catalytic activity. Qiu *et al.*³⁰ built a Z-scheme mesoporous α -MnO₂/Mn₃O₄ heterojunction to promote its photocatalytic oxidation activity by using the characteristics of Mn having multivalent states. It is proposed that the heterojunction of multivalent manganese oxide can efficiently prevent the recombination of photo-generated carriers.

Furthermore, MnIn₂S₄, which belongs to the AB₂X₄ (A = Cu, Cd, and Mn; B = Al, Ga, and In; X = S, Se, and Te) family,^{31–33} has also received considerable attention in the field of photocatalysis due to its suitable band gap and simple preparation. Chen *et al.*³⁴ synthesized a direct Z-scheme 2D/2D MnIn₂S₄/g-C₃N₄ structure, which has high photocatalytic activity for the treatment of pharmaceutical wastewater and hydrogen evolution.

In this work, α -MnO₂ nanocubes were obtained through the guiding effect of surfactant and annealing treatment, and then a multi-stage α -MnO₂@MnIn₂S₄ composite photocatalyst was synthesized *in situ* by a solvothermal reaction after adsorption of Mn²⁺ and In²⁺. SEM and TEM showed that the synthesized catalyst was MnIn₂S₄ nanosheets coated on α -MnO₂ nanocubes making an α -MnO₂@MnIn₂S₄ core-shell structure. Compared with the single materials, the composite catalyst exhibits excellent catalytic activity and structural stability in reducing 4-NA and Cr(VI).

2 Experimental

2.1 Materials

All reagents were used without any further purification in this research. Manganese sulfate monohydrate (MnSO₄·H₂O, AR), cyclohexane (C₆H₁₂), manganese chloride tetrahydrate (MnCl₂·4H₂O, AR), indium chloride (InCl₃·4H₂O, AR) and ammonium formate (CH₅NO₂) were bought from Aladdin Industrial Corporation. Potassium dichromate (KCr₂O₇), potassium permanganate (KMnO₄), cetyltrimethylammonium bromide (CTAB), ammonium bicarbonate (NH₄HCO₃), and thioacetamide (TAA) were purchased from Sinopharm Group Chemical Reagent Co., Ltd.

2.2 Catalyst synthesis

2.2.1 Preparation of α -MnO₂ nanocubes. The α -MnO₂ nanocubes were prepared by calcining the δ -MnO₂ precursor,

where the δ -MnO₂ precursor comes from the etching of MnCO₃ nanocubes using KMnO₄. Typically,^{30,35} 4 g of CTAB was added to 100 mL of cyclohexane, and then 5 mL of *n*-butanol and 0.8 M NH₄HCO₃ were added. After stirring until the solution became clear, 5 mL of 0.4 M MnSO₄ solution was added dropwise. The conversion of the solution from a clear to a milky white suspension indicated the formation of MnCO₃ nanocubes. After centrifugal drying and washing with water and ethanol, 1 g of MnCO₃ was dispersed in a certain amount of 0.032 M KMnO₄ solution with continuous magnetic stirring. The reaction process is as follow (1):



After filtering and separation of the above suspension, it was dried at 60 °C for 8 hours to obtain a δ -MnO₂ precursor. Furthermore, the obtained δ -MnO₂ precursor was heat-treated in an air atmosphere at 500 °C for 2 h at a heating rate of 5 °C min⁻¹ to obtain α -MnO₂ nanocubes.

2.2.2 Preparation of α -MnO₂@MnIn₂S₄. The α -MnO₂@MnIn₂S₄ composite photocatalyst was prepared by a solvothermal method. A certain amount of α -MnO₂ was added to 60 mL of ethylene glycol and sonicated for 60 min. Immediately, 1 mmol MnCl₂, 2 mmol InCl₃, and 4 mmol TAA were added to the above suspension in sequence and kept under stirring for 1 h. The reaction precursor solution was transferred into a Teflon-lined stainless steel autoclave with a capacity of 100 mL, which was heated and maintained at 180 °C for 24 h. The α -MnO₂@MnIn₂S₄ composite photocatalyst was obtained by natural cooling to room temperature, washing with deionized water and ethanol, centrifugation, and drying at 60 °C overnight. According to the added amount of α -MnO₂ (0.25 mmol, 0.5 mmol, 0.75 mmol, and 1 mmol, respectively), the as-prepared photocatalysts were named M-MIS-*x* (*x* = 1, 2, 3, and 4). For comparison, the pure MnIn₂S₄ (MIS) sample can be synthesized without adding any α -MnO₂ nanocubes by the same fabrication methods.

2.3 Characterization of the catalyst

The morphology of the fabricated samples was demonstrated by using an electron microscope (SEM and TEM, JEM-2100). X-ray diffraction (XRD, D8 Advance X-ray diffractometer, Cu K, λ = 1.5406 Å, tube voltage 40 kV, tube current 40 mA) was used to characterize the crystallinity and structure of the catalyst. The surface element composition and valence state of the sample were determined by X-ray photoelectron spectroscopy. The ultraviolet-visible diffuse reflection spectrum (DRS) was used to show the light absorption capacity of the catalyst using BaSO₄ as the substrate.

2.4 Evaluation of photocatalytic performance

The photocatalytic activity of the catalyst was evaluated by reducing 4-nitroaniline (4-NA) and Cr(VI). In a typical catalytic process, 50 mg of the sample was dispersed in 50 mL 4-NA (10 mg L⁻¹) ultrasonically, and ammonium formate (150 mg) was added as a hole trapping agent. Before the irradiation, the



adsorption-desorption equilibrium was achieved in the dark and in a nitrogen environment for 1 h. At regular intervals, 3 mL of the photocatalytic reaction solution was taken out. The catalyst in the solution was removed with centrifugation and a 0.22 μm drainage filtration membrane. The product detection by using a UV-Vis spectrophotometer shows the characteristic absorption peak of 4-NA at 381.7 nm, while 293.7 nm and 238.1 nm belong to the characteristic absorption peaks of 4-phenylenediamine (4-PDA). In addition, for the photocatalytic reduction of Cr(VI), 50 mg of the sample was dispersed in 50 mL of 50 mg L⁻¹ Cr(VI) solution and adsorbed in a dark room for 1 h. Under visible light irradiation, 1 mL of the solution was taken at regular intervals of 20 min and centrifuged to remove the catalyst. The concentration of Cr(VI) was detected at 540 nm by the normal diphenylcarbazide (DPC) method.

2.5 Electrochemical performance test

The transient photocurrent and electrochemical impedance spectra (EIS) were obtained on a CHI 760E electrochemical workstation. The standard three-electrode system was composed of the photocatalyst (0.2 mg sample loaded on a 1.0 cm² conductive glass area), a platinum wire, and a saturated Ag/AgCl electrode, which were used as the working electrode, counter electrode, and reference electrode in the test system. The transient photocurrent test of the sample was carried out in 0.5 M Na₂SO₄ aqueous solution with a 300 W xenon lamp with a 420 nm cutoff filter. The electrochemical impedance spectroscopy test of the sample was carried out in 0.1 M KCl solution containing 2.5 mM K₃[Fe(CN)₆]/K₄[Fe(CN)₆] as the electrolyte.

3 Results and discussion

In this work, we constructed a heterojunction photocatalyst by growing MnIn₂S₄ nanosheets *in situ* on the surface of α -MnO₂ nanocubes to efficiently reduce 4-NA and Cr(VI), constructing a core-shell-like α -MnO₂@MnIn₂S₄ hierarchical heterostructure as illustrated schematically in Fig. 1. Firstly, MnCO₃ nanocubes were synthesized by room temperature solvent orientation. After an oxidation-reduction reaction with KMnO₄ and annealing treatment, the α -MnO₂ nanocubes are obtained. Finally, MnIn₂S₄ nanosheets were grown *in situ* on the surface of α -MnO₂ nanocubes by a solvothermal reaction. It should be noted that Mn²⁺ will be preferentially adsorbed to the α -MnO₂ surface due to the negative charge of the α -MnO₂ surface³⁶ and the Fajans rules during the stirring process before the solvothermal reaction,³⁷ which is beneficial in the *in situ* growth of

MnIn₂S₄. The photocatalytic functions of α -MnO₂ and MnIn₂S₄ are concentrated in a hierarchical core-shell structure, which facilitates the separation and migration of photogenerated electrons and holes, and provides a large number of active sites for the reduction reaction.

3.1 Structure and morphology

The powder X-ray diffraction (XRD) technology was used to investigate the crystal structure and phase purity of pure MnIn₂S₄ nanosheets, α -MnO₂ nanocubes, and all M-MIS nanocomposites with different α -MnO₂ contents. Fig. S1† displays the XRD pattern of the as-prepared MnCO₃ precursor in which all the characteristic crystal peaks are consistent with the cubic phase (JCPDS no: 44-1472).³⁸ As shown in Fig. S2,† the pure MnO₂ sample shows diffraction peaks at 12.8°, 18.1°, 28.8°, 36.7°, 37.5°, and 49.9°, which can be assigned to the (110), (200), (310), (400), (211) and (411) crystallographic planes of cubic α -MnO₂ (JCPDS no: 44-0141).³⁹ Along with the coupling of MnIn₂S₄ nanosheets and α -MnO₂ to form the α -MnO₂@MnIn₂S₄ nanocomposite catalyst, the M-MIS-*x* series samples show diffraction peaks at 14.3°, 27.6°, 33.4°, and 48°, respectively, which correspond to the (111), (311), (400) and (440) crystal planes of cubic phase (JCPDS no: 65-7474) MnIn₂S₄.³⁴ It is worth noting that there is no impurity diffraction characteristic peak of In₂S₃ or MnS in the pure MnIn₂S₄ and M-MIS-*x* composite catalysts. The hierarchical composite catalyst assembled using α -MnO₂ and MnIn₂S₄ is beneficial to the carrier transport between the interfaces and improves the photocatalytic activity.^{40,41}

The morphology of the as-prepared materials was characterized by SEM and TEM. Fig. 2a-c shows the morphology of MnCO₃ as the precursor of α -MnO₂. The size of the cube with a smooth surface is about 200 nm. As shown in Fig. S3,† after the reaction of MnCO₃ with KMnO₄, the morphology of the nanocube is still maintained, and “small thorns” of nanosheets are grown on the smooth surface. After the annealing treatment, the surface of the δ -MnO₂ nanocube becomes rough due to the etching of the ions and voids are generated inside, which will facilitate the adsorption of Mn²⁺ and In³⁺ to form a heterojunction *in situ* (Fig. 2d-f). The HRTEM image of a single nanocube (Fig. 2i-II) exhibits the lattice lines with a spacing of about 0.692 nm and 0.245 nm, which are in good agreement with the (110) and (400) crystal plane distances of α -MnO₂,^{42,43} respectively, and the angle between the two crystal planes is 45°. MnIn₂S₄ was crystallized in three-dimensional nanoflakes as shown in Fig. 2g and h, and this two-dimensional structure can provide more catalytic sites. Similarly, the lattice fringes attributed to the (222) and (400) crystal planes of MnIn₂S₄ are 0.309 nm and 0.268 nm, respectively (Fig. 2i-I).³⁴ Next, the thin MnIn₂S₄ nanosheets were grown on the surface of the α -MnO₂ cubes by a solvothermal method. After the reaction, MnIn₂S₄ nanosheets uniformly covered the surface of the α -MnO₂ cube and had a good overall 3D morphology (Fig. 2j). The layered cubic core-shell structure of α -MnO₂@MnIn₂S₄ is shown by the TEM images (Fig. 2k). No visible interlayer gaps were found between the MnIn₂S₄ shell and the α -MnO₂ core (Fig. S4†),

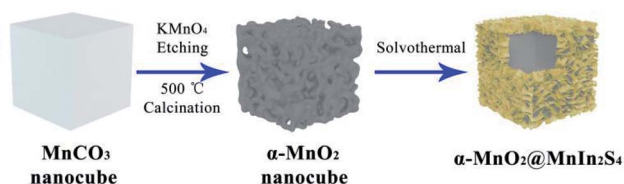


Fig. 1 Schematic illustration of the synthetic process of hierarchical α -MnO₂@MnIn₂S₄.



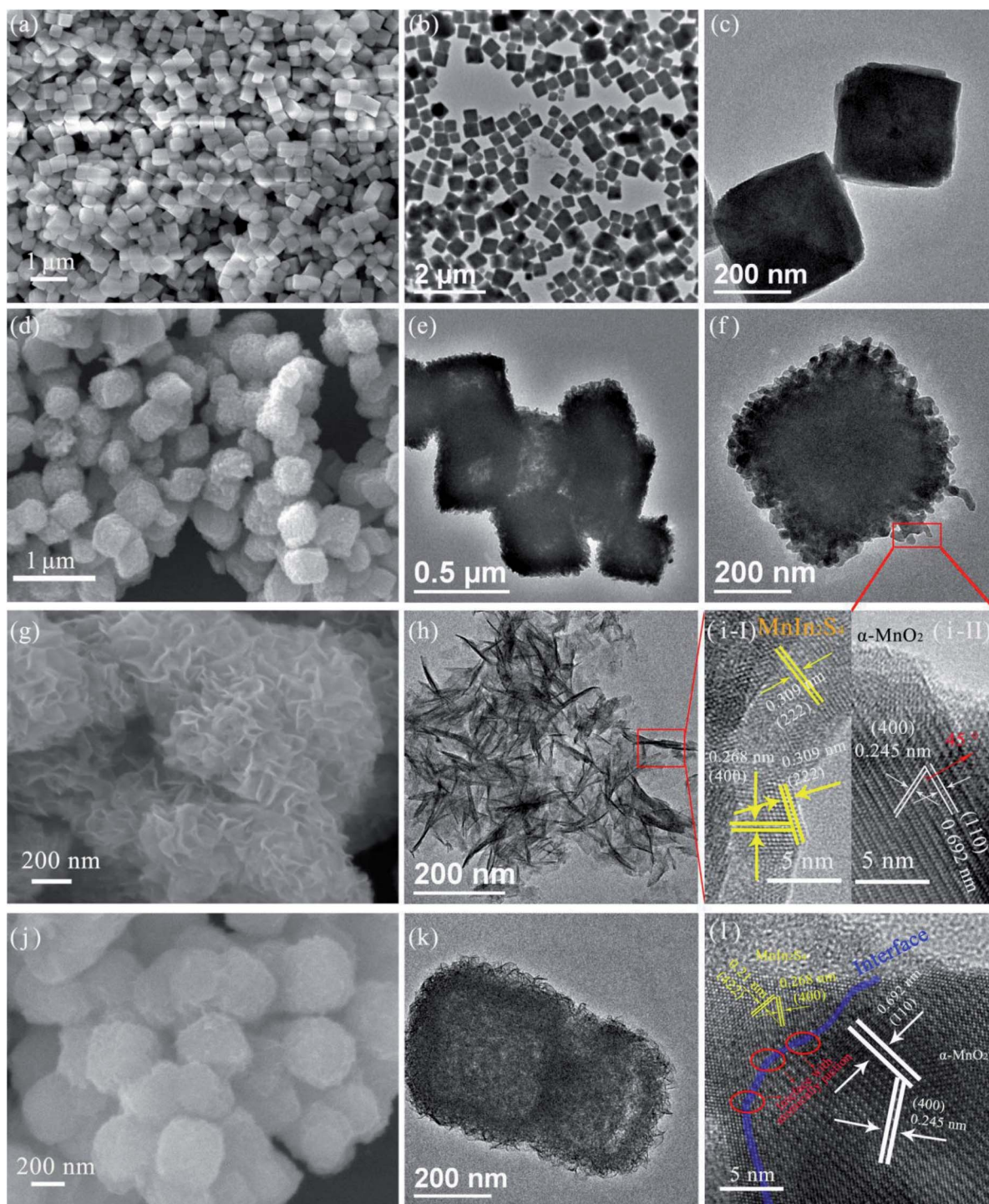


Fig. 2 SEM and TEM images of (a–c) MnCO_3 nanocubes, (d–f) $\alpha\text{-MnO}_2$ nanocubes, (g and h) MnIn_2S_4 nanosheets, and (j and k) M–MIS-2, and HRTEM images of (i) $\alpha\text{-MnO}_2$ (II) and MnIn_2S_4 (I) and (l) M–MIS-2.

indicating close contact between the two components. The HRTEM image shows the heterojunction between layered MnIn_2S_4 and $\alpha\text{-MnO}_2$ (Fig. 2l), in which lattice fringes can be assigned to cubic $\alpha\text{-MnO}_2$ and hexagonal MnIn_2S_4 crystal planes can be seen. In particular, the red elliptical area in Fig. 2l clearly shows the interface with an atomic junction between the crystal planes

of the two materials. These results provide clear evidence of the formation of an $\alpha\text{-MnO}_2@\text{MnIn}_2\text{S}_4$ core-shell structure.

Besides, the SEM elemental mapping of M–MIS-2 is shown in Fig. 3a. It can be seen from the images that Mn, O, In, and S elements all appear on the hierarchical nanocubes, which verifies the successful recombination of MnIn_2S_4 and $\alpha\text{-MnO}_2$ from the side. In order to clearly understand the combination of

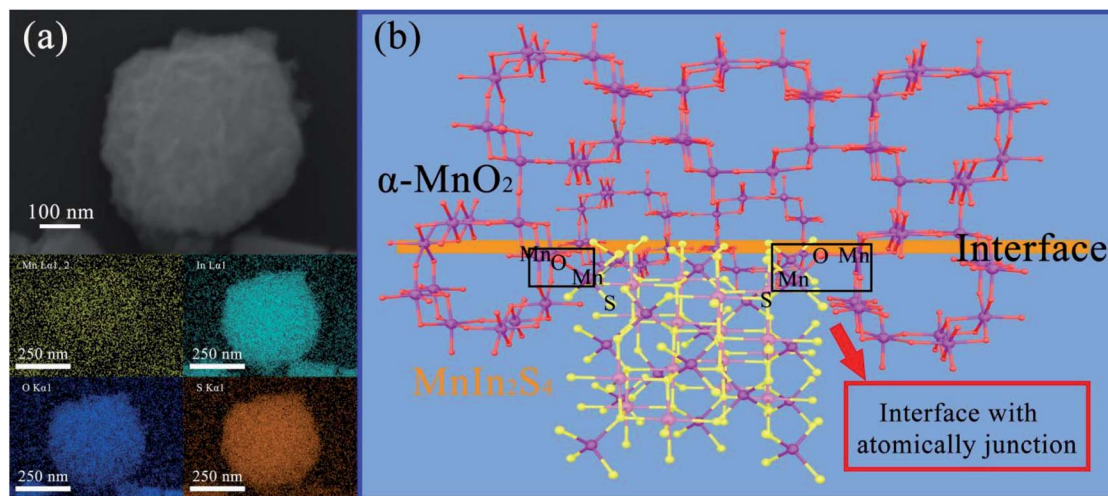


Fig. 3 (a) SEM elemental mapping of M-MIS-2 and (b) schematic illustration of the interface between α -MnO₂ and MnIn₂S₄.

the two components at the interface, a schematic illustration of the interface and atomic-junction combination of α -MnO₂ and MnIn₂S₄ is shown based on HR-TEM and SEM elemental mapping (Fig. 3b). The negative charge oxygen atoms in α -MnO₂ crystals are exposed on the surface.^{44,45} Mn²⁺ and In³⁺ exist in the suspension system at the same time, and Mn²⁺ will be preferentially anchored on the surface of α -MnO₂ according to the Fajans rules.³⁷ These atomically-defined-junction and close interface contact facilitate the transport and separation of photo-generated electrons and improve the catalytic activity of the photocatalyst.

3.2 Surface element analysis

The surface elements and valence states of the as-prepared photocatalyst were characterized by XPS. Herein, the interaction mechanism between α -MnO₂ and MnIn₂S₄ can be further confirmed by Mn 2p_{3/2}, 2p_{1/2}, and S 2p XPS analysis.⁴⁶ Fig. S5† shows the full-scale XPS spectra for α -MnO₂@MnIn₂S₄ (M-MIS-2) and single α -MnO₂ and MnIn₂S₄. Among them, the M-MIS-2 XPS spectrum contains Mn, In, O, and S elements, which confirms the coexistence of α -MnO₂ and MnIn₂S₄ in the composite materials. This result is completely consistent with SEM elemental mapping (Fig. 3a). The Mn 2p spectra of M-MIS-2 are displayed in Fig. 4a and show the characteristic peaks assigned to Mn²⁺ (641.48 eV and 652.32 eV) and Mn⁴⁺ (643.92 eV and 653.71 eV).^{47,48} Compared with composite materials, the position of the characteristic peak of Mn 2p_{3/2} and Mn 2p_{1/2} in single α -MnO₂ and MnIn₂S₄ is shifted, which is closely related to the heterojunction formation by the two materials.⁴⁹ It can be seen from Fig. 4b that the O 1s high-resolution spectra of α -MnO₂ and M-MIS-2 samples show the characteristic peak at 529.43 eV, which belongs to Mn–O–Mn.⁵⁰ Moreover, the characteristic peaks located at 531.37 eV and 530.31 eV in α -MnO₂ and M-MIS-2 can be assigned to Mn–O–H, respectively.⁵¹ The Mn–O–H shift to low binding energy is because the chemical state of Mn on the surface of α -MnO₂ is 4+, which is higher than that (2+) of Mn in the MnIn₂S₄ on the surface of M-MIS-2. As shown in Fig. 4c, the binding energy of S 2p in MnIn₂S₄

was detected at 161.30 eV and 162.37 eV, respectively. Correspondingly, the characteristic peak of the composite catalyst (M-MIS-2) has a slight shift towards high binding energy (161.25 eV and 162.25 eV, respectively).⁵² The In 3d spectra of MnIn₂S₄ and M-MIS-2 are illustrated in Fig. 4d. The binding energies of the doublets are 3d_{5/2} (444.88 eV and 444.77 eV) and 3d_{3/2} (452.50 eV and 452.37 eV), respectively. As mentioned earlier, the shift in binding energy of the heterojunction sample indicates a strong interaction between MnIn₂S₄ nanosheets and α -MnO₂ nanocubes, which enhances the migration efficiency of photogenerated carriers, thereby improving the catalytic activity of the catalyst.

3.3 Optical properties, band structure and carrier separation efficiency

Many factors affect the activity of photocatalysts, such as the band gap, crystal structure, crystal plane orientation and optical properties.^{53,54} Among them, optical properties are the most important characteristics that determine the application of photocatalytic materials in the field of photocatalysis. Fig. 5a shows the UV-Vis diffuse reflectance spectra of M-MIS with different molar ratios and single materials. Compared with other materials, α -MnO₂ has higher band-edge absorption in the visible and near-infrared regions, indicating that it is a narrow band gap semiconductor. The absorption band edge of MnIn₂S₄ is about 560 nm. With the increase of the content of α -MnO₂ in the α -MnO₂@MnIn₂S₄ heterostructure, the absorption of the composite materials in visible light gradually increases, which is beneficial for the easy excitation of the semiconductor under visible-light and generates more carriers. Furthermore, the band gap width of α -MnO₂ and MnIn₂S₄ calculated using the Kubelka–Munk formula is determined to be 1.53 eV and 2.23 eV, respectively, which are consistent with those in the reported literature.^{30,34} Similarly, the positions of the bottom of the valence band (VB) of α -MnO₂ and MnIn₂S₄ are determined from the VB-XPS analysis and it was found to be 2.61 eV and 1.42 eV, respectively (Fig. 5c). Furthermore, the CB potential of α -MnO₂ and MnIn₂S₄ at the point of zero charges was theoretically calculated from eqn (2):⁵⁵



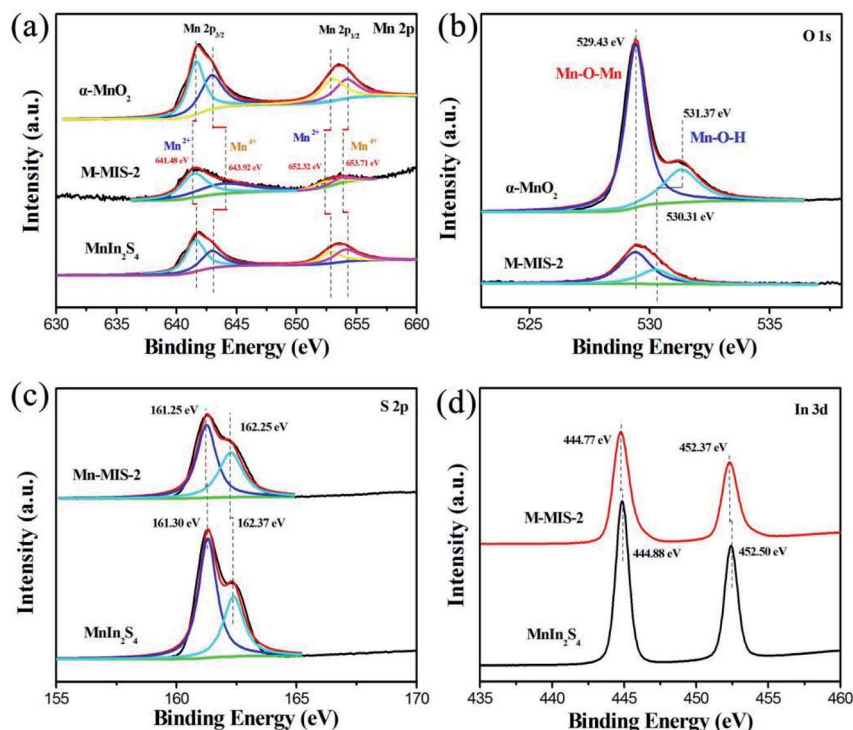


Fig. 4 High-resolution (a) Mn 2p, (b) O 1s, (c) S 2p and (d) In 3d XPS spectrum for the α -MnO₂@MnIn₂S₄ (M-MIS-2), and pure α -MnO₂, and MnIn₂S₄.

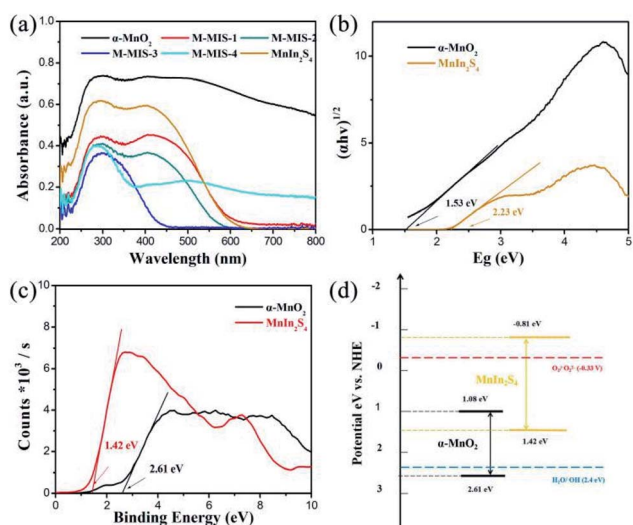


Fig. 5 (a) UV-vis diffuse reflectance spectra of different samples, (b) calculated band gap patterns based on UV-vis diffuse reflectance spectra of α -MnO₂ and MnIn₂S₄, (c) XPS valence band spectrum and (d) energy band structure of α -MnO₂ and MnIn₂S₄.

$$E_g = |E_{CB} - E_{VB}| \quad (2)$$

where E_g is the band gap energy. The calculated E_{CB} values of α -MnO₂ and MnIn₂S₄ are 1.08 eV and −0.81 eV, respectively (Fig. 5d).

Moreover, the migration rate of photogenerated carriers is also an important factor of the photocatalytic activity of the

catalyst. To further compare the migration and separation efficiency of photogenerated electrons-holes of the different materials, the transient photocurrent responses were measured during several on-off intermittent irradiation cycles.⁵⁶ As shown in Fig. 6a, the M-MIS-2 exhibited a much higher photocurrent response than other single photocatalysts (α -MnO₂ and MnIn₂S₄), indicating that the composite material has a higher charge generation and transfer efficiency. This carrier migration in the composite catalyst leads to an effective $e_{CB}^- - h_{VB}^+$ separation and a lower recombination rate in the M-MIS-2 sample, ultimately leading to enhanced photocatalytic performance.⁵⁷ Furthermore, EIS measurements are used to gain a deeper understanding of the photogenerated charge carrier transfer process and recombination behavior. Fig. 6b displays a typical Nyquist plot of different samples under dark conditions. It can be seen that the α -MnO₂ electrode forms a partial

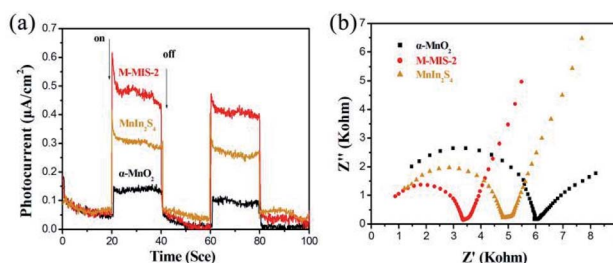


Fig. 6 (a) Photocurrent density and (b) EIS Nyquist plots of α -MnO₂, MnIn₂S₄ and M-MIS-2.

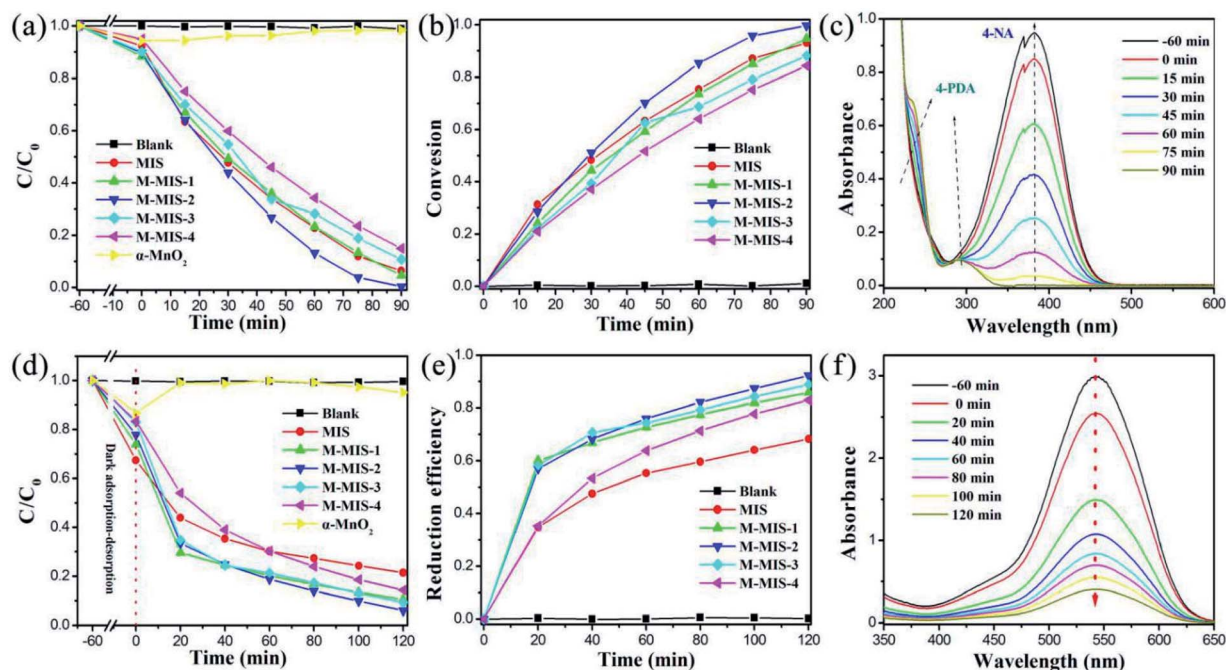


Fig. 7 (a and d) Time-course variation of the 4-NA and Cr(vi) solution under visible light irradiation with photocatalysts, (b) the conversion of selective photo-reduction from 4-NA to 4-PDA in the presence of different catalysts, (c) UV-vis absorption spectra of 4-NA at different irradiation times using M-MIS-2 with the addition of ammonium formate and N₂, (e) Cr(vi) reduction efficiency of 50 mg of different catalysts under visible light for 120 min with different initial Cr(vi) concentrations, and (f) UV-vis absorption spectra of aqueous Cr(vi) (50 mg L⁻¹) treated with M-MIS-2.

semicircle with a diameter larger than that of the other samples. The semicircle in the high-frequency region reflects the charge transfer resistance and constant phase elements at the photocatalyst/electrolyte interface. The radius of the semicircle of M-MIS-2 is the smallest at any frequency, indicating that the charge transfer resistance is smaller, which leads to a better electronic conductivity and charge separation efficiency.⁵⁸ Furthermore, the high electron-hole separation efficiency in the M-MIS composite photocatalyst was analyzed by steady-state photoluminescence (PL) spectroscopy. As can be seen from Fig. S6,[†] M-MIS-2 exhibits a lower PL emission intensity compared to α -MnO₂ and MnIn₂S₄, which means that there is a faster charge transfer rate at the contact interface of the composite catalyst, and it has a better separation efficiency than a single component.

3.4 Photocatalytic performance

The photocatalytic performance of the synthesized α -MnO₂@-MnIn₂S₄ heterojunctions was characterized by the photoreduction from 4-NA to 4-PDA with a visible light source ($\lambda > 420$ nm). Obviously, pure α -MnO₂ has no activity towards 4-PDA evolution from 4-NA, and the dark adsorption on the surface of α -MnO₂ gradually desorbs 4-NA molecules during the light process. Different amounts of α -MnO₂ were compounded with MnIn₂S₄, and the photocatalytic performance of M-MIS on 4-NA was greatly improved (Fig. 7a). In particular, M-MIS-2 has the best reduction efficiency, with a conversion efficiency of 99.6% for photoreduction of 4-NA to 4-PDA within 90 min (Fig. 7b). As shown in Fig. 7c, with the gradual decrease in the intensity of

the main absorption peak of 4-nitroaniline at 381.7 nm, at the same time the characteristic absorption peak intensities of 4-phenylenediamine at 293.7 nm and 238.1 nm consecutively increase. As shown in Fig. S7,[†] the color of the reaction solution gradually changed from light yellow to colorless after 90 minutes of visible light irradiation, indicating the highly selective conversion efficiency of 4-NA to 4-PDA.

Similarly, the M-MIS heterojunction catalyst also shows higher photocatalytic efficiency towards the photoreduction of Cr(vi). As shown in Fig. 7d, when single α -MnO₂ is used as a catalyst, the content of Cr(vi) changes little, which is attributed to the adsorption effect of the high specific surface area of the mesoporous α -MnO₂. The composite catalysts showed higher catalytic activity than pure MIS in reducing Cr(vi). M-MIS-2 showed the best photocatalytic reduction efficiency among all the catalysts, reaching 94% (Fig. 7e and f). Therefore, the above results indicate that the M-MIS-2 sample exhibits excellent visible-light-driven photocatalytic reduction performance, indicating that the α -MnO₂@MnIn₂S₄ heterojunction is an attractive material for environmental remediation and organic conversion.

Besides the photocatalytic activity of the target material, its catalytic stability is also an important factor in measuring the application prospects of the catalyst. It can be seen in Fig. 8a that the photoreduction reactions of 4-NA over M-MIS-2 can still maintain the conversion rate above 92%. The slight decrease may be due to the loss of the catalyst. The sample after four reduction experiments was characterized by XRD (Fig. 8b), and the results show that its crystal shape did not change



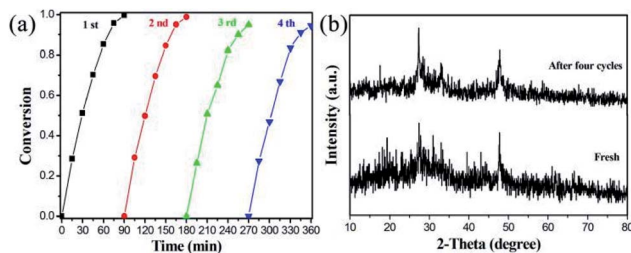


Fig. 8 (a) Cycles of photocatalytic reduction of 4-NA by M-MIS-2; (b) XRD patterns of M-MIS-2 after four cycle experiments.

significantly and had good stability. In addition, the SEM images of the M-MIS-2 catalyst before and after the reaction are shown in Fig. S8.† The morphology of the sample before and after the reaction is still maintained.

3.5 Possible photocatalytic mechanism

As we all know, the active species (such as superoxide radicals ($\cdot\text{O}_2^-$), holes (h^+) and hydroxyl radicals ($\cdot\text{OH}$)) in the photocatalytic reaction system play an important role. In order to detect these species, we use AgNO_3 , isopropanol (IPA), and disodium ethylenediaminetetraacetic acid (EDTA-2Na) with a concentration of 1 mM as the e^- , $\cdot\text{OH}$ and h^+ capture agents.^{14,59,60} Obviously, it can be observed from Fig. 9 that with the addition of AgNO_3 , the reduction efficiency of the catalyst towards $\text{Cr}(\text{vi})$ significantly decreases, indicating that e^- is an important active species participating in the reduction reaction. After IPA was introduced into the photoreduction system, the catalytic efficiency decreased slightly, which may be because some of the $\cdot\text{OH}$ was produced by e^- and competed with the $\text{Cr}(\text{vi})$ reduction process. When the EDTA-2Na is added, it can be found that the catalytic reduction efficiency was significantly improved, which is due to the addition of the hole trapping agent and it inhibits the recombination of photogenerated electrons-holes, consequently, improving the rate and

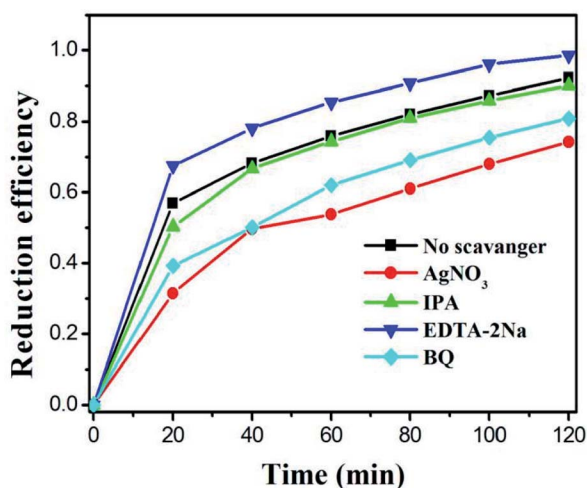


Fig. 9 Effects of different scavengers on $\text{Cr}(\text{vi})$ reduction in the presence of M-MIS-2.

utilization of electron migration to the catalyst surface. After adding BQ in the system, the photocatalytic reduction efficiency also decreased, which is due to the generation of $\cdot\text{O}_2^-$ from photogenerated electrons. When the $\cdot\text{O}_2^-$ was captured, the balance of the reaction ($e^- + \text{O}_2 \rightarrow \cdot\text{O}_2^-$) moves to the right, reducing the photogenerated electrons in the reaction system. Therefore, during the photocatalytic $\text{Cr}(\text{vi})$ reduction, the electrons are the main participating species.

To further demonstrate the photocatalytic reaction mechanism of $\alpha\text{-MnO}_2/\text{MnIn}_2\text{S}_4$, the detection of superoxide radicals ($\cdot\text{O}_2^-$) and hydroxyl radicals ($\cdot\text{OH}$) can be carried out. The enhanced photocatalytic ability was confirmed by conversion experiments of nitro blue tetrazolium (NBT) and terephthalic acid (TA). Due to the exclusive reaction between $\cdot\text{O}_2^-$ and NBT, NBT can be used to detect the presence of $\cdot\text{O}_2^-$ in the photo-reaction.⁶¹ Therefore, in the photocatalytic reaction, a low NBT concentration means that a large amount of $\cdot\text{O}_2^-$ is produced in the system. It can be seen from Fig. 10a that, as the photocatalytic reaction time increases, the NBT concentration continues to decrease, confirming that the $\alpha\text{-MnO}_2/\text{MnIn}_2\text{S}_4$ (represented by M-MIS-2) system can produce $\cdot\text{O}_2^-$ during the photoreaction. The NBT measurement of all samples after 60 min is shown in Fig. 11b. Taking into account the effect of error, no change in the concentration of NBT was observed in the single $\alpha\text{-MnO}_2$ system, indicating that $\cdot\text{O}_2^-$ was not produced during the photoreaction. In addition, M-MIS-2 produced the highest level of $\cdot\text{O}_2^-$. The $\cdot\text{O}_2^-$ is produced by the reaction of O_2 with photogenerated electrons and more O_2 indicates high separation efficiency of photogenerated carriers in the $\alpha\text{-MnO}_2/\text{MnIn}_2\text{S}_4$ system.

At the same time, high $\cdot\text{O}_2^-$ production concentration are accompanied by high $\cdot\text{OH}$ concentrations.⁶² The $\cdot\text{OH}$ can be

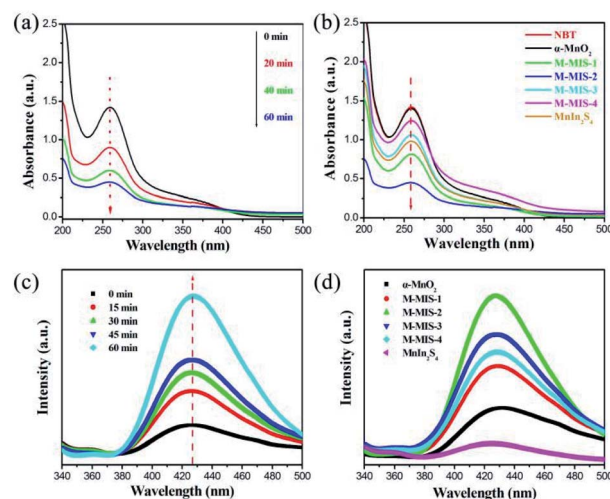


Fig. 10 Temporal UV-vis absorption spectral changes of NBT over (a) M-MIS-2 with different times and (b) all different photocatalysts (300 W Xe lamp, illumination time = 60 min, NBT dosage = 0.025 mM); PL spectral variations in 2 mM NaOH solution with the addition of 0.5 mM terephthalic acid using a (c) M-MIS-2 Z-scheme photocatalyst at different times and (d) single $\alpha\text{-MnO}_2$, MnIn_2S_4 and M-MIS- x ($x = 1, 2, 3, 4$).

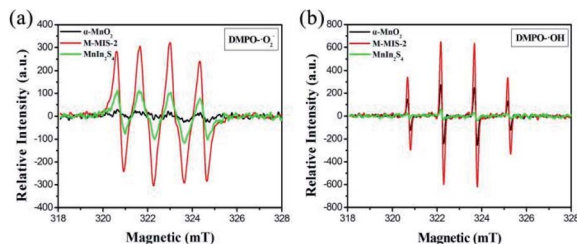


Fig. 11 ESR spectra of radical adducts trapped by DMPO: (a) DMPO- O_2^- and (b) DMPO- OH in the $\alpha\text{-MnO}_2$, MnIn_2S_4 and M-MIS-2 dispersion solution.

equally divided with terephthalic acid (TA) to form 2-hydroxyterephthalic acid (TAOH), which exhibits a characteristic peak at 425 nm when excited at a wavelength of 325 nm, so TA can be applied to detect OH . As shown in Fig. 10c, the gradual increase of the fluorescence intensity over time indicates that OH is gradually generated during the illumination process. In addition, the fluorescence signal of M-MIS- x is stronger than that of single $\alpha\text{-MnO}_2$ and MnIn_2S_4 after 60 min of irradiation (Fig. 10d). The M-MIS-2 has the highest signal intensity, which further proves that the composite material can produce higher active oxygen performance. This is consistent with the NBT detection result, which indicates that the composite catalyst improves the separation efficiency of carriers and produces more active components.

Furthermore, as shown in Fig. 11a, the ESR signal for DMPO- O_2^- can be observed for MnIn_2S_4 and M-MIS-2, which means that O_2^- was generated in the reaction system during visible light irradiation. Moreover, the characteristic peak of DMPO- OH for pure MnIn_2S_4 does not appear under visible light irradiation (Fig. 11b) but four weak peaks were detected for pure $\alpha\text{-MnO}_2$. As for M-MIS-2 heterojunction composites, the DMPO- OH signal intensities of OH are obviously much higher

than those in pure $\alpha\text{-MnO}_2$, which can be attributed to the construction of the atomically defined heterojunction to improve OH production.

Based on the above test results and experimental phenomenon, a possible Z-scheme photocatalytic reduction mechanism of 4-NA and $\text{Cr}(\text{vi})$ using $\alpha\text{-MnO}_2@ \text{MnIn}_2\text{S}_4$ is proposed. As illustrated in Fig. 12, under visible light irradiation, the electrons located on the valence band (VB) of $\alpha\text{-MnO}_2$ and MnIn_2S_4 are excited and then transfer to the conduction band (CB), while holes remain on the VB. At the interface of $\alpha\text{-MnO}_2$ and MnIn_2S_4 , electrons in the CB of $\alpha\text{-MnO}_2$ are transferred along the interface to the VB of MnIn_2S_4 to recombine with holes. The electrons remaining in the CB of MnIn_2S_4 gain relatively more negative potential energy and have stronger reducibility, and then react with 4-NA and $\text{Cr}(\text{vi})$, hence reducing them to 4-PDA and $\text{Cr}(\text{III})$. In contrast to the traditional type-II electron transport mechanism (Fig. S9†), the e^- in the CB of MnIn_2S_4 migrates to the CB of $\alpha\text{-MnO}_2$ and the h^+ of $\alpha\text{-MnO}_2$ is transferred to the VB of MnIn_2S_4 . Since the CB potential of $\alpha\text{-MnO}_2$ is more positive than that of O_2/O_2^- ($E = -0.33 \text{ eV}$)⁶³ and $\text{Cr}(\text{vi})/\text{Cr}(\text{II})$ ($E = 0.51 \text{ eV}$),⁶⁴ the VB potential of MnIn_2S_4 is more negative than that of $\text{H}_2\text{O}/\text{OH}$ ($E = 2.4 \text{ eV}$).⁶⁵ In this case, the catalyst cannot only reduce $\text{Cr}(\text{vi})$ to $\text{Cr}(\text{III})$, but also cannot generate O_2^- and OH in the reaction system, which is contrary to the experimental results (NBT and TA detection test). Therefore, the electron transmission in $\alpha\text{-MnO}_2@ \text{MnIn}_2\text{S}_4$ is a novel Z-scheme mechanism instead of the traditional type-II mechanism.

4 Conclusions

In summary, a core-shell $\alpha\text{-MnO}_2@ \text{MnIn}_2\text{S}_4$ hierarchical cubic Z-scheme heterostructure was designed and assembled by growing MnIn_2S_4 nanosheets *in situ* on the surface of $\alpha\text{-MnO}_2$ nanocubes as a heterojunction photocatalyst, which was used to efficiently reduce 4-NA and $\text{Cr}(\text{vi})$ under visible light. The composition and structure of the final heterojunction were controlled by facile annealing treatment and a hydrothermal reaction. The interface contact heterojunction defined at the atomic-level has proven the effective and rapid migration of photogenerated carriers on the catalytic surface, and the complex multilayer structure also exposes abundant catalytic sites, providing a large specific surface area for the adsorption of 4-NA and $\text{Cr}(\text{vi})$, and consequently, highly improving the photocatalytic reduction reaction. As a result, Z-scheme $\alpha\text{-MnO}_2@ \text{MnIn}_2\text{S}_4$ with the optimal composition ratio showed excellent reduction performance and cycle stability. The selective reduction of 4-NA to 4-PDA was almost complete (99.6%) within 90 minutes, and the $\text{Cr}(\text{vi})$ reduction efficiency reached more than 93% within 120 minutes. This work reveals the basis of the photocatalytic performance of $\alpha\text{-MnO}_2@ \text{MnIn}_2\text{S}_4$ composite materials and has potential application in organic photosynthesis and pollutant purification.

Credit author statement

Min Zhang: Conceptualization, methodology, software, investigation, and writing-original draft. Muhammad Arif and

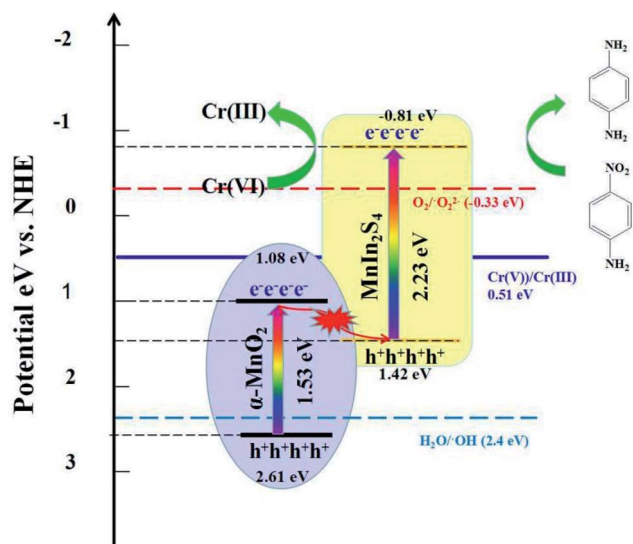


Fig. 12 Schematic illustration of the photocatalytic mechanism of $\alpha\text{-MnO}_2@ \text{MnIn}_2\text{S}_4$.



Yuxiang Hua: Writing – review & editing, supervision, and visualization. Bo Qiu: Resources, software, supervision, and data curation. Yue Mao: Resources and visualization. Xiaoheng Liu: Writing-review & editing and supervision. All the authors discussed the results and commented on the article.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the National Natural Science Foundation of China (51872141). The TEM part of this work was supported by Mr Chenyao Hu from Nanjing Forestry University.

References

- 1 D. A. Nicewicz and D. W. C. MacMillan, *Science*, 2008, **322**, 77–80.
- 2 H. Seo, M. H. Katcher and T. F. Jamison, *Nat. Chem.*, 2017, **9**, 453–456.
- 3 H. B. Yu, L. B. Jiang, H. Wang, B. B. Huang, X. Z. Yuan, J. H. Huang, J. Zhang and G. M. Zeng, *Small*, 2019, **15**, 1901008.
- 4 B. Weng, S. Q. Liu, N. Zhang, Z. R. Tang and Y. J. Xu, *J. Catal.*, 2014, **309**, 146–155.
- 5 R. W. Liang, F. F. Jing, G. Y. Yan and L. Wu, *Appl. Catal., B*, 2017, **218**, 452–459.
- 6 A. L. Chibac, V. Melinte, V. Brezová, E. A. Brosseau, V. Langlois and D. L. Versace, *ChemCatChem*, 2019, **11**, 3307–3317.
- 7 K. Mishra, T. N. Poudel, N. Basavegowda and T. R. Lee, *J. Catal.*, 2016, **344**, 273–285.
- 8 Z. H. Farooqi, R. Khalid, R. Begum, U. Farooq, Q. S. Wu, W. T. Wu, M. Ajamal, A. Irfan and K. Naseem, *Environ. Technol.*, 2019, **40**, 2027–2036.
- 9 M. A. Pandit, S. H. K. N. Dasari, S. Billakanti, M. Ramadoss and K. Muralidharan, *J. Phys. Chem. C*, 2020, **124**, 18010–18019.
- 10 T. Liu, Y. Sun, B. Jiang, W. Guo, W. Qin, Y. M. Xie, B. Zhao, L. Zhao, Z. Q. Liang and L. Jiang, *ACS Appl. Mater. Interfaces*, 2020, **12**, 28100–28109.
- 11 W. Chen, T. Y. Liu, T. Huang, X. H. Liu and X. J. Yang, *Nanoscale*, 2016, **8**, 3711–3719.
- 12 P. Viswanathan and R. Ramaraj, *J. Mol. Catal. A: Chem.*, 2016, **424**, 128–134.
- 13 M. Misra, S. R. Chowdhury and T. I. Lee, *Appl. Catal., B*, 2020, **272**, 118991.
- 14 L. Hou, Y. Q. Zhang, Y. Y. Ma, Y. L. Wang, Z. F. Hu, Y. Z. Gao and Z. G. Han, *Inorg. Chem.*, 2019, **58**, 16667–16675.
- 15 B. Jiang, Y. F. Gong, J. N. Gao, T. Sun, Y. J. Liu, N. Oturan and M. A. Oturan, *J. Hazard. Mater.*, 2019, **365**, 205–226.
- 16 J. Ding, L. T. Pu, Y. F. Wang, B. D. Wu, A. Q. Yu, X. L. Zhang, B. C. Pan, Q. X. Zhang and G. D. Gao, *Environ. Sci. Technol.*, 2018, **52**, 12602–12611.
- 17 H. L. Ma, Y. M. Zhang, Q. H. Hu, D. Yan, Z. Z. Yu and M. L. Zhai, *J. Mater. Chem.*, 2012, **22**, 5914–5916.
- 18 X. Zheng, D. Yuan, Y. Li and C. G. Liu, *Environ. Pollut.*, 2019, **254**, 113042.
- 19 C. P. Xu, P. R. Anusuyadevi, C. Aymonier, R. Luque and S. Marre, *Chem. Soc. Rev.*, 2019, **48**, 3868–3902.
- 20 L. Buzzetti, G. E. M. Crisenza and P. Melchiorre, *Angew. Chem., Int. Ed.*, 2019, **58**, 3730–3747.
- 21 A. Li, W. Zhu, C. C. Li, T. Wang and J. L. Gong, *Chem. Soc. Rev.*, 2019, **48**, 1874–1907.
- 22 R. Xiao, C. X. Zhao, Z. Y. Zou, Z. P. Chen, L. Tian, H. T. Xu, H. Tang, Q. Q. Liu, Z. X. Lin and X. F. Yang, *Appl. Catal., B*, 2020, **268**, 118382.
- 23 H. H. Zeng, B. L. Xing, C. T. Zhang, L. J. Chen, H. H. Zhao, X. F. Han, G. Y. Yi and G. X. Huang, *Energy Fuels*, 2020, **34**, 2480–2491.
- 24 S. I. Han, S. W. Lee, M. G. Cho, J. M. Yoo, M. H. Oh, B. Jeong, D. Kim, O. K. Park, J. C. Kim, E. Namkoong, J. W. Jo, N. Lee, C. H. Lim, M. Soh, Y. E. Sung, J. M. Yoo, K. Park and T. W. Hyeon, *Adv. Mater.*, 2020, **32**, 2001566.
- 25 X. Sun, G. L. Zhang, R. H. Du, R. Xu, D. W. Zhu, J. C. Qian, G. Bai, C. Yang, Z. Y. Zhang, X. Zhang, D. H. Zou and Z. Y. Wu, *Biomaterials*, 2019, **194**, 151–160.
- 26 P. Miao, J. X. Chen, Y. S. Tang, K. J. Chen and J. Kong, *Sci. China Mater.*, 2020, 1–12.
- 27 D. L. Chao, C. Ye, F. X. Xie, W. H. Zhou, Q. H. Zhang, Q. F. Gu, K. Davey, L. Gu and S. Z. Qiao, *Adv. Mater.*, 2020, **32**, 2001894.
- 28 F. W. Boyom-Tatchemo, F. Devred, G. Ndiiffo-Yemeli, S. Laminsi and E. M. Gaigneaux, *Appl. Catal., B*, 2020, **260**, 118159.
- 29 Y. G. Zhang, M. Y. Wu, Y. H. Kwok, Y. F. Wang, W. Zhao, X. L. Zhao, H. B. Huang and D. Y. C. Leung, *Appl. Catal., B*, 2019, **259**, 118034.
- 30 P. Wu, S. Q. Dai, G. X. Chen, S. Q. Zhao, M. L. Fu, P. R. Chen, Q. Chen, X. J. Jin, Y. C. Qiu, S. H. Yang and D. Q. Ye, *Appl. Catal., B*, 2020, **268**, 118418.
- 31 S. B. Wang, Y. Wang, S. L. Zhang, S. Q. Zang and X. W. Lou, *Adv. Mater.*, 2019, **31**, 1903404.
- 32 C. Y. Liu, X. Li, J. Z. Li, Y. J. Zhou, L. L. Sun, H. Q. Wang, P. W. Huo, C. C. Ma and Y. S. Yan, *Carbon*, 2019, **152**, 565–574.
- 33 D. Santamaria-Perez, M. Amboage, F. J. Manjon, D. Errandonea, A. Muñoz, P. Rodríguez-Hernández, A. Mújica, S. Radescu, V. V. Ursaki and I. M. Tiginyanu, *J. Phys. Chem. C*, 2012, **116**, 14078–14087.
- 34 W. Chen, Z. C. He, G. B. Huang, C. L. Wu, W. F. Chen and X. H. Liu, *Chem. Eng. J.*, 2019, **359**, 244–253.
- 35 X. Q. Chen, Z. G. Cao, L. D. Xing, Y. H. Liao, Y. C. Qiu and W. S. Li, *Nanoscale*, 2017, **9**, 18467–18473.
- 36 H. Zhang, Q. Qi, P. G. Zhang, W. Zheng, J. Chen, A. G. Zhou, W. B. Tian, W. Zhang and Z. M. Sun, *ACS Appl. Energy Mater.*, 2018, **2**, 705–714.
- 37 H. Fang and P. Jena, *J. Mater. Chem. A*, 2016, **4**, 4728–4737.
- 38 W. Wang and L. Ao, *Cryst. Growth Des.*, 2008, **8**, 358–362.
- 39 B. K. Wu, G. B. Zhang, M. Y. Yan, T. F. Xiong, P. He, L. He, X. Xu and L. Q. Mai, *Small*, 2018, **14**, 1703850.



- 40 Y. Shiraishi, J. Imai, N. Yasumoto, H. Sakamoto, S. Tanaka, S. Ichikawa and T. Hira, *Langmuir*, 2019, **35**, 5455–5462.
- 41 C. K. Yao, A. L. Yuan, Z. S. Wang, H. Lei, L. Zhang, L. M. Guo and X. P. Dong, *J. Mater. Chem. A*, 2019, **7**, 13071–13079.
- 42 R. Bi, G. X. Liu, C. Zeng, X. P. Wang, L. Zhang and S. Z. Qiao, *Small*, 2019, **15**, 1804958.
- 43 B. X. Li, G. X. Rong, Y. Xie, L. F. Huang and C. Q. Feng, *Inorg. Chem.*, 2006, **45**, 6404–6410.
- 44 H. Chen, M. Q. Wang, Y. N. Yu, H. Liu, S. Y. Lu, S. J. Bao and M. W. Xu, *ACS Appl. Mater. Interfaces*, 2017, **9**, 35040–35047.
- 45 C. Revathi and R. T. R. Kumar, *Electroanalysis*, 2017, **29**, 1481–1489.
- 46 L. B. Ni, Z. Wu, G. L. Zhao, C. Y. Sun, C. Q. Zhou, X. X. Gong and G. W. Diao, *Small*, 2017, **13**, 1603466.
- 47 D. L. Chao, W. H. Zhou, C. Ye, Q. H. Zhang, Y. G. Cheng, L. Gu, K. Davey and S. Z. Qiao, *Angew. Chem., Int. Ed.*, 2019, **58**, 7823–7828.
- 48 X. Shi, H. Zheng, A. M. Kannan, P. Salcedo and B. Escobar, *Inorg. Chem.*, 2019, **58**(8), 5335–5344.
- 49 M. Zhang, J. C. Yao, M. Arif, B. Qiu, H. F. Yin, X. H. Liu and S. M. Chen, *Appl. Surf. Sci.*, 2020, **526**, 145749.
- 50 Y. Y. Shi, S. M. Gao, Y. F. Yuan, G. J. Liu, R. C. Jin, Q. Y. Wang, H. Xu and J. Lu, *Nano Energy*, 2020, **77**, 105153.
- 51 Y. Chen, X. Zhang, C. Xu and H. Xu, *Electrochim. Acta*, 2019, **309**, 424–431.
- 52 R. Sha, N. Vishnu and S. Badhulika, *Sens. Actuators, B*, 2019, **279**, 53–60.
- 53 C. X. Zhao, L. Tian, Z. Y. Zou, Z. P. Chen, H. Tang, Q. Q. Liu, Z. X. Lin and X. F. Yang, *Appl. Catal., B*, 2020, **268**, 118445.
- 54 Y. C. Zhang, N. Afzal, L. Pan, X. W. Zhang and J. J. Zou, *Adv. Sci.*, 2019, **6**, 1900053.
- 55 J. Zhou, J. X. Cui, M. Dong, C. Y. Sun, S. Q. You, X. L. Wang, Z. Y. Zhou and Z. M. Su, *Chem. Commun.*, 2020, **56**, 7261–7264.
- 56 S. M. Ghoreishian, K. S. Ranjith, H. Lee, H. I. Ju, S. Z. Nikoo, Y. K. Han and Y. S. Huh, *J. Hazard. Mater.*, 2020, **391**, 122249.
- 57 L. H. Zheng, H. R. Su, J. Z. Zhang, L. S. Walekar, B. X. Zhou, M. C. Long and Y. H. Hu, *Appl. Catal., B*, 2018, **239**, 475–484.
- 58 M. M. Fang, J. X. Shao, X. G. Huang, J. Y. Wang and W. Chen, *J. Mater. Sci. Technol.*, 2020, **56**, 133–142.
- 59 P. J. Li, W. Cao, Y. Zhu, Q. Y. Teng, L. Peng, C. Y. Jiang, C. S. Feng and Y. P. Wang, *Sci. Total Environ.*, 2020, **715**, 136809.
- 60 H. F. Yin, Y. Cao, T. L. Fan, M. Zhang, J. C. Yao, P. F. Li, S. M. Cheng and X. H. Liu, *Sci. Total Environ.*, 2020, **754**, 141926.
- 61 Y. C. Deng, J. Liu, Y. B. Huang, M. M. Ma, K. Liu, X. M. Dou, Z. J. Wang, S. C. Qu and Z. G. Wang, *Adv. Funct. Mater.*, 2020, **30**, 2002353.
- 62 M. Li, H. W. Huang, S. X. N. Tian and Y. H. Zhang, *ChemCatChem*, 2018, **10**, 4477–4496.
- 63 T. Huang, Y. T. Luo, W. Chen, J. C. Yao and X. H. Liu, *ACS Sustainable Chem. Eng.*, 2018, **6**, 4671–4679.
- 64 C. Cheng, D. Y. Chen, N. J. Li, Q. F. Xu, H. Li, J. H. He and J. M. Lu, *J. Hazard. Mater.*, 2020, **391**, 122205.
- 65 W. Chen, L. Chang, S. B. Ren, Z. C. He, G. B. Huang and X. H. Liu, *J. Hazard. Mater.*, 2020, **384**, 121308.

